

Reduction of UH^+ formation for $^{236}\text{U}/^{238}\text{U}$ isotope ratio measurements at ultratrace level in double focusing sector field ICP-MS using D_2O as solvent

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The main factors affecting the accurate and precise determination of ^{236}U using ICP-MS are instrumental background, the isobaric interference of $^{235}\text{UH}^+$ molecular ion on $^{236}\text{U}^+$ analyte ions, and the presence of $^{238}\text{U}^+$ and $^{235}\text{U}^+$ peak tails. An optimized analytical method for attenuating the influence of these factors on uranium isotope ratio measurements at ultratrace level of environmental samples has been developed. In order to reduce $^{235}\text{UH}^+$ formation, D_2O (heavy water) is used as a solvent for the dissolution and dilution of uranium samples. Abundance sensitivity was improved by use of medium mass resolution ($m/\Delta m = 4450$) in comparison with low mass resolution in double-focusing sector field ICP-MS (ICP-SFMS). For solution introduction the performances of several different sample introduction systems (Meinhard, Aridus and ultrasonic nebulizer) were studied. It has been shown, that for all nebulization systems, a diminution in UH^+/U^+ is observed in D_2O as compared with H_2O as solvent. Optimum results were obtained in ICP-SFMS for a desolvating microconcentric nebulizer system (Aridus) with a minimum hydride formation rate of 9×10^{-7} and a limit for $^{236}\text{U}/^{238}\text{U}$ isotopic ratio measurements of $3 - 5 \times 10^{-7}$. A comparison was performed of three commercially available sector field ICP-MS devices, with good agreement found between single collector and multiple collector ICP-MS (MC-ICP-MS).

Introduction

Determination of the concentration and isotopic analysis of uranium at trace and ultratrace levels is of increasing interest for environmental monitoring, nuclear safeguards and nuclear forensic studies.^{1–3} In addition to analysis of natural uranium isotopes, determination of ^{236}U is a challenging task.^{4–6} In nature, uranium occurs in the Earth's crust and in sea-water, at around 2.7 ppm and 3.0 ppb, respectively.⁷ The uranium isotopes of primordial origin are ^{238}U and ^{235}U with abundances of 99.27% and 0.72%, respectively. ^{234}U in nature is produced from the ^{238}U decay series, with an abundance of 0.005%, and is approximately in radioactive equilibrium with ^{238}U . ^{236}U , which is generated by neutron capture from ^{235}U (the neutrons are obtained from (α, n) reactions, spontaneous fission of ^{238}U) and from interaction of cosmic rays with matter, is present in the natural Earth's crust at ultratrace levels.⁸ The natural abundance of ^{236}U is highest in U-rich materials, and several studies have demonstrated that the $^{236}\text{U}/^{238}\text{U}$ isotopic ratio is of the order of 10^{-10} in U ores.^{9–11} A $^{236}\text{U}/^{238}\text{U}$ isotopic ratio of the order of 10^{-13} – 10^{-14} is expected in the Earth's crust.¹² However, enhanced levels of ^{236}U are produced (in the past ~ 50 years) in nuclear reactors by the neutron irradiation of ^{235}U : $^{235}\text{U} + {}^1_0\text{n} \rightarrow ^{236}\text{U}$ (with a cross section of 95 barns for thermal neutrons *versus* fission cross section of 586 barns). As a result, the $^{236}\text{U}/^{238}\text{U}$ isotopic ratio increases to up to 10^{-3} – 10^{-2} , and the ^{236}U has a "poisoning" effect upon the fission of ^{235}U .⁸ Recovered reactor-irradiated U contains ^{236}U and its presence has been confirmed in depleted uranium.^{12–14} Artificial ^{236}U has been identified in Chernobyl-vicinity soils,⁴ several other environmental sites,⁵

