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Fate of organic compounds during transformation of ferrihydrite in iron formations

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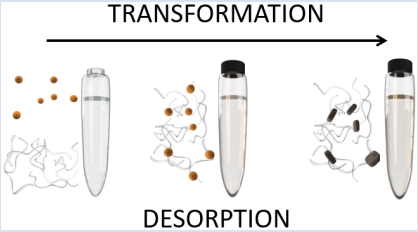
Abstract



The absence of organic compounds from Precambrian iron formations (IF) challenges the hypothesis of the biogenic origin of IF. Here we address the fate of adsorbed organic compounds during transformation from ferrihydrite to hematite. We determined the binding energy between hematite and common molecular terminations found in extracellular polymeric substances and biofilms: carboxylic, hydroxyl and phosphate functional groups. We found that the bond between hematite and hydroxyl groups is approximately 2 times stronger than the bond between hematite-carboxyl and -phosphate groups. We transformed synthetic ferrihydrite to hematite at 90 °C in presence of glycerol, which has a high density of hydroxyl groups, and measured the amount of mineral associated glycerol before and after the transformation. We show that the transformation releases glycerol highlighting that organic compounds adsorbed at precursor ferrihydrite could be desorbed during the process of IF sedimentation and diagenesis. Our results suggest that the absence of organic compounds in IF should not be used as evidence against their biogenic origin.

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Introduction

Traditionally, iron formations (IFs) have been considered as abiotically generated chemical sediments but an increasing body of evidence suggests the active role of microbes in their precipitation (Koehler *et al.*, 2010). The role of anoxygenic phototrophic bacteria in the formation of IFs has been the object of speculation since the work of Garrels *et al.* (1973). An improved understanding of microbially induced Fe(II) oxidation highlights that oxidising bacteria, such as *Rhodobacter ferrooxidans*, could likely drive the formation of IFs (Kappler *et al.*, 2005). Further, mass balance calculations have shown that Fe(II)-oxidising phototrophic bacteria have the capacity to oxidise all Fe(II) from the Precambrian ocean, causing formation of iron (oxyhydr)oxide (FeOx) (Hegler *et al.*, 2008). However, the low concentration of organic compounds in IFs and their diagenetic or metamorphic derivatives, challenges the hypothesis of their biogenic origin (Klein, 2005).

Extracellular polymeric substances (EPS) promote FeOx nucleation (Sand *et al.*, 2020) where microbially formed FeOx are often found in close association with the EPS (Chan *et al.*, 2004). When EPS is encrusted with FeOx, the polymers are shed and new EPS are formed. This prevents the encrustation of the microbe itself (Phoenix *et al.*, 2000; Chan *et al.*, 2004). During IF formation, FeOx-EPS composites would have settled through the water column and been deposited on the sea floor. Some of the organic compounds were degraded by diagenetic and metamorphic processes (Köhler *et al.*, 2013; Posth *et al.*, 2013; Halama *et al.*, 2016). In addition, there is evidence that the depositional environment was most likely scarce in organic compounds (Dodd *et al.*, 2019). Thus, the majority of the organic compounds was likely lost prior to deposition.

Hematite is a major component of IFs found today (Konhauser *et al.*, 2017) but, initially, the FeOx in the FeOx-polymer complexes were most likely composed of ferrihydrite (Chan *et al.*, 2004). The transformation to hematite would have happened in aqueous solutions both before (Sun *et al.*, 2015) and after deposition on the seabed. Depending on the solution conditions, ferrihydrite to hematite transformation can involve intermediary phases such as lepidocrocite and goethite, *e.g.*, through Fe(II) catalysed transformation (Hansel *et al.*, 2003), or it can be direct (Cudennec and Lecerf, 2006). Transformation involving intermediary phases is a dissolution-precipitation process (Schwertmann and Murad, 1983) implying that the interface between ferrihydrite and organic compounds is eliminated. In this scenario, organic compounds would have been released to solution where they could have been re-adsorbed but less strongly bound to newly formed phases (Chen *et al.*, 2015), or released in the water column and subsequently degraded by various biotic and abiotic processes (Kleber *et al.*, 2015). The direct transformation from ferrihydrite to hematite is a solid-state transition where atoms move only locally to occupy new structural positions (Cudennec and Lecerf, 2006), without the loss of an interface with adsorbed complexes. Thus, the direct transformation of ferrihydrite to hematite is not necessarily accompanied by a removal of organic compounds. However, the Gibbs free energy of binding (ΔG_{bu})

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between ferrihydrite and EPS is larger than between hematite and EPS (Sand *et al.*, 2020) implying that the polymers were more likely to desorb from hematite than from ferrihydrite. In such a case, the absence of organic compounds in IFs cannot be an argument against their biogenic origin.

