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COMMUNICATION

# Samarium Coordinated Polymer: Structural, Vibrational and Thermal Studies of $[\text{Sm}_2(\text{C}_3\text{H}_2\text{O}_4)_3(\text{H}_2\text{O})_6]_n$

B. H. Doreswamy · M. Mahendra ·  
J. Shashidhara Prasad · P. A. Varughese ·  
George Varghese

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**Abstract** The title compound,  $[\text{Sm}_2(\text{C}_3\text{H}_2\text{O}_4)_3(\text{H}_2\text{O})_6]$ , was investigated by X-ray diffraction. It crystallizes in the monoclinic space group  $C2/c$  with cell parameters  $a = 17.1650(8) \text{ \AA}$ ,  $b = 12.3010(5) \text{ \AA}$ ,  $c = 11.1420(4) \text{ \AA}$ ,  $\beta = 127.5161(10)^\circ$ ,  $Z = 4$  and  $V = 1866.04(14) \text{ \AA}^3$ . The Sm atom lies on a two-fold axis and has nine-coordination with six oxygen atoms from carboxylate groups and three water molecules. The compound forms a layer-type polymeric structure. The layers are formed by samarium and one independent malonate group to give a three-dimensional framework. The extensive network of hydrogen bonds and bridge bonds observed in this structure enhances the structural stability. The thermal dehydration of the compound was investigated by thermogravimetric analysis.

**Keywords** Samarium · Malonate hydrates · FTIR · TGA · X-ray crystal structure

## 1 Introduction

In the vast field of materials science and technology, rare earths or rare earth-related compounds are of immense

interest due to their many interesting and peculiar properties [1–6]. Many researchers have carried out extensive work on dicarboxylates of rare earth elements (e.g., oxalates [7–9], malonates [10, 11], maleates [12, 13], tartrates [14] owing to their important applications. However, little has been done on rare earth malonates. In our current study, we have investigated a samarium complex with malonate as a ligand and different water of hydrations. An interesting feature of the malonate ligand is that it exhibits flexible stereochemistry and variable modes of binding with metal ions in the crystalline state [15–25]. It is also characterized by an active methylene ( $-\text{CH}_2-$ ) group between two carboxylate groups. Three bonding modes, S1, S2 and S3 (Fig. 1) have been observed. In this report, the flexible coordination ability of samarium is discussed; it adopts the S2 and S3 bonding mode; i.e., monodentate, chelating and bridging [15–25]. In view of the above, we report here the growth of  $[\text{Sm}_2(\text{C}_3\text{H}_2\text{O}_4)_3(\text{H}_2\text{O})_6]_n$  (**1**) crystals by the gel technique [26] and the X-ray crystal study to discern the coordination geometry around the rare earth. In addition, we report the Fourier transform infrared (FT-IR) spectrum and thermogravimetric analysis (TGA) of **1**.

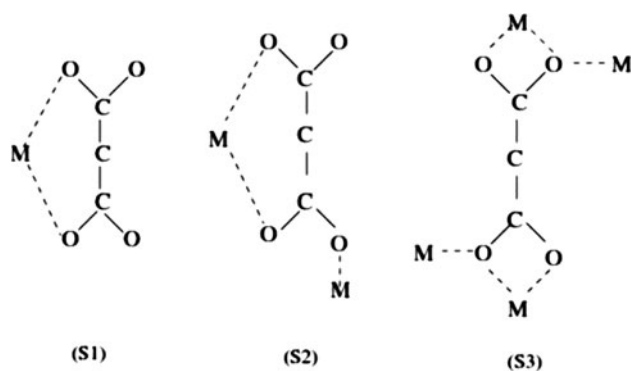
## 2 Experimental

Pale yellow colored single crystals of **1** were prepared using the diffusion-reaction process inside a non-reacting gel medium. The growth medium of hydrosilica gel was prepared by titrating sodium metasilicate solution (specific gravity =  $1.03 \text{ g cm}^{-3}$ ) with malonic acid (1 M, purity 99.99%) to pH 5.5. This solution was then poured into glass tubes having an internal diameter of 18 mm and kept undisturbed for gel formation. After 3 days, the samarium chloride solution (0.2 M, purity 99.999%) was poured over

