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# Utilization of Promising Calcium Manganite Oxygen Carriers for Potential Thermochemical Energy Storage Application

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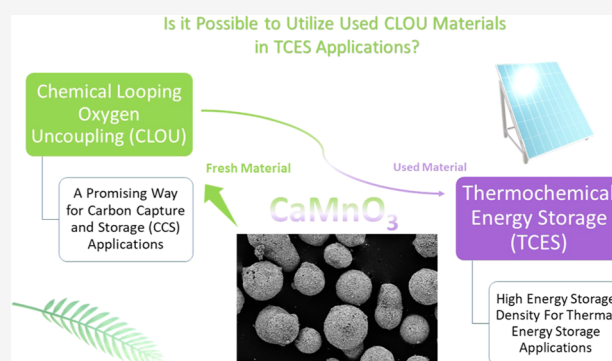
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**ABSTRACT:** In this study, oxygen release/consumption behavior of calcium manganese-based oxides ( $\text{CaMn}_{1-x}\text{B}_x\text{O}_3$ , where B: Cu, Fe, Mg and  $x = 0.1$  or  $0.2$ ) used in a chemical looping oxygen uncoupling (CLOU) application was investigated. The effect of B-site dopants such as Fe, Mg, and Cu on the oxygen release behavior was also investigated with the aim to use these materials in thermal energy storage (TES). Previous literature studies about CLOU performance of doped calcium manganites were taken into consideration for dopants selection. Calcium manganite-based oxides have been used in chemical looping oxygen uncoupling (CLOU) applications owing to their oxygen release behavior to the gas phase. Studies have revealed that calcium manganite-based oxides show a promising nonstoichiometry over a range of temperatures and oxygen partial pressures, which makes them useful for thermochemical energy storage applications. However, the related literature studies have been mainly focused on their nonstoichiometric characteristics related to temperature and oxygen partial pressure and thermodynamic properties. In this work, thermal analysis and fluidized bed tests were carried out as complementary techniques.  $\text{CaMn}_{0.8}\text{Cu}_{0.2}\text{O}_3$  showed the highest oxygen release performance in fluidized bed tests, while  $\text{CaMn}_{0.9}\text{Mg}_{0.1}\text{O}_3$  had the best cyclic stability overall among the samples used in the study.



## INTRODUCTION

Thermal energy storage (TES) is a necessary tool to reach the sustainable energy goals via green and sustainable energy resources.<sup>1</sup> Continuous electricity production is important, and it only becomes possible with storing the energy during on-sun hours and releasing the stored energy during off-sun hours.<sup>2</sup> Among the energy storage techniques, thermochemical energy storage (TCES) exhibits a significantly good performance regarding energy storage density (0.5–1 kWh/kg),<sup>3</sup> theoretically unlimited storage period or storage within a large temperature range (25–1000 °C).<sup>4</sup> Besides, there is no need to use specific and complex storage procedures as the heat is stored in the form of chemical energy in TCES.<sup>5</sup> To store a high amount of energy, redox reactions of metal oxides can be useful since their operational temperatures are relatively high.<sup>6</sup> From metal oxides,  $\text{Co}_3\text{O}_4$ ,  $\text{CuO}$ ,  $\text{Fe}_2\text{O}_3$ , and  $\text{Mn}_2\text{O}_3$  show remarkable performance to meet the expectations for TCES applications.<sup>7</sup> Among these oxides,  $\text{Mn}_2\text{O}_3$  is very promising since it is less expensive, more abundant, more cyclically stable, and environmentally friendly.<sup>8,9</sup> Since the reoxidation rates of pure  $\text{Mn}_3\text{O}_4$  to  $\text{Mn}_2\text{O}_3$  are slightly slow, the combination of the Mn oxides was also tested in some studies to increase the performance of the Mn oxides for thermochemical energy storage applications.<sup>4,10–15</sup> Especially, Cu/Mn combined oxides showed a lower redox temperature than the pure Cu or Mn oxides.<sup>4,16</sup> In addition, Fe/Mn combined oxides also

showed much faster reoxidation rates compared to pure Mn oxides.<sup>10,13,16</sup>

In the literature, there are a few studies about  $\text{CaMnO}_3$  as a thermochemical energy storage material.<sup>17,18</sup> These studies were focused on redox characteristics, redox chemistry, and oxygen nonstoichiometry of the material.<sup>17–20</sup> However, there is no available systematic study about the effect of dopants on the oxygen release behavior of  $\text{CaMnO}_3$  so far.

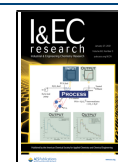
However, Mn-based systems require performance improvement as they feature slow kinetics of the oxidation step.<sup>14</sup> Slow kinetics of Mn-based system have also been seen in chemical looping oxygen uncoupling applications.<sup>21</sup> To tackle this problem in chemical looping oxygen uncoupling (CLOU), Mn oxides have been combined with other oxides such as  $\text{SiO}_2$ ,  $\text{Fe}_2\text{O}_3$ ,  $\text{MgO}$ , and  $\text{CaO}$ .<sup>22–24</sup> Especially calcium manganese-based materials have shown great promise as the reactivity is high and estimated lifetime is long enough to be used in chemical looping oxygen uncoupling (CLOU) applica-

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tions.<sup>25–27</sup> Calcium manganite,  $\text{CaMnO}_{3-\delta}$ , has a perovskite-type structure and shows continuous nonstoichiometry with respect to oxygen.<sup>17</sup> The  $\delta$ -factor defines the degree of oxygen deficiency in the structure.<sup>19,28</sup> This characteristic enhances the oxygen mobility in the lattice, and it can release oxygen to the gas phase over a range of temperatures.<sup>29</sup> These features raise  $\text{CaMnO}_3$  to be a perfect fit as a thermochemical energy storage material.

In this study, doped calcium manganite-based oxygen carriers used in methane reactivity tests<sup>29</sup> were investigated as potential thermochemical energy storage materials. The effects of dopants such as Cu, Fe, and Mg on the oxygen release behavior of  $\text{CaMnO}_3$  were investigated systematically. The amount of dopant (A) was kept as  $x = 0.1$  or  $0.2$  where the sample composition was  $\text{CaMn}_{1-x}\text{A}_x\text{O}_3$  (A: Fe, Mg, Cu). The effect of dopant combination such as Cu–Fe and Mg–Ti was also investigated.

## EXPERIMENTAL SECTION

The oxygen carriers used in this study were produced by VITO (Flemish Institute for Technological Research) by the spray-drying method, and the materials were designed to have a molar composition of  $\text{CaMn}_{1-x}\text{A}_x\text{O}_3$  (A: Fe, Mg, Cu, Cu&Fe, Mg&Ti). The detailed synthesis conditions and the parameters related to the chemical looping oxygen uncoupling application, where the chosen calcium manganates were used as oxygen carriers, can be found elsewhere.<sup>29</sup> The starting chemical compositions of the powders are given in Table 1. The phase

**Table 1.** Designed Chemical Composition and Notation of Samples Used in This Study<sup>a</sup>

sample <sup>ref work</sup>	composition
CM <sup>30</sup>	$\text{CaMnO}_3$
CMF1 <sup>31</sup>	$\text{CaMn}_{0.9}\text{Fe}_{0.1}\text{O}_3$
CMF2 <sup>31</sup>	$\text{CaMn}_{0.8}\text{Fe}_{0.2}\text{O}_3$
CMM1 <sup>26,29,32</sup>	$\text{CaMn}_{0.9}\text{Mg}_{0.1}\text{O}_3$
CMM2 <sup>29,32</sup>	$\text{CaMn}_{0.9}\text{Mg}_{0.2}\text{O}_3$
CMC1	$\text{CaMn}_{0.9}\text{Cu}_{0.1}\text{O}_3$
CMC2	$\text{CaMn}_{0.9}\text{Cu}_{0.2}\text{O}_3$
CMCF	$\text{CaMn}_{0.8}\text{Cu}_{0.1}\text{Fe}_{0.1}\text{O}_3$
CMMT <sup>26,33</sup>	$\text{CaMn}_{0.775}\text{Mg}_{0.1}\text{Ti}_{0.125}\text{O}_3$

<sup>a</sup>Reference work (ref work) refers to the published results with using the same composition in a CLOU operation.

analysis of the samples after having been used in CLOU applications (called fresh) and after the fluidized bed tests (called used) is given in the Supporting Information. The samples after having been used in CLOU applications were sieved into the 180–250  $\mu\text{m}$  size range to get rid of possible destroyed particles.

