

Crystal chemistry of the mendipite-type system $\text{Pb}_3\text{O}_2\text{Cl}_2\text{—Pb}_3\text{O}_2\text{Br}_2$

Oleg I. Siidra^{*I}, Sergey V. Krivovichev^I, Thomas Armbruster^{II} and Wulf Depmeier^{III}

^I Department of Crystallography, St. Petersburg State University, University Emb. 7/9, 190034 St. Petersburg, Russia

^{II} Laboratorium für chemische und mineralogische Kristallographie, Universität Bern, Freiestraße 3, 3102 Bern, Switzerland

^{III} Institut für Geowissenschaften, Universität zu Kiel, Olshausenstraße 40, 24118 Kiel, Germany

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Abstract. The crystal structures of the mendipite series $\text{Pb}_3\text{O}_2\text{Cl}_2\text{—Pb}_3\text{O}_2\text{Br}_2$ have been refined. The structures are based upon $[\text{O}_2\text{Pb}_3]^{2+}$ double chains of edge-sharing OPb_4 tetrahedra. There are three symmetrically independent Pb^{2+} cations. The number of nonequivalent halogen sites is two (X1, X2). Short Pb—O bonds are located on one side of the Pb^{2+} cations and weak Pb—X bonds are located on the other side of the Pb^{2+} coordination sphere. The evident strong distortion of the Pb^{2+} coordination polyhedra is due to the stereoactivity of the $6s^2$ lone electron pairs of the Pb^{2+} cations. Pb1—X2 and Pb2—X2 bonds are the most sensitive to the X site occupancy, which is in agreement with the non-linear behavior of the *a* and *c* parameters. Determination of unit-cell parameters by single crystal studies showed strong deviation from Vegard's rule. Nonlinearity of the lattice parameters is caused by selective ordering of the halide anions over X1 and X2 sites. Br atoms prefer the X2 position, whereas Cl prefers the X1 site. The angle between two adjacent OPb_4 tetrahedra was determined to analyze the influence of halogen atoms on the structure of the $[\text{O}_2\text{Pb}_3]^{2+}$ chain. Different occupancy of the X1 site by Cl and Br atoms leads to most pronounced angular changes. These observations may be interpreted as adaptation of the $[\text{O}_2\text{Pb}_3]^{2+}$ double chains to the large halide ions in the crystal structures of the mendipite series compounds.

Introduction

Mendipite $\text{Pb}_3\text{O}_2\text{Cl}_2$ is the most abundant mineral among the natural lead oxyhalides. It was first described by Spencer and Mountain (1923) from the Mendip Hills, England. The crystal structure was solved by Gabrielson (1957) and refined in $P2_12_12_1$ space group. Later, Vincent and Perault (1971) refined the crystal structure of synthetic $\text{Pb}_3\text{O}_2\text{Cl}_2$ in *Pnma*. In the same space group, the structure of natural mendipite was refined by Pasero and Vacchiano

(2000). The most recent single-crystal studies of synthetic $\text{Pb}_3\text{O}_2\text{Cl}_2$ were published by Krivovichev and Burns (2001). The crystal structure of $\text{Pb}_3\text{O}_2\text{Br}_2$ was determined using powder samples by Berdonosov *et al.* (1996) and later by Noren *et al.* (2002).

The environmental importance of lead oxyhalides was pointed out by many authors. Pb oxychlorides were detected in dust particles emitted from a lead smelter (Sobanska *et al.*, 1999; Wu, Biswas, 2000). Pb halides (chloride-bromides) as well as oxy- and hydroxyhalides were observed in automobile exhaust gases (Post, Buseck, 1985) and roadside soils (Smith, 1976). However, Pb oxyhalides have not only environmental importance and are also of interest from the viewpoint of material science as anisotropic materials with a variety of physical properties that depend strongly on the crystallographic direction. Sigman and Korgel (2005) recently described the synthesis and properties of highly birefringent nanostructures with mendipite composition. To date, detailed chemical and structural information is available for pure oxy- and hydroxychloride (Krivovichev, Burns, 2001a; 2002; 2006; Siidra *et al.*, 2007 a, b, c), oxy- and hydroxybromide (Krivovichev, Burns, 2001b; Siidra *et al.*, 2007d; Keller, 1983; Riebe, Keller, 1989) and oxyiodide (Welch *et al.*, 2001) systems, whereas little is known about mixed halide systems such as Cl—Br (Krivovichev *et al.*, 2006).

