

Surface chemistry of alanine on Ni{111}

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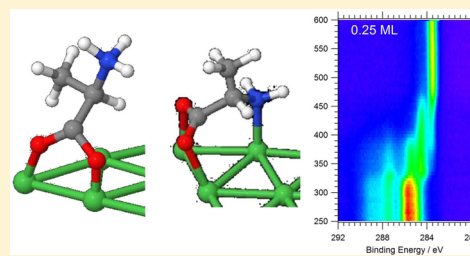
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1 **Surface Chemistry of Alanine on Ni{111}**2 Richard E. J. Nicklin,[†] Alix Cornish,[†] Andrey Shavorskiy,[‡] Silvia Baldanza,[†] Karina Schulte,[§] Zhi Liu,[‡]
3 Roger A. Bennett,[†] and Georg Held^{*,†,||}4 [†]Department of Chemistry, University of Reading, Reading RG6 6AD, U.K.5 [‡]Advanced Light Source, Lawrence Berkeley National Laboratory, Berkeley, California 94720, United States6 [§]MAX-lab, Lund University, 22100 Lund, Sweden7 ^{||}Diamond Light Source Ltd, Oxfordshire OX11 0DE, U.K.8 **S** Supporting Information

9 **ABSTRACT:** The adsorption of L-alanine on Ni{111} has been studied as a
10 model of enantioselective heterogeneous catalysts. Synchrotron-based X-ray
11 photoelectron spectroscopy and near-edge X-ray absorption fine structure
12 (NEXAFS) spectroscopy were used to determine the chemical state, bond
13 coordination, and out-of-plane orientation of the molecule on the surface.
14 Alanine adsorbs in anionic and zwitterionic forms between 250 and ≈ 320 K.
15 NEXAFS spectra exhibit a strong angular dependence of the π^* resonance
16 associated with the carboxylate group, which is compatible with two distinct
17 orientations with respect to the surface corresponding to the bidentate and
18 tridentate binding modes. Desorption and decomposition begin together at
19 ≈ 300 K, with decomposition occurring in a multistep process up to ≈ 450 K. Comparison with previous studies of amino acid
20 adsorption on metal surfaces shows that this is among the lowest decomposition temperatures found so far and lower than typical
21 temperatures used for hydrogenation reactions where modified Ni catalysts are used.

22 **INTRODUCTION**

23 Amino acid adsorption on metals is a burgeoning field of research
24 due to its relevance to chemical sensing, biomedical research, and
25 enantioselective catalysis. Heterogeneous enantioselective catal-
26 ysis using alanine as a chiral modifier for Ni has been known since
27 1962.¹ The best studied reaction in this context is the
28 enantioselective hydrogenation of methyl acetoacetate,^{2–4} with
29 surface sensitivity results pointing to the involvement of the
30 low-Miller-index planes.^{5,6}

31 The orientation and bonding of amino acids on transition-
32 metal surfaces has been probed by a variety of techniques,
33 notably low-energy electron diffraction (LEED), temperature-
34 programmed desorption (TPD), near-edge X-ray absorption
35 fine structure (NEXAFS) spectroscopy, reflection absorption
36 IR spectroscopy (RAIRS), high-resolution electron energy loss
37 spectroscopy (HREELS), photoelectron diffraction (PhD),
38 scanning tunneling microscopy (STM), and X-ray photoelectron
39 spectroscopy (XPS).^{7–47} LEED, NEXAFS, STM, XPS, and PhD
40 have been used to investigate the adsorption of glycine and
41 alanine on Cu surfaces. It was demonstrated that the adsorbate
42 binds as the anion, adopting a tridentate binding mode on the
43 surface, with single Cu atoms bonding to both O atoms
44 and the N atom.^{11,19,21,28,46,47} This three-point binding has
45 received theoretical support.^{28,48–51} It breaks the surface
46 symmetry and imposes a chiral environment on the surface, a
47 prerequisite for chiral catalysis.^{7,52,53} For adsorption on Ni{111},
48 however, the work by Ghiringhelli et al.⁵⁴ suggests that alanine
49 will adsorb via the N atom and a single carboxyl group O
50 atom (the other O atom being protonated). Recent work

51 by Baldanza et al.⁴⁵ has highlighted the possibility of a bidentate
52 binding mode for alanine on the Cu{111} surface, where the
53 alanine adsorbs on the surface as the zwitterion, with single Cu
54 atoms bonding to the two O atoms only.

55 The present work examines the chemical nature and
56 orientation of the smallest chiral modifier, alanine, bound to
57 the Ni{111} surface. Synchrotron-based XPS and NEXAFS
58 spectroscopy are used to determine the chemical state and out-
59 of-plane orientation of the adsorbed molecules. Temperature-
60 programmed (TP-)XPS experiments were performed to
61 investigate changes in the chemical state and decomposition
62 pathways followed by the molecules upon annealing. The
63 experimental results indicate that alanine initially adsorbs in its
64 anionic form with a tridentate binding mode at low coverages.
65 With increasing coverage, alanine begins to bind as the zwitterion
66 in a bidentate mode, with surface saturation at 0.25 ML and both
67 forms of alanine coexisting on the surface. Desorption and
68 decomposition are seen to begin together at around 300 K, with
69 decomposition occurring in a multistep process below 450 K.
70 NEXAFS spectra exhibit strong angular dependence of the π^*
71 resonances associated with the carboxylate group, allowing the
72 establishment of two distinct orientations with respect to the
73 surface, which are presumed to correspond to the bidentate and
74 tridentate binding modes.

