

Physico-Chemical Characteristics of the Products Derived from the Thermolysis of Waste *Abies alba Mill* Wood

Félix Antonio López^{1*}, Olga Rodríguez¹, Andrea Urien¹, Belén Lobato², Teresa Álvarez Centeno² and Francisco José Alguacil¹

¹Centro Nacional de Investigaciones Metalúrgicas, Consejo Superior de Investigaciones Científicas, Madrid, Spain; ²Instituto Nacional del Carbón, Consejo Superior de Investigaciones Científicas, Oviedo, Spain.
Email: flopez@cenim.csic.es

Received Month Day, Year (2012).

ABSTRACT

This paper reports the physico-chemical characteristics of the products derived from the thermolysis (thermolytic distillation) of waste silver fir (*Abies alba Mill.*) wood at different temperatures (400-600°C) in a pilot scale plant. Depending on the thermolysis temperature, the procedure yielded 45-53 wt.% pyrolygneous acid with a high water content (80-86 wt.%) and pH \approx 3.6. The process also produced a carbonaceous solid or biochar (23-26 wt.%), its properties strongly dependent on the thermolysis temperature. Gases (20-31 wt.%) were also produced; these were transformed into electrical energy via a gas turbine. The pyrolygneous acid was centrifuged to isolate a subfraction composed mostly of phenols (phenol, mequinol and furfural) with a total C content of 68-74 wt.%. The remainder was subjected to fractionated distillation at laboratory scale. The distillate was subjected to liquid-liquid extraction using diethyl ether in two stages to obtain a bio-oil composed mainly of acetic acid (\approx 47%), aldehydes, ketones and alcohols (\approx 31%), phenols (\approx 18%) and aliphatic alcohols. The characteristics of the bio-oil depended on the thermolysis temperature.

Keywords: Therrmolysis; *Abies alba Mill.*; Bio-oil; Biochar.

1. Introduction

Lignocellulose biomass is the only renewable source of fixed carbon that can be converted into liquid, solid and gaseous fuels with no contribution to net CO₂ emissions. The pyrolysis of biomass is a promising method of producing solid (charcoal), liquid (tar and other organic compounds) and gaseous products [1-3] that could provide alternative sources of fuel and chemicals [4]. The liquids produced during pyrolysis, widely known as pyrolygneous acids (PAs), are complex mixtures of oxygenated aliphatic and aromatic compounds [5], including CH₃OH, C₂H₅O, C₃H₆O, phenols, and water. CH₃OH arises from the methoxyl groups of uronic acid and from the breakdown of methyl esters and/or ethers generated during the decomposition of pectin-like compounds. C₂H₅O is derived from the acetyl groups of hemicelluloses. Tars contain native resins, intermediate carbohydrates, phenols, aromatic compounds, aldehydes, their condensation products, and other derivatives. Pyrolysis gas is mainly composed of CO₂, CO, CH₄, H₂, C₂H₆, C₂H₄, minor amounts of higher gaseous organics, and

water vapor [6]. The pyrolysis products are mainly generated between 300 and 600°C, their composition being highly dependent on the temperature. The aim of the present work was to study the influence of temperature (400-600°C) on the distribution and physico-chemical characteristics of the products obtained from the thermolysis of waste silver fir (*Abies alba Mill.*) wood. This tree species is one of the most commonly cultivated in Europe.

2. Materials and methods

2.1. Materials

The silver fir wood used in the present work came from the clearing of forests in the Catalan Pyrenees in Spain's northeast. All wood was reduced to a size of 250x100x50 mm before thermolysis. Proximate and ultimate analysis of this waste wood (WW) was performed using a LECO TGA 701 analyser and a LECO CHNS 923 analyser respectively. Its higher heating value (HHV) was estimated using an IKAWEEEME C4000 automatic bomb calorime-

ter. The lignin content was determined using the Tappi method T22 [7]. The cellulose content was determined using the Kushner procedure [8]. **Table 1** shows the main characteristics of this WW; **Table 2** reports its lignocellulose composition.