The practical importance of mendipite-related phases prompted us to take a closer look at the crystal chemistry of the $\text{Pb}_3\text{O}_2\text{Cl}_2\text{—Pb}_3\text{O}_2\text{Br}_2$ system. Single-crystal studies of these phases also provide important information pertinent to an understanding of the transport of lead and the crystallization of lead compounds in natural systems.

Experimental

Synthesis

Single crystals of mendipite phases were obtained by the solid-state reactions method. PbO (Merck, 99%), PbCl_2 (Aldrich, 99.9%) and PbBr_2 (Aldrich, 99.9%) were used as received. Reactants were mixed in an agate mortar in quantities according to the given stoichiometry. Syntheses were carried out with increment of 10% in the Cl:Br ratio

* Correspondence author (e-mail: siidra@mail.ru)

Table 1. Crystallographic data and chemical composition of mendipite-related compounds.

Sample ^a	Cl:Br ^b	Formula ^c	Formula ^d	<i>a</i> , Å	<i>b</i> , Å	<i>c</i> , Å	<i>V</i> , Å ³
M1	0:1	$\text{Pb}_{3.04}\text{O}_2\text{Br}_{1.92}$	$\text{Pb}_3\text{O}_2\text{Br}_2$	12.244(5)	5.872(2)	9.799(4)	704.6(5)
M2	1:9	$\text{Pb}_{2.98}\text{O}_2(\text{Cl}_{0.15}\text{Br}_{1.86})_{2.01}$	$\text{Pb}_3\text{O}_2\text{Cl}_{0.19}\text{Br}_{1.81}$	12.1949(7)	5.8705(5)	9.7968(9)	701.4(7)
M3	2:8	$\text{Pb}_{2.98}\text{O}_2(\text{Cl}_{0.38}\text{Br}_{1.63})_{2.01}$	$\text{Pb}_3\text{O}_2\text{Cl}_{0.46}\text{Br}_{1.54}$	12.100(9)	5.855(5)	9.755(2)	691.1(8)
M4	3:7	$\text{Pb}_{3.02}\text{O}_2(\text{Cl}_{0.56}\text{Br}_{1.42})_{1.98}$	$\text{Pb}_3\text{O}_2\text{Cl}_{0.52}\text{Br}_{1.48}$	12.0518(11)	5.8556(5)	9.7526(9)	688.25(11)
M5	4:6	$\text{Pb}_{3.03}\text{O}_2(\text{Cl}_{0.79}\text{Br}_{1.18})_{1.97}$	$\text{Pb}_3\text{O}_2\text{Cl}_{0.81}\text{Br}_{1.19}$	11.9818(19)	5.8485(9)	9.7273(15)	681.65(18)
M6	5:5	$\text{Pb}_{3.02}\text{O}_2(\text{Cl}_{0.94}\text{Br}_{1.04})_{1.98}$	$\text{Pb}_3\text{O}_2\text{Cl}_{1.03}\text{Br}_{0.97}$	11.922(5)	5.835(2)	9.701(4)	674.8(4)
M7	6:4	$\text{Pb}_{3.01}\text{O}_2(\text{Cl}_{1.17}\text{Br}_{0.82})_{1.99}$	$\text{Pb}_3\text{O}_2\text{Cl}_{1.09}\text{Br}_{0.91}$	11.917(9)	5.819(5)	9.663(8)	670.1(9)
M8	7:3	$\text{Pb}_{2.98}\text{O}_2(\text{Cl}_{1.39}\text{Br}_{0.63})_{2.02}$	$\text{Pb}_3\text{O}_2\text{Cl}_{1.41}\text{Br}_{0.59}$	11.8957(28)	5.8244(14)	9.6441(23)	668.2(3)
M9	8:2	$\text{Pb}_{3.01}\text{O}_2(\text{Cl}_{1.61}\text{Br}_{0.41})_{2.02}$	$\text{Pb}_3\text{O}_2\text{Cl}_{1.61}\text{Br}_{0.39}$	11.9077(17)	5.8264(8)	9.6117(13)	666.85(16)
M10	9:1	$\text{Pb}_{3.02}\text{O}_2(\text{Cl}_{1.76}\text{Br}_{0.22})_{1.98}$	$\text{Pb}_3\text{O}_2\text{Cl}_{1.84}\text{Br}_{0.16}$	11.8928(17)	5.8163(9)	9.5653(14)	661.65(17)
M11	1:0	$\text{Pb}_{3.02}\text{O}_2\text{Cl}_{1.98}$	$\text{Pb}_3\text{O}_2\text{Cl}_2$	11.808(8)	5.7790(41)	9.4784(68)	646.8(8)

a: M = mendipite; b: Cl:Br ratio used in the synthesis; c: obtained by electron microprobe analysis; d: obtained by single-crystal X-ray analysis

