Acta Crystallographica Section E

Structure Reports Online

ISSN 1600-5368

Methyl 6-amino-6-oxohexanoate

Tobias Gruber,[a]‡ Christopher J. Schofield[a] and Amber L. Thompson[b]*

[a]Chemistry Research Laboratory, Department of Chemistry, University of Oxford, Mansfield Road, Oxford OX1 3TA, England, and [b]Chemical Crystallography, Chemistry Research Laboratory, Department of Chemistry, University of Oxford, Mansfield Road, Oxford OX1 3TA, England

Correspondence e-mail: amber.thompson@chem.ox.ac.uk

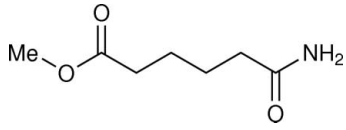
Received 18 November 2011; accepted 25 January 2012

Key indicators: single-crystal X-ray study; T = 150 K; mean σ(C–C) = 0.002 Å; R factor = 0.044; wR factor = 0.128; data-to-parameter ratio = 17.7.

The title compound, $C_7H_{13}NO_3$, adopts an approximately planar conformation. The torsion angles in the aliphatic chain between the carbonyl group C atoms range from 172.97 (14) to 179.38 (14)° and the r.m.s. deviation of all non-H atoms is 0.059 Å. The crystal packing is dominated by two strong N—H···O hydrogen bonds involving the amide groups and forming $R_2^2(8)$ rings and $C(4)$ chains. Overall, a two-dimensional network parallel to (100) is formed. A weak intermolecular C—H···O interaction is also present.

Related literature

For the synthesis of the title compound, see: Kulikova *et al.* (1960); Nishitani *et al.* (1982); Micovic *et al.* (1988). For information on the solid-state characteristics of different polymorphs of adipic acid, see: Fun & Chantrapromma (2009); Ranganathan *et al.* (2003); Srinivasa Gopalan *et al.* (1999, 2000); Pfefer & Boistelle (2000); Housty & Hospital (1965); Arevalo & Canut (1961); Hirokawa (1950); Morrison & Robertson (1949); MacGillavry (1941). For details on co-crystals of the title compound, see: Goswami *et al.* (2010); Delori *et al.* (2008); Bucar *et al.* (2007); Childs & Hardcastle (2007); Duan *et al.* (2005); Li *et al.* (2001); Urbanczyk-Lipkowska & Gluzinski (1996). For other reports of adipic acid derivatives, see: Li & Goddard (2002); Seaton & Tremayne (2002); Hospital & Housty (1966). For uses of the title compound in heterocycle synthesis, see: Jungheim *et al.* (2005); Fukumoto *et al.* (2007). For hydrogen-bond motifs, see: Bernstein *et al.* (1995). For details of the H-atom treatment, see: Cooper *et al.* (2010).



Experimental

Crystal data

$C_7H_{13}NO_3$
M_r = 159.19
Monoclinic, $P2_1/c$
a = 12.896 (3) Å
b = 7.2143 (8) Å
c = 9.6324 (12) Å
β = 106.474 (17)°
V = 859.4 (2) Å^3
Z = 4
Cu $K\alpha$ radiation
μ = 0.80 mm^{-1}
T = 150 K
0.18 × 0.12 × 0.02 mm

Data collection

Agilent SuperNova Dual (Cu at zero) diffractometer with an Atlas detector
Absorption correction: multi-scan (*CrysAlis PRO*; Agilent, 2010)
T_{min} = 0.48, T_{max} = 0.98
7240 measured reflections
1771 independent reflections
1426 reflections with $I > 2\sigma(I)$
R_{int} = 0.035
Standard reflections: 0

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.044
$wR(F^2)$ = 0.128
S = 1.00
1770 reflections
100 parameters
H-atom parameters constrained
$\Delta\rho_{max}$ = 0.24 e Å^{-3}
$\Delta\rho_{min}$ = −0.24 e Å^{-3}

Table 1
Hydrogen-bond geometry (Å, °).

D—H···A	D—H	H···A	D···A	D—H···A
N3—H31···O1^{i}	0.86	2.07	2.929 (2)	173 (1)
N3—H32···O1^{ii}	0.86	2.09	2.922 (2)	162 (1)
C10—H101···O11^{iii}	0.95	2.61	3.486 (3)	153 (1)

Symmetry codes: (i) $-x, -y+3, -z+1$; (ii) $x, -y+\frac{5}{2}, z+\frac{1}{2}$; (iii) $x, -y+\frac{1}{2}, z-\frac{1}{2}$.