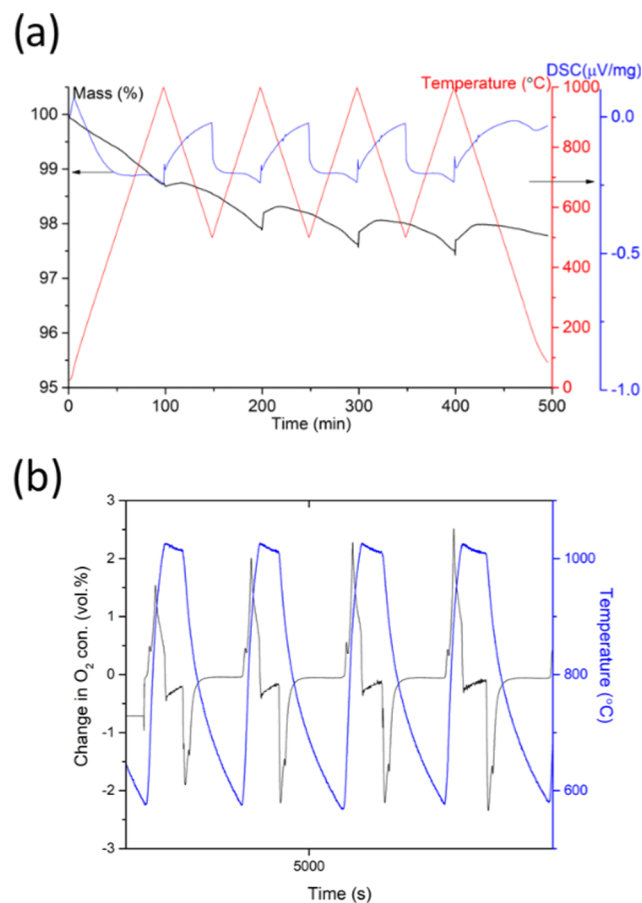
Characterization of the synthesized samples for phase analysis was done using an X-ray diffractometer (XRD, Bruker D8 Advance, Cu  $K\alpha$ , 40 kV, 40 mA) in the  $2\theta$  range of 20–75° with a step size of 0.01. Phase analysis was conducted on samples before and after being used in the cyclic experiments. A scanning electron microscope (SEM, Zeiss LEO Ultra 55 FEG) was used to study the particle morphologies. Detailed information is given in the Supporting Information. For thermal analysis, a simultaneous thermal analyzer (STA)-(Netzsch-STA 409 PC Luxx) was used to determine thermal behavior. For each experiment, the amount of sample was 70 mg. The samples were exposed to air and  $\text{N}_2$  during all

oxidation cycles ( $p\text{O}_2$ : 0.21 atm) and reduction cycles, respectively. To study the oxygen release/consumption behavior of the perovskites, several redox cycles including heating and cooling steps were carried out in a fluidized bed reactor with a length of 820 mm and a porous quartz plate of diameter 22 mm placed 370 mm from the bottom. For each experiment, the amount of sample was 15 g. The reduction and oxidation cycles were initiated by changing the temperature between 600 and 1000 °C for each cycle. The system was flushed by air ( $p\text{O}_2$ : 0.21 atm) during the oxidation cycles and by  $\text{N}_2$  during the reduction cycles. The  $\text{O}_2$  level in the outlet gas was detected by a Rosemount NGA 2000 multicomponent gas analyzer.

## RESULTS

To understand the redox characteristics of the doped calcium manganites used in this study, the properties of the pure calcium manganite were also investigated. The  $\text{CaMnO}_3$  system was chosen as it showed good performance as an oxygen carrier in a chemical looping oxygen uncoupling study.<sup>34</sup> In the literature, thermochemical energy storage experiments are mainly carried out with either thermal analysis setups or packed bed reactors.<sup>6,35</sup> In this study, the redox characteristics of the samples were investigated via both thermal analysis and fluidized bed experiments.

Figure 1a,b shows the thermal behavior of CM under controlled atmosphere and the oxygen release/consumption

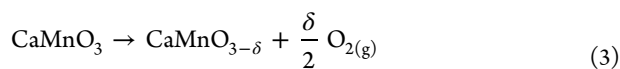
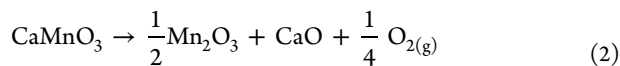
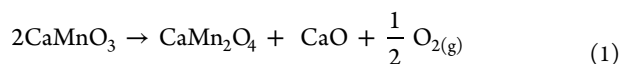


**Figure 1.** Thermal behavior of CM under controlled atmosphere (a) obtained by STA and  $\text{O}_2$  release/consumption ability of CM in the fluidized bed test (b).

Table 2. Redox Properties of the Samples

sample	thermal analysis				fluidized bed experiments					
	$T_{\text{peak,ox}} (^{\circ}\text{C})$	$T_{\text{peak,red}} (^{\circ}\text{C})$	$\Delta w_{\text{ox}} (\%)$ (+)	$\Delta w_{\text{red}} (\%)$ (-)	$T_{\text{peak,ox}} (^{\circ}\text{C})$	$T_{\text{peak,red}} (^{\circ}\text{C})$	$\Delta VO_{2,\text{peak ox}} (\% \text{ av.})$ (+%)	$\Delta VO_{2,\text{peak red}} (\% \text{ av.})$ (-%)	$\Delta mO_{2,\text{ox}} (\text{+g, av.})$	$\Delta mO_{2,\text{red}} (\text{-g, av.})$
CM	795	970	0.2	0.4	885	970	2.20	2.22	0.23	0.31
CMF1	580	970	1.21	1.06	840	980	2.06	1.78	0.28	0.34
CMF2	610	975	1.06	1.02	830	980	1.69	1.64	0.24	0.32
CMM1	750	975	1.35	1.38	895	985	2.62	2.66	0.33	0.39
CMM2	770	980	1.32	1.37	885	980	2.08	2.05	0.29	0.37
CMC1	730	975	0.75	1.15	870	945	2.11	2.33	0.29	0.35
CMC2	740	970	0.9	1.20	860	950	2.66	2.74	0.34	0.41
CMCF	710	980	0.8	0.75	850	975	1.64	1.54	0.29	0.34
CMMT	770	980	0.5	1.05	830	980	1.05	1.31	0.15	0.21

behavior of CM in the fluidized bed test, respectively. The completion of a cycle involving heating and cooling takes 6000 s in STA, while it takes 1650 s in the fluidized bed test. These different time spans of applied thermocycles may affect the reaction kinetics, and different temperatures may be observed for redox cycles (Table 2). Possible reactions are given below (reactions 1–3). The expected mass loss for both reactions 1 and 2 is calculated as 5.6 wt % on average. The mass loss for reaction 3 may vary. Thermal analysis results showed that there was a mass loss of around 1.4 wt % within the first heating step, and the mass loss reached 2 wt % when the first cycle was completed. After the first cycle, the mass loss was around 0.4 wt % on average for each of the following cycles, while the mass gain was around 0.2 wt %. The reduction onset ( $T_{\text{red,on}}$ ), where the reduction starts, and the peak temperature ( $T_{\text{red,peak}}$ ), where the reduction was observed as maximum, were recorded as 700 and 970 °C, respectively. The oxidation onset temperature and the oxidation peak temperature were observed as 975 and 795 °C, respectively. It was reported that  $\delta$  for  $\text{CaMnO}_{3-\delta}$  is lower than 0.01 below 750 °C when the  $p\text{O}_2$  is above 0.01 in the system.<sup>36</sup> In thermal analysis,  $p\text{O}_2$  was set to 0.21 atm. From this point of view, most likely, only reactions 1 and 2 partly occurred simultaneously during the cycles in STA.