(Table 1). Syntheses marked in this table as M1 and M11 are bromine- and chlorine end-member of the $\text{Pb}_3\text{O}_2\text{Cl}_2\text{--Pb}_3\text{O}_2\text{Br}_2$ series, respectively. The produced stoichiometric mixtures were loaded into platinum crucibles and heated in a furnace Carbolite 1200. The following syntheses conditions were found to be optimal to grow single crystals suitable for single crystal X-ray diffraction studies: 1) keeping the mixture at 715 °C for 1 h in air; 2) cooling to 685 °C with a cooling rate of 1 °C min⁻¹ and keeping at this temperature for 15 min; 3) cooling down to room temperature with a cooling rate of 30 °C h⁻¹. The products of the syntheses consisted of yellowish elongated transparent crystals up to 10 mm in size. Powder diffraction patterns were obtained with a Siemens XP18 2 diffractometer for each synthesis to confirm the absence of any by-products. Quantitative electron microprobe analysis provided the chemical formulas for all synthesis products (Table 1). Camscan-4DV electron-scan microscope and AN-10000 semiconductor spectrometer were used at 20 kV and 0.7 nA. $\text{PbL}\alpha$, $\text{BrL}\alpha$, $\text{ClK}\alpha$ were used as analytic lines. Spectrum of the $\text{PbM}\alpha$ line was subtracted to get the correct chemical formula because of its overlapping with the $\text{ClK}\alpha$ line. PbCl_2 and PbBr_2 were used as standards. All calculations were made using AF4/FLS software.

X-ray data collection

Suitable crystals of the mendipite-related compounds were mounted on a Bruker 1 K three-circle CCD based X-ray diffractometer operated at 50 kV and 40 mA. More than a hemisphere of three-dimensional data was collected for each crystal using monochromatic $\text{MoK}\alpha$ X-radiation, with frame widths of 0.3° in 2θ , and with 20 seconds spent counting for each frame. The unit-cell parameters were refined using least-squares techniques. The intensity data were integrated and corrected for Lorentz, polarization, and background effects using the Bruker program SAINT. All the crystals were modeled as ellipsoids for semi-empirical absorption-corrections.

Structure solution and refinement

SHELX-97 programs within the WINGX package were used for the determination and refinement of the structures of all compounds (Table 2). The refinements were made on the basis of parameters taken from Krivovichev and Burns (2001a). The final models included atomic positional parameters and anisotropic-displacement parameters for all atoms, and a weighting scheme of the structure

Table 2. Crystallographic data and refinement parameters for mendipite-related compounds.

Sample	Crystal size (mm)	<i>F</i> (000)	Total reflections	Unique $ F_0 \geq 4\sigma_F$	$2T_{\text{max}}$	<i>GooF</i>	<i>D</i> _{calc} (g/cm ³)	<i>R</i> ₁	<i>wR</i> ₂
M1	0.10 × 0.03 × 0.05	1328	3848	596	27.98	0.931	7.668	0.056	0.093
M2	0.08 × 0.01 × 0.01	1315	3735	770	27.81	1.069	7.624	0.065	0.069
M3	0.07 × 0.01 × 0.02	1295	3061	565	26.98	1.091	7.622	0.081	0.121
M4	0.09 × 0.01 × 0.01	1284	3720	754	28.07	1.039	7.585	0.025	0.031
M5	0.08 × 0.04 × 0.02	1270	3561	746	27.68	1.100	7.576	0.047	0.051
M6	0.07 × 0.02 × 0.01	1253	3638	622	27.95	0.972	7.553	0.083	0.104
M7	0.09 × 0.01 × 0.01	1250	3428	639	27.78	1.028	7.584	0.053	0.072
M8	0.08 × 0.03 × 0.01	1227	3566	722	27.89	1.068	7.467	0.038	0.044
M9	0.08 × 0.02 × 0.01	1212	3608	739	28.06	1.105	7.388	0.051	0.055
M10	0.09 × 0.05 × 0.04	1198	3532	678	27.77	1.054	7.361	0.053	0.061
M11	0.06 × 0.01 × 0.01	1184	3551	583	28.00	0.926	7.440	0.040	0.069