To understand better the fate of adsorbed organic compounds during direct transformation from ferrihydrite to hematite, we used dynamic force spectroscopy (DFS) to measure the energy of binding between hematite and representative organic functional groups found in EPS. Subsequently we identified the functional group least likely to desorb during transformation and carried out transformation experiments where we used thermogravimetric analysis (TGA) to measure the loss of organic compounds during transformation.

Materials and Methods

DFS. We used an Asylum MFP3D atomic force microscope (AFM) and functionalised gold coated AFM tips (MSCT, Bruker) with carboxyl (COO^-), hydroxyl (OH), and phosphate (HPO_4^-) headgroups pointing away from the AFM probe (Fig. 1a). Tips were functionalised following the protocol described in Jelavić *et al.* (2017) (details in SI).

We used the {0001} face for DFS because it is one of the most common hematite surfaces in the environment (Hartman, 1989) and because it is a predominant face on hematite produced by transformation of ferrihydrite (Cornell and Schwertmann, 2004). The specular hematite monocrystal {0001} face was cleaned prior to analysis following the protocol described in Jelavić *et al.* (2017) (details in the SI). For DFS measurements, we collected >500 force curves *per* experiment. The tip approaching velocity was set to 500 nms^{-1} and the retraction velocity varied between 10–10000 nms^{-1}. The trigger force was set to 100 pN and dwell time to 0.5 s. During the measurement, the head groups bound to the hematite surface (Fig. 1a) and we measured the forces applied to break the bond (*i.e.* the force curve, Fig. 1b). We fit the rupture forces *vs.* the loading rate to a multibond model (Friddle *et al.*, 2012) enabling calculation of ΔG_{bu} (Eq. S-1 to Eq. S-4). The measurements were done in 10 mM NaCl solution at pH = 5.6. We chose these conditions, rather than conditions at which the transformation experiments were conducted, to maximise the interaction forces between the tip and hematite surface (Lützenkirchen *et al.*, 2013; Newcomb *et al.*, 2017).

Transformation experiments. Ferrihydrite was synthesised using the method of Schwertmann and Cornell (2000). We used 0.5 M NaCl solution and pH = 7 as a proxy for Precambrian seawater. We chose glycerol (CH_2**OH**-CH**OH**-CH_2**OH**) as a model for an OH-rich molecule. Adsorption of polymers is a complex function of degree of branching, length and hydrophobicity (van Oss, 1997). The aim here is not to account for such variations but to isolate the effect of ΔG_{bu}, which is independent of those parameters, on the magnitude of desorption during transformation. Coprecipitation of ferrihydrite-polymer aggregates results in a high organic content compared to adsorbed polymers (Mikutta *et al.*, 2014) and can cause variations in grain size and aggregation. To avoid any influence from such variations and to avoid low transformation rates related to high surface coverage, we adsorbed glycerol to ferrihydrite in this study. We mixed 15 mg of ferrihydrite with 15 ml of 0.5 M NaCl and added glycerol to a final concentration of 0.5 % (Fig. 1c). In control samples, we omitted glycerol. Samples were shaken at 100 rpm overnight to equilibrate. The next day, a batch of samples was rinsed with 20 ml of rinsing stock solution (0.5 M NaCl adjusted to pH = 7 with 1 M NaOH) by ultra-centrifugation (equilibrated sample). Another batch was placed in the oven heated to 90 °C. The samples from the oven were sampled at specific time steps to follow the transformation pathway using X-ray diffraction (XRD). When sampled, a batch was rinsed with 20 ml of 0.5 M NaCl that was adjusted to pH = 7 (aged sample). All samples were freeze dried after rinsing.

XRD. Samples for XRD were washed with ultra-deionised water (resistivity >18.2 MΩcm). A volume of 1.5 ml of suspension was pipetted on the zero background Si holders and left to dry at room temperature. Such sample preparation results in large preferred orientation of anisotropic grains such as goethite and lepidocrocite and allows us to detect minute quantities of crystalline material in a poorly ordered matrix of ferrihydrite. We collected diffractograms in reflection mode on the Bruker D8 Advance instrument using Cu Kα radiation (λ = 1.5418 Å) operated at 40 kV and 40 mA, and a LynxEye detector. Diffractograms were collected from 10–90° 2θ with step size of 0.02° and 1.7 s counting time *per* step. The sample was spun at 20 rpm. We used 0.3° divergence and 3° antiscatter slits, 2.5° Soller slits on incident and diffracted beams and a 0.02 mm thick Ni-filter. The opening of detector window was 2.94°.

