



Theoretical free energies of electron transfer, electrochemical properties, electron transfer kinetic and quantitative structural relationships studies of alkynyldihydrofullerene in [X-UT-Y][R-C₆₀-M⁺] supramolecular complexes

Avat Arman Taherpour ^{a,*}, Masomeh Tayebi-Suraki^b and Nosratollah Mahdizadeh^a

^a Department of Organic Chemistry, Faculty of Chemistry, Razi University, P.O. Box: 67149-67346, Kermanshah, Iran

^b Chemistry Department, Faculty of Science, Islamic Azad University, 38135-567, Arak, Iran

*Corresponding author at: Department of Organic Chemistry, Faculty of Chemistry, Razi University, P.O. Box: 67149-67346, Kermanshah, Iran. Tel.: +98.831.4274559; fax: +98.831.4274559. E-mail address: avatarman.taherpour@gmail.com (A.A. Taherpour).

ARTICLE INFORMATION

Received: 13 April 2012
Received in revised form: 10 July 2012
Accepted: 11 July 2012
Online: 30 September 2012

KEYWORDS

Fullerenes
Marcus theory
Non-IPR Carbon Cage
Rehm-Weller equation
Unsaturated thiocrown ethers
Alkynyldihydrofullerene derivatives

ABSTRACT

The isolated pentagon rule (IPR) states that all pentagonal carbon rings are isolated in the most stable fullerenes. Fullerenes (buckminsterfullerene) are a class of spherical carbon allotrope group with unique properties. Electron transfer between fullerene C₆₀ derivatives such as alkynyldihydrofullerene (1-alkynyl-C₆₀ carbanion) and other molecules are thought to involve the transfer of electrons between molecules surrounding the fullerene cage. One class of electron-transfer molecules has introduced as [X-UT-Y][R-C₆₀-M⁺] (R=*tert*-Bu- & H-C≡C-; M=Li & K). The supramolecular complexes [X-UT-Y] (**1-9**) and [R-C₆₀-M⁺] (R=*tert*-Bu- & H-C≡C-; M=Li & K) are shown to possess a previously unreported host-guest interaction for electron transfer processes. The unsaturated, thiocrown ethers (**1-9**, with *cis*-geometry) (described as [X-UT-Y], where X and Y indicate the numbers of carbon and sulfur atoms, respectively) are a group of crown ethers that display interesting physicochemical properties in light of their conformational restriction compared to a corresponding saturated system, as well as the sizes of their cavities. Topological indices have been successfully used to construct mathematical methods that relate the structural data to the various chemical and physical properties. To establish a good relationship between the structures of **1-9** with derivatives of alkynyldihydrofullerene (1-alkynyl-C₆₀ carbanion) as [R-C₆₀-M⁺] (R=*tert*-Bu- & H-C≡C-; M=Li & K) in DMSO and THF solvents **12-38**, an index (μ_{cs}) is utilized. This index is the ratio of the sum of the number of carbon atoms (n_c) and the number of sulfur atoms (n_s) with the product of these two numbers for **1-9**. In this study, we investigated the relationships between this index and the first to third free energies of electron transfer ($\Delta G_{et(n)}$; n=1-3, which is given by the Rehm-Weller equation) between **1-9** and [R-C₆₀-M⁺] (R=*tert*-Bu- & H-C≡C-; M=Li & K) as [X-UT-Y][R-C₆₀-M⁺] (R=*tert*-Bu- & H-C≡C-; M=Li & K) supramolecular complexes in DMSO and THF solvents. The first to third free energies of electron transfer and the kinetic rate constants of the electron transfers, $\Delta G_{et(n)}$ and k_{et} (n=1-3), respectively, were also calculated for [X-UT-Y][R-C₆₀-M⁺] (R=*tert*-Bu- & H-C≡C-; M=Li & K) in DMSO and THF, in accordance with the Marcus theory.

1. Introduction

In 1985 fullerenes were demonstrated as one of the uniquely stable molecular allotropic forms of carbon [1,2]. Fullerenes are more reactive than planar aromatics, acts as an important driving force for the additional reactions. This property returns to the reduction of the strain energies, which results from pyramidalization in the *sp*²-carbon network [3,4]. One of the simplest compounds in the huge family of fullerenes is C₆₀ [3]. Since 1985, metal-containing endohedral fullerenes (EMFs) have attracted special attention as a new class of technologically relevant material due to their combined fullerene-like and metallic properties. In most EMFs, introduction of metal atoms into carbon cages leads to an increase in the electron affinity relative to the corresponding empty-cages [2,5,6]. Endohedral metallofullerenes (EMFs) have shown very interesting applications in optoelectronic devices since varying the encapsulated metal cluster can change the optical and electronic properties, without changing the

structural features of the outer carbon shell. Enhancement of third-order nonlinear optical susceptibility observed in EMFs further establishes them as potential candidates for nonlinear optical devices [2,3]. In early studies, La@C₆₀ was examined by Kroto *et al.*, although it was never isolated in solid form [1,2]. A variety of endohedral metallofullerenes have been reported, but their investigations have been limited because they were typically formed in extremely low yields [2]. Various theoretical techniques have been employed to study the electronic, structural and vibrational properties of fullerenes and their derivatives [2]. These studies have revealed the structure of the fullerene cages, the electronic states of metal atoms and fullerene cages, and how the electronic properties and chemical reactivity of empty fullerene cages change upon endohedral metal doping [2]. EMFs of carbon cage size in the range between C₆₆ and C₈₄ were shown to be stable structures although the corresponding empty fullerenes are either not stable or have been isolated only in minor quantities in the fullerene soot [2].

Among fullerenes, many of them do not follow the IPR rule. There are several ways in which the IPR can be violated; one way is to generate fused-pentagons in which the pentagons are adjacent to one another [2]. All pentagonal carbon rings are isolated in the most stable fullerenes, in accordance with the isolated pentagon rule (IPR). The IPR has proved valuable in understanding the stability of cage structures of fullerenes and metallofullerenes [3]. The most abundant fullerenes, C_{60} and C_{70} and all pure carbon fullerenes larger than C_{70} follow the IPR [8-11]. Non-IPR fullerenes, which contain adjacent pentagons (APs), have been stabilized experimentally in cases where it is topologically impossible to fully isolate all the pentagons, in accordance with Euler's theorem [8]. Alcami *et al.* have shown that, apart from strain, the most important physical property that governs the relative stabilities of fullerenes is the charge distribution within the cage [8]. This charge distribution is controlled by the number and location of APs and pyrene motifs. Alcami *et al.* have also shown that, when these motifs are uniformly distributed and well-separated from one other, stabilization of non-IPR endohedral and exohedral derivatives, as well as pure carbon fullerene anions and cations, becomes the rule, rather than the exception. This suggests that non-IPR derivatives might even be more common than IPR derivatives [8].

Fullerenes violating the IPR are only obtained in derivatized form, since the [5,5]-bond carbons readily react to release bond strain. However, non-IPR fullerenes still have unsaturated sp^2 carbons at the [5,5] bond junctions, which allow their chemical properties to be probed [8-11]. It is concluded that although the fused-pentagon sites are very reactive toward a carbene, the carbons forming the [5,5] junctions are less reactive than the adjacent ones; this confirms that these carbons interact strongly with the encaged metals and are stabilized by them [11].

For C_{60} and C_{72} , only one IPR structure is consistent with their symmetry. Theoretical studies on C_{60} and C_{72} have shown that the non-IPR-satisfying structures are more stable than the IPR-satisfying structures [8-15]. It is believed that the fused-pentagon pairs in non-IPR EMFs are stabilized by electron transfer from the encapsulated metals to the carbon cage [16]. Actually, both experimental results and theoretical calculations have indicated that the encaged metals are associated with the fused-pentagon pairs in non-IPR EMFs, but more direct experimental proof is still lacking [16,17-24]. Furthermore, unlike the exohedrally derivatized non-IPR fullerenes, non-IPR EMFs still bear unsaturated sp^2 carbons at the [5,5] bond junctions, and thus, it is of special interest to elucidate the properties and chemical reactivities of the fused pentagons. However, little is known about the properties of non-IPR fullerenes and non-IPR EMFs [16]. The non-IPR fullerenes are always attractive not only because of their unique structures, which contain fused pentagons, but also because of their unusual properties resulting from the high curvatures around the adjacent pentagons [16,25,26].

