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Wastewater treatment valorisation by simultaneously removing and recovering phosphate and ammonia from municipal effluents using a mechano-thermo activated magnesite technology

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Abstract

Phosphate and nitrate enrichment largely impair aquatic ecosystem functions and services, thus comprising an emerging problem of environmental concern. The problem pertains to developing countries where phosphate and nitrate discharge to surface water is on the rise due to a rapid growth in population. Herein, these pollutants (phosphate and ammonia) were removed from real municipal wastewater using a simple, fast, and cost-effective process. Raw cryptocrystalline magnesite, a mineral abundant in South Africa, was simply milled and calcined (mechano-thermo processing) in order to produce the activated magnesite (feed). The feed was then used in batch processing for pollutants adsorption and precipitation from real wastewater. The process was optimised by varying the treatment or contact time, feed dosage, concentration, pH, and temperature. The feed and product mineral (produced sludge) were characterised using X-ray Diffraction (XRD), field emission scanning electron microscopy (FESEM) compatible with energy dispersive spectroscopy (EDS), and Fourier Transform Infrared Spectrometer (FTIR). It was identified that the optimal conditions differed for each pollutant, highlighting the importance of tailoring the process to fit the local wastewater characteristics and as part of a treatment train system. Specifically, maximum P removal was achieved after 5 min of mixing, using 1 g L^{-1} of feed, 123 mg L^{-1} initial phosphate concentration, pH 8 – 10, and was not affected by temperature variations; whereas, for ammonia removal, optimal conditions were 180 min, 16 g L^{-1} feed dosage, 80 mg L^{-1} initial concentration, pH 10 and temperature > 45 °C. The optimal conditions for the removal of both pollutants from real wastewater were 30 min, 8 g L^{-1} dosage, 7.5 pH, 35 ^{o}C. Furthermore, Mg and Ca concentration was found to influence the process. Reduction in total dissolved solids and electrical conductivity suggest an attenuation of chemical species. Characterisation revealed that the product mineral obtained under the optimal conditions for pollutants removal is rich in quartz, periclase, brucite, calcite, magnesite, and struvite. This was further supported by the FTIR results, which indicated the presence of Mg-O, PO_4^{3-}, N-H and –OH stretches. In addition, the EDS verified the presence of Mg, Ca and P in product mineral. Results are suggestive of the high efficiency of the mechano-thermo activated magnesite treatment process for P and N removal and struvite crystallization. Thus, this technology could valorise municipal wastewater effluents and open new horizons for the

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effective and sustainable management of wastewater effluents, since struvite can replace the mined phosphate fertilizers, which are rapidly depleting, in the agriculture industry.

Keywords: Mechano-thermo treatment; Pre-treated magnesite; Phosphorus; Nitrogen; municipal effluents; magnesium ammonium phosphate (struvite), sustainable

1 Introduction

Nowadays, research has focused on identifying feasible and practical approaches to manage high levels of phosphate and ammonia contained in wastewater streams, especially in developing countries where wastewater could be poorly treated. A wide range of technologies, based on different techniques and mechanisms, have been developed and explored, with biological processes, adsorption, filtration, bio-sorption, phytoremediation and crystallisation being popular for phosphate and ammonia removal from wastewater (Li et al., 2017; Peng et al., 2018; Satoshi et al., 2013). Adsorption has been used to remove phosphate and ammonia from wastewater, typically at bench and semi-industrial scale using clay minerals, metals and their composites, while filtration techniques have been explored as well (Goh et al., 2008; Huang et al., 2017; Satoshi et al., 2013; Yagi and Fukushi, 2012; Zulfiqar et al., 2014). Adsorption followed by crystals settling is particularly popular (Li et al., 2015; Li et al., 2017), with Ca^{2+} or Mg^{2+} based materials being typically used as precursors (Li et al., 2017; Peng et al., 2018; Stolzenburg et al., 2015). Other adsorbents include elements that have high affinity to phosphate and ammonia, mainly trivalent metals, such as Fe^{3+} and Al^{3+}, which have the added benefit of regeneration (Benyoucef and Amrani, 2012; Chiban et al., 2012; Deng et al., 2016; Egle et al., 2016; Gao et al., 2018; Goh et al., 2008).

South Africa is well-endowed with cryptocrystalline magnesite, which makes it a perfect candidate material for sustainable wastewater treatment and phosphate and ammonia removal (Magagane et al., 2019). This material is rich in Mg and it is well established that treating wastewater with Mg^{2+} based minerals can result to the formation of magnesium ammonium phosphate (struvite), which has a number of industrial applications (Bayuseno and Schmahl, 2018; Kozik et al., 2014; Stolzenburg et al., 2015; Sutiyono et al., 2016; Wang et al., 2017); albeit its production depends on a number of factors such as treatment or contact time, dosage, pH, and temperature (Li et al., 2017). A number of salts that bear Mg^{2+} ions have been used for struvite crystallisation, including $MgCl_2$, Mg $(OH)_2$ and MgO (Peng et al., 2018; Stolzenburg et al., 2015; Wang et al., 2017); however, these are mainly commercially

available salts, typically associated with high manufacturing costs which hinders the commercial application of the process (Muhmood et al., 2019).

For the case of South Africa, a developing country, a low cost and easily accessible replacement Mg source is required to make the process commercially attractive. Due to its high Mg content, amorphous and cryptocrystalline magnesite could be a possible source of Mg^{2+} and Ca^{2+} for the removal of phosphate and ammonia from aqueous matrices. These materials have been explored for mine water treatment and the management of other contaminants (Masindi, 2017; Masindi et al., 2016a; Masindi et al., 2018a; Masindi et al., 2018b; Masindi et al., 2016b). Specifically, Masindi et al. (2016a) highlighted that cryptocrystalline magnesite, and its pre-treated derivative (Magagane et al., 2019), are characterised of Mg, Ca, Fe and Si sandwiched together, hence giving it an up-hand due to hybrid properties therein. The mineral phases can also aid in the removal of phosphate and ammonia from wastewater, eventually leading to the formation of struvite. It dould be noted tha the precipitation of struvite from wastewater streams is gaining the attention of the scientific community due to its potential for phosphorus recovering for fertilization (Yan and Shih, 2016), and particularly owing to the fact that phosphate rock, a non-renewable resource, is predicted to be depleted within the next 100 years (Li et al., 2019). Furthermore, apart from calcium and magnesium, (Torzewska and Różalski, 2015) noted that certain bacteria, such as the genus Proteus, can also play a role in inducing struvite crystallization. (Wąsik et al., 2017) also noted that coliform bacteria, in the presence of magnesium and calcium ions, among others, can created suitable conditions for the formation of struvite crystals. (Yan and Shih, 2016) studied the effects of calcium and ferric ions on struvite precipitation, while (Liu and Wang, 2019) noted that wastewater with the Ca/Mg ratio above 1/2 is not suitable for phosphorus recovery, due to the low proportion of struvite recovery. Finally, Li et al. (2019) reviewed the key aspects for struvite process design and development, the research trends, product application and process economics.

Herein, based on preliminary results of our group (Mavhungu et al., 2018), we comprehensively examine the feasibility of South African pre-treated magnesite, rather than commercial salts, for phosphate and ammonia removal from real wastewater. This study also sheds light on the mechanisms governing the removal of pollutants from aquatic matrices through struvite precipitation. Furthermore, struvite, a process co-product, has an economic value, since if recovered could be suitable for reuse in agriculture (Taddeo et al., 2018).

Therefore, its recovery could both reduce the cost and environmental impact of the wastewater treatment process, which will be examined in future works of our group.

2 Materials and methods

2.1 Wastewater and raw cryptocrystalline magnesite collection

Real untreated wastewater, rich in phosphate and ammonia, was collected from a municipal wastewater treatment facility in Pretoria, South Africa. The treatment facility receives wastewater from a number of activities from the surrounding area. High-density polyethylene (HDPE) wide-mouth bottles were used for sample collection. Solids and debris were removed by filtration, using Macherey-Nagel filter papers (MN 615. Ø125mm), and the samples were used immediately thereafter. In the context of this work, the wastewater was diluted, as required, to reach the desired phosphate and ammonia concentration. Raw cryptocrystalline magnesite samples were collected from an abandoned magnesite mine in Folovhodwe, Limpopo Province, South Africa.

2.2 Production of the mechano-thermo activated magnesite

The collected raw cryptocrystalline magnesite was first milled into a fine powder, using a vibratory ball miller at 500 rpm for 15 min. Thence, the fine powder was calcined at 1,000 °C, using a furnace as reported in our previous study (Magagane et al., 2019). Finally, the samples were milled again, under the same conditions, and then sieved through 32 microns perforated sieve. The thermo-mechano (i.e. calcined and milled) activated magnesite samples, feed thereafter, were then stored in zip-lock plastic bags until utilisation for simultaneous adsorption of phosphate and ammonia.

2.3 Quality control and quality assurance

Experiments were performed in triplicate, giving similar results, and are reported as mean average values. For quality control and assurance, ISO-accredited laboratories and NIST standards were used for sample analyses. To ensure reliable, accurate, and high precision results inter-laboratory analysis was also undertaken, i.e. for validation and verification purposes.

2.4 Process optimisation

For process optimization and to obtain an in-depth understanding of the factors influencing phosphate and ammonia removal, a number of operational parameters were evaluated. These include: i) contact time (i.e. mixing duration of the feed-wastewater solution); ii) feed dosage (i.e. the amount of mechano-thermo activated magnesite used during treatment); iii) chemical species concentration (i.e. phosphate and ammonia initial concertation); iv) pH; and v) temperature. Experiments were performed in 500 mL volumetric flasks (low-form beaker), stirred at 500 rpm using an overhead stirrer. To gain insight, the one-factor-at-a-time (OFAT) method was used, i.e. each time one parameter was varied while the others were fixed, as shown in Table 1.

Table 1: Process parameters assayed for the optimisation of phosphate and ammonia removal from wastewater.

Parameter	**Units**	**Data variation range**
Contact time	minutes (min)	1; 5; 10; 30; 60;120; 150; 180; 240; and 300
Dosage	grams*	0.1; 0.5; 1; 2; 3; 4; 5; 8; and 10
Concentration	mg L^{-1} or ppm	2; 4; 8; 12; 16; 21; 33; 41; 82; and 123
pH of solution	N/A	1, 2, 3, 4, 5, 6, 7, 8, 9, 10, 11, and 12
Temperature	°C	35; 45; 55; 65; and 75
*per 500 mL of wastewater		

The effect of the initial phosphate and ammonia concentration was studied by diluting the raw municipal wastewater to the desired concentration (Table 1). The pH was adjusted using 0.1 M NaOH and/or 0.1 M HCl, as required. A 4-decimal place analytical balance was used to calculate feed dosages. Finally, the temperature was regulated using a hotplate stirrer and the effect of contact time was examined by agitating the wastewater-feed solution at the time intervals specified in Table1.