and in the NBS standard reference materials (SRM) 950a and 950b.¹⁰ The determination of ^{236}U , therefore, can be used as a powerful tool for "fingerprinting" of artificial uranium in environmental samples. Large increases in the $^{236}\text{U}/^{238}\text{U}$ isotopic ratio represent a much more sensitive indicator of the presence of irradiated uranium than small deviations in $^{235}\text{U}/^{238}\text{U}$ from its value in nature (0.00725). Thus, according to the different studies,^{15,16} the estimated $^{235}\text{U}/^{238}\text{U}$ isotopic ratio in spent uranium of the Chernobyl reactor varies from 0.0067 to 0.0084, which makes difficult accurate determination of the spent uranium in contaminated soil samples by measuring $^{235}\text{U}/^{238}\text{U}$ isotopic ratio. The accurate measurement of $^{236}\text{U}/^{238}\text{U}$ isotopic ratio, therefore, is of great importance as evidence of contamination of environmental samples with uranium from nuclear reactors, reprocessing plants, or other anthropogenic sources.

In recent years, several approaches have been utilized for $^{236}\text{U}/^{238}\text{U}$ isotopic ratio measurements. Since ^{236}U decays by alpha-emission, with alpha-particle energies, E_α , of 4.494 MeV (74%) and 4.445 MeV (26%),¹⁷ alpha spectrometry has been used in most laboratories.^{6,14,18} However, ^{236}U has a relatively low specific activity, and its alpha particle energies are relatively close to ^{235}U , which exhibits decay energies of 4.502 MeV (relative intensity 1.7%) and 4.435 MeV (relative intensity 0.7%).¹⁷ Compared with decay-counting techniques, mass spectrometry is advantageous due to its high sensitivity and the possibility of measuring isotope ratios with good accuracy and precision for determination of long-lived nuclides.^{1,19} Thermal ionization mass spectrometry (TIMS)^{20–22} and accelerator mass spectrometry (AMS)^{7,8,23} are well-established techniques for uranium isotope ratio measurements at the ultratrace level.

Measurements of $^{236}\text{U}/^{238}\text{U}$ isotopic ratio down to 10^{-12} are feasible by AMS,²⁴ but the capital costs and centralized placement of AMS facilities restrict its use to specialized applications. Inductively coupled plasma mass spectrometry (ICP-MS) exhibits high sensitivity, good accuracy and precision of isotopic measurement and a relatively simple sample preparation procedure;^{25,26} arguably, ICP-MS is one of the most suitable methods for routine uranium isotope ratio measurement in environmental samples.^{27–29} However, the determination of ^{236}U by ICP-MS is constrained by isobaric interference from $^{235}\text{U}^1\text{H}^+$,^{4,5,30} as well as the limited abundance sensitivity of ICP-MS instruments.^{31–33} Both of these effects result in increases in the background signal at m/z 236. The abundance sensitivity arises from overlap of ^{235}U and ^{238}U peak tails, and is especially problematic in sector field ICP-MS at low mass resolution. As an example, Boulyga and Becker³⁴ found an abundance sensitivity at $m/z = 236$ of 6×10^{-8} in quadrupole ICP-MS with hexapole collision cell, exhibiting an abundance sensitivity up to two orders of magnitude better, which was ruled to be preferential. It is, however, possible to improve the abundance sensitivity in ICP-SFMS at medium or high mass resolution; Ketterer *et al.*⁵ obtained an abundance sensitivity of 5×10^{-8} at $m/z = 236$ in the presence of ^{238}U by using medium mass resolution ($m/\Delta m = 3440$) in ICP-SFMS. In a recent study,³⁵ an ICP-QMS device equipped with a hexapole collision cell was used to investigate uranium isotopic compositions in soils from the relocation zone of Belarus near the Chernobyl reactor. Application of ultrasonic nebulizer and membrane desolvator allowed the achievement of a relatively low uranium hydride formation rate of UH^+/U^+ of $\sim 2 \times 10^{-6}$. This performance was more than adequate for analysis of Chernobyl-area samples, which exhibited $^{236}\text{U}/^{238}\text{U}$ isotopic ratio ranging from 0.058 to 0.972×10^{-3} .

The determination of low $^{236}\text{U}/^{238}\text{U}$ isotopic ratios would benefit from attenuation of UH^+ formation: therefore, in the present study, D_2O (heavy water) as a solvent was investigated because a suppression of UH^+ formation using D_2O relative to H_2O was expected. The abundance sensitivity and figures of merit of ICP-SFMS were investigated in both solvents using a U standard reference, and the developed method in D_2O was applied to measurements of $^{236}\text{U}/^{238}\text{U}$ isotopic ratios in natural samples from Israel.