B. H. Doreswamy  
Department of Physics, SJB Institute of Technology,  
Kengeri, Bangalore 560 060, India

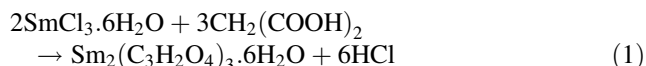
M. Mahendra (✉) · J. Shashidhara Prasad  
Department of Studies in Physics, University of Mysore,  
Mysore 570 006, India  
e-mail: mahendra@physics.uni-mysore.ac.in

P. A. Varughese · G. Varghese  
Crystal Physics Centre, St. Berchmans' College,  
Changanachery 686 101, Kerala, India



**Fig. 1** Bonding mode, S1, S2 and S3, of the malonate ligand where M = Cu, Co, Eu, Cr for S1, M = Pr, Eu, Cu, Sc, Zn, Nd for S2 and M = Eu, Pr, Na for S3

the gel column and kept tranquil to initiating the diffusion process and subsequent chemical reaction for the crystallization of samarium malonate (Eq. 1). Faceted crystals of average size  $4 \times 1 \times 0.6 \text{ mm}^3$  were obtained in 4–5 weeks after initiation of the reaction process.



The infrared spectra from  $4000\text{--}400 \text{ cm}^{-1}$  were carried out with a Bruker IFS66 spectrophotometer as a KBr pellet.

**Table 1** Crystallographic data of **1**

Empirical formula	$\text{C}_9 \text{H}_{18} \text{O}_{18} \text{Sm}_2$
Formula weight	714.93
Temperature	293(2) K
Wavelength	0.71073 Å
Crystal system, space group	Monoclinic, C2/c
Unit cell dimensions	$a = 17.1650(8) \text{ Å}$ $b = 12.3010(5) \text{ Å}$ $c = 11.1420(5) \text{ Å}$ $\beta = 127.5161(10)^\circ$
Volume	$1866.04(14) \text{ Å}^3$
Z, Calculated density	4, 2.545 $\text{Mg/m}^3$
Absorption coefficient	$6.324 \text{ mm}^{-1}$
$F(000)$	1360
Crystal size	$0.2 \times 0.2 \times 0.2 \text{ mm}$
Theta range for data collection	$21.97\text{--}32.48^\circ$
Limiting indices	$25 \leq h \leq 25, -13 \leq k \leq 13,$ $-16 \leq l \leq 16$
Reflections collected/unique	3279/1748 [ $R(\text{int}) = 0.0201$ ]
Refinement method	Full-matrix least-squares on $F^2$
Data/restraints/parameters	1748/0/132
Goodness-of-fit on $F^2$	1.142
Final $R$ indices [ $I > 2\sigma(I)$ ]	$R_1 = 0.0425, wR_2 = 0.1214$
$R$ indices (all data)	$R_1 = 0.0438, wR_2 = 0.1229$
Largest diff. peak and hole	1.524 and $-1.408 \text{ e.Å}^{-3}$

The resolution of the instrument was better than  $4 \text{ cm}^{-1}$ . All the characteristic bands of the compound are observed in the spectrum.

Thermogravimetric analysis of **1** was performed on a STA 1500 thermal analysis system. The compound was heated from ambient temperature to  $900 \text{ °C}$  at a heating rate of  $10 \text{ °C}$  per minute under nitrogen.