Data collection: *CrysAlis PRO* (Agilent, 2010); cell refinement: *CrysAlis PRO*; data reduction: *CrysAlis PRO*; program(s) used to solve structure: *SIR92* (Altomare *et al.*, 1994); program(s) used to refine structure: *CRYSTALS* (Betteridge *et al.*, 2003); molecular graphics: *CAMERON* (Watkin *et al.*, 1996); software used to prepare material for publication: *CRYSTALS* and *PLATON* (Spek, 2009).

TG thanks Deutsche Forschungsgemeinschaft (DFG), Germany, for generous funding (GR 3693/1–1:1).

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: LH5383).

References

Agilent (2010). *CrysAlis PRO*. Agilent Technologies, Yarnton, Oxfordshire, England.
Altomare, A., Cascarano, G., Giacovazzo, C., Guagliardi, A., Burla, M. C., Polidori, G. & Camalli, M. (1994). *J. Appl. Cryst.* **27**, 435.
Arevalo, M. I. & Canut, M. L. (1961). *Bol. R. Soc. Esp. Hist. Nat. (Sec. Geol.)*, **59**, 37–40.
Bernstein, J., Davis, R. E., Shimoni, L. & Chang, N.-L. (1995). *Angew. Chem. Int. Ed. Engl.* **34**, 1555–1573.

‡ Present address: Institut für Organische Chemie, TU Braunschweig, Hagenring 30, D-38106 Braunschweig, Germany.

Betteridge, P. W., Carruthers, J. R., Cooper, R. I., Prout, K. & Watkin, D. J. (2003). *J. Appl. Cryst.* **36**, 1487.
Bucar, D.-K., Henry, R. F., Lou, X., Borchardt, T. B. & Zhang, G. G. Z. (2007). *Chem. Commun.* pp. 525–527.
Childs, S. L. & Hardcastle, K. I. (2007). *Cryst. Growth Des.* **7**, 1291–1304.
Cooper, R. I., Thompson, A. L. & Watkin, D. J. (2010). *J. Appl. Cryst.* **43**, 1100–1107.
Delori, A., Suresh, E. & Pedireddi, V. R. (2008). *Chem. Eur. J.* **14**, 6967–6977.
Duan, L.-M., Ye, L., Liu, Y.-B., Xie, F.-T., Yu, J.-H. & Xu, J.-Q. (2005). *Pol. J. Chem.* **79**, 1835–1842.
Fukumoto, S., Matsunaga, N., Ohra, T., Ohyabu, N., Hasui, T., Motoyaji, T., Siedem, C. S., Tang, T. P., Demeese, L. A. & Gauthier, C. (2007). PCT Int. Appl. 2007077961.
Fun, H.-K. & Chantrapromma, S. (2009). *Acta Cryst.* E**65**, o624.
Goswami, S., Hazra, A. & Fun, H.-K. (2010). *J. Inclusion Phenom.* **68**, 461–466.
Hirokawa, S. (1950). *Bull. Chem. Soc. Jpn*, **23**, 91–94.
Hospital, M. & Housty, J. (1966). *Acta Cryst.* **20**, 626–630.
Housty, J. & Hospital, M. (1965). *Acta Cryst.* **18**, 693–697.
Jungheim, L. N., McGill, J. M., Trasher, K. J., Herr, R. J. & Muralikrishna, V. (2005). PCT Int. Appl. 2005019184.
Kulikova, A. E., Zil'berman, E. N. & Sazanova, N. A. (1960). *Zh. Obshch. Khim.* **30**, 2180–2183.
Li, Y. & Goddard, W. A. (2002). *Macromolecules*, **35**, 8440–8455.
Li, W., Zhang, J.-P., Tong, M.-L. & Chen, X.-M. (2001). *Aust. J. Chem.* **54**, 213–217.
MacGillavry, C. H. (1941). *Recl Trav. Chim. Pays-Bas*, **60**, 605–617.
Micovic, V., Ivanovic, M. D. & Piatak, D. M. (1988). *J. Serb. Chem. Soc.* **53**, 419–426.
Morrison, J. D. & Robertson, J. M. (1949). *J. Chem. Soc.* pp. 987–992.
Nishitani, T., Horikawa, H., Iwasaki, T., Matsumoto, K., Inoue, I. & Miyoshi, M. (1982). *J. Org. Chem.* **47**, 1706–1712.
Pfefer, G. & Boistelle, R. (2000). *J. Cryst. Growth*, **208**, 615–622.
Ranganathan, A., Kulkarni, G. U. & Rao, C. N. R. (2003). *J. Phys. Chem. A*, **107**, 6073–6081.
Seaton, C. C. & Tremayne, M. (2002). *Chem. Commun.* pp. 880–881.
Spek, A. L. (2009). *Acta Cryst.* D**65**, 148–155.
Srinivasa Gopalan, R., Kumaradhas, P. & Kulkarni, G. U. (1999). *J. Solid State Chem.* **148**, 129–134.
Srinivasa Gopalan, R., Kumaradhas, P., Kulkarni, G. U. & Rao, C. N. R. (2000). *J. Mol. Struct.* **521**, 97–106.
Urbanczyk-Lipkowska, Z. & Gluzinski, P. (1996). *Supramol. Chem.* **7**, 113–118.
Watkin, D. J., Prout, C. K. & Pearce, L. J. (1996). *CAMERON*. Chemical Crystallography Laboratory, Oxford, England.