It is a well known fact that fullerene C_{60} is a highly electronegative molecule. So, it readily undergoes nucleophilic reactions [27,28]. After the interesting work by Wudl *et al.* [29], Hirsch and co-workers [30] studied the reaction with various organolithium and Grignard reagents, which gave monoalkyl-dihydrofullerenes with well-defined structures after protonation. Fagan *et al.* [31] reported the synthesis and properties of *tert*-butyl- C_{60} anion by means of nucleophilic addition of *tert*-butyllithium to C_{60} [27]. The fullerene proton in the *tert*-butyl derivative was shown to be highly acidic with the pKa value of 5.7 [31]. Since an ethynyl group has electron-withdrawing inductive effects [32] as compared with alkyl groups, it was expected that C_{60} bearing an ethynyl group, $R-C\equiv C-C_{60}-H$, would give even more stable carbanion than *tert*-Bu- C_{60}^- [27,33]. However, these derivatives were not quite suitable for the study of their anions, since the phenylethynyl derivative was only soluble in common organic solvents and

the trimethylsilylethynyl derivative was rather unstable under basic conditions. In 1996, Komatsu *et al.* prepared a 1-octynyl derivative, Hex- $C\equiv C-C_{60}-H$ (**1**), which was expected to be more soluble and chemically stable. The synthesis, properties, and reactions of 1-octynyl- C_{60} anion (2^-) bearing 60 π -electrons on the C_{60} core, which was readily generated from **1** ($R-C_{60}-M^+$; $R=tert\text{-}Bu\text{-} & H-C\equiv C\text{-}$) were reported by Komatsu *et al.* [27,33].

The unsaturated, *cis*-thiocrown ethers (**1-9**), comprise a group of compounds with interesting physicochemical properties, in light of their conformational restrictions compared to corresponding saturated systems and the sizes of their cavities. The presence of sulfur atoms gives rise to the unique properties of thiocrown ethers. The *cis*-unsaturated thiocrown ethers (described as [X-UT-Y], where X and Y indicate the numbers of carbon and sulfur atoms, respectively), **1-9**, were synthesized and their structures were confirmed [34-43]. 1,4-Dithiin is the smallest member of compounds **1-9** that has been widely studied [44-54]. In 2001, the structures of [X-UT-Y] (X = 6, 9, 12, 15, 18, 21, 24 and 27 and Y = 2-9) **1-9** were reported by Tsuchiya *et al.* [34]. In that report, the 1H and ^{13}C NMR spectra, x-ray crystallographic data, ORTEP drawings, cavity size, and UV spectra of [X-UT-Y] **1-9** were carefully considered [34]. The x-ray crystal structures and ORTEP drawings for some members of **1-9** [X-UT-Y], namely X = 15, 18, 21, 24 and 27 and Y = 5-9, show the presence of cavities and a nearly coplanar arrangement of sulfur atoms [34]. The average radii of the cavities for **4-8** were found to be 1.76, 2.34, 3.48, 4.43 and 5.36 Å, respectively [34]. The previously reported ^{13}C and 1H -NMR spectra in $CDCl_3$ showed that compound **4** has the highest chemical shifts. The electron densities of the C=C bonds increase with the increasing ring size from **4** to **9** and decrease from **4** to **1** with the decreasing ring size [34,35]. In 2006, the oxidation potential ($^{ox}E_i$), cyclic voltammograms (Fc/Fc $^+$), and free energies of electron transfer (ΔG_{et}) of supramolecular complexes of [X-UT-Y][C_{60}] with *cis*-unsaturated thiocrown ethers, **1-9**, were considered by Tsuchiya *et al.* [35]. The EMFs and their complexes with the thiocrown ethers have shown interesting properties for applications and basic research studies.

The wide variety of useful applications of graph theory shows that this branch of discrete mathematics can benefit various fields of sciences. Graph theory has been found to be an effective tool in QSAR and QSPR [55-60]. A graph is a topological concept rather than a geometrical concept, and hence Euclidean metric lengths, angles and three-dimensional spatial configurations have no meaning. Numerous studies have related these fields using topological indices (*TI*) [60]. Numerous applications have shown that one can utilize the topological index as a very useful tool for molecular structural studies, which can be used to describe and predict reactivity and structural properties, in addition to biological and toxicological characters of compounds. One group of *TI* was founded by Randić, who introduced the molecular branching index [61]. In 1975, Randić proposed a topological index that has become one of the most widely used in both QSAR and QSPR studies. However, the most important contribution of this stage is probably the development of a great number of applications of *TIs* in several fields of chemistry. The *TIs* are based on Randić's original idea of molecular branching, but have been extended to account for contributions coming from path clusters, clusters, and chains of different lengths [62-69]. A burst in the research into *TIs* began during the 1990s, and is marked by an increase in the number of studies and applications of *TIs* in chemistry [70,78]. Among the successful *TIs* in these applications, it is worth noting the molecular connectivity indices [70,71] (including the Randić index [61]), the Randić index [61, 65-69,79-83], the Kier indices [73,74], the electro-topological state indices [75], the Balaban index [75], and the Wiener index [72]. Trinajstić and coworkers have reported that 39 topological indices are presently available in the literature [78,84,85].

Table 1. The reduction potentials ${}^{\text{red}}E_n$ ($n=1-3$ in Volt) of $R-C_{60}M^+$; $R=tert\text{-}Bu\text{-}$ & $Hex-C\equiv C\text{-}$, $M=Li$ & K ^a

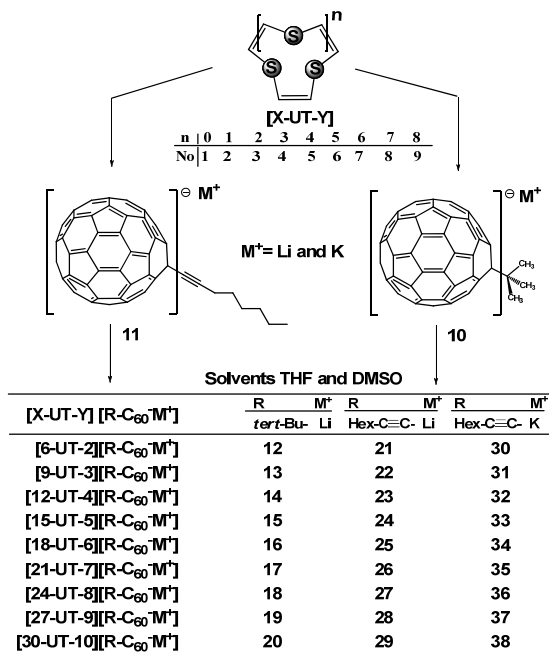
Compounds	Solv.	${}^{\text{red}}E_1$	${}^{\text{red}}E_2$	${}^{\text{red}}E_3$
<i>Tert</i> -Bu- $C_{60}Li^+$	THF	-1.53	-2.04	-2.61
<i>Tert</i> -Bu- $C_{60}Li^+$	DMSO	-1.27	-1.77	-
Hex-C \equiv C- $C_{60}K^+$	THF	-1.55	-2.08	-2.75
Hex-C \equiv C- $C_{60}Li^+$	DMSO	-1.21	-1.72	-2.33

^aSee reference [27].

Estrada has performed important studies of generalized *TIs* with several topological indices in the graph invariant [79]. In 1993 and 1997, the Wiener and Harary indices were applied to studies of fullerenes [85-87]. The use of effective mathematical methods in making strong correlations between chemical properties and the indices has been reported [80-90], which is an important area of development. The ratio of the sum of the number of carbon atoms (n_c) to the number of sulfur atoms (n_s) to the product of these two numbers (μ_{cs}) was a useful numerical and structural value in the studies reported here on the unsaturated thiocrown ethers **1-9** [86-90].

Quantitative structural relationships studies of the μ_{cs} index with respect to the oxidation potentials (${}^{\text{ox}}E_1$) of thiocrown ethers (**1-9**), as well as the free energy of electron transfer (ΔG_{et}) between **1-9** with fullerenes and some of the EMFs were previously reported [16,86-91].

Here, the relationships between this index and oxidation potential (${}^{\text{ox}}E_1$) of **1-9**, as well as the first and second free energies of electron transfer ($\Delta G_{et(n)}$, for $n=1, 2$, which is given by the *Rehm-Weller* equation) between **1-9** and $[R-C_{60}M^+]$ ($R=tert\text{-}Bu\text{-}$ & $H-C\equiv C\text{-}$; $M=Li$ & K) and their supramolecular complexes derivatives as $[X-UT-Y][R-C_{60}M^+]$ ($R=tert\text{-}Bu\text{-}$ & $H-C\equiv C\text{-}$; $M=Li$ & K) in DMSO and THF solvents, are presented and investigated. The first to third free energies of electron transfer and kinetic rate constants of the electron transfers, $\Delta G_{et(n)}$ and k_{et} ($n=1-3$), respectively, were also calculated in this study for $[X-UT-Y][R-C_{60}M^+]$ ($R=tert\text{-}Bu\text{-}$ & $H-C\equiv C\text{-}$; $M=Li$ & K) in DMSO and THF solvents and in accordance with the *Marcus* theory [16,88-93], **Figure 1**.

**Figure 1.** The structures of unsaturated thiocrown ethers **1-9** with $[R-C_{60}M^+]$ ($R=tert\text{-}Bu\text{-}$ & $H-C\equiv C\text{-}$; $M=Li$ & K) **10** and **11**, to produce supramolecular complexes $[X-UT-Y][R-C_{60}M^+]$ ($R=tert\text{-}Bu\text{-}$ & $Hex-C\equiv C\text{-}$; $M=Li$ & K) **12-38**.