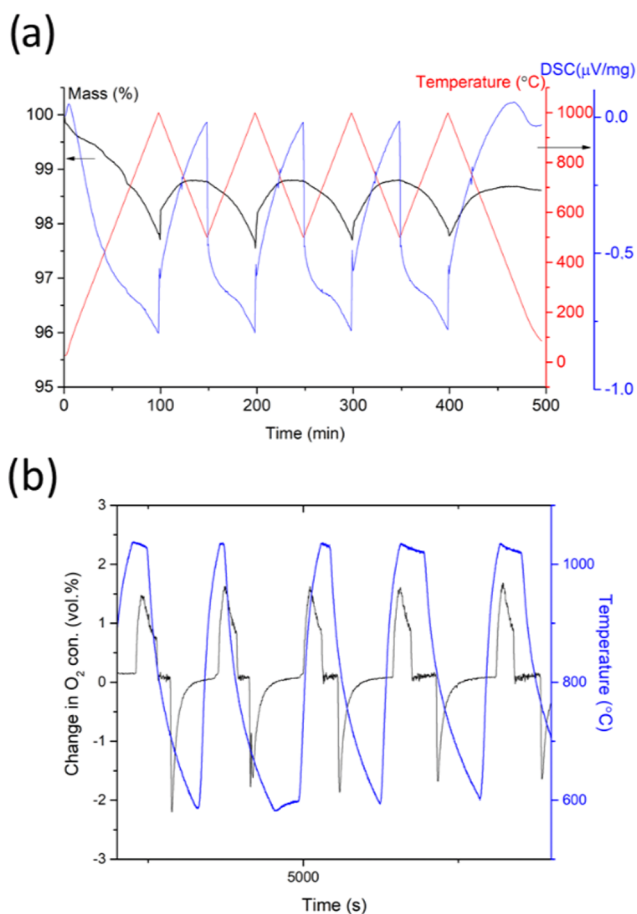


In the fluidized bed test, the result was slightly different. There was no significant temperature difference for reduction peaks between thermal analysis and fluidized bed test; however, the oxidation peak temperature observed in the fluidized bed test was around 90 °C higher than the thermal analysis. Since the heating and cooling rates in the fluidized bed are higher than those in the thermal analysis, different kinetic behavior may be expected for thermal analysis and fluidized bed test.<sup>37,38</sup> The peak point of  $\text{O}_2$  release/consumption during the cycles was obtained from the gas analyzer data (Figure 1b). The total  $\text{O}_2$  release and consumption during the cycles were obtained via calculating the area under the oxidation and reduction peaks (Figure 1b). Two different phase transitions during both oxidation and reduction were reckoned to occur as there was a protrusion observed on the peaks. Theoretical  $\text{O}_2$  release and consumption during redox reaction via reaction 1 are

calculated to be around 5.3 and 100 g of  $\text{CaMnO}_3$ , respectively. Even though these values vary depending on  $\delta$  for reaction 3, probably their averages will be lower than 1.12 and 100 g of  $\text{CaMnO}_3$ , respectively.<sup>36,36</sup> Reactions 1 and 3 may be pathways for the redox reactions in the fluidized bed test, as average  $\text{O}_2$  consumption and release were detected around 0.23 and 0.31 g during oxidation and reduction, respectively. Most likely, the point where the protrusions were observed represents reaction 3, while the other peak was representative for where reaction 1 occurs. These formations were also revealed by XRD phase analysis after the fluidized bed tests (Figure S1), where  $\text{CaMnO}_{3-\delta}$  along with a low amount of  $\text{CaMn}_2\text{O}_4$  were observed. Even though  $\text{CaMnO}_3$  has shown very promising performance for thermochemical energy storage applications in other studies,<sup>17,19,20,39</sup> it did not show the best performance among the samples that were used in our study. This was already expected since the sample was already used as an oxygen carrier in a chemical looping oxygen uncoupling application, which may end up with decreased reactivity. The partial  $\text{O}_2$  pressure in this study was 0.21 bar, which is close to the real applications, where air is used without dilution. It is known that the amount of released oxygen would increase when the  $p\text{O}_2$  in the system decreases.<sup>18</sup> As the other studies suggested that  $\text{CaMnO}_3$  may be reduced via reaction 3, where  $p\text{O}_2$  is lower than 0.18 atm,<sup>17,39</sup> our study revealed that reaction 1 can also occur, where  $p\text{O}_2$  is above 0.21 atm.

Promising results were obtained for both CLOU and TCES applications with Fe doping for  $\text{CaMnO}_3$ .<sup>18,30,39</sup> It has been revealed that Fe doping enhances the reversible redox reactions of  $\text{CaMnO}_3$  and allows operation at higher temperatures than pure  $\text{CaMnO}_3$ .<sup>39</sup> In this study, we used two different  $\text{CaMn}_{1-x}\text{Fe}_x\text{O}_3$  ( $x = 0.1$  and  $0.2$ ), which have been used previously as an oxygen carrier in a CLOU application. The starting-phase compositions of  $\text{CaMn}_{0.9}\text{Fe}_{0.1}\text{O}_3$  and  $\text{CaMn}_{0.8}\text{Fe}_{0.2}\text{O}_3$  can be seen in Figures S2 and S3 in the Supporting Information, respectively. After the CLOU experiments, some  $\text{CaMn}_2\text{O}_4$  formation was observed along with perovskite in both samples. In thermal analysis, the peak oxidation temperature was recorded as 580 °C, which is 215 °C lower than pure  $\text{CaMnO}_3$  (Figure 2). A similar result was also observed in a related study, where the oxygen release onset temperature was lowered below 400 °C by Fe doping to  $\text{CaMnO}_3$ .<sup>39,39</sup> Like  $\text{CaMn}_{0.9}\text{Fe}_{0.1}\text{O}_3$ ,  $\text{CaMn}_{0.8}\text{Fe}_{0.2}\text{O}_3$  also performed (Figure 3) a lower peak oxidation temperature (610 °C) than pure  $\text{CaMnO}_3$  (795 °C). The reduction peak temperature was observed as 970 and 975 °C for  $\text{CaMn}_{0.9}\text{Fe}_{0.1}\text{O}_3$  and  $\text{CaMn}_{0.8}\text{Fe}_{0.2}\text{O}_3$ , respectively. However, almost 6 times higher mass gain was observed for



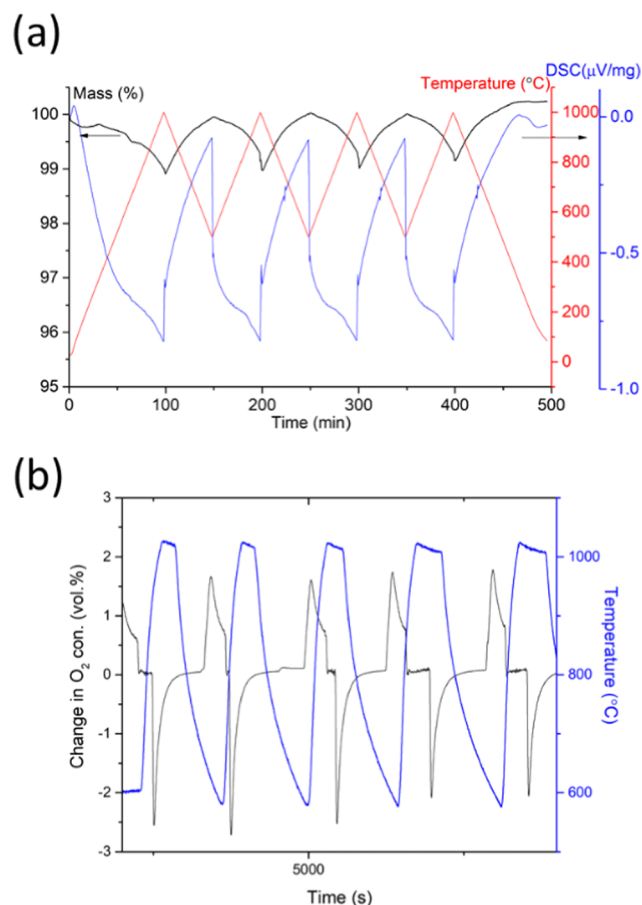


**Figure 2.** Thermal behavior of CMF1 under controlled atmosphere (a) obtained by STA and O<sub>2</sub> release/consumption ability of CMF1 in the fluidized bed test (b).

CaMn<sub>0.9</sub>Fe<sub>0.1</sub>O<sub>3</sub> and the mass loss was 2.65-fold higher than CaMnO<sub>3</sub> during oxidation and reduction, respectively. The mass gain/loss was recorded as 1.06 wt % during oxidation and 1.02 wt % during reduction.