TGA. We used Netzsch TG 209 F1 Libra. Samples were heated at a rate of 10 °C min^{-1} from 30 to 1000 °C under N_2 atmosphere. ~15 mg of sample were placed into the Pt crucible and the weight loss was measured as a function of temperature.

TEM. Images were taken with a Philips CM 20 TEM equipped with a thermionic LaB_6 filament. We used an accelerating voltage of 200 kV. Samples were prepared by placing a droplet of sample suspension on a formvar coated TEM grid and left for 5–10 s. Subsequently, the grid was washed in ultra-deionised water and water droplets were removed with the edge of a paper towel.

Results and Discussion

Energy of binding. The force spectra (Fig. 2) show that the rupture force between hematite and the OH group is ~6 times higher than between COO^- and HPO_4^- groups (Fig. 2, Table S-1). Recalculated to ΔG_{bu} (Eq. S-4), that is an increase in ~2 kT. These binding energies suggest that the interaction is dominated by strong Van der Waals or weak electrostatic forces (Israelachvili, 2011). Considering that the {0001} hematite surface is completely hydroxylated in water (Trainor *et al.*, 2004), our results follow the trend where OH-OH bonds are twice as strong as COO^--OH bonds at circumneutral pH (Vezenov *et al.*, 1997) present in Precambrian ocean. Thus, irrespective of the transformation pathway, biopolymers rich in acidic groups desorb more easily from hematite than those rich in hydroxyl groups.

Transformation pathway. To investigate the fate of a ferrihydrite-associated OH-containing molecule during transformation to hematite, we transformed the ferrihydrite-glycerol complex, monitored the transformation pathway with XRD and determined the weight loss with TGA. The ferrihydrite-glycerol complex transformed directly to hematite without any intermediate phases detectable by XRD (Fig. S-3) suggesting a solid-state transformation pathway. As a control, we transformed ferrihydrite in absence of glycerol using the same solution conditions. Even though the hematite was the first phase to occur after 17 hr, goethite started forming after 43 hr (Fig. S-2) indicating that, in pure systems, some ferrihydrite dissolves and reprecipitates as goethite, as previously shown (Das *et al.*, 2011).

The transformation is accompanied by glycerol release. In general, the ferrihydrite-gycerol complex lost more weight than the pure ferrihydrite (Fig. 3, Table 1). Ferrihydrite is more