2.5 Percentage removal

The amount of phosphate and ammonia removed from municipal wastewater was estimated using equation (1), i.e.:

$$\% \text{ Removal} = \frac{C_0 - C_e}{C_0} \times 100 \quad (1)$$

where C_0 is the initial phosphate or ammonia concentration and C_e is the final concentration, after treatment, of phosphate or ammonia.

2.6 Analytical techniques

The pH, Electrical conductivity (EC) and Total Dissolved Solids (TDS) were monitored using a multi-parameter probe (HANNA instrument, HI9828). Chemical species concentrations were determined by means of inductively coupled plasma mass spectrometry (ICP-MS) (7500ce, Agilent, Alpharetta, GA, USA). The raw magnesite, the feed (mechano-thermo activated magnesite) and product mineral (i.e. the produced sludge that contains struvite) properties were examined by the analytical techniques detailed in Table 2.

Table 2: The equipment used for characterisation of feed and product mineral.

Parameter	Analytical Technique	Model
Mineralogical properties	X-ray Diffraction (XRD)	PANalytical X'Pert PRO-diffractometer equipped with Philips PW 1710 Diffractometer with graphite secondary monochromatic source
Functional groups	Fourier Transform Infrared Spectrometer (FTIR)	Perkin-Elmer Spectrum 100 Fourier Transform Infrared Spectrometer (FTIR) equipped with Perkin-Elmer Precisely Universal Attenuated Total Reflectance (ATR) sampling accessory with a diamond crystal.
Morphology, Mapping and elements	Field Emission Scanning Electron Microscope (FESEM) equipped with Energy-dispersive X-ray spectroscopy (EDS)	Auriga Cobra FIB-FESEM, Carl Zeiss FE-SEM, Germany

3 Results and discussions

3.1 Characterisation of the raw magnesite, feed, and product mineral

3.1.1 Mineralogical analysis

The mineralogical properties of the raw magnesite, the feed (mechano-thermo activated magnesite) and product mineral (produced sludge) were estimated by means of X-ray Diffraction (XRD) and are shown in Figure 1.

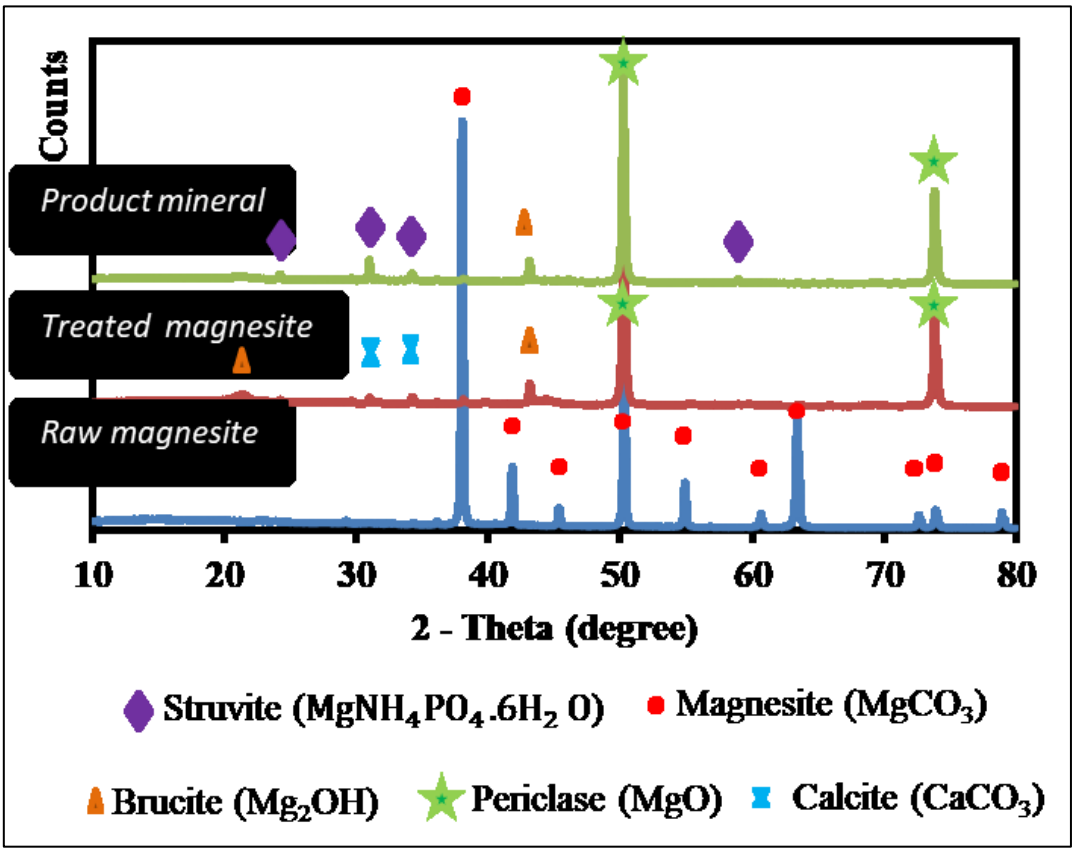

Figure 1: The mineralogical properties of raw magnesite, feed, and product mineral.

As shown in Figure 1 there is a change in mineral phases after calcination. Initially, it was observed that the raw magnesite contains pure magnesium carbonate. However, after calcination, new phases were observed in the feed and these include periclase, brucite, and calcite, in addition to amorphous fractions. Finally, the produced sludge (product mineral) was found to contain struvite, periclase, and brucite. This is an indication that phosphate and ammonia are adsorbed by the feed and form struvite, as denoted in equation (2):

$$Mg^{2+} + NH_4^+ + PO_4^{3-} + 6H_2O \rightarrow MgNH_4PO_4.6H_2O \quad (2)$$

Similar results were reported by Sutiyono et al. (2016). Furthermore, the obtained peaks at 21, 23, 31, and 34 2theta degrees are in good agreement with those reported by Li et al. (2017) and Gao et al. (2018). This confirms that the reaction between the feed and the wastewater leads to the formation of magnesium ammonium phosphate (struvite). It should

be noted that since the activated magnesite (feed) reacts with the wastewater to produce struvite, it cannot be regenerated or recycled in the process. The other components observed, such as the periclase and brucite, could be also beneficial since it is envisaged that the product mineral (struvite) would be used for agriculture purposes. Finally, the presence of Mg and Ca in the feed contributes towards increasing the pH value of the wastewater and generating Mg^{2+} and Ca^{2+}. This is shown in eq. (3) – (4):

$$MgO + H_2O \rightarrow Mg^{2+} + 2OH^- \quad (3)$$

$$CaO + H_2O \rightarrow Ca^{2+} + 2OH^- \quad (4)$$

Mg^{2+} will then scavenge the NH_4^+ and PO_4^{3-} to form a complex, as denoted in equation 2.

3.1.2 Morphological properties

The morphological properties of the feed and product mineral were measured by a Field Emission Scanning Electron Microscope (FESEM). The clear, ultra-high resolution and low electrostatically distorted images obtained by the FESEM are shown in Figure 2.

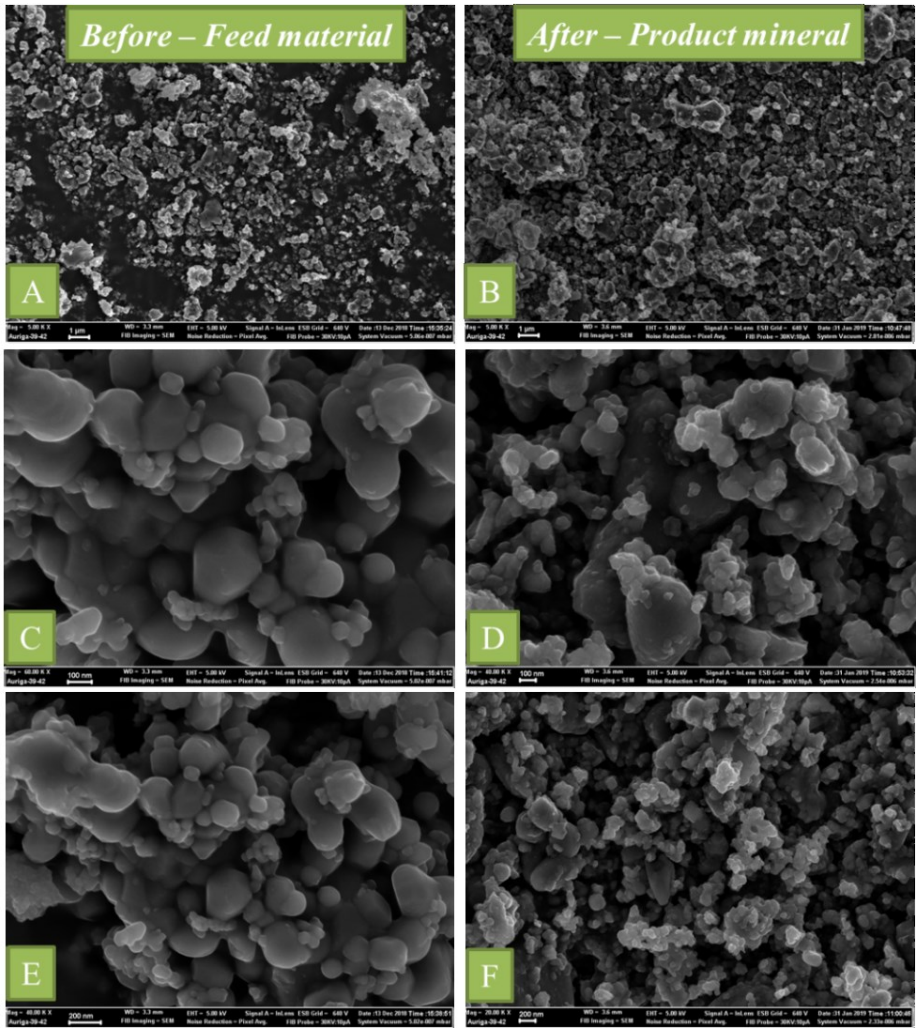

Figure 2: The morphological properties of the feed (left-hand-side) and product mineral (right-hand-side) obtained by FESEM at 1 µm (A and B), 100 nm (C and D), and 200 nm (E and F).