Experimental

ICP-MS instrumentation

A double-focusing sector field ICP-MS with reverse Nier–Johnson geometry (ICP-SFMS, ELEMENT, Thermo Finnigan MAT, Bremen, Germany) was used for $^{236}\text{U}/^{238}\text{U}$ isotopic ratio measurements in the Central Division of Analytical Chemistry at Research Centre Juelich. For solution introduction, three nebulizers were utilized: Meinhard concentric glass nebulizer (CGN-A, JE Meinhard Associates, Inc., Santa Ana, CA, USA); microconcentric nebulizer (MCN) equipped with membrane desolvation system (Aridus, Cetac Technologies, Inc., Omaha, NE, USA); ultrasonic nebulizer with a membrane desolvation system (USN U-6000AT⁺, Cetac Technologies, Inc). Aqueous solutions were introduced in the continuous flow mode using a peristaltic pump (Perimax 12, Spetec GmbH, Erding, Germany). The ICP torch was shielded with a grounded platinum electrode (GuardElectrode[®], Finnigan MAT).

Isotopic measurements were conducted at the Geological Survey of Israel with a double-focusing sector field MC-ICP-MS with a forward Nier–Johnson geometry (Nu-Instruments, UK). For solution introduction a microconcentric nebulizer equipped with a desolvation system (Cetac Aridus) was used. Measurements at Northern Arizona University were conducted using a VG Axiom double focusing sector field MC-ICP-MS, operated in the single collector (electron multiplier) mode.

The Axiom MC-ICP-MS is equipped with a Cetac U-5000AT ultrasonic nebulizer with 5 °C water condenser, and is operated at an uptake rate of 0.5 mL min^{-1} .

Standards and reagents

NIST standard reference materials U020, U350 and U500 (National Institute of Standards and Technology, USA) and solutions of uranium CCLU-500 (Laboratory Standard, Nuclear Research Center, Prague, Czech Republic³⁶) were used for the optimization of isotope ratio measurement of uranium. Uranium isotope ratios for the CCLU-500 standard have been established ($^{234}\text{U}/^{238}\text{U} = 0.011122$, $^{235}\text{U}/^{238}\text{U} = 0.99991$, $^{236}\text{U}/^{238}\text{U} = 0.002789$) by calibration against the NIST-500 SRM using TIMS. Single-element standard stock solutions of U and Th were obtained from Merck (Darmstadt, Germany). The solutions were diluted either with high purity deionized water (18 M Ω), obtained from a Millipore Milli-Q-Plus water purifier (Millipore Bedford, MA, USA) or with deuterium oxide (obtained from Merck, purity 99.95%) to the concentration of up to 100 ng ml^{-1} for determining the isotopic ratios of uranium and thorium by ICP-MS. All solutions were acidified to 1% HNO_3 with sub-boiled nitric acid. In case of dilution with D_2O , the final purity of deuterium after adding the standards and acidifying was $\sim 99.90\%$.

Samples and sample preparation

Two samples with natural uranium isotopic pattern were received from Israel. Sample I was NIST SRM NBL112a pure uranium metal from the natural uranium mineral. Sample II was a natural uranium mineral pitchblende (U_2O_3) from the precambrian mineralization in Timna mines, southern Israel. The investigated samples were diluted with heavy water.

Optimization of ICP-SFMS and measurement protocol

The Element ICP-MS instrument was optimized for maximum $^{238}\text{U}^+$ signal and minimum background at m/z 236 u using a 1 ng ml^{-1} U solution of natural isotopic composition. Instrument operation conditions are summarized in Table 1. Further details about the instrumentation and measurement procedure used are described elsewhere.^{27,37} The measured uranium isotopic ratio was corrected to account for UH^+ formation, mass bias, and detector dead-time. The mass bias factor (assuming an exponential correction²⁹) was determined using a 5 ng ml^{-1} NIST U500 standard solution as described elsewhere.^{28,38} A dead-time detector is of great importance for accurate measurements of isotope ratios and was calculated using the method described by Ketterer *et al.*³⁹ The detector dead-time of the ICP-SFMS “Element” and the MC-ICP-MS “Axiom” were found to be 45 ns and 20 ns, respectively.

