

Supporting Information

Synthesis and Characterization of 3,5-Bis(di-*tert*-butylphosphinito)pyridine Pincer Complexes

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Nuclear Magnetic Resonance Spectra

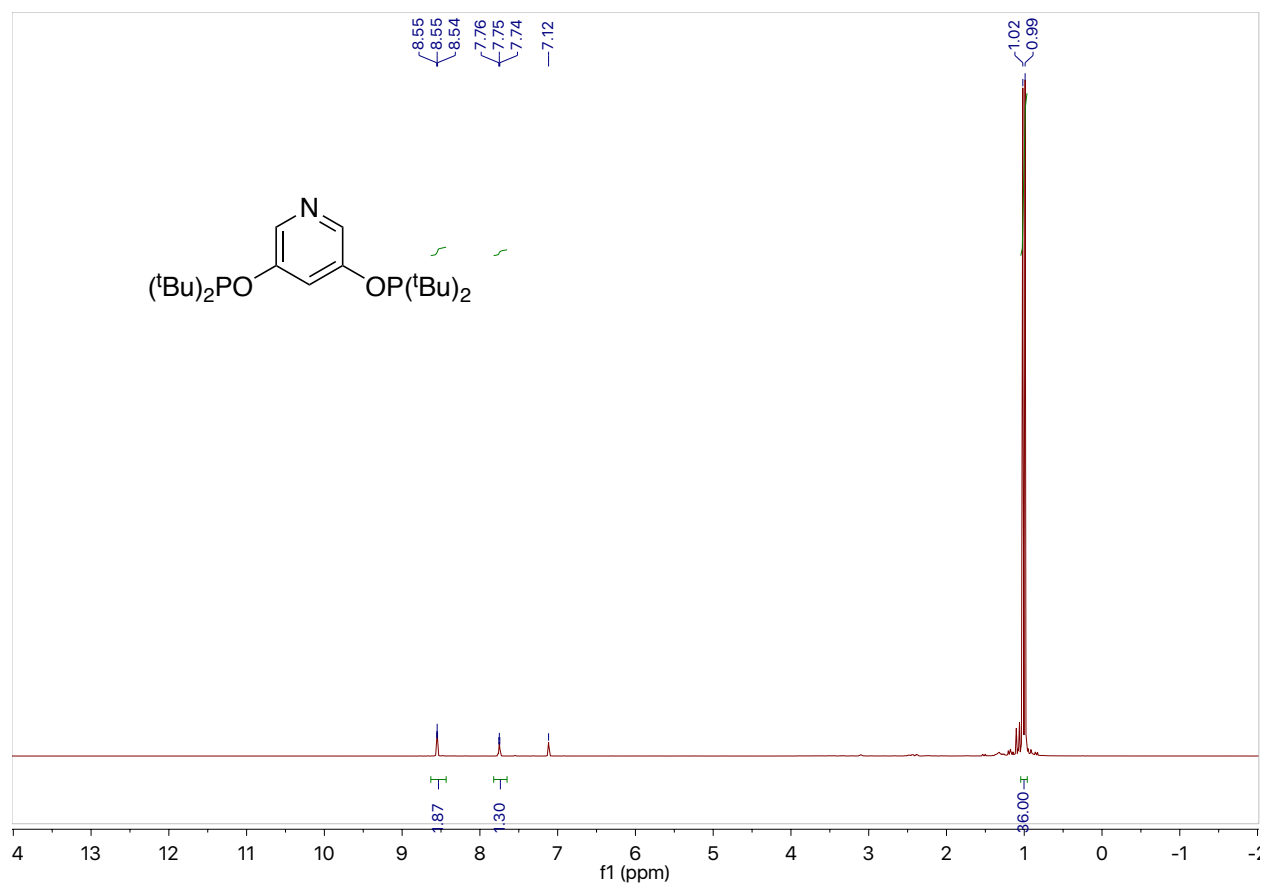


Figure S1. ^1H NMR (400 MHz, C_6D_6) spectrum of 3,5-bis(di-*tert*-butylphosphinito)pyridine, **1**.

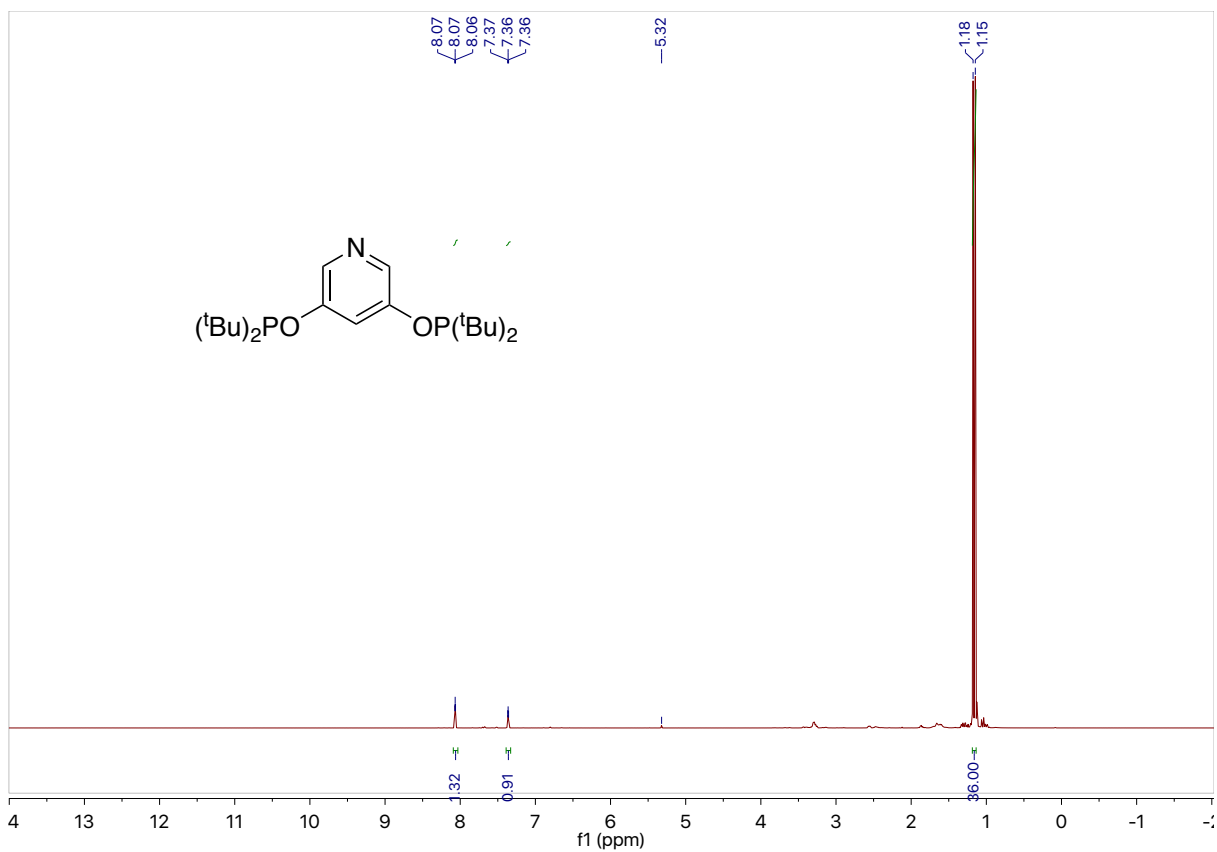


Figure S2. ¹H NMR (400 MHz, CD₂Cl₂) spectrum of 3,5-bis(di-*tert*-butylphosphinito)pyridine, **1**.

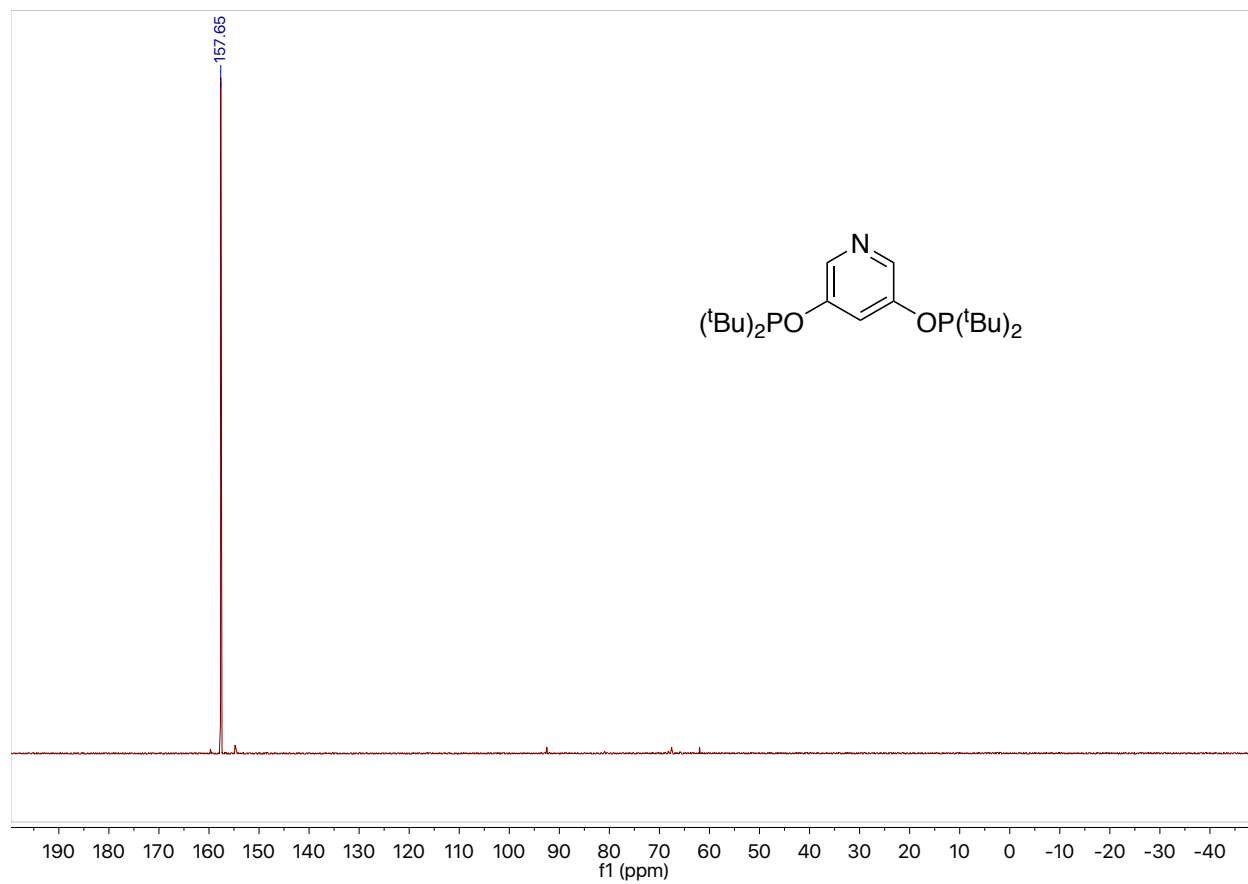


Figure S3. $^{31}\text{P}\{^1\text{H}\}$ NMR (162 MHz, C_6D_6) spectrum of 3,5-bis(di-*tert*-butylphosphinito)pyridine, **1**.

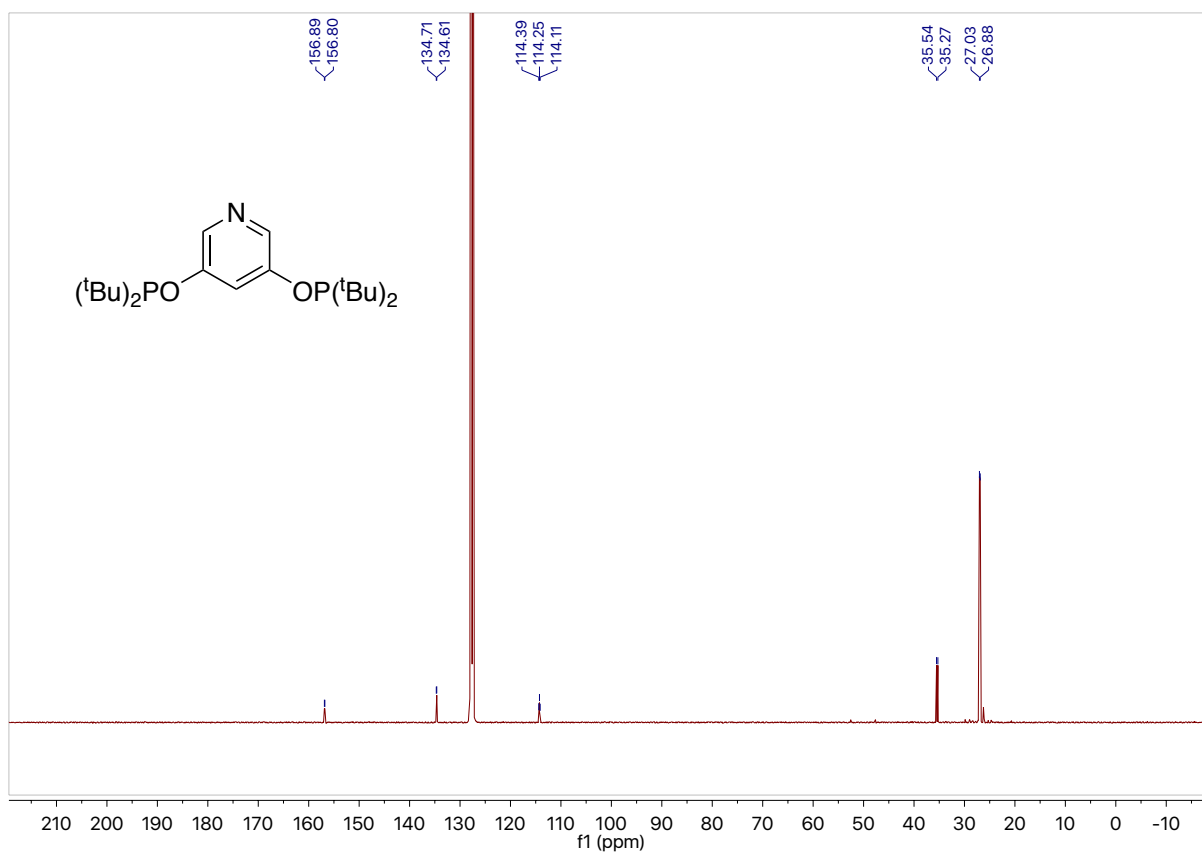


Figure S4. $^{13}\text{C}\{^1\text{H}\}$ NMR (101 MHz, C_6D_6) of 3,5-bis(di-*tert*-butylphosphinito)pyridine, **1**.

