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## Synthesis and Molecular Properties of Some Imine Complexes with Iron Supported by Phosphine Ligands

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### Research Article

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#### ABSTRACT

The compounds which were synthesized and presented in this project serve as stable models of reactive intermediates, proposed in catalytic transformations with the ruthenium counterparts. It is also of general interest to study the reactivity of trimethylphosphine stabilized iron compounds toward aromatic imines and ketones. Other than some cyclomanganated products, there are no examples of cyclometalated products of 3d row transition metals with imines. In this section, a review of reactions in the activation and functionalization of C-H bonds by solution-phase transition metal-based systems are presented, with an emphasis on the activation of aromatic C-H bonds. Phenyl ketimines react smoothly under mild conditions with low-valent iron (0) adducts to form five-membered metallacycles. The reaction of Diphenylketimine and *tert*-butylphenylketimine with  $\text{Fe}(\text{PMe}_3)_4$  generates a hydrido-iron (II) complex 1,2 respectively. Iodomethane reactions of hydrido-iron compound form iodo-iron (II) complex 3. Benzylic imine react with  $\text{Fe}(\text{CH}_3)_2(\text{PMe}_3)_4$  by elimination of methane to afford hexacoordinate methyl-iron(II) complex 4. All complexes are given from 1- 4 were characterized through IR, NMR and X-ray diffraction was discussed.

#### INTRODUCTION

Organometallic chemistry lies at the interface between classical organic and inorganic chemistry because it looks at the interaction between inorganic metal ions and organic molecules. The field has provided some powerful new synthetic methods in organic chemistry and is beginning to make links with biochemistry with the discovery of several metallo-enzymes that involve organometallic intermediates [1,2,3,4]. Organometallic ideas have been useful in interpreting the chemistry of metal surfaces and of metal colloids. The controlled pyrolysis of organometallic species has proved to be a useful way of preparing solid-state material with unusual properties [5,6]. Public concern for the environment has led to the rise of "green chemistry", the purpose of which is to minimize the production of chemical waste in industry and commerce. One way to do this is to use catalysts rather than stoichiometric reagents to bring about reactions. Many commercially important processes that rely on transition metal organometallic complexes as catalysts have been developed, and such applications are likely to gain more importance in the future [7]. The carbon - hydrogen bond of alkanes cannot usually be regarded as functional group. Its unique position in organic chemistry is well illustrated by the standard representation of organic molecules: the presence of C-H bonds is indicated simply by the absence of any other bond. This "invisibility" of C-H bonds reflects both their ubiquitous nature and their lack of reactivity. With these characteristics in mind it is clear that if the ability to selectively functionalize C-H bonds were well developed, it could potentially constitute the most broadly applicable and powerful class of transformations in organic synthesis. Realization of such potential could revolutionize the synthesis of organic molecules ranging in complexity from methanol to the most elaborate natural or unnatural products. The activation of C-H bonds by "organometallic routes" (i.e., those involving the formation of a bond between carbon and a metal center) is a very large, diverse and highly active field [8,9,10,11,12,13,14]. The first reported of "C-H activation" by a transition metal complex is often attributed to *Chatt*. [15] C-H bond activations are often classified as proceeding via either *nucleophilic* (oxidative addition) or *electrophilic* modes. However, these two classes of C-H activation have much in common, perhaps even more so than the corresponding modes of  $\text{H}_2$  activation. Both reactions appear to proceed through a  $\sigma$ -bond intermediate [16,17,18]. Even more striking, putatively *electrophilic* activations, in most cases, proceed via complete oxidative addition (followed by deprotonation of the resulting metal hydride) [19].

## EXPERIMENTAL

## General procedures and materials

All syntheses and manipulations were carried out under an inert atmosphere of argon or nitrogen by using modified Schlenk techniques [20]. This type of apparatus offers an opportunity of carrying out a complete synthesis in a closed system, and in one run. Synthesis, transport and storage of chemicals were done under an atmosphere of purified Argon (BTS catalyst) [21]. Solvents (THF, diethyl ether, pentane) were dried according to known procedures and were freshly distilled prior to use. All reagents (Aldrich, Acros, Fluka, or Lancaster) were used as purchased without further purification.

Trimethylphosphine  $\text{PMe}_3$ [22],  $\text{Fe}(\text{PMe}_3)_4$ [23],  $\text{FeMe}_2(\text{PMe}_3)_4$ [24], (*tert*- Butylphenylketimine and Diphenylketimine [25,26], and were synthesized according to literature procedures. C, H, P, N analyses of air sensitive solids were carried by H. Kolbe micro-analytical laboratory, Mulheim /Ruhr. Infrared spectra (4000-400  $\text{cm}^{-1}$ ), as obtained from Nujol mulls between KBr discs, were recorded on a Bruker FRA106 spectrometer.  $^1\text{H}$ ,  $^{13}\text{C}$  and  $^{31}\text{P}$  NMR spectra were obtained from Bruker AVANCE 500, ARX 300 and AM 200 spectrometers.  $^{13}\text{C}$  and  $^{31}\text{P}$  NMR resonances were obtained with broad-band proton decoupling. Assignment of  $^{13}\text{C}$  signals was supported by DEPT trace. Melting points were measured in capillaries sealed under argon atmosphere.

