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Mechanistic study of the selective hydrogenation of carboxylic acid derivatives over supported rhenium catalysts

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Abstract

The structure and performance of TiO₂-supported Re (Re/TiO₂) catalysts for selective hydrogenation of carboxylic acid derivatives have been investigated. Re/TiO₂ promotes selective hydrogenation reactions of carboxylic acids and esters that form the corresponding alcohols, and of amides that generate the corresponding amines. These processes are not accompanied by reduction of aromatic moieties. A Re loading amount of 5 wt% and a catalyst pretreatment with H₂ at 500 °C were identified as being optimal to obtain the highest catalytic activity for the hydrogenation processes. The results of studies using various characterization methods, including X-ray diffraction (XRD), X-ray absorption fine structure (XAFS), X-ray photoelectron spectroscopy (XPS), and scanning transmission electron microscopy (STEM), indicate that the Re species responsible for the catalytic hydrogenation processes have sub-nanometer to a few nanometer sizes and average oxidation states higher than 0 and below +4. The presence of either a carboxylic acid and/or its corresponding alcohol is critical for preventing the Re/TiO₂ catalyst from promoting production of dearomatized byproducts. Namely, although Re/TiO₂ is intrinsically capable of hydrogenating aromatic rings, carboxylic acids, alcohols, amides, and amines strongly adsorb on the Re species, which leads to suppression of this process. Moreover, the developed catalytic system was applied to selective hydrogenation of triglycerides that form the corresponding alcohols.

1. Introduction

Hydrogenation reactions of carboxylic acids, esters, and amides are receiving increased attention because of their use in a variety of applications in the pharmaceutical and fine-chemical industry, as well as for upgrading bio-based chemical feedstocks.¹⁻⁵ Although significant progress has been made in recent years, the development of catalysts that promote these processes remains a formidable task. The difficulties are predominantly a consequence of the fact that carboxylic acid derivatives are kinetically unreactive and thermodynamically stable as a result of the low electrophilicity of the carbonyl carbons.⁶ Although stoichiometric amounts of strong reducing agents are typically and effectively used to accomplish these reactions, use of these agents has serious drawbacks in terms of reagent/functional group compatibility, poor atom economy and safety especially when the processes are performed on an industrial scale.⁷ Alternative approaches that utilize catalytic reduction with H₂ are undoubtedly more attractive because they are considerably more atom economic, environmentally benign and cost-effective. Therefore, substantial effort has been made for the development of these hydrogenation reactions, and some are industrially viable.⁷ However, because carboxylic acid derivatives are the most difficult substrates to hydrogenate, harsh reaction conditions are usually required.⁸ For instance, successful operation of a conventional Cu-Cr-based catalytic system for these hydrogenations requires the use of highly severe conditions ($T = 250\text{ }^{\circ}\text{C}$, $p_{\text{H}_2} > 20\text{ MPa}$) and high active metal loadings (ca. 20 wt%).⁹ Moreover, these processes often are unselective and inefficient as a consequence of over-reduction. In particular, most catalytic systems are limited because they preferentially promote hydrogenation of aromatic systems over carboxylic acid groups.¹⁰⁻¹⁴ Because aromatic systems are present in many organic substances, this problem seriously limits the use of those hydrogenation processes in synthetic organic chemistry.

Catalysts comprised of platinum-group-metal complexes have played a leading role in carrying out selective hydrogenations.¹⁵⁻²⁸ However, the chemical and thermal stabilities of these catalysts are typically poor, and separation of catalysts from products is relatively difficult. Another drawback of these homogeneous catalysts is that they often require the use of stoichiometric quantities of additives. As a result, the design and development of highly versatile heterogeneous catalysts for selective hydrogenation of carboxylic acids and their derivatives are highly desirable goals in organic chemistry.

In this context, we have recently reported that TiO₂-supported Re (Re/TiO₂) is a heterogeneous catalyst that, in the absence of additives, promotes selective hydrogenation reactions of various carboxylic acid derivatives that contain aromatic moieties.^{29,30} In the processes, carboxylic acids and esters are converted to corresponding alcohols, while amides are transformed to the corresponding amines. Importantly, this heterogeneous catalytic system does not promote hydrogenation of aromatic groups but preferentially hydrogenate carboxylic acid derivative groups. In addition to these simple hydrogenation processes, this catalyst can be employed in *N*-alkylation reactions of amines, which proceed via the conversion of carboxylic acid (or esters) to the corresponding alcohols followed by reaction of the amine with the generated alcohol. Furthermore, low temperature ($T = 150$ °C) hydrogenation of CO₂ to methanol³¹ and the *N*-methylation reactions of amines, using CO₂ and H₂ as a methylation reagent, can be promoted using Re/TiO₂.

Although the earlier studies led to the development of a highly active and selective catalyst for hydrogenation of carboxylic acid derivatives, they did not reveal the precise nature of the catalytically active Re species. Moreover, little is known about the nature of the active species in a number of valuable Re promoted transformation^{32–42} despite having been the subject of scientific investigations for years.^{43–55} The main reasons for this are that Re often is highly dispersed on supports and is difficult to be reduced, and it can exist in a variety of oxidation states.^{56–65} This deficiency makes it highly desirable to carry out a comprehensive study of supported Re catalysts that focuses on gaining insight into a fundamental understanding of the catalytically active Re species.

In this study, we describe the results of a detailed mechanistic study of selective hydrogenation reactions of carboxylic acids and their derivatives catalyzed by Re/TiO₂. Important information was gained in this investigation using scanning transmission electron microscopy (STEM) and *in situ* methods such as X-ray diffraction (XRD), X-ray absorption fine structure (XAFS), and X-ray photoelectron spectroscopy (XPS). The results show that for effective hydrogenation processes the Re species in the catalyst should have sub-nanometer to a few nanometer sizes and an average oxidation state that is higher than 0 and below +4. Furthermore, the co-presence of either the carboxylic acid derivatives and/or the corresponding products is important for preservation of the high selectivity of these catalytic reactions so that byproducts arising from reduction of aromatic moieties are not formed. The results of this

investigation have not only provided mechanistic insight into Re/TiO₂-promoted hydrogenations, they also led to a better understanding of the behavior of Re on the solid surfaces, which are known to be interesting catalytically active systems.

2. Experimental

Materials and preparation of the catalysts

Chemicals were obtained from commercial suppliers and used as received. TiO₂ (ST-01, 188 m² g⁻¹) was purchased from Ishihara Sangyo Co., Ltd. Al₂O₃ (124 m² g⁻¹) was prepared by calcination of γ -AlOOH (Catapal B) obtained from Sasol for 3 h at $T = 900$ °C. SiO₂ (Q-10, 300 m² g⁻¹) was purchased from Fuji Silysia Chemical Ltd. The carbon support (296 m² g⁻¹) was supplied by Kishida Chemical.

Precursors of Re(5)/TiO₂ (Re = 5 wt%) and Re(5)/Support (Support = Al₂O₃, SiO₂, and Carbon) were prepared by mixing the corresponding support materials with appropriate metal sources, i.e., aqueous solution of NH₄ReO₄. In a typical preparation of Re(5)/TiO₂, NH₄ReO₄ (0.72 g) was added to a round-shaped glass vessel (500 mL) that contain 100 mL of deionized water. After sonication for 1 min, TiO₂ (9.5 g) was added and the resulting solution was stirred at 200 rpm for 15 min at r.t. The reaction mixture was then evaporated at 50 °C, and the residue was dried at 90 °C in air for 12 h. The obtained material was calcined at 500 °C in air for 3 h. Prior to each experiment, the catalyst was prepared by reducing those precursors in a glass tube ($T = 500$ °C, $t = 0.5$ h) under a flow of H₂. Other catalysts were also prepared in the same manner.

Characterization of the catalysts

STEM images were recorded on a JEM-ARM200F microscope (JEOL). The Cs-corrector CESCOR (CEOS GmbH, Germany) was used.

In situ XRD (Rigaku Ultima IV) measurements were conducted using a CuK α radiation source. Samples were heated under a flow of 5% H₂/N₂ (30 mL min⁻¹) at temperatures up to 500 °C with a heating rate of 10 °C/min. XRD patterns were recorded under a flow of 5% H₂/N₂ (30 mL min⁻¹) at temperatures between r.t.-500 °C. *Ex situ* XRD was performed by using a CuK α radiation source with a Rigaku Miniflex. Temperature programmed reduction by H₂ (H₂-TPR) was performed by using BELCAT II (MicrotracBEL).

Ti K-edge and Re L₃-edge XAFS measurements were performed at the BL01B1 line (SPring-8) using a Si(111) double crystal monochromator (proposal No. 2018B1126). For *in situ* analysis, samples in pellet form (ϕ 10 mm) were introduced into a quartz cell equipped with Kapton film windows and gas lines. Samples were heated under a flow of 5% H₂/He (100 mL min⁻¹) at

temperatures up to 700 °C with a heating rate of 10 °C/min. Spectra were recorded under a flow of 5% H₂/He (100 mL min⁻¹) at temperatures between 100-700 °C. A measurements on Re metallic powder as a reference was made using a sample that is sealed in a glove bag filled with N₂ after H₂ reduction at 900 °C in order to ensure that surface ReO_x is not formed. The extended XAFS (EXAFS) analysis was carried out using the REX ver. 2.5 program (RIGAKU). The Fourier transformation of the *k*³-weighted EXAFS from *k* space to *R* space was performed over a *k* range of 3–12 Å⁻¹. Curve-fitting analysis was carried out using the parameters for the Re–O and Re–Re shells provided by FEFF6.