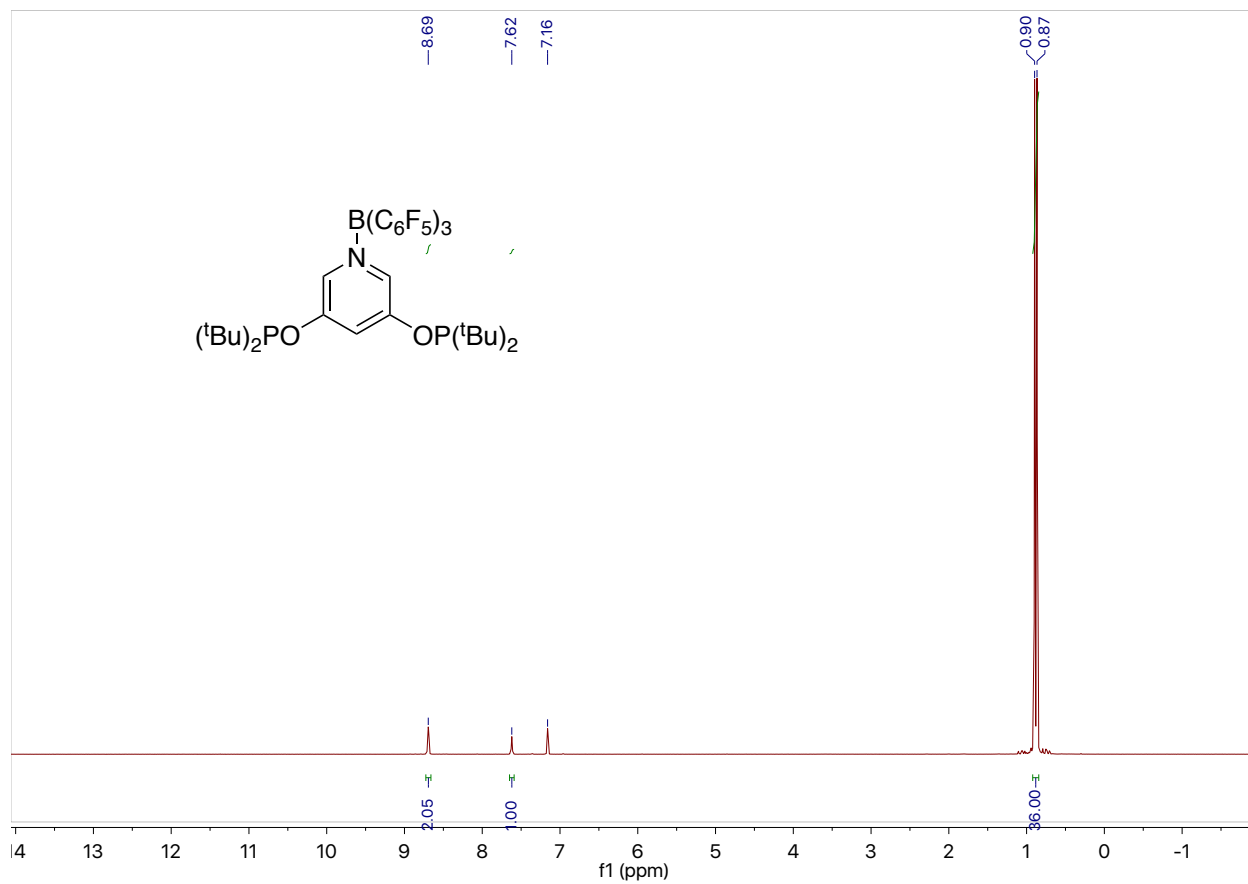


Figure S5. ¹H NMR (400 MHz, C₆D₆) spectrum of *N*-tris(pentafluorophenyl)borane-3,5-bis(di-*tert*-butylphosphinito)pyridine, **2**.

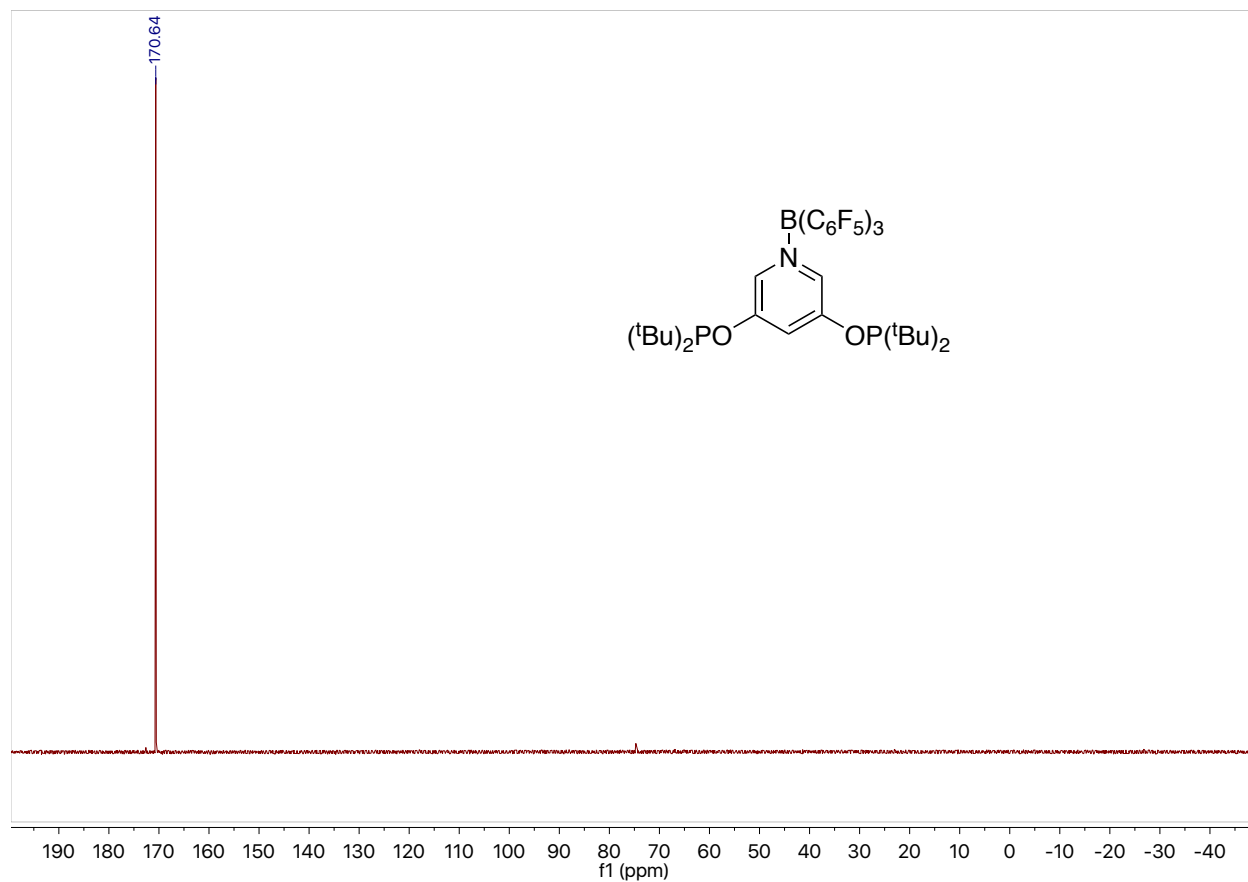


Figure S6. $^{31}\text{P}\{^1\text{H}\}$ NMR (162 MHz, C_6D_6) spectrum of *N*-tris(pentafluorophenyl)borane-3,5-bis(di-*tert*-butylphosphinito)pyridine, **2**.

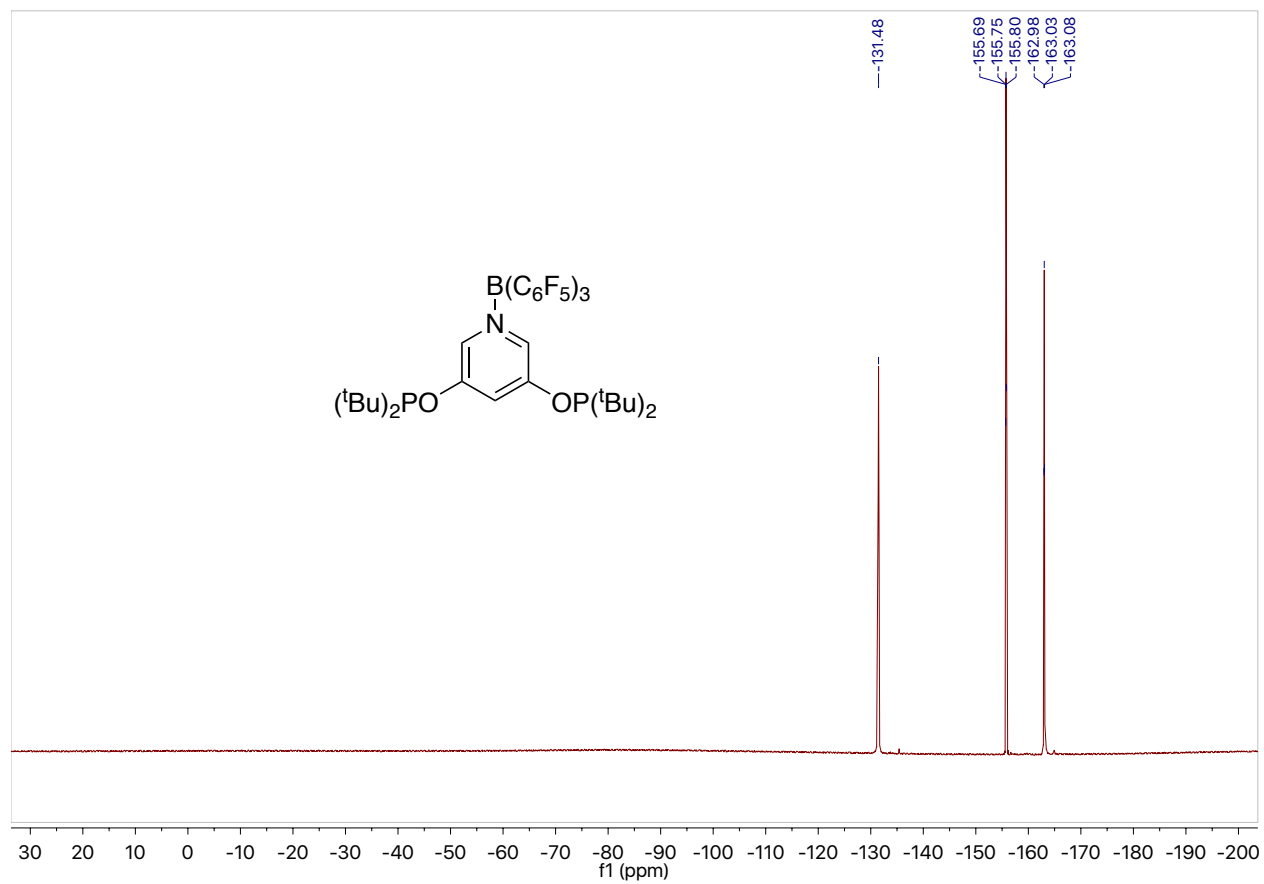


Figure S7. $^{19}\text{F}\{^1\text{H}\}$ NMR (376 MHz, C_6D_6) spectrum of *N*-tris(pentafluorophenyl)borane-3,5-bis(di-*tert*-butylphosphinito)pyridine, **2**.