supplementary materials

Acta Cryst. (2012). E**68**, o593–o594 [doi:10.1107/S1600536812003303]

Methyl 6-amino-6-oxohexanoate

Tobias Gruber, Christopher J. Schofield and Amber L. Thompson

Comment

Adipic acid has importance in various industrial applications including the production of polyamides and polyurethanes. The solid state characteristics of different polymorphs of adipic acid have already been investigated intensively (Fun & Chantrapromma, 2009; Ranganathan *et al.*, 2003; Srinivasa Gopalan *et al.*, 2000; Pfefer & Boistelle, 2000; Srinivasa Gopalan *et al.*, 1999; Housty & Hospital, 1965; Arevalo & Canut, 1961; Hirokawa, 1950; Morrison & Robertson, 1949; MacGillavry, 1941), as have adipic acid co-crystals (Goswami *et al.*, 2010; Delori *et al.*, 2008; Bucar *et al.*, 2007; Childs & Hardcastle, 2007; Duan *et al.*, 2005, Li *et al.*, 2001; Urbanczyk-Lipkowska & Gluzinski, 1996). Reports on single-crystal X-ray structures of adipic acid derivatives have focused on the important nylon-based materials (Li & Goddard, 2002). Here we describe the structure of a simple adipic acid derivative, *viz.* methyl 6-amino-6-oxohexanoate (**I**), an approved starting material for heterocyclic synthesis (Jungheim *et al.*, 2005; Fukumoto *et al.*, 2007).

Methyl 6-amino-6-oxohexanoate (**I**) crystallizes from methanol as colourless crystals in the monoclinic space group $P2_1/c$ (Fig. 1). The molecule is approximately planar; the largest deviation from the mean plane defined by the non-hydrogen atoms is 0.116 Å for carbonyl oxygen O1 and the aliphatic chain between the carbonyl carbons is only slightly twisted with torsion angles ranging from 172.97 (14) to 179.38 (14)°. The crystal packing is dominated by two strong N—H···O hydrogen bonds (see Table 1), similar to those seen in the two polymorphs of adipamide (monoclinic: Hospital & Housty, 1966; triclinic: Seaton & Tremayne, 2002). In (**I**), the the amide nitrogen in serves as a double intermolecular hydrogen donor: N3—H31···O1^{i} forms an $R_2^2(8)$ amide dimer around an inversion centre, while N3—H32···O1^{ii} connects pairs of dimers to form $C(4)$ chains parallel to the *c* axis. The combination of the C(4) and $R_2^2(8)$ motifs generates a secondary network of $R_{10}^6(24)$ as described for related compounds including benzamide *etc*. (Bernstein *et al.* (1995); Fig. 2).

Notably, the methyl ester carbonyl group is not involved in hydrogen bonding, however, it is in a suitable position to engage in a weak C—H···O intermolecular interaction with an ester methyl group [d(H···O) = 2.614 (3) Å].

In conclusion, the structure of (**I**), together with those similar and previously reported, suggest that the variation in the carbonyl substituent at adipic acid does not cause substantial changes to the conformation of the molecule.