XPS was recorded on an EA 125 X-ray photoelectron spectrometer (Omicron) having a modified UHV chamber employing Mg Kα radiation (1253.6 eV). Binding energies calibration were accomplished by using the O 1s peak energy (532.0 eV) of TiO₂. Samples were subjected to measurements after reduction under a flow of H₂ in a glove box filled with N₂. The glove box was connected directly to the XPS chamber to avoid exposing the reduced samples to air.

Catalytic hydrogenation reactions

The following procedure is representative for reductions of carboxylic acids and amides. After being reduced with H₂ at *T* = 500 °C (cf. Preparation of the Catalysts), the catalyst (2 mol% with respect to the loaded metal), and a mixture of 3-phenylpropionic acid (1.0 mmol) and *n*-dodecane (0.29 mmol) were added to an autoclave with 10 mL volume. The mixture was stirred magnetically (*T* = 140 °C, *p*_{H₂} = 5 MPa). The resultant mixture was diluted with 2-propanol and then analyzed by using GC (Shimadzu GC2014 having a capillary column UA⁺-1 from Frontier Laboratories Ltd.) and GC-MS (Shimadzu GCMS-QP2010 with the same capillary column). GC analyses with *n*-dodecane as the internal standard were employed for yield determination. Yields are given relative to the starting substrate. Conversion was calculated using the following equation (1).

$$\text{Conversion (\%)} = \left(1 - \frac{\text{Amount of detected substrate (mmol)}}{\text{Amount of loaded substrate (mmol)}}\right) \times 100 \quad (1)$$

The following procedure is representative for reduction of triglycerides. After being reduced with H₂ at *T* = 500 °C (cf. Preparation of the Catalysts), the catalyst (2 mol% with respect to the loaded Re for the number of carboxylic acid ester moieties), and a mixture of tristearin (0.33 mmol) and *n*-dodecane (0.29 mmol) were added to the autoclave. The reaction mixture was stirred magnetically (*T* = 220 °C, *p*_{H₂} = 5 MPa). The resultant mixture was analyzed by using the

GC and GC-MS. GC analyses with *n*-dodecane as the internal standard were employed for yield determination. Yields are given relative to the number of carboxylic acid ester moieties. Detailed procedures for product separations and NMR analysis is given in the Supporting Information.

3. Results and discussion

Characterization of the supported Re catalysts

Re(5)/TiO₂, Re(5)/Al₂O₃, Re(5)/SiO₂, and Re(5)/Carbon catalysts were characterized by using *in situ* XRD, STEM, and XAFS analysis. **Figure 1** shows the XRD patterns for the catalysts arising from measurements carried out under a flow of 5% H₂/N₂ (30 mL min⁻¹). The XRD pattern of the Re(5)/TiO₂ catalyst is essentially identical to that of pristine TiO₂ even after H₂ reduction at 500 °C. Peaks from Re species were also not observed for Re(5)/Al₂O₃ and Re(5)/Carbon. On the other hand, peaks from Re species are present in the XRD pattern of Re(5)/SiO₂ after pretreatment with H₂ at 500 °C. The results indicate that Re forms crystalline and sufficiently large species that can be observed by using XRD only on the SiO₂ support. The *ex situ* XRD patterns for these catalysts (**Figure S1** in the Supporting Information) are consistent with those arising from the *in situ* measurements.

Figure 2 shows bright field STEM (BF-STEM) images and high-angle annular dark-field STEM (HAADF-STEM) images of Re(5)/TiO₂, following exposure to air as a consequence of the experimental setup. Inspection of the images shows that Re on TiO₂ is highly dispersed and has particle sizes below 2 nm. Also, some Re species exist as clusters with a few Re atoms and single atoms in the TiO₂ matrix. Viewing the HAADF-STEM images of Re(5)/Al₂O₃, Re(5)/SiO₂, and Re(5)/Carbon, displayed in **Fig. 3**, shows that Re is relatively dispersed on the Al₂O₃ surface and Re clusters with sizes below 1 nm exist. In addition to these highly dispersed species, aggregated Re species with 2-5 nm sizes are present. In Re(5)/SiO₂ and Re(5)/Carbon, Re exists mainly in the form of aggregated particles with sizes larger than 5 nm although there still are some atomically dispersed Re.

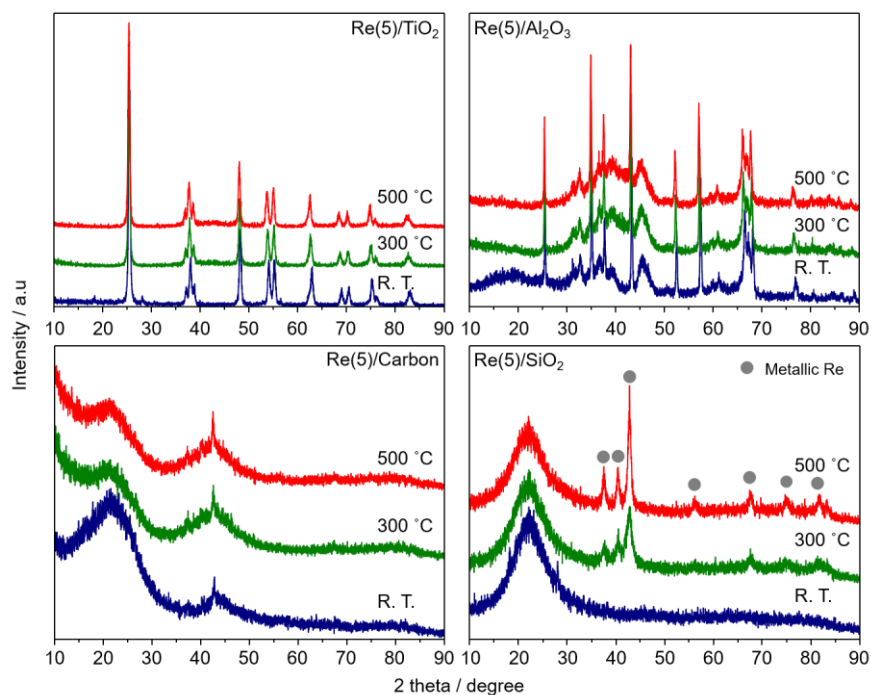


Figure 1. *In situ* XRD patterns of Re(5)/TiO₂, Re(5)/SiO₂, Re(5)/Al₂O₃, and Re(5)/Carbon measured under a flow of 5% H₂/N₂ (30 mL min⁻¹) up to 500 °C with a heating rate of 10 °C/min.

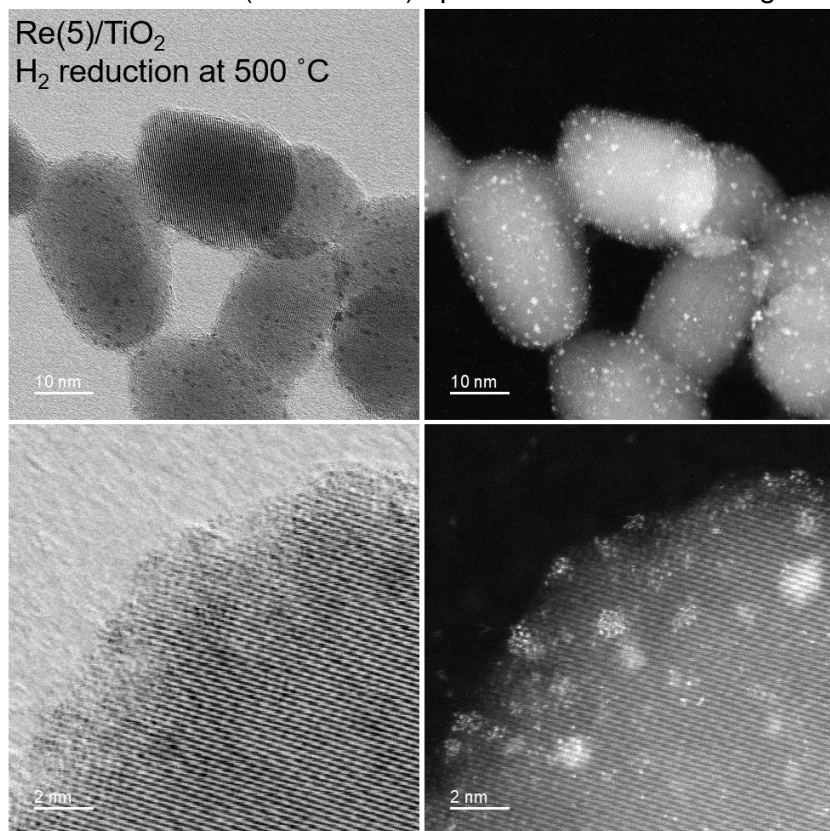


Figure 2. BF(left)- and HAADF(right)-STEM images for Re(5)/TiO₂. The Re(5)/TiO₂ sample was reduced at 500 °C under a flow of H₂. Due to the experimental setup, the sample was exposed to air before measurements.