From the comparative analysis between the feed's and the product mineral's morphological properties shown in Figure 2, it is evident that the feed contains nano-sheets, with hexagonal structures uniformly distributed across the surface. Contrary, the product mineral portrayed nanosheet-like structures with reduced size. Therefore, the results are suggestive of a possible dissolution of the feed material or, more likely, the formation of new mineral phases, as will be discussed further down. These results are also consistent with those reported by Heraldy et al. (2017).

3.1.3 Elemental composition of feed and product mineral

As mentioned above, the FESEM was equipped with an Energy-dispersive X-ray spectroscopy (EDS) detector, which was used towards identifying the elemental composition of the feed and of the product mineral.

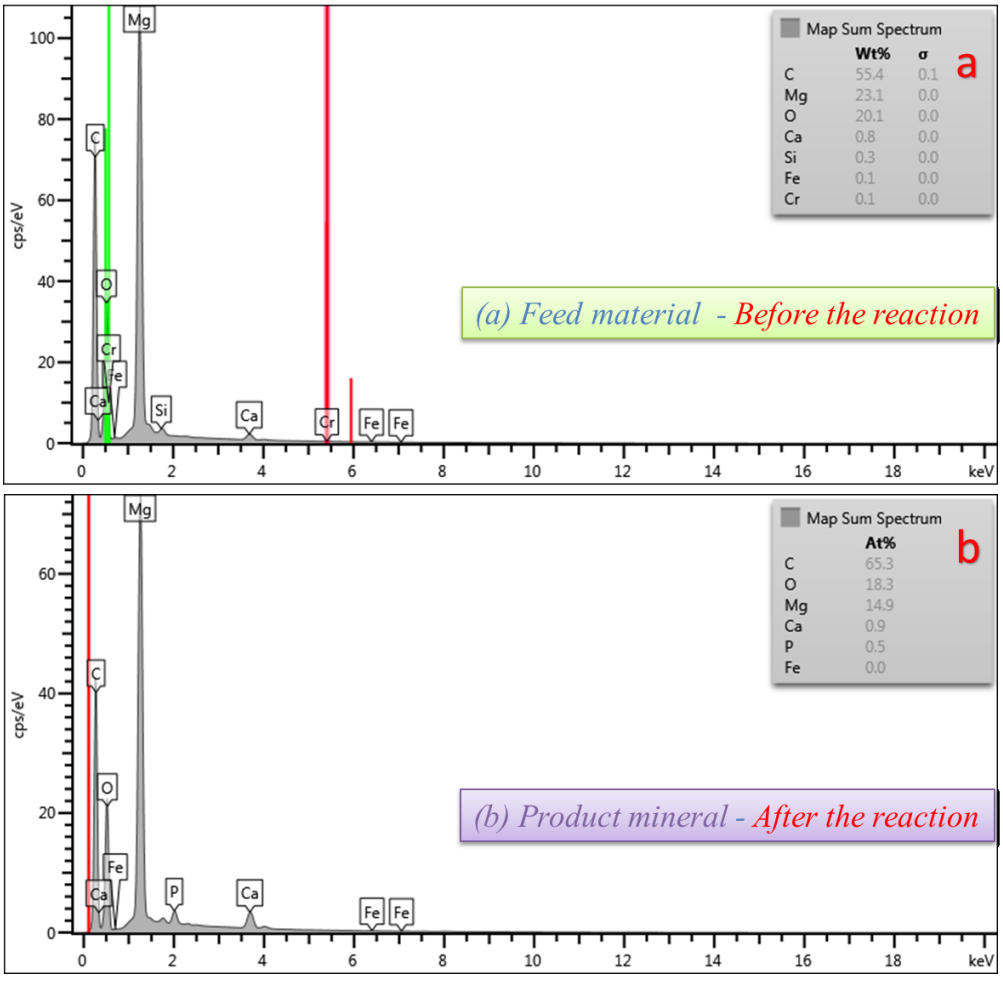

Figure 3: Elemental composition of the: (a) feed and (b) product mineral.

Specifically, the feed material (mechano-thermo activated magnesite) was found to contain Mg, O, and C, in addition to traces of Ca, and Si (Figure 3). This composition will accommodate an increase in the pH of the wastewater-feed mixture. Furthermore, as discussed above, Mg contributes to the removal of NH_4^+ and PO_4^{3-} from wastewater, which is the main target of this work. Finally, the product mineral (Figure 3b) contains Mg, O, and C, as major components, while traces of Ca and P were also observed. This, along with the observed water quality of the treated water (Table B1 Appendix), suggests that P has been removed from wastewater and fixed in the product mineral (sludge).

3.1.4 Elemental composition analysis

Here the FESEM and the corresponding EDS elemental mapping images for the feed and product mineral are given and discussed, as to gain an in-depth understanding and verify or refute the elemental composition obtained by the EDS element analysis. Specifically, EDS element analysis revealed the existence of Mg, O, Ca, and Si, in the feed and Mg, O, K, P, in the product mineral and therefore FESEM-EDS elemental mapping will focus on them. In Figure 4i the FESEM-EDS layered images of the feed material and for Mg, O, Ca, and Si are shown.

The elemental mapping confirmed the existence of Mg, O, Ca, and Si in the feed, which was identified in the elemental composition of feed and product mineral section. These elements contribute to the removal of phosphate and ammonia from wastewater, while both Mg and Ca contribute to an increase of the pH, as discussed in the mineralogical analysis section. Furthermore, the identified elemental composition of product mineral was also confirmed by the FESEM-EDS imagery, as shown in Figure 4ii. Specifically, the elemental mapping of the product mineral revealed that indeed the product mineral contains Mg, O, Ca and P, as shown in Figure 4ii (b), (c), (d), and (e), respectively. This is an indication that Si was released as the pH was increasing and Mg, Ca and P formed struvite, hence achieving a high removal capacity as will be discussed further down.

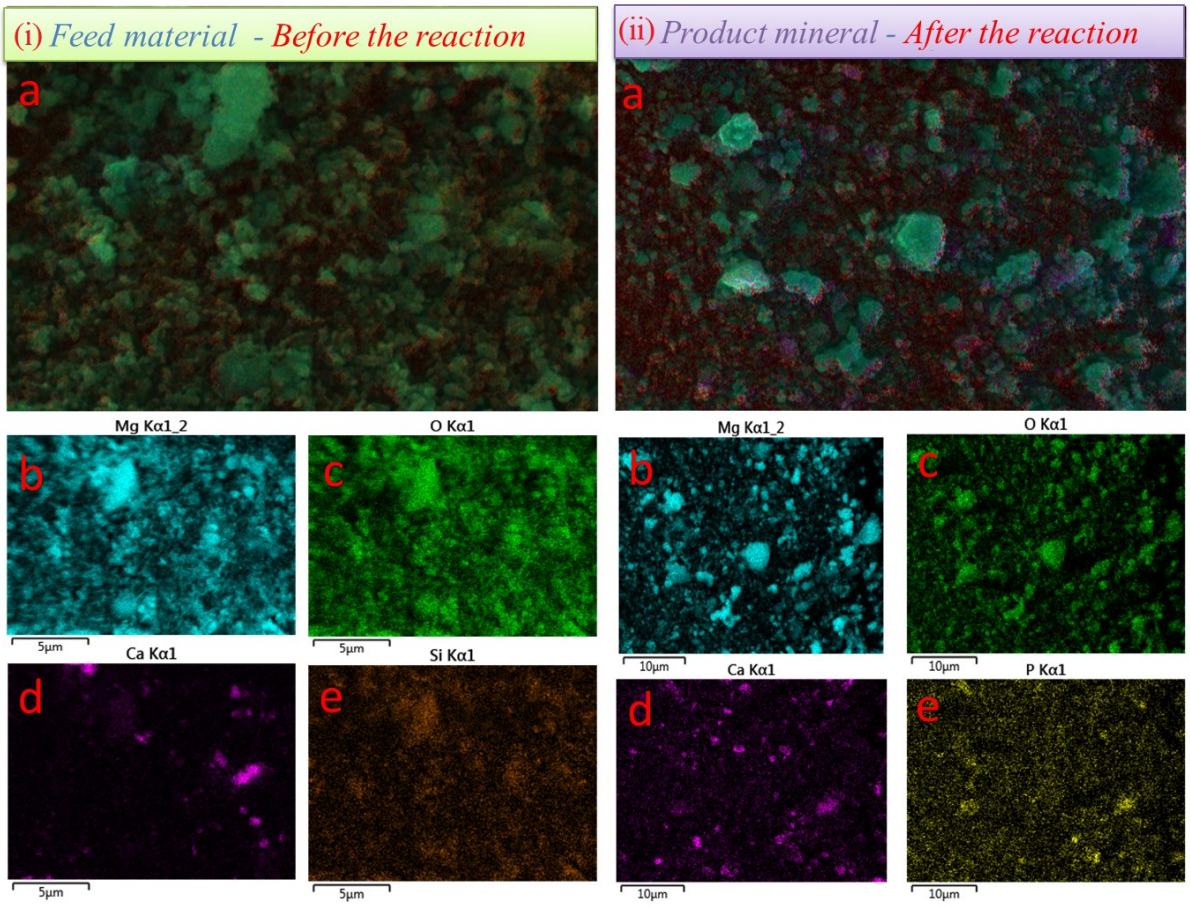

Figure 4: (i) The FESEM image of the: (a) feed material and the FESEM-EDS layered images of the feed material for: (b) Mg, (c) O, (d) Ca, and (e) Si; and (ii) the FESEM image for the: (a) product mineral, and the FESEM-EDS layered images of the feed material for (b) Mg, (c) O, (d) Ca, and (e) P.

3.1.5 Fourier transform infrared spectroscopy analysis

The results of the Fourier transform infrared spectroscopy (FTIR) analysis, both for the feed and the product mineral are shown in Figure 5. Furthermore, the identified functional groups and wavenumber of feed and product mineral are given.