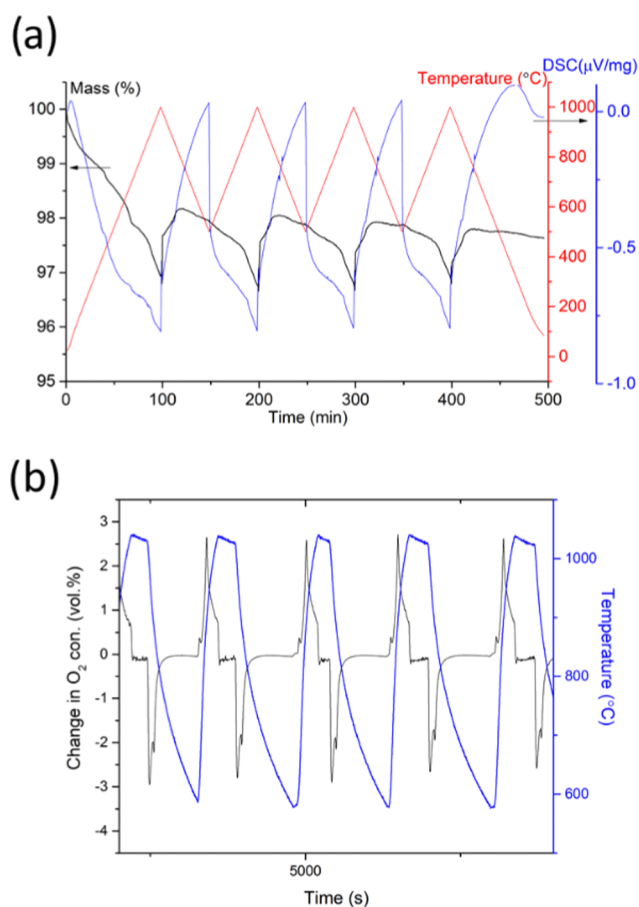
Fe-doped CaMnO<sub>3</sub> samples showed higher redox peak temperatures in fluidized bed experiments than in thermal analysis. This result was not surprising considering the slow reduction kinetics of Fe oxides. In addition, pO<sub>2</sub> was set to 0.21 bar in the fluidized bed, which makes the reduction more difficult compared to lower pO<sub>2</sub> values since Fe oxides do not have oxygen release ability under these conditions. The oxidation peak temperature was recorded as 840 and 830 °C for CaMn<sub>0.9</sub>Fe<sub>0.1</sub>O<sub>3</sub> and CaMn<sub>0.8</sub>Fe<sub>0.2</sub>O<sub>3</sub>, respectively, while the reduction peak temperature was 980 °C for both. In fluidized bed tests, CaMn<sub>0.9</sub>Fe<sub>0.1</sub>O<sub>3</sub> showed better performance than CaMn<sub>0.8</sub>Fe<sub>0.2</sub>O<sub>3</sub> with regard to the amount of O<sub>2</sub> released/consumed. Therefore, Fe-doped calcium manganese oxides can be easily put forward to enhance the redox properties of CaMnO<sub>3</sub> systems.

It is known that B-site dopant will have a higher impact than the A-site dopant on vacancy creation, which is related to the oxygen release ability of the compound.<sup>36</sup> However, the effect of MgO addition to the B site was tested for a CLOU application.<sup>32,40</sup> A study revealed that MgO-doped CaMnO<sub>3</sub> showed good stability in a 50-cycle test and a small increase in reactivity was also observed.<sup>32</sup> Both in thermal analysis and fluidized bed experiments (Figures 4 and 5 for CMM1 and



**Figure 3.** Thermal behavior of CMF2 under controlled atmosphere (a) obtained by STA and O<sub>2</sub> release/consumption ability of CMF2 in the fluidized bed test (b).

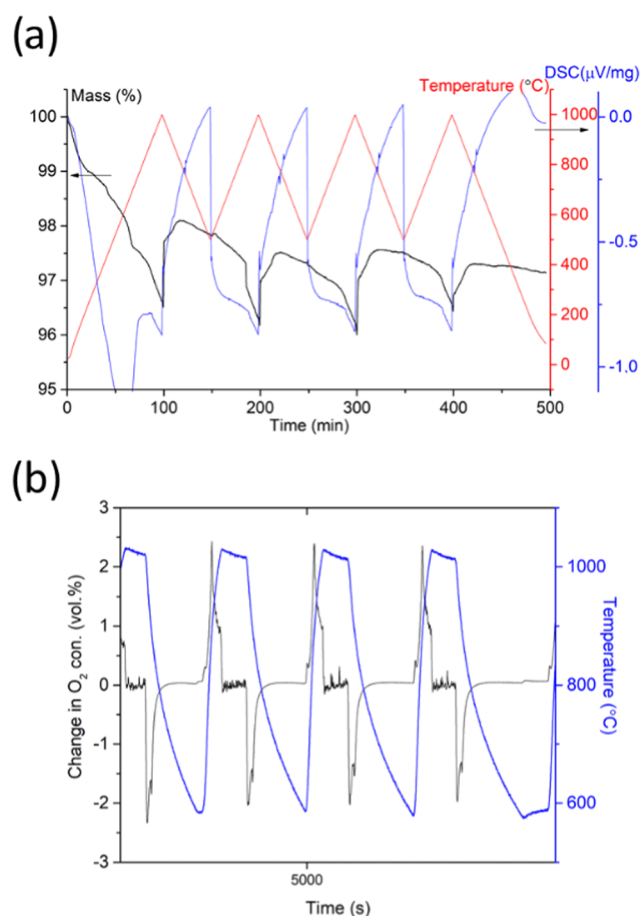
CMM2, respectively), the reduction peak temperature was observed around 975–985 °C for both MgO-doped CaMnO<sub>3</sub>. However, the oxidation peak temperature was obtained as 750 °C in thermal analysis, while it was 895 °C in fluidized bed test for CaMn<sub>0.9</sub>Mg<sub>0.1</sub>O<sub>3</sub>. This difference most likely occurred due to the slow reoxidation kinetics of the CaMnO<sub>3</sub>, like the other doped CaMnO<sub>3</sub> samples. This temperature difference between the experiments was also observed for CaMn<sub>0.8</sub>Mg<sub>0.2</sub>O<sub>3</sub>, as expected. The phase analysis of the fresh and used samples of CMM1 and CMM2 revealed that there was free MgO in the system along with the perovskite structure (Figures S4 and S5). Most likely, a part of the perovskite dissociated into MgO and Ca<sub>2</sub>MnO<sub>4</sub> during a previous CLOU application. There was no free MgO in the starting sample before it had been used in a CLOU application. However, as the intensity ratio of the MgO was different in the fresh sample (after CLOU) than in the used sample (after fluidized bed tests), MgO dissociation was thought to have occurred during the experiments. Even though there was only a 25–45 °C temperature difference for the oxidation peak temperature of MgO-doped CaMnO<sub>3</sub> samples and pure CaMnO<sub>3</sub> in the thermal analysis, the mass gain was 6.75 times higher for CaMn<sub>0.9</sub>Mg<sub>0.1</sub>O<sub>3</sub> and 5 times higher for CaMn<sub>0.8</sub>Mg<sub>0.2</sub>O<sub>3</sub> than for pure CaMnO<sub>3</sub>. A similar result was obtained in the fluidized bed tests as well. The average oxygen peak release was recorded as 2.66 vol % for CaMn<sub>0.9</sub>Mg<sub>0.1</sub>O<sub>3</sub> for reduction steps. The peak oxygen consumption was recorded around 2.62 vol % for oxidation steps. Among the MgO-doped samples, the best results were obtained with



**Figure 4.** Thermal behavior of CMM1 under controlled atmosphere (a) obtained by STA and O<sub>2</sub> release/consumption ability of CMM1 in the fluidized bed test (b).

CaMn<sub>0.9</sub>Mg<sub>0.1</sub>O<sub>3</sub>. The different redox characteristics of the MgO-doped CaMnO<sub>3</sub> samples and the pure CaMnO<sub>3</sub> supported that MgO was still present in the perovskite system despite the observed phase dissociation.