Results and discussion

Figures of merit of ICP-SFMS

Abundance sensitivity of ICP-SFMS. Owing to molecular hydride formation of $^{235}\text{U}^1\text{H}^+$ and a strong peak tailing of $^{238}\text{U}^+$ on mass 236 u ICP-SFMS measurements at low mass resolution are not sufficient for accurate $^{236}\text{U}/^{238}\text{U}$ determination. In this work, in order to improve the abundance sensitivity of ICP-SFMS for ^{236}U measurements, the medium mass resolution mode ($m/\Delta m = 4450$) was applied.

For the ICP-SFMS (Element), the abundance at mass 236 u was estimated based upon measurements performed with a ^{232}Th standard solution. All experimental parameters were first optimized in respect of the maximum of $^{238}\text{U}^+$ ion intensity. A $0.1 \text{ } \mu\text{g ml}^{-1}$ Th standard solution (^{232}Th abundance 100%) was used to obtain a statistically “true” peak tail for studying abundance sensitivity. This avoids measuring the combined

Table 1 Optimized operation condition for ICP-SFMS and MC-ICP-MS with several sample introduction systems

	ICP-SFMS (Element)			MC-ICP-MS Nu Plasma	MC-ICP-MS VG Axiom
	Meinhard	Aridus	USN	Aridus	USN
Solution uptake rate/mL min ⁻¹	0.58	0.1	2.0	0.08	0.50
Rf power/W	1250	1100	1100	1300	1280
Cooling gas flow rate/L min ⁻¹		18		13	13
Auxiliary gas flow rate/L min ⁻¹		1.4		1.1	0.85
Nebulizer gas flow rate/L min ⁻¹	0.91	1.2	1.2	0.7	0.74
USN heating temperature/°C	—	—	140	—	140
USN cooling temperature/°C	—	—	3	—	5
Spray chamber temperature/°C	—	70	—	70	20
Desolvator heating temperature/°C	—	160	160	160	—
Sweep gas flow/L min ⁻¹	—	3.6	5.3	3.2	—
Extraction lens potential/V		2000		—	4320
Mass window (%)		60		—	150
Scanning mode		Peak hopping		—	E-scan, Peak jump
Mass resolution, <i>m</i> / Δ <i>m</i>		4450		300	3450

influences of ²³⁵U⁺ and ²³⁸U⁺ at *m/z* 236. In order to minimize the effect of other limiting factors, such as possible contamination of the blank and molecular ion formation, the abundance sensitivity for the isotope with mass 232.0375 u for ²³²Th was studied at masses *m* ± 0.5, *m* ± 1.5, *m* ± 2.5, etc. This approach has the advantage that abundance sensitivity can be measured even when isobaric interferences are presented, so the peak tail is not affected by possible interferences at masses *m* ± 1 u, *m* ± 2 u, *m* ± 3 u, etc. Measured in medium resolution, the abundance ratio sensitivity for ²³²Th is presented in Fig. 1. The intensity of the ²³²Th⁺ was 16 Mcps.

The abundance sensitivity for two mass units below the tailing peak in ICP-SFMS was calculated using eqn. 1.

$$\text{Abundance sensitivity} \left(\frac{m}{m+2} \right) = \frac{\text{intensity on } m/z=230}{\text{intensity on } m/z=232} \quad (1)$$

Besides the peak tailing from ²³⁸U⁺, abundance sensitivity of

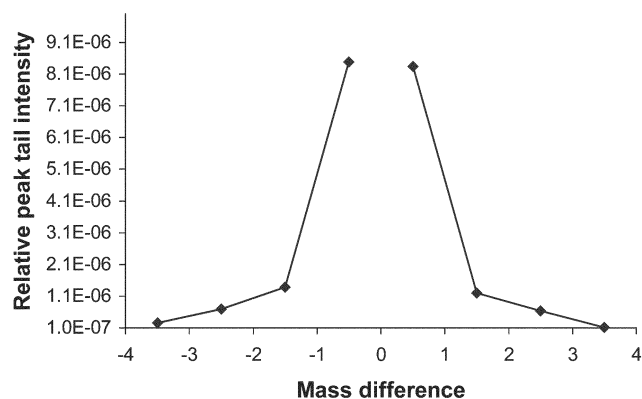


Fig. 1 Measured peak tail intensities at masses *m* ± *x_n* (*x_n* = 0.5 u, 1.5 u and 2.5 u) to peak intensity at mass *m* (*m* = 232.0375 u for ²³²Th).