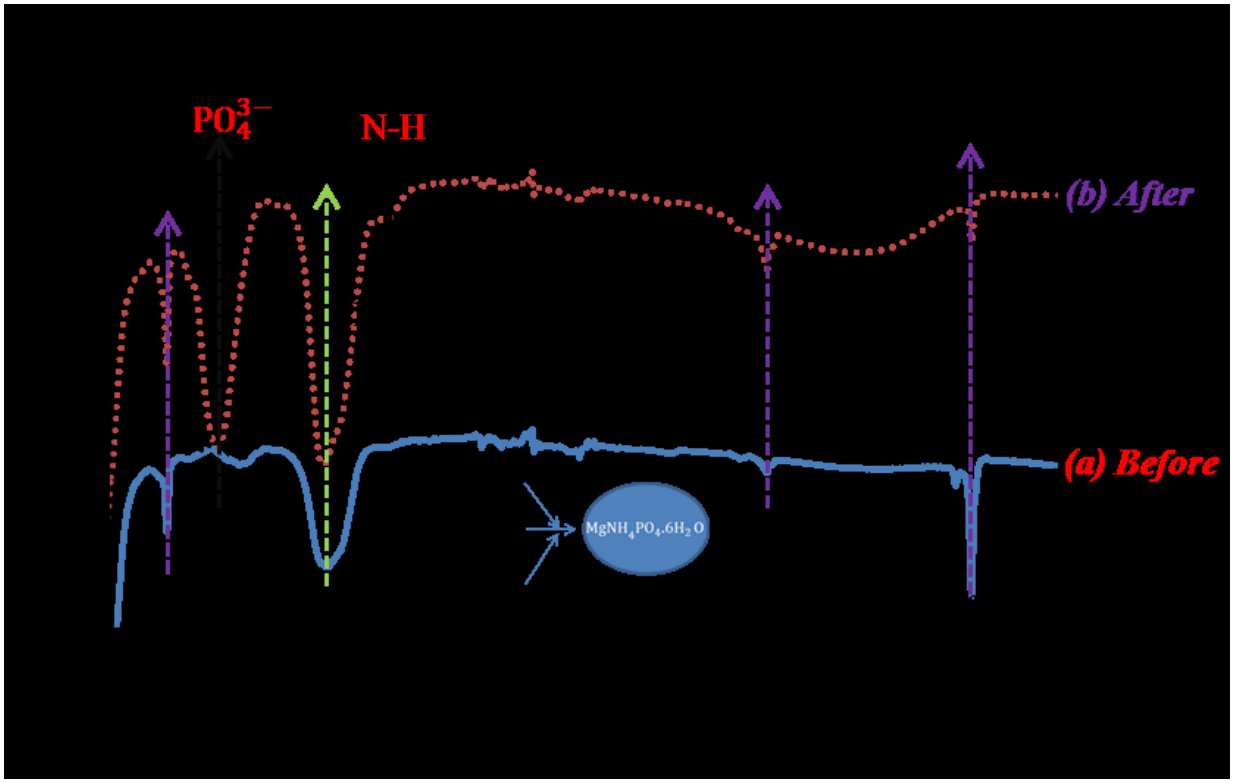

Figure 5: The FTIR results for the feed and product mineral.

The wavenumbers of the peaks of the feed and the product mineral (Figure 5) were compared with those obtained from the literature (Heraldy et al., 2017; Magagane et al., 2019) and were found to be in good agreement. Furthermore, the FTIR spectrum of the feed confirms the existence of water hydration, Mg-O metal-oxygen bond, and carbonate. This corroborates the XRD results and is an indication of the presence of periclase and traces of brucite in the feed. Furthermore, in the FTIR spectrum of the feed, the presence of carbonate was identified, which was expected since it is a residue of the calcination process. Finally, the FTIR spectrum of the product mineral was observed to contain water hydration, Mg-O metal-oxygen bond, N-H bond, and PO_4^{3-}, hence confirming that struvite has been formed. Overall, the obtained results are in agreement with those reported in the literature (Heraldy et al., 2017) and further support the results obtained by the EDS and XRD analyses.

3.2 Operational parameters optimisation

3.2.1 Variation of pollutants removal as a function of contact time

First the effect of contact time, i.e. mixing duration, on total dissolved solids (TDS) removal was examined and then on phosphate and ammonia removal. As mentioned in Table 1a wide spectrum of mixing durations, i.e. 1, 5, 10, 30, 60, 120, 150, 180, 240, and 300 min, was

considered. A sharp decrease in the TDS level/concentration was observed from 0 to 5 min (Figure A1, in the appendix). Thereafter, only trivial fluctuations were observed with increasing contact time. This is a good indication that, in terms of TDS removal, the system reaches equilibrium within the first 5 min of contact time.

Moreover, the percentage removal of the targeted pollutants under study, i.e. phosphate and ammonia, along with the measured level of calcium and magnesium, are shown in Figure A2 in the appendix, and for the abovementioned contact times. A steep increase in the percentage removal of ammonia was observed during the first 10 min (~68 % removal). Then a decrease in the percentage removal is observed up to the contact time 120 min and thereafter the ammonia percentage removal increases up to the 300 min contact time (~76 % removal). Overall, it can be claimed that the optimal contact time is 180 min, where around 75 % of ammonia is removed, since thereafter ammonia is removed commenced to stabilize, hence suggesting that the reaction has reached equilibrium. On the other hand, the system is highly efficient at removing phosphate, since starting at 1 min, the phosphate removal is ~95 % and then maximized at 5 min (~98 %). Thereafter, it remained constant, with more than 98 % of phosphate being removed. This suggests that the system has reach equilibrium in phosphate removal at 5 min contact time. The abovementioned percentage removals are attributed to the reaction of the feed with phosphate and ammonia. A similar trend has been reported by Yagi and Fukushi (2012).

Furthermore, the level of magnesium was observed to follow a somewhat similar pattern with the one observed in ammonia percentage removal, with a step increase in its concentration during the first 30 min, then a decrease at contact time 60 min, then an increase at contact time 90 min and thereafter it remains practically stable (Figure A2 in the appendix). On the other hand, the level of calcium increases at contact time 5 min, then it decreases and thereafter trivial fluctuations are observed.

The above indicate that the reaction of $Mg^{2+}, Ca^{2+}, NH_4^+,$ and PO_4^{3-} (see eq. 2) is leading to the formation of struvite. To afford the reaction adequate time, 5 min was taken as the optimum contact time for the removal of phosphate (more than 98 % is removed) and 180 min was taken as the optimum time for the removal of ammonia (around 75 % of the ammonia content is removed). Furthermore, 5 min was considered as the ideal time for the synthesis of struvite using the mechano-thermo treated magnesite. Thus, this study attained a fast struvite formation time, compared to other studies in literature (Sutiyono et al., 2016).

3.2.2 Variation of pollutants removal as a function of feed dosage

In this section the effect of feed dosage on TDS removal and on the removal of the targeted pollutants was examined by varying both the feed dosage (from 0.1 to 10 g). Three different contact times were examined, i.e. 10, 30, and 60 min, while the pollutants' initial concentration (contact time 0 min) was also measured.

Regarding TDS removal, it was found that its level/concentration is directly affected by the contact time and feed dosage (Figure A3, appendix). It appears that with increasing mixing time and dosage, TDS concetration decreases. Specifically, apart from the value for the 30 min contact time and dosage 1 g per 500 ml (i.e. 2 g L^{-1}) and the value for the 60 min contact time and dosage 5 g, where a small increase in the TDS level is observed compared to the previously measured value (see Figure A3, appendix), the pattern that was observed was that the regardless of the contact time the TDS level decreases with increasing feed dosage, until it is practically near-zero. To be more specific, for contact time 10, 30, and 60 min the TDS is practically near-zero (less than 5 mg L^{-1}) when the dosage is equal or higher than 8 g, 3 g, and 8 g, respectively. Therefore, 30 min and 3 g (i.e. 6 g L^{-1}) were observed to be the most efficient range for the reduction of TDS (Figure A3, appendix).

Furthermore, the effectiveness of the process in removing the targeted pollutants was examined by correlating their percentage removal with varying feed dosages, as described in Table 1. The calcium and magnesium levels were also measured at 30 min contact time and for the examined feed dosages. It was observed that the process is very efficient in phosphate removal (percentage removal> 95 %) even at the lowest dosage, i.e. 0.1 g. For the 0.5 g feed dosage phosphate is practically removed (> 98.3 %) and thereafter remains constant (Figure A4, appendix). This indicates the high affinity of the feed material to phosphate and suggests that the system can operate efficiently even at very low feed dosages. For ammonia removal, a different pattern was observed. The system presents a steep increase in the percentage of ammonia removal from the 0.1 to 0.5 g feed dosage, thence the percentage removal decreases up to 5 g dosage, while a sharp increase in ammonia removal in is observed for higher values. The fluctuations in ammonia removal could be explained by the competition to reactive sites.

At the same time, it is observed that the Ca level remained practically constant up to the 4 g feed dosage and thereafter it sharply increase with increasing feed dosage values. On the other hand, the Mg level is observed to decrease up until the 4 g feed dosage, with a sharp decrease observed between 3 g and 4 g, while for 4 g dosage values and higher the Mg level

is minuscule. This large reduction in the Mg level could be attributed to the increase in pH, traced back to increasing Ca concentration as discussed above, leading to the formation and precipitation of brucite. After the 4 g dosage, the Ca levels were observed to steeply increase with an increase in dosage. This could be attributed to possible dissolution of Ca and Mg from the feed into the wastewater and could be precursors for struvite synthesis.

In light of the obtained results, it is concluded that with 0.5 g (i.e. 1 g L^{-1}) dosage phosphate removal is optimized, whereas for ammonia removal optimisation 8 g (i.e. 16 g L^{-1}) dosage is required. Results suggest the importance of tailoring the process parameters according to the characteristics and treatment quality requirements of the targeted wastewater effluent. Here, 3 g (i.e. 6 g L^{-1}) was considered as the optimum dosage for the removal of both phosphate and ammonia.

3.2.3 Pollutant removal as a function of their initial concentration

First, the effect of phosphate initial concentration on TDS removal was examined by varying its initial concentration from 2 to 123 mg L^{-1}, as described in Table 1. An increase in TDS removal is observed with increasing initial phosphate concentration, while pH was observed to decrease and particularly a steep decrease is observed for phosphate values higher than 33 mg L^{-1} (Figure A5, appendix). The high pH values observed with low phosphate concentration indicate the over-saturation of the system, with alkaline generating materials (MgO and CaO) elevating the pH. As the phosphate concentration increases, more chemical components are present in the system, which then react with the feed's dissolved chemical components, hence a reduction in pH is observed. As a result, for initial phosphate concentrations between 2 to 21 mg L^{-1}, the TDS level of the system (wastewater-feed) was higher than the initial TDS of the wastewater (Figure A5, appendix), however, in inverse proportion, i.e. as the initial phosphate concentration, and by extension the TDS level of the wastewater, was increasing (up to 33 mg L^{-1} phosphate concentration) the TDS level of the system was decreasing. This is an indication that the feed addition is increasing the TDS concentration due to the dissolution of chemical components in its matrices. Nevertheless, for phosphate initial concentration higher than 33 mg L^{-1}, and up to 123 mg L^{-1}, the TDS levels of the system were constantly lower than the initial TDS of the wastewater, in direct proportion, which indicate pollutants removal from the wastewater-feed system.