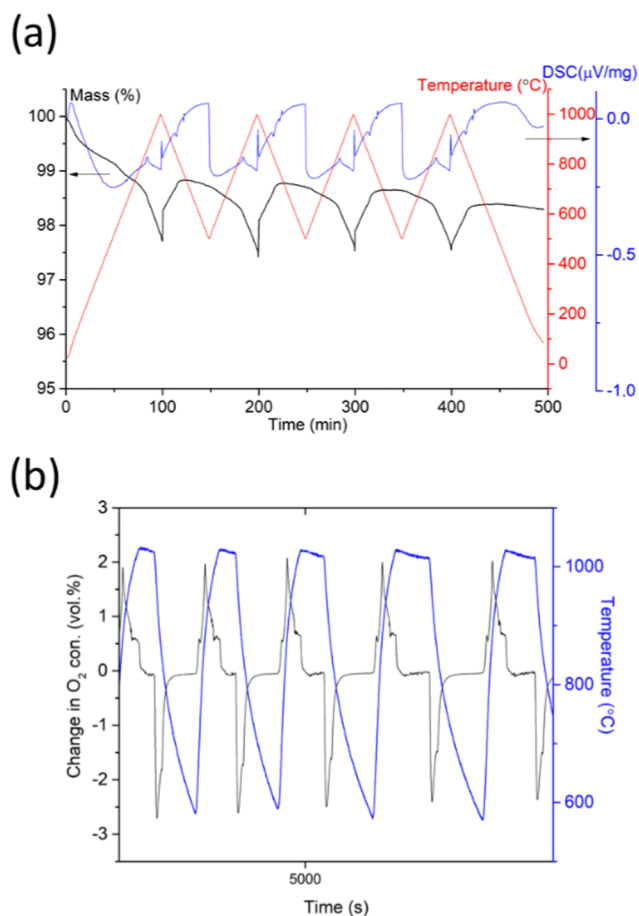
Cu is one of the most well-known materials both for both chemical looping oxygen uncoupling and thermochemical energy storage applications.<sup>14,41,42</sup> However, a support material to combine with CuO is generally proposed due to the low melting point and the relatively high cost of the pure copper compounds.<sup>43</sup> From this point of view, CuO-doped CaMnO<sub>3</sub> samples were also tested in this study. The phase analyses of CMC1 and CMC2 are presented in Figures S6 and S7, respectively. In Figure 6, the simultaneous thermal analysis and fluidized bed test results were given for CaMn<sub>0.9</sub>Cu<sub>0.1</sub>O<sub>3</sub>. The reduction peak temperature was obtained as 975 and 945 °C in the thermal analysis and the fluidized bed test, respectively. There was only a 25 °C difference observed between pure CaMnO<sub>3</sub> and CaMn<sub>0.9</sub>Cu<sub>0.1</sub>O<sub>3</sub> for the reduction peak temperature in the fluidized bed test. This difference most likely occurred due to the CuO addition. This hypothesis was also supported by the results of CaMn<sub>0.2</sub>Cu<sub>0.2</sub>O<sub>3</sub> (Figure 7). The difference for the reduction peak temperature was obtained to be around 20 °C between pure CaMnO<sub>3</sub> and CaMn<sub>0.8</sub>Cu<sub>0.2</sub>O<sub>3</sub>. The oxidation peak temperatures were observed to be 730 and 740 °C, while it was found to be 795 °C for pure CaMnO<sub>3</sub>. The difference for the recorded oxidation peak temperature was smaller in the fluidized bed tests between CuO-doped CaMnO<sub>3</sub> (870 and 860 °C, respectively) and pure CaMnO<sub>3</sub>.



**Figure 5.** Thermal behavior of CMM2 under controlled atmosphere (a) obtained by STA and O<sub>2</sub> release/consumption ability of CMM2 in the fluidized bed test (b).

Copper affected the mass gain/mass loss positively in the thermal analysis. During the oxidation cycles, 0.75 wt % mass gain was observed for CaMn<sub>0.9</sub>Cu<sub>0.1</sub>O<sub>3</sub>. For CaMn<sub>0.2</sub>Cu<sub>0.2</sub>O<sub>3</sub>, it was even higher and recorded as 0.9 wt %. During the reduction cycles, the mass loss was obtained as 1.15 and 1.2 wt % for CaMn<sub>0.9</sub>Cu<sub>0.1</sub>O<sub>3</sub> and CaMn<sub>0.2</sub>Cu<sub>0.2</sub>O<sub>3</sub>, respectively. Fluidized bed tests also showed a similar result to thermal analysis. During the oxidation cycles, average O<sub>2</sub> consumption by CaMn<sub>0.9</sub>Cu<sub>0.1</sub>O<sub>3</sub> and CaMn<sub>0.2</sub>Cu<sub>0.2</sub>O<sub>3</sub> was 0.29 and 0.34 g, respectively. During the reduction cycles, 0.35 g of O<sub>2</sub> was released by CaMn<sub>0.9</sub>Cu<sub>0.1</sub>O<sub>3</sub> and the value was 0.41 g for CaMn<sub>0.2</sub>Cu<sub>0.2</sub>O<sub>3</sub>. Among the tested samples, Cu-doped CaMnO<sub>3</sub> showed one of the best performances together with Mg-doped CaMnO<sub>3</sub>.

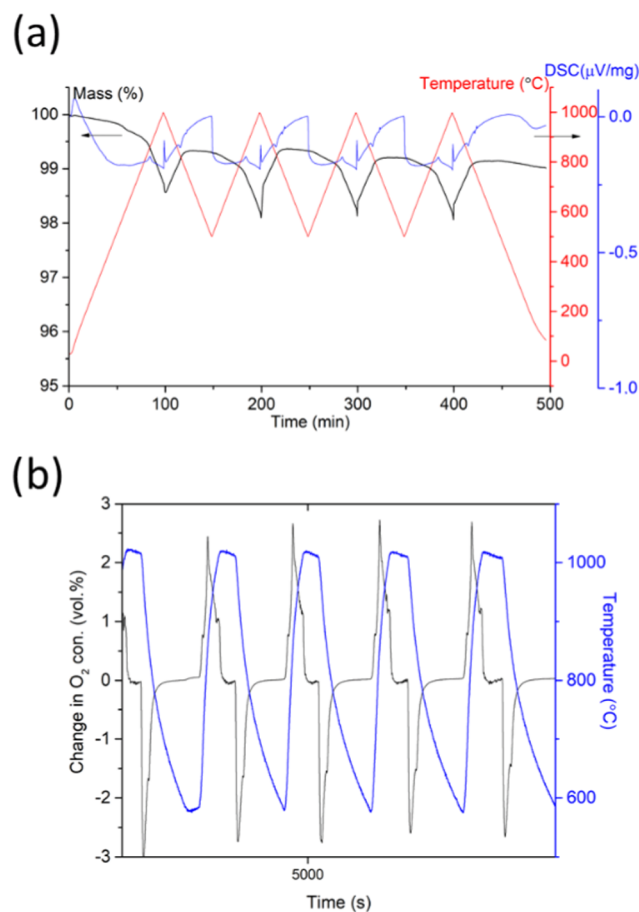
The combined transitional-metal doping to CaMnO<sub>3</sub> has also been studied.<sup>29,30</sup> As both Fe and Cu dopings show better performance than pure CaMnO<sub>3</sub>, the effect of the combined doping of Cu and Fe was tested as well (Figure 8). In addition, a well-studied oxygen carrier CMMT<sup>26,32,33</sup> was also tested after being used in a CLOU application. Despite the individual high performance of Cu- and Fe-doped CaMnO<sub>3</sub> samples, Cu- and Fe-doped CaMnO<sub>3</sub> could not show the same performance in the thermal analysis tests. This most likely occurred due to the phase dissociation after the CLOU operation (Figure S8). However, as the dissociated compounds may also exhibit redox behavior, the performance of the material was still promising. The peak reduction temperature obtained from the thermal



**Figure 6.** Thermal behavior of CMC1 under controlled atmosphere (a) obtained by STA and O<sub>2</sub> release/consumption ability of CMC1 in the fluidized bed test (b).

analysis was similar (980 °C) to CMF and CMC (970–975 °C). However, the oxidation peak temperature of CMCF was recorded as 710 °C, which was similar to that of the CMC samples (730–740 °C) but quite higher than that of the CMF samples (580–610 °C). The mass gain during oxidation and the mass loss during the reduction were recorded as 0.8 and 0.75 wt %, respectively. Even though CMCF did not exhibit the best performance, it was more stable than other samples in a manner of cyclic mass change. For the fluidized bed tests, there was no significant temperature difference observed between the CMCF, CMF, and CMC. However, CMCF also showed a way better performance than pure CM sample. CMC, CMF, and CMCF consumed and released similar amounts of O<sub>2</sub> during the redox cycles in the fluidized bed tests. However, CMCF exhibited slower kinetics, which can be explained by the wider and shorter O<sub>2</sub> consumption/releasing peaks than CMC and CMF samples.