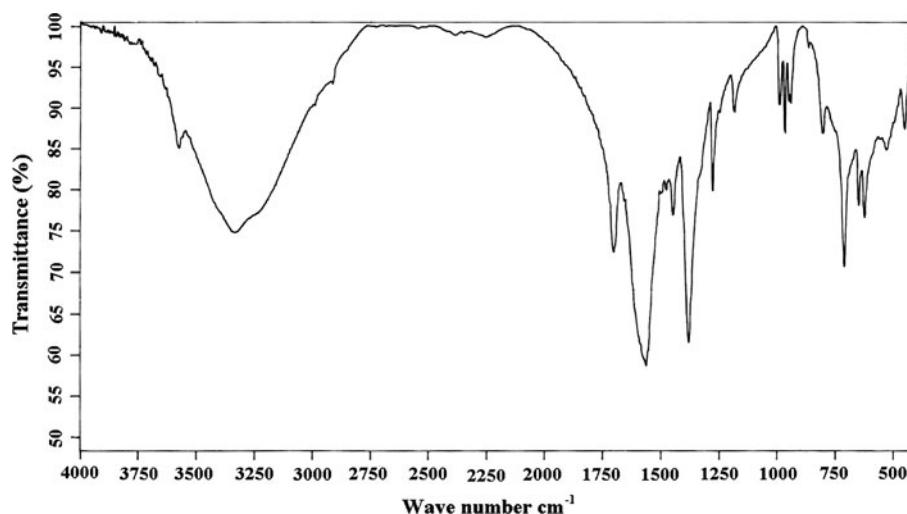
A single crystal of **1** of dimensions  $0.2 \times 0.2 \times 0.2 \text{ mm}^3$  was chosen for X-ray diffraction studies. The measurements were made on a DIPLabo Imaging Plate system with graphite monochromated  $\text{MoK}_\alpha$  radiation ( $\lambda = 0.71073 \text{ Å}$ ). Thirty-six frames of data were collected using the oscillation method. Successive frames were scanned in steps of  $5^\circ/\text{min}$  with an oscillation range of  $5^\circ$ . Image processing and data reduction were done using Denzo [27]. The structure was solved and refined using the maXus [28–30] program. All the non-hydrogen atoms were revealed in the first map. Initially, full-matrix least-squares refinement was performed using 1,748 observed reflections ( $I > 2\sigma(I)$ ) for all the non-hydrogen atoms with isotropic temperature factors; the residual at this stage converged to  $R = 0.0897$ . The hydrogen atoms were placed at calculated positions and were refined with isotropic temperature factors. Refinement of non-hydrogen atoms with anisotropic thermal parameters was started at this stage. After ten cycles of refinement, the residuals saturated at  $R = 0.0425$ . The details of crystal data, data collection and refinement are given in Table 1.

## 3 Results and Discussion

### 3.1 Spectral Analysis

The FT-IR spectral study reveals the bonding structure of the **1**. The spectrum (Fig. 2) was interpreted by comparing with those of malonic acid and related compounds [31–34]. In **1**, the internal vibrations are mainly due to carboxylate group, methylene group and water molecules. The most interesting frequencies with respect to structure are those of C–O vibrations. These are assigned (Table 2) to the asymmetric and symmetric stretching modes of the carboxylate group. The highly intense bands at  $1564$  and  $1381 \text{ cm}^{-1}$  are ascribed to the asymmetric stretching and symmetric stretching modes of the carboxylate group, respectively. The lowering of both the stretching frequencies is due to the drainage of the electron density from the carboxylate group to the metal to form chelates or bridges between carboxylates. The medium bands at  $711$  and  $650 \text{ cm}^{-1}$  are ascribed to the bending modes of the carboxylate group.

The asymmetric and symmetric stretching modes of methylene group are observed as weak bands at  $2986$  and