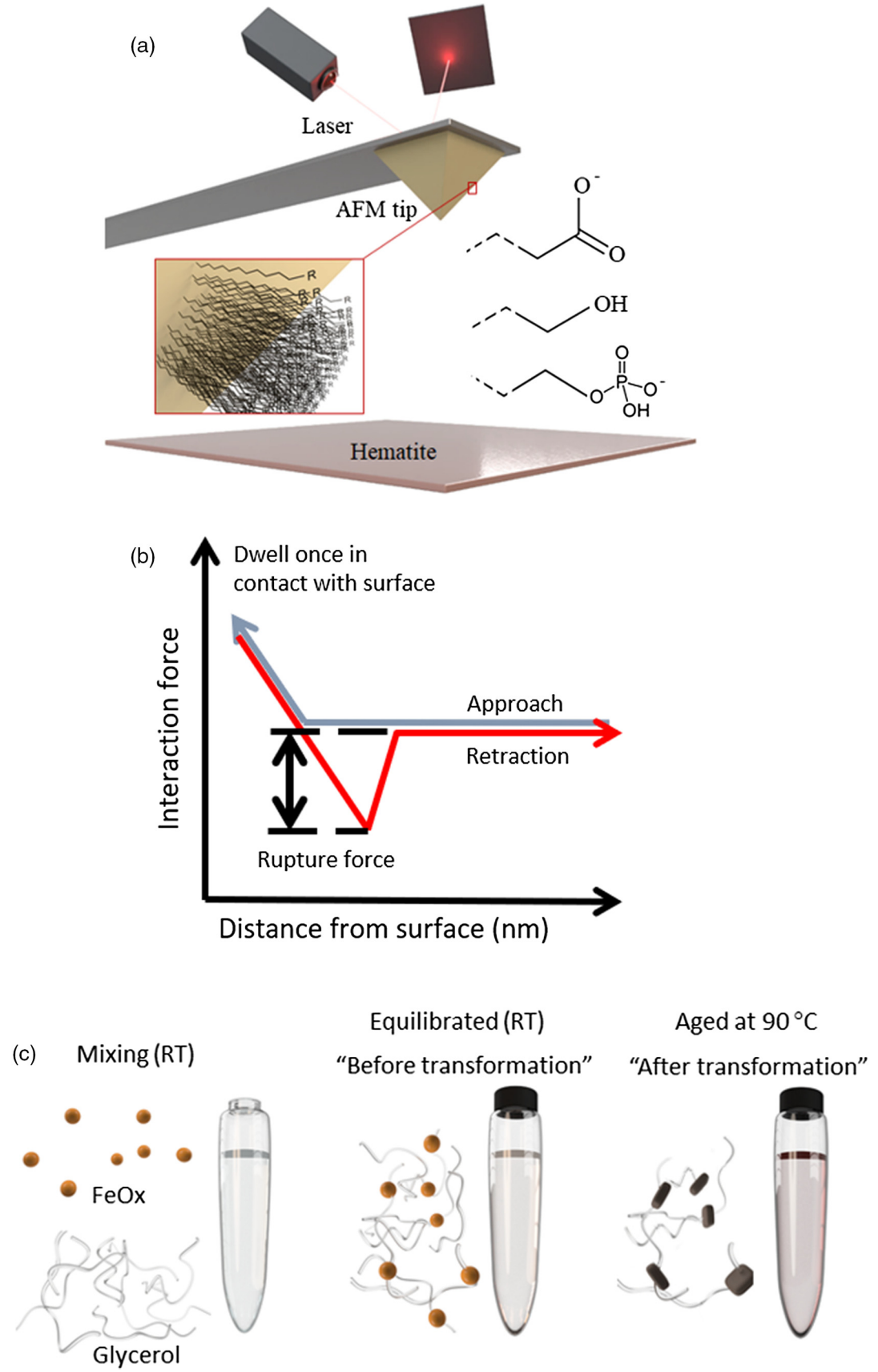

Figure 1 **(a)** Schematics of the DFS (Dynamic Force Spectroscopy): self assembled monolayers with carboxyl, hydroxyl and phosphate head-groups covalently bonded to AFM tip. **(b)** A scheme of a force curve. **(c)** In transformation experiments, ferrihydrite and glycerol were mixed at room temperature and left to equilibrate overnight. One sample was then taken for TGA (equilibrated sample) and the rest was placed in the oven at 90 °C until the transformation was complete (aged sample).

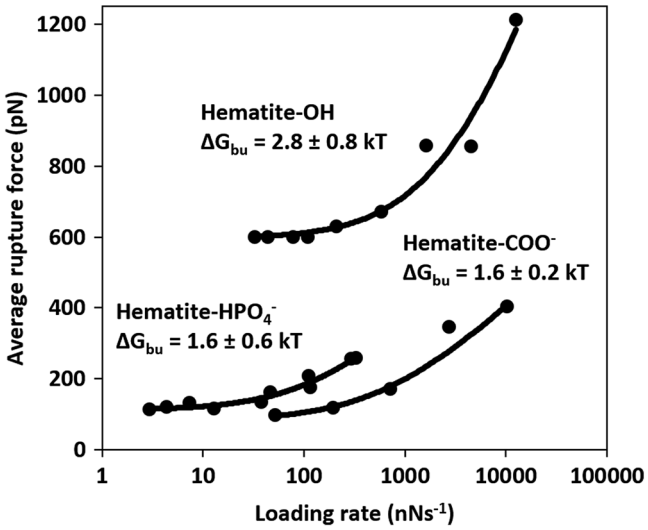

Figure 2 Multibond fit (Friddle *et al.*, 2012) to the dynamic force spectra between hematite surface and alkyl thiol self assembled monolayers with OH, HPO_4^- and COO^- head groups. The uncertainty represents the error propagated from the standard deviations of the fit. Loading rate is a nominal retraction velocity (nms^{-1}) multiplied with the spring constant of the cantilever (nNnm^{-1}).