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75 ■ EXPERIMENTAL SECTION

76 The experiments were carried out in two different ultrahigh
77 vacuum (UHV) systems. XPS and NEXAFS experiments were
78 performed at both the UHV endstation of beamline I-311 at the
79 MAXlab Synchrotron Radiation Facility in Lund, Sweden,⁵⁵ and
80 the near-ambient pressure endstation of beamline 9.3.2 at the
81 ALS Synchrotron in Berkeley, CA, USA,⁵⁶ which are equipped
82 with Scienta SES200 and 4000R-Hipp electron energy analysers,
83 respectively.

84 Standard procedures including Ar⁺ ion sputtering and a final
85 annealing step to 1000 K in UHV were applied for sample
86 cleaning. Cleanliness was checked by XPS. The sample
87 temperature was measured by means of K-type thermocouples
88 spot-welded to the sample. L-Alanine (>98% from Aldrich) was
89 dosed by using a home-built evaporator for organic molecules.
90 The evaporator design consists of one or more resistively heated
91 stainless steel crucibles each containing a glass tube filled with
92 the organic material to be evaporated. K-type thermocouples are
93 attached to the crucibles for accurate and reproducible
94 temperature control. L-Alanine was dosed by heating the crucible
95 to temperatures between 413 and 428 K.

96 For the experiments performed at MAXLab and ALS, L-
97 alanine was adsorbed at sample temperatures of 250 and 300 K,
98 respectively. XP spectra in the C 1s, N 1s, and O 1s regions were
99 recorded using photon energies of 410, 525, and 650 eV,
100 respectively, and a pass energy of 50 eV. Spectra of the Fermi
101 edge were measured after each change in the beamline settings
102 with the same monochromator and analyzer parameters (photon
103 energy, pass energy) to calibrate the offset of the binding
104 energies. In order to study the temperature dependence of the
105 adsorbed alanine layers, fast XP spectra were recorded at a rate of
106 typically 27 s per spectrum while annealing the sample at a
107 constant rate of 0.2 K s⁻¹ (5.4 K per spectrum).

108 NEXAFS spectra at the O, C, and N K-edges were recorded at
109 MAXlab in Auger-yield mode using the same analyzer as for XPS
110 and with the photon beam at angles of 0° or 65° with respect to
111 the surface normal. O, C, and N K-edge spectra were recorded
112 in the kinetic energy ranges 490–525 eV, 250–295 eV, and
113 365–390 eV, while the photon energy was scanned from 525 to
114 560 eV, 280 to 310 eV, and 395 to 415 eV, respectively, in steps of
115 0.15 eV. The photon flux, I_0 , was recorded by measuring the
116 current from a photodiode inserted into the beamline. NEXAFS
117 spectra were measured for the clean surface and after adsorption
118 of L-alanine at 250 K. The data presented here are integrated
119 over a smaller kinetic energy range of 511.3–513.3 eV (262.5–
120 265.0 eV, 380.5–383.5 eV) in order to avoid artificial structures
121 due to photoemission peaks. These spectra are then divided by I_0 ,
122 normalized with respect to the pre-edge (low photon energy)
123 background, and have the clean-surface spectra subtracted, which
124 were treated in the same way. In a final step, the spectra were
125 normalized with respect to the postedge (high photon energy)
126 background.

127 ■ RESULTS AND DISCUSSION

128 **Chemical State.** Figure 1 shows the XP spectra recorded in
129 the C 1s, N 1s, and O 1s regions after L-alanine deposition onto
130 the surface at 250 and 300 K. Coverage calibration was per-
131 formed via comparison of the O 1s and C 1s XPS signal areas with
132 those for a saturated CO layer of known coverage $\Theta = 0.57$ ML⁵⁷
133 using an excitation energy of 950 eV (not shown). With lower
134 photon/kinetic energies, photoelectron (as used for recording
135 the data in Figure 1) diffraction can lead to variations in peak

intensities which are not proportional to the surface coverage. 136
The coverage, Θ , is subsequently expressed in terms of L-alanine 137
molecules per surface Ni atom, assuming that all of the carbon 138
and oxygen present on the surface are incorporated in alanine 139
molecules. Strong photoelectron diffraction effects are seen at 140
lower photon energies, with relative peak intensities varying 141
significantly. Whereas previous work on Cu surfaces has shown 142
relatively slight change in overall spectral structure with coverage, 143
the C 1s spectra shown in Figure 1 show notable variation 144
with coverage within the 283–286 eV binding energy envelope. 145
This is attributed to the generation of decomposition fragments 146
at low coverages of alanine, and their presence complicates the 147
fitting of the spectra and the calibration of the surface coverage of 148
intact alanine. 149