To gain insight into the process removal efficacy, the effect of both phosphate (concentration range 2 to 123 mg L^{-1}) and ammonia (concentration range 4 to 239 mg L^{-1}) initial concentration was examined, since it has been inferred that changes in the chemical composition of wastewater can affect the simultaneous recovery of both phosphate and ammonia (Muhmood et al., 2019). Results are shown in Figure 6.

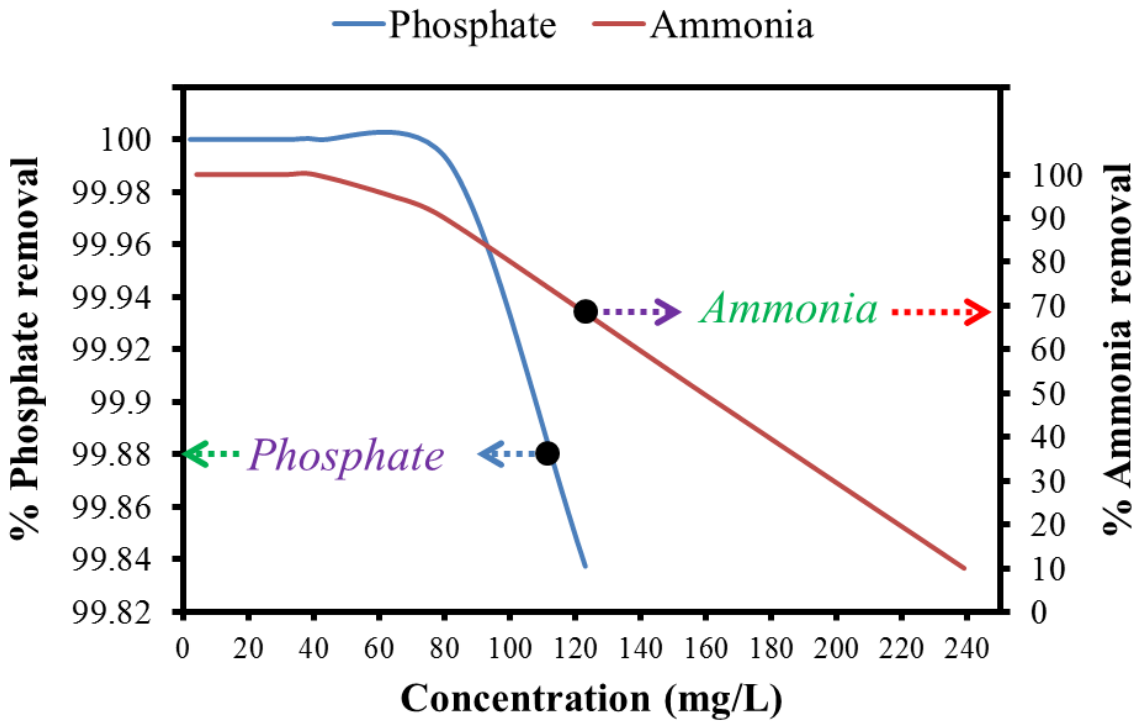

Figure 6: Variation in percentage removal of ammonia and phosphate with varying initial concentration (conditions: room temperature, 8 g of dosage in 500 mL, 30 min of mixing, and 500 rpm mixing speed).

As shown in Figure 6, pollutants removal was observed to decrease with an increase in their initial concentration, particularly in the case of ammonia. This is an indication that the feed's precursors chemical compounds are getting depleted with an increase in pollutant concentration, thus leading to a reduction in their removal efficiency. It should be mentioned that the system is highly efficient in removing phosphate, as shown in the corresponding Y-axis (left-hand-side) of Figure 6 (phosphate removal >99.80% for the concentration spectrum 2 - 123 mg L^{-1} and with 16 g L^{-1} feed dosage). More specifically, when phosphate initial concentration is between 0 to 41 mg L^{-1}, the system performance is optimised, with 100 % phosphate removal. For values higher than 41 mg L^{-1} the system's removal efficiency appears to sharply reduces, nonetheless for practical applications it remains more or less constant

since for 82 mg L^{-1} initial concentration phosphate removal is 99.99 % and for the 123 mg L^{-1} value phosphate removal is 99.84 %.

On the other hand, ammonia removal is optimised for initial ammonia concentration in the range of 8 to 41 mg L^{-1}, beyond that, the removal efficiency was observed to gradually decrease Specifically, for values 8 to 40 mg L^{-1} ammonia removal is totally reduces (100 % removal), then it start to decrease, with a sharp decrease in ammonia removal efficiency being observed for values higher than 80 mg L^{-1}. In that regard, it can be inferred that 16 g L^{-1} of mechano-thermo activated magnesite (feed) is sufficient to remove ≤ 123 and 80 mg L^{-1} of phosphate and ammonia, respectively. Results also suggest that, as with all wastewater treatment methods, the proposed method could operate as part of a treatment train, for the efficient removal of both phosphate and particularly ammonia.

3.2.4 Pollutants removal as a function of pH

The effect of the pH on TDS, phosphate, and ammonia removal was examined by taking into account a wide spectrum of pH values, ranging from as low as 1 to as high as 12 (Table 1). The results for TDS are shown in Figure A6, in the appendix, where the initial and final (after treatment) pH level, along with the TDS level, are shown. It was observed that for pH values 1 and 2 the TDS level is very low, even at contact time 0 min, and it becomes minuscule after 10 min treatment (Figure A6, appendix). Furthermore, for pH values in the range of 3 to 8, TDS decreases with increasing pH, however, for pH values higher than 8 TDS increases with increasing pH values. This could be attributed to Mg precipitation from the aqueous system, hence making the system deficient of one of its main components. Nonetheless, for pH values higher than 10 the system is dominated by Ca ions, thus making them the best candidate for anions removal. Therefore, the treatment system was observed to be dependent on pH. The pH of the final solution was also observed to increase with an increase in initial pH. The most effective conditions were observed to be in the range of pH 8.5 to 10. This was also corroborated by examining the pollutants percentage removal as a function of pH, while also the Mg and Ca concentration/level were measured (Figure A7, appendix).

Specifically, in Figure A7 it is observed that pollutants percentage removal increases with increasing pH in the range of 2 to 10. Thereafter, i.e. for pH values higher than 10, a large reduction in phosphate removal is observed, however, ammonia removal keeps increasing with increasing pH. More specifically, regarding the efficiency of the system in phosphate

removal, this was found to be very efficient since it was observed to be > 99% for pH values in the range of 1 to 10, however, it drops drastically for pH values between 10 to 12. Furthermore, the levels of Mg and Ca were also observed to decrease with an increase in pH, while at pH >10 they appeared to have been depleted and this is directly proportional to the level of phosphate in solution. A rapid decrease in Mg and Ca concentration was also observed for pH values 2 – 3, which is proportional to the levels of TDS in the system. As mentioned above the percentage of ammonia removal keeps increasing with increasing pH, however the observed steep increase in ammonia removal after the value 10 probably cannot be attributed to the mechano-thermo activated magnesite process per se, since pH > 10 promote the loss of ammonia to air due to stripping (Sutiyono et al., 2016). Therefore, pH in the range 8 - 10 was observed to be suitable for the removal of both phosphate and ammonia (Figure A7, appendix). Similar results were reported by Stolzenburg et al. (2015).

3.2.5 Variation of pollutants removal as a function of temperature

The influence of temperature on TDS and pollutants removal efficiency is examined here. As detailed in Table 1, the examined temperature range is from 35 to 75 ^{o}C, while ambient pH was used (7.56 to 7.79). The effect of temperature on TDS removal was first examined and it is observed that TDS removal practically remains constant, regardless of temperature variations (Figure A8, appendix). However, some small variations in the TDS levels were observed, nonetheless these can be mainly attributed to pH changes. This suggested that TDS, and, by extension pollutants removal is practically independent of the temperature. As such, the results suggest that process scale-up could be conducted under South Africa's temperate climate, to make this technology cost- and energy-efficient as well as environmentally friendly.

To examine this assumption, the effect of temperature on pollutants removal was also examined (Figure A9, appendix). It was observed that phosphate removal was independent of temperature, nevertheless, at first glance, it appears that ammonia removal is somewhat dependent on temperature, since a steep increase in ammonia percentage removal is observed from the 35 °C to 45 °C and thereafter the percentage removal slightly improves with increasing temperature (Figure A9, appendix). However, this is most probably attributed to the fact that ammonia evaporates at temperatures higher than 30 °C, as reported by Sutiyono et al. (2016), and not to the mechano-thermo activated magnesite process per se. The Mg levels were observed to be dependent on temperature, as they rapidly decreased with an

453 increase in temperature, hence denoting a rapid reaction with the pollutants. This is an
454 indication that struvite formation is endothermic, because it requires additional energy for
455 optimal removal of struvite. Furthermore, the Ca levels were observed to gradually increase
456 with an increase in temperature, hence offsetting the TDS of the defined system. In light of
457 that, 55 °C appears to be the optimum temperature for pollutants, and particularly for
458 ammonia, removal. However, for industrial deployment applications and to make this
459 technology simpler to apply and energy friendly, and by extension commercially appealing,
460 the system should be scaled up under South Africa's temperate climate, rather than 55 °C
461 temperature. We should also note here that in our future works we plan to assess the
462 environmental footprint of the process, examining also the effect of additional energy and
463 ammonia emissions when the process operates at 55 °C temperature.

464

465 **4 Treatment of municipal effluents at optimised conditions**

466 Finally, the real wastewater effluent was treated under the optimum conditions, identified in
467 sections 3.2.1 – 3.2.5 (30 min and 4 g: 500 mL dosage, ambient temperature and pH). In
468 Table B1 in the appendix, the main characteristics of the raw municipal wastewater used
469 throughout this work, as well as the main characteristics of the treated effluent, under the
470 identified optimumconditions, are reported.