CMMT was mentioned in the literature as a stable and high-performance oxygen carrier for CLOU.<sup>26,32,33</sup> However, CMMT could not show a high TES performance in this study. There was no significant temperature difference between CMMT, CMM, and CM for the redox peak temperatures when comparing the thermal analysis and the fluidized bed tests (Figure 9). The mass loss/gain performance of CMMT was almost 40% lower than that of CMM samples in the thermal analysis. However, CMMT showed 2.5 times higher mass gain/loss than CM. The fluidized bed test performance of

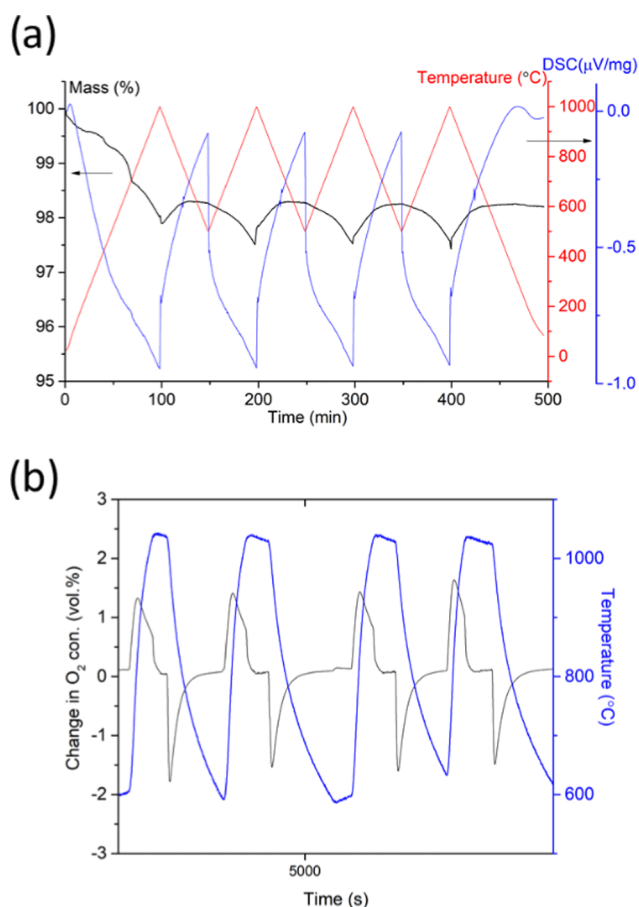


**Figure 7.** Thermal behavior of CMC2 under controlled atmosphere (a) obtained by STA and O<sub>2</sub> release/consumption ability of CMC2 in the fluidized bed test (b).

CMMT was slightly different from that of the thermal analysis. In the fluidized bed tests, CMMT exhibited even lower O<sub>2</sub> uptake/release performance than CM. The decrease in the performance most likely occurred due to a new phase formation, Ti<sub>0.9</sub>Mg<sub>0.1</sub>O<sub>5</sub>, which was detected via phase analysis (Figure S9).

## DISCUSSION

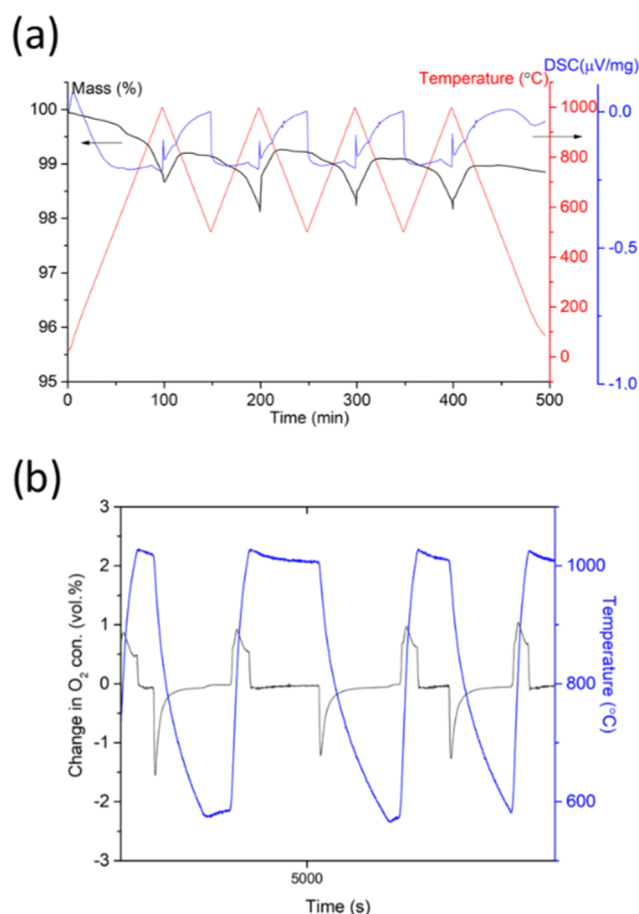
Thermochemical energy storage (TCES) applications are crucial to store energy in a clean way and to conserve the primary energy sources.<sup>44</sup> Therefore, it offers an opportunity to help reduce greenhouse gas emissions.<sup>44</sup> Carbon capture and storage (CCS) technologies are also important to tackle climate change.<sup>45</sup> Recently, materials can be used in both thermochemical energy storage applications and carbon capture and storage applications.<sup>46</sup> For both applications, the candidate materials should be cost-effective, environmentally friendly, and thermochemically stable.<sup>47,48</sup> Since a high amount of materials are needed in industrial applications for both methods, it would be valuable to recycle the “used” materials by another application or to use a low-cost material such as waste or ore. From this point of view, promising doped CaMnO<sub>3</sub>-based oxygen carriers used in chemical looping oxygen uncoupling (CLOU) applications were tested for potential TCES. Calcium manganite-based samples were designed to have a certain composition, which is given in Table 1 when they were synthesized for CLOU applica-



**Figure 8.** Thermal behavior of CMCF under controlled atmosphere (a) obtained by STA and O<sub>2</sub> release/consumption ability of CMCF in the fluidized bed test (b).

tions.<sup>26,29–33</sup> A fresh material that is not used in any CLOU operation would most likely exhibit a higher performance than that obtained in this study, since perovskite-based materials are well known as good candidates for thermochemical energy storage applications.<sup>49</sup> Even though fresh materials were beyond the scope of this study, phase analysis was carried out on both fresh samples and used samples in a CLOU application to reveal the effect of the CLOU operation on the phase structure of the samples. However, no difference was observed regarding phase structure, which was also observed in other related studies.<sup>29,31</sup> Although the metal oxide-based CLOU and TCES applications are similar in the oxidation manner, CLOU has a harsher reduction step, where a fuel is used as the reduction tool. Therefore, we have tested the used material in a CLOU application to mimic the worst-case scenario in a manner of thermochemical stability.

Fluidized bed experiments were used to investigate the oxygen release/consumption behavior of the samples related to temperature changes, whereas thermal analysis combined with differential scanning calorimetry could reveal the possible phase transitions without weight loss/gain in the system. It is important to use a sufficiently large amount of sample in the tests to mimic the real applications. However, the amount of sample used is very limited in the thermal analysis experiments for practical reasons. But thermal analysis results are important to reveal thermodynamic characteristics and to mimic the ideal conditions, where the heating and cooling rates are predefined and strictly controlled. From this point of view, simultaneous



**Figure 9.** Thermal behavior of CMMT under controlled atmosphere (a) obtained by STA and O<sub>2</sub> release/consumption ability of CMMT in the fluidized bed test (b).

thermal analysis and fluidized bed tests were used as complementary techniques in this study.

Decomposition of the particles into a smaller size can be observed in a reduction step, and an oxidation step may cause some agglomeration due to the sintering effect in CLOU applications.<sup>50</sup> Therefore, before testing the particles in fluidized bed experiments or in thermal analysis, the samples were sieved into the 180–250 μm size range to ensure that all particles have the same size range without any fragmented particles. After the fluidized bed experiments, all samples were investigated under a scanning electron microscope and secondary electron images were recorded to investigate the thermomechanical and morphological stabilities of the particles (Figure S10). It is worth noting that the number of redox cycles applied should be increased to have better understanding of the thermomechanical stability of the oxygen carriers. However, the poorest thermomechanical stability after the applied cycles was observed on Fe-doped CaMnO<sub>3</sub> samples, while the Cu- and Fe-doped CaMnO<sub>3</sub> samples showed the highest thermomechanical stability.

Among the samples, the highest O<sub>2</sub> release/consumption in the fluidized bed tests was obtained with the CaMn<sub>0.8</sub>Cu<sub>0.2</sub>O<sub>3</sub> sample. For thermal analysis experiment, CaMn<sub>0.9</sub>Mg<sub>0.1</sub>O<sub>3</sub> showed the best performance for the highest mass change. From DSC curves, the heat storage capacities were calculated for the samples that showed the best performance. The heat storage capacity was calculated as 280 kJ/kg for



CaMn<sub>0.9</sub>Mg<sub>0.1</sub>O<sub>3</sub> and 430 kJ/kg for CaMn<sub>0.8</sub>Cu<sub>0.2</sub>O<sub>3</sub>. The heat storage capacity of pure CaMnO<sub>3</sub> was also calculated (240 kJ/kg) to reveal the difference of CM, CMM1, and CMC2. The obtained heat storage capacity of pure CM was slightly lower than the reported values in the literature.<sup>39</sup> Most likely, this difference occurred due to the formation of CaMn<sub>2</sub>O<sub>4</sub> in the pure CM sample after the CLOU operation.