Figure 1a,d shows C 1s spectra recorded after dosing at 250 150
and 300 K, respectively. Six distinct peaks can be identified at low 151
coverages, with binding energies of 283.5, 284.5, 285.3, 285.7, 152
287.3, and 288.9–289.1 eV (see Tables S1 and S2 in the 153
Supporting Information for a list of fitting parameters). As 154
coverage increases above 0.11 ML at 250 K (0.20 ML at 300 K), 155
the peak at 285.3 eV appears to shift to 285.0 eV, while, in the 156
multilayer, the peak at 285.7 eV shifts to 286.1 eV. At multilayer 157
coverage, only three peaks can be clearly resolved at 289.1, 287.3, 158
and 286.1 eV. The multilayer is only seen to grow for deposition 159
at 250 K; at 300 K, no multilayer could be generated. The two 160
low binding energy peaks have been assigned to decomposition 161
fragments based on previously published results: atomic carbon 162
or H_xCN are assigned at 283.5 eV^{58–60} and –C=C– groups at 163
284.5 eV.^{61–67} The 285.3 eV peaks incongruously high intensity 164
at low coverage, coupled with its apparent movement to 285.0 eV 165
with higher coverage, is consistent with known behavior of CO 166
when coadsorbed with hydrocarbons,⁵⁷ although it should be 167
noted that dehydrogenated H_xCN species on Pd{111} also 168
generate a C 1s signal at 285.0 eV.⁶⁸ The remaining peaks at 169
285.7–286.1 eV, 287.3 eV, and 288.9–289.1 eV are assigned in 170
a similar fashion to previous work on Cu{110} and Cu{111}.^{28,45} 171
The low coverage 285.7 eV signal is assigned to the superposition 172
of peaks arising from the methyl and α -C atoms in the anionic 173
alaninate. At the multilayer coverage, the methyl C atom peak 174
shifts to 286.1 eV. This shift is a result of reduced screening by the 175
metal in the multilayer as opposed to the monolayer, a similar 176
effect being seen for the carboxylate group shifting from 288.9 to 177
289.1 eV. The 287.3 eV peak is assigned to the α -C atom of 178
zwitterionic alanine (bound to a –NH₃ group). 179

Figure 1b,e shows N 1s spectra for alanine dosed onto the Ni 180
surface at 250 and 300 K, respectively. Three peaks are observed, 181
at 397.2 eV (only at 300 K), 399.7 eV, and 401.6 eV. These peaks 182
are assigned to a H_xCN decomposition species, –NH₂ from 183
anionic alaninate and –NH₃⁺ from zwitterionic alanine. 184

The O 1s spectra of Figure 1c,f shows two chemical 185
environments, a dominant peak at 531.5 eV and a small feature 186
at 533.5 eV. These are assigned to the deprotonated carboxylic 187
acid group and the protonated –COOH, respectively.⁴¹ The 188
peak at 533.5 eV is only presented above 0.11 ML at 250 K and 189
above 0.19 ML at 300 K. It is likely that its presence is connected 190
to the abundance of hydrogen on the surface. For submonolayer 191
coverages (i.e., <0.25 ML), the 533.5 eV peak appears to have a 192
fixed relative intensity of 10% of the main peak at 531.5 eV, 193
suggesting a degree of hydrogen bonding between the adsorbed 194
molecules. At coverages above 0.25 ML, the O 1s signal shifts to 195
progressively higher binding energies, indicating the decoupling 196
of the carboxylate group from the surface and the formation of 197
the multilayer. The relative intensity of the 533.5 eV signal also 198