471 It can be inferred from Table B1 (appendix) that the optimised system can effectively remove
472 pollutants from real municipal effluents, by means of a simple and easily accessible
473 mechano-thermo treated cryptocrystalline magnesite technology. Specifically, the main
474 pollutants under study, i.e. phosphate and ammonia, were drastically reduced, particularly
475 phosphate which was practically removed. Furthermore, an increase in the pH, from 7.5 to
476 10.8, indicates that a reaction took place between the feed (mechano-thermo activated
477 magnesite) and the wastewater, leading to struvite formation. EC and TDS were also
478 observed to decrease, which further support that there is a large attenuation of pollutants from
479 the aqueous media. Specifically, sulphate was reduced from 150 to 40 mg L^{-1}; phosphate was
480 practically removed, from 120 to 0 mg L^{-1}; whereas ammonia was reduced from 135 to 30
481 mg L-1, hence suggesting struvite formation (Table B1, appendix). The levels of magnesium
482 and calcium were also observed to have decreased significantly, which implies a possible co-
483 precipitation during the struvite formation.

5 Mechanisms governing pollutants removal

Taking into account the abovementioned results, it appears that the removal of phosphate and ammonia from wastewater is mainly attributed to a combination of adsorption and precipitation. Specifically, adsorption precedes precipitation and then struvite crystallization, (Li et al., 2017) leading to the formation of a number of valuable substances, by the mechanisms shown below. Specifically, equations (5) – (9) depict the route for the formation of the final products, eventually leading to struvite formation:

$$Mg/CaO + H_2O \rightarrow Mg/Ca^{2+} + 2OH^- \quad (5)$$

$$5Ca^{2+} + 3HPO_4^{2-} + OH^- = Ca_5(PO_4)_3OH \quad (6)$$

$$Ca^{2+} + HPO_4^{2-} = CaHPO_4 \quad (7)$$

$$3Ca^{2+} + 2PO_4^{2-} = Ca_3(PO_4)_2 \quad (8)$$

$$Mg^{2+} + H_nPO_4^{n-3} + NH_4^+ + 6H_2O \rightarrow MgNH_4PO_4.6H_2O + nH^+ \quad (9)$$

where n = 0,1,2, etc., and it corresponds with the solution pH (Peng et al., 2018).

Therefore, as is suggested by equations (5) – (9) the presence of Mg and Ca will lead to the effective treatment of the municipal effluents via crystallization of calcium phosphate and struvite synthesis process. Similar inferences were reported by Peng et al. (2018). Finally, it should be noted that this system cannot be purely defined by adsorption models, since the process is not reversible and it is not utterly a surface phenomenon, hence warranting its classification as a precipitation process that leads to a crystallization process.

6 Conclusions

Mechano-thermo treated cryptocrystalline magnesite (feed) was successfully employed for the removal of phosphate and ammonia, among other pollutants, from real municipal wastewater effluents. Operational parameters of the treatment process were assessed and the optimised parameters were observed to vary for phosphate (5 min of mixing, 0.5 g of feed dosage in 500 mL or 1 g L^{-1}, 123 mg L^{-1} initial phosphate concentration, and pH 8 – 10) and for ammonia (180 min, 16 g L^{-1} dosage, 80 mg L^{-1} concentration, pH 10 and > 45°C) removal. The optimal conditions for the removal of both phosphate and ammonia were identified to be 30 min, 3 g L^{-1} dosage, ambient pH and temperature (7.5 pH and 35 ^oC respectively), and 120 and 135 mg L^{-1} initial phosphate and ammonia concentration,

513 respectively. The above suggest that the system can be part of a treatment train and/or be
514 tailored according to the initial wastewater characteristics and the targeted water quality of
515 the treated effluent.

516 Furthermore, XRD confirmed the presence of struvite in the produced sludge, hence
517 indicating that the removed pollutants react with the feed, leading to struvite formation, a
518 valuable process co-product. This was further confirmed by the FT-IR analysis, which
519 revealed the presence of N-H, Mg-O and P groups in the matrices of the product mineral
520 (sludge). Struvite can be used as a replacement of mined phosphate fertilizers and it is also
521 rich in nitrogen and magnesium, suggesting its production can valorise the wastewater
522 treatment process and promote wastewater sustainable management. Furthermore, the water
523 quality analyses revealed that the treated effluent is suitable for irrigation purposes, since
524 most of the pollutants that may trigger algal growth have been effectively removed by the
525 mechano-thermo activated magnesite driven system. Overall, with the proposed treatment
526 process not only phosphate and ammonia are effectively removed, but struvite is also
527 produced, thus leading to the valorisation of wastewater treatment process and opening new
528 horizons for the effective and sustainable management of wastewater effluents.

529 **Acknowledgments**

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535

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628 **Appendix A – Operational parameters optimisation figures**

629

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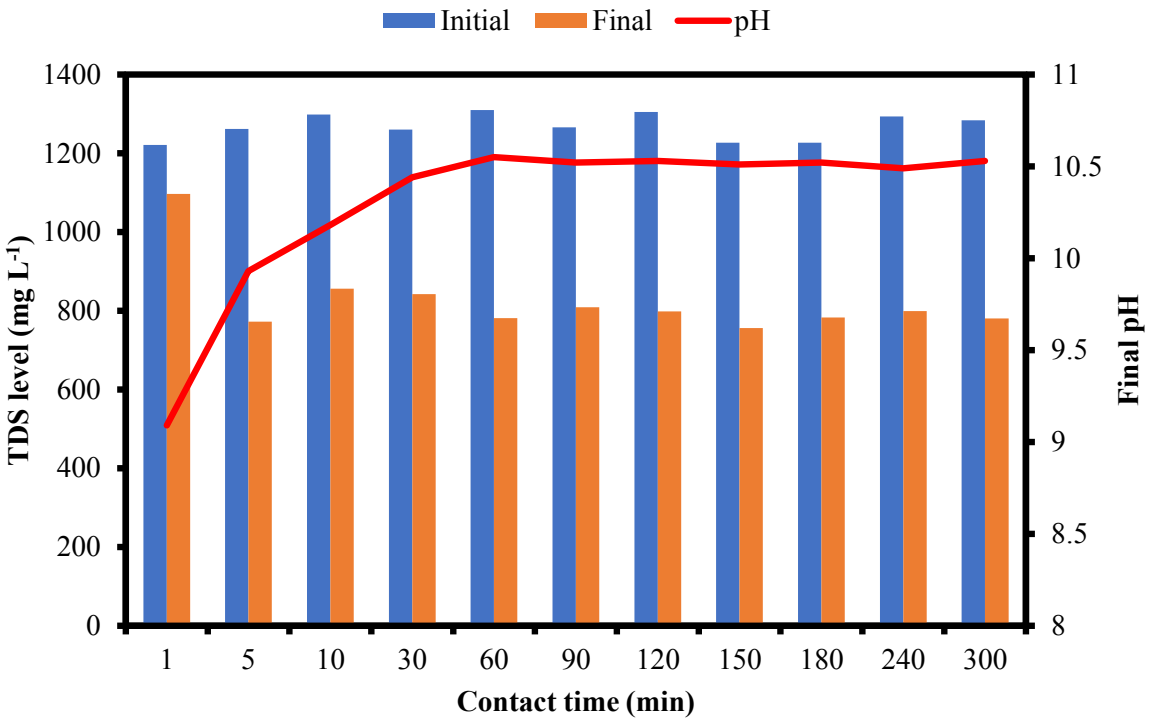

631

632 **Figure A1:** Variation in TDS with an increase in contact time (conditions: 1 g: 100 mL S/L
633 ratios, room temperature, 1 g of pre-treated magnesite and 500 rpm mixing speed).

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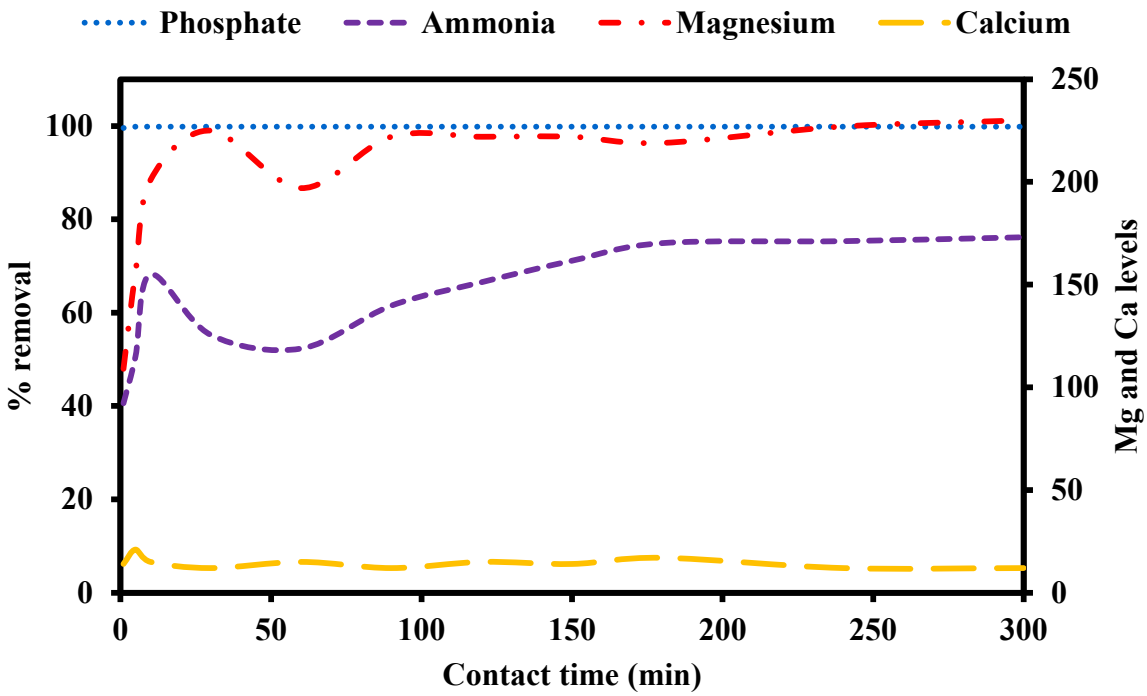

635

636 **Figure A2:** The variation in percentage removal of ammonia and phosphate along with the
637 levels of calcium and magnesium, with varying contact time (conditions: 1 g: 100 mL S/L
638 ratios, room temperature, 1 g of pre-treated magnesite, and 500 rpm mixing speed).