*tert*- Butylphenylketimine

To a stirred THF solution of 5.2 g (50.4 mmol) of benzonitrile and 31.5 ml (50.4 mmol, 1.6 M in THF) of *tert*-butyl magnesium chloride, 250 mg (1.31 mmol) of copper (I) iodide at  $-78^\circ\text{C}$  were added. The reaction mixture turned dark brown immediately. Progress of the reaction was monitored by thin layer chromatography. After 48 h, the reaction was completed, and the mixture was then filtered off from magnesium chloride followed by evaporation of the solvent in vacuo. The residue was distilled under reduced pressure to afford 4.2 g of an oily liquid (52%, boiling at  $84^\circ\text{C}$ ).  $^1\text{H}$  NMR (300 MHz,  $[\text{D}_8]\text{THF}$ , 297 K):  $\delta$  = 1.22(s, 9H, C(CH<sub>3</sub>)<sub>3</sub>); 7.18 – 7.24 (m, 2H, Ar-H); 7.32-7.36 (m, 3H, Ar-H); 9.43. (s, CH); 128.1 (s, CH); 142.9 (s, C); 180.4 (s, NC).

Hydrido[2-[(imino- $\kappa$ N)phenylmethyl]phenyl- $\kappa$ C]tris(trimethylphosphine)iron(II)(1)

Diphenylketimine (0, 72 g, 3.99 mmol) in 20 ml of pentane were combined with  $\text{Fe}(\text{PMe}_3)_4$  (1.44 g, 3.99 mmol) at  $-27^\circ\text{C}$  to afford violet crystals (1). Yield 1.451g(78%); M.P 105 – 107  $^\circ\text{C}$ . (dec.). IR (Nujol)  $\nu$  3281 w ( $\nu$  H-N=C); 1730vs ( $\nu$  Fe-H).  $^1\text{H}$  NMR (500 MHz,  $[\text{D}_8]\text{THF}$ , 300 K):  $\delta$  = - 16.4 (dt,  $^2J_{\text{P,H}}$  = 22.8 Hz,  $^2J_{\text{P,H}}$  = 81.3 Hz, 1H, Fe-H); 0.94 (s, 18 H, PCH<sub>3</sub>); 1.42 (d,  $^2J_{\text{P,H}}$  = 4.7 Hz, 9H, PCH<sub>3</sub>); 6.72 (t,  $^3J_{\text{H,H}}$  = 7.0 Hz, 1H, Ar-H); 6.82 (t,  $^3J_{\text{H,H}}$  = 7.1 Hz, 1H, Ar-H); 7.38 (t,  $^3J_{\text{H,H}}$  = 7.2 Hz, 2H, Ar-H); 7.47 (t,  $^3J_{\text{H,H}}$  = 7.8 Hz, 3H, Ar-H); 7.89 (d,  $^3J_{\text{H,H}}$  = 6.4 Hz, 1H, Ar-H); 9.40 (s, 1H, N-H).  $^{31}\text{P}$  NMR (202 MHz,  $[\text{D}_8]\text{THF}$ , 296 K):  $\delta$  = 18.2 (d,  $^2J_{\text{P,P}}$  = 37.4 Hz, 2P, PCH<sub>3</sub>), 22.5 (t,  $^2J_{\text{P,P}}$  = 37.4 Hz, 1P, PCH<sub>3</sub>). C<sub>22</sub>H<sub>38</sub>FeNP<sub>3</sub> (465.29): Calc.: C, 56.79; H, 8.23; N 3.01; P 19.97 Found: C, 57.06; H, 7.58; N 2.96; P 20.09 %.

Hydrido[2-[(imino- $\kappa$ N)phenylmethyl]-1,1-dimethylethyl- $\kappa$ C]tris(trimethylphosphine)iron(II) (2)

1.27g ( 3.53 mmol) of  $\text{Fe}(\text{PMe}_3)_4$  in pentane were combined with 0.57g (3.53 mmol) of *tert*-butylphenylketimine to afford reddish violet crystals (2) at  $4^\circ\text{C}$ . Yield 0.753g (48%); M.P 118 - 120  $^\circ\text{C}$ . (dec.). IR (Nujol)  $\nu$  3335 m ( $\nu$  H-N=C); 3060 m, 3026 m ( $\nu$  H-C=C); 1795 m ( $\nu$  Fe-H); 1565 m ( $\nu$  C=C); 1433 s ( $\delta_{\text{as}}$  PCH<sub>3</sub>); 1389 s; 1357 m; 1293 m, 1270 s ( $\delta_{\text{s}}$  PCH<sub>3</sub>); 770 w, 730 s ( $\nu$  C-H<sub>arom</sub>); 940 vs ( $\rho_1$ PCH<sub>3</sub>).  $^1\text{H}$  NMR (500 MHz,  $[\text{D}_8]\text{THF}$ , 300 K):  $\delta$  = - 17.5 (dt,  $^2J_{\text{P,H}}$  = 21.8 Hz,  $^2J_{\text{P,H}}$  = 80.0 Hz, 1 H, Fe-H); 0.88 (s, 18 H; PCH<sub>3</sub>), 1.40 (d,  $^2J_{\text{P,H}}$  = 4.5 Hz, 9H, PCH<sub>3</sub>); 1.46 (s, 9H; C(CH<sub>3</sub>)<sub>3</sub>); 6.71 (dd,  $^3J_{\text{H,H}}$  = 6.5 Hz,  $^3J_{\text{H,H}}$  = 7.6 Hz, 2H, Ar-H); 7.74 (d,  $^3J_{\text{H,H}}$  = 7.4 Hz, 1H, Ar-H); 7.85 (d,  $^3J_{\text{H,H}}$  = 6.5 Hz, 1H; Ar-H); 9.13 (s, 1H, N-H).  $^{13}\text{C}$  NMR ( 125 MHz,  $[\text{D}_8]\text{THF}$ , 300 K):  $\delta$  = 22.3 (tt,  $^3J_{\text{P,C}}$  = 2.6 Hz,  $^1J_{\text{P,C}}$  = 12.8 Hz, PCH<sub>3</sub>); 23.3 (td,  $^1J_{\text{P,C}}$  = 12.8 Hz,  $^3J_{\text{P,C}}$  = 2.6 Hz, PCH<sub>3</sub>); 29.8 (s, C(CH<sub>3</sub>)<sub>3</sub>); 40.1 (s, C(CH<sub>3</sub>)<sub>3</sub>); 117.6 (s, CH); 186.1 (m, C=N), 200.1 (m, Fe-C).  $^{31}\text{P}$  NMR (202 MHz,  $[\text{D}_8]\text{THF}$ , 300 K):  $\delta$  18.2 (d,  $^2J_{\text{P,P}}$  = 36.5 Hz, 2P, PCH<sub>3</sub>); 23.7 (t,  $^2J_{\text{P,P}}$  = 36.5 Hz, 1P, PCH<sub>3</sub>). C<sub>20</sub>H<sub>42</sub>FeNP<sub>3</sub> (445.31): Calc.: C, 53.94; H, 9.51; N 3.15; P 20.87 Found: C, 53.89; H, 9.35; N 3.61; P 19.03 %