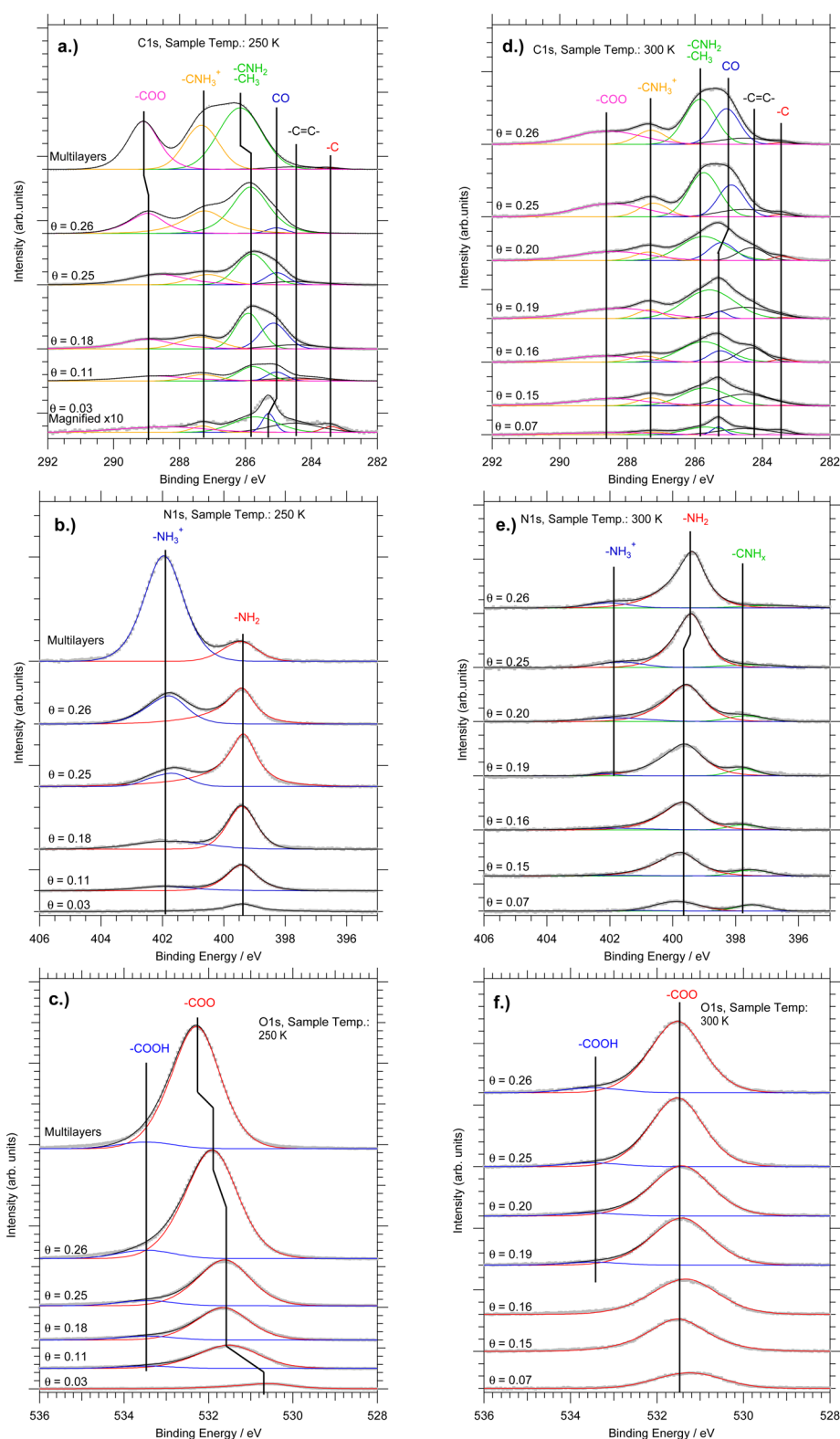


Figure 1. C 1s (a, d), N 1s (b, e), and O 1s (c, f) XP spectra of L-alanine overlayers on Ni{111}, recorded at 410, 525, and 650 eV, respectively. Gray dots represent raw data and continuous lines fitted peaks. Spectra (a–c) adsorbate dosed onto crystal held at 250 K, and (d–f) adsorbate dosed onto crystal held at 300 K.

199 decreases at high coverages, which leads to the conclusion that
 200 alanine is deprotonated in the multilayer. Throughout this
 201 discussion, it is assumed that the O 1s signal generated by
 202 coadsorbed CO is masked by the intensity of the signal generated
 203 by the alanine.

Figure 2 shows a plot of peak areas (in terms of sur- 204
 face coverage) of the C 1s and N 1s peaks discussed above 205
 against the total alanine surface coverage for adsorption at 250 K. 206
 The 287.3 eV C 1s peak and the 401.6 eV N 1s peak, which 207
 are both associated with zwitterionic alanine, are observed 208

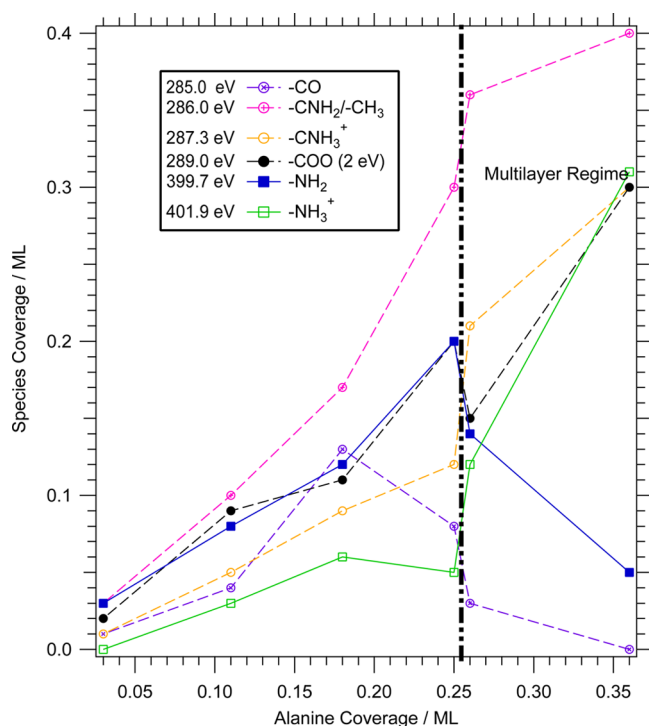


Figure 2. Plot of peak area converted to surface coverage for identified chemical species at varying alanine coverages.

well below the onset of multilayer growth. This raises three possibilities:

- Decomposition fragments are coadsorbed at submonolayer coverages, generating XP signals overlapping with these two peaks.
- A second L-alanine layer begins to grow on the surface before completion of the first chemisorbed layer, producing the NH_3^+ signal from localized multilayers, as suggested by Tysoe for the L-alanine Pd{111} system.³³
- Bidentate, zwitterionic alanine is present on the surface at temperatures above 250 K, in contrast to alanine on Cu{111}, where the bidentate form has only been observed at 200 K.⁴⁵

It appears unlikely that a decomposition product would mimic the zwitterion quite so precisely. On the other hand, the presence of the 287.3 eV peak at very low surface coverages forces us to consider that a fragment may, in fact, be the dominant contributor to this peak in this coverage regime at temperatures above 320 K. Considering the question of multilayer growth before completion of the monolayer, it is noted that the 401.6 eV peak (associated with the zwitterion) begins to grow in the N 1s XP spectrum at below $\Theta = 0.11$ ML. It seems unlikely that a second layer would begin to form at such a low coverage; however, above $\Theta = 0.26$ ML, the 399.7 eV peak begins to shrink compared to the 401.2 eV peak, suggesting the growth of a second layer over the monolayer, attenuating the 399.7 eV peak signal. We, therefore, favor the third possibility that molecules bound to the surface exist in both the bidentate zwitterionic and the tridentate anionic states at coverages above 0.11 ML, with the anionic state being populated first, and second layers beginning to grow at coverages ≥ 0.25 ML. Consequently, it is suggested that 0.25 ML corresponds to the saturation coverage for Ni{111}. This value is supported by the observation of a (2×2) LEED pattern for a saturated glycine layer on Pt{111}⁴⁴

and a (2×2) -like for a saturated layer of alanine on Cu{111} observed by STM.⁶⁹ It should be noted, however, that no superstructure was observed in LEED for any of the L-alanine layers on Ni{111} prepared in this work.

Molecular Orientation. Figure 3 shows O and N K-edge NEXAFS spectra of alanine layers on Ni{111} close to saturation coverage ($\Theta = 0.25$ ML) recorded at 250 K with the X-ray beam at 0° and 65° with respect the surface normal. Carbon K-edge NEXAFS data were also recorded (see the Supporting Information); however, strong background features overlapping with the adsorbate signal do not allow a quantitative analysis of these spectra.

The oxygen K-edge spectra (Figure 3a,b) exhibit four peaks. The π^* resonance consists of two peaks at 532.1 and 532.9 eV, which are attributed to COO groups of intact alanine in two different conformations. Figure 3b highlights the fact that the maximum of the π^* resonance is blue-shifted by 0.15 eV from the 65° to the 0° spectra, confirming that the resonance consists of two distinct peaks with different angle dependencies. The angular dependencies of the high and low energy peaks, respectively, return tilt angles of the COO groups of 56° and 64° to the surface (see the Supporting Information and ref 70 for details). In addition, two broad σ resonances can be resolved at 539.0 and 543.2 eV, which are attributed to C–C and C–O bonds in accordance with the assignments made for glycine and alanine on Cu{110}.^{19,28}

Figure 3c shows nitrogen K-edge NEXAFS spectra. The main peak at 405.5 eV is assigned to a $\sigma(\text{C-N})$ shape resonance. The weak peak at 399.5 eV is assigned to a π^* transition of $\text{C}=\text{N}$ decomposition fragments, in agreement with prior work identifying X-ray induced decomposition fragments of amino acids.^{28,71,72} Such low energy peaks are also seen to develop as alanine-covered Cu surfaces are annealed.⁴²

A broad conclusion from these NEXAFS data is that alanine is adsorbed on the surface in two conformations discriminated by differing tilt angles of the COO plane. It is presumed that these two conformations correspond to tridentate (low tilt angle) and bidentate (high tilt angle) adsorption.

Temperature Dependence. Figure 4 summarizes the temperature-programmed XPS experiments examining the behavior of adsorbed alanine during annealing. The TP-XPS diagrams (Figure 4a–c) comprise a series of fast XP spectra recorded as the temperature of the sample was ramped at 0.2 K s^{-1} (5.4 K per spectrum; for examples, see Figure 4d–f) and merged to produce a 2D plot of temperature vs binding energy, where false color is used to show spectrum intensity. Vertical line profiles taken at fixed binding energies are shown to illustrate the evolution of the chemical state of the adsorbed alanine (Figure 4g–i).

The behavior starting from multilayer coverage (>0.4 ML), dosed at 200 K, is tracked through the N 1s spectra in Figure 4a,d,g, while the N 1s and C 1s spectra in Figure 4b,e,h and c,f,i show the behavior of the saturated chemisorbed layer ($\Theta = 0.25$ ML) dosed at 250 K. Significant changes take place between 300 and 480 K. Complete decomposition of surface-bound alanine is associated with the appearance of C 1s and N 1s peaks at around 283.5 and 397.2 eV, respectively. Alanine desorption commences at around 300 K, as the N 1s peak at binding energy 399.7 eV and the C 1s peaks at 286.0 and 289.0 eV, previously assigned to intact alanine, decrease dramatically, initially without a corresponding growth of the peaks assigned to decomposition products. Decomposition starts at around 320 K and proceeds to