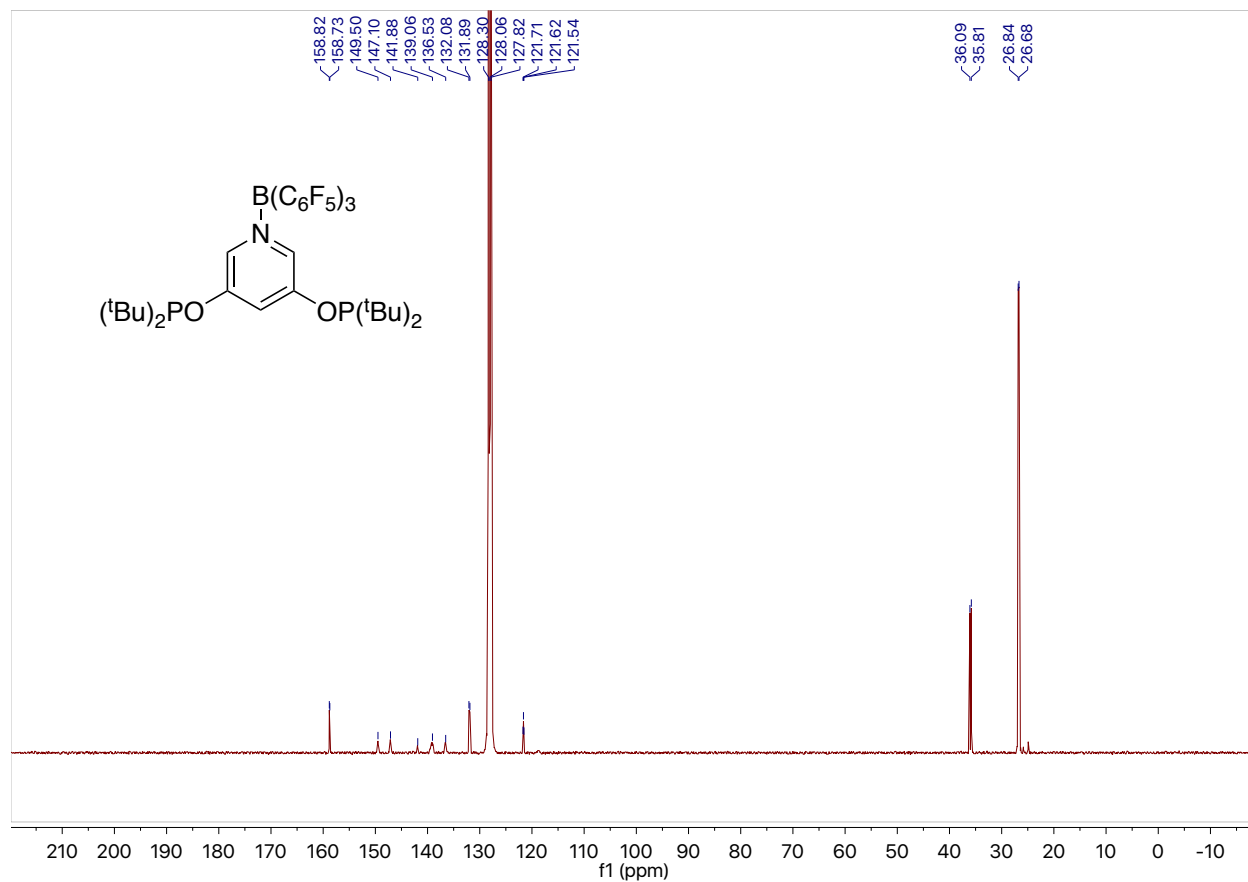


Figure S8. ¹³C{¹H} NMR (101 MHz, C₆D₆) of *N*-tris(pentafluorophenyl)borane-3,5-bis(di-*tert*-butylphosphinito)pyridine, **2**.

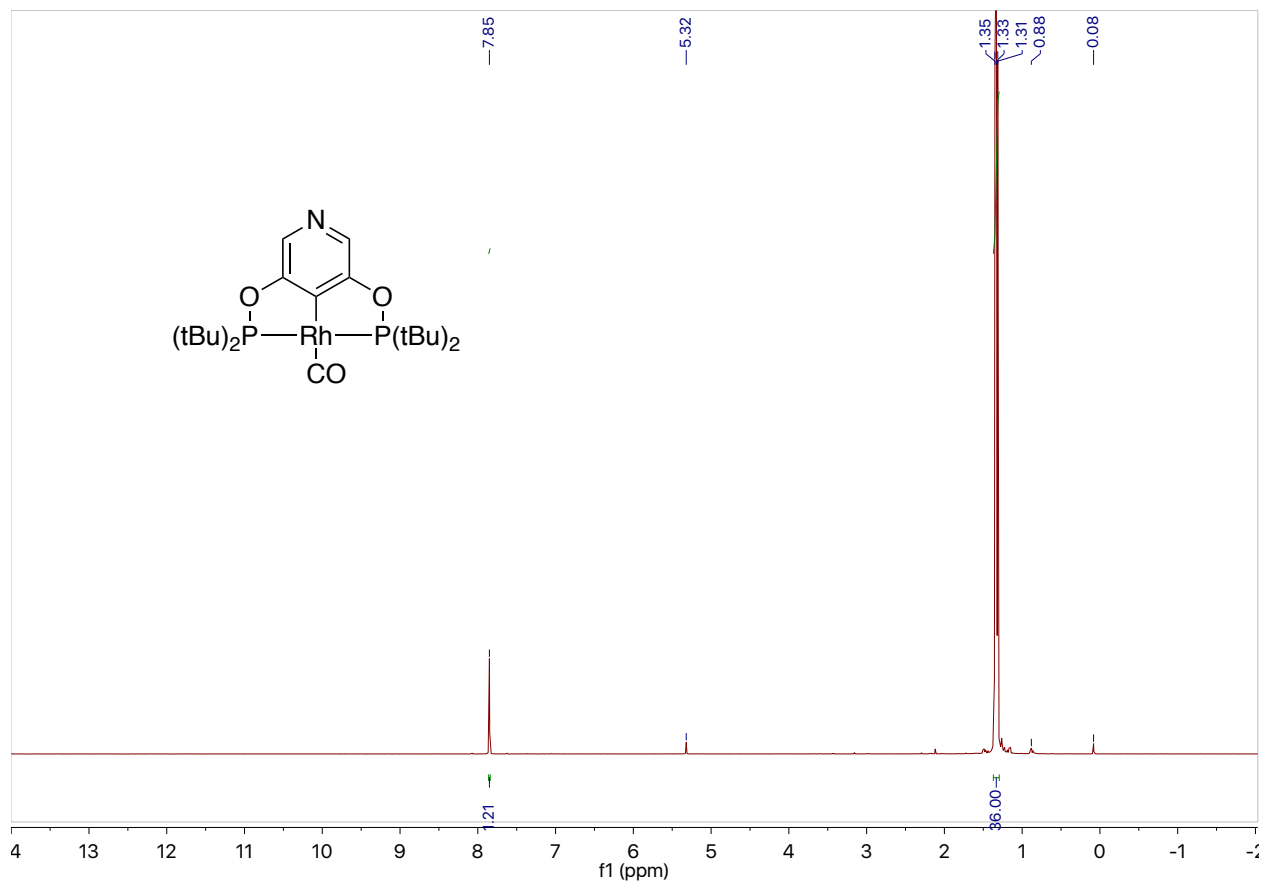


Figure S9. ^1H NMR (400 MHz, CD_2Cl_2) spectrum of (PyPOCOP)Rh(CO), **3**. Contamination by small amounts of silicone grease and pentane is present.

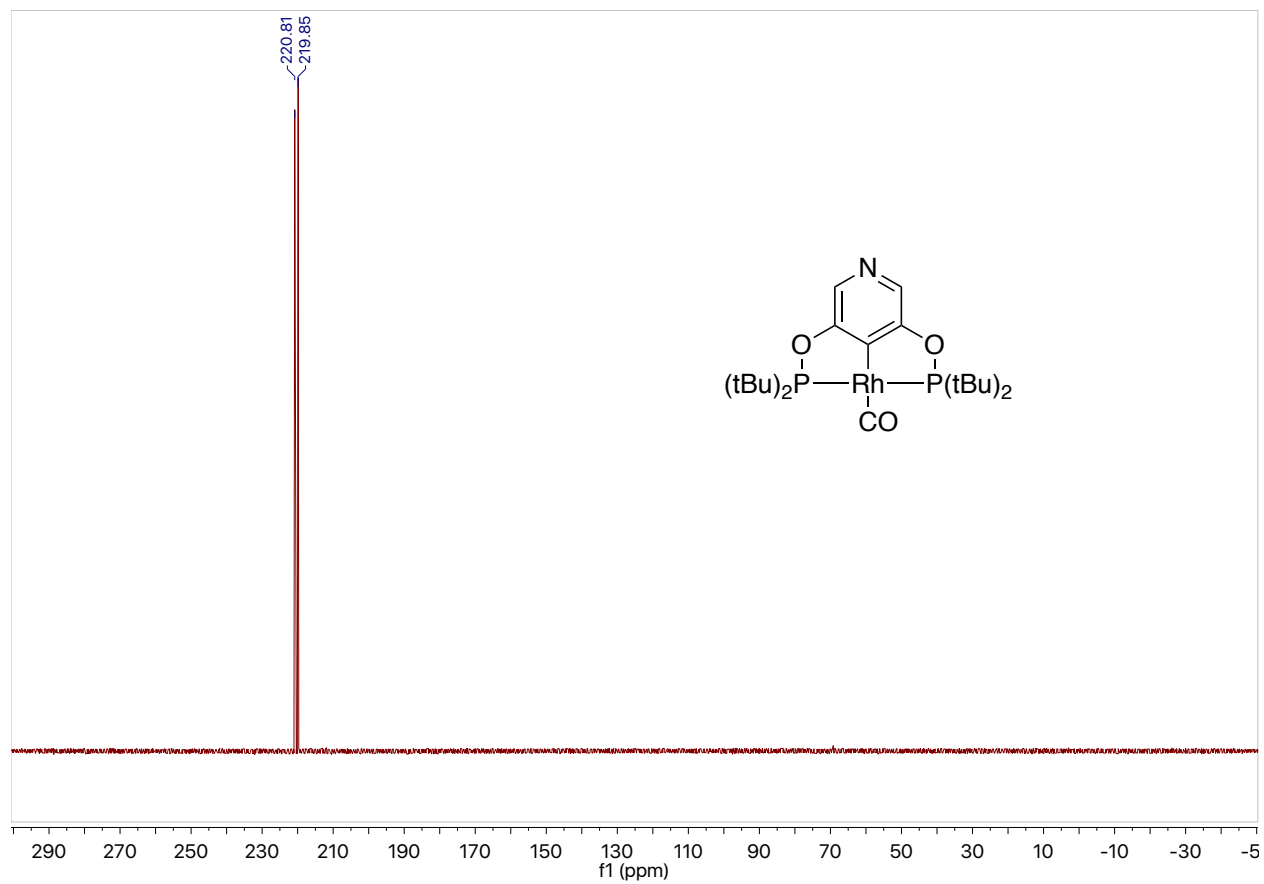


Figure S10. $^{31}\text{P}\{^1\text{H}\}$ NMR (162 MHz, CD_2Cl_2) spectrum of (PyPOCOP)Rh(CO), **3**.

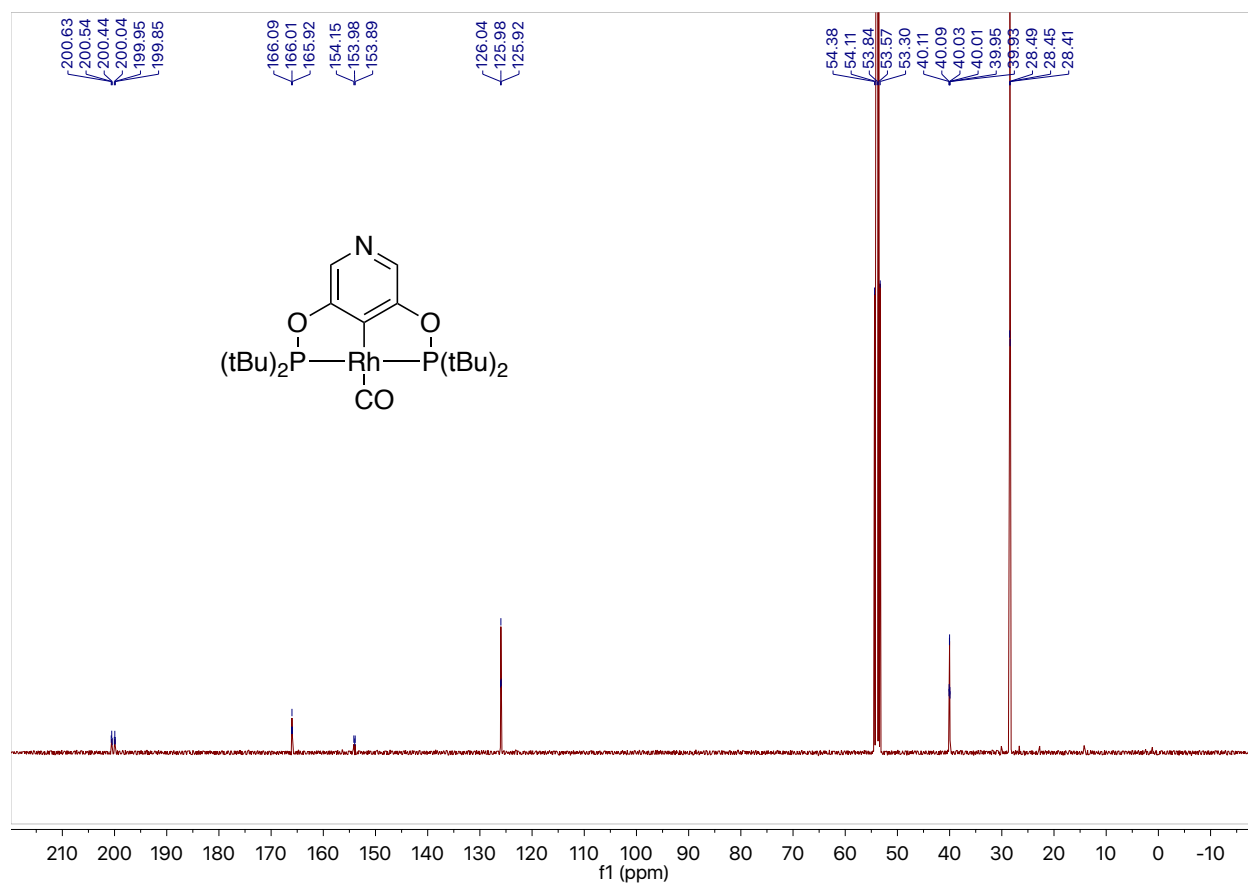


Figure S11. $^{13}\text{C}\{^1\text{H}\}$ NMR (101 MHz, CD_2Cl_2) of (PyPOCOP)Rh(CO), 3.