Iodo [2-[(imino- $\kappa$ N) phenyl methyl]-2-methylphenyl- $\kappa$ C] tris-(trimethylphosphine) iron (II) (3)

A sample containing 0.598g (1.285mmol) of compound (1) were combined with  $\text{CH}_3\text{I}$  0.227 g (1.00 ml, 1.285 mmol) in 50 mL of THF after 18 h formed a pink solution. The volatiles were removed in vacuo, and the resulting solid was extracted with pentane. Crystallization at  $-27^\circ\text{C}$  afforded dark pink cubes: yield 0.635 g of 3 (77%); M.P < 143  $^\circ\text{C}$ . (dec.). IR (Nujol)  $\nu$  3320 w ( $\nu$  H-N=C); 3070 vw, 1565 m ( $\nu$  H-C=C); 1556 w ( $\nu$  C=C); 1525 vw ( $\nu$  C=N); 941 vs ( $\rho_1$  PCH<sub>3</sub>); 738 s ( $\nu$  C-H<sub>arom</sub>); 716 s, ( $\nu_{\text{as}}$ PC<sub>3</sub>); 660 m ( $\nu_{\text{s}}$ PC<sub>3</sub>).  $^1\text{H}$  NMR (500 MHz,  $[\text{D}_8]\text{THF}$ , 300 K):  $\delta$  = 1.06 (s, 18H, PCH<sub>3</sub>), 1.61 (d,  $^2J_{\text{P,H}}$  = 5.9 Hz, PCH<sub>3</sub>); 6.69 (m, 2H, Ar-H); 7.51 (d,  $^4J_{\text{P,H}}$  = 7.4 Hz, 1H, Ar-H); 7.76 (d,  $^3J_{\text{H,H}}$  = 6.7 Hz, 1H, Ar-H); 9.50 (s, 1H, N-H).  $^{31}\text{P}$  NMR (202 MHz,  $[\text{D}_8]\text{THF}$ , 300 K):  $\delta$  13.16 (

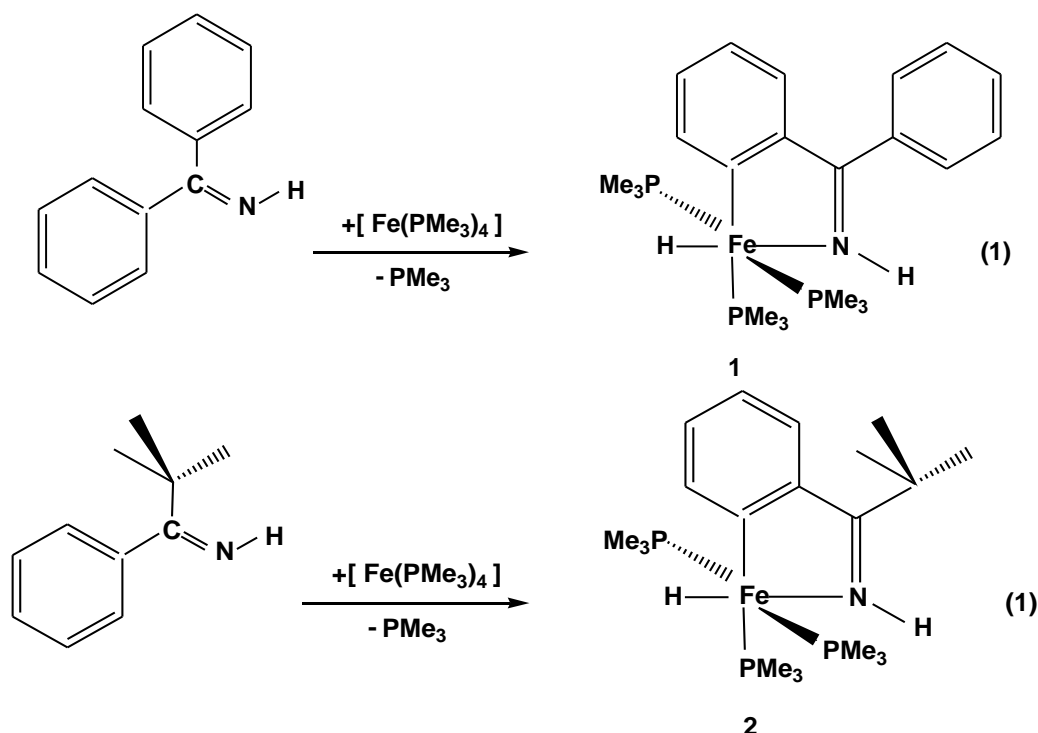
d,  $^2J_{P,P} = 55.9$  Hz, 2P, PCH<sub>3</sub>); 19.9 (t,  $^2J_{P,P} = 55.9$  Hz, 1P, PCH<sub>3</sub>). C<sub>22</sub>H<sub>37</sub>FeNP<sub>3</sub> (591.19): Calc.: C, 44.69; H, 6.31; N 2.37; P 15.72 Found: C, 44.39; H, 5.95; N 2.21; P 15.03 %.