2. Graphing and mathematical method

All mathematical and graphing operations were performed using *MATLAB-7.4.0 (R2007a)* and *Microsoft Office Excel-2003* programs. The ratio of the sum of the number of carbon atoms (n_c) and the number of sulfur atoms (n_s) to the product of these two numbers (μ_{cs}) is a useful numerical and structural value for the unsaturated thiocrown ethers **1-9** that were investigated [75,76,86-90].

$$\mu_{cs} = n_s + n_c / (n_s \cdot n_c) \quad (1)$$

If $n_c = 2n_s$, the coefficient of μ_{cs} is given by

$$\mu_{cs} = 3/(2n_s) \quad (2)$$

For modeling, both linear (MLR: Multiple Linear Regressions) and nonlinear (ANN: Artificial Neural Network) models were examined. Other indices were examined and the best results and equations for extending the physicochemical and electrochemical data were chosen.

The *Rehm-Weller* equation estimates the free energy change between an electron donor (*D*) and an acceptor (*A*) as: [92]

$$\Delta G_{et}^{\ominus} = e[E_D^{\ominus} - E_A^{\ominus}] - \Delta E^* + \omega_1 \quad (3)$$

where "*e*" is the unit electrical charge, E_D^{\ominus} and E_A^{\ominus} are the reduction potentials of the electron donor and acceptor, respectively, ΔE^* is the energy of the singlet or triplet excited state, and ω_1 is the work required to bring the donor and acceptor within the electron transfer (ET) distance. The work term in this expression can be considered to be '0' in so far as an electrostatic complex exists before the electron transfer [92].

The *Marcus* theory of electron transfer implies rather weak (<0.05eV) electronic coupling between the initial (locally excited, LE) and final (electron transfer, CT) states and presumes that the transition state is close to the crossing point of the LE and CT terms [93-100]. The value of the electron transfer rate constant k_{et} is controlled by the activation free energy $\Delta G_{et}^{\#}$, which is a function of the reorganization energy ($I/4$) and the electron transfer driving force ΔG_{et} [93-100].

$$\Delta G_{et}^{\#} = (I/4)(1 + \Delta G_{et}/I)^2 \quad (4)$$

$$k_{et} = k_0 \exp(-\Delta G_{et}^{\#}/RT) \quad (5)$$

The reorganization energy of organic molecules ranges from 0.1-0.3 eV. In this study, was used the minimum amount of reorganization energy [93-100].

3. Discussion

The reduction potentials and redox behavior of 1-octynyl- C_{60} anion (2^-) were examined before by the use of cyclic voltammetry in THF and DMSO solvents [27]. Characteristic voltammogram results were observed by Komatsu *et al.* and their data were presented in **Table 1**, together with those reported for *tert*-Bu- C_{60} anion [27].

Table 2. Structural coefficients of unsaturated thiocrown ethers [X-UT-Y] **1-9** and the values of the free energy of electron transfer ($\Delta G_{et(n)}$, n=1-3), in kcal mol⁻¹, between unsaturated thiocrown ethers, **1-9**, with [R-C₆₀-M⁺] (R=*tert*-Bu- & Hex-C≡C-; M=Li & K) **10-11** in supramolecular [[X-UT-Y][R-C₆₀-M⁺]] (R=*tert*-Bu- & Hex-C≡C-; M=Li & K) **12-38** complexes ^a [27].

No	Formula of [X-UT-Y]	μ_{cs}	${}^{ox}E_1$ (Volt)	[X-UT-Y][<i>tert</i> -Bu-C ₆₀ -Li ⁺] 12-20 solved in THF			[X-UT-Y][<i>tert</i> -Bu-C ₆₀ -Li ⁺] 12-20 solved in DMSO		
				$\Delta G_{et(1)}$	$\Delta G_{et(2)}$	$\Delta G_{et(3)}$	$\Delta G_{et(1)}$	$\Delta G_{et(2)}$	$\Delta G_{et(3)}$
1	6-UT-2(1,4-dithiin)	0.7500	1.02	58.05 (58.11)	69.82 (69.87)	82.96 (83.02)	52.06 (52.12)	63.59 (63.65)	-
2	9-UT-3	0.5000	0.97	57.11 (56.96)	68.87 (68.72)	82.01 (81.86)	51.11 (50.96)	62.64 (62.49)	-
3	12-UT-4	0.3750	0.89	55.20 (55.11)	66.96 (66.87)	80.10 (80.02)	49.20 (49.12)	60.73 (60.65)	-
4	15-UT-5	0.3000	0.82	53.60 (53.50)	65.35 (65.26)	78.49 (78.40)	47.59 (47.50)	59.12 (59.03)	-
5	18-UT-6	0.2500	0.79	52.33 (52.81)	64.08 (64.57)	77.29 (77.71)	46.33 (46.81)	57.86 (58.34)	-
6	21-UT-7	0.2143	0.73	51.34 (51.48)	63.09 (63.18)	76.23 (76.33)	45.33 (45.43)	56.86 (56.96)	-
7	24-UT-8	0.1875	0.69	50.54 (50.50)	62.29 (62.26)	75.43 (75.41)	44.53 (44.51)	56.06 (56.04)	-
8	27-UT-9	0.1667	0.66	49.89 (49.81)	61.64 (61.57)	74.78 (74.71)	43.88 (43.81)	55.41 (55.34)	-
9	30-UT-10	0.1500	0.63	49.35 (49.12)	61.10 (60.88)	74.24 (74.02)	43.33 (43.12)	54.87 (54.65)	-

No	Formula of [X-UT-Y]	μ_{cs}	${}^{ox}E_1$ (Volt)	[X-UT-Y][Hex-C≡C-C ₆₀ -Li ⁺] 21-29 solved in DMSO			[X-UT-Y][Hex-C≡C-C ₆₀ -K ⁺] 30-38 solved in THF		
				$\Delta G_{et(1)}$	$\Delta G_{et(2)}$	$\Delta G_{et(3)}$	$\Delta G_{et(1)}$	$\Delta G_{et(2)}$	$\Delta G_{et(3)}$
1	6-UT-2(1,4-dithiin)	0.7500	1.02	50.68 (50.73)	62.43 (62.49)	76.50 (76.56)	58.51 (58.57)	70.74 (70.79)	86.18 (86.24)
2	9-UT-3	0.5000	0.97	49.73 (49.58)	61.49 (61.34)	75.56 (75.41)	57.57 (57.42)	69.79 (69.64)	85.24 (85.09)
3	12-UT-4	0.3750	0.89	47.82 (47.73)	59.58 (59.50)	73.64 (73.56)	55.66 (55.58)	67.88 (67.80)	83.33 (83.25)
4	15-UT-5	0.3000	0.82	46.20 (46.12)	57.97 (57.88)	72.03 (71.95)	54.05 (53.96)	66.27 (66.18)	81.72 (81.63)
5	18-UT-6	0.2500	0.79	44.94 (45.43)	56.70 (57.19)	70.77 (71.25)	52.78 (53.27)	65.01 (65.49)	80.46 (80.94)
6	21-UT-7	0.2143	0.73	43.95 (44.05)	55.71 (55.80)	69.78 (69.87)	51.79 (51.88)	64.01 (64.11)	79.46 (79.56)
7	24-UT-8	0.1875	0.69	43.15 (43.12)	54.91 (54.88)	68.98 (68.95)	50.99 (50.96)	63.21 (63.18)	78.66 (78.64)
8	27-UT-9	0.1667	0.66	42.50 (42.43)	54.26 (54.19)	68.33 (68.26)	50.34 (50.27)	62.56 (62.49)	78.09 (77.94)
9	30-UT-10	0.1500	0.63	41.96 (41.74)	53.71 (53.50)	67.78 (67.57)	49.79 (49.58)	62.02 (61.80)	77.47 (77.25)

^a The data for the compounds and their complexes have not been previously reported. The supramolecular [X-UT-Y][R-C₆₀-M⁺] (R=*tert*-Bu- & Hex-C≡C-; M=Li & K) **12-38** complexes have not been synthesized nor reported previously.

They reported that the anion 2- in THF was oxidized to its radical at -0.39 V vs ferrocene/ferrocenium (Fc/Fc⁺: 0.1 V/s; supporting electrolyte, Bu₄NBF₄; 0.1 M), the corresponding reduction was shifted to -1.20 V, the difference of peak potentials being 0.81 V [27]. They investigated that this large negative shift of the cathodic peak was ascribed to some chemical process associated with the electron transfer, which is most probably a very rapid dimerization of the 1-octynyl-C₆₀ radical [27]. The reduction of the dimer to the original monomeric anion 2- required extra energy for dissociation and also for structural change in C₆₀ framework, which could have caused the cathodic shift as has been observed by Komatsu *et al.* [27]. The second and third reductions were nearly reversible as observed in individual redox waves [27]. In order to examine the structure and properties of anion 2-, semiempirical MO calculations (AM1) were conducted as has been done by Hirsch and co-workers for *tert*-Bu-C₆₀ anion [27,30] (0.1 V/s; supporting electrolyte, Bu₄NBF₄; 0.1 M).