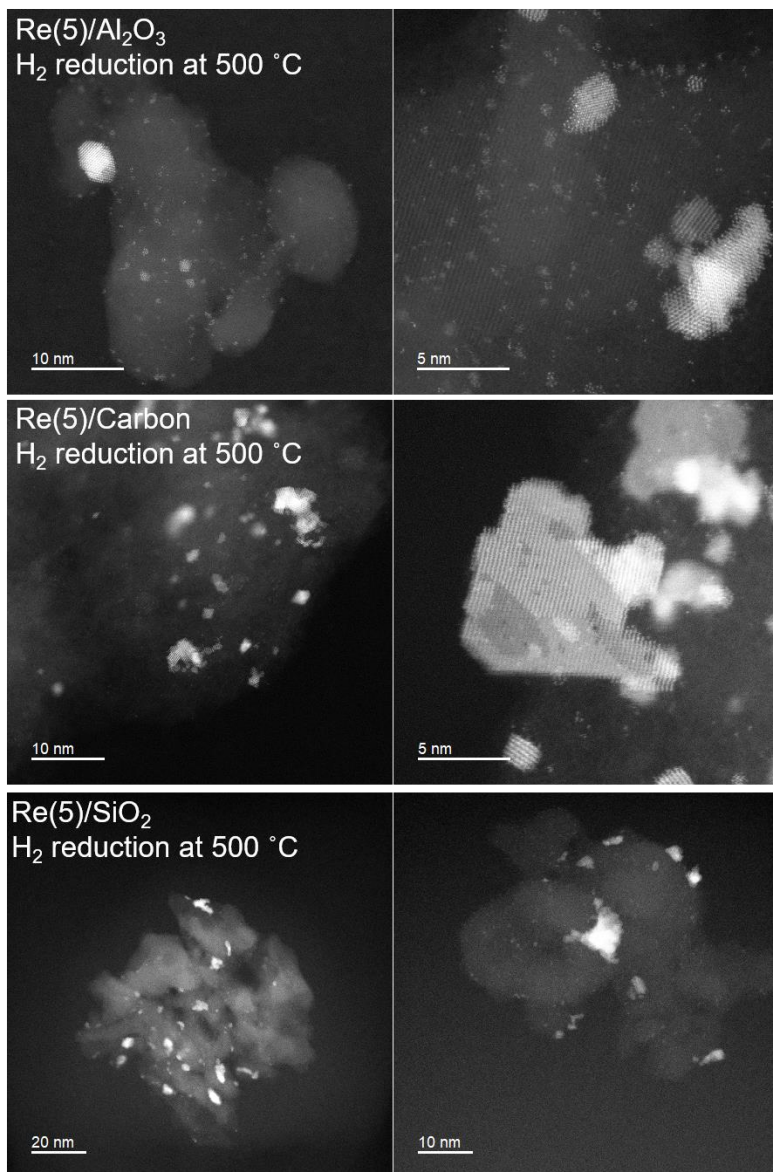


Figure 3. HAADF-STEM images for Re(5)/SiO₂, Re(5)/Al₂O₃, and Re(5)/Carbon. All the catalysts were reduced at 500 °C under a flow of H₂. Due to the experimental setup, the samples were exposed to air before the measurements.

Figure 4 shows Re L₃-edge XANES and EXAFS spectra. The *in situ* XANES spectra of the Re samples (**Fig. 4A-E**) were recorded under a flow of 5% H₂/He (100 mL min⁻¹) at various temperatures between 100-700 °C together with the spectra of reference compounds measured at room temperature. The XANES spectra of all the catalysts treated with H₂ at temperatures up to 200 °C corresponded to that of Re₂O₇, indicating that the Re species in these samples were in the highest oxidation state (VII). The white line intensity for Re(5)/TiO₂ decreased by the H₂ reduction at 300 °C while those of the other Re species decreased at 400 °C. These results

indicate that the valence states of the Re species decreased at 200-400 °C. It is worth mentioning here that Re on TiO₂ starts to be reduced at the lowest temperature (below 300 °C) among the samples measured in this study. A comparison of XANES spectra of Re on different supports is given in **Fig. 4E**. The spectra were recorded using samples treated with H₂ at 500 °C because catalysts treated in this manner have the highest activities in hydrogenation reactions of carboxylic acid derivatives.^{29,30} The results show the edge position lie between those for metallic Re and ReO₂ as references while the white line intensities are all higher than that of ReO₂. These results indicate that all the catalysts contain oxidized Re species. In addition, as shown in **Fig. S2**, a prominent peak was observed at ca. 10549 eV and appeared with the order of Re(5)/TiO₂ < Re(5)/Al₂O₃ < Re(5)/Carbon < Re(5)/SiO₂. It was reported that the peak intensity increases with increase in the Re particle sizes.⁵⁰

Fourier transforms (FT) of *k*³-weighted EXAFS spectra of the supported Re catalysts are displayed in **Fig. 4F**. The spectra were recorded at room temperature on samples not exposed to air after H₂ pretreatment at 500 °C. The distances and coordination numbers (CNs) for the Re–O and Re–Re shells were determined by curve-fitting analysis of the EXAFS data (**Table 1**). The EXAFS spectra of Re(5)/TiO₂ and Re(5)/Al₂O₃ contain features associated with both Re–O and Re–Re. Re(5)/TiO₂ has large Re–O contributions, suggesting that the Re species on TiO₂ is highly dispersed. On the other hand, the spectra of Re(5)/SiO₂ and Re(5)/Carbon are comprised of a single Re–Re contribution at 2.75 Å with a CN of 10.4 and 10.7, respectively. The high CNs of Re–Re observed for Re(5)/SiO₂ and Re(5)/C indicate that most of the Re species form large Re particles. These observations are consistent with results from XRD and STEM measurements.

XANES spectra showed that all the Re species have valence higher than 0 whereas EXAFS spectra showed the Re–Re contribution which suggests formation of metallic Re particles. The Re species in those catalysts would be mixture of Re particles and highly dispersed Re species which should have Re–O bond. Although the fraction of oxidized Re species estimated from the XANES spectra was not so different among the studied catalysts, results of EXAFS showed strong support dependence on the dispersion of Re. True reasons for this inconsistency are not clear. However, origins of the observed XANES spectra would be related to not only the valence states but also the complex structures of Re.⁵⁰

Ti K-edge XANES spectra of Re(5)/TiO₂ were also recorded on samples treated with H₂ at temperatures up to 700 °C (**Figure 5**). Inspection of the spectra shows that the edge position shifts slightly to lower energy with increasing reduction temperature, which indicates that Ti species in the support are also reduced by the H₂ pretreatment. **Figure S3** shows H₂-TPR profiles of Re(5)/TiO₂, Re(5)/SiO₂, Re(5)/Al₂O₃, and Re/Carbon before the H₂ reduction treatment. The peaks corresponding to the H₂ consumption appeared around temperatures of 200-400 °C for all the samples. It was also found that the H₂ consumption peak for Re(5)/TiO₂ appeared at the lowest temperature and that for Re(5)/Al₂O₃ appeared at the highest temperature among the samples tested. These features are consistent with the results of *in situ* Re L₃-edge XANES spectra measured during the H₂ reduction.