### Methyl [2-[(methyl imino-κN) methylphenyl-κC] tris-(trimethylphosphine)] iron (II) (4)

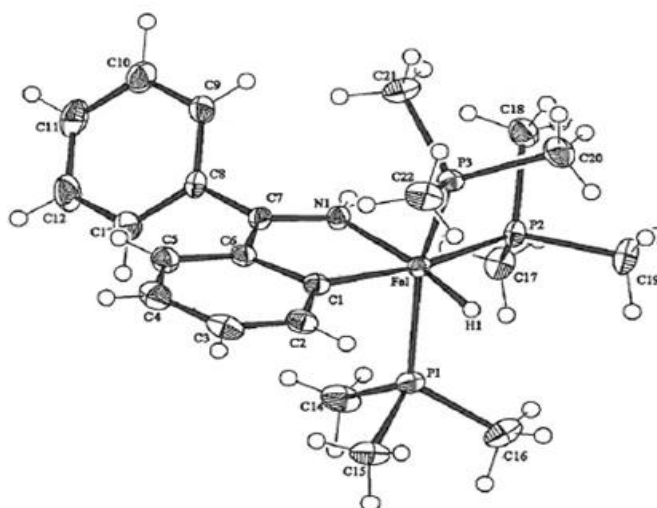
1.0 g (2.56 mmol) of Fe(CH<sub>3</sub>)<sub>2</sub>(PMe<sub>3</sub>)<sub>4</sub> in THF (20 ml) were combined with 0.31 g (2.56 mmol) of *N*-methylbenzylideneimine in diethyl ether (50 ml) at -80 °C. The mixture was warmed to 20 °C and stirred for 18 h. During this period the solution turned red violet. The volatiles were removed in vacuo and the residue was extracted with pentane. Crystallization at 4 °C, afforded violet crystals (4) Yield 753 mg (65%); M.P. 115 - 117 °C (dec.). IR (Nujol):  $\nu = 3053$  vw, 3034 vw ( $\nu$  H-C=C); 1572 m ( $\nu$  C=C); 1521 m ( $\nu$  C=N); 1416 w ( $\delta_{as}$  PCH<sub>3</sub>); 1294 m, 1276 s ( $\delta_s$  PCH<sub>3</sub>); 1221 w; 1180 w ( $\delta_s$  Fe-CH<sub>3</sub>); 1118 w; 1020 w; 996 w; 937 vs ( $\rho_1$  PCH<sub>3</sub>); 844 m ( $\rho_2$  PCH<sub>3</sub>); 746 w, 730 m ( $\gamma$  C-H<sub>arom</sub>); 712 vw, 702 w ( $\nu_{as}$  PC<sub>3</sub>); 677 w, 658 w ( $\nu_s$  PC<sub>3</sub>). <sup>1</sup>H NMR (500 MHz, [D<sub>8</sub>] THF, 298 K):  $\delta = -1.38$  (s, 3H, Fe-CH<sub>3</sub>); 0.71(s, 18H, PCH<sub>3</sub>); 1.49 (s, 9H, PCH<sub>3</sub>); 3.65 (s, 3H, N-CH<sub>3</sub>); 6.74 (s, 1H, Ar-H); 6.85(s, 1H, Ar-H); 7.30 (s, 1H, Ar-H); 7.87(s, 1H, Ar-H); 8.48 (s, 1H, N=C-H). C<sub>18</sub>H<sub>38</sub>FeNP<sub>3</sub> (417.25): Calc.: C, 51.81; H, 9.18; N 3.36; P 22.27 Found: C, 51.52; H, 9.38; N 3.37; P 22.27 %.

## RESULTS AND DISCUSSION

Reacting pentane solutions of Fe (PMe<sub>3</sub>)<sub>4</sub> with diphenyl - and *tert*-butylphenylketimine (Eq.(1)) gives from which upon cooling dark violet prisms of crystal 1 and 2. The yields are 78% for 1 and 48% for 2 respectively. Compound 2 is moderately sensitive while compound 1 is extremely hygroscopic. Less than 1 bar argon decomposes between 105 - 118 °C.