The reduction potentials of **1-9**, shown in Table 2, demonstrate that the μ_{cs} index decreases with increasing the molecular size. In Table 2, related values for the supramolecular complexes of [X-UT-Y] **1-9** with R-C₆₀-M⁺; R=*tert*-Bu- & H-C≡C- (M=Li and K) (**10-12**) were also shown. Table 2 shows the values of oxidation potential (${}^{ox}E_1$), as well as the calculated first and second free energies of electron transfer ($\Delta G_{et(n)}$, n=1-3) between some of the [X-UT-Y] and the complexes **10-12** for supramolecular **13-15** complexes. The

${}^{red}E_n$ (n=1-3) data for **10**, **11** and **12** are presented in Table 1 [16].

The oxidation potentials (${}^{ox}E_1$) of **4-7** were found to be 0.82, 0.79, 0.73 and 0.69 V, respectively [34,35,86-90]. The free energies of electron transfer ($\Delta G_{et(n)}$, n=1-3) between **1-9** and **10-12** for making [X-UT-Y][R-C₆₀-M⁺] (R=*tert*-Bu- & H-C≡C-; M=Li & K) (complexes **13-15**) were calculated using the Rehm-Weller equation [86-90].

In Figure 1 supramolecular [X-UT-Y][R-C₆₀-M⁺] (R=*tert*-Bu- & Hex-C≡C-; M=Li & K) complexes were shown as groups **12-38** in THF and DMSO solvents.

Table 3 shows the equations of the relationships between the μ_{cs} index and the first, second and third free energies of electron transfer ($\Delta G_{et(n)}$, n=1-3) between **1-9** with the molecules [R-C₆₀-M⁺] (R=*tert*-Bu- & H-C≡C-; M=Li & K) **10** and **11** to produce [X-UT-Y][R-C₆₀-M⁺] (R=*tert*-Bu- & Hex-C≡C-; M=Li & K) in THF and DMSO solvents.

Figure 2 shows the plots of the relationship between μ_{cs} and the three free energies ($\Delta G_{et(n)}$, n=1-3 in kcal mol⁻¹) for electron transfer between **1-9** and [*tert*-But-C₆₀-Li⁺] to produce supramolecular [X-UT-Y][*tert*-But-C₆₀-Li⁺] (compounds **12-20**) solved in THF. The related curves for other complexes [X-UT-Y][R-C₆₀-M⁺] (R=*tert*-Bu- & Hex-C≡C-; M=Li & K) **21-38** which they solved in THF and DMSO have similar structures to Figure 2. The calculated values and the related equations of the three free energies of electron transfer ($\Delta G_{et(n)}$, n=1-3) were shown in Table 2 and 3.

Table 3. Second order polynomial equations (4-16) that indicate the relationship between the index μ_{cs} and the first and second free energies of electron transfer ($\Delta G_{et(n)}$, $n=1,2$) between unsaturated thiocrown ethers **1-9** with **10** and **11** in the structures $[X-UT-Y][R-C_{60}M^+]$ ($R=tert\text{-}Bu\text{-}$ & $Hex\text{-}C\equiv C\text{-}$; $M=Li$ & K) **12-38**.

Complexes $[X-UT-Y][R-C_{60}M^+]$ ($R=tert\text{-}Bu\text{-}$ & $Hex\text{-}C\equiv C\text{-}$; $M=Li$ & K)			Equations	R ²	$\Delta G_{et(n)} = a(\mu_{cs})^2 + b(\mu_{cs}) + c$		
Number of complexes					a	b	c
12-20	THF	1	Eq.-6	0.995	-30.691	42.137	43.716
12-20	THF	2	Eq.-7	0.995	-30.707	42.172	55.460
12-20	THF	3	Eq.-8	0.995	-30.683	42.158	68.606
12-20	DMSO	1	Eq.-9	0.995	-30.683	42.158	37.706
12-20	DMSO	2	Eq.-10	0.996	-30.683	42.158	49.236
21-29	DMSO	1	Eq.-11	0.996	-30.754	42.214	44.154
21-29	DMSO	2	Eq.-12	0.996	-30.751	42.208	53.377
21-29	DMSO	3	Eq.-13	0.996	-30.735	42.190	71.831
30-38	THF	1	Eq.-14	0.996	-30.705	42.166	36.323
30-38	THF	2	Eq.-15	0.996	-30.754	42.214	48.074
30-38	THF	3	Eq.-16	0.996	-30.694	42.163	62.150

The values were obtained using the *Rehm-Weller* equation [92], for other complexes of **1-9** with $[R-C_{60}M^+]$ ($R=tert\text{-}Bu\text{-}$ & $Hex\text{-}C\equiv C\text{-}$; $M=Li$ & K) **10-11** as $[X-UT-Y][R-C_{60}M^+]$ ($R=tert\text{-}Bu\text{-}$ & $Hex\text{-}C\equiv C\text{-}$; $M=Li$ & K) **12-38**.

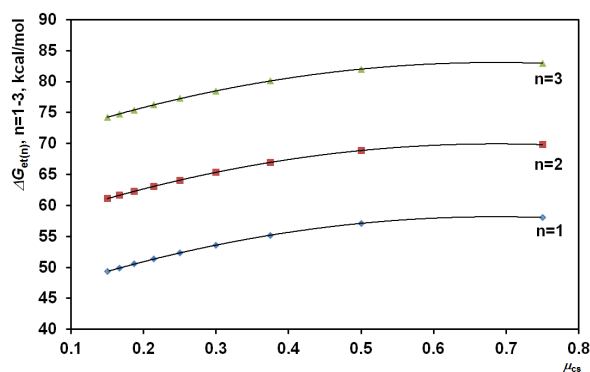


Figure 2. Plots of the relationship between the μ_{cs} index versus the first to third free energies of electron transfer ($\Delta G_{et(n)}$, $n=1-3$ in kcal mol^{-1}) between **1-9** with $[tert\text{-}But\text{-}C_{60}\text{-}Li^+]$ (compounds **12-20**) to produce supramolecular $[X-UT-Y][tert\text{-}But\text{-}C_{60}\text{-}Li^+]$ solved in THF. The related curves for other complexes $[X-UT-Y][R-C_{60}M^+]$ ($R=tert\text{-}Bu\text{-}$ & $Hex\text{-}C\equiv C\text{-}$; $M=Li$ & K) **21-38** which they solved in THF and DMSO have similar structures to **Figure 2**.

Using equations 6-16 (**Table 3**), it is possible to obtain good approximations for the $\Delta G_{et(n)}$ of supramolecular complexes **12-20**, **21-29** and **30-38** groups, in the first to third reduction potential states of $[X-UT-Y][R-C_{60}M^+]$ ($R=tert\text{-}Bu\text{-}$ & $Hex\text{-}C\equiv C\text{-}$; $M=Li$ & K) **12-38** complexes, in THF and DMSO solvents. Equations 6-8 describe **Figure 2**. They show a quadratic relationship between $\Delta G_{et(n)}$ ($n=1-3$) and μ_{cs} for $[X-UT-Y][tert\text{-}But\text{-}C_{60}\text{-}Li^+]$ (compounds **12-20**) solved in THF. In the complex groups, the R^2 values for the graphs (**Figure 2**) are 0.995 and the appropriate equations are given in **Table 3**. In light of the good correlations between μ_{cs} and the free energies of electron transfer, it is possible to use μ_{cs} to calculate the $\Delta G_{et(n)}$ of **12-38**. The values of the ΔG_{et} decrease with increasing group size (**1-9**) and decreasing μ_{cs} indices, as indicated in **Table 2** and **3**. Equations 9 and 10 return to the quadratic second order relationship between $\Delta G_{et(n)}$ ($n=1$ and 2) and μ_{cs} for $[X-UT-Y][tert\text{-}But\text{-}C_{60}\text{-}Li^+]$ (compounds **12-20**) solved in DMSO. In accordance with the data of **Table 1**, there was not any reported result for the third reduction potential (${}^{red}E_3$) of $tert\text{-}Bu\text{-}C_{60}\text{-}Li^+$ in DMSO.

The values of the first to third free energies were shown in **Table 2** for the electron transfer process between unsaturated thiocrown ethers **1-9** and the reduction potential of **10** and **11** as the complexes $[X-UT-Y][Hex\text{-}C\equiv C\text{-}C_{60}\text{-}Li^+]$ (in THF and DMSO solvents) (groups **21-29**). The predicted values of $\Delta G_{et(n)}$ for the complexes **21-29** and other complexes were calculated using the *Rehm-Weller* equation. The equations of the relationship between values of μ_{cs} and the free energies of

electron transfer $\Delta G_{et(n)}$ of **21-29** were shown in **Table 3**. Equations 11 to 13 show quadratic polynomial structures. The R^2 values that indicate correlation between μ_{cs} and the free energies of electron transfer of **21-29** (in ${}^{red}E_n$, $n=1-3$) state of **11** in DMSO solvent) are all equal to 0.996. These good correlations between μ_{cs} and the free energies of electron transfer suggest that it is possible to use μ_{cs} to calculate the $\Delta G_{et(n)}$ for $[X-UT-Y][Hex\text{-}C\equiv C\text{-}C_{60}\text{-}Li^+]$ (in DMSO solvents) (groups **21-29**) by the equations of this model. This case is found to be similar to the other complexes, in which the values of the $\Delta G_{et(n)}$ decrease with the increasing size (**1-9**) and decreasing μ_{cs} indices (see **Table 2**).