Atom	Sample	x	y	z	$U_{eq}, \text{\AA}^2$	Occupancy
Pb1	M1	0.42101(11)	0.2500	0.42105(11)	0.0161(4)	
	M2	0.42100(7)	0.2500	0.42054(10)	0.0157(4)	
	M3	0.42058(15)	0.2500	0.41974(16)	0.0156(6)	
	M4	0.42055(3)	0.2500	0.41893(4)	0.01764(15)	
	M5	0.42032(5)	0.2500	0.41794(7)	0.0172(3)	
	M6	0.42017(14)	0.2500	0.41750(16)	0.0150(6)	
	M7	0.41996(9)	0.2500	0.41742(10)	0.0166(4)	
	M8	0.41965(5)	0.2500	0.41734(6)	0.0192(2)	
	M9	0.41928(5)	0.2500	0.41756(8)	0.0162(3)	
	M10	0.41890(8)	0.2500	0.41778(9)	0.0157(4)	
	M11	0.41861(7)	0.2500	0.41795(10)	0.0149(3)	
Pb2	M1	0.71266(12)	0.2500	0.36806(12)	0.0196(4)	
	M2	0.71352(8)	0.2500	0.36924(12)	0.0190(4)	
	M3	0.71497(15)	0.2500	0.37037(19)	0.0189(7)	
	M4	0.71677(4)	0.2500	0.37223(5)	0.02139(16)	
	M5	0.71849(5)	0.2500	0.37371(8)	0.0209(3)	
	M6	0.71929(14)	0.2500	0.37437(19)	0.0189(6)	
	M7	0.71973(10)	0.2500	0.37392(11)	0.0206(4)	
	M8	0.71953(5)	0.2500	0.37297(7)	0.0231(2)	
	M9	0.71917(6)	0.2500	0.37150(9)	0.0202(3)	
	M10	0.71874(8)	0.2500	0.37002(10)	0.0194(4)	
	M11	0.71822(7)	0.2500	0.36829(10)	0.0185(3)	
Pb3	M1	0.44199(12)	0.2500	0.78803(11)	0.0188(4)	
	M2	0.44186(8)	0.2500	0.78784(10)	0.0183(4)	
	M3	0.44065(18)	0.2500	0.78772(17)	0.0193(7)	
	M4	0.44006(4)	0.2500	0.78750(4)	0.02066(16)	
	M5	0.43939(5)	0.2500	0.78762(7)	0.0197(3)	
	M6	0.43879(15)	0.2500	0.78780(16)	0.0182(6)	
	M7	0.43882(10)	0.2500	0.78874(11)	0.0192(4)	
	M8	0.43894(5)	0.2500	0.79015(6)	0.0216(2)	
	M9	0.43918(6)	0.2500	0.79184(8)	0.0189(3)	
	M10	0.43963(8)	0.2500	0.79388(9)	0.0184(4)	
	M11	0.44019(8)	0.2500	0.79587(10)	0.0173(3)	
X1	M1	0.6894(3)	0.2500	0.6960(3)	0.0236(8)	Br ₁
	M2	0.6890(2)	0.2500	0.6957(3)	0.0210(10)	Br _{0.815(18)} Cl _{0.185(18)}
	M3	0.6882(5)	0.2500	0.6958(6)	0.014(2)	Br _{0.54(4)} Cl _{0.46(4)}
	M4	0.68820(15)	0.2500	0.69496(17)	0.0223(6)	Br _{0.450(9)} Cl _{0.550(9)}
	M5	0.6879(2)	0.2500	0.6949(4)	0.0232(11)	Br _{0.289(14)} Cl _{0.711(14)}
	M6	0.6886(8)	0.2500	0.6924(10)	0.024(3)	Br _{0.15(3)} Cl _{0.85(3)}
	M7	0.6873(6)	0.2500	0.6936(7)	0.027(2)	Br _{0.14(2)} Cl _{0.86(2)}
	M8	0.6890(3)	0.2500	0.6930(4)	0.0259(13)	Br _{0.043(12)} Cl _{0.957(12)}
	M9	0.6892(3)	0.2500	0.6931(6)	0.0219(16)	Br _{0.020(16)} Cl _{0.980(16)}
	M10	0.6895(5)	0.2500	0.6934(7)	0.0214(12)	Cl ₁
	M11	0.6897(5)	0.2500	0.6940(7)	0.0219(13)	Cl ₁
X2	M1	0.6291(4)	0.2500	0.0752(3)	0.0295(10)	Br ₁
	M2	0.6301(3)	0.2500	0.0750(3)	0.0284(7)	Br ₁
	M3	0.6293(6)	0.2500	0.0747(5)	0.0318(15)	Br ₁
	M4	0.63057(12)	0.2500	0.07453(13)	0.0292(5)	Br _{0.934(9)} Cl _{0.066(9)}
	M5	0.63077(18)	0.2500	0.0735(2)	0.0289(8)	Br _{0.907(13)} Cl _{0.093(13)}
	M6	0.6302(5)	0.2500	0.0741(6)	0.0259(19)	Br _{0.81(3)} Cl _{0.19(3)}
	M7	0.6314(4)	0.2500	0.0733(4)	0.0298(14)	Br _{0.77(2)} Cl _{0.23(2)}
	M8	0.6320(2)	0.2500	0.0745(2)	0.0308(9)	Br _{0.553(13)} Cl _{0.447(13)}