Table 1. Main characteristics of the waste wood before thermolysis.

| Proximate analysis ^a (%) | | | | Ultimate analysis (%) | | | | | High Heating Value (MJkg ⁻¹) |
|----------------------------------------|-----------------|-----|---------------------------|--------------------------|-----|-----|---|------|---------------------------------------------|
| Moisture content | Volatile matter | Ash | Fixed carbon ^b | C | H | N | S | O | HHV |
| 14.1 | 78.7 | 0.4 | 6.8 | 51.2 | 6.4 | 0.2 | - | 42.2 | 19.9 |

a. Weight percentage (in terms of dry mass); b. By difference

Table 2. Lignocellulose composition of the waste wood.

| Hemicellulose ^a | Cellulose ^a | Lignin ^a | Extractive ^a | Ash ^a |
|----------------------------|------------------------|---------------------|-------------------------|------------------|
| 15.0 | 52.1 | 29.9 | 2.4 | 0.5 |

a. Weight percentage (in terms of dry mass)

2.2. Experiments

Figure 1 shows a diagram of the procedure for the production of biochar and bio-oil. Thermolysis (thermolytic distillation) (1) was performed in a pilot plant making use of GRAUTHERMIC-CSIC technology [9]. The gases produced were condensed using a system of cryogenically cooled condensers, giving rise to a pyroligneous acid (PA) fraction (2).

The non-condensed gases were filtered and the gas led to a TOTEM (Total Energy Module) turbine (3) for the production of electricity. Batches of WW were heated for 4 h (treatment capacity 18 kg h⁻¹) at 400, 500 and 600°C. After completing the thermolytic process, the reactors were cooled for 4 h, opened, and the biochar removed by aspiration. The composition and yield (by weight) of the PA and biochar fractions were then determined; the difference between the weight of the initial WW sample and the combined weights of the PA and biochar is equal to the weight of the gas produced. The PA (2) was then centrifuged, producing a heavier tar fraction (4) and a lighter fraction (5). The latter fraction was separated into its components in a Vitreux fractionated distillation column. The residue left over (7) was added to the tar fraction obtained by centrifugation (4) to produce a combined fraction (8). The distillate from the fractionated distillation process (6) was extracted with diethyl ether (proportion 1:3 by volume) in two stages to obtain an aqueous (9) and an organic fraction (10). The organic fraction was evaporated in a rotary evaporator at

42°C to recover the ether and obtain a final bio-oil (11).

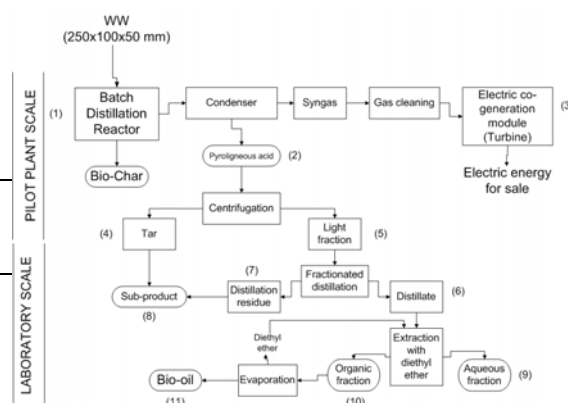


Figure 1. Flow chart showing the thermolytic process.

2.3. Characterization of the pyroligneous acid, bio-char, gases and bio-oil.

The proximate and ultimate compositions of the PAs, bio-chars and bio-oils produced at each thermolysis temperature were determined using an automated LECO CHNS 923 analyser. Their HHV was determined using an IKAWEEEM C4000 automatic bomb calorimeter. The water content of these products was determined by Karl-Fischer titration using Hydranal Composite 5K [Fluka] and employing a Karl-Fischer MKS-520 Moisture Titrator apparatus. The total acid number (TAN) was determined in a KEM Mod. AT-500N automated potentiometric titrator with two electrodes (glass and reference). The titration agent was a standard solution of KOH (F=1). The bio-oils were also analysed by gas chromatography/mass spectroscopy (GC/MS) using a SHIMADZU GCMS-QP2010-Plus apparatus equipped with a 10 m x 0.1 mm x 0.1 µm TRB-5MS TEKNOKROMA capillary column (95% dimethyl-5% diphenyl polysiloxane). Samples of the gases produced at each thermolysis temperature were collected from the uncondensed gas exit tube after having passed through water, activated carbon and lead carbonate solution filters; all samples were captured in TEDLAR® bags. The samples were then analysed by gas chromatography using a HEWLETT-PACKARD 5890 gas chromatograph equipped with a Porapak N and Molecular Sieve multi-column system, a thermal conductivity detector (TCD) and a flame ionisation detector (FID). The porous characteristics of the chars were determined using a Beckman

Coulter SA1100 automatic adsorption analyser. Their surface areas were determined using the BET equation in the p/p_0 range 0.015-0.15 ($r^2 > 0.9999$).