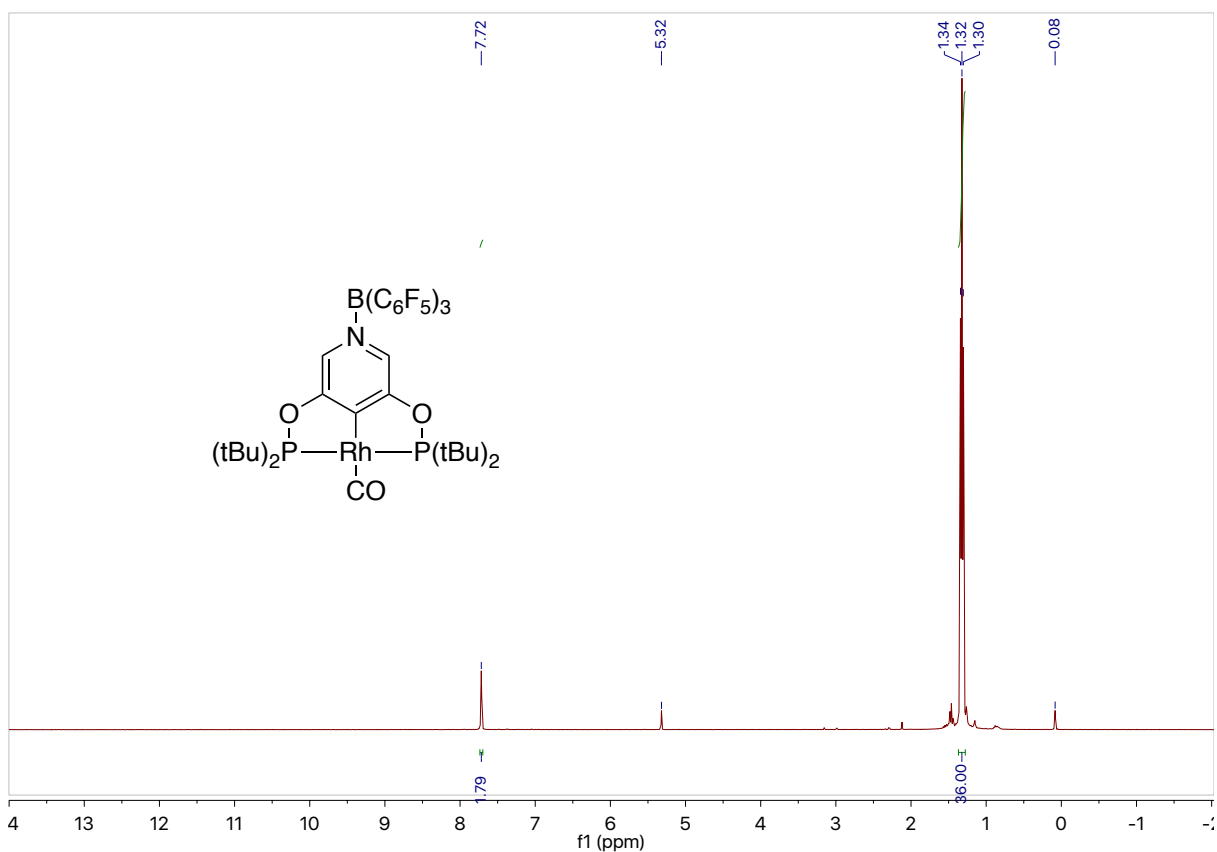


Figure S12. ^1H NMR (400 MHz, CD_2Cl_2) spectrum of (BCF-PyPOCOP)Rh(CO), **4**. Contamination by a small amount of silicone grease is present.

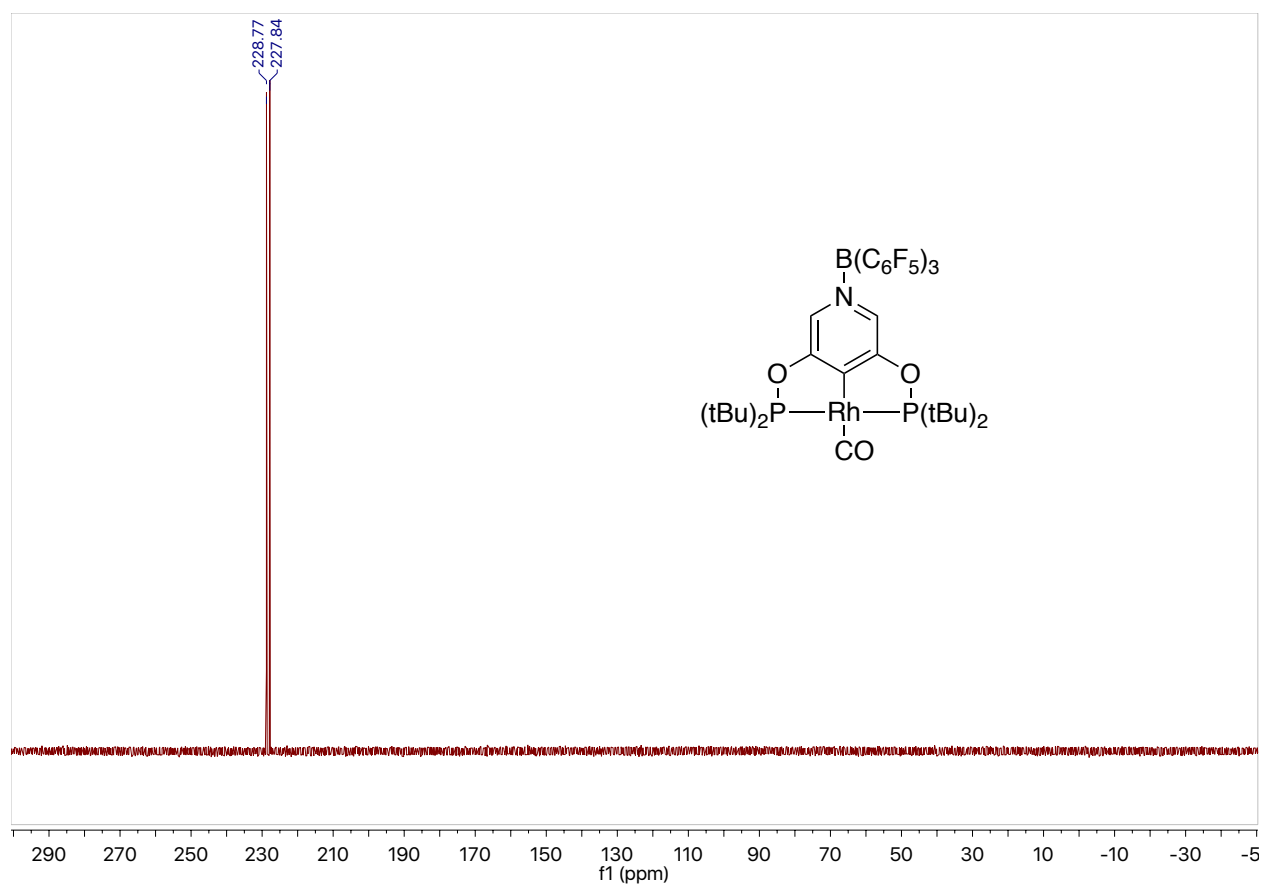


Figure S13. $^{31}\text{P}\{^1\text{H}\}$ NMR (162 MHz, CD_2Cl_2) spectrum of (BCF-PyPOCOP)Rh(CO), **4**.

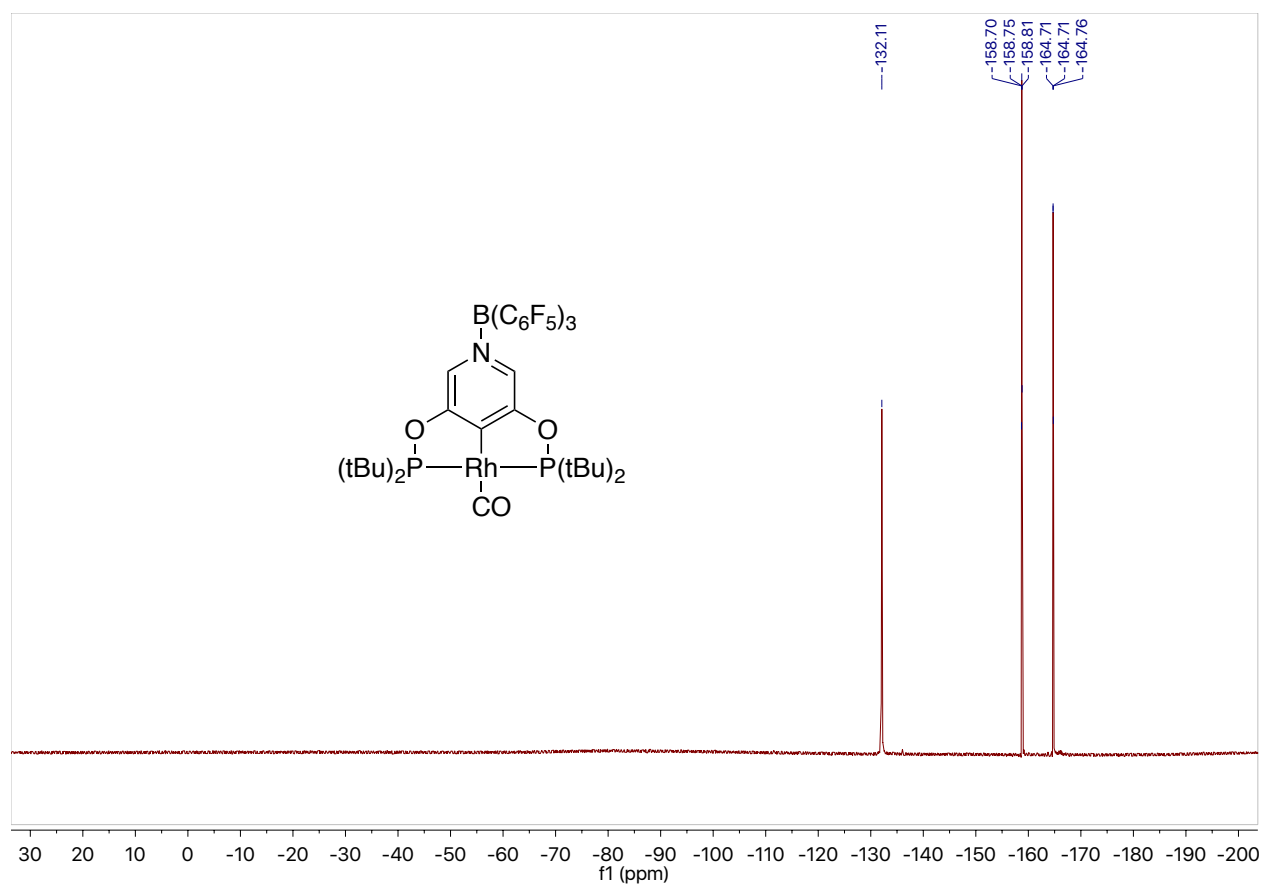


Figure S14. $^{19}\text{F}\{^1\text{H}\}$ NMR (376 MHz, CD_2Cl_2) spectrum of (BCF-PyPOCOP)Rh(CO), **4**.