Table 3. Atomic coordinates and displacement parameters for mendipite-related compounds.

Table 3. Continued.

Atom	Sample	x	y	z	$U_{\text{eq}}, \text{\AA}^2$	Occupancy
	M9	0.6328(3)	0.2500	0.0747(4)	0.0287(12)	$\text{Br}_{0.369(16)}\text{Cl}_{0.631(16)}$
	M10	0.6337(5)	0.2500	0.0774(6)	0.0312(19)	$\text{Br}_{0.186(19)}\text{Cl}_{0.814(19)}$
	M11	0.6357(6)	0.2500	0.0811(7)	0.0295(15)	Cl_1
O	M1	0.5778(15)	-0.006(2)	0.3908(12)	0.021(4)	
	M2	0.5800(9)	-0.0053(19)	0.3916(11)	0.013(2)	
	M3	0.5830(19)	-0.002(5)	0.3916(19)	0.015(5)	
	M4	0.5815(4)	-0.0049(9)	0.3916(5)	0.0184(12)	
	M5	0.5828(6)	-0.0079(16)	0.3921(8)	0.0190(19)	
	M6	0.5840(15)	-0.005(4)	0.3956(17)	0.014(5)	
	M7	0.5830(11)	-0.002(2)	0.3906(12)	0.020(3)	
	M8	0.5823(6)	-0.0047(15)	0.3907(7)	0.0218(18)	
	M9	0.5814(6)	-0.0039(16)	0.3903(9)	0.021(2)	
	M10	0.5810(9)	-0.005(2)	0.3900(10)	0.016(3)	
	M11	0.5793(9)	-0.0052(15)	0.3873(11)	0.014(3)	

factors. Each was refined on the basis of F^2 for all unique reflections. The final atomic coordinates and anisotropic displacement parameters are given in Table 3, and selected interatomic distances are in Table 4.

Results

The crystal structure (Fig. 1a) of mendipite-type phases is based upon $[\text{O}_2\text{Pb}_3]^{2+}$ double chains of edge-sharing OPb_4 tetrahedra. The $[\text{O}_2\text{Pb}_3]^{2+}$ double chains extend along the b axis (Fig. 1b). Such chains have previously been ob-

served in the crystal structures of many natural and synthetic lead oxocentered compounds (Krivovichev *et al.*, 2004; Siidra *et al.*, 2007e). In crystal structures of the studied compounds, the chains occur in two mutually perpendicular orientations with the angle varying from $82.48(12)^\circ$ (M11) to $89.36(15)^\circ$ (M6). The halogen ions connect the chains through weak Pb-X ($\text{X}=\text{Cl}, \text{Br}$) bonds only.