Experimental

The title compound was recovered as a side product in 0.5% yield from the cyclization reaction of amino pimelic acid methylester in *p*-cymene *via* a redox process (Nishitani *et al.*, 1982). Crystals suitable for X-ray diffraction were obtained by slow evaporation of a solution of the compound in methanol.

Alternatively, the title compound can also preprared by reaction of the respective acid chloride with ammonia (Micovic *et al.*, 1988) and the partial hydrolysis of the corresponding nitrile (Kulikova *et al.*, 1960).

Refinement

The structure was refined by full-matrix least-squares. H atoms were treated in the usual manner: positioned geometrically (aliphatic) or located in the difference map (amide) and refined prior to inclusion in the model using riding constraints (Cooper *et al.*, 2010).

Dihedral angles were calculated with *PLATON* (Spek, 2009); all other standard uncertainties calculated from the full variance co-variance matrix within *CRYSTALS* (Betteridge *et al.*, 2003).

Computing details

Data collection: *CrysAlis PRO* (Agilent, 2010); cell refinement: *CrysAlis PRO* (Agilent, 2010); data reduction: *CrysAlis PRO* (Agilent, 2010); program(s) used to solve structure: *SIR92* (Altomare *et al.*, 1994); program(s) used to refine structure: *CRYSTALS* (Betteridge *et al.*, 2003); molecular graphics: *CAMERON* (Watkin *et al.*, 1996); software used to prepare material for publication: *CRYSTALS* (Betteridge *et al.*, 2003) and *PLATON* (Spek, 2009).

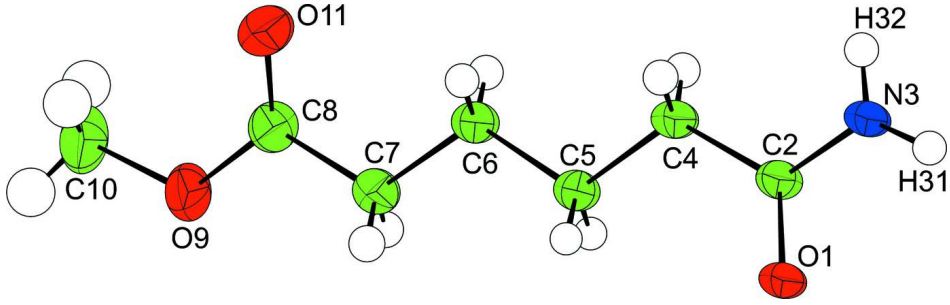


Figure 1

Molecular structure of (**I**) with displacement ellipsoids drawn at 50% probability.

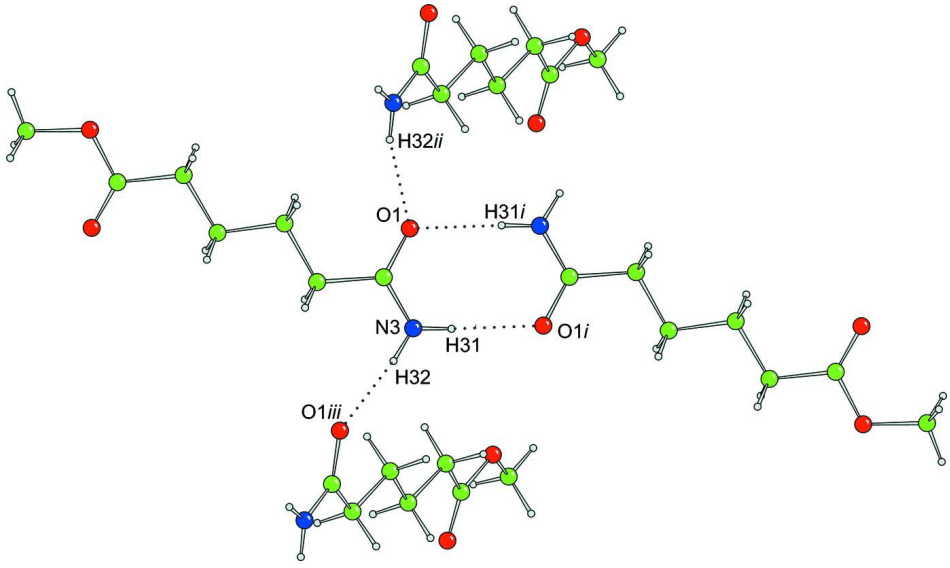


Figure 2

Hydrogen bonding in the crystal structure of (**I**) [*i*: -*x*,3 - *y*,1 - *z*; *ii*: *x*,5/2 - *y*,-1/2 + *z*; *iii*: *x*,5/2 - *y*,1/2 + *z*].