**Fig. 2** FTIR spectrum of **1****Table 2** Absorption wave numbers and proposed assignments of **1**

Wave number (cm <sup>-1</sup> )	Assignments
3332(svbr)	$\nu$ OH
2986(w)	$\nu_{as}$ CH <sub>2</sub>
2912(w)	$\nu_s$ CH <sub>2</sub>
1703(s)	$\delta$ H <sub>2</sub> O
1564(vsbr)	$\nu_{as}$ OCO
1502(m)	$\nu_{as}$ OCO
1478(m)	$\delta$ CH <sub>2</sub>
1449(m)	$\delta$ CH <sub>2</sub>
1381(vs)	$\nu_s$ OCO
1278(m)	$\omega$ CH <sub>2</sub>
1244(w)	$\omega$ CH <sub>2</sub>
1185(w)	$\nu_{as}$ CC
967(w)	$\nu_s$ CC
950(w)	$\nu_s$ CC
990(w)	$\rho$ CH <sub>2</sub>
941(w)	$\rho$ CH <sub>2</sub>
866(vw)	$\delta$ CC
803(w)	$\rho$ H <sub>2</sub> O
711(s)	$\delta$ OCO
650(m)	$\delta$ OCO
625(m)	$\delta$ CCO
531(w)	$\omega$ , $t$ H <sub>2</sub> O
482(m)	$\omega$ , $t$ H <sub>2</sub> O
453(w)	$\nu$ M–OH <sub>2</sub>

Note.  $\nu$  stretching,  $\nu_{as}$  asymmetric stretching,  $\nu_s$  symmetric stretching,  $\delta$  bending,  $\omega$  wagging,  $\rho$  rocking,  $t$  twisting,  $s$  strong,  $svbr$  strong and very broad,  $vs$  very strong,  $vsbr$  very strong and broad,  $m$  medium,  $w$  weak,  $vw$  very weak

2912 cm<sup>-1</sup>, while the bending modes are associated with the medium bands at 1478 and 1449 cm<sup>-1</sup>. The medium band at 1278 cm<sup>-1</sup> and the weak band at 1244 cm<sup>-1</sup> are

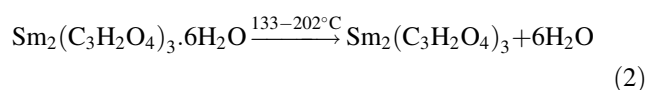
due to –CH<sub>2</sub>– wagging. In addition, the rocking modes of this group are observed as weak bands at 990 and 941 cm<sup>-1</sup>.

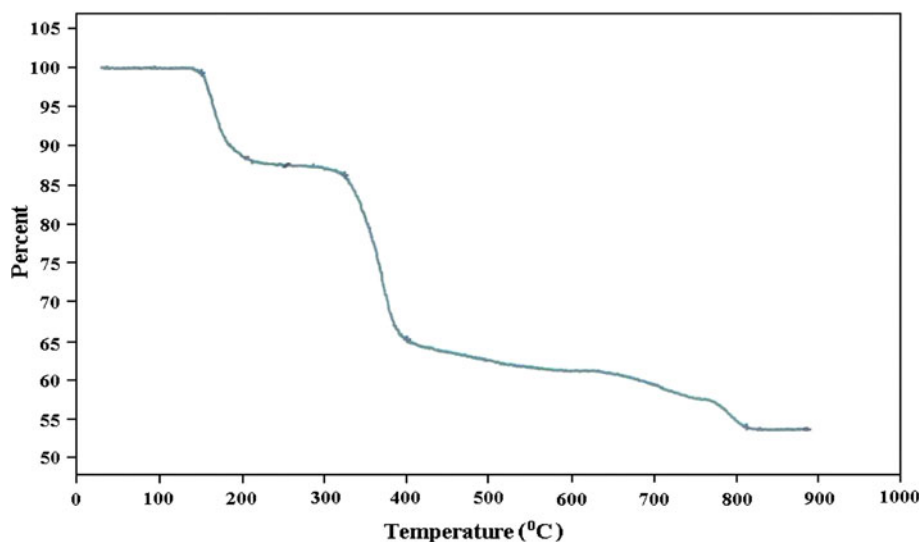
The weak bands at 1185, 967 and 950 cm<sup>-1</sup> are attributed to the C–C stretching modes. The C–C bending mode is seen as a very weak band at 866 cm<sup>-1</sup>. A medium band at 625 cm<sup>-1</sup> is assigned to the C–CO bending mode.