ICP-SFMS for ²³⁶U determination with respect to ²³⁵U⁺ was also considered. Obtained values for two mass units below and one mass unit up the tailing peak were about 1.0 × 10⁻⁶ and 4.8 × 10⁻⁶, respectively (see Table 2). These values were further applied for the correction of ²³⁶U⁺.

Sensitivity for uranium measurement and limit of detection.

Sensitivities for uranium achieved on the ICP-SFMS in medium resolution mode using different nebulization system are summarized in Table 2. For the samples diluted with heavy water a maximum ²³⁸U⁺ intensity of 1770 Mcps ppm⁻¹ was observed when the ultrasonic nebulizer with membrane desolvator was applied. About a one order of magnitude lower signal (200 Mcps ppm⁻¹) was obtained using Meinhard nebulizer. Application of a microconcentric nebulizer with the Aridus membrane desolvation system yielded an intensity about 400 Mcps ppm⁻¹ for ²³⁸U⁺. Approximately the same values were observed when for dilution of the samples Milli-Q water was used.

The detection limits of ²³⁶U (3σ of the blank on the mass 236 u divided by the ²³⁸U⁺ ion intensity) calculated for different nebulization system are summarized in Table 2. The lowest limit of detection of ²³⁶U, 0.04 × 1 fg ml⁻¹, was achieved on the diluted with D₂O samples when the Aridus nebulizer with desolvator was used.

In order to evaluate the capability of ICP-SFMS for determining the ²³⁶U/²³⁸U isotopic ratio, minimum detectable ratio criteria were applied (see eqn. 2).

$$\text{Minimum detectable } ^{236}\text{U}/^{238}\text{U isotopic ratio} = \frac{\text{signal on } m/z=236 + 3\sigma}{\text{intensity on } m/z=238} \quad (2)$$

Fig. 2 presents the minimum detectable ratio (3σ) for the ²³⁶U/²³⁸U isotopic ratio of uranium, for different nebulizers,

Table 2 Figures of merit of ICP-SFMS (*m*/ Δ *m* = 4450) for several solution introduction devices measured on the samples diluted with heavy water

	Solution uptake rate [ml min ⁻¹]	Sensitivity of ²³⁸ U [Mcps ppm ⁻¹]	Uranium hydride rate, UH ⁺ /U ⁺	Abundance sensitivity		LOD (3σ) for ²³⁶ U [10 ⁻¹⁵ g ml ⁻¹]
				$\frac{m-2}{m}$	$\frac{m+1}{m}$	
Samples diluted with Milli-Q water						
Meinhard	0.58	205	1.05 × 10 ⁻⁴	1.06 × 10 ⁻⁶	4.9 × 10 ⁻⁶	0.41
USN with desolvator	2.0	1800	1.20 × 10 ⁻⁵	0.98 × 10 ⁻⁶	4.8 × 10 ⁻⁶	0.16
Aridus	0.1	400	1.00 × 10 ⁻⁵	0.98 × 10 ⁻⁶	4.8 × 10 ⁻⁶	0.13
Samples diluted with heavy water						
Meinhard	0.58	200	6.05 × 10 ⁻⁶	1.02 × 10 ⁻⁶	4.8 × 10 ⁻⁶	0.19
USN with desolvator	2.0	1770	1.10 × 10 ⁻⁶	0.98 × 10 ⁻⁶	4.8 × 10 ⁻⁶	0.09
Aridus	0.1	400	9.02 × 10 ⁻⁷	0.98 × 10 ⁻⁶	4.8 × 10 ⁻⁶	0.04