Equations 14-16 (**Table 3**) demonstrate the relationships between the free energies of electron transfer between **1-9** and **11** in $[X-UT-Y][Hex\text{-}C\equiv C\text{-}C_{60}\text{-}K^+]$ (in THF solvent) (groups **30-38**) complexes with the μ_{cs} indices for the unsaturated thiocrown ethers, **1-9**. These data were fitted using regression with a second-order polynomial. The amounts of R^2 values for these graphs are 0.996. Using Equations 3 and 14-16, it is possible to calculate the values of $\Delta G_{et(1)}$ to $\Delta G_{et(3)}$ of $[X-UT-Y][Hex\text{-}C\equiv C\text{-}C_{60}\text{-}K^+]$ complexes (groups **30-38**) in THF solvent. As shown in **Table 2**, similar to complexes **12-20** and **21-29** (in THF and DMSO solvents) the values of $\Delta G_{et(n)}$ in the complexes decrease with increasing the size of **1-9** and decreasing μ_{cs} indices. In THF solvent the values of $\Delta G_{et(1)}$ to $\Delta G_{et(3)}$ show greater amounts than DMSO in each group of **12-20**, **21-29** and **30-38** complexes. It may returns back to the greater dipole moment and structural properties of DMSO than THF. It seems that these properties were susceptible the HOMO and LUMO level energies and the HOMO-LUMO energy gaps of the complexes in the groups **12-20**, **21-29** and **30-38**.

The ratio of sum of the number of carbon atoms (n_c) and the number of sulfur atoms (n_s) to the product of these two numbers, given by the index μ_{cs} , shows a good relationship with structural values of the unsaturated thiocrown ethers, **1-9**. These results show the calculated values of the free energies of electron transfer $\Delta G_{et(1)}$ to $\Delta G_{et(3)}$. On the basis of the first to third reduction potentials (${}^{red}E_1$ to ${}^{red}E_3$ states) the free energies of electron transfer were calculated for the supramolecular $[X-UT-Y][R-C_{60}M^+]$ ($R=tert\text{-}Bu\text{-}$ & $Hex\text{-}C\equiv C\text{-}$; $M=Li$ & K) complexes as groups **12-38** in THF and DMSO solvents.

The *Marcus* theory is currently the dominant theory of electron transfer in chemistry. The *Marcus* theory is widely accepted because it accurately predicts electron transfer rates. The most significant prediction is that the rate of electron transfer will increase as the electron transfer reaction becomes more exergonic, but only to a point [94-100].

Electron transfer (ET) is one of the most important chemical processes in nature and plays a central role in many biological, physical and chemical (both organic and inorganic) systems. Solid-state electronics depends on controlling ET in semiconductors. Current molecular electronics depends critically on understanding and controlling the transfer of electrons in and between molecules and nanostructures.

Table 4. Structural coefficients of unsaturated thiocrown ethers [X-UT-Y] **1-9** and the values of the free energy of electron transfer ($\Delta G_{et(n)}^\#$, n=1-3), in kcal mol⁻¹, between unsaturated thiocrown ethers, **1-9**, with [R-C₆₀-M⁺] (R=*tert*-Bu- & Hex-C≡C-; M=Li & K) **10** and **11** in supramolecular [X-UT-Y][R-C₆₀-M⁺] (R=*tert*-Bu- & Hex-C≡C-; M=Li & K) **12-38** ^a [27].

No	Formula of [X-UT-Y]	[X-UT-Y][<i>tert</i> -Bu-C ₆₀ -Li ⁺] 12-20 solved in THF			[X-UT-Y][<i>tert</i> -Bu-C ₆₀ -Li ⁺] 12-20 solved in DMSO ^b		
		$\Delta G_{et(1)}^\#$	$\Delta G_{et(2)}^\#$	$\Delta G_{et(3)}^\#$	$\Delta G_{et(1)}^\#$	$\Delta G_{et(2)}^\#$	$\Delta G_{et(3)}^\#$
1	6-UT-2(1,4-dithiin)	122.88	169.55	230.62	101.99	143.94	-
2	9-UT-3	118.72	164.66	224.86	98.17	139.39	-
3	12-UT-4	112.18	156.94	215.86	92.26	132.33	-
4	15-UT-5	106.63	150.36	208.10	87.21	126.26	-
5	18-UT-6	104.30	147.59	204.83	85.10	123.72	-
6	21-UT-7	99.88	142.08	198.38	80.96	118.72	-
7	24-UT-8	96.68	138.50	194.14	78.26	115.44	-
8	27-UT-9	94.45	135.84	190.94	76.24	112.98	-
9	30-UT-10	92.26	133.20	187.81	74.26	110.58	-

No	Formula of [X-UT-Y]	[X-UT-Y][Hex-C≡C-C ₆₀ -Li ⁺] 21-29 solved in DMSO			[X-UT-Y][Hex-C≡C-C ₆₀ -K ⁺] 30-38 solved in THF		
		$\Delta G_{et(1)}^\#$	$\Delta G_{et(2)}^\#$	$\Delta G_{et(3)}^\#$	$\Delta G_{et(1)}^\#$	$\Delta G_{et(2)}^\#$	$\Delta G_{et(3)}^\#$
1	6-UT-2(1,4-dithiin)	97.42	139.39	199.45	124.57	173.52	247.00
2	9-UT-3	93.72	134.95	194.14	120.38	168.57	241.09
3	12-UT-4	87.92	128.01	185.74	113.82	160.79	231.77
4	15-UT-5	83.02	122.04	178.59	108.20	154.10	223.72
5	18-UT-6	80.96	119.55	175.52	105.85	151.29	220.34
6	21-UT-7	76.92	114.60	169.55	101.19	145.76	213.64
7	24-UT-8	74.26	111.38	165.63	98.17	142.08	209.64
8	27-UT-9	72.31	108.99	162.72	95.93	139.39	205.24
9	30-UT-10	70.40	106.63	159.84	93.72	136.72	202.67

^a The data for the compounds and their complexes were not reported, previously. The compounds [X-UT-Y][R-C₆₀-M⁺] (R=*tert*-Bu- & Hex-C≡C-; M=Li & K) **12-38** complexes were neither synthesized nor reported previously.

^b This derivative has not $k_{et(3)}$ and $\Delta G_{et(3)}^\#$ in DMSO.

Table 5. The values of the first to third free activation energies of electron transfer rate constants ($k_{et(n)}$, n=1-3) of [X-UT-Y][R-C₆₀-M⁺] (R=*tert*-Bu- & Hex-C≡C-; M=Li & K) **12-38** supramolecular complexes between **1-9** and **10** and **11**.

No	Formula of [X-UT-Y]	[X-UT-Y][<i>tert</i> -Bu-C ₆₀ -Li ⁺] 12-20 solved in THF			[X-UT-Y][<i>tert</i> -Bu-C ₆₀ -Li ⁺] 12-20 solved in DMSO ^a		
		$k_{et(1)}$	$k_{et(2)}$	$k_{et(3)}$	$k_{et(1)}$	$k_{et(2)}$	$k_{et(3)}$
1	6-UT-2(1,4-dithiin)	4.32×10^{-78}	2.54×10^{-112}	4.12×10^{-157}	9.07×10^{-63}	1.57×10^{-93}	-
2	9-UT-3	4.87×10^{-75}	9.87×10^{-109}	6.96×10^{-153}	5.76×10^{-60}	3.38×10^{-90}	-
3	12-UT-4	3.07×10^{-70}	4.56×10^{-103}	2.75×10^{-146}	1.25×10^{-55}	5.10×10^{-85}	-
4	15-UT-5	3.58×10^{-66}	3.01×10^{-98}	1.36×10^{-140}	6.33×10^{-52}	1.43×10^{-80}	-
5	18-UT-6	1.84×10^{-64}	3.25×10^{-96}	3.38×10^{-138}	2.23×10^{-50}	1.04×10^{-78}	-
6	21-UT-7	3.24×10^{-61}	3.56×10^{-92}	1.82×10^{-133}	2.42×10^{-47}	4.88×10^{-75}	-
7	24-UT-8	7.19×10^{-59}	1.52×10^{-89}	2.36×10^{-130}	2.32×10^{-45}	1.23×10^{-72}	-
8	27-UT-9	3.06×10^{-57}	1.36×10^{-87}	5.23×10^{-128}	7.11×10^{-44}	7.91×10^{-71}	-
9	30-UT-10	1.25×10^{-55}	1.17×10^{-85}	1.03×10^{-125}	1.98×10^{-42}	4.57×10^{-69}	-