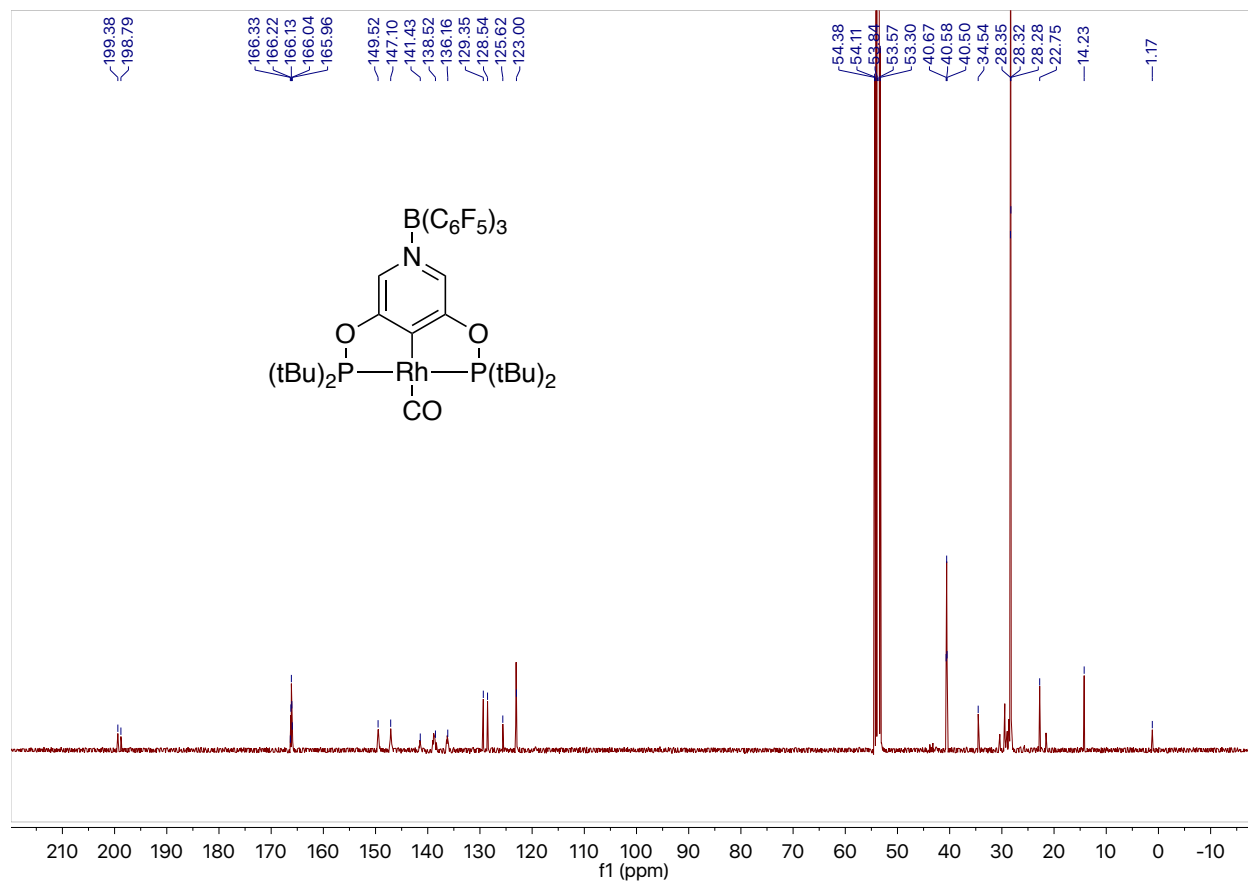


Figure S15. $^{13}\text{C}\{^1\text{H}\}$ NMR (101 MHz, CD_2Cl_2) of (BCF-PyPOCOP)Rh(CO), 4.

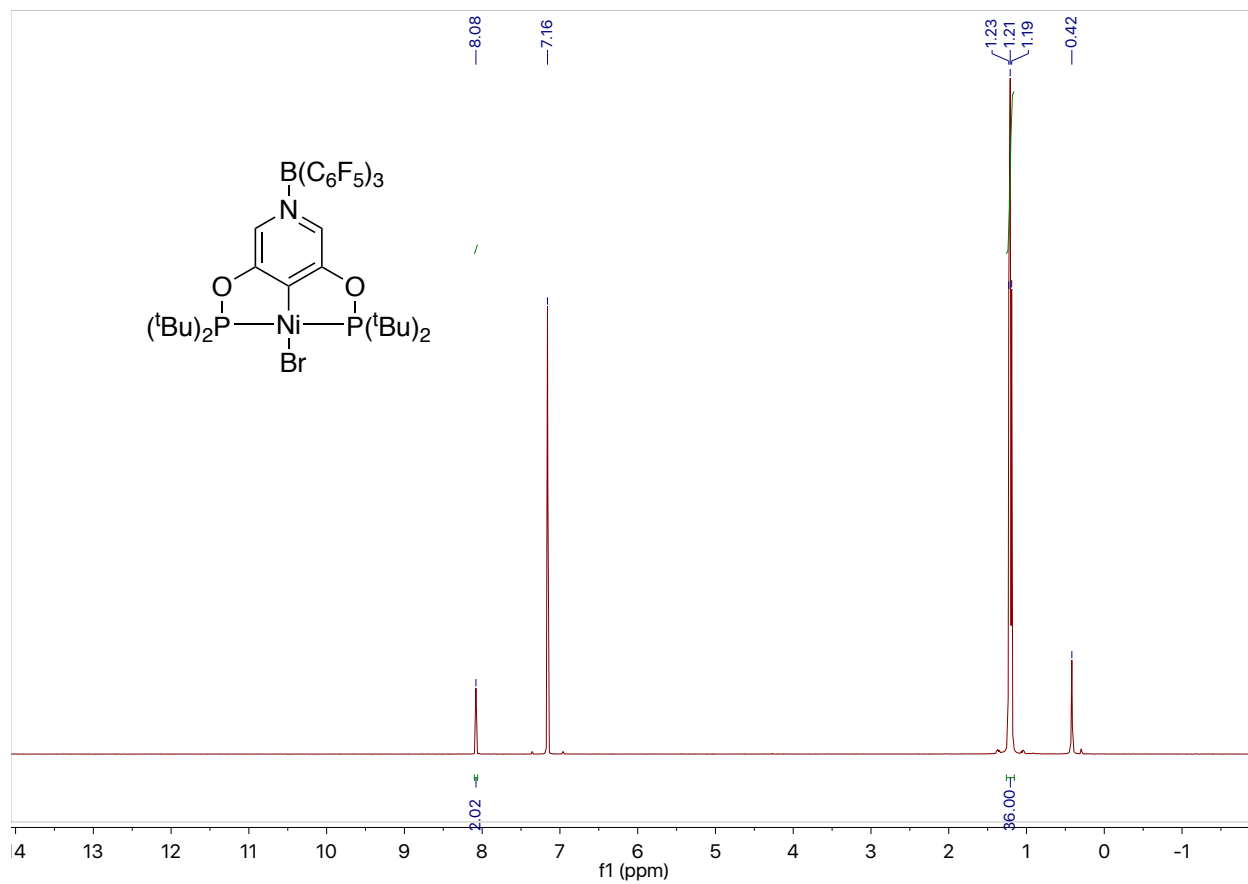


Figure S16. ^1H NMR (400 MHz, C_6D_6) spectrum of (BCF-PyPOCOP)NiBr, **5**. Contamination by water is present.

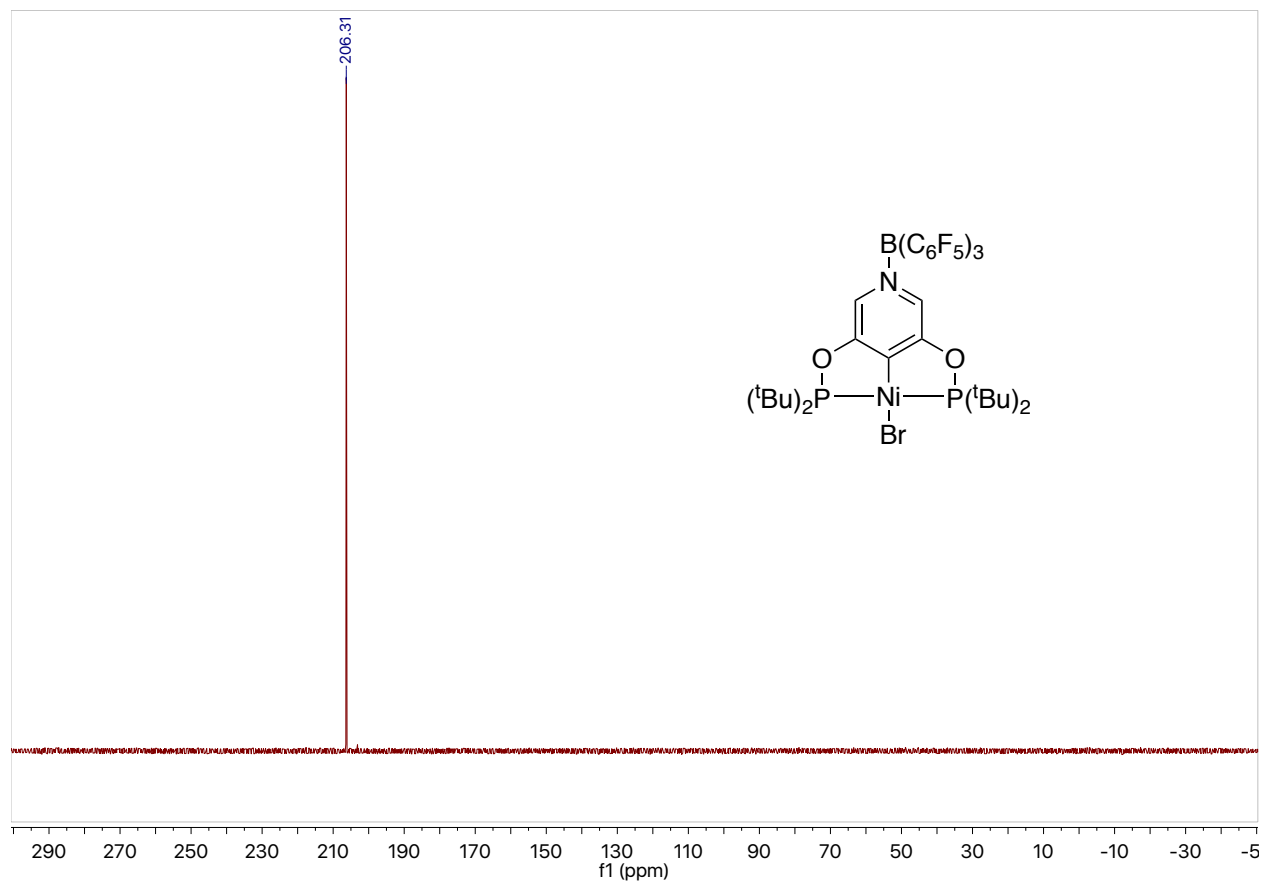


Figure S17. $^{31}\text{P}\{^1\text{H}\}$ NMR (162 MHz, C_6D_6) spectrum of (BCF-PyPOCOP)NiBr, **5**.

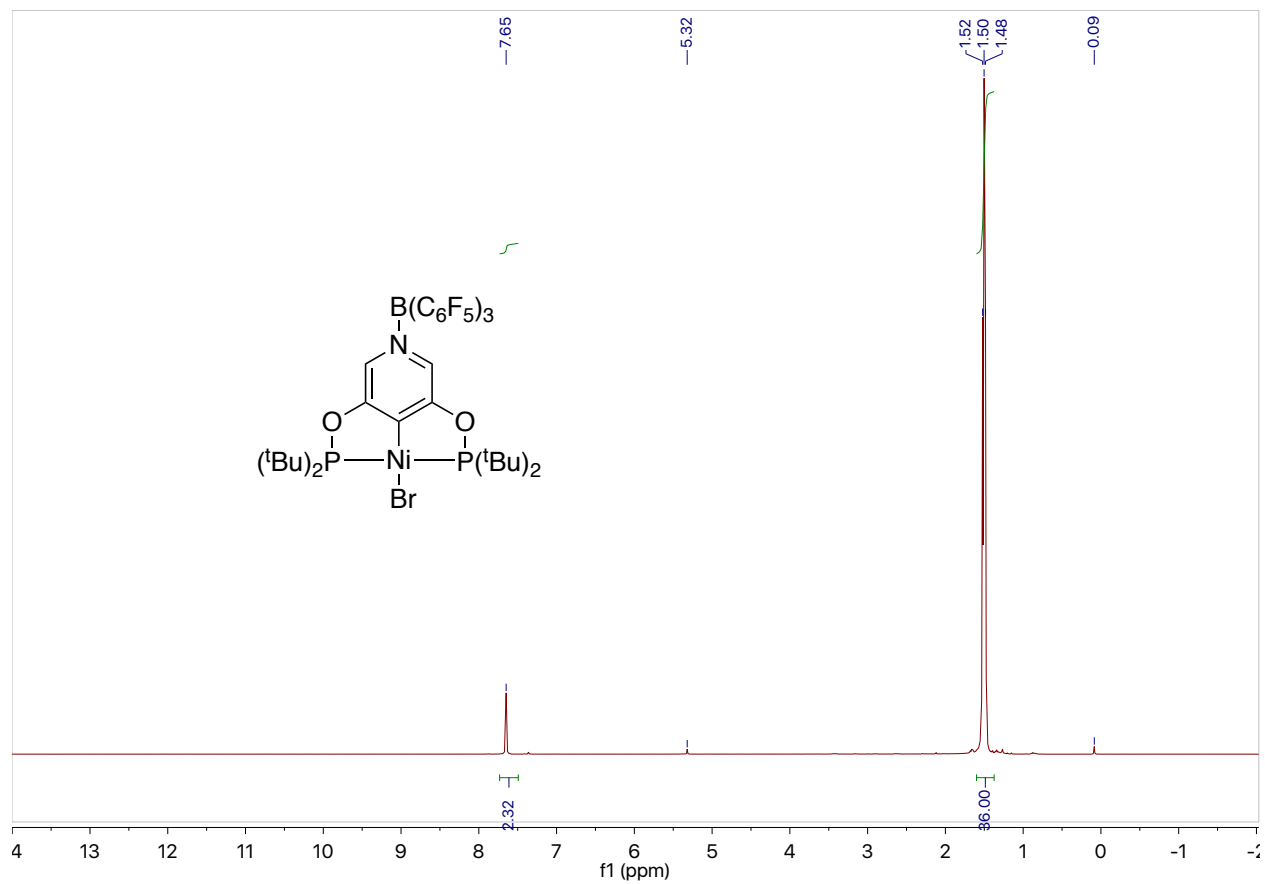


Figure S18. ^1H NMR (400 MHz, CD_2Cl_2) spectrum of (BCF-PyPOCOP)NiBr, **5**. Contamination by a small amount of silicone grease is present.