## CONCLUSIONS

In this work, calcium manganite-based oxides used in chemical looping oxygen uncoupling (CLOU) applications were tested as a potential thermochemical energy storage material. The dopant effects of Fe<sub>2</sub>O<sub>3</sub>, MgO, CuO, Fe<sub>2</sub>O<sub>3</sub>&CuO, and MgO&TiO<sub>2</sub> were also investigated to compare their performance with pure CaMnO<sub>3</sub>. Even though materials may end up with decreased reactivity after being used in CLOU applications, the results showed that promising calcium manganites for CLOU applications showed good performance also for thermochemical energy storage applications. The highest O<sub>2</sub> release/reoxidation performance among the tested samples was exhibited by CaMn<sub>0.8</sub>Cu<sub>0.2</sub>O<sub>3</sub>, and the heat storage capacity was calculated as 430 kJ/kg CaMn<sub>0.8</sub>Cu<sub>0.2</sub>O<sub>3</sub>. However, CaMn<sub>0.9</sub>Mg<sub>0.1</sub>O<sub>3</sub> showed the best cyclic stability, which is also important for materials' lifetime. This study showed that the utilization of recycled CLOU materials in a thermochemical energy storage application may be possible if a chosen oxygen carrier shows high cyclic stability.

## ASSOCIATED CONTENT

### Supporting Information

The Supporting Information is available free of charge at <https://pubs.acs.org/doi/10.1021/acs.iecr.0c05182>.

Detailed characterization of the samples (PDF)

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### Notes

The authors declare no competing financial interest.

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## REFERENCES

- (1) Ervin, G. Solar Heat Storage Using Chemical Reactions. *J. Solid State Chem.* **1977**, *22*, 51–61.
- (2) Wentworth, W. E.; Chen, E. Simple Thermal Decomposition Reactions for Storage of Solar Thermal Energy. *Sol. Energy* **1976**, *18*, 205–214.
- (3) André, L.; Abanades, S.; Flamant, G. Screening of Thermochemical Systems Based on Solid-Gas Reversible Reactions for High Temperature Solar Thermal Energy Storage. *Renewable Sustainable Energy Rev.* **2016**, *64*, 703–715.
- (4) André, L.; Abanades, S.; Cassayre, L. Experimental Investigation of Co-Cu, Mn-Co, and Mn-Cu Redox Materials Applied to Solar Thermochemical Energy Storage. *ACS Appl. Energy Mater.* **2018**, *1*, 3385–3395.
- (5) Michel, B.; Mazet, N.; Mauran, S.; Stitou, D.; Xu, J. Thermochemical Process for Seasonal Storage of Solar Energy: Characterization and Modeling of a High Density Reactive Bed. *Energy* **2012**, *47*, 553–563.
- (6) Pardo, P.; Deydier, A.; Anxionnaz-Minvielle, Z.; Rougé, S.; Cabassud, M.; Cognet, P. A Review on High Temperature Thermochemical Heat Energy Storage. *Renewable Sustainable Energy Rev.* **2014**, *32*, 591–610.
- (7) Block, T.; Schmücker, M. Metal Oxides for Thermochemical Energy Storage: A Comparison of Several Metal Oxide Systems. *Sol. Energy* **2016**, *126*, 195–207.
- (8) Carrillo, A. J.; Serrano, D. P.; Pizarro, P.; Coronado, J. M. Thermochemical Heat Storage Based on the Mn<sub>2</sub>O<sub>3</sub>/Mn<sub>3</sub>O<sub>4</sub> Redox Couple: Influence of the Initial Particle Size on the Morphological Evolution and Cyclability. *J. Mater. Chem. A* **2014**, *2*, 19435–19443.
- (9) Carrillo, A. J.; Serrano, D. P.; Pizarro, P.; Coronado, J. M. Improving the Thermochemical Energy Storage Performance of the Mn<sub>2</sub>O<sub>3</sub>/Mn<sub>3</sub>O<sub>4</sub> Redox Couple by the Incorporation of Iron. *ChemSusChem* **2015**, *8*, 1947–1954.
- (10) Wokon, M.; Bauer, T.; Linder, M. Investigations on Thermochemical Energy Storage Based on Manganese-Iron Oxide in a Lab-Scale Reactor. *AIP Conf. Proc.* **2017**, *1850*, No. 090008.
- (11) Carrillo, A. J.; Moya, J.; Bayón, A.; Jana, P.; De La Peña O'Shea, V. A.; Romero, M.; Gonzalez-Aguilar, J.; Serrano, D. P.; Pizarro, P.; Coronado, J. M. Thermochemical Energy Storage at High Temperature via Redox Cycles of Mn and Co Oxides: Pure Oxides versus Mixed Ones. *Sol. Energy Mater. Sol. Cells* **2014**, *123*, 47–57.
- (12) Carrillo, A. J.; Serrano, D. P.; Pizarro, P.; Coronado, J. M. Manganese Oxide-Based Thermochemical Energy Storage: Modulating Temperatures of Redox Cycles by Fe-Cu Co-Doping. *J. Energy Storage* **2016**, *5*, 169–176.
- (13) André, L.; Abanades, S.; Cassayre, L. Experimental and Thermodynamic Study of Co-Fe and Mn-Fe Based Mixed Metal Oxides for Thermochemical Energy Storage Application. *AIP Conf. Proc.* **2017**, *1850*, No. 090002.
- (14) André, L.; Abanades, S.; Cassayre, L. Mixed Metal Oxide Systems Applied to Thermochemical Storage of Solar Energy: Benefits of Secondary Metal Addition in Co and Mn Oxides and Contribution of Thermodynamics. *Appl. Sci.* **2018**, *8*, No. 2618.
- (15) Carrillo, A. J.; Serrano, D. P.; Pizarro, P.; Coronado, J. M. Thermochemical Heat Storage at High Temperatures Using Mn<sub>2</sub>O<sub>3</sub>/Mn<sub>3</sub>O<sub>4</sub> System: Narrowing the Redox Hysteresis by Metal Co-Doping. *Energy Procedia* **2015**, *73*, 263–271.
- (16) Block, T.; Schmücker, M. Metal Oxides for Thermochemical Energy Storage: A Comparison of Several Metal Oxide Systems. *Sol. Energy* **2016**, *126*, 195–207.
- (17) Bulfin, B.; Vieten, J.; Starr, D. E.; Azarpira, A.; Zachäus, C.; Hävecker, M.; Skorupska, K.; Schmücker, M.; Roeb, M.; Sattler, C. Redox Chemistry of CaMnO<sub>3</sub> and Ca<sub>0.8</sub>Sr<sub>0.2</sub>MnO<sub>3</sub> Oxygen Storage Perovskites. *J. Mater. Chem. A* **2017**, *5*, 7912–7919.
- (18) Imponenti, L.; Albrecht, K. J.; Kharait, R.; Sanders, M. D.; Jackson, G. S. Redox Cycles with Doped Calcium Manganites for Thermochemical Energy Storage to 1000 °C. *Appl. Energy* **2018**, *230*, 1–18.
- (19) Rørmark, L.; Mørch, A. B.; Wiik, K.; Stølen, S.; Grande, T. Enthalpies of Oxidation of CaMnO<sub>3-δ</sub>, Ca<sub>2</sub>MnO<sub>4-δ</sub> and SrMnO<sub>3-δ</sub> - Deduced Redox Properties. *Chem. Mater.* **2001**, *13*, 4005–4013.