3. Results and discussion

Table 3 shows the product yields from the thermolysis of the WW in relation to the thermolysis temperatures. The biochar yield fell slightly with temperature, the opposite seen for the non-condensable gases. The PA yield represented about 50 wt.% of the thermolytic products, falling slightly with temperature. The light fraction obtained following PA centrifugation (5) was composed mainly of water; this water content increased slightly with the thermolysis temperature (81 wt.% at 400°C and 83 wt.% at 600°C).

The TAN for this light fraction remained practically constant at around 13 g/l. The pH also remained constant at approximately 3.6. Fractionated distillation of the light fraction and subsequent extraction showed water soluble compounds to be the major component (9). The tar (8) and bio-oil (11) yields fell with increasing thermolysis temperature.

Table 3. Effect of the temperature on product distribution.

| Product yield (wt. %) | Thermolysis temperature (°C) | | |
|-----------------------------------|------------------------------|------|------|
| | 400 | 500 | 600 |
| Biochar | 26.4 | 24.5 | 23.6 |
| Non-condensable gases (bio-gases) | 20.8 | 26.5 | 30.9 |
| Pyroligneous acids | 52.8 | 49.1 | 45.5 |
| Tar (sub-product) (8) | 13.7 | 8.8 | 8.3 |
| Water-soluble fraction (9) | 31.8 | 34.0 | 33.0 |
| Bio-oil (11) | 7.3 | 6.3 | 4.2 |

The yield of biochar decreased with increasing thermolysis temperature, due either to the greater primary decomposition of the biomass at higher temperature, or secondary decomposition of the char residue. Such secondary decomposition at higher temperatures may also have given rise to some non-condensable gaseous products that contributed to an increasing gas yield with the thermolysis temperature.

Table 4 summarizes the chemical composition of the combustible gases, their HHVs, and the amount of electricity produced from them with respect to the thermolysis temperature. The gases produced were rich in CH₄, H₂, CO₂ and CO, and contained small quantities of C₂H₄, C₂H₆, C₃H₆, C₃H₈, C₄H₈, nC₄H₁₀ and isoC₄H₁₀.

The percentage volume of CO and CO₂ increased remarkably with temperature, while those of H₂ and CH₄ decreased. The total gas volume (η) also increased with temperature (100 Nm³ t⁻¹ at 400°C and 150 Nm³ t⁻¹ at 600°C). The HHV of the gases fell with the thermolysis temperature, a consequence of a reduction in H₂ and CH₄. However, the electrical energy production (Θ) via the gas turbine increased with temperature (range 434 - 562 kWh•t), a consequence of the relatively greater CO volume. The mean yield of the turbine was 70%.

Table 4. Effect of the thermolysis temperature on gas composition (vol. %) and properties.

| | Thermolysis temperature (°C) | | |
|-------------------------------------------|------------------------------|-------|-------|
| | 400 | 500 | 600 |
| H ₂ | 17.0 | 11.2 | 10.6 |
| CH ₄ | 34.2 | 26.9 | 21.2 |
| CO | 11.7 | 25.9 | 27.0 |
| $\Sigma(C_2-C_4)$ | 5.0 | 7.0 | 5.6 |
| CO ₂ | 16.9 | 25.1 | 28.2 |
| HHV (MJNm ⁻³) | 22.3 | 21.6 | 19.1 |
| η (Nm ³ t ⁻¹) | 100 | 130 | 150 |
| Θ (kWh t ⁻¹) | 434.2 | 551.6 | 562.4 |

C₂ = C₂H₄, C₂H₆; C₃ = C₃H₆, C₃H₈, C₄ = C₄H₈, nC₄H₁₀, isoC₄H₁₀

Table 5 reports the change in the biochar composition with the thermolysis temperature. The volatile material content decreased with increasing temperature, corresponding to an increment in the fixed carbon content. The biochar obtained at 600°C had a high fixed carbon content - 88.9% - this corresponds to a very high C content (92%).