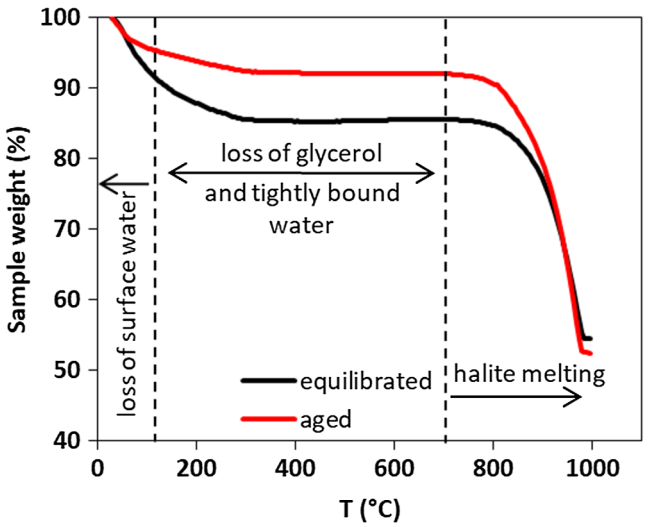

Figure 3 Comparison of thermogravimetry curves of ferrihydrite-glycerol (equilibrium-black curve) and resulting hematite-glycerol (aged-red curve). The smaller mass loss in the region 125–700 °C (dashed vertical lines) for the aged sample compared to the equilibrium sample suggests bigger loss of tightly bound water and glycerol during the transformation.

Table 1 Weight loss during the TGA analysis in <125 °C and 125–700 °C regions.

Sample		% weight loss	
		<125 °C	125–700 °C
Glycerol	equilibrated	8.9	5.4
	aged	4.8	3.2
Control	equilibrated	9.8	4.7
	aged	n.a.*	n.a.

* transformation yielded both goethite and hematite making comparison impossible

hydrated than hematite so the loss of loosely adsorbed water (<125 °C) and tightly bound water (125–700 °C) is higher for ferrihydrite than hematite (Hiemstra and Van Riemsdijk, 2009). Comparing the weight loss of ferrihydrite-glycerol (5.4 %; Table 1) to pure ferrihydrite (4.7 %) in the 125–700 °C range, shows that the tightly bound water accounts only for 87 % of the weight loss in this region, indicating that the remaining 13 % is the loss of glycerol and associated interfacial water. This implies that hematite produced by aging ferrihydrite-glycerol contains less glycerol than the original ferrihydrite-glycerol (equilibrated), and that glycerol is released during transformation. Considering that the initial amount of adsorbed glycerol is ~5 mg/g of ferrihydrite (Fig. S-1), TGA indicates that all glycerol has been desorbed during the transformation.

The reason for the loss of the glycerol during transformation can also be the smaller specific surface area (SSA) of the produced hematite (Fig. S-4) and not only the lower affinity for glycerol compared to ferrihydrite. From TEM images, we estimated the SSA of ferrihydrite to be 374–790 m^2g^{-1} and of hematite to be 10–46 m^2g^{-1}. Thus, the decrease in the SSA because of grain coarsening during transformation is between 9–70 times which alone might explain the decrease in glycerol content. However, the loss must be amplified by the lower ΔG_{bu} for hematite-OH than for ferrihydrite-OH system. In addition, our thermodynamic results suggest that the loss of organic compounds would have happened because of decreased ΔG_{bu} even if the subsequent grain coarsening does not occur, *e.g.*, in case of hematite growth inhibition caused by organic compounds. However, both scenarios are likely to have contributed to the loss of organic compounds during the formation of IFs.

We have demonstrated that glycerol, a molecule with a high ΔG_{bu} to hematite, desorbs during the transformation of ferrihydrite to hematite. We propose that a significant mass of organic compounds from FeOx-EPS composites is desorbed early in the process of FeOx sedimentation because of the transformation from ferrihydrite to hematite. This loss of organic compounds is probably further enhanced by the grain coarsening of hematite during diagenetic and metamorphic processes and concomitant reactions in the microenvironment that result in secondary mineralisation, *e.g.*, Posth *et al.* (2013). Our experiments were designed to determine the loss of strongly bound molecules from FeOx during the direct transformation, which is least likely to result in desorption of adsorbed organic compounds. The less strongly bound organic compounds or organic compounds adsorbed to ferrihydrite that transformed *via* dissolution-precipitation pathway would likely have desorbed in higher proportion than that we report here. Thus, our results highlight that the absence of organic compounds in IF should not be used as evidence against their biogenic origin.

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Additional Information

Supplementary Information accompanies this letter at http://www.geochemicalperspectivesletters.org/article2030.

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