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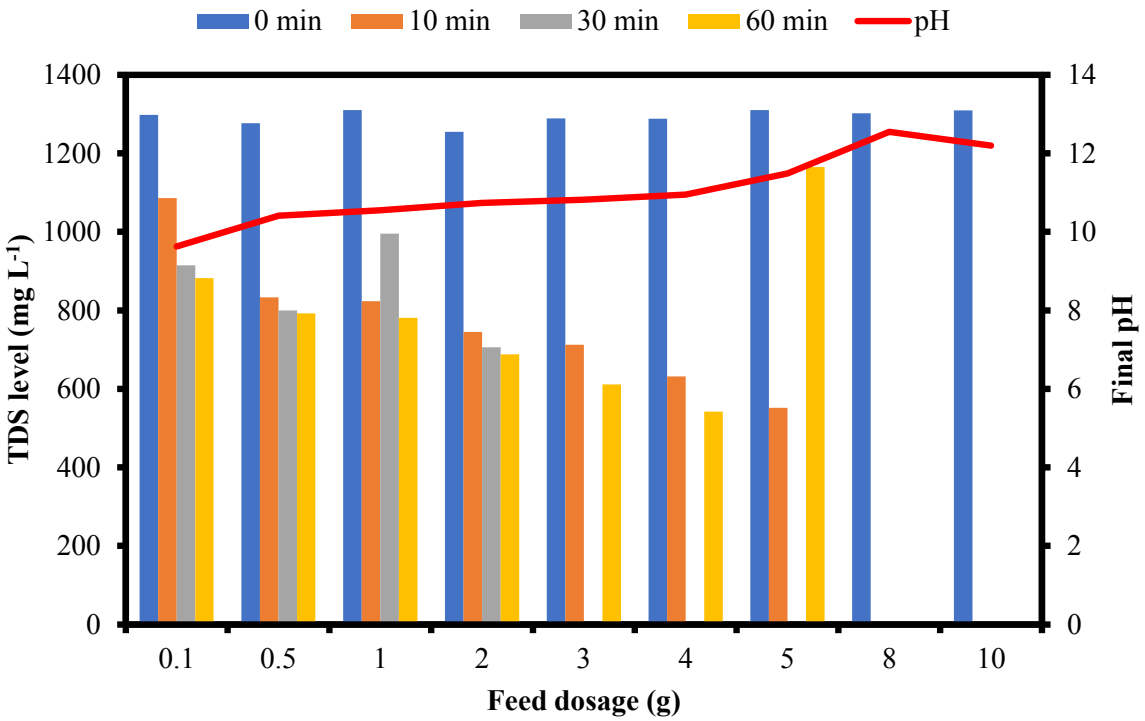

640

641 **Figure A3:** Variation in TDS concentration with increasing feed dosage and contact time
642 (conditions: room temperature, pH > 6.5, 500 rpm mixing speed, time : 10 to 60 min, 123
643 ppm for phosphate, 80 ppm for ammonia, and room temperature)

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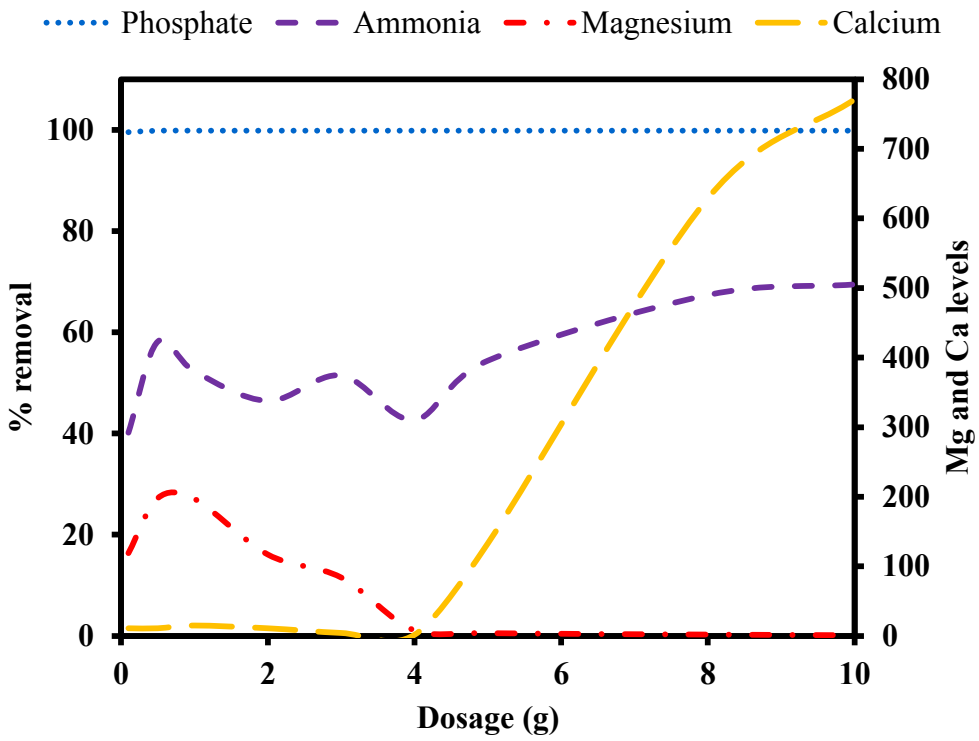

645

646 **Figure A4:** Variation in percentage removal of ammonia and phosphate, along with the
647 levels of calcium and magnesium, with varying feed dosage (conditions: 500 mL volume,
648 room temperature, contact time 30 min and 500 rpm mixing speed).

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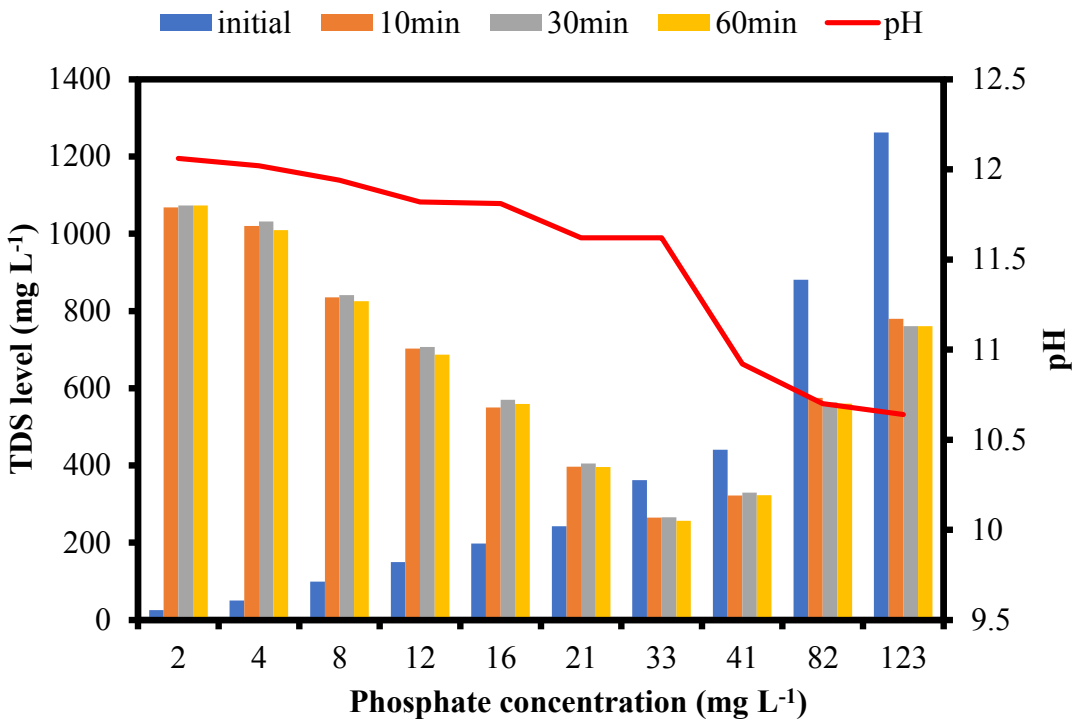

650

651 **Figure A5:** Variation in TDS with a variation in phosphate initial concentration (conditions:
652 room temperature, 30 min of mixing, pH > 6.5, 8 g: 500 mL, and 500 rpm mixing speed).

653

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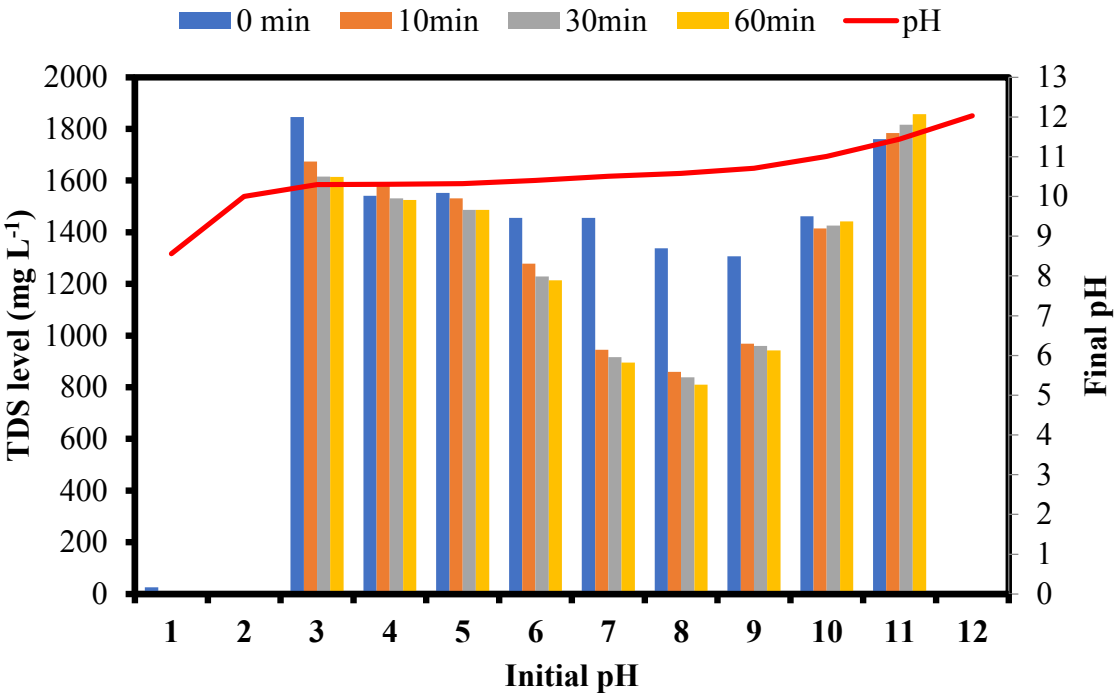
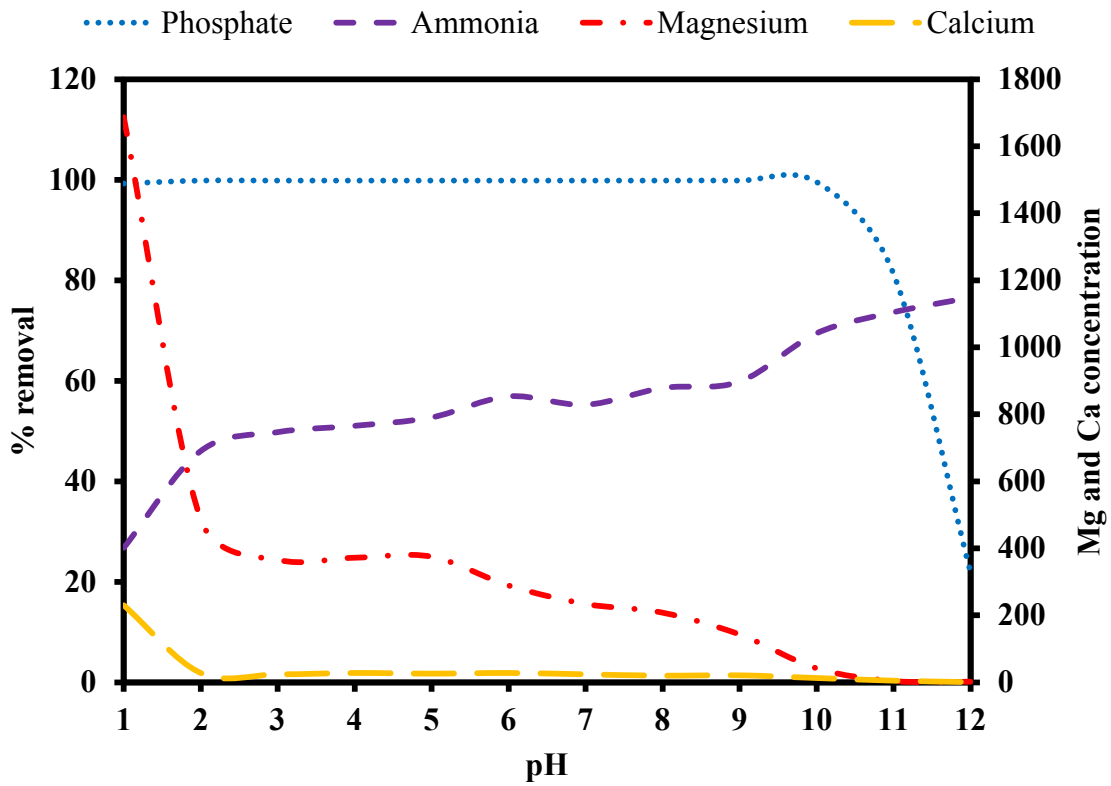