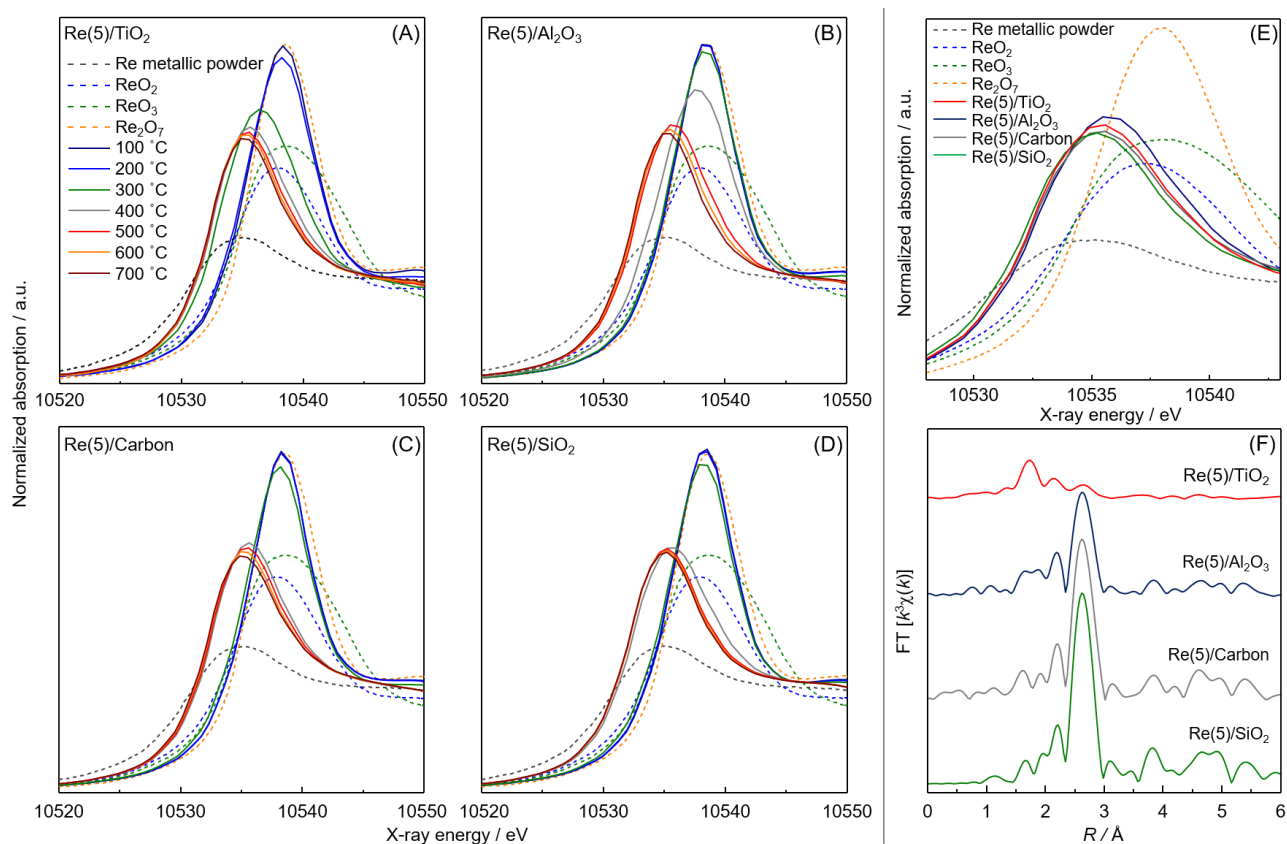


Figure 4. *In situ* Re L₃-edge XANES spectra of (A) Re(5)/TiO₂, (B) Re(5)/Al₂O₃, (C) Re(5)/Carbon, and (D) Re(5)/SiO₂. Samples were heated under a flow of 5% H₂/He (100 mL min⁻¹) up to 700 °C with a heating rate of 10 °C/min. XANES spectra for reference samples were recorded at room temperature. (E) Comparison of XANES spectra of Re on different supports which were measured at 500 °C under the flow of 5% H₂/He (100 mL min⁻¹). (F) Fourier transforms of k³-weighted EXAFS oscillations recorded at room temperature. The samples were not exposed to air after the H₂ reduction at 500 °C by sealing the samples in a glove bag filled with N₂.

Table 1. Curve-fitting analysis of the Re L₃-edge EXAFS of the Re(5)/TiO₂, Re(5)/Al₂O₃, Re(5)/Carbon, and Re(5)/SiO₂ catalysts.

Sample	Shell	CN ^a	R (Å) ^b	σ (Å) ^c	R _f (%) ^d
Re(5)/TiO ₂	Re	3.4	2.62	0.09	4.1
	O	2.7	2.02	0.09	
Re(5)/Al ₂ O ₃	Re	8.6	2.74	0.08	3.3
	O	2.0	1.98	0.09	
Re(5)/Carbon	Re	10.4	2.75	0.07	2.3
Re(5)/SiO ₂	Re	10.7	2.75	0.07	1.7

^a Coordination number. ^b Bond distance. ^c Debye-Waller factor. ^d Residual factor.

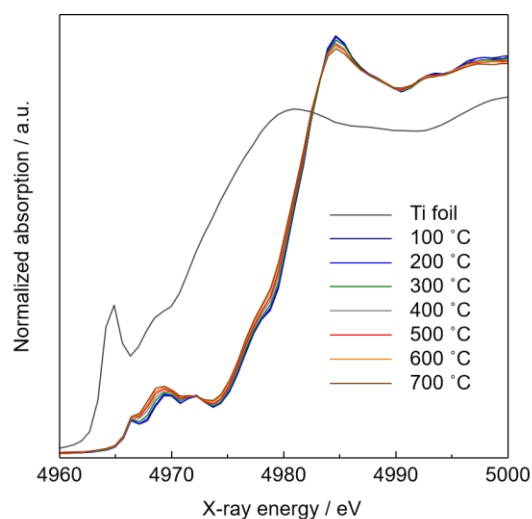


Figure 5. *In situ* Ti K-edge XANES spectra of Re(5)/TiO₂. A sample was heated under a flow of 5% H₂/He (100 mL min⁻¹) up to 700 °C with a heating rate of 10 °C/min. The spectrum of Ti foil was recorded at room temperature.

XPS analysis was performed in order to gain further information about the oxidation states of the Re species on TiO₂. **Figure 6** shows spectra of the Re 4f region of Re(5)/TiO₂ before and after reduction with H₂ at 500 °C, along with a spectrum after H₂ reduction at 700 °C. It should be noted that spectra were recorded on reduced samples directly after H₂ reduction without being exposed to air by using a glove box that is connected to the XPS chamber. The spectra are comprised of signals associated with both 4f_{7/2} and 4f_{5/2} core levels. The spectrum of Re(5)/TiO₂ prior to H₂ reduction contains two Re components in 4f_{7/2} peaks at ca 46.1 eV and ca 44.0 eV, which are attributed to Re⁷⁺ and Re⁶⁺ species, respectively.³⁰ In contrast, a peak at ca 40.7 eV is present in the spectrum of Re(5)/TiO₂ after H₂ reduction at 500 °C, which is assigned to a 4f_{7/2} peak of a Re⁰ species.⁶¹ In addition, contributions from oxidized Re species such as Re⁴⁺, Re²⁺, and Re^{δ+} (δ ~ 1) can be present in the spectra of the samples even after H₂

reduction.⁴⁷ Moreover, analysis of the spectra shows that Re species in Re(5)/TiO₂ exist in further reduced states after H₂ reduction at 700 °C. These results indicate that, in a manner that is consistent with the XAFS observations, H₂ reduction induces formation of low-valent Re species including those in the metallic state. Importantly, however, not all the Re species are reduced to the metallic state even when H₂ reduction is carried out at 700 °C.³⁰

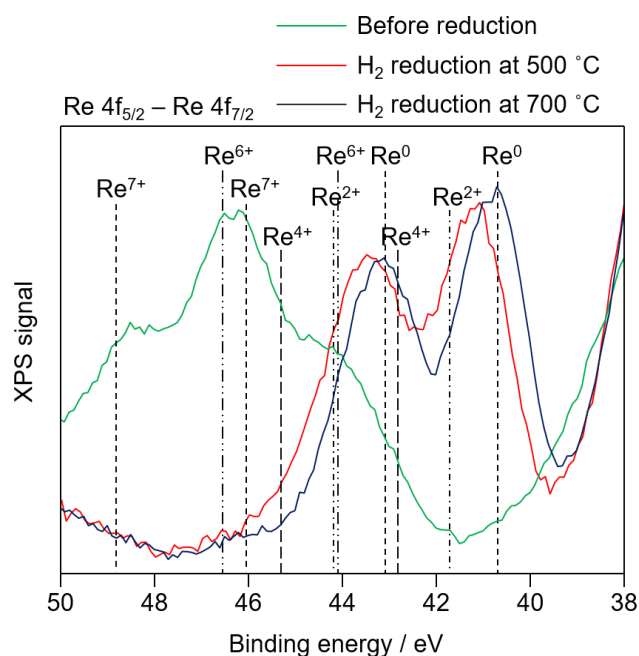
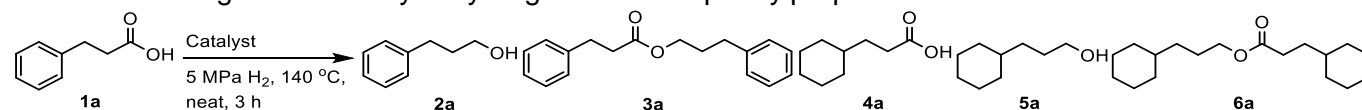


Figure 6. XP spectra of the Re 4f region of Re(5)/TiO₂ before and after the reduction with H₂ at 500 and 700 °C. The spectra before and after the reduction 700 °C are from our previous report.³⁰

Selective hydrogenation of carboxylic acids and their derivatives

Hydrogenations reactions of 3-phenylpropionic acid (**1a**) were explored in order to assess the activities of the various catalysts investigated by using various characterization methods. Following pretreatment of the catalysts with H₂ at 500 °C, reactions were conducted using 1 mmol of **1a** and 2 mol% (based on metal content) of the catalysts in a stainless autoclave (10 mL) under 5 MPa H₂ at 140 °C for 3 h. The results, summarized in **Table 2**, show that the reaction promoted by Re(5)/TiO₂ produces 3-phenylpropanol (**2a**) and the corresponding ester (**3a**), which is formed via an esterification reaction of the starting carboxylic acid with the formed alcohol (Entry 1). Hydrogenation reaction of **1a** catalyzed by Pt(5)/TiO₂ also generates **2a** and **3a** in relatively high yields. In contrast, use of Pt(5)/TiO₂ as a catalyst leads to competitive formation of undesired products such as 3-cyclohexyl-propionic acid (**4a**), 3-cyclohexyl-propan-1-ol (**5a**) and 3-cyclohexyl-propyl 3-cyclohexyl-propionate (**6a**), which are produced by benzene ring reductive dearomatization reactions (Entry 2). These products are also formed in reactions promoted by other precious metal (Ir, Ru, Rh, and Pd) containing catalysts and propylcyclohexane is generated when the reaction time is extended (Entries 3-6).²⁹ In processes in which Re catalysts such as Re(5)/SiO₂, Re(5)/Al₂O₃, and Re(5)/Carbon are employed, only **2a** and **3a** are produced (Entries 7-9), suggesting that Re does not catalyze reactions that form dearomatized products. Although reactions catalyzed by these Re-based catalysts display good selectivity for reduction of the carboxylic acid group rather than the aromatic moiety, the yields are highest when Re(5)/TiO₂ is used. We also found that the desired alcohol **2a** is efficiently generated via a pathway involving initial formation of 3-phenylpropyl 3-phenyl-propionate (**3a**) in yields that reach 97% after 24 h.²⁹ The results clearly demonstrate that the combination of Re as the catalytically active species and TiO₂ as the support leads to a catalytic system that promotes selective and efficient hydrogenations of carboxylic acids to form alcohols, and that this system does not promote generation of undesired side-products such as alkanes and dearomatized compounds. It should also be noted that we have confirmed that the hydrogenation reactions proceed with heterogeneous nature in our previous studies.^{29,30}