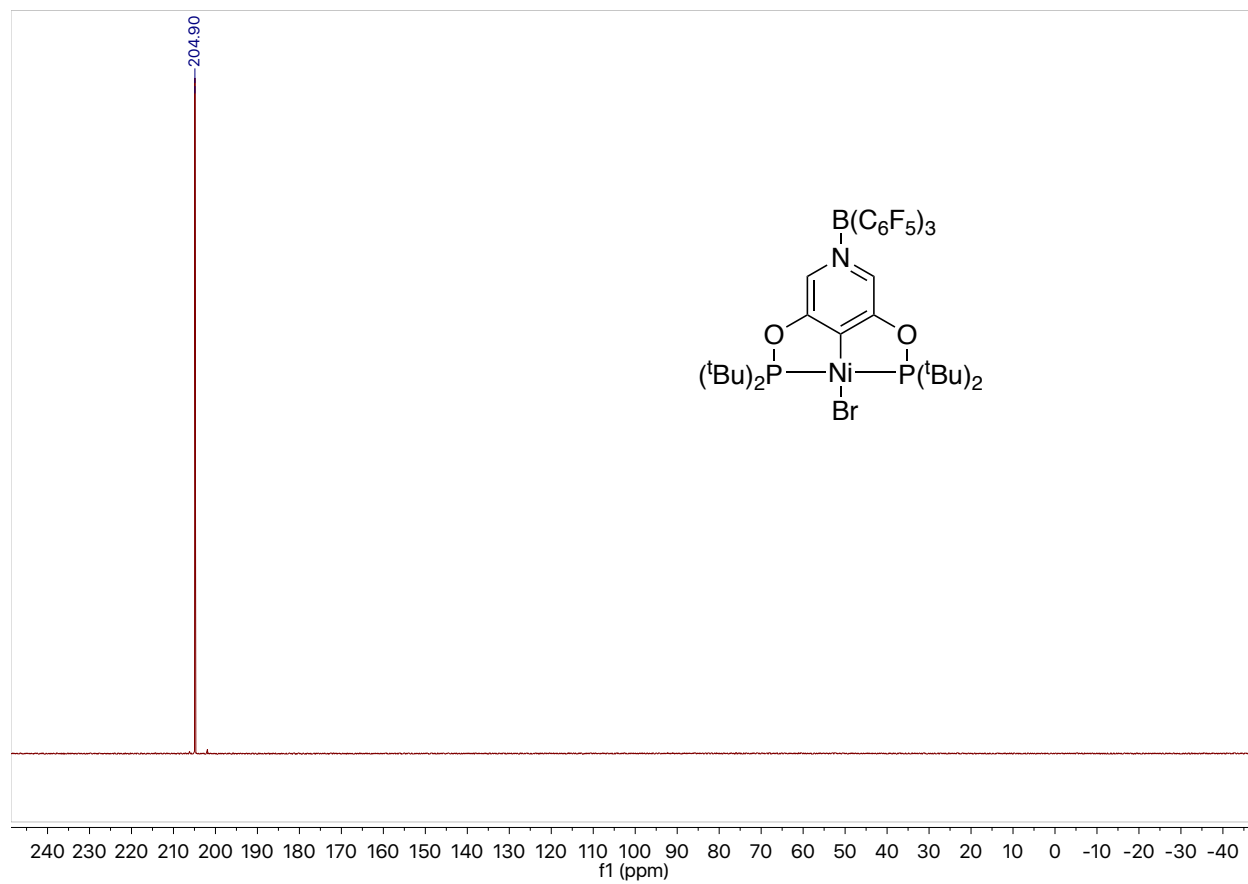


Figure S19. $^{31}\text{P}\{^1\text{H}\}$ NMR (162 MHz, CD_2Cl_2) spectrum of (BCF-PyPOCOP)NiBr, **5**.

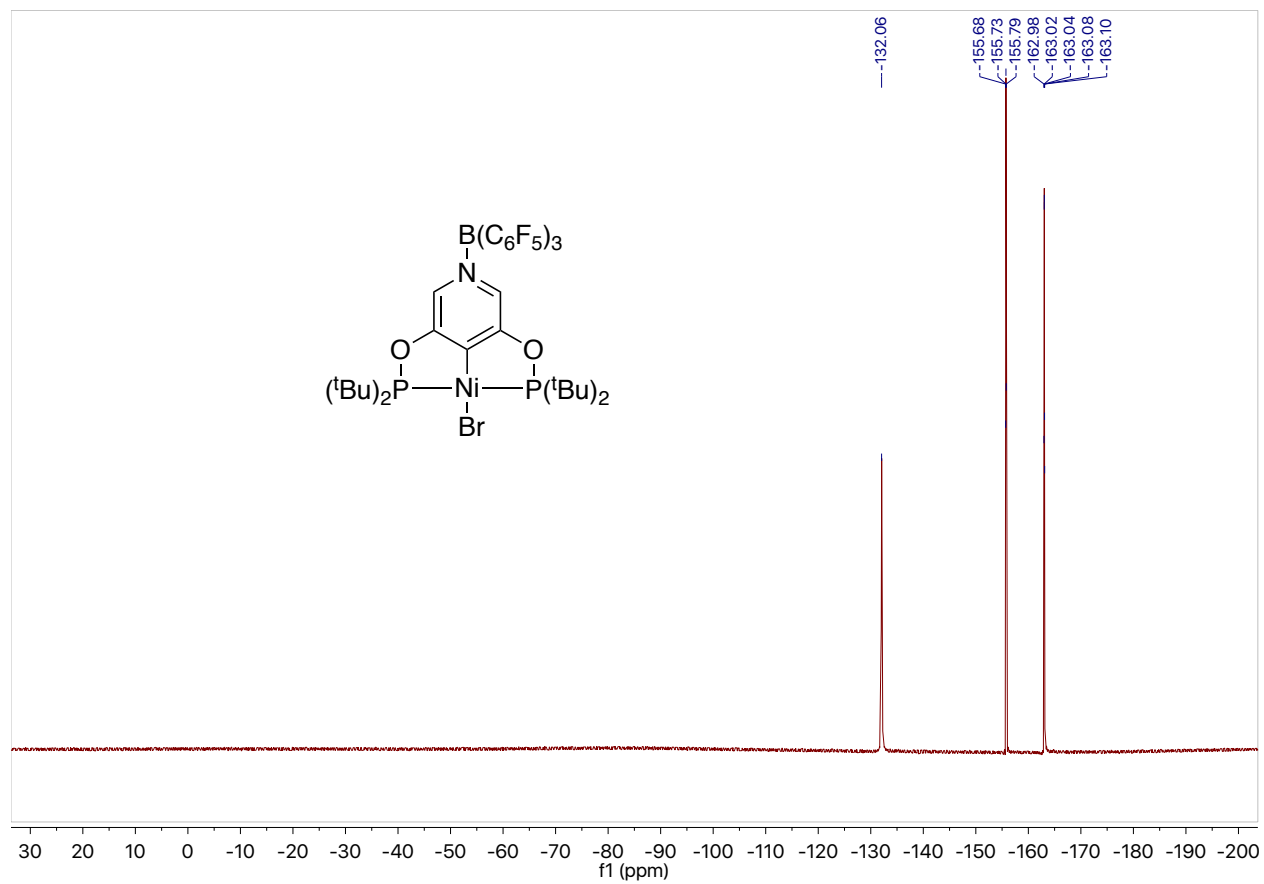


Figure S20. $^{19}\text{F}\{^1\text{H}\}$ NMR (376 MHz, CD_2Cl_2) spectrum of (BCF-PyPOCOP)NiBr, **5**.

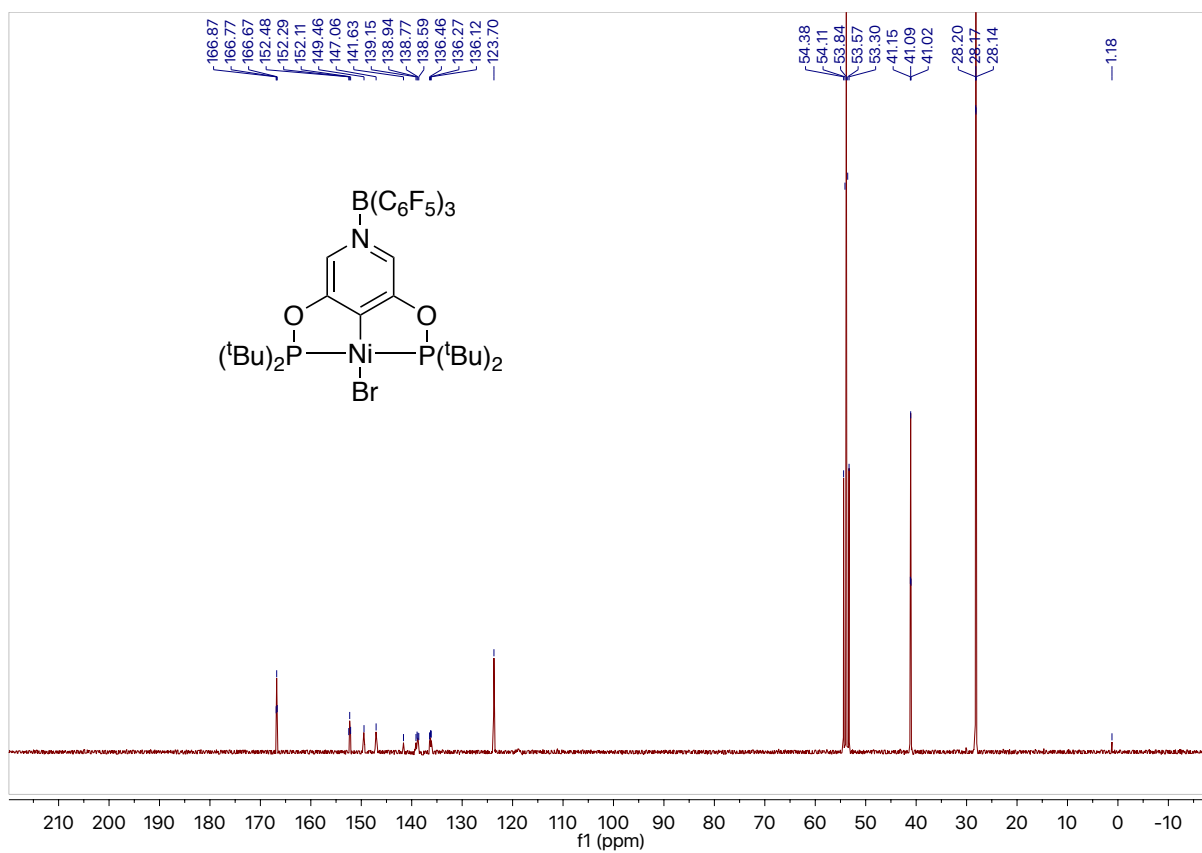


Figure S21. $^{13}\text{C}\{^1\text{H}\}$ NMR (101 MHz, CD_2Cl_2) of (BCF-PyPOCOP)NiBr, **5**. Contamination by a small amount of silicone grease is present.