There are three symmetrically independent Pb^{2+} cations in the structures of the mendipite series $\text{Pb}_3\text{O}_2\text{Cl}_2\text{--Pb}_3\text{O}_2\text{Br}_2$. The number of nonequivalent halogen sites is two ($\text{X1}, \text{X2}$). Pb1 is coordinated by four O atoms, two X1 and one X2 halogen atoms. Pb2 is coordinated by two

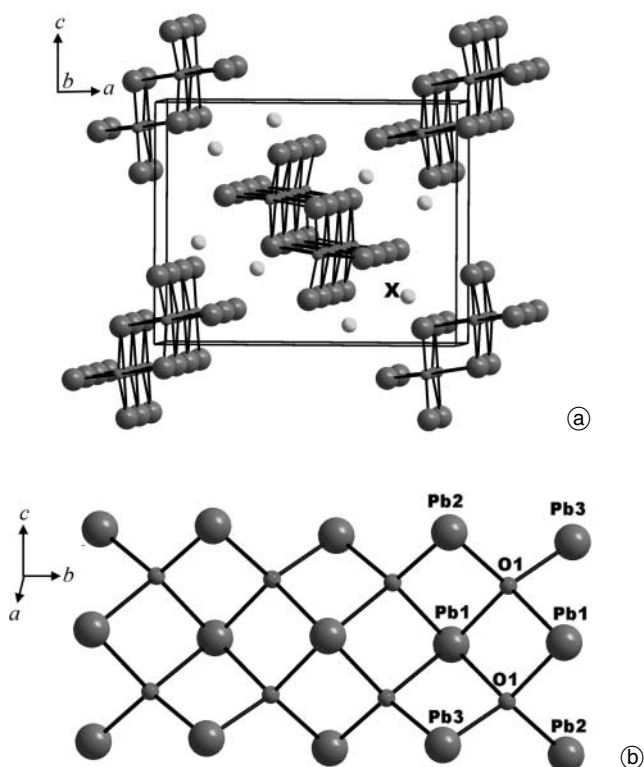


Fig. 1. Projection of the crystal structure of the mendipite-related compounds. Only the Pb–O bonds are shown. Pb – large dark circles, X ($\text{X}=\text{Cl}, \text{Br}$) – light circles, O – small grey circles (a). $[\text{O}_2\text{Pb}_3]^{2+}$ double chain of OPb_4 oxocentered tetrahedra shown in ball-and-stick representation (b).

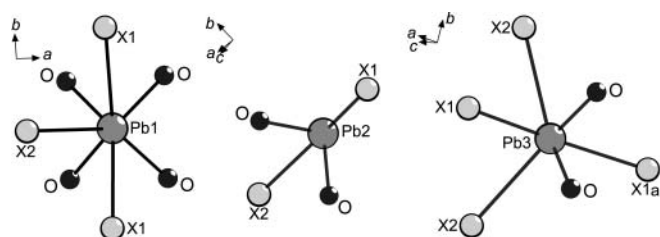


Fig. 2. Coordination of Pb atoms in the structures of the mendipite series $\text{Pb}_3\text{O}_2\text{Cl}_2\text{--Pb}_3\text{O}_2\text{Br}_2$.

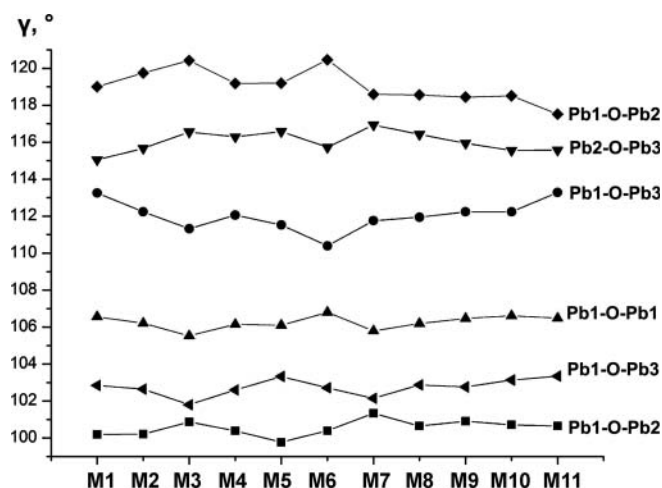


Fig. 3. Variation of angles in the OPb_4 tetrahedra in the crystal structures of the mendipite series $\text{Pb}_3\text{O}_2\text{Cl}_2\text{--Pb}_3\text{O}_2\text{Br}_2$.

Table 4. Selected interatomic distances (Å) for mendipite-related compounds.