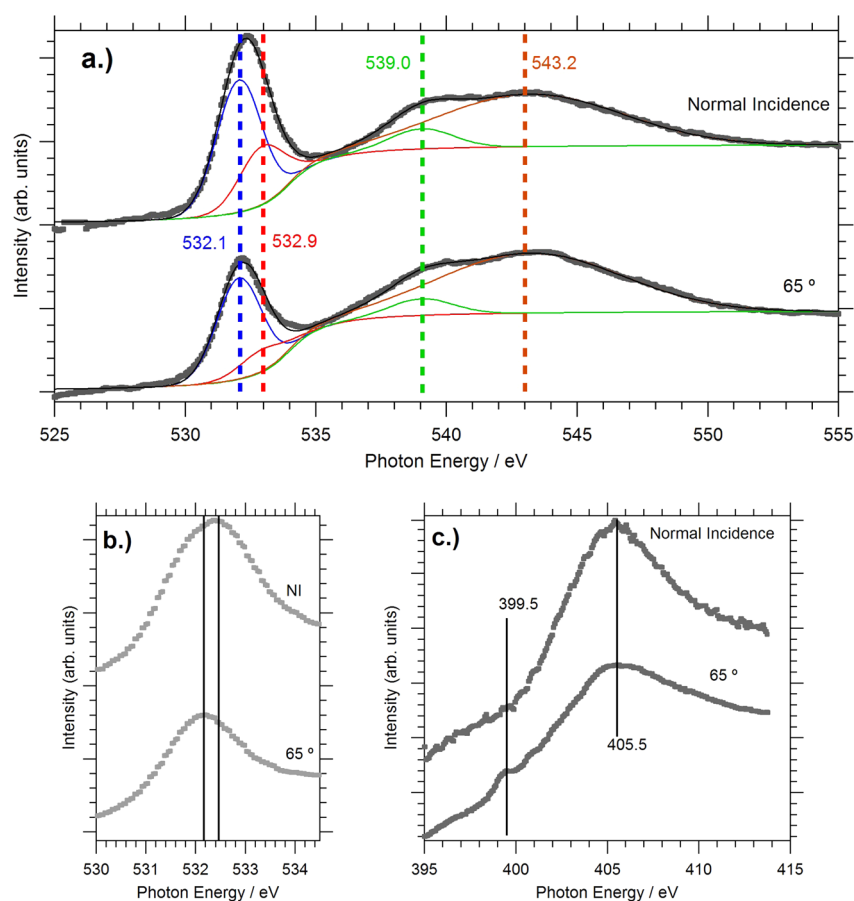


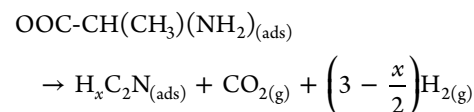
Figure 3. Experimental data (dots) and fits (lines) of Auger yield NEXAFS (a, b: O K-edge data; c: N K-edge data) for a saturated layer of L-alanine on Ni{111} at 250 K. Angles indicated are those between the surface normal and the X-ray beam.

306 completion in at least three steps at ≈ 320 K, ≈ 380 K,
 307 and ≈ 460 K, which are most clearly seen in the C 1s spectrum
 308 (Figure 4b,h). Above 480 K, the N 1s and C 1s signals cease to
 309 change in binding energy or intensity until they disappear
 310 completely at ≈ 730 K.

311 Temperature-programmed desorption experiments (see the
 312 Supporting Information) show evolution of H₂ and CO₂
 313 between 300 and 450 K with two peaks at 350 and 420 K and
 314 two desorption peaks of relative mass 28, which are assigned to
 315 CO (420 K) and N₂ (570 K). Considering the higher heating rate
 316 in the TPD experiments (1 K s⁻¹), the desorption/decom-
 317 position processes are expected to be shifted to higher
 318 temperatures. We, therefore, correlate the TPD peaks to the
 319 changes observed in TP-XPS as follows: 350 K (TPD) \leftrightarrow 320 K
 320 (TP-XPS), 420 K (TPD) \leftrightarrow 380 K (TP-XPS), 570 K (TPD) \leftrightarrow
 321 460 K (TP-XPS). The assignment of the mass 28 TPD peaks is
 322 based on the observation that there is no oxygen signal visible in
 323 XPS above 400 K, whereas a desorption of a small fraction of
 324 nitrogen is compatible with the N 1s signal.

325 The first two decomposition steps include dehydrogenation
 326 and lead to the release of hydrogen and the breaking of the C–C
 327 “backbone” of the amino acid, which is accompanied by the loss
 328 of CO₂. Both processes occur in parallel; however, the different
 329 relative intensities in the H₂ and CO₂ desorption features
 330 indicate that dehydrogenation dominates in the first step, leaving
 331 behind an intermediate, which undergoes C–C bond scission in
 332 the second step. Overall, the surface reaction taking place