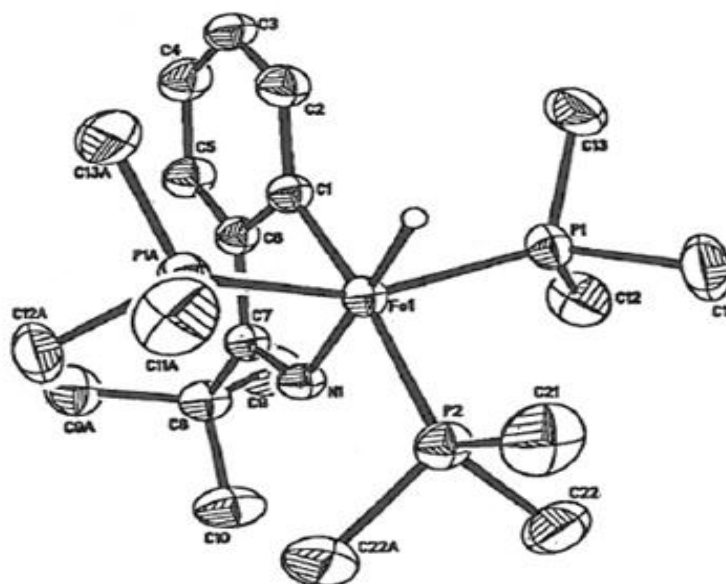


The FeH function is recognized in the infrared spectra by conspicuous  $\nu(\text{FeH})$  bands (1: 1730 cm<sup>-1</sup>, 2:1795cm<sup>-1</sup>) and  $\nu(\text{N-H})$  bands are observed with hypsochromic shifts between 45-70cm<sup>-1</sup> indicating coordination through the N atom. Both spectra show all characteristic bands of coordinated PMe<sub>3</sub> and the coordinated aromatic backbone, which in the <sup>1</sup>H NMR spectra, Fe-H resonance appear at -16.4 for 1 and -17.5 for 2 as doublets of triplets due to the coupling of hydride nucleus with *trans* and *cis* disposed trimethylphosphine. Both compounds show two sets of environmentally different PMe<sub>3</sub> groups which indicate that the iron centers have an octahedral coordination. The <sup>31</sup>P NMR spectra of the two compounds, which are temperature invariant, show coupling of 37 Hz for two sets of PMe<sub>3</sub> groups. The PMe<sub>3</sub> group *trans* to the metalated carbon gives a triplet with a resonance at around 23 ppm and the two isochronic *trans* phosphines appears at 18 ppm as doublets in both cases. The molecular structure of 1 (Fig. 1) shows a slightly distorted octahedral fram of donor atoms centered by an iron atom that bears three meridional P donor groups. The angle P3-Fe1-P1 = 152.878 Å shows a significant deviation from 180°. The Fe-P distances are in the typical range of 2.18-2.22 Å.[131] Fe1-P2 distance is lengthened due to *trans* influence of the carbon donor.N1-Fe1 and C1-Fe1 distance s are also in typical ranges [27]. The Fe1-H1 1.480(19) Å lies in the expected range of a coordinated H to an Fe atom [28,29].



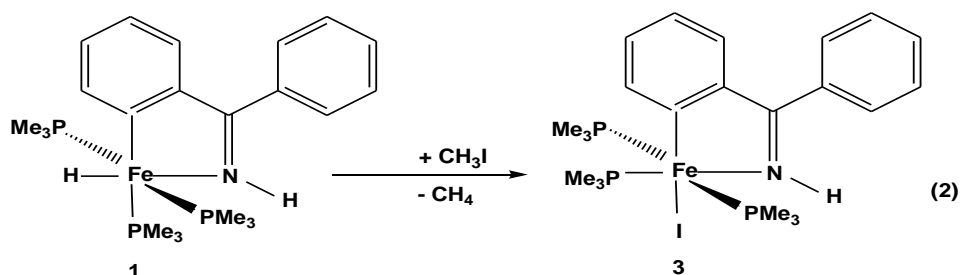
**Figure 1:** Molecular structure of 1. Selected bond lengths (Å) and angles(°): Fe1-H1 1.480(19), Fe1-N1 1.9434(13), Fe1-C1 1.9860(15), Fe1-P3 2.1803(5), Fe1-P1 2.1874(5), Fe1-P2 2.2146(5), N1-C7 1.314(2); N1-Fe1-C1 79.0(6); N1-Fe1-H1 172.8(7), P3-Fe1-P1 152.878(18), P3-Fe1-P2 95.320(19), P1-Fe1-P2 95.66(2), N1-Fe1-P3 103.01(4), C1-Fe1-P2 167.36(5).

The X-ray diffraction study on a single crystal of 2 (Fig. 2) confirms the structure proposed in eq. 1. The iron atom is coordinated in a distorted octahedral fashion with two trimethylphosphines in *trans* positions P1-Fe1-P1= 153.16(3)°, An ideal equatorial plane is formed by the atoms C1 and N1 of the *ortho*-metalated imine coordinating with the iron atom to form a five-membered ring with a bite angle N1-Fe1-C1 = 78.93(9)°. and with a sum of internal angles (539.9°) close to the ideal value (540°) is planar. The hydrido-iron (II) complex 1 is good candidates for oxidative addition reactions. Potential ligands that do not have a lone pair of filled  $\pi$  type orbitals are still able to interact with transition metal complexes by breaking a  $\sigma$  bond. Using iodomethane as a substrate, the interesting question arises to whether attack by the metal would proceed as a regioselective addition or promote ring opening through a subsequent reductive C, C- coupling reaction.

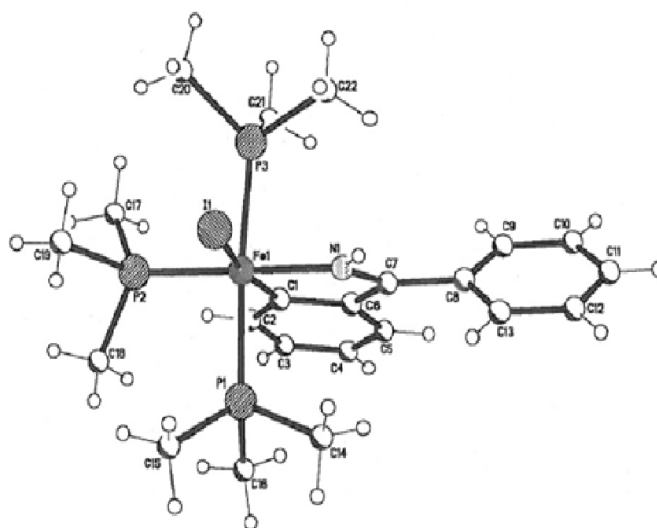


**Figure 2:** Molecular structure of 2 ( ORTEP plot with hydrogen atoms omitted). Selected bond lengths (Å) and angles(°): Fe1-N1 1.9316(19), Fe1-C1 1.966(2), Fe1-P1 2.1739(5), Fe1-P1A 2.1739(5), Fe1-P2 2.2064(8), N1-C7 1.297(3) ; N1-Fe1-C1 78.93(9), C1-Fe1-P2 168.98(7), N1-Fe1-P1 102.105(17), C1-Fe1-P1 68.71(2), P1-Fe1-P2.