No	Formula of [X-UT-Y]	[X-UT-Y][Hex-C≡C-C ₆₀ -Li ⁺] 21-29 solved in DMSO			[X-UT-Y][Hex-C≡C-C ₆₀ -K ⁺] 30-38 solved in THF		
		$k_{et(1)}$	$k_{et(2)}$	$k_{et(3)}$	$k_{et(1)}$	$k_{et(2)}$	$k_{et(3)}$
1	6-UT-2(1,4-dithiin)	2.04×10^{-59}	3.38×10^{-90}	3.00×10^{-134}	2.51×10^{-79}	3.12×10^{-115}	3.98×10^{-169}
2	9-UT-3	1.06×10^{-56}	6.04×10^{-87}	2.36×10^{-130}	2.97×10^{-76}	1.34×10^{-111}	8.68×10^{-165}
3	12-UT-4	1.91×10^{-52}	7.51×10^{-82}	3.39×10^{-124}	1.91×10^{-71}	6.73×10^{-106}	5.90×10^{-158}
4	15-UT-5	7.51×10^{-49}	1.78×10^{-77}	5.99×10^{-119}	2.53×10^{-67}	5.46×10^{-101}	4.73×10^{-152}
5	18-UT-6	2.42×10^{-47}	1.21×10^{-75}	1.06×10^{-116}	1.34×10^{-65}	6.25×10^{-99}	1.44×10^{-149}
6	21-UT-7	2.21×10^{-44}	5.17×10^{-72}	2.54×10^{-112}	3.48×10^{-62}	7.20×10^{-95}	1.17×10^{-144}
7	24-UT-8	1.98×10^{-42}	1.19×10^{-69}	1.91×10^{-109}	5.76×10^{-60}	3.56×10^{-92}	1.98×10^{-141}
8	27-UT-9	5.30×10^{-41}	6.67×10^{-68}	2.60×10^{-107}	2.52×10^{-58}	3.33×10^{-90}	5.40×10^{-139}
9	30-UT-10	1.35×10^{-39}	3.58×10^{-66}	3.40×10^{-105}	1.06×10^{-56}	3.06×10^{-88}	1.30×10^{-136}

^a This derivative has not $k_{et(3)}$ and $\Delta G_{et(3)}^\#$ in DMSO.

Electron transfer is a very simple chemical reaction, which can be used to gain insight into other kinds of chemistry and biochemistry. Electron transfer is fundamental in chemistry [93-100].

The free energy of electron transfer ΔG_{et} is the difference between the reactants and the products, and $\Delta G_{et}^\#$ is the activation energy. The reorganization energy is the energy required to force the reactants to have the same nuclear configuration as the products without permitting the electron transfer. If the entropy changes are ignored, the free energy becomes energy or potential energy [93-100].

Using Equations 4 and 5, it is possible to calculate the activate free energies of electron transfer and kinetic rate constants of the electron transfers, $\Delta G_{et(n)}^\#$ and $k_{et(n)}$ (n=1-3), respectively, for **12-38** in accordance with Marcus theory. Figure 3 shows the surfaces of the free energies of electron transfer $\Delta G_{et(n)}$ and $\Delta G_{et(n)}^\#$ (n=1-3) between [X-UT-Y] **1-9** and [R-C₆₀-M⁺] (R=*tert*-Bu- & Hex-C≡C-; M=Li & K) **10-11** in supramolecular [X-UT-Y][R-C₆₀-M⁺] (R=*tert*-Bu- & Hex-C≡C-; M=Li & K) **12-38** complexes. The values of the first to third activation free energies of electron transfer, $\Delta G_{et(n)}^\#$ (n=1-3) for

12-20 and **21-29** in THF and DMSO solvents, decrease with $\Delta G_{et(n)}$ and μ_{cs} descriptor, while the kinetic rate constants of the electron transfers $k_{et(n)}$ (n=1-3), increase with decreasing $\Delta G_{et(n)}$ and $\Delta G_{et(n)}^\#$ (n=1-3) for **12-38**. See Tables 4 and 5 and Figure 3. In THF solvent the values of $\Delta G_{et(n)}^\#$ to $\Delta G_{et(3)}^\#$ show greater amounts than DMSO in each group of **12-20**, **21-29** and **30-38** complexes. In notice to the reported results of the electron transfer rate constant $k_{et(n)}$ (n=1-3) in Table 5, the rate of electron transfer in DMSO is greater than THF.

Figure 3 shows the surfaces of the free energies of electron transfer $\Delta G_{et(n)}$ and $\Delta G_{et(n)}^\#$ (n=1 to 3) between **1-9** with **10** and **11** in the structures **12-38** complexes. The values were calculated by equations 1-16, which they are shown in Tables 2-5. With the appropriate equations and in light of the good correlations (see Figures 1-3 and equations 1-16), it is possible to calculate the values of the first to third free energies of electron transfer (ΔG_{et} in kcal/mol), the first to third activation free energies of electron transfer and kinetic rate constants of the electron transfers, $\Delta G_{et(n)}^\#$ and $k_{et(n)}$ (n=1-3), respectively, for **12-38** in THF and DMSO, in close accordance with the results of Marcus theory.

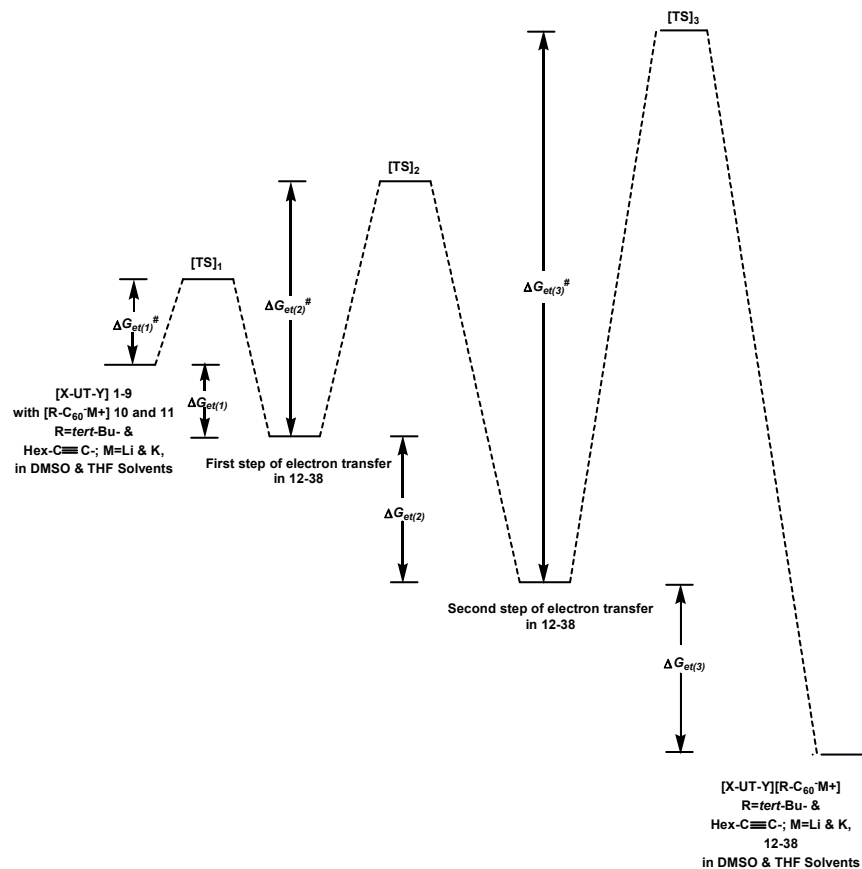


Figure 3. The surfaces of the free energies of electron transfer $\Delta G_{et(n)}$ and $\Delta G_{et(n)}^{\ddagger}$ ($n=1-3$) between **1-9**, **10** and **11** in **12-38**.

The supramolecular complexes of unsaturated thiocrown ethers **1-9** with **10** and **11** as [X-UT-Y][Hex-C≡C-C₆₀-Li⁺] (in THF and DMSO solvents) (groups **21-29**) and their electrochemical data $\Delta G_{et(n)}$, $\Delta G_{et(n)}^{\ddagger}$ and $k_{et(n)}$ ($n=1-3$) have neither been synthesized nor reported before.

4. Conclusion

There have been three reduction potentials (${}^{red}E_1$ to ${}^{red}E_3$) states reported for [R-C₆₀M⁺] (R=tert-Bu- & Hex-C≡C-; M=Li & K) **10-11** in THF and DMSO solvents before. The *cis*-unsaturated thiocrown ethers **1-9** have important physicochemical properties. Here, we reported the first to third free energies of electron transfer $\Delta G_{et(n)}$ ($n=1-3$) on the basis of the electrochemical data of [X-UT-Y] **1-9** and the first to third reduction potential (${}^{red}E_1$ to ${}^{red}E_3$) of the [R-C₆₀M⁺] (R=tert-Bu- & Hex-C≡C-; M=Li & K) **10-11**, first and second activate free energies of electron transfer and kinetic rate constants of the electron transfers, $\Delta G_{et(n)}^{\ddagger}$ and $k_{et(n)}$ ($n=1-3$), respectively, in the supramolecular complexes **12-38**. The predicted values of $\Delta G_{et(n)}$ ($n=1-3$) were calculated for **12-38** groups in THF and DMSO solvents using the *Rehm-Weller* equation. Using the ratio of sum of the number of carbon atoms (n_c) and the number of sulfur atoms (n_s) and the product of these two values, μ_{cs} , equations were derived that yield good structural relationships with the aforementioned physicochemical data. These equations allow one to calculate $\Delta G_{et(n)}$ ($n=1-3$), $\Delta G_{et(n)}^{\ddagger}$ and $k_{et(n)}$ ($n=1-3$) on the basis of the first and second reduction potential (${}^{red}E_1$ and ${}^{red}E_2$) of **10** & **11** for the **12-38** supramolecular complex groups. The values were calculated using the *Rehm-Weller* equation and equations 2-3

concern to the *Marcus* theory. The group of supramolecular complexes [X-UT-Y][R-C₆₀M⁺] (R=tert-Bu- & Hex-C≡C-; M=Li & K) **12-38** were neither synthesized nor previously reported.