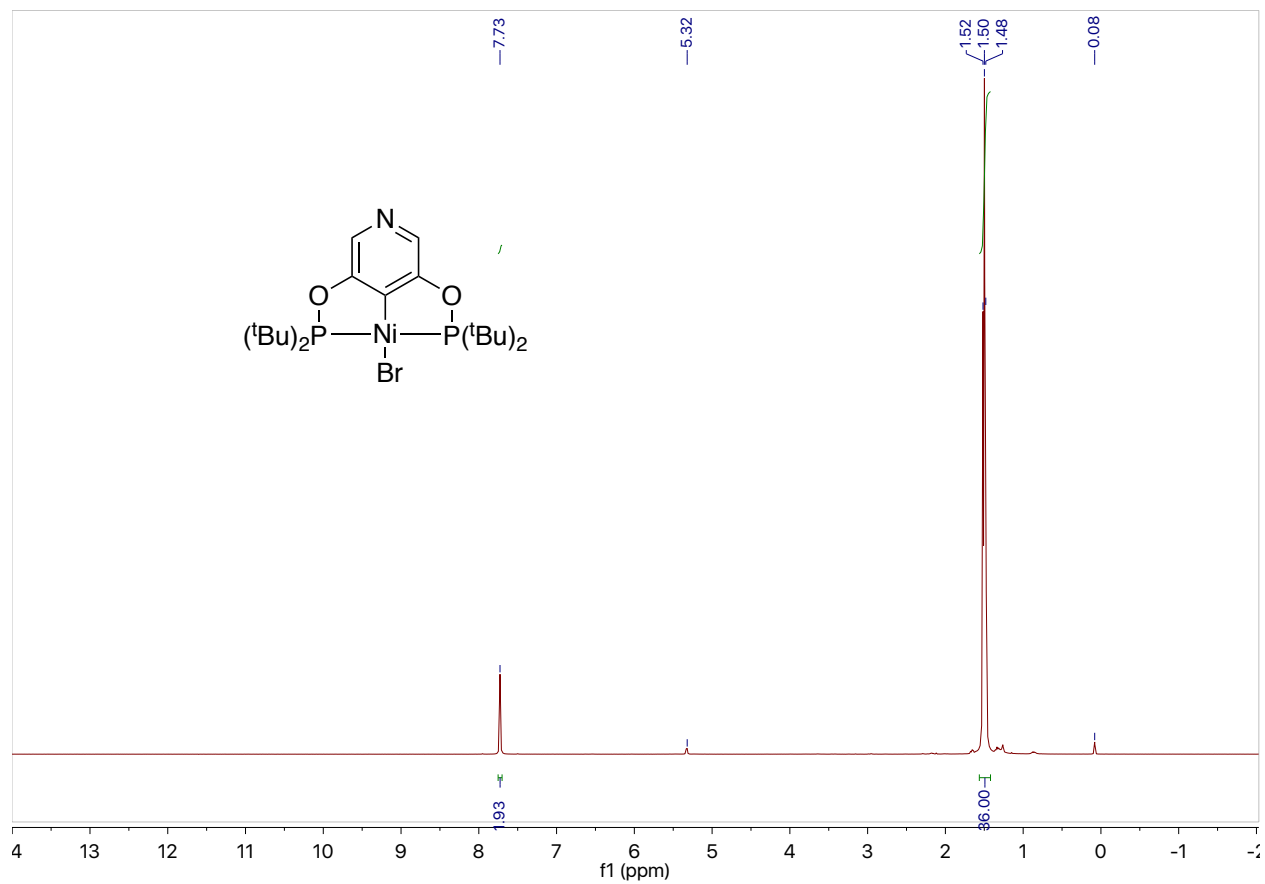


Figure S22. ^1H NMR (400 MHz, CD_2Cl_2) spectrum of (PyPOCOP)NiBr, **6**. Contamination by a small amount of silicone grease is present.

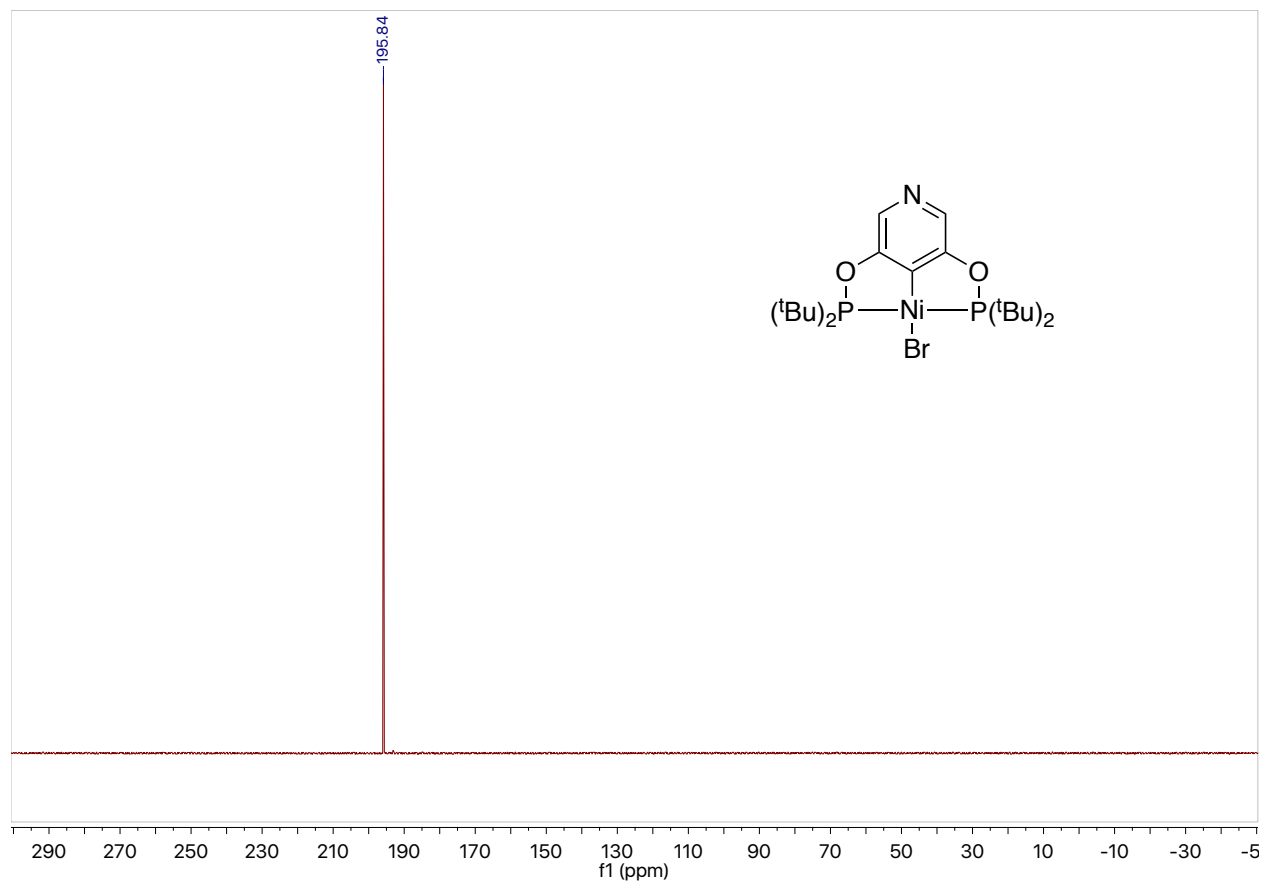


Figure S23. $^{31}\text{P}\{^1\text{H}\}$ NMR (162 MHz, CD_2Cl_2) spectrum of (PyPOCOP)NiBr, **6**.

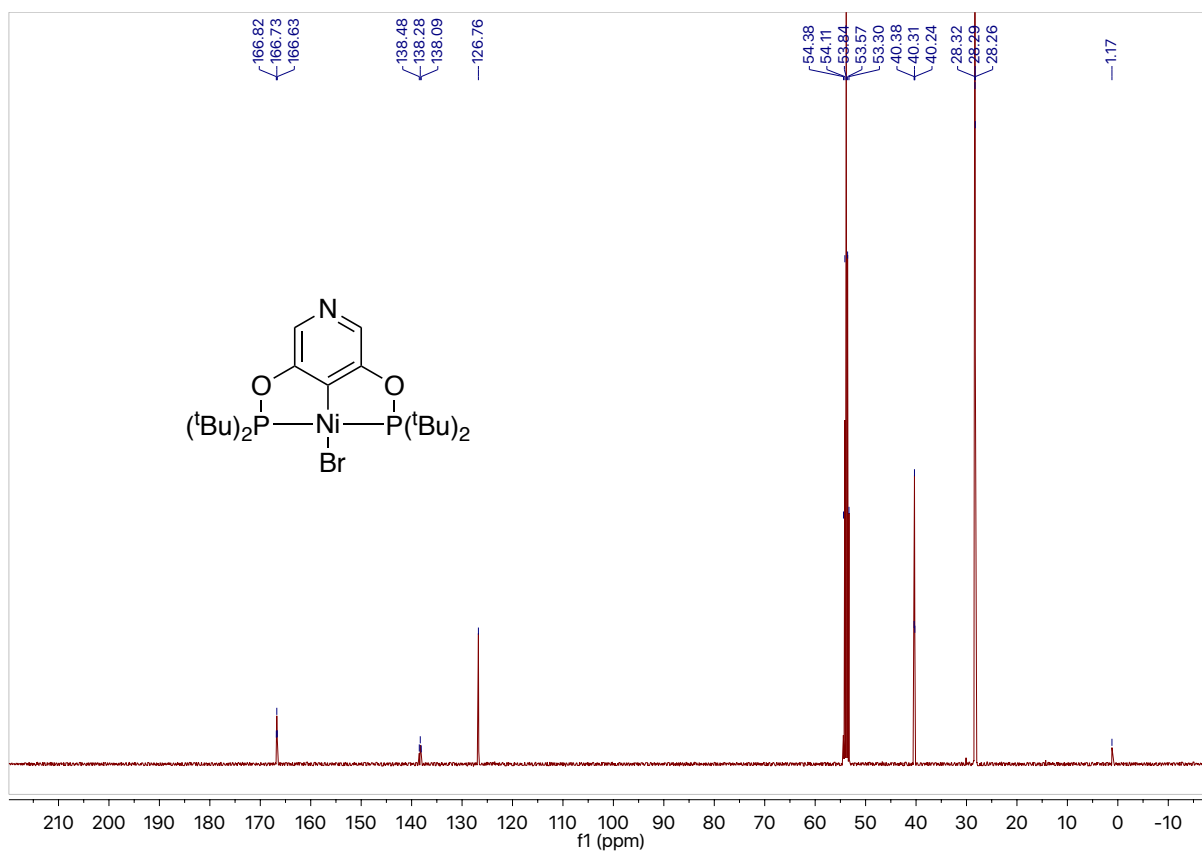


Figure S24. $^{13}\text{C}\{^1\text{H}\}$ NMR (101 MHz, CD_2Cl_2) of (PyPOCOP)NiBr, **6**. Contamination by a small amount of silicone grease is present.

Infrared Spectra

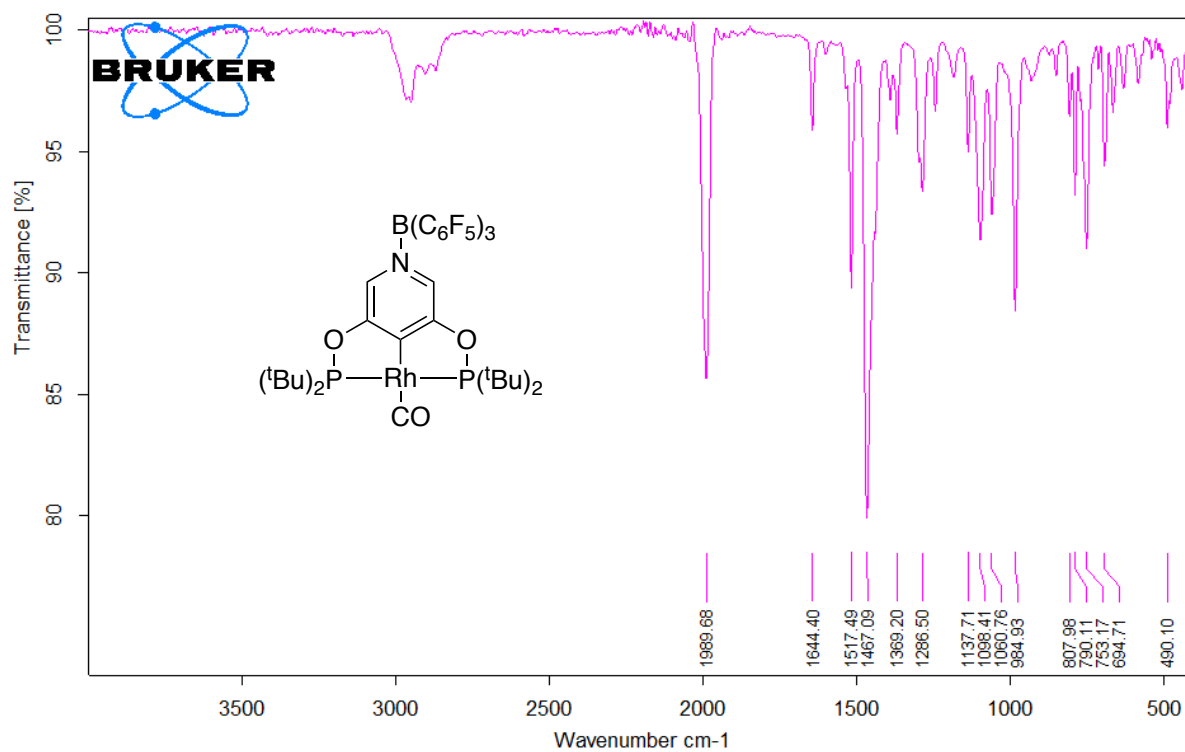


Figure S25. IR spectrum (ATR-IR, thin film THF) of (BCF-PyPOCOP)Rh(CO), **4**.