Table 5. Main characteristics of the biochar produced at different thermolysis temperatures.

| | Thermolysis temperature (°C) | | | |
|-------------------------------------|------------------------------|------|------|------|
| | 400 | 500 | 600 | |
| Proximate analysis ^a (%) | Moisture content (%) | 2.5 | 1.2 | 1.3 |
| | Volatile matter | 27.6 | 14.6 | 7.9 |
| | Ash | 1.4 | 1.7 | 1.9 |
| | Fixed carbon ^b | 68.4 | 82.4 | 88.9 |
| Ultimate analysis ^a (%) | C | 80.8 | 88.9 | 92.0 |
| | H | 3.3 | 3.4 | 2.6 |
| | N | 0.5 | 0.5 | 0.6 |
| | S | 0.00 | 0.00 | 0.02 |

| | O^b | 15.4 | 7.2 | 4.8 |
|----------------------|----------------------------|------|------|------|
| Higher Heating Value | HHV (MJ kg ⁻¹) | 30.4 | 34.2 | 34.9 |

a. Weight percentage (in terms of dry mass); b. By difference

The biochars had essentially a macroporous structure (Table 6). The total pore volume increased with the thermolysis temperature from 0.05 cm³g⁻¹ at 400°C to 0.17 cm³g⁻¹ at 600°C. The mean pore size fell from 2 nm at 400°C to 0.82 nm at 600°C; however, the BET area increased from 124 m²g⁻¹ at 400 °C to 405 m²g⁻¹ at 600°C.

Table 6. Surface area and pore volume of biochars.

| | Thermolysis temperature | | |
|------------------------------------------------|-------------------------|------|------|
| | (°C) | | |
| | 400 | 500 | 600 |
| V_T (cm ³ g ⁻¹) | 0.05 | 0.09 | 0.17 |
| V_{micro} (cm ³ g ⁻¹) | 0.05 | 0.09 | 0.16 |
| V_{meso} (cm ³ g ⁻¹) | 0.00 | 0.00 | 0.01 |
| L_0 (nm) | >2 | >2 | 0.82 |
| E_0 (kJmol ⁻¹) | 16.2 | 13.1 | 24.7 |
| S_{micro} (m ² g ⁻¹) | 0 | 0 | 390 |
| S_{ext} (m ² g ⁻¹) | 2 | 2 | 3 |
| S_{BET} (m ² g ⁻¹) | 124 | 199 | 405 |

V_T : total pore volume estimated at $P/P_0 = 0.95$ (cm³g⁻¹); V_{micro} : micropore volume (cm³g⁻¹); V_{meso} : mesopore volume (cm³g⁻¹); L_0 : average width of the micropores (nm); S_{micro} : micropore surface area (m²g⁻¹); S_{ext} : external surface area (m²g⁻¹); S_{BET} : BET surface area determined by N₂ adsorption at 77 K (m²g⁻¹)

Table 7 illustrates the variation in the chemical composition of the heavy hydrocarbon (tar) and bio-oil fractions with respect to the thermolysis temperature. The tar fraction was composed mainly of phenols, acetones, aldehydes (derivatives of ethanone, propanone, pentanone, cyclopentanone and benzofuranone) and alcohols (mequinol, eugenol, glycerine and hydroquinone).

The bio-oil was composed mostly of carboxylic acids (acetic acid ≈ 40% [in terms of the area of the gas chromatography peak], propanoic acid ≈ 4%, and butanoic acid ≈ 2%) derived from the excision of the hemicellulose acetyl groups, aldehydes and ketones (furfural, 2-propanone 1-hydroxy and 2-cyclopenta-1-one) derived from the breakage of the cellulose bonds. Phenols (phenol, 2-methoxy phenol and 1,2 2-methoxy-5-methylphenol) produced by the depolymerisation of lignin and alcohols (mequinol and glycerine) were present in smaller proportions. The quantity of aldehydes and ketones fell with increasing

thermolysis temperature (39% at 400°C compared to 30% at 600°C). The content in carboxylic acids and phenols did not vary significantly, remaining at around 44% and 17% respectively. The alcohol content rose slightly with temperature from 1% at 400°C to 4% at 600°C. The bio-oils obtained at the different temperatures had pH and density values, O and H contents and HHVs similar to those of other pyrolytic bio-oils, though their C content was slightly lower [10].