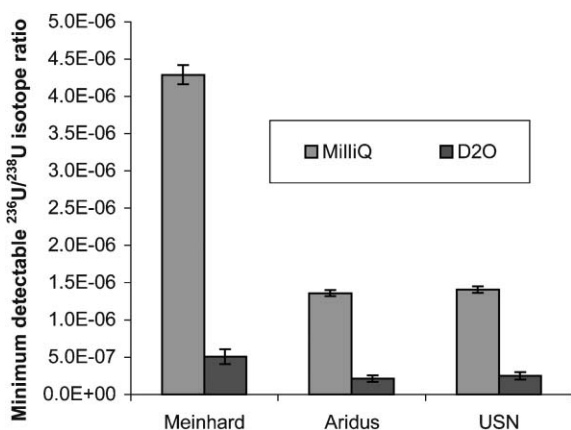


Fig. 2 Minimum detectable ratio criteria (3σ) for $^{236}\text{U}/^{238}\text{U}$ isotopic ratio of natural uranium for different nebulizers using H_2O and D_2O for dilution of the samples ($R = 4450$).

using Milli-Q and D_2O waters for dilution of the samples. A decrease in the minimum detectable ratio about of one order of magnitude was observed for all nebulizers. The largest effect was found with a Meinhard nebulizer (because no desolvator is used) due to significant elimination of $^{235}\text{U}^1\text{H}^+$ ions when heavy water for dilution is used.

Precision and accuracy of isotopic ratio measurement. Precision and accuracy of the method for $^{236}\text{U}/^{238}\text{U}$ isotopic ratio measurement ($n = 6$) was determined on isotope standard reference materials CCLU-500 ($^{236}\text{U}/^{238}\text{U} = 0.002789$) and NIST U-350 ($^{236}\text{U}/^{238}\text{U} = 0.002598$) using different nebulization systems (see Table 3). A precision in the range 0.14–0.45% (RSD) was yielded when a Meinhard nebulizer was applied. An improvement of one order of magnitude is possible using MC-ICP-MS.^{40,41} Measured $^{236}\text{U}/^{238}\text{U}$ isotopic ratios were ranged from 2.778×10^{-3} to 2.803×10^{-3} whereby accuracies -0.08% and -0.07% were achieved for CCLU-500 diluted with Milli-Q and heavy waters, respectively.

In general, a small deterioration in precision and accuracy was observed, when for dilution of samples D_2O water is used instead of Milli-Q water. This can be explained by insufficient acidification of heavy water that leads to adsorption of atoms of uranium on the walls of ICP parts and, therefore, less-stable solution introduction into the plasma.

Hydride formation rate of uranium. The hydride formation rate of uranium (UH^+/U^+) was studied in ICP-SFMS using H_2O and D_2O solvents. A typical ICP mass spectrum of uranium in the mass range of 237.5–241 u measured at medium mass resolution ($m/\Delta m = 4450$) for a $0.1 \mu\text{g ml}^{-1}$ solution of

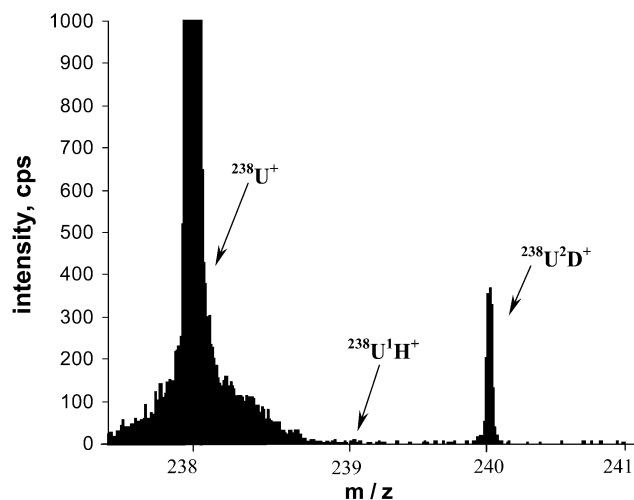


Fig. 3 ICP-MS spectrum of $^{238}\text{U}^1\text{H}$ at mass resolution of 4450 measured for natural uranium, diluted with D_2O using a Meinhard nebulizer.