Methyl 6-amino-6-oxohexanoate

Crystal data

$C_7H_{13}NO_3$
$M_r = 159.19$
Monoclinic, $P2_1/c$
Hall symbol: -P 2ybc
$a = 12.896$ (3) Å
$b = 7.2143$ (8) Å
$c = 9.6324$ (12) Å
$\beta = 106.474$ (17)°
$V = 859.4$ (2) Å^3
$Z = 4$
$F(000) = 344$
$D_x = 1.230$ Mg m^{-3}
Melting point: not measured K
Cu $K\alpha$ radiation, $\lambda = 1.54184$ Å
Cell parameters from 2081 reflections
$\theta = 4$–$76°$
$\mu = 0.80$ mm^{-1}
$T = 150$ K
Lath, clear_pale_colourless
0.18 × 0.12 × 0.02 mm

Data collection

Agilent SuperNova Dual (Cu at zero) diffractometer with an Atlas detector
Graphite monochromator
ω scans
Absorption correction: multi-scan (*CrysAlis PRO*; Agilent, 2010)
$T_{min} = 0.48$, $T_{max} = 0.98$
7240 measured reflections
1771 independent reflections
1426 reflections with $I > 2\sigma(I)$
$R_{int} = 0.035$
$\theta_{max} = 76.0°$, $\theta_{min} = 3.6°$
$h = -16 \rightarrow 15$
$k = -9 \rightarrow 7$
$l = -12 \rightarrow 11$

Refinement

Refinement on F^2
Least-squares matrix: full
$R[F^2 > 2\sigma(F^2)] = 0.044$
$wR(F^2) = 0.128$
$S = 1.00$
1770 reflections
100 parameters
0 restraints
Primary atom site location: structure-invariant direct methods
Hydrogen site location: inferred from neighbouring sites
H-atom parameters constrained
Method = Modified Sheldrick $w = 1/[\sigma^2(F^2) + (0.07P)^2 + 0.22P]$, where $P = (\max(F_o^2,0) + 2F_c^2)/3$
$(\Delta/\sigma)_{max} = 0.001$
$\Delta\rho_{max} = 0.24$ e Å^{-3}
$\Delta\rho_{min} = -0.24$ e Å^{-3}

Special details

Experimental. Agilent Technologies (2010). CrysAlisPro. Version 1.171.35.4 (release 09-12-2010 CrysAlis171 .NET) (compiled Dec 9 2010,10:47:41) Empirical absorption correction using spherical harmonics, implemented in SCALE3 ABSPACK scaling algorithm.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	*y*	*z*	U_{iso}*/U_{eq}
O1	0.07129 (11)	1.29877 (15)	0.44865 (11)	0.0413
C2	0.09105 (12)	1.2640 (2)	0.58008 (14)	0.0323
N3	0.05751 (12)	1.37331 (17)	0.66941 (12)	0.0372
H31	0.0242	1.4758	0.6389	0.0445*
H32	0.0717	1.3426	0.7589	0.0449*
C4	0.15519 (14)	1.0967 (2)	0.64774 (15)	0.0380
C5	0.18119 (13)	0.9615 (2)	0.54141 (15)	0.0339
C6	0.24703 (14)	0.7987 (2)	0.62007 (15)	0.0373
C7	0.27433 (15)	0.6613 (2)	0.51664 (17)	0.0417

C8	0.34505 (14)	0.5046 (2)	0.59123 (17)	0.0411
O9	0.36947 (12)	0.39151 (19)	0.49492 (14)	0.0571
C10	0.43990 (18)	0.2384 (3)	0.5546 (2)	0.0604
H101	0.4474	0.1631	0.4773	0.0899*
H103	0.5097	0.2837	0.6123	0.0882*
H102	0.4069	0.1646	0.6164	0.0903*
O11	0.37696 (14)	0.48159 (19)	0.71881 (14)	0.0605
H71	0.3122	0.7249	0.4570	0.0507*
H72	0.2080	0.6062	0.4549	0.0501*
H62	0.3138	0.8436	0.6854	0.0443*
H61	0.2066	0.7354	0.6763	0.0458*
H52	0.2233	1.0258	0.4878	0.0396*
H51	0.1133	0.9163	0.4743	0.0412*
H41	0.2227	1.1408	0.7133	0.0470*
H42	0.1152	1.0285	0.7024	0.0471*