Table 2. Heterogeneous catalytic hydrogenation of 3-phenylpropionic acid.^a

Entry	Catalysts	Conv. (%)	Yield (%)				
			2a	3a	4a	5a	6a
1	Re(5)/TiO ₂	36	18	18	0	0	0
2	Pt(5)/TiO ₂	66	13	24	20	3	6
3	Ir(5)/TiO ₂	6	0	6	0	0	0
4	Ru(5)/TiO ₂	1	0	1	0	0	0
5	Rh(5)/TiO ₂	66	0	0	65	0	1
6	Pd(5)/TiO ₂	52	0	0	52	0	0
7	Re(5)/Al ₂ O ₃	13	7	6	0	0	0
8	Re(5)/Carbon	13	5	8	0	0	0
9	Re(5)/SiO ₂	3	1	2	0	0	0

^aReaction conditions: 2 mol% metal, 1 mmol phenylpropionic acid, no solvent, $p_{\text{H}_2} = 5 \text{ MPa}$, $T = 140 \text{ }^\circ\text{C}$, $t = 3 \text{ h}$. Yields were determined by GC using *n*-dodecane as the internal standard.

Control experiments were conducted to identify the origin of the high selectivity of Re(5)/TiO₂ in catalyzing hydrogenation reactions that form alcohols rather than those lead to reduction of benzene rings. In our previous studies, we employed density functional theory (DFT) calculations to show that the selectivity of hydrogenations of carboxylic acid derivatives strongly depends on the relative adsorption affinities of substances containing benzene rings, and COOH or COOCH₃ groups.^{29,30} In the present effort, we performed individual catalytic hydrogenation reactions of (i) propylbenzene, (ii) propylbenzene + lauric acid (dodecanoic acid), and (iii) propylbenzene + lauryl alcohol (dodecanol) using Re(5)/TiO₂ as the catalyst. The results displayed in **Fig. 7** show that Re(5)/TiO₂ promotes hydrogenation (dearomatization) of propylbenzene to form propylcyclohexane. However, when a mixture of propylbenzene and lauric acid is used, dearomatization reaction of the former substance does not take place while lauric acid is preferentially reduced to generate the corresponding alcohol and lauryl laurylate. Moreover, no reaction occurs using a mixture of propylbenzene and lauryl alcohol. These results show that Re(5)/TiO₂ is capable of promoting hydrogenation of benzene rings but this ability is suppressed in the presence of carboxylic acids and alcohols. This finding suggests that, owing to its oxophilic nature,³⁷ Re species in the catalytic system strongly adsorb carboxylic acids and alcohols, which suppresses its participation in benzene ring reductive dearomatization reactions.

In order to gain further information about the substrates governed selectivity phenomenon, control hydrogenation reactions of amides were explored (**Figure 8**). Individual catalytic hydrogenation reactions of (i) propylbenzene, (ii) propylbenzene + *N,N*-dimethyl-2-phenylacetamide, and (iii) propylbenzene + *N,N*-Dimethylphenethylamine were performed using Re(5)/TiO₂ under more vigorous conditions ($p_{H_2} = 5$ MPa and $T = 200$ °C). The results show that catalytic hydrogenation of propylbenzene occurs under these conditions to produce propylcyclohexane. On the other hand, hydrogenations of a mixture of propylbenzene and *N,N*-dimethyl-2-phenylacetamide, and a mixture of propylbenzene + *N,N*-Dimethylphenethylamine do not lead to formation of products arising by reductive dearomatization. Rather, selective hydrogenation of the amide moiety occurs to generate the corresponding amine. The results, which are in accord with those coming from studies of hydrogenation reactions of propylbenzene, lauric acid, and lauryl alcohol described above, show that the presence of amide or amine groups suppresses undesired dearomatization reactions.

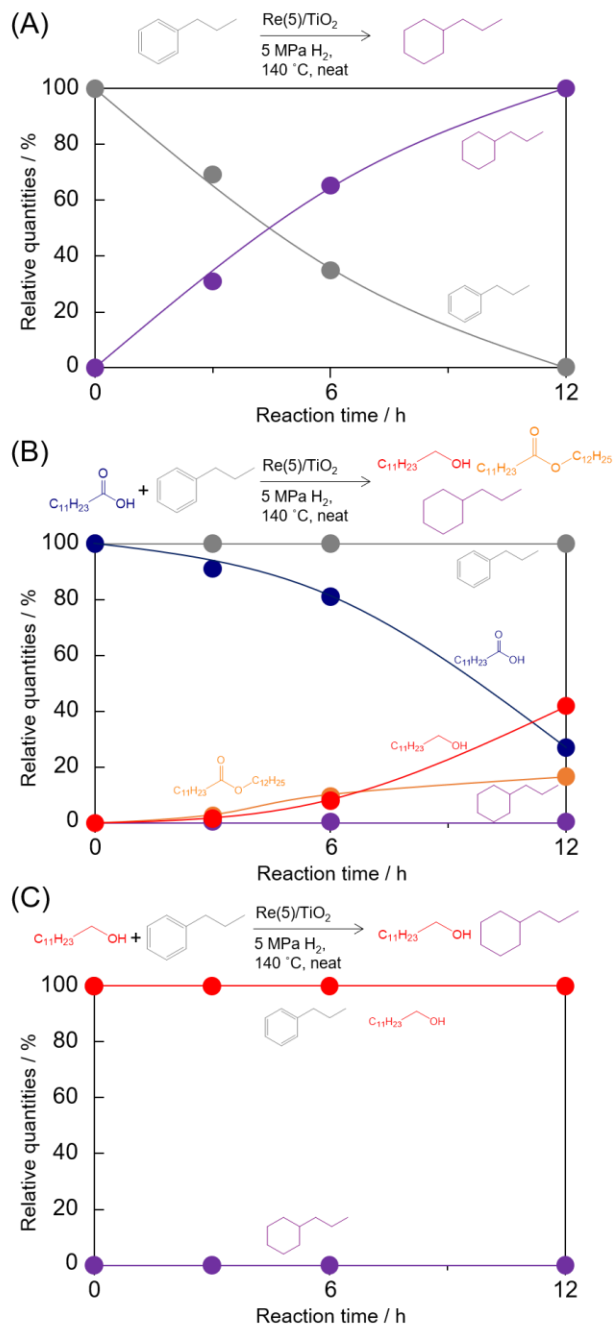


Figure 7. Catalytic hydrogenation of (A) propylbenzene, (B) propylbenzene + lauric acid (dodecanoic acid), and (C) propylbenzene + lauryl alcohol (dodecanol) over Re(5)/TiO₂. Reaction conditions: 2 mol% Re, 1 mmol of each substrate, no solvent, $p_{H_2} = 5$ MPa, $T = 140$ °C. Yields were determined by GC using *n*-octane as the internal standard.

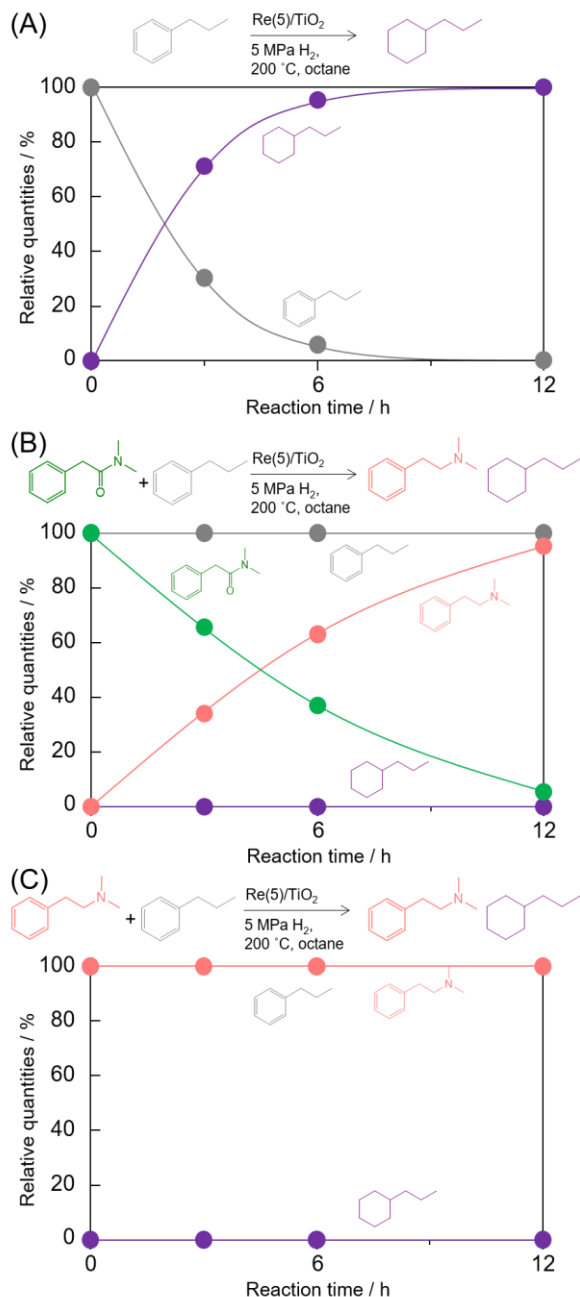


Figure 8. Catalytic hydrogenation of (A) propylbenzene, (B) propylbenzene + *N,N*-dimethyl-2-phenylacetamide, and (C) propylbenzene + *N,N*-Dimethylphenethylamine over Re(5)/TiO₂. Reaction conditions: 2 mol% Re, 1 mmol of each substrate, 1 mL of octane, $p_{\text{H}_2} = 5 \text{ MPa}$, $T = 200 \text{ }^\circ\text{C}$. Yields were determined by GC using *n*-octane as the internal standard.