Acknowledgments

The corresponding author gratefully acknowledges the colleagues in the Chemistry Department of The University of Queensland (UQ)-Australia, for their useful suggestions.

References

- [1]. Kroto, H. W.; Heath, J. R.; O'Brien, S. C.; Curl, R. F.; Smalley, R. E. *Nature* **1985**, *318*, 162-163.
- [2]. Guhaa, S.; Nakamoto, K. *Coord. Chem. Rev.* **2005**, *249*, 1111-1132.
- [3]. Murphy, T. A.; Pawlik, T. H.; Weidinger, A.; Hohne, M.; Alcalá, R.; Spath, J. M. *Phys. Rev. Lett.* **1996**, *77*, 1075-1078.
- [4]. Billups, W. E. *J. Am. Chem. Soc.* **2005**, *127*(33), 11876.
- [5]. Wang, L. S.; Conceicao, C.; Jin, C.; Smalley, R. E. *Chem. Phys. Lett.* **1991**, *182*, 5-11.
- [6]. Boltalina, O. V.; Ioffe, I. N.; Sorokin, I. D.; Sidorov, L. N. *J. Phys. Chem.* **1997**, *101*(50), 9561-9563.
- [7]. Heflin, J. R.; Marciu, D.; Figura C.; Wang, S.; Burbank, P.; Stevenson, S.; Dorn, H. C. *Appl. Phys. Lett.* **1998**, *72*, 2788.
- [8]. Zettergren, H.; Alcamí, M.; Martín, F. *Chem. Phys. Chem.* **2008**, *9*(6), 861-866.
- [9]. Kato, H.; Taninaka, A.; Sugai, T.; Shinohara, H. *J. Am. Chem. Soc.* **2003**, *125*(26), 7782-7783.
- [10]. Slanina, Z.; Chen, Z.; Schleyer, P. V. R.; Uhlik, F.; Lu, X.; Nagase, S. *J. Phys. Chem. A.* **2006**, *110*(6), 2231-2234.
- [11]. Lu, X.; Nikawa, H.; Nakahodo, T.; Tsuchiya, T.; Ishitsuka, M. O.; Maeda, Y.; Akasaka, T.; Toki, M.; Sawa, H.; Slanina, Z.; Mizorogi, N.; Nagase, S. *J. Am. Chem. Soc.* **2008**, *130*(28), 9129-9136.