between 300 and 400 K is



334 Figure 4g,h,i shows that the integrated area under the N 1s XPS
 335 signal does not change much above 350 K, whereas the inte-
 336 grated C 1s signal falls by more than 50% between 350 and 480 K,
 337 supporting the above mechanism. The amine residue will then
 338 undergo sequential loss of hydrogen until a highly unsaturated
 339 hydrocarbon is left on the surface. This transition causes the
 340 changes in binding energy seen around 460 K. Figure 4h shows a
 341 drop to 56% in the integrated N 1s signal (1/1.8 of the initial inten-
 342 sity) for the saturated chemisorbed layer between 300 and 500 K,
 343 whereas the integrated C 1s drops to 34% (1/2.9 in Figure 4i).
 344 Considering that the C:N ratio for intact alanine (at 300 K) is 3:1
 345 and disregarding possible intensity variations due to diffraction
 346 effects, this would lead to a C:N ratio of 2:1 after complete
 347 decomposition. The C 1s and N 1s signals of the decomposition
 348 products disappear at the same temperature, 730 K (N₂ desorbs
 349 into the gas phase, and C dissolves into the Ni bulk). This is a
 350 strong indication that the C and N atoms are still associated in a
 351 molecular surface species at higher temperatures rather than
 352 isolated chemisorbed atoms. Additional support for this
 353 assumption comes from the fact that the nitrogen K-edge
 354 NEXAFS spectrum of this species is different from published
 355 spectra of atomic N on Ni surfaces (see Figure S2 in the
 356 Supporting Information and ref 73). An extended N-doped

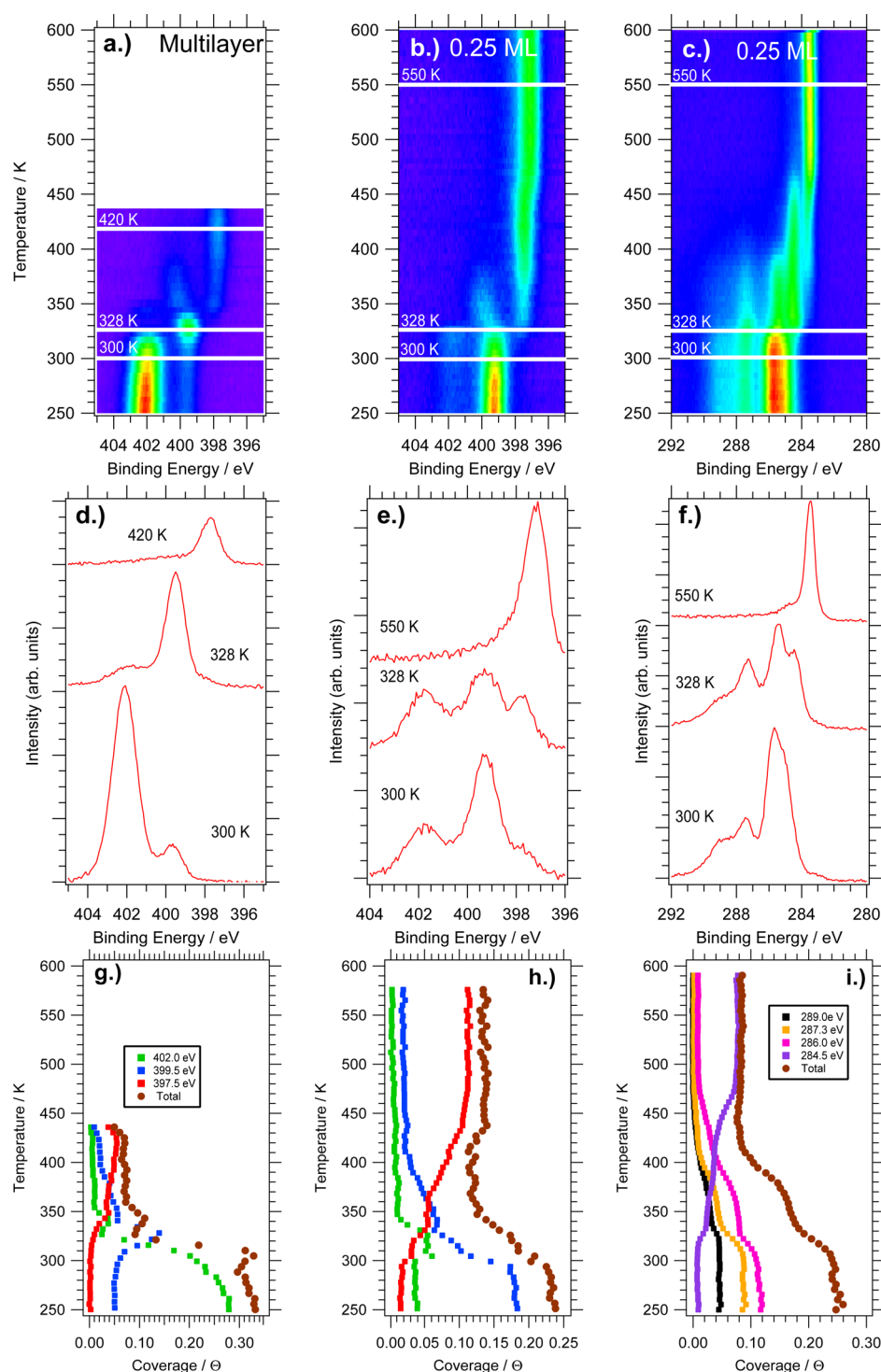


Figure 4. (a–c) TP-XPS data for L-alanine adsorbed on Ni{111}, (d–f) selected spectra extracted from (a–c), (g–i) integrated line profiles showing integrated intensity in 2 eV wide bands. (a, d, g) N 1s (excitation energy 525 eV) of >0.4 ML dosed at 200 K. (b, e, h) N 1s (excitation energy 525 eV) of 0.25 ML dosed at 250 K. (c) C 1s (excitation energy 410 eV) of 0.25 ML dosed at 250 K. Heating rate $\beta = 0.2 \text{ K s}^{-1}$; data acquisition time per spectrum, 27 s.