Next, the effect of H₂ pressure and concentration of the carboxylic acid on the rate of the Re(5)/TiO₂ promoted hydrogenation of 3-phenylpropionic acid in *n*-octane was investigated (**Fig. 9**). The hydrogenation rate was defined as the sum of the alcohol formation rate and one half of the ester formation rate. The results show that the rate of hydrogenation increases with

increasing H₂ pressure. The reaction order with respect to H₂ pressure was calculated to be +0.88, which suggests that H₂ is involved in a kinetically important step. On the other hand, the reaction order with respect to 3-phenylpropionic acid concentration was calculated to be +0.03, indicating that the carboxylic acid is strongly adsorbed on the surface of the Re(5)/TiO₂ catalyst. This observation is in agreement with those made in the selectivity experiments described above.

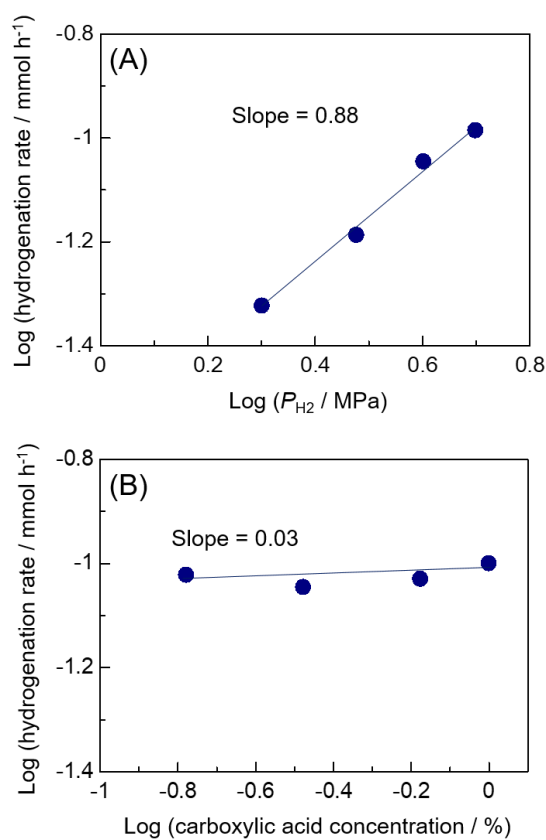


Figure 9. Effect of (A) the H₂ pressure and (B) 3-phenylpropionic acid concentration on the rate of hydrogenation of 3-phenylpropionic acid catalyzed by Re(5)/TiO₂. Hydrogenation rate = alcohol formation rate plus one half the ester formation rate. Pre-treatment: H₂ (30 mL min⁻¹), 500 °C, 0.5 h; reaction conditions: 2 mol% Re, 1 mmol 3-phenylpropionic acid, 1 mL *n*-octane as a solvent, 140 °C, 3 h.

Hydrogenation reactions of 3-phenylpropionic acid using Re(*x*)/TiO₂ (*x* = 0.2, 1, 5, 10, and 20 wt%) were conducted in order to determine the optimal catalyst loading and to investigate the effect of the size of the Re species on the process (**Figure 10A**). The highest rate of hydrogenation was achieved using Re(5)/TiO₂. The results indicate that the hydrogenation process occurs more efficiently when the catalytic system contains Re species of sub-nanometer to a few nanometer sizes (see HAADF-STEM images of Re(*x*)/TiO₂ in **Fig. 11**). We also explored hydrogenation reactions of 3-phenylpropionic acid promoted by Re(5)/TiO₂ following

pretreatment with H₂ at different temperatures (200, 300, 500, 700, and 900 °C). The results displayed in **Fig. 10B** show that Re(5)/TiO₂, pre-treated at 500 °C with H₂, is optimal for this process. These observations along with those made in the catalyst characterization studies using techniques such as XANES and XPS demonstrate that the average oxidation state of the Re species that promote hydrogenation should be higher than 0 and below +4. At this time, we do not know the precise nature of the catalytically active Re species, and the catalyst may contain a variety of Re species in different oxidation states. We also cannot exclude the possibility that these mixtures of Re species in different oxidation states and the interfaces between them are required for the efficient progression of the reaction. It should also be noted that reductive pretreatment at temperatures beyond 500 °C leads to sintering of the Re species (see HAADF-STEM images in **Fig. 12**) and a lowering of the activity of the catalyst.

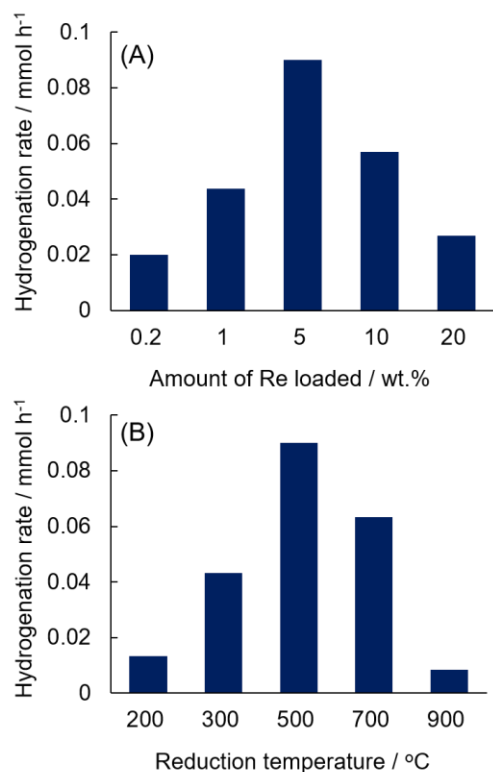


Figure 10. Effect of (A) the Re loading and (B) pretreatment reduction temperature on the hydrogenation of 3-phenylpropionic acid catalyzed by Re(x)/TiO₂ (Re(5)/TiO₂ was used for the reduction temperature dependence). Hydrogenation rate = the sum of the alcohol formation rate and one half the ester formation rate. Pre-treatment: H₂ (30 mL min⁻¹), 500 °C, 0.5 h; reaction conditions: 2 mol% Re, 1 mmol 3-phenylpropionic acid, no solvent, 5 MPa H₂, 140 °C, 3 h.

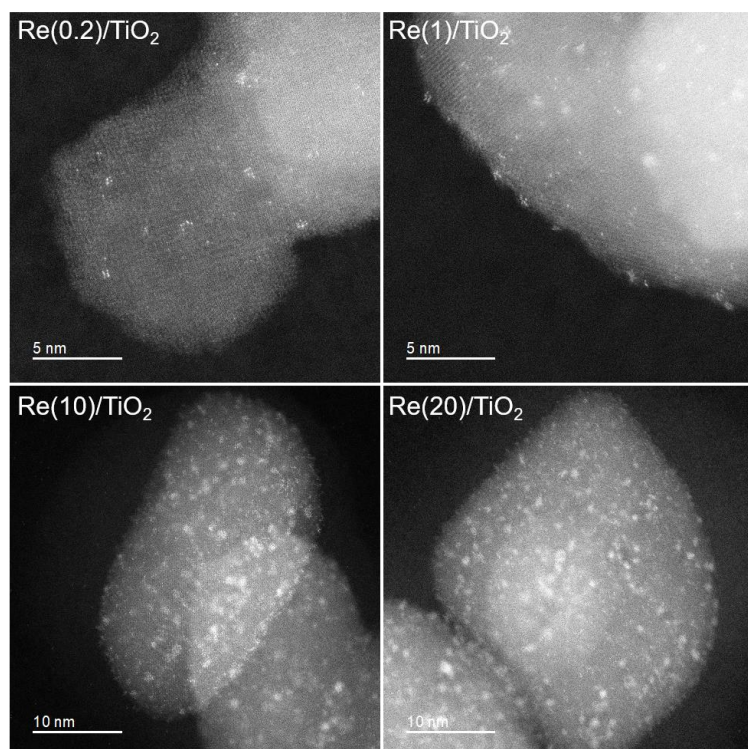


Figure 11. HAADF-STEM images for $\text{Re}(x)/\text{TiO}_2$ ($x = 0.2, 1, 10, 20$ wt%) reduced at $T = 500$ °C under a flow of H_2 . Due to the experimental setup, the samples were exposed to air before images were recorded.