95.705(18). Reaction of 1 in pentane proceeds as described in [Eq.2].



When iodomethane is added to hydrido-iron (II) complexes at  $-70^{\circ}\text{C}$ . Compound 3 crystallize from pentane as dark pink cubes yield (77%) and decompose at  $< 143^{\circ}\text{C}$ . The complete transformation of 1 into the iodo complex 3 can be monitored in the IR spectrum by the disappearance of the Fe-H band at  $1730\text{ cm}^{-1}$  and there is a hypsochromic shift in all characteristic bands. In the  $^1\text{H}$  NMR spectra, there is no Fe-H resonance at high field and trimethylphosphine  $^1\text{H}$  and  $^{13}\text{C}$  resonances are shifted to low field. Phosphorus nuclei *trans* to each other resonance around 1.1 ppm and the third around 1.6 ppm. The observed patterns in the  $^{31}\text{P}$  NMR spectra are consistent with a meridional arrangement of the three trimethylphosphines. On the basis of chemical shifts, the electron density at the iron nucleus is subject to the  $\sigma$ -electron withdrawing power of the halides [30].

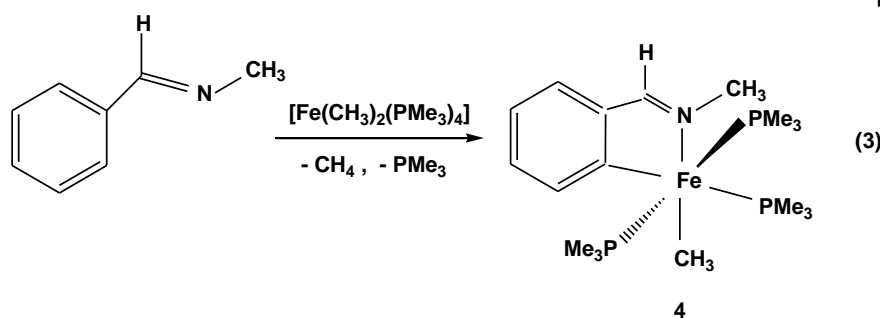


**Figure 3. Molecular structure of 3. Selected bond lengths (Å) and angles(°): Fe1-C1 1.957(3), Fe1-N1 1.940(2), Fe1-I1 2.7616(5), N1-C8 1.294(4), Fe1-P1 2.2587(9), Fe1-P2 2.2350(9), Fe1-P3 2.2568(9); N1-Fe1-C1 80.05(12), C1-Fe1-I1 166.44(9), N1-Fe1-P1 88.68(8); N1-Fe1-P2 175.24(9); N1-Fe1-P3 85.93(8); N1-Fe1-I 86.64(8); C1-Fe1-P2 95.51(9).**

A molecule of the compound 3 is an octahedral geometry, a meridional arrangement of the three trimethylphosphines is recognized. The remaining three coordination sites are occupied by an iodide and a chelating imine ligand. The iodide and one trimethylphosphine P-donor are coplanar with the metallacycles forming an angle P2-Fe1-I1  $97.87(3)^{\circ}$ . Because of its weak *trans* influence the iodide is coordinated *trans* to the metallated carbon C1-Fe1-I1  $166.44(9)^{\circ}$ . The five-membered metallacycle has a bite angle of  $80.05(12)^{\circ}$  and the sum of internal angles ( $539.85^{\circ}$ ) approach the ideal value of a planar five-membered ring. Complex 3 is distorted than 1 because iodine ligand restricts the flexibility of the trimethylphosphine groups. N1-Fe1 and C1-Fe1 distances are  $1.940(2)\text{ Å}$  and  $1.957(3)\text{ Å}$ , respectively and fall in the expected range.

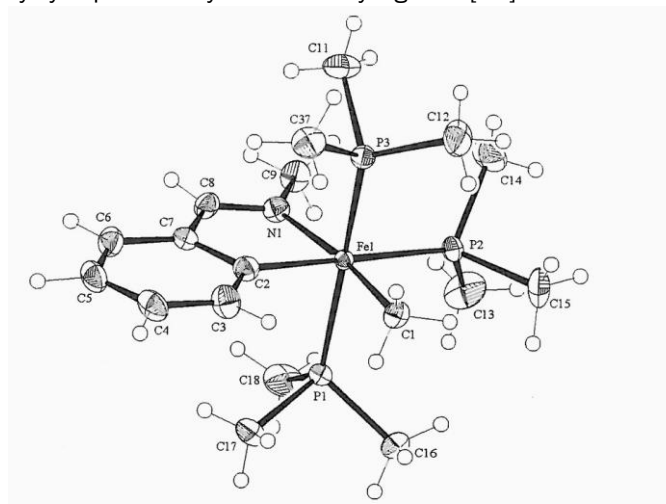
The Fe1-I1 distance  $2.7616(5)\text{ Å}$  agrees with the Fe-I  $2.6646\text{ Å}$  distance in an *ortho*-metallated octahedral Fe(II) compound  $\text{Fe}(\text{CO})_2\{\text{P}(\text{O}^i\text{Ph})_3\}\{(\text{PhO})_2\text{POC}_6\text{H}_4\}(\text{I})$  [31]. The two mutually *trans* trimethylphosphine groups have long Fe-P distances compared with the one that is *trans* influence of trialkylphosphines. All bond lengths are well-matched with the other structural features [32,33].

Benzylic imines(N-methyl) react with  $\text{Fe}(\text{CH}_3)_2(\text{PMe}_3)_4$  in THF at  $-70^{\circ}\text{C}$  through the elimination of methane to afford the hexacoordinate iron(II) complex 4 [Eq. 3].