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
O1	0.0675 (8)	0.0363 (6)	0.0230 (5)	0.0117 (5)	0.0176 (5)	0.0041 (4)
C2	0.0438 (8)	0.0298 (7)	0.0251 (6)	−0.0007 (6)	0.0126 (6)	0.0004 (5)
N3	0.0586 (8)	0.0333 (6)	0.0223 (5)	0.0073 (6)	0.0154 (5)	0.0023 (4)
C4	0.0549 (9)	0.0355 (8)	0.0247 (6)	0.0069 (7)	0.0133 (6)	0.0033 (5)
C5	0.0441 (8)	0.0312 (7)	0.0269 (6)	0.0011 (6)	0.0109 (5)	−0.0001 (5)
C6	0.0521 (9)	0.0317 (7)	0.0281 (7)	0.0030 (6)	0.0112 (6)	0.0006 (5)
C7	0.0532 (9)	0.0378 (8)	0.0327 (7)	0.0066 (7)	0.0100 (7)	−0.0031 (6)
C8	0.0461 (9)	0.0330 (8)	0.0408 (8)	−0.0015 (6)	0.0070 (7)	−0.0036 (6)
O9	0.0645 (8)	0.0499 (7)	0.0489 (7)	0.0206 (6)	0.0031 (6)	−0.0128 (5)
C10	0.0552 (11)	0.0471 (10)	0.0707 (13)	0.0130 (9)	0.0046 (9)	−0.0113 (9)
O11	0.0857 (10)	0.0492 (7)	0.0424 (7)	0.0185 (7)	0.0111 (7)	0.0068 (5)

Geometric parameters (Å, º)

O1—C2	1.2441 (18)	C6—H62	0.967
C2—N3	1.3269 (19)	C6—H61	0.966
C2—C4	1.503 (2)	C7—C8	1.501 (2)
N3—H31	0.864	C7—H71	0.969
N3—H32	0.858	C7—H72	0.977
C4—C5	1.5193 (19)	C8—O9	1.338 (2)
C4—H41	0.972	C8—O11	1.192 (2)
C4—H42	0.970	O9—C10	1.442 (2)
C5—C6	1.520 (2)	C10—H101	0.949
C5—H52	0.967	C10—H103	0.971
C5—H51	0.984	C10—H102	0.981
C6—C7	1.516 (2)		
O1—C2—N3	121.98 (13)	C7—C6—H62	108.5
O1—C2—C4	122.14 (13)	C5—C6—H61	109.3
N3—C2—C4	115.88 (12)	C7—C6—H61	108.8
C2—N3—H31	120.7	H62—C6—H61	108.4

C2—N3—H32	118.8	C6—C7—C8	113.61 (13)
H31—N3—H32	120.4	C6—C7—H71	109.3
C2—C4—C5	115.01 (11)	C8—C7—H71	107.5
C2—C4—H41	107.5	C6—C7—H72	109.9
C5—C4—H41	108.7	C8—C7—H72	107.0
C2—C4—H42	109.3	H71—C7—H72	109.5
C5—C4—H42	107.1	C7—C8—O9	110.96 (14)
H41—C4—H42	109.2	C7—C8—O11	125.70 (15)
C4—C5—C6	111.05 (12)	O9—C8—O11	123.34 (16)
C4—C5—H52	108.5	C8—O9—C10	115.85 (15)
C6—C5—H52	108.5	O9—C10—H101	108.6
C4—C5—H51	109.3	O9—C10—H103	110.4
C6—C5—H51	109.8	H101—C10—H103	110.9
H52—C5—H51	109.7	O9—C10—H102	109.0
C5—C6—C7	112.25 (12)	H101—C10—H102	108.9
C5—C6—H62	109.5	H103—C10—H102	109.1

Hydrogen-bond geometry (Å, º)

D—H···*A*	*D*—H	H···*A*	*D*···*A*	*D*—H···*A*
N3—H31···O1^{i}	0.86	2.07	2.929 (2)	173 (1)
N3—H32···O1^{ii}	0.86	2.09	2.922 (2)	162 (1)
C10—H101···O11^{iii}	0.95	2.61	3.486 (3)	153 (1)

Symmetry codes: (i) −*x*, −*y*+3, −*z*+1; (ii) *x*, −*y*+5/2, *z*+1/2; (iii) *x*, −*y*+1/2, *z*−1/2.