357 graphene-like C_y-N network can be excluded as the XPS binding
 358 energies reported for such structures are in the ranges of 284.5–
 359 285.0 eV (C 1s) and 398.7–400.7 eV (N 1s), i.e., significantly
 360 higher than the ones found here, and they are stable up to >900 K.⁷⁴

361 ■ CONCLUSIONS AND SUMMARY

362 Thermal decomposition of small amino acids, such as alanine and
 363 glycine, has previously been observed on a suite of surfaces by

means of TPD and TP-XPS, including Cu{111},⁴⁵ Cu{110},^{7,42} 364
 Cu{531},³⁹ Pt{111},^{41,44} and Pd{111}.^{32,33} In all surveyed cases, 365
 decomposition proceeds via a β -hydride elimination step (or 366
 possibly γ -elimination in the case of alanine) prior to C–C bond 367
 scission between the carboxylate and the α -carbon, which leads 368
 to the release of H₂ and CO₂ into the gas phase. This is followed 369
 by sequential dehydrogenation and collapse of the alkyl amine, 370
 releasing H₂ into the gas phase. From the list of decomposition 371
 temperatures in Table 1, it is clear that Ala/Ni{111}, alongside 372

Table 1. Onset of Decomposition for Glycine and Alanine Adsorbed on Different Single Crystal Surfaces under UHV Conditions

	onset (K)	method	ref
Ala/Cu{111}	440	TPD	45
	425	TP-XPS	45
Ala/Cu{110}	500	TPD	7
	470	TP-XPS	42
Ala/Ni{111}	320	TP-XPS	this work
Ala/Pd{111}	350	TPD	33
Gly/Cu{110}	460	TP-XPS	42
Gly/Cu{531}	490	TPD	39
Gly/Pd{111}	320	TPD	32
Gly/Pt{111}	360	TPD	44
	350	TP-XPS	41

Gly/Pd{111}, is among the least stable adsorption systems studied so far. The onset of decomposition around 320 K is more than 100 K lower than that for the close-packed Cu{111} surface. In general, dissociation is far more facile on the group-10 transition metals than on Cu. Therefore, Cu surfaces are not likely to be good models for enantioselective modification of the Ni, Pd, or Pt catalysts that are used in enantioselective hydrogenation reactions. Important in this context is also the fact that the onset of decomposition for Ala/Ni{111} is within the range of temperatures used for chiral modification of Raney nickel (273–373 K) and below the typical range of reaction temperatures for the enantioselective hydrogenation of MAA (333–343 K).^{2,3} This is a clear indication that the active modifier under reaction conditions is not the intact alaninate surface species that we observed below 320 K in UHV in the present study. Under reaction conditions, decomposition of alanine may be hindered by the presence of solvent and/or the reactants on the surface; however, this would almost certainly also involve a change in the geometry and bond coordination of the adsorption complex. A recent near-ambient pressure study of Gly/Pt{111} has shown that an aqueous environment up to 0.2 Torr does not stabilize the intact adsorbate on this system significantly, although the stability and reaction path of the decomposition products change.⁴¹ Payne et al.⁷⁵ have shown that exposure to water vapor in the range of 10⁹ L (= 10³ mbar s) leads to significant surface oxidation on polycrystalline Ni samples, as one would also expect under catalytic reaction conditions.⁷⁶ Neither the behavior of relevant chiral molecules on Ni oxides nor the influence of other solvents or the reactants of the hydrogenation reaction on the surface chemistry of amino acids have been studied so far. Another possibility is that the H_xC₂N_(ads) surface species resulting from the first decomposition step is the actual modifier. This species is not chiral, but it could stabilize a chiral reconstruction of the Ni surface and/or form a chiral complex with the solvent or one of the reactants.

In summary, the chemical state and molecular orientation of L-alanine on Ni{111} have been investigated between 250 and 600 K. Synchrotron-based XPS and NEXAFS spectroscopy show that alanine is adsorbed in both the anionic (tridentate) and the zwitterionic (bidentate) states up to room temperature, with both states coexisting in the saturated layer. The chemisorbed layer saturates at $\Theta = 0.25$ ML. At low coverages, decomposition of the alanine molecule is significant even at 250 K; however, at higher coverages, surface congestion appears to prevent break up of the molecule. Desorption and decomposition begin together

at ≈ 300 K, with decomposition occurring in a multistep process up to 450 K.

ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/acs.jpcc.5b08814.

Data analysis and additional data (PDF)

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Notes

The authors declare no competing financial interest.

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