Sample	Pb1–O1 2x	Pb1–O1 2x	Pb1–X1 2x	Pb1–X2	Pb2–X1	Pb2–O1 2x
M1	2.335(13)	2.455(17)	3.430(2)	3.574(5)	3.226(4)	2.244(16)
M2	2.335(11)	2.467(11)	3.4224(16)	3.548(3)	3.212(3)	2.224(11)
M3	2.35(2)	2.47(2)	3.402(4)	3.525(7)	3.191(6)	2.18(3)
M4	2.340(5)	2.462(5)	3.3946(9)	3.4953(15)	3.1662(17)	2.219(5)
M5	2.328(9)	2.476(8)	3.3819(16)	3.470(2)	3.146(4)	2.225(8)
M6	2.309(19)	2.46(2)	3.366(5)	3.458(6)	3.107(10)	2.20(2)
M7	2.350(12)	2.449(14)	3.354(4)	3.440(5)	3.113(7)	2.199(14)
M8	2.338(7)	2.451(8)	3.3591(19)	3.422(3)	3.107(4)	2.213(8)
M9	2.338(9)	2.446(8)	3.360(2)	3.412(4)	3.112(5)	2.217(8)
M10	2.327(10)	2.446(11)	3.354(3)	3.392(6)	3.113(6)	2.217(11)
M11	2.326(10)	2.421(10)	3.333(4)	3.340(7)	3.106(7)	2.213(10)
$\langle \text{Pb–O, X} \rangle$ 2,33	2,45	3,38	3,46	3,15	2,21	

Sample	Pb2–X2	Pb3–O1 2x	Pb3–X1	Pb3–X1a	Pb3–X2 2x
M1	3.047(4)	2.277(13)	3.097(4)	3.160(4)	3.343(2)
M2	3.057(3)	2.286(11)	3.088(3)	3.146(3)	3.3455(17)
M3	3.065(6)	2.29(3)	3.059(7)	3.127(7)	3.330(4)
M4	3.0836(13)	2.275(5)	3.0402(18)	3.1238(18)	3.3327(7)
M5	3.103(2)	2.265(9)	3.018(3)	3.111(3)	3.3290(13)
M6	3.100(5)	2.298(19)	2.989(10)	3.119(10)	3.314(3)
M7	3.090(4)	2.269(13)	3.002(7)	3.101(7)	3.308(3)
M8	3.060(2)	2.268(8)	2.978(4)	3.119(4)	3.3010(14)
M9	3.032(4)	2.276(9)	2.980(4)	3.125(4)	3.2963(19)
M10	2.976(5)	2.278(10)	2.977(6)	3.123(6)	3.276(3)
M11	2.892(7)	2.251(10)	2.960(7)	3.100(7)	3.243(4)
$\langle \text{Pb–O, X} \rangle$ 3.05	2.28	3.02	3.12	3.31	

O atoms, one X1 and one X2 atom. Pb3 is coordinated by two O atoms, two X1 atoms and two X2 atoms.

Coordination polyhedra for the Pb atoms are shown in Fig. 2. The short Pb–O bonds located on one side of the Pb²⁺ cations have lengths from 2.18 (Pb2–O1 in M3) to 2.48 Å (Pb1–O1 in M5) with the average $\langle \text{Pb–O} \rangle$ value of 2.32 Å, which is in a good agreement with the values suggested by Krivovichev and Filatov (2001). The average

$\langle \text{Pb–O–Pb} \rangle$ angles vary from 109.41° (M5) to 109.48° (M1), which is very close to the value of 109.5° for a regular tetrahedron. The reduction of one or several Pb–O–Pb angle values caused by edge sharing is compensated by increase of the other bond angles (Fig. 3). Average distances between the lead atoms within the OPb₄ tetrahedra ($\langle \text{Pb} \cdots \text{Pb} \rangle$) vary from 3.60 Å ($\langle \text{P3} \cdots \text{Pb1} \rangle$, $\langle \text{Pb2} \cdots \text{Pb1} \rangle$) to 3.92 Å ($\langle \text{Pb1} \cdots \text{Pb3} \rangle$, $\langle \text{Pb1} \cdots \text{Pb2} \rangle$)

Table 5. The Pb \cdots Pb distances in OPb₄ tetrahedra in the crystal structures of the mendipite series Pb₃O₂Cl₂–Pb₃O₂Br₂.