- [12]. Wakahara, T.; Nikawa, H.; Kikuchi, T.; Nakahodo, T.; Aminur Rahman, G. M.; Tsuchiya, T.; Maeda, Y.; Akasaka, T.; Yoza, K.; Horn, E.; Yamamoto, K.; Mizorogi, N.; Slanina Z.; Nagase, S. *J. Am. Chem. Soc.* **2006**, *128(44)*, 14228-14229.
- [13]. Kobayashi, K.; Nagase, S.; Yoshida M.; Osawa, E. *J. Am. Chem. Soc.* **1997**, *119(51)*, 12693-12694.
- [14]. Slanina, Z.; Kobayashi, K.; Nagase, S. *Chem. Phys. Lett.* **2003**, *372(5-6)*, 810-814.
- [15]. Slanina, Z.; Ishimura, K.; Kobayashi K.; Nagase, S. *Chem. Phys. Lett.* **2004**, *384*, 114-118.
- [16]. Lu, X.; Nikawa H.; Nakahodo T.; Tsuchiya, T.; Ishitsuka, M. O.; Maeda, Y.; Akasaka, T.; Toki, M.; Sawa, H.; Slanina, Z.; Mizorogi N.; Nagase, S. *J. Am. Chem. Soc.* **2008**, *130(28)*, 9129-9136.
- [17]. Wang, C.; Kai, T.; Tomiyama, T.; Yoshida, T.; Kobayashi, Y.; Nishibori, E.; Takata, M.; Sakata M.; Shinohara, H. *Nature* **2000**, *408*, 426.
- [18]. Stevenson, S.; Fowler, P. W.; Heine, T.; Duchamp, J. C.; Rice, G.; Glass, T.; Harich, K.; Hajdu, E.; Bible, R.; Dorn, H. C. *Nature* **2000**, *408*, 427-428.
- [19]. Shi, Z.; Wu, X.; Wang, C.; Lu X.; Shinohara, H. *Angew. Chem. Int. Ed.* **2006**, *45*, 2107-2111.
- [20]. Beavers, C. M.; Zuo, T.; Duchamp, J. C.; Harich, K.; Dorn, H. C.; Olmstead, M. M.; Balch, A. L. *J. Am. Chem. Soc.* **2006**, *128(35)*, 11352-11353.
- [21]. Yang, S.; Popov, A. A.; Dunsch, L. *Angew. Chem. Int. Ed.* **2007**, *46*, 1256-1259.
- [22]. Slanina, Z.; Chen, Z.; Schleyer, P. R.; Uhlik, F.; Lu, X.; Nagase, S. *J. Phys. Chem. A* **2006**, *110*, 2231-2234.
- [23]. Popov, A. A.; Dunsch, L. *J. Am. Chem. Soc.* **2007**, *129(38)*, 11835-11849.
- [24]. Park, S. S.; Liu, D.; Hagelberg, F. *J. Phys. Chem. A* **2005**, *109(39)*, 8865-8873.
- [25]. Curry, J. D. *J. Exp. Biol.* **1999**, *202*, 3285-3294.
- [26]. Kamat, S.; Su, X.; Ballarini, R.; Heuer, A. H. *Nature* **2000**, *405*, 1036-1040.
- [27]. Murata, Y.; Motoyama, K.; Komatsu, K.; Wan, T. S. M. *Tetrahedron*, **1996**, *52(14)*, 5077-5090.
- [28]. Bhyrappa, P. D. W.; Paul, P.; Stinchcombe, P.; Bolskar, J.; Sun, R. D.; Reed, C. A. *J. Am. Chem. Soc.* **1995**, *117*, 2907-2914.
- [29]. Wudl, F.; Hirsh, A.; Khemani, K. C.; Suzuki, T.; Allemand, P. M.; Koch, A.; Eckert, H. ; Srdanov, G.; Webb, H. M. In *Fullerenes Synthesis, Properties, and Chemistry of Large Carbon Clusters*; Hammond, G. S.; Kuck, V. J.; Eds.; ACS Symposium Series 48, American Chemical Society: Washington, DC, 1992, 161-175.
- [30]. Hirsch, A.; Soi, A.; Karfunkel, H. R. *Angew. Chem., Int. Ed. Engl.* **1992**, *31*, 766-768.
- [31]. Fagan, P. J.; Krusic, P. J.; Evans, D. H.; Lerke, S. A.; Johnston, E. *J. Am. Chem. Soc.* **1992**, *114*, 9697-9699.
- [32]. Charton, M. In *Progress in Physical Organic Chemistry*, Ed.; rw Taft, Wiley: New York, 1981, Vol. 13, p. 119-251.
- [33]. Anderson; H. L.; Faust, R.; Rubin, Y.; Diederich, F. *Angew. Chem.* **1994**, *106*, 1427-1429.
- [34]. Tsuchiya, T.; Shimizu, T.; Kamigata, N. *J. Am. Chem. Soc.* **2001**, *123(47)*, 11534-11538.
- [35]. Tsuchiya, T.; Kurihara, H.; Sato, K.; Wakahara, T.; Akasaka, T.; Shimizu, T.; Kamigata, N.; Mizorogi N.; Nagase, S. *Chem. Commun.* **2006**, *20*, 3585-3587.
- [36]. Anderson, M. R.; Dorn, H. C.; Stevenson, S. A. *Carbon* **2000**, *38*, 1663-1670.
- [37]. Cooper, S. R. *Acc. Chem. Res.* **1988**, *21(4)*, 141-146.
- [38]. Blake, A. J.; Schroder, M.; *Advances in Inorganic Chemistry*, Ed. Sykes, A. G., Academic Press. Inc.: New York, 1990, Vol. 35, p 2 and references therein.
- [39]. Rawle, S. C.; Cooper, S. R. *J. Chem. Soc., Chem. Commun.* **1987**, *4*, 308-309.
- [40]. Parker, D. *Macrocyclic Synthesis: A Practical Approach*, Ed. , Oxford University Press: New York, 1996.
- [41]. Pedersen, C. J. *J. Org. Chem.* **1971**, *36(2)*, 254-257.
- [42]. Murray, S. G.; Hartley, F. R. *Chem. Rev.* **1981**, *81(4)*, 365-414.
- [43]. Nakayama, J.; Kaneko, A.; Sugihara, Y.; Ishii, A. *Tetrahedron.* **1999**, *55(33)*, 10057-10066.
- [44]. Weaver, J. H.; Chai, Y.; Kroll, G. H.; Jin, C.; Ohno, T. R.; Haufler, R. E.; Guo, T.; Alford, J. M.; Conceicao, J.; Chibante, L. P. F.; Jain, A.; Palmer, G.; Smalley, R. E. *Chem. Phys. Lett.* **1992**, *190(5)*, 460-464.
- [45]. Smalley, R. E.; Hamond, G. S.; Kuck, V. J.; Editor, *Fullerenes*, Washington DC: American Chemical Society, pp. 1992, 141-59.
- [46]. Yannoni, C. S.; Hoinkis, M.; De Vries, M. S.; Bethune, D. S.; Salem, J. R.; Crowder, M. S.; Johnson R. D.; Robert, D. *Science* **1992**, *256*, 1191-1192.
- [47]. Ruoff, R. S.; Kadish, K. M.; Boulas, P.; Chen, E. C. M. *J. Phys. Chem.* **1995**, *99(21)*, 8843-8850.
- [48]. Fowler, P. W. Manolopoulos, D. E.; In: *An Atlas of Fullerenes*, Vol. 30, Oxford: Clarendon Press, 1995.
- [49]. Hoffman, K. R.; Delapp, K.; Andrews, H.; Sprinkle, P.; Nickels, M.; Norris, B. J. *Lumin.* **1995**, *66-67(1-6)*, 244-248.
- [50]. Dennis, T. J. S.; Kai, T.; Tomiyama, T.; Shinohara, H. *Chem. Commun.* **1998**, *5*, 619-620.
- [51]. Stevenson, S.; Dorn, H. C.; Burbank, P. M.; Harich, K.; Haynes, J. Kiang, C. H.; Salem, J. R.; de Vries, M. S.; Van Loosdrecht, P. H. M.; Johnson, R. D.; Yannoni, C. S.; Bethune, D. S. *Anal. Chem.* **1994**, *66(17)*, 2675-2679.
- [52]. Hiduka, Y.; Wakahara, T.; Nakajima, K.; Tsuchiya, T.; Nakahodo, T.; Maeda, Y.; Akasaka, T.; Mizorogi N.; Nagase, S. *Chem. Commun.* **2006**, *19*, 2057-2059.
- [53]. Slanina, Z.; Kobayashi, K.; Nagase, S. *J. Chem. Phys.* **2004**, *120*, 3397-3400.
- [54]. Nagase, S.; Kobayashi, K. *Chem. Phys. Lett.* **1994**, *231(2-3)*, 319-324.
- [55]. Hansen, P. J.; Jurs, P. *J. Chem. Edu.* **1988**, *65*, 574-580.
- [56]. Hosoya, H. *Bull. Chem. Soc. Jpn.* **1971**, *44*, 2332-2339.
- [57]. Randic, M. *Acta Chim. Slov.* **1998**, *4*, 239-252.
- [58]. Rucker G.; Rucker, C. *J. Chem. Inf. Comput. Sci.* **1999**, *39*, 788-802.
- [59]. Wiener, H. *J. Am. Chem. Soc.* **1947**, *69*, 17-20.
- [60]. Du, Y. P.; Liang, Y. Z.; Li B. Y.; Xu, C. J. *J. Chem. Inf. Comput. Sci.* **2002**, *42*, 1128-1138.
- [61]. Randic, M. *J. Am. Chem. Soc.* **1975**, *97*, 6609-6615.
- [62]. Sabljic A.; Trinajstic, N. *Acta Pharm. Ugosl.* **1981**, *31*, 189-214.
- [63]. Sybold, P. G.; May, M.; Bagal, U. A. *J. Chem. Edu.* **1987**, *64(7)*, 575-582.
- [64]. Kier L. B.; Hall, L. H.; *Molecular Connectivity in Chemistry and Drug Research*, Academic Press, New York, 1976.
- [65]. Randic, M. *J. Math. Chem.* **1991**, *7*, 155-168.
- [66]. Randic, M.; and Mills, D.; Basak, S. C. *Int. J. Quantum Chem.* **2000**, *80*, 1199-1209.
- [67]. Randic, M.; Plavsic, D.; Lers, N. *J. Chem. Inf. Comput. Sci.* **2001**, *41*, 657-662.
- [68]. Randic, M.; Basak, S. C. *J. Chem. Inf. Comput. Sci.* **2001**, *41*, 614-618.
- [69]. Randic, M.; Pompe, M. *J. Chem. Inf. Comput. Sci.* **2001**, *41*, 575-581.
- [70]. Kier, L. B.; Hall, L. H. *Molecular Connectivity in Structure-Activity*, Research Studies Press, John Wiley and Sons: Letchworth, England, 1986.
- [71]. Gutman, I.; Randic, M. *Chem. Phys. Lett.* **1977**, *47*, 15-19.
- [72]. Wiener, H. *J. Am. Chem. Soc.* **1947**, *69(1)*, 17-20.
- [73]. Kier, L. B. *Quant. Struc. Act. Relat.* **1985**, *4*, 109-116.
- [74]. Kier L. B.; Hall, L. H. *Molecular Structure Description: The Electropotological State*, Academic Press, New York, 1999.
- [75]. Balaban, A. *Chem. Phys. Lett.* **1982**, *89(5)*, 399-404.
- [76]. Hu, Q. N.; Liang, Y. Z. *Internet Electron. J. Mol. Des.* **2004**, *3(6)*, 335-349.
- [77]. Barysz, M.; Plavsic, D.; Trinajstic, N. *Match.* **1986**, *19*, 89-116.
- [78]. Estrada, E. *Chem. Phys. Lett.* **2008**, *463(4-6)*, 422-425.
- [79]. Taherpour A. A.; Shafiee, F. *J. Mol. Struct., Theochem.* **2005**, *726*, 183-188.
- [80]. Hansch, C.; Leo, A.; Hoekman, D. *Exploring QSAR: Hydrophobic, Electronic, Steric Constants*, ACS, Washington, DC, USA, 1995.
- [81]. Bundy, J. G.; Morriss, A. W. J.; Durham, D. G.; Campbell C. D.; Paton, G. I. *Chemosphere.* **2001**, *42*, 885-892.
- [82]. Li, A.; Yalkowsky, S. H. *Ind. Eng. Chem. Res.* **1998**, *37*, 4470-4475.
- [83]. Bolboaca, S. D.; Jantschi, L. *Int. J. Mol. Sci.* **2007**, *8*, 335-345.
- [84]. Slanina, Z.; Chao, M. C.; Lee, S. L.; Gutman, I. *J. Serb. Chem. Soc.* **1997**, *62(3)*, 211-217.
- [85]. Plavsic, D.; Nikolic, S.; Trinajstic, N.; Mihalic, Z. *J. Math. Chem.* **1993**, *12*, 235-250.
- [86]. Taherpour, A. A. *Full., Carb. Nanot., Carb. Nanostruct.* **2007**, *15*, 405-415.
- [87]. Taherpour, A. A. *Full., Carb. Nanot., Carb. Nanostruct.* **2008**, *16(2)*, 142-153.
- [88]. Taherpour, A. A. *Full., Carb. Nanot., Carb. Nanostruct.* **2009**, *17(2)*, 171-186.
- [89]. Taherpour, A. A.; Asadi, T. *Full., Carb. Nanot., Carb. Nanostruct.* **2011**, *19*, 166-181.
- [90]. Taherpour, A. A. *Phosph. Sulf. Silic.* **2010**, *185*, 422-432.
- [91]. Murata, Y.; Motoyama, K.; Komatsu, K.; Wan, T. S. M. *Tetrahedron* **1996**, *52*, 5077-5090
- [92]. Rehm, A.; Weller, A. *Isr. J. Chem.* **1970**, *8*, 259-271.
- [93]. Marcus, R. A. *Modern Phys.* **1993**, *65(3)*, 599-610.
- [94]. Andrea, M. Marcus Theory for Electron Transfer a short introduction MPIP-Journal Club-Mainz-January 29, 2008.
- [95]. Barbara P. F. *J. Phys. Chem.* **1996**, *100*, 13148-13161.
- [96]. Newton, M. D. *Chem. Rev.* **1991**, *91*, 767-792.
- [97]. Jortner, J.; Freed, K. F. *J. Chem. Phys.* **1970**, *52*, 6272-6291.
- [98]. Marcus, R. A. *J. Chem. Phys.* **1965**, *43*, 679-701.
- [99]. Marcus, R. A.; Sutin, N. *Biochim. Biophys. Acta.* **1985**, *811*, 265-322.
- [100]. Kuzmin, M. G. XVIIth IUPAC Symposium on Photochemistry, Dresden, German, July 22-27, 2000, Book of Abstracts, p. 372.