

Optical Sensing and Switching with Designed Molecules

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Abstract: Some of the recent research in optical sensor and switch molecules at Queen's University of Belfast is outlined. The ion-induced absorption spectral changes of molecules with ICT (Internal Charge Transfer) excited states are used to demonstrate superposed logic where multiple logic configurations are simultaneously observable. Similar spectral effects in multi-receptor structures give rise to XOR logic. The ion-induced luminescent spectral changes of molecules with ICT excited states can be combined with the phenomenon of Photoinduced Electron Transfer (PET) to exemplify tristable systems. Insights can also be gained into the luminescence effects initiated in designed molecules by transition metal ions in non-aqueous media. The ion-induced luminescent spectral changes of PET systems with one or two receptors display YES and AND logic respectively. The latter yield coincidence sensors for the direct detection of ion pairs. Overall, interesting optical sensors and switches can be produced from rather small molecules when the latter are fitted with ion-receptors.

Keywords: Fluorescent ICT sensors · Fluorescent PET sensors · Indicators · Molecular logic gates

1. Introduction

Since molecules or atoms operate in a far smaller domain than ours, we need special molecules on location so that scientists can learn the secrets of the molecular world. Optically empowered molecules do this job very well. Suitable signals are wavelength shifts in electronic absorption spectra or in fluorescence emission spectra or even better, fluorescence quantum yield changes in the case of the latter. Combining differently functional units in different ways is a productive approach to obtaining chemically-induced optical signals [1]. Wavelength shifts are produced from integrated 'Chromo/Fluorophore-Receptor' formats accommodating internal charge transfer (ICT) excited states [2]. 'Fluorophore-Spacer-Receptor' formats involve pho-

toinduced electron transfer (PET) which can switch 'off' fluorescence emission. This can be chemically reversed so that the fluorescence quantum yield becomes high again [2]. Expansion of the above formats to accommodate two receptors produces 'Receptor₁-Chromophore-Receptor₂' and 'Receptor₁-Spacer₁-Fluorophore-Spacer₂-Receptor₂' systems. These two basic formats can also be fused to a 'Fluorophore-Receptor₁-Spacer-Receptor₂' system. Each of these systems has its sensory/switching uses, as we will demonstrate.

2. 'Chromophore-Receptor' Systems

1 [3] is a 'chromophore-receptor' system containing a dye structure with a conjugated π -electron set, with clear donor and acceptor terminals. **1** also contains a tetra-anionic chelating site for multivalent cations. So it is not at all surprising that Ca^{2+} perturbs the ICT excited state of **1**, with a resulting blue shift in the UV-Vis absorption spectrum (Fig. 1). From a molecular logic viewpoint [4], let's take Ca^{2+} as the

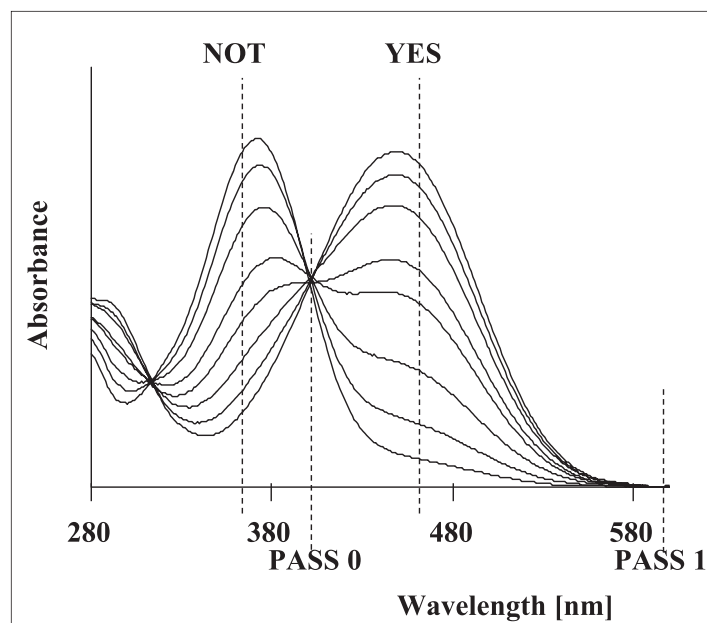


Fig. 1. Absorption spectra of **1** with varying pCa, showing YES, NOT, PASS 0 and PASS 1 logic. pCa values in order of increasing absorbance at 470 nm; 2.3, 5.2, 5.7, 6.0, 6.2, 6.5, 7.0, infinity