uranium with natural isotope composition diluted in D_2O is shown in Fig. 3. Instead of UH^+ at mass 239 u (interferent in correct $^{239}\text{Pu}^+$ measurements⁴²), the formation of UD^+ ions was observed: therefore, using heavy water, the determination of ^{240}Pu will be disturbed. Results of the hydride formation rate UH^+/U^+ for different nebulizers using Milli-Q and heavy waters are summarized in Table 2. In the case of Milli-Q water as solvent, application of an ultrasonic nebulizer with a microporous Teflon membrane desolvator allowed reduction of the UH^+/U^+ ratio down to 1.20×10^{-5} , compared with 1.05×10^{-4} for the Meinhard nebulizer (without desolvator). The lowest hydride formation rate 1.0×10^{-5} was achieved with a microconcentric nebulizer with Aridus desolvator, reducing effectively the formation rate of uranium hydride ions UH^+ or possible isobaric interference of $^{235}\text{U}^1\text{H}^+$ on $^{236}\text{U}^+$ by factor of ten in comparison with the Meinhard nebulizer.

Significant reduction of the hydride formation rate for the Meinhard nebulizer was found when, instead of Milli-Q water, D_2O was applied for dilution of the samples. The obtained value of the UH^+/U^+ ratio was 6.05×10^{-6} , which was almost two orders of magnitude lower than the hydride formation rate, achieved with Milli-Q water. The lowest formation of uranium hydride molecular ions [$9.02 \pm 0.3 \times 10^{-7}$] was observed for the Aridus microconcentric nebulizer with membrane desolvator. The decreasing formation ratio for nebulizers with desolvators (USN and Aridus, 10 fold and 11 fold, respectively) is lower than for the Meinhard nebulizer (17 fold) for the samples diluted with heavy water. In our

Table 3 Results of $^{236}\text{U}/^{238}\text{U}$ isotopic ratios of uranium on CCLU-500 and NIST U350 standard reference materials diluted with Milli-Q and heavy water and their relative standard deviation ($n = 6$)

Standard reference material	Nebulizer	Measured $^{236}\text{U}/^{238}\text{U}$ ratio	RSD (%)	Recommended $^{236}\text{U}/^{238}\text{U}$ ratio	Accuracy (%)
Dilution with Milli-Q water					
CCLU-500	Meinhard	2.787×10^{-3}	0.26	2.789×10^{-3}	-0.08
	Aridus	2.777×10^{-3}	0.88		-0.42
	USN with desolvator	2.775×10^{-3}	0.92		-0.49
NIST U350	Meinhard	2.594×10^{-3}	0.14	2.598×10^{-3}	-0.15
	Aridus	2.592×10^{-3}	0.36		-0.22
	USN with desolvator	2.606×10^{-3}	1.01		0.29
Dilution with heavy water					
CCLU-500	Meinhard	2.787×10^{-3}	0.40	2.789×10^{-3}	-0.07
	Aridus	2.766×10^{-3}	0.96		-0.84
	USN with desolvator	2.772×10^{-3}	1.19		-0.61
NIST U350	Meinhard	2.590×10^{-3}	0.45	2.598×10^{-3}	-0.32
	Aridus	2.585×10^{-3}	0.54		-0.50
	USN with desolvator	2.630×10^{-3}	1.12		1.23

measurements was found (see Table 2) that even with application of nebulizers with a desolvation system (USN or Aridus) for measurement of the samples diluted with heavy water, the complete elimination of hydride formation is still not possible. The most probable reason for this could be a not "100%-pure" deuterium oxide solution (in present work 99.9%) used for dilution samples. Moreover, the formation of hydride ions is caused by hydrogen, as well as water, which are presented as an impurity in argon and residue gas, respectively. Further study of this effect will be of interest in order to decrease the uranium hydride formation and, therefore, improve ability to detect ^{236}U .