- (20) Yu, N.; Nair, M. M.; Mahinpey, N. Structure and Phase Evolution of CaMnO<sub>3</sub> Perovskite during Isothermal Redox Cycles. *Can. J. Chem. Eng.* **2019**, *97*, 2131–2136.
- (21) Mattisson, T. Materials for Chemical-Looping with Oxygen Uncoupling. *ISRN Chem. Eng.* **2013**, *2013*, 1–19.
- (22) Shulman, A.; Cleverstam, E.; Mattisson, T.; Lyngfelt, A. Chemical - Looping with Oxygen Uncoupling Using Mn/Mg-Based Oxygen Carriers - Oxygen Release and Reactivity with Methane. *Fuel* **2011**, *90*, 941–950.
- (23) Jing, D.; Hermans, E. Y. S. I. M.; Leion, H.; Rydén, M.; Mattisson, T.; Van Noyen, J.; Lyngfelt, A. In *Manganese-Silica Combined Oxides as Oxygen Carrier for Chemical-Looping Combustion*, 2nd International Conference on Chemical Looping, 2012.
- (24) Mohammadpour, N.; Azimi, G.; Leion, H.; Rydén, M.; Mattisson, T.; Lyngfelt, A. Investigation of Manganese-Iron Oxide Materials Based on Manganese Ores as Oxygen Carriers for Chemical Looping with Oxygen Uncoupling (CLOU). *Energy Technol.* **2014**, *2*, 469–479.
- (25) Abad, A.; Gayán, P.; De Diego, L. F.; García-Labiano, F.; Adánez, J. Modelling Chemical-Looping Assisted by Oxygen Uncoupling (CLaOU): Assessment of Natural Gas Combustion with Calcium Manganite as Oxygen Carrier. *Proc. Combust. Inst.* **2019**, *37*, 4361–4369.
- (26) Moldenhauer, P.; Hallberg, P.; Biermann, M.; Snijkers, F.; Albertsen, K.; Mattisson, T.; Lyngfelt, A. Oxygen-Carrier Development of Calcium Manganite-Based Materials with Perovskite Structure for Chemical-Looping Combustion of Methane. *Energy Technol.* **2020**, *8*, No. 2000069.
- (27) Hallberg, P.; Jing, D.; Rydén, M.; Mattisson, T.; Lyngfelt, A. Chemical Looping Combustion and Chemical Looping with Oxygen Uncoupling Experiments in a Batch Reactor Using Spray-Dried CaMn<sub>1-x</sub>Mg<sub>x</sub>O<sub>3-δ</sub> (M = Ti, Fe, Mg) Particles as Oxygen Carriers. *Energy Fuels* **2013**, *27*, 1473–1481.
- (28) Goldyreva, E. I.; Leonidov, I. A.; Patrakeev, M. V.; Kozhevnikov, V. L. Thermodynamics of Oxygen in CaMnO<sub>3-δ</sub>. *J. Solid State Electrochem.* **2013**, *17*, 3185–3190.
- (29) Jing, D.; Snijkers, F.; Hallberg, P.; Leion, H.; Mattisson, T.; Lyngfelt, A. Effect of Production Parameters on the Spray-Dried Calcium Manganite Oxygen Carriers for Chemical-Looping Combustion. *Energy Fuels* **2016**, *30*, 3257–3268.
- (30) Arjmand, M.; Hedayati, A.; Azad, A. M.; Leion, H.; Rydén, M.; Mattisson, T. Ca<sub>x</sub>La<sub>1-x</sub>Mn<sub>1-y</sub>MyO<sub>3-δ</sub> (M = Mg, Ti, Fe, or Cu) as Oxygen Carriers for Chemical-Looping with Oxygen Uncoupling (CLOU). *Energy Fuels* **2013**, *27*, 4097–4107.
- (31) Hallberg, P.; Jing, D.; Rydén, M.; Mattisson, T.; Lyngfelt, A. Chemical Looping Combustion and Chemical Looping with Oxygen Uncoupling Experiments in a Batch Reactor Using Spray-Dried CaMn<sub>1-x</sub>Mg<sub>x</sub>O<sub>3-δ</sub> (M = Ti, Fe, Mg) Particles as Oxygen Carriers. *Energy Fuels* **2013**, *27*, 1473–1481.
- (32) Jing, D.; Mattisson, T.; Leion, H.; Rydén, M.; Lyngfelt, A. Examination of Perovskite Structure CaMnO<sub>3-δ</sub> with MgO Addition as Oxygen Carrier for Chemical Looping with Oxygen Uncoupling Using Methane and Syngas. *Int. J. Chem. Eng.* **2013**, *2013*, No. 679560.
- (33) Hallberg, P.; Källén, M.; Jing, D.; Snijkers, F.; Van Noyen, J.; Rydén, M.; Lyngfelt, A. Experimental Investigation of CaMnO<sub>3-δ</sub> Based Oxygen Carriers Used in Continuous Chemical-Looping Combustion. *Int. J. Chem. Eng.* **2014**, *2014*, No. 412517.
- (34) Rydén, M.; Leion, H.; Mattisson, T.; Lyngfelt, A. Combined Oxides as Oxygen-Carrier Material for Chemical-Looping with Oxygen Uncoupling. *Appl. Energy* **2014**, *113*, 1924–1932.
- (35) Tescari, S.; Agrafiotis, C.; Breuer, S.; De Oliveira, L.; Neises-Von Puttkamer, M.; Roeb, M.; Sattler, C. Thermochemical Solar Energy Storage via Redox Oxides: Materials and Reactor/Heat Exchanger Concepts. *Energy Procedia* **2014**, *49*, 1034–1043.
- (36) Mishra, A.; Li, T.; Li, F.; Santiso, E. E. Oxygen Vacancy Creation Energy in Mn-Containing Perovskites: An Effective Indicator for Chemical Looping with Oxygen Uncoupling. *Chem. Mater.* **2019**, *31*, 689–698.
- (37) Peña, M. A.; Fierro, J. L. G. Chemical Structures and Performance of Perovskite Oxides. *Chem. Rev.* **2001**, *101*, 1981–2017.
- (38) Ji, Q.; Bi, L.; Zhang, J.; Cao, H.; Zhao, X. S. The Role of Oxygen Vacancies of ABO<sub>3</sub> Perovskite Oxides in the Oxygen Reduction Reaction. *Energy Environ. Sci.* **2020**, *13*, 1408–1428.
- (39) Mastronardo, E.; Qian, X.; Coronado, J. M.; Haile, S. Fe-Doped CaMnO<sub>3</sub> for Thermochemical Heat Storage Application. *AIP Conf. Proc.* **2019**, *2126*, No. 210005.
- (40) Keller, M.; Leion, H.; Mattisson, T. Use of CuO/MgAl<sub>2</sub>O<sub>4</sub> and La<sub>0.8</sub>Sr<sub>0.2</sub>FeO<sub>3</sub>/c-Al<sub>2</sub>O<sub>3</sub> in Chemical Looping Reforming System for Tar Removal from Gasification Gas. *AIChE J.* **2016**, *62*, 38–45.
- (41) Alonso, E.; Pérez-Rábago, C.; Licurgo, J.; Fuentealba, E.; Estrada, C. A. First Experimental Studies of Solar Redox Reactions of Copper Oxides for Thermochemical Energy Storage. *Sol. Energy* **2015**, *115*, 297–305.
- (42) Abad, A.; Adánez-Rubio, I.; Gayán, P.; García-Labiano, F.; de Diego, L. F.; Adánez, J. Demonstration of Chemical-Looping with Oxygen Uncoupling (CLOU) Process in a 1.5kW Th Continuously Operating Unit Using a Cu-Based Oxygen-Carrier. *Int. J. Greenhouse Gas Control* **2012**, *6*, 189–200.
- (43) Rydén, M.; Jing, D.; Källén, M.; Leion, H.; Lyngfelt, A.; Mattisson, T. CuO-Based Oxygen-Carrier Particles for Chemical-Looping with Oxygen Uncoupling - Experiments in Batch Reactor and in Continuous Operation. *Ind. Eng. Chem. Res.* **2014**, *53*, 6255–6267.
- (44) Kerskes, H. Thermochemical Energy Storage. *Storing Energy* **2016**, 345–372.
- (45) Tregambi, C.; Montagnaro, F.; Salatino, P.; Solimene, R. Directly Irradiated Fluidized Bed Reactors for Thermochemical Processing and Energy Storage: Application to Calcium Looping. *AIP Conf. Proc.* **2017**, *1850*, No. 090007.
- (46) Humphries, T. D.; Möller, K. T.; Rickard, W. D. A.; Sofianos, M. V.; Liu, S.; Buckley, C. E.; Paskevicius, M. Dolomite: A Low Cost Thermochemical Energy Storage Material. *J. Mater. Chem. A* **2019**, *7*, 1206–1215.
- (47) Imtiaz, Q.; Hosseini, D.; Müller, C. R. Review of Oxygen Carriers for Chemical Looping with Oxygen Uncoupling (CLOU): Thermodynamics, Material Development, and Synthesis. *Energy Technol.* **2013**, *1*, 633–647.
- (48) Wu, S.; Zhou, C.; Doroodchi, E.; Nellore, R.; Moghtaderi, B. A Review on High-Temperature Thermochemical Energy Storage Based on Metal Oxides Redox Cycle. *Energy Convers. Manage.* **2018**, *168*, 421–453.
- (49) André, L.; Abanades, S. Investigation of Metal Oxides, Mixed Oxides, Perovskites and Alkaline Earth Carbonates/Hydroxides as Suitable Candidate Materials for High-Temperature Thermochemical Energy Storage Using Reversible Solid-Gas Reactions. *Mater. Today Energy* **2018**, *10*, 48–61.
- (50) Mattisson, T.; Lyngfelt, A.; Leion, H. Chemical-Looping with Oxygen Uncoupling for Combustion of Solid Fuels. *Int. J. Greenhouse Gas Control* **2009**, *3*, 11–19.