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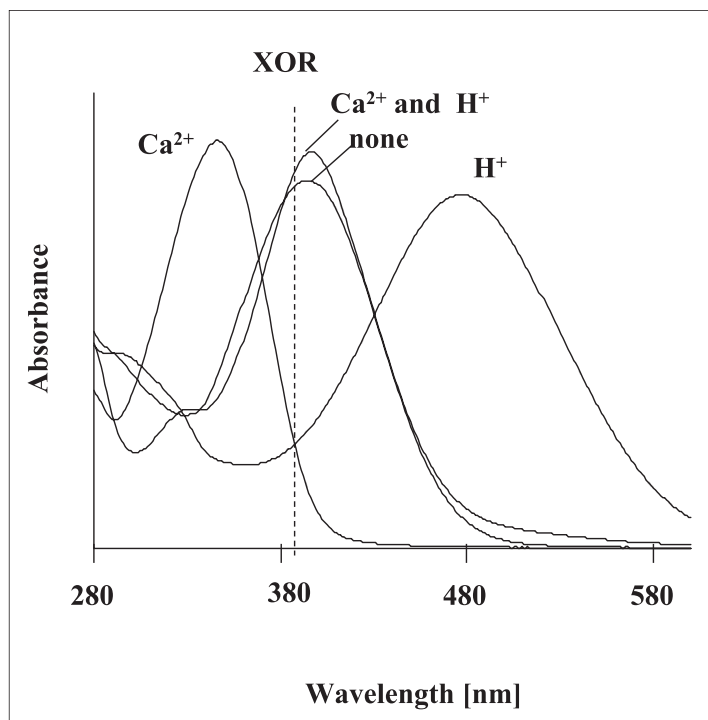


Fig. 2. XOR logic with **2**, transmittance observed at 387 nm

Table 2. Truth table showing XOR logic with **2**

Input ₁	Input ₂	Output (transmittance) 387 nm
Ca ²⁺	H ⁺	XOR
0(low, <10 ⁻⁹ M)	0(low, 10 ^{-9.7} M)	0(low, 16)
0(low, <10 ⁻⁹ M)	1(high, 10 ^{-5.6} M)	1(high, 59)
1(high, 10 ^{-2.3} M)	0(low, 10 ^{-9.7} M)	1(high, 58)
1(high, 10 ^{-2.3} M)	1(high, 10 ^{-5.6} M)	0(low, 15)

4. 'Fluorophore-Receptor₁-Spacer-Receptor₂' Systems

3 [10] has an amine as a receptor for H⁺ and a popular biochemical fluorophore. But the latter also carries a deprotonatable N-H group which swings into action at sufficiently alkaline pH values. Therefore **3** fits the 'Fluorophore-Receptor₁-Spacer-Receptor₂' system. The absorption spectra of **3** are strongly dependent on pH in the alkaline range with a large H⁺-induced red shift (77 nm) but much less so in the acidic/neutral pH range showing only a small H⁺-induced blue shift (8 nm). The former behaviour illustrates ICT sensory behaviour as in common pH indicators [6], whereas the mild response in the acidic/neutral range is observed for fluorescent PET pH sensors which have significant electrostatic interactions across the spacer [11]. Of course, ideal PET sensors should show no distur-

bance of the absorption spectrum at all [1]. The fluorescence emission spectra of **3** are strongly dependent on pH in both the alkaline range and in the acidic/neutral pH range with a clear plateau in-between. There is a strong switching 'on' of fluorescence with an FE (fluorescence enhancement) of 30 with a corresponding pK_a value of 9.3. The acidic/neutral region also shows a further large increase but the FE value is only 2.7 simply because it is calculated from a far higher baseline of the plateau reached following the protonation within the alkaline region. The corresponding pK_a value is 6.7. **3** is a fortunate case where the ICT and the PET behaviours are clearly separated in pH space because the pK_a value of the morpholine group is rather low due to the electron withdrawing oxygen. Furthermore, the fluorescence quenching by PET is kept at a manageably low rate by using the rather long trimethylene spacer. Sufficient fluorescence quantum yield is then left over for quenching when the 4-NH is deprotonated in the alkaline range. Such systems show tristability: three pH-independent signal regions at acidic, alkaline and intermediate pH ranges (Fig. 3) [12].