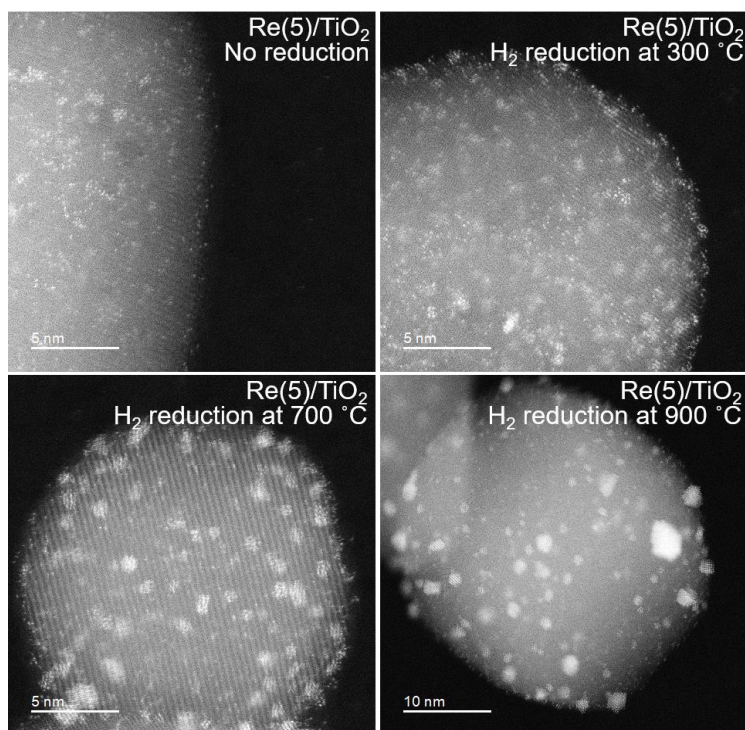
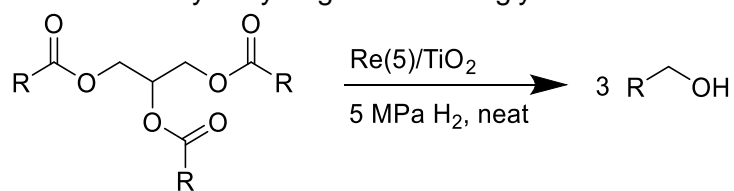


Figure 12. HAADF-STEM images for $\text{Re}(5)/\text{TiO}_2$ without reduction and reduced at $T = 300, 700,$ and 900 °C under a flow of H_2 . Due to the experimental setup, the samples were exposed to air before images were recorded.

Selective hydrogenation of triglycerides

Because it is highly desired to replace fossil fuels as chemical and energy sources, much attention has been given to the conversion of biomass materials into fuels and chemicals.⁶⁶⁻⁷² Among biomass materials, triglycerides serve as the most important feedstocks for the synthesis of fatty amines, amides, nitriles and alcohols used in the production of surfactants, polymers and other useful commodity chemicals.⁷³ However, benign manufacturing processes for the direct transformation of triglycerides into these chemicals are not well developed. In this study, we demonstrated the potential applicability of Re(5)/TiO₂ catalyzed hydrogenations to direct transformation of triglycerides into the corresponding alcohols. As shown in **Table 3**, hydrogenation reactions of a variety of triglycerides with different chain lengths (C4-C18), using Re(5)/TiO₂ pretreated with 5 MPa H₂ at 500 °C, leads to formation of the corresponding alcohols without any solvents.

Table 3. Catalytic hydrogenation of triglycerides to form alcohols using Re(5)/TiO₂.^a

Entry	Substrate	Product	Temperature (°C)	Time (h)	Alcohol yield (%)
1		C ₁₇ H ₃₅ -OH	230	30	89 (82)
2		C ₁₅ H ₃₁ -OH	230	30	89 (84)
3		C ₁₃ H ₂₇ -OH	230	24	90 (86)
4		C ₁₁ H ₂₃ -OH	230	30	84 (81)
5		C ₉ H ₁₉ -OH	230	24	86 (80)
6		C ₇ H ₁₅ -OH	230	24	66 (62)
7		C ₅ H ₁₁ -OH	220	36	82 (78)
8		C ₃ H ₇ -OH	220	24	87 (83)

^aReactions were performed under $p_{\text{H}_2} = 5$ MPa. Yields were determined by GC using *n*-dodecane as the internal standard. Yields of isolated products are given in parenthesis.

4. Conclusions

In the effort described above, we carried out a detailed mechanistic study of selective hydrogenation reactions of carboxylic acids and their derivatives promoted by TiO₂-supported Re (Re/TiO₂) catalysts. The results show that the optimal Re loading and H₂ pretreatment temperature are 5 wt% and 500 °C, respectively. Detailed investigations using XRD, XAFS, and XPS as well as the STEM measurements were carried out to identify the catalytically active Re species. The results indicate that the active Re species should have sub-nanometer to a few nanometer sizes and average oxidation states higher than 0 and below +4. Also, we found that Re on a TiO₂ support is more highly dispersed than it is on the other supports such as SiO₂, Al₂O₃, and Carbon. Furthermore, we observed that the Re/TiO₂ promotes hydrogenation of propylbenzene to form propylcyclohexane, indicating this catalytic system has the intrinsic ability to promote hydrogenation of aromatic rings. However, when either lauric acid or lauryl alcohol (or an amide or amine) is present in the reaction mixture, reductive dearomatization reactions do not take place. These observations suggest that strong adsorption of carboxylic acids and alcohols (or amides and amines) to the catalyst, caused by the oxophilic nature of Re, suppresses the dearomatization process. This finding is in accord with results of kinetic investigations, which showed that the order of the reaction with respect to the carboxylic acid is +0.03. Finally, in a study aimed at demonstrating that the catalytic system can be used for conversion of biomass materials into value-added chemicals, we observed that Re/TiO₂ can be employed to convert triglycerides into the corresponding alcohols.

Conflicts of interest

There are no conflicts to declare.

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Note and references

- 1 A. M. Smith and R. Whyman, *Chem. Rev.*, 2014, **114**, 5477–5510.
- 2 J. Pritchard, G. A. Filonenko, R. van Putten, E. J. M. Hensen and E. A. Pidko, *Chem. Soc. Rev.*, 2015, **44**, 3808–3833.
- 3 K. Tomishige, Y. Nakagawa and M. Tamura, *Green Chem.*, 2017, **19**, 2876–2924.
- 4 T. Toyao, S. M. A. Hakim Siddiki, K. Kon and K. Shimizu, *Chem. Rec.*, 2018, **18**, 1374–1393.
- 5 J. Verduyck, R. Coeck and D. E. De Vos, *ACS Sustain. Chem. Eng.*, 2017, **5**, 3290–3295.
- 6 P. A. Dub and T. Ikariya, *ACS Catal.*, 2012, **2**, 1718–1741.
- 7 D. S. Thakur and A. Kundu, *J. Am. Oil Chem. Soc.*, 2016, **93**, 1575–1593.
- 8 Y. Pouilloux, *J. Catal.*, 1998, **176**, 215–224.
- 9 H. Adkins and B. Wojcik, *J. Am. Chem. Soc.*, 1934, **56**, 2419–2424.
- 10 C. Hirose, N. Wakasa and T. Fuchikami, *Tetrahedron Lett.*, 1996, **37**, 6749–6752.
- 11 G. Beamson, A. J. Papworth, C. Philipps, A. M. Smith and R. Whyman, *Adv. Synth. Catal.*, 2010, **352**, 869–883.
- 12 M. Stein and B. Breit, *Angew. Chemie - Int. Ed.*, 2013, **52**, 2231–2234.
- 13 J. Coetzee, H. G. Manyar, C. Hardacre and D. J. Cole-Hamilton, *ChemCatChem*, 2013, **5**, 2843–2847.
- 14 K. Shimizu, W. Onodera, A. S. Touchy, S. M. A. H. Siddiki, T. Toyao and K. Kon, *ChemistrySelect*, 2016, **1**, 736–740.
- 15 R. A. Grey, G. P. Pez, A. Wallo and J. Corsi, *J. Chem. Soc. Chem. Commun.*, 1980, 783–784.
- 16 Y. Hara, H. Inagaki, S. Nishimura and K. Wada, *Chem. Lett.*, 1992, **21**, 1983–1986.
- 17 X. Cui, Y. Li, C. Topf, K. Junge and M. Beller, *Angew. Chemie - Int. Ed.*, 2015, **54**, 10596–10599.
- 18 L. Deng, B. Kang, U. Englert, J. Klankermayer and R. Palkovits, *ChemSusChem*, 2016, **9**, 177–180.
- 19 X. Fang, M. Sun, J. Zheng, B. Li, L. Ye, X. Wang, Z. Cao, H. Zhu and Y. Yuan, *Sci. Rep.*, 2017, **7**, 1–9.
- 20 S. Yoshioka and S. Saito, *Chem. Commun.*, 2018, **54**, 13319–13330.
- 21 H. T. Teunissen and C. J. Elsevier, *Chem. Commun.*, 1998, 1367–1368.
- 22 K. Nomura, H. Ogura and Y. Imanishi, *J. Mol. Catal. A Chem.*, 2001, **166**, 345–349.
- 23 J. Zhang, G. Leitner, Y. Ben-David and D. Milstein, *Angew. Chemie - Int. Ed.*, 2006, **45**, 1113–1115.
- 24 W. Kuriyama, Y. Ino, O. Ogata, N. Sayo and T. Saito, *Adv. Synth. Catal.*, 2010, **352**, 92–96.
- 25 M. Ito, T. Ootsuka, R. Watari, A. Shiibashi, A. Himizu and T. Ikariya, *J. Am. Chem. Soc.*, 2011, **133**, 4240–4242.
- 26 T. vom Stein, M. Meuresch, D. Limper, M. Schmitz, M. Hölscher, J. Coetzee, D. J. Cole-Hamilton, J. Klankermayer and W. Leitner, *J. Am. Chem. Soc.*, 2014, **136**, 13217–13225.
- 27 M. Naruto and S. Saito, *Nat. Commun.*, 2015, **6**, 8140.
- 28 T. J. Korstanje, J. Ivar van der Vlugt, C. J. Elsevier and B. de Bruin, *Science*, 2015, **350**, 298–302.
- 29 T. Toyao, S. M. A. H. Siddiki, A. S. Touchy, W. Onodera, K. Kon, Y. Morita, T. Kamachi, K. Yoshizawa and K. Shimizu, *Chem. - A Eur. J.*, 2017, **23**, 1001–1006.