Sample	Pb1–Pb1, Å	Pb1–Pb3, Å	Pb3–Pb1, Å	Pb1–Pb2, Å	Pb1–Pb2, Å	Pb2–Pb3, Å	$\langle \text{Pb} \cdots \text{Pb} \rangle$, Å
M1	3.841(8)	3.954(6)	3.605(15)	3.609(9)	3.946(7)	3.814(7)	3.80
M2	3.841(11)	3.947(11)	3.607(9)	3.603(15)	3.943(9)	3.818(8)	3.79
M3	3.836(9)	3.935(8)	3.598(8)	3.595(14)	3.931(13)	3.807(12)	3.78
M4	3.839(14)	3.930(8)	3.602(10)	3.599(9)	3.932(11)	3.817(11)	3.79
M5	3.840(9)	3.921(9)	3.603(11)	3.598(11)	3.927(9)	3.820(11)	3.79
M6	3.834(10)	3.912(7)	3.599(12)	3.591(8)	3.918(15)	3.813(9)	3.78
M7	3.828(8)	3.907(10)	3.595(9)	3.597(8)	3.912(12)	3.809(10)	3.78
M8	3.831(11)	3.913(9)	3.603(11)	3.593(6)	3.913(14)	3.809(13)	3.78
M9	3.833(9)	3.922(12)	3.605(7)	3.598(17)	3.913(7)	3.809(7)	3.78
M10	3.828(8)	3.923(8)	3.606(10)	3.595(9)	3.906(8)	3.803(9)	3.78
M11	3.803(13)	3.903(8)	3.591(8)	3.569(7)	3.881(8)	3.777(16)	3.75
$\langle \text{Pb} \cdots \text{Pb} \rangle$, Å	3.83	3.92	3.60	3.60	3.92	3.81	

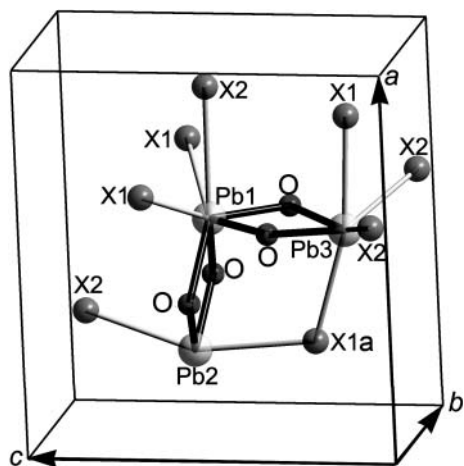


Fig. 4. Pb-X and Pb-O bonds orientation in the crystal structures of the mendipite series $\text{Pb}_3\text{O}_2\text{Cl}_2\text{-Pb}_3\text{O}_2\text{Br}_2$.

(Table 5). The shortest $\text{Pb1} \cdots \text{Pb2}$ distance between $[\text{O}_2\text{Pb}_3]^{2+}$ chains is 3.59 Å (M11). Edge sharing between two tetrahedra leads to repulsion of oxygen atoms and, as a result, Pb-Pb distances become shorter than those corresponding to the unshared edges.

Note the strong shortening of all $\text{Pb} \cdots \text{Pb}$ distances in the structure of the end-member $\text{Pb}_3\text{O}_2\text{Cl}_2$ (Table 5).

The weak Pb-X bonds located on the opposite side of the Pb^{2+} coordination sphere vary from 2.89 Å (Pb2-X2 in M11) to 3.57 Å (Pb1-X2 in M1). The evident strong distortion of the Pb^{2+} coordination polyhedra is due to the stereoactivity of the $6s^2$ lone electron pairs of the Pb^{2+} cations. The X1 halogen site is coordinated by five Pb^{2+} cations, but the X2 halogen site is coordinated by only four Pb^{2+} cations. The Pb-X bonds are oriented in space as follows (Fig. 4): Pb2-X1, Pb2-X2 are oriented along the *c* axis; Pb1-X2, Pb1-X1, Pb3-X1 – along the *b* axis; Pb1-X1, Pb3-X2 – along the *a* axis. The variations for the Pb-X bonds by the exchange of Br for Cl are: Pb1-X1 $\Delta = 0.1$; Pb1-X2 $\Delta = 0.23$; Pb2-X2 $\Delta = 0.2$; Pb2-X1 $\Delta = 0.12$; Pb3-X1 $\Delta = 0.14$; Pb3-X1a $\Delta = 0.06$; Pb3-X2 $\Delta = 0.11$ (Table. 4). Pb1-X2 and Pb2-X2 bonds (Fig. 4) are the most sensitive ones to the X site occupancy, which