The violet solution of compound 4 in pentane deposit crystals at  $-27^{\circ}\text{C}$ . The solids have moderate air-sensitivity and under argon decompose at  $115^{\circ}\text{C}$ , elemental composition shows that the loss of methane and one of the trimethylphosphine groups during the reaction. The infrared spectrum of compound 4 shows the absorption band  $\nu(\text{C}=\text{N})$  of coordinated imine ligand at  $1572\text{ cm}^{-1}$  which occur at significantly lower energies than the free imine  $\text{C}=\text{N}$  stretching absorption, the ( $\nu\text{C}=\text{N}$ ,  $1651\text{ cm}^{-1}$ ). The  $^1\text{H}$  NMR spectrum of 4 in  $d_8\text{THF}$  exhibits a broad singlet resonance at 1.38 ppm due to the metal-coordinated methyl group along with the expected resonances for the protons of the metalated phenyl ring from 6.74 ppm to 7.87 ppm. The  $^{31}\text{P}$  NMR spectrum shows a triplet at 7.6 ppm ( $^2J_{p,p} = 35.6\text{ Hz}$ ) and a doublet at 16.7 ppm ( $^2J_{p,p} = 35.6\text{ Hz}$ ) with P-P coupling as expected for a mer configuration of 4.

The definitive characterization of 4 as an *ortho*-metalated imine complex is derived from an X-ray diffraction experiment on a single crystal. The coordination geometry around the iron atom octahedral with two P-donor atoms of the occupying *trans* positions:  $\text{P1-Fe-P3} = 168.04(4)^{\circ}$ . The perpendicular plane is formed by the atoms C2 and N1 of the *ortho* metalated *N*-methyl benzylidene imine ( $\text{N1-Fe1-C2} = 80.34(13)^{\circ}$ ), defining a five-membered ring with the iron atom, the P2 atom of the trimethylphosphine group disposed *trans* to C2 ( $\text{C2-Fe1-P2} = 176.97(11)^{\circ}$ ) and the  $\text{CH}_3$  ligand located in opposite position to the nitrogen atom ( $\text{N1-Fe1-C1} = 169.91(16)^{\circ}$ ). The  $\text{Fe1-N1}$  bond length of  $2.01(3)\text{ \AA}$  and the  $\text{Fe1-C2}$  bond distance of  $1.987(4)\text{ \AA}$  are typical for Fe-N and Fe-C<sub>aryl</sub> single bonds. The Fe-P distances range between  $2.20 - 2.26\text{ \AA}$  and are in agreement with the values previously found in complexes 1 – 3. Similar to the hydrido-iron(II) complexes, the reaction is believed to begin with coordination of the iron adduct to the nitrogen by substitution of one trimethylphosphine, bringing the metal closer to the *ortho* C-H bond. This chelation assistance facilitates highly selective C-H bond cleavage to give a cyclometalated iron (IV) intermediate. Reductive elimination of  $\text{CH}_4$  results in an *ortho*-metalated methyl-iron (II) complex with three trimethylphosphine ligands. Alkyl complexes of iron (II) are usually coordinatively saturated, 18 electron species supported by cyclopentadienyl and carbonyl ligands.[34]



**Figure 4.** Molecular structure of 4. Selected bond lengths ( $\text{\AA}$ ) and angles ( $^{\circ}$ ):  $\text{Fe1-C1 } 2.085(4)$ ,  $\text{Fe1-C2 } 1.987(4)$ ,  $\text{Fe1-N1 } 2.014(3)$ ,  $\text{N1-C8 } 1.290(5)$ ,  $\text{Fe1-P1 } 2.2241(11)$ ,  $\text{Fe1-P2 } 2.2568(11)$ ,  $\text{Fe1-P3 } 2.2037(11)$ ;  $\text{N1-Fe1-C1 } 169.91(16)$ ,  $\text{N1-Fe1-C2 } 80.34(13)$ ,  $\text{C2-Fe1-P2 } 176.97(11)$ ,  $\text{P1-Fe1-P3 } 168.04(4)$ ,  $\text{N1-Fe1-P1 } 93.76(9)$ ,  $\text{C1-Fe1-P3 } 85.08(15)$ .

**Table 1: Crystal data for compound 1.**

Empirical formula	$\text{C}_{22}\text{H}_{38}\text{FeN}_3\text{P}_3$
Formula weight	465.29
Temperature	$293(2)\text{ K}$
Wavelength	$0.71073\text{ \AA}$

Crystal system	Triclinic
Space group	P-1
Unit cell dimensions	a = 9.2765(12) Å      α = 65.433(9)°. b = 11.4028(13) Å      β = 78.456(10)°. c = 13.4638(16) Å      γ = 74.800(10)°.
Volume	1243.2(3) Å <sup>3</sup>
Z	2
Density (Calculated)	1.243 Mg/m <sup>3</sup>
Absorption coefficient	0.807 mm <sup>-1</sup>
F (000)	496
Crystal size	0.15 x 0.2 x 0.5 mm
Theta range for data collection	1.67 to 26.92°.
Index ranges	-11<=h<=11, -41<=k<=41, 17<=l<=17
Reflections collected	18486
Independent reflections	5253 [R (int) = 0.0990]
Completeness to theta =26.92 °	97.7%
Refinement method	Full-matrix least-squares on F <sup>2</sup>
Data / restraints / parameters	5253 / 0 / 396
Goodness-of-fit on F <sup>2</sup>	1.109
Final R indices [I>2sigma (I)]	R1 = 0.0294, wR2 = 0.0778
R indices (all data)	R1 = 0.0317, wR2 = 0.0791
Largest diff. peak and hole	0.461 and -0.517 e. Å <sup>-3</sup>