The presence of water of hydration in **1** is evidenced by the strong and very broad band centered at 3332 cm<sup>-1</sup>. This is due to the O–H stretching vibration. There is a decrease in frequency, which indicates hydrogen bonding. The bending mode of water is observed as a strong band at 1703 cm<sup>-1</sup> whereas the rocking mode is a weak band at 803 cm<sup>-1</sup>. The weak band at 531 cm<sup>-1</sup> is ascribed to the wagging and twisting modes of water. Also, a weak band at 453 cm<sup>-1</sup> is attributed to the M–OH<sub>2</sub> stretching vibration.

### 3.2 Thermogravimetric Analysis

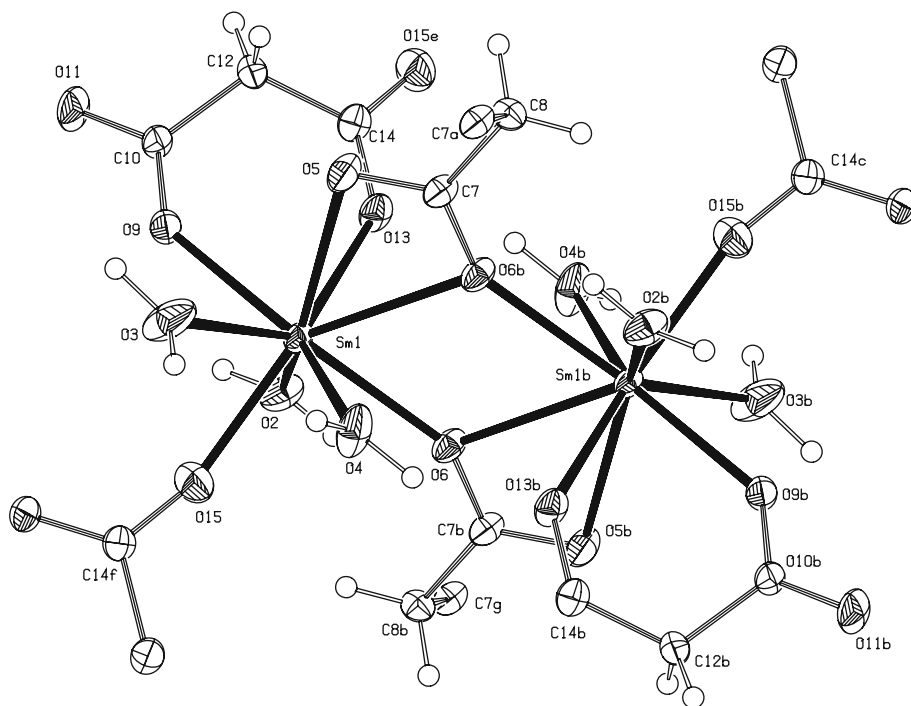
Thermogravimetric (TG) study reveals that **1** is stable to 133 °C. The TG curve of **1** (Fig. 3) shows a three-stage thermal decomposition. Dehydration occurs in the first stage (133–202 °C) in which all the six waters of hydration are lost. Above 133 °C, the TG of anhydrous malonate is a plateau and is transformed into the oxycarbonate of the metal and elementary carbon (323–399 °C). In the final decomposition stage (399–810 °C) the oxycarbonate is decomposed to yield the metal oxide. The residue comprises a mixture of elementary carbon and the metal oxide. The thermal decomposition occurring in each stage may be described as follows:

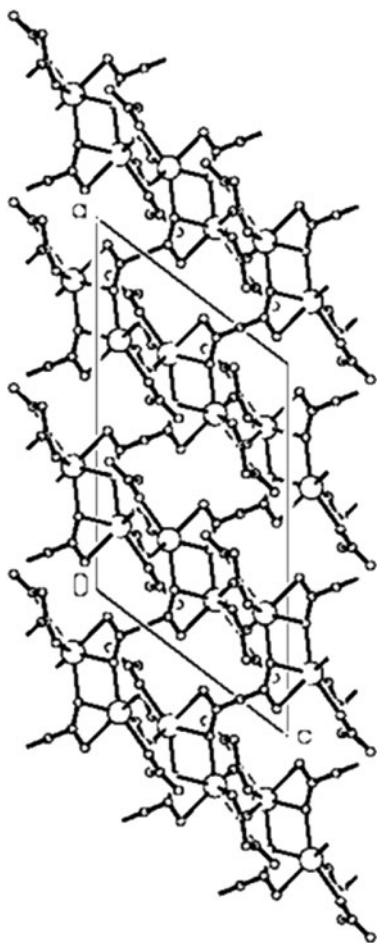


**Fig. 3** TGA curve of **1****Table 3** Thermal decomposition data of **1**

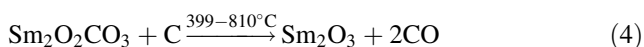
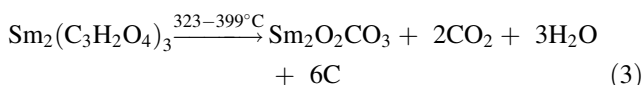
Stage	Temperature range (°C)	Mass loss %		Leaving groups	% of metal oxide + carbon	
		Observed	Calculated		Observed	Calculated
1	133–202	14.50	15.13	6H <sub>2</sub> O		
2	323–399	19.50	19.89	2CO <sub>2</sub> , 3H <sub>2</sub> O		
3	399–810	12.05	11.76	2CO	53.95	53.22

**Fig. 4** Perspective ORTEP [45] diagram of **1** with thermal ellipsoids at 50% probability. H atoms have been omitted for clarity [Symmetry codes:  $a = (1/2-x, 1/2-y, -z)$ ,  $b = (-1/2+x, 1/2-y, -1/2+z)$ ,  $c = (x, 1-y, -1/2+z)$ ,  $d = (-1/2+x, -1/2+y, -1+z)$ ,  $e = (1/2-x, -1/2+y, 1/2-z)$ ,  $f = (1-x, y, 1/2-z)$  and  $g = (-x, 1-y, -z)$ ]





**Fig. 5** Packing of the molecules along the *b* axis. Water molecules and H atoms have been omitted for the sake of clarity



The percentage of mass losses observed is in good agreement with the calculated values (Table 3). The weight losses due to hydration, coordinated water molecules and decomposed organic moiety of the organic ligand are in agreement with those found for other compounds described in the literature [35–41].

### 3.3 Crystal Structure Analysis

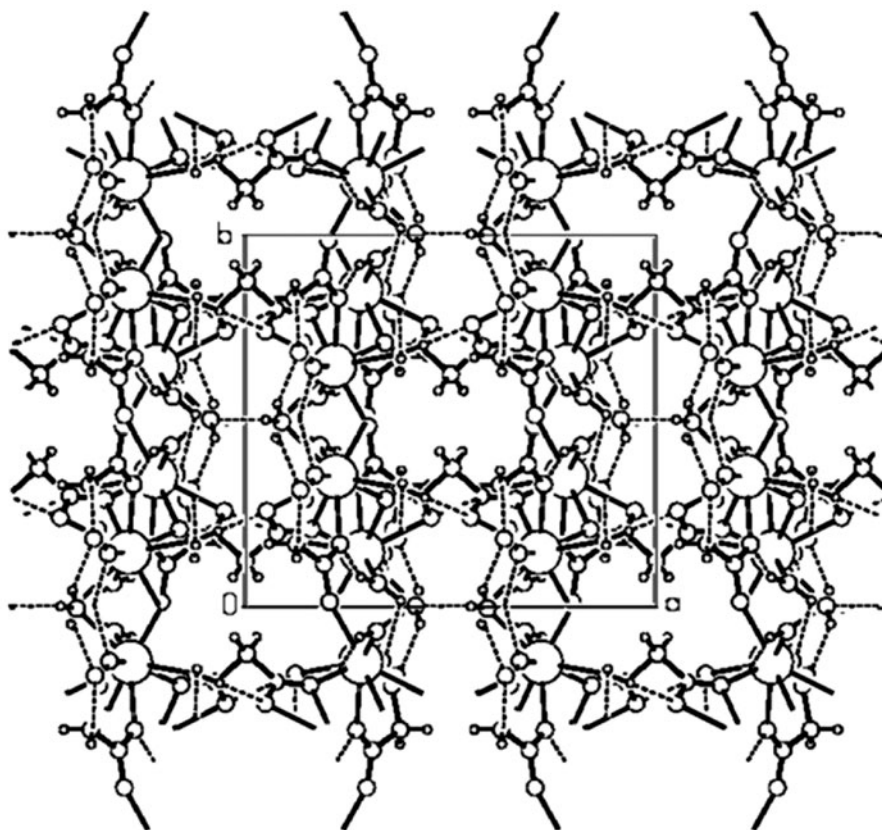
A perspective drawing showing the coordination geometry, labeling scheme and 50% probability displacement ellipsoid is shown in Fig. 4. The packing of the molecules in **1** along the axes exhibit a three-dimensional layer-type polymeric framework (Figs. 5, 6).