- 30 T. Toyao, S. M. A. H. Siddiki, Y. Morita, T. Kamachi, A. S. Touchy, W. Onodera, K. Kon, S. Furukawa, H. Ariga, K. Asakura, K. Yoshizawa and K. Shimizu, *Chem. - A Eur. J.*, 2017, **23**, 14848–14859.
- 31 K. W. Ting, T. Toyao, S. M. A. H. Siddiki and K. I. Shimizu, *ACS Catal.*, 2019, **9**, 3685–3693.
- 32 N. D. Spencer and G. A. Somorjai, *J. Catal.*, 1982, **72**, 142–146.
- 33 L. Wang, R. Ohnishi and M. Ichikawa, *J. Catal.*, 2000, **190**, 276–283.
- 34 J. Ullrich and B. Breit, *ACS Catal.*, 2018, **8**, 785–789.
- 35 H. S. Lacheen, P. J. Cordeiro and E. Iglesia, *Chem. - A Eur. J.*, 2007, **13**, 3048–3057.
- 36 X. She, H. M. Brown, X. Zhang, B. K. Ahring and Y. Wang, *ChemSusChem*, 2011, **4**, 1071–1073.
- 37 M. Chia, Y. J. Pagán-Torres, D. D. Hibbitts, Q. Tan, H. N. Pham, A. K. Datye, M. Neurock, R. J. Davis, J. A. Dumesic, Y. J. Pag, Y. J. Pagán-Torres, D. D. Hibbitts, Q. Tan, H. N. Pham, A. K. Datye, M. Neurock, R. J. Davis and J. A. Dumesic, *J. Am. Chem. Soc.*, 2011, **133**, 12675–12689.
- 38 S. Lwin, C. Keturakis, J. Handzlik, P. Sautet, Y. Li, A. I. Frenkel and I. E. Wachs, *ACS Catal.*, 2015, **5**, 1432–1444.
- 39 N. Ota, M. Tamura, Y. Nakagawa, K. Okumura and K. Tomishige, *Angew. Chemie - Int. Ed.*, 2015, **54**, 1897–1900.
- 40 L. Sandbrink, E. Klindtworth, H. U. Islam, A. M. Beale and R. Palkovits, *ACS Catal.*, 2016, **6**, 677–680.
- 41 S. Tazawa, N. Ota, M. Tamura, Y. Nakagawa, K. Okumura and K. Tomishige, *ACS Catal.*, 2016, **6**, 6393–6397.
- 42 M. Tamura, N. Yuasa, J. Cao, Y. Nakagawa and K. Tomishige, *Angew. Chemie - Int. Ed.*, 2018, **57**, 8058–8062.
- 43 M. F. L. Johnson and V. M. LeRoy, *J. Catal.*, 1974, **35**, 434–440.
- 44 S. Tatarenko, P. Dolle, R. Morancho, M. Alnot, J. J. Ehrhardt and R. Ducros, *Surf. Sci.*, 1983, **134**, L505-L512.
- 45 F. Yang, D. Liu, H. Wang, X. Liu, J. Han, Q. Ge and X. Zhu, *J. Catal.*, 2017, **349**, 84–97.
- 46 S. R. Bare, F. D. Vila, M. E. Charochak, S. Prabhakar, W. J. Bradley, C. Jaye, D. A. Fischer, S. T. Hayashi, S. A. Bradley and J. J. Rehr, *ACS Catal.*, 2017, **7**, 1452–1461.
- 47 M. T. Greiner, T. C. R. Rocha, B. Johnson, A. Klyushin, A. Knop-Gericke and R. Schlögl, *Zeitschrift für Phys. Chemie*, 2014, **228**, 521–541.
- 48 H. Peled and M. Asscher, *Surf. Sci.*, 1987, **183**, 201–215.
- 49 A. S. Y. Chan, W. Chen, H. Wang, J. E. Rowe and T. E. Madey, *J. Phys. Chem. B*, 2004, **108**, 14643–14651.
- 50 S. R. Bare, S. D. Kelly, F. D. Vila, E. Boldingh, E. Karapetrova, J. Kas, G. E. Mickelson, F. S. Modica, N. Yang and J. J. Rehr, *J. Phys. Chem. C*, 2011, **115**, 5740–5755.
- 51 S. Koso, H. Watanabe, K. Okumura, Y. Nakagawa and K. Tomishige, *J. Phys. Chem. C*, 2012, **116**, 3079–3090.
- 52 K. Hahn and M. Mavrikakis, *Top. Catal.*, 2014, **57**, 54–68.
- 53 Y. Takeda, M. Tamura, Y. Nakagawa, K. Okumura and K. Tomishige, *ACS Catal.*, 2015, **5**, 7034–7047.

- 54 S. Lwin, Y. Li, A. I. Frenkel and I. E. Wachs, *ACS Catal.*, 2015, **5**, 6807–6814.
- 55 N. Ota, M. Tamura, Y. Nakagawa, K. Okumura and K. Tomishige, *ACS Catal.*, 2016, **6**, 3213–3226.
- 56 B. Mitra, X. Gao, I. E. Wachs, A. M. Hirt and G. Deo, *Phys. Chem. Chem. Phys.*, 2001, **3**, 1144–1152.
- 57 J. Okal, L. K. Episki, L. Krajczyk and W. Tylus, *J. Catal.*, 2003, **219**, 362–371.
- 58 J. Okal, *Appl. Catal. A Gen.*, 2005, **287**, 214–220.
- 59 X. She, J. H. Kwak, J. Sun, J. Hu, M. Y. Hu, C. Wang, C. H. F. Peden and Y. Wang, *ACS Catal.*, 2012, **2**, 1020–1026.
- 60 D. D. Falcone, J. H. Hack, A. Y. Klyushin, A. Knop-Gericke, R. Schlögl and R. J. Davis, *ACS Catal.*, 2015, **5**, 5679–5695.
- 61 B. K. Ly, B. Tapin, M. Aouine, P. Delichere, F. Epron, C. Pinel, C. Especel and M. Besson, *ChemCatChem*, 2015, **7**, 2161–2178.
- 62 K. Liu, J. Pritchard, L. Lu, R. Van Putten, M. W. G. M. Verhoeven, M. Schmitkamp, X. Huang, L. Lefort, C. J. Kiely, E. J. M. Hensen and E. A. Pidko, *Chem. Commun.*, 2017, **53**, 9761–9764.
- 63 S. T. Thompson and H. H. Lamb, *J. Catal.*, 2017, **350**, 111–121.
- 64 X. Di, C. Li, G. Lafaye, C. Especel, F. Epron and C. Liang, *Catal. Sci. Technol.*, 2017, **7**, 5212–5223.
- 65 Y. Nakagawa, S. Tazawa, T. Wang, M. Tamura, N. Hiyoshi, K. Okumura and K. Tomishige, *ACS Catal.*, 2018, **8**, 584–595.
- 66 B. Peng, Y. Yao, C. Zhao and J. A. Lercher, *Angew. Chemie - Int. Ed.*, 2012, **51**, 2072–2075.
- 67 R. Beerthuis, G. Rothenberg and N. R. Shiju, *Green Chem.*, 2015, **17**, 1341–1361.
- 68 A. Suknev, V. Zaikovskii, V. Kaichev, E. Paukshtis, E. Sadovskaya and B. Bal'Zhinimaev, *J. Energy Chem.*, 2015, **24**, 646–654.
- 69 B. Rozmysłowicz, A. Kirilin, A. Aho, H. Manyar, C. Hardacre, J. Wärnå, T. Salmi and D. Y. Murzin, *J. Catal.*, 2015, **328**, 197–207.
- 70 A. M. Robinson, J. E. Hensley and J. Will Medlin, *ACS Catal.*, 2016, **6**, 5026–5043.
- 71 S. Liu, T. Simonetti, W. Zheng and B. Saha, *ChemSusChem*, 2018, **11**, 1446 – 1454.
- 72 M. Wang, M. Liu, H. Li, Z. Zhao, X. Zhang and F. Wang, *ACS Catal.*, 2018, **8**, 6837–6843.
- 73 M. A. R. Jamil, S. M. A. H. Siddiki, A. S. Touchy, M. N. Rashed, S. S. Poly, Y. Jing, K. W. Ting, T. Toyao, Z. Maeno and K. Shimizu, *ChemSusChem*, 2019, **12**, 3115–3125.