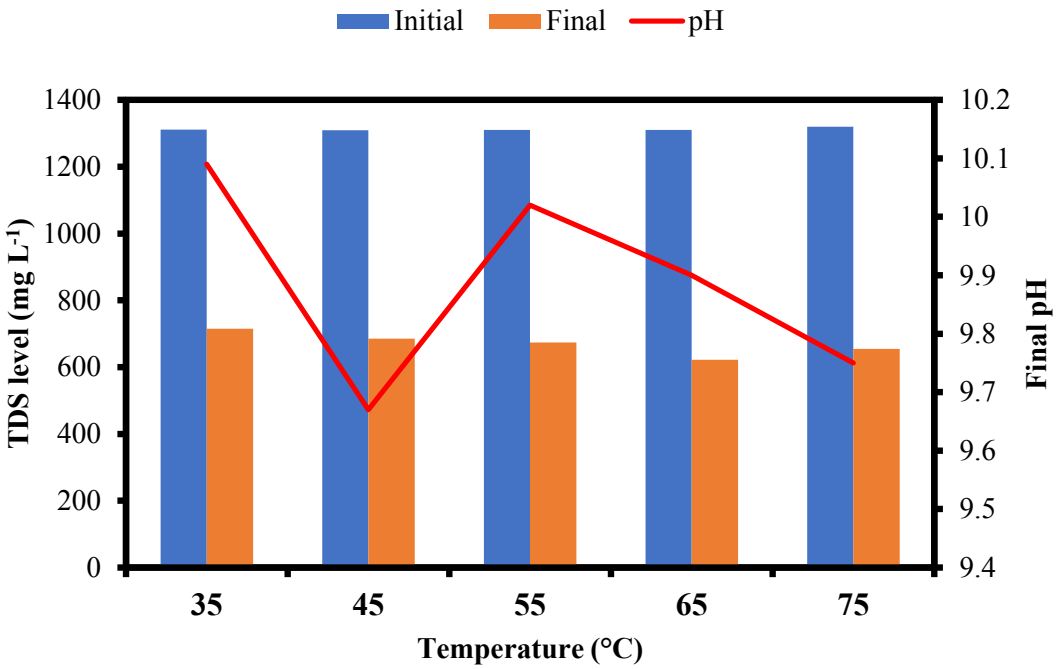
Figure A6: Influence of the initial pH level on TDS removal and on the final pH level (conditions: room temperature, feed dosage 8 g: 500 mL, 500 rpm mixing speed, 123 ppm for phosphate, and 80 ppm for ammonia).



660

661 **Figure A7:** Variation in percentage removal of ammonia and phosphate with varying pH and
662 influence on Mg and Ca concentration (conditions: room temperature, 30 min contact time, 8
663 g : 500 mL of dosage, ≤ 123 and 80 mg/L of phosphate and ammonia, and 500 rpm mixing
664 speed).

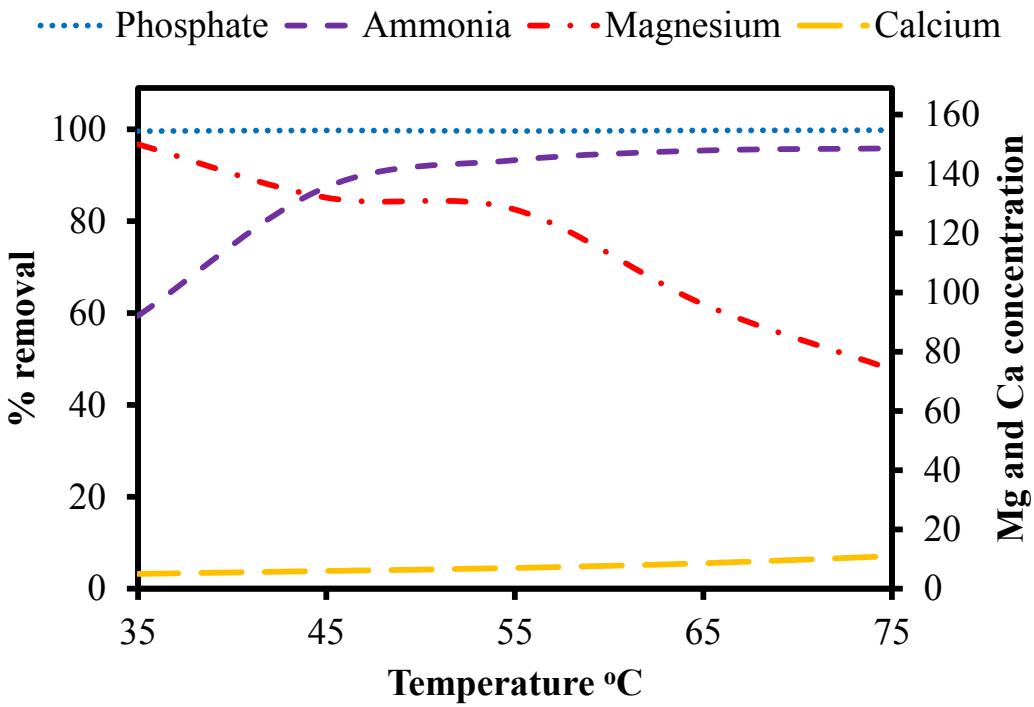
665



666

667 **Figure A8:** Influence of temperature on TDS removal (conditions: 30 min, 8 g: 500 mL, 500
668 rpm mixing speed, 123 ppm for phosphate, and 80 ppm for ammonia, pH between 7.56 to
669 7.79).

670



671

672 **Figure A9:** Variation in percentage removal of phosphate and ammonia under different
673 temperature gradients (conditions: 30 min, room temperature, 4 g: 500 mL of dosage, ≤ 123
674 and 80 mg/L of phosphate and ammonia, and 500 rpm mixing speed).

675

676

677

678 **Appendix B – Raw wastewater and treated effluent characteristics**

679 **Table B1:** Main characteristics of the raw wastewater and the treated effluent under optimal
680 conditions, i.e. 30 min and 4 g: 500 mL dosage.

Parameters	Units	Raw municipal effluents	Treated effluent
Aluminium	mg L^{-1}	<0.00073	<0.00073
Ammoniacal Nitrogen	mg L^{-1}	135	30
Antimony	mg L^{-1}	0.0017	0.0011
Arsenic	mg L^{-1}	0.0035	0.0034
Barium	mg L^{-1}	0.011	0.0058
Boron	mg L^{-1}	<0.16	<0.16
Cadmium	mg L^{-1}	<0.0002	<0.0002
Calcium	mg L^{-1}	200	0.2
Chloride	mg L^{-1}	76	81
Chlorine (free)	mg L^{-1}	0.07	0.04
Chromium	mg L^{-1}	0.00037	0.0013
Colour	mg L^{-1}	44	33
Copper	mg L^{-1}	0.0013	0.0007
Cyanide total	mg L^{-1}	<0.010	<0.010
Electrical conductivity (EC)	mS/cm	200	120
Flouride	mg L^{-1}	<0.2	1.2
Iron	mg L^{-1}	0.015	<0.00088
Lead	mg L^{-1}	<0.00011	<0.00011
Magnesium	mg L^{-1}	1600	0.5
Manganese	mg L^{-1}	0.022	<0.00025
Mercury	mg L^{-1}	<0.00045	<0.00045
Monochloramine	mg L^{-1}	0.03	<0.010
Nickel	mg L^{-1}	0.017	0.013
Nitrate + Nitrite	mg L^{-1}	1.5	1.3
Nitrate Nitrogen	mg L^{-1}	1.4	1.3
Nitrite Nitrogen	mg L^{-1}	<0.2	<0.2
pH	N/A	7.5	10.82
Phenols	mg L^{-1}	<0.01	<0.01
Phosphate	mg L^{-1}	120	<0.001
Selenium	mg L^{-1}	<0.0021	<0.0021
Sodium	mg L^{-1}	77	74
Sulphate	mg L^{-1}	150	40
TDS	mg L^{-1}	2900	756
Turbidity	NTU	10	2.9
Uranium	mg L^{-1}	<0.00008	<0.00008
Zinc	mg L^{-1}	0.014	<0.00057

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