The layers are formed by three types of rings; i.e., Sm–O–C–O and Sm–O–C–C–O chelate rings formed by bidentate coordination and Sm–O–Sm–O rings formed by the bridging structure. The distance between two Sm atoms is 4.379(2) Å. Each samarium atom is coordinated by six carboxylate and three water molecules. The water molecules are involved in an extensive network of hydrogen bonds amongst themselves and with the non-bridging atoms of the ligands (O5 and O11). It appears that, in addition to the oxygen bridge bonds involving the malonate ligand, the crystal structure is stabilized through extensive hydrogen bonding involving the carboxylate groups and water molecules to form a three-dimensional network (Fig. 6). In all the hydrogen bonds, the water oxygens (O2, O3, O4) act as donor atoms and the carboxylate or other water oxygens act as acceptor atoms. Furthermore, additional hydrogen bonds occur between a carboxyl group and a water molecule. The structure has both intra- and inter-molecular hydrogen bonds of the type O–H···O. The lengths and angles along with the symmetry codes are given in Table 4.

For the malonate ligand, the dihedral angle between the planes O9–C10–O11–C12 and C12–O13–C14–O15 is 36.1(4)°. Two malonate units are crystallographically independent and with different arrangements; e.g., S2 and S3 (Fig. 1). The S2 mode is characterized by the malonate ligand bridging two samarium atoms, acting as both a bidentate chelate ( $^-\text{O}_2\text{CCH}_2\text{CO}_2^-$ ) and monodentate ( $^-\text{O}_2\text{C}-$ ) ligand to samarium. The chelate ring has a boat conformation while the monodentate function adopts an *anti-syn* conformation. In the S3 arrangement, the malonate ligand bridges four Sm atoms through two bidentate four-membered chelate-rings as well as two monodentate interactions of *anti-anti* conformation. The C–O bond distances and O–C–O bond angles in both malonate ligands are 1.253(5)–1.270(7) Å and 120.8(5)–123.6(6)°, respectively. These values are in agreement with previously reported malonate-lanthanide complexes [15, 42–44].

The structure of **1** shows a three-dimensional arrangement of samarium units bridged by malonate groups (Figs. 5, 6), which result from the cross-linking of the single chains that run parallel to *ac* plane (Fig. 7) and the double twisted zig-zag chains that extend in the *ab* plane (Fig. 8).

The samarium-to-water oxygen bond lengths are 2.487(5)–2.528(4) Å, which are slightly longer than the samarium-to-carboxylate oxygen lengths (2.370(4)–2.445(3) Å). This is consistent with the coordination ability of water being weaker than that of the carboxylates.