**Table 2: Crystal data for compound 2.**

Empirical formula	C <sub>20</sub> H <sub>42</sub> FeN <sub>3</sub> P <sub>3</sub>
Formula weight	445.31
Temperature	298(2) K
Wavelength	0.71073 Å
Crystal system	Orthorhombic
Space group	Pnma
Unit cell dimensions	a = 9.6581(7) Å      α = 90°. b = 13.5905(10) Å      β = 90°. c = 18.7569(14) Å      γ = 90°.
Volume	2462.0(3) Å <sup>3</sup>
Z	4
Density (Calculated)	1.201 Mg/m <sup>3</sup>
Absorption coefficient	0.812 mm <sup>-1</sup>
F (000)	960
Crystal size	0.40 x 0.20 x 0.18 mm <sup>3</sup>
Theta range for data collection	1.85 to 28.20°.
Index ranges	-11<=h<=12, -18<=k<=18, -24<=l<=24
Reflections collected	29358
Independent reflections	3145 [R (int) = 0.0990]
Completeness to theta =28.20 °	99.8 %
Refinement method	Full-matrix least-squares on F <sup>2</sup>
Data / restraints / parameters	3145 / 0 / 143
Goodness-of-fit on F <sup>2</sup>	0.997
Final R indices [I>2sigma (I)]	R1 = 0.0378, wR2 = 0.1010
R indices (all data)	R1 = 0.0483, wR2 = 0.1098
Largest diff. peak and hole	0.493 and -0.299 e. Å <sup>-3</sup>

**Table 3: Crystal data for compound 3.**

Empirical formula	C <sub>22</sub> H <sub>37</sub> FeNP <sub>3</sub>
Formula weight	591.19
Temperature	293(2) K
Wavelength	0.71073 Å
Crystal system	Orthorhombic
Space group	P2(1)2(1)2(1)
Unit cell dimensions	a = 10.4504(9) Å      α = 90°. b = 13.7867(12) Å      β = 90°. c = 18.3930(15) Å      γ = 90°.
Volume	2650.0(4) Å <sup>3</sup>

Z	4
Density (Calculated)	1.482Mg/m <sup>3</sup>
Absorption coefficient	1.924 mm <sup>-1</sup>
F (000)	1200
Crystal size	0.45 x 0.40 x 0.35 mm <sup>3</sup>
Theta range for data collection	1.85 to 28.34°.
Index ranges	-13<=h<=13,-18<=k<=18, -23<=l<=24
Reflections collected	27374
Independent reflections	6596 [R (int) = 0.0426]
Completeness to theta =28.34 °	99.9 %
Refinement method	Full-matrix least-squares on F <sup>2</sup>
Data / restraints / parameters	6596 / 0 / 262
Goodness-of-fit on F <sup>2</sup>	0.965
Final R indices [I>2sigma (I)]	R1 = 0.0347, wR2 = 0.0739
R indices (all data)	R1 = 0.0416, wR2 = 0.0768
Largest diff. peak and hole	0.971 and -0.412 e. Å <sup>-3</sup>

Table 4: Crystal data for compound 4.

Empirical formula	C <sub>18</sub> H <sub>38</sub> FeNP <sub>3</sub>	
Formula weight	417.25	
Temperature	150(2) K	
Wavelength	0.71073 Å	
Crystal system	Triclinic	
Space group	P-1	
Unit cell dimensions	a = 8.6470(14) Å	α = 103.140(16)°.
	b = 14.882(4) Å	β = 95.330(13)°.
	c = 17.814(3) Å	γ = 90.060(17)°.
Volume	2222.1(7) Å <sup>3</sup>	
Z	4	
Density (Calculated)	1.247 Mg/m <sup>3</sup>	
Absorption coefficient	0.895 mm <sup>-1</sup>	
F (000)	896	
Crystal size	0.25 x 0.2 x 0.1 mm <sup>3</sup>	
Theta range for data collection	1.62 to 26.86°.	
Index ranges	-10<=h<=10,-18<=k<=18, -22<=l<=22	
Reflections collected	26214	
Independent reflections	9308 [R (int) = 0.0727]	
Completeness to theta =26.86 °	97.7 %	
Refinement method	Full-matrix least-squares on F <sup>2</sup>	
Data / restraints / parameters	9308 / 0 / 778	
Goodness-of-fit on F <sup>2</sup>	1.053	
Final R indices [I>2sigma (I)]	R1 = 0.0626, wR2 = 0.1186	
R indices (all data)	R1 = 0.0744, wR2 = 0.1274	
Largest diff. peak and hole	1.753 and -0.440 e. Å <sup>-3</sup>	

Transition metal alkyl complexes are of practical importance due to their role in numerous catalytic transformations such as olefin hydrogenation, hydroformylation, and polymerization [35]. It is the existence of several decomposition pathways that makes many transition metal alkyls unstable. Early efforts to prepare simple alkyls or aryls of transition metals showed that such compounds were generally unstable under ambient conditions. Despite the synthesis of [Me<sub>3</sub>Pt]<sub>4</sub> by Pope and Peachy, it was assumed until quite recently that metal-carbon bonds are weak [36]. In fact, we now know that such M-C bonds are reasonably strong ranging 165-350 kJ mol<sup>-1</sup>. When it was found that the presence of ligands such as η-C<sub>5</sub>H<sub>5</sub>, CO or PR<sub>3</sub> allowed the synthesis of thermally stable compounds, it was considered that the presence of such ligands provided better overlap between sp<sup>3</sup> hybridized carbon and metal orbitals implying increased bond strength. A lacking β-decomposition pathway by choice of a Fe-CH<sub>3</sub> group, a blocking imine ligand and a PMe<sub>3</sub> ligand are in favor of stable compound 4. The principle governing their stability make a useful starting point in order to understand some of the most important organometallic reactions.

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