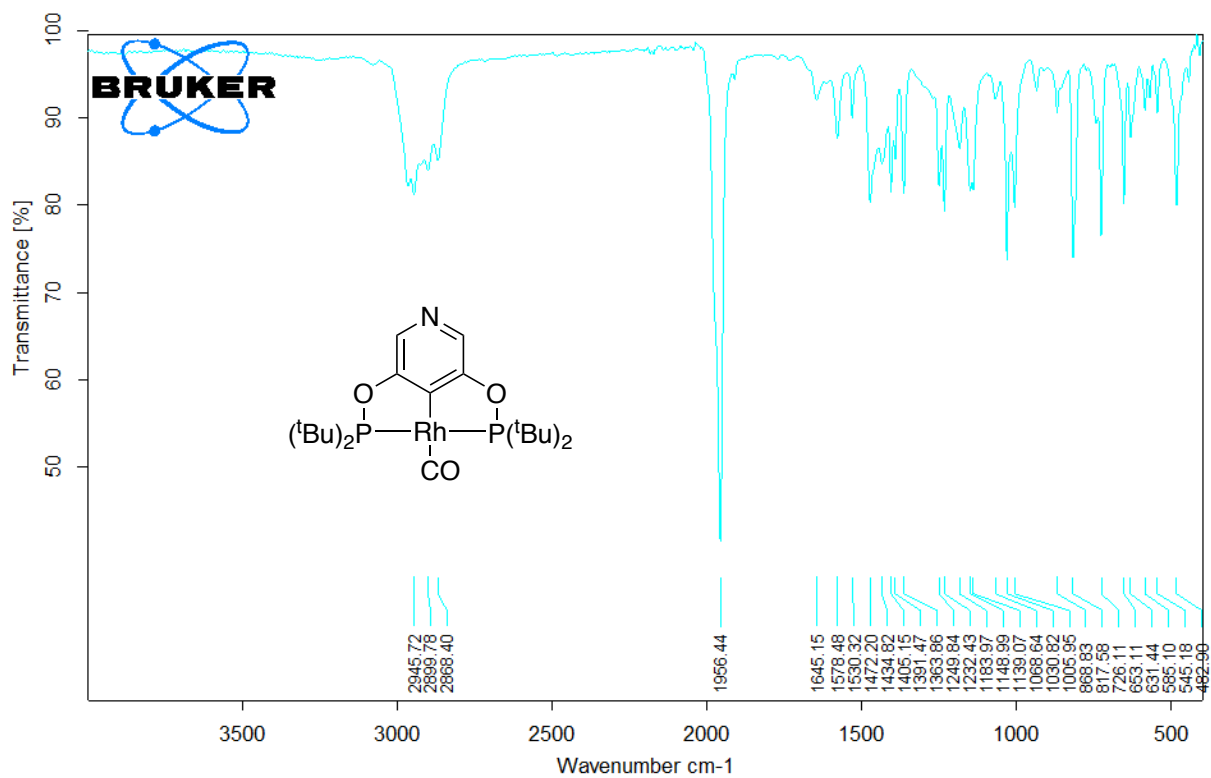


Figure S26. IR spectrum (ATR-IR, thin film THF) of (PyPOCOP)Rh(CO), **3**.

Cyclic Voltammetry

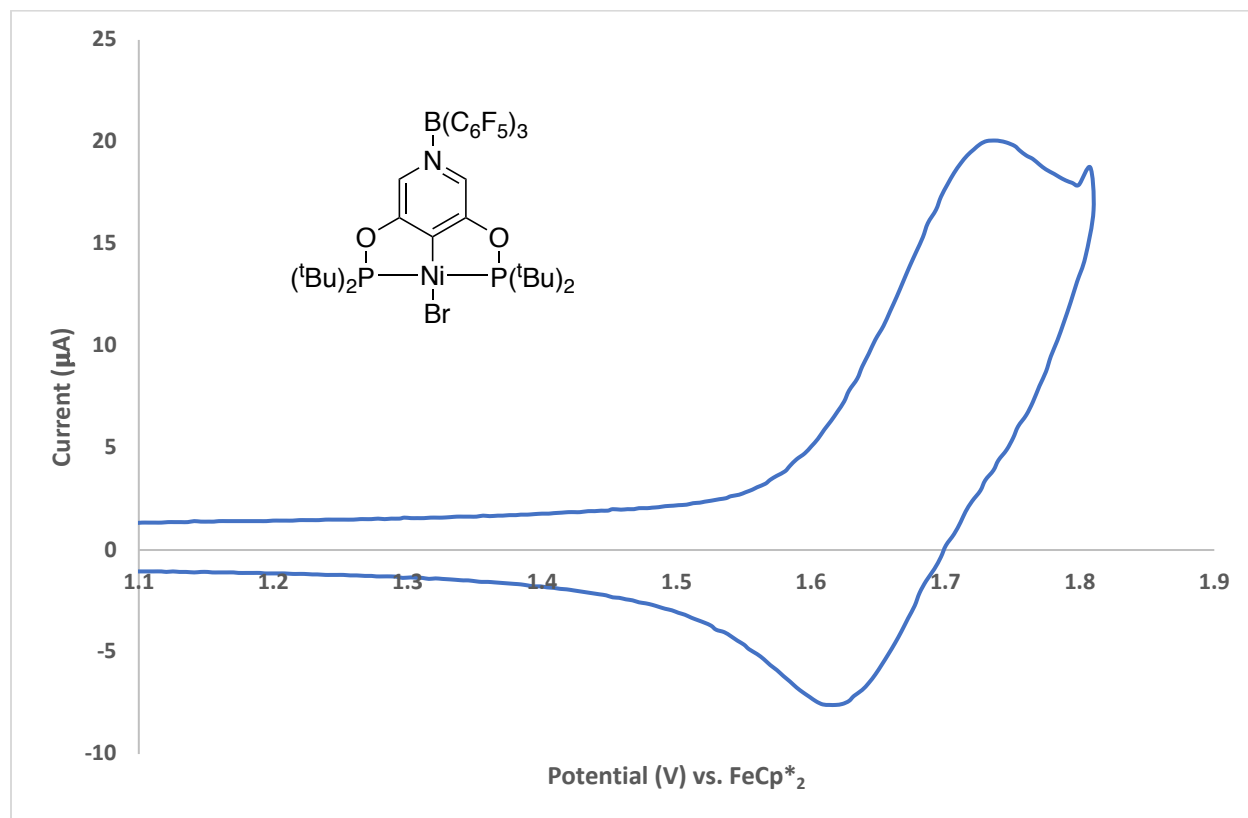


Figure S27. Cyclic voltammogram of complex **5** in CH_2Cl_2 with 0.1 M $(n\text{-Bu}_4\text{N})(\text{PF}_6)$ electrolyte. $E_{1/2} = 1.68$ V and $E_{\text{ox}} = 1.74$ V vs. decamethylferrocene.

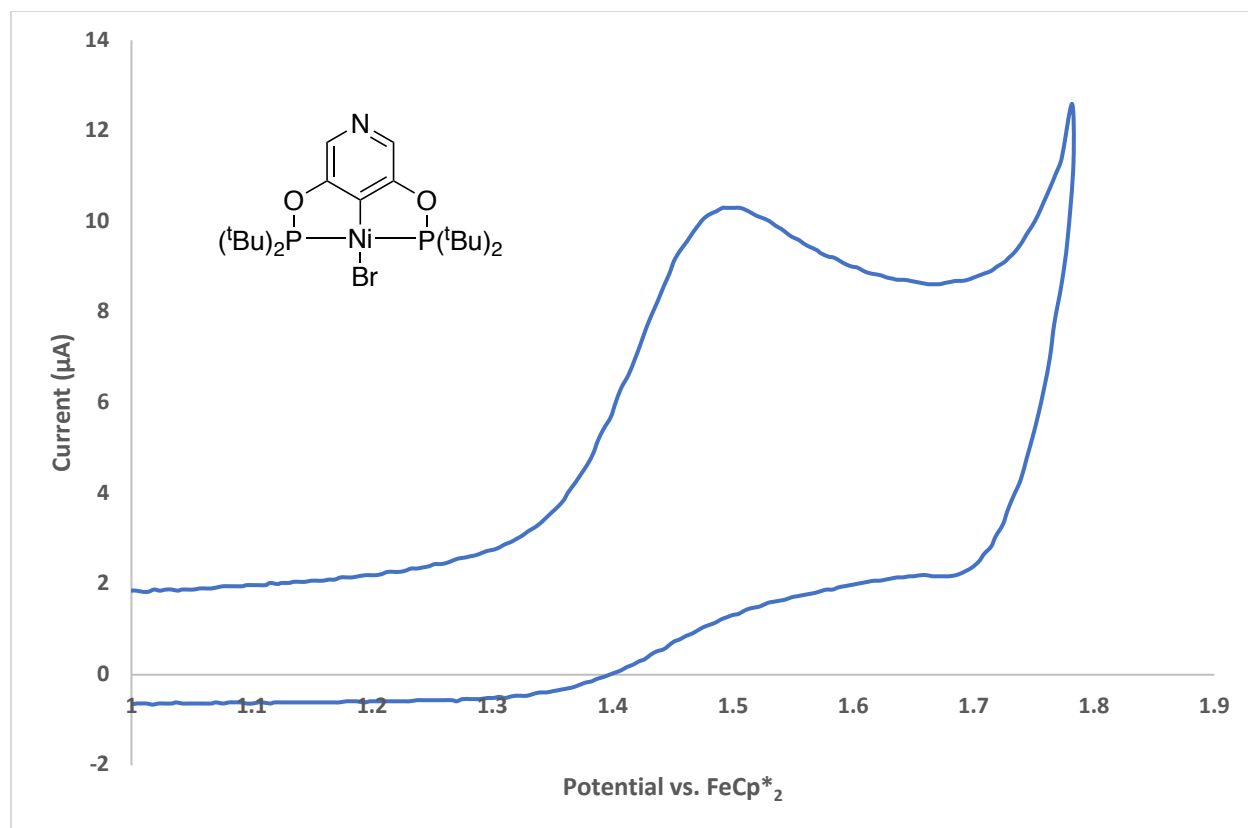


Figure S28. Cyclic voltammogram of complex **6** in CH_2Cl_2 with 0.1 M $(n\text{-Bu}_4\text{N})(\text{PF}_6)$ electrolyte. $E_{\text{ox}} = 1.749\text{V}$ vs. decamethylferrocene.

X-Ray Structure Determination

Compound **5** (BCF-PyPOCOP)NiBr

Low-temperature diffraction data (ϕ - and ω -scans) were collected on a Bruker AXS D8 VENTURE KAPPA diffractometer coupled to a PHOTON 100 CMOS detector with Mo K_{α} radiation ($\lambda = 0.71073 \text{ \AA}$) from an I μ S micro-source for the structure of compound **5**. The structure was solved by direct methods using SHELXS and refined against F^2 on all data by full-matrix least squares with SHELXL-2016 using established refinement techniques.¹⁻³ All non-hydrogen atoms were refined anisotropically. All hydrogen atoms were included into the model at geometrically calculated positions and refined using a riding model. The isotropic displacement parameters of all hydrogen atoms were fixed to 1.2 times the U value of the atoms they are linked to (1.5 times for methyl groups).

Compound **5** crystallizes in the monoclinic space group $P2_1/n$ with one molecule in the asymmetric unit.

Empirical formula	C39 H38 B Br F15 N Ni O2 P2
Formula weight	1049.07
Temperature	100(2) K
Wavelength	0.71073 Å
Crystal system	Monoclinic
Space group	P2 ₁ /n
Unit cell dimensions	a=18.331(2) Å
	a= 90°.
	b= 99.077(5)°.
	g= 90°.
b = 12.0823(15) Å	4346.9(9) Å ³
c = 19.875(3) Å	
Volume	
Z	4
Density (calculated)	1.603 Mg/m ³
Absorption coefficient	1.539 mm ⁻¹
F(000)	2112
Crystal size	0.450 x 0.450 x 0.400 mm ³
Theta range for data collection	2.250 to 36.417°.
Index ranges	-28<=h<=30, -20<=k<=20, -33<=l<=33
Reflections collected	151341
Independent reflections	21162 [R(int) = 0.0554]
Completeness to theta = 25.242°	99.9 %
Absorption correction	Semi-empirical from equivalents
Max. and min. transmission	0.7471 and 0.6178
Refinement method	Full-matrix least-squares on F ²
Data / restraints / parameters	21162 / 0 / 571
Goodness-of-fit on F ²	1.025
Final R indices [I>2sigma(I)]	R1 = 0.0357, wR2 = 0.0693
R indices (all data)	R1 = 0.0611, wR2 = 0.0765
Extinction coefficient	n/a
Largest diff. peak and hole	0.918 and -0.964 e.Å ⁻³

Table S1. Crystal data and structure refinement for complex **5**.

References

1. Sheldrick, G.M. Phase Annealing in SHELX-90: Direct Methods for Larger Structures. *Acta. Cryst.* **1990**, *A46*, 467-473.
2. Sheldrick, G.M. A Short History of SHELX. *Acta. Cryst.* **2008**, *A64*, 112-122.
3. Müller, P. Practical Suggestions for Better Crystal Structures. *Crystallography Reviews* **2009**, *15*, 57-83.