A relative **4** allows us to test the ability of rather basic sensors to respond to transition metal ions in aprotic media – a subfield which has received quite a lot of attention in recent years [13][14]. Indeed, **4** produces large FE values in response to chromium (III) chloride hexahydrate in acetonitrile [13]. However, a classical PET sensor **5** for H⁺ which was especially sterically crowded for the occasion also produced a FE value almost as large. Metal ion coordination is expected to be more sensitive to steric bulk in the ligand than protonation. **5** is expected

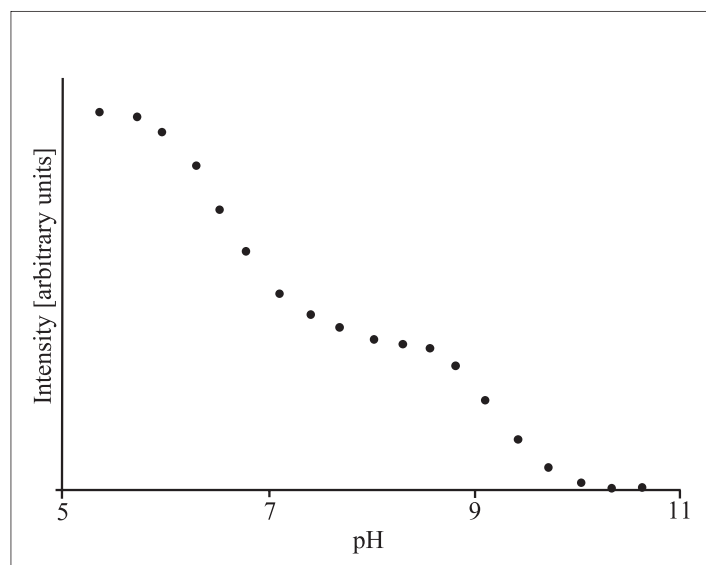


Fig. 3. Intensity (at the fluorescence band maximum) – pH profile of **3** (when excited at an isosbestic point)

to be much poorer than 4 for metal binding also because the former lacks chelation possibilities. Equally importantly, Cr^{3+} is known to be a kinetically inert ion which will not exchange members of its first coordination sphere within the timescale of a fluorescence spectroscopic experiment [15]. So it appears that the large FE values are a result of acid–base reactions rather than metal coordination to sensors such as 4. Where do the acid–base reactions come from? Water ligands in the coordination sphere of transition metal ions of high charge are known to be quite acidic [15]. So these can interact with the basic receptor units of the sensors. This means that strong basicity can be a pitfall which needs to be avoided when designing PET sensors for transition metal ions. Classical analytical chemistry accounts for protonation of ligands in water by employing conditional formation constants [16], for instance. So sensor designers would be wise to follow the analytical chemists into water or at least mixed aqueous solvents as the medium with pH-buffering.

5. 'Receptor₁-Spacer₁-Fluorophore-Spacer₂-Receptor₂' Systems

6 [17] is a 'Receptor₁-Spacer₁-Fluorophore-Spacer₂-Receptor₂' system which combines a textbook hydrocarbon fluorophore with a benzo-15-crown-ether unit as 'Receptor₁' and a partially protonatable tetraamine as 'Receptor₂'. Then, Na^+ and anionic phosphate species serve as inputs into the two receptors respectively. Each of these block pre-existing PET channels according to literature precedents in simpler 'Fluorophore-Spacer-Receptor' systems [18][19]. So strong fluorescence is visible only when the ion-pair of sodium phosphate is presented to 6 in methanol/water 1:1 (v/v) at pH 8. This is AND logic behaviour. AND logic was the first indication that molecules could perform some of the operations [20] that modern computer technology was built on [21]. Direct detection of ion-pairs brings AND logic into useful analytical sensing situations. The experimental observation in the present case is clouded by the fact that 'innocent' ions like Cl^- show significant interactions with the protonated tetraamine units of 6 in the mixed aqueous solution employed in the experiment. Nevertheless, next-generation systems can improve on this aspect so that sensing of coincident species becomes practical.

Acknowledgement

We close this thumbnail sketch of business at Belfast by thanking the European Social Fund, the Department of Employment and Learning in Northern Ireland, Invest NI (RTD COE 40) and

the Japan Society for the Promotion of Science for their support.

Received: March 1, 2005

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