Table 7. Main characteristics of residual tar and bio-oils.

| | Tar | | | Bio-oil | | | |
|--------------------------------------------|------------------|------|------|---------|------|------|------|
| | Temperature (°C) | 400 | 500 | 600 | 400 | 500 | 600 |
| Ultimate analysis ^a (%) | C | 72.8 | 71.2 | 68.2 | 45.7 | 47.0 | 47.5 |
| | H | 7.7 | 7.9 | 8.9 | 8.8 | 9.0 | 8.5 |
| | N | 0.4 | 0.3 | 0.3 | 0.1 | 0.1 | 0.2 |
| | S | 0.05 | 0.05 | 0.17 | 0.00 | 0.01 | 0.02 |
| | O ^b | 19.0 | 20.6 | 22.3 | 45.3 | 43.9 | 43.8 |
| | H/C Ratio | 1.3 | 1.3 | 1.6 | 2.3 | 2.3 | 2.1 |
| pH | - | - | - | 2.2 | 3.2 | 3.0 | |
| H ₂ O (wt. %) | - | - | - | 10.3 | 5.2 | 4.3 | |
| Higher Heating Value (MJkg ⁻¹) | 31.9 | 31.6 | 30.3 | 24.3 | 30.9 | 27.5 | |

a. Weight percentage (in terms of dry mass); b. By difference

4. Conclusions

This work focuses in the valorisation of waste silver fir wood via a simple thermolysis process. Its success lies mainly in the large amount of energy recovered from the gases generated. The biochars produced show great aptitude for CO₂ capture, and could be used as activated compounds in animal feed for the removal of toxins from the intestinal tract. However, unlike that seen with other pyrolytic techniques, the present procedure produces large quantities of pyroligneous acids, the transformation of which into bio-oils is poor in terms of yield. Nonetheless, the properties of the bio-oils thus obtained are similar to those obtained by other pyrolytic techniques.

Acknowledgements

Dr. O. Rodríguez acknowledged for the contract JAE-Doc_09-00121 (CSIC), Co-funded under the FSE Operational Programme 2007-2013 Multiregional Adaptability and Employment Operational Programme. The experimental work undertaken was performed with the financial support of ENRECO 2000 Ltd.

REFERENCES

- [1] M. Amutio, G. Lopez, R. Aguado, M. Artetxe, J. Bilbao and M. Olazar, "Effect of Vacuum on Ligno-cellulosic Biomass Flash Pyrolysis in a Conical Spouted Bed Reactor", *Energy & Fuels*, Vol. 25, No. 9, 2011, pp. 3950-3960.
doi:10.1021/ef200712h
- [2] A. V. Bridgwater, "Review of fast pyrolysis of biomass and product upgrading", *Biomass and Bioenergy*, Vol. 38, No. 0, 3// 2012, pp. 68-94.
doi:10.1016/j.biombioe.2011.01.048
- [3] M. Balat, "An Overview of the Properties and Applications of Biomass Pyrolysis Oils", *Energy Sources Part A-Recovery Utilization and Environmental Effects*, Vol. 33, No. 7, 2011 2011, pp. 674-689.
doi:10.1080/15567030903228914
- [4] A. Demirbas, "Recovery of oily products from organic fraction of black liquor via pyrolysis", *Energy Sources Part A-Recovery Utilization and Environmental Effects*, Vol. 30, No. 20, 2008, pp. 849-855.
doi:10.1080/15567030701457368
- [5] N. a. T. K. Prakash, "Advances in modeling and simulation of biomass pyrolysis", *Asian Journal of Scientific Research*, Vol. 2, 2009, pp. 1-27.
doi:10.3923/ajsr.2009.1.27
- [6] P. T. Williams and A. R. Reed, "Pre-formed activated carbon matting derived from the pyrolysis of biomass natural fibre textile waste", *Journal of Analytical and Applied Pyrolysis*, Vol. 70, No. 2, 2003, pp. 563-577.
doi:10.1016/s0165-2370(03)00026-3
- [7] C. W. Dence. "The determination of lignin." In: Dence CW, ed. *Methods in Lignin Chemistry*. Berlin: Springer-Verlag 1992, pp. 33-40.
- [8] A. V. Obolenskaya, Z. P. Elnitskaya and A. A. Leonovitch. "Laboratory Works in Wood Cellulose Chemistry". Moscow: Ecologya; 1991, pp. 176-179.
- [9] F. A. Lopez, T. A. Centeno, F. J. Alguacil and B. Lobato, "Distillation of granulated scrap tires in a pilot plant", *Journal of Hazardous Materials*, Vol. 190, No. 1-3, 2011, pp. 285-292.
doi:10.1016/j.jhazmat.2011.03.039
- [10] Q. Zhang, J. Chang, T. Wang and Y. Xu, "Review of biomass pyrolysis oil properties and upgrading research", *Energy Conversion and Management*, Vol. 48, No. 1, 2007, pp. 87-92.
doi:10.1016/j.enconman.2006.05.010