Experiments conducted using the Axiom ICP-SFMS also demonstrated significant diminution of the attenuation of UH^+/U^+ in D_2O solvent. The ratio $^{238}\text{U}^1\text{H}^+ / ^{238}\text{U}^+$ was measured for natural U solutions prepared in H_2O and D_2O ; the instrument was operated in the single collector (electron multiplier) mode with a mass resolution of $m/\Delta m = 3450$. The U concentrations were adjusted to produce $^{238}\text{U}^+$ signals of $\sim 1.2 \times 10^6$ cps; under these conditions the abundance sensitivity effects at m/z 339 are essentially negligible (< 1 cps), and the $^{238}\text{U}^1\text{H}^+ / ^{238}\text{U}^+$ can be measured without influence from peak tails. Data were collected in 400 s blocks using a rapid electrostatic scanning program. These experiments revealed a $^{238}\text{U}^1\text{H}^+ / ^{238}\text{U}^+$ ratio of $(3.3 \pm 0.4) \times 10^{-5}$ ($n = 6$) in H_2O solution; in D_2O solution a $^{238}\text{U}^1\text{H}^+ / ^{238}\text{U}^+$ ratio of $(1.1 \pm 0.1) \times 10^{-5}$ ($n = 16$) was observed. This represents a highly significant three-fold decrease in UH^+ formation when D_2O is used as the solvent. Ongoing work includes examining the potential for improvement in the determination of ^{239}Pu in D_2O solvent, where the interference of $^{238}\text{U}^1\text{H}^+$ is a well known problem.

$^{236}\text{U}/^{238}\text{U}$ isotope ratio in natural samples from Israel

The optimized method was applied for measurement of the $^{236}\text{U}/^{238}\text{U}$ isotopic ratio of two natural samples, received from Israel, which were diluted with heavy water. Comparative measurements were performed by ICP-SFMS and MC-ICP-MS. All results are summarized in Table 4.

Measured by ICP-SFMS, $^{236}\text{U}/^{238}\text{U}$ isotopic ratios ranged from $(4.8 \pm 0.9) \times 10^{-7}$ to $(5.3 \pm 0.8) \times 10^{-7}$ and from $(6.5 \pm 1.2) \times 10^{-7}$ to $(7.4 \pm 1.5) \times 10^{-7}$ for samples I and II, respectively, whereas MC-ICP-MS measurements of $^{236}\text{U}/^{238}\text{U}$ isotopic ratio yielded $5.00 \pm 0.09 \times 10^{-7}$ for both samples.

Conclusion

This study evaluates the capability of ICP-MS to measure the $^{236}\text{U}/^{238}\text{U}$ isotope ratio of uranium down to the level of 10^{-7} . In order to decrease the formation rate of UH^+/U^+ samples were diluted with heavy water. The measured uranium hydride formation rate for samples diluted with D_2O was about one order of magnitude lower for all nebulization systems in comparison with the samples diluted with conventional Milli-Q

Table 4 $^{236}\text{U}/^{238}\text{U}$ isotope ratio in natural samples, received from Israel, diluted with heavy water and measured by ICP-SFMS using different nebulizers and by MC-ICP-MS with Aridus nebulizer

Sample	ICP-MS	Nebulizer	Measured $^{236}\text{U}/^{238}\text{U}$ isotopic ratio	
I	ICP-SFMS	Meinhard	$(5.3 \pm 0.8) \times 10^{-7}$	
		Aridus	$(4.8 \pm 0.9) \times 10^{-7}$	
		USN with desolvator	$(5.0 \pm 1.3) \times 10^{-7}$	
II	MC-ICP-MS	Aridus	$(5.00 \pm 0.08) \times 10^{-7}$	
		ICP-SFMS	Meinhard	$(7.1 \pm 0.9) \times 10^{-7}$
			Aridus	$(6.5 \pm 1.2) \times 10^{-7}$
MC-ICP-MS	Aridus	USN with desolvator	$(7.4 \pm 1.5) \times 10^{-7}$	
		Aridus	$(5.00 \pm 0.09) \times 10^{-7}$	

water. Instead of UH^+ , formation of UD^+ was observed. The abundance sensitivity of ICP-MS was improved using the medium mass resolution mode in comparison to low mass resolution. Measurement of $^{236}\text{U}/^{238}\text{U}$ isotopic ratio in standard reference materials diluted with heavy water showed good accuracy and low relative standard deviation. Further study will be of interest to decrease the minimum hydride formation rate, using for acidification of the solution deuterium-based acid, e.g., D_2NO_3 .

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