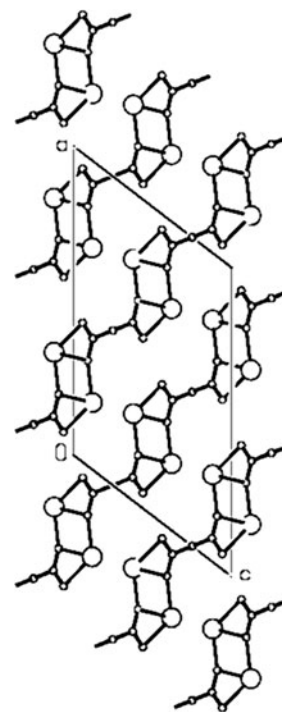
**Fig. 6** Projection of the structure down *c* axis. Dashed lines correspond to hydrogen bridging

**Table 4** Intra- and intermolecular hydrogen bonds of **1**; lengths, angles and symmetry codes

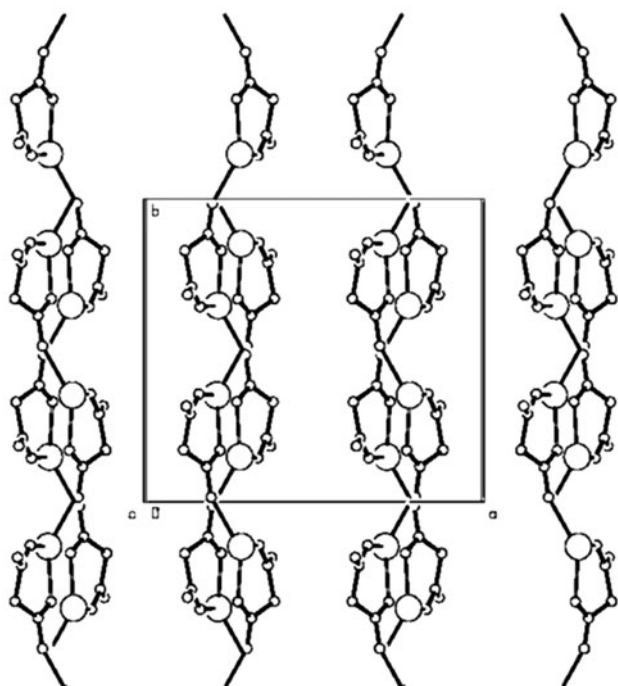
D–H···A	D–A(Å)	D–H···A(°)	Symmetry codes
O2–H2A···O11	2.7158	114	$1/2-x, 1/2-y, 1-z$
O2–H2B···O5	2.8472	174	$1/2+x, 1/2-y, 1/2+z$
O4–H15B···O13	2.7288	153	$1/2-x, 1/2-y, -z$
O3–H18A···O3	2.8474	148	$-x, 1-y, -z$
O4–H15A···O9	2.8117	113	$x, 1-y, -1/2+z$
O3–H18B···O11	2.7327	137	$x, 1-y, -1/2+z$

#### 4 Conclusion

A good quality single crystal of samarium malonate hydrate suitable for structural studies was produced by the gel technique. Single crystal studies revealed that the compound is  $[\text{Sm}_2(\text{C}_3\text{H}_2\text{O}_4)_3(\text{H}_2\text{O})_6]_n$ . All the characteristic bands expected for the compound are observed in the FT-IR spectrum. Thermogravimetric study showed that the compound is thermally stable up to 133 °C. Above 133 °C, **1** undergoes three-stages of thermal decomposition and ultimately gets reduced to samarium oxide at 810 °C.



**Fig. 7** Projection of the single chain down *b* axis. Malonate ligand S2, water molecules and H atoms have been omitted for clarity



**Fig. 8** Projection of the independent single chain down *c* axis. Malonate ligand S3, water molecules and H atoms have been omitted for clarity

## 5 Supplementary Material

CCDC no. 215588 contains the supplementary crystallographic data for this paper. These data can be obtained free of charge at [www.ccdc.cam.ac.uk/conts/retrieving.html](http://www.ccdc.cam.ac.uk/conts/retrieving.html) [or from the Cambridge Crystallographic Data Centre (CCDC), 12 Union Road, Cambridge CB2 1EZ, UK; fax: +44(0)1223-336033; email: deposit@ccdc.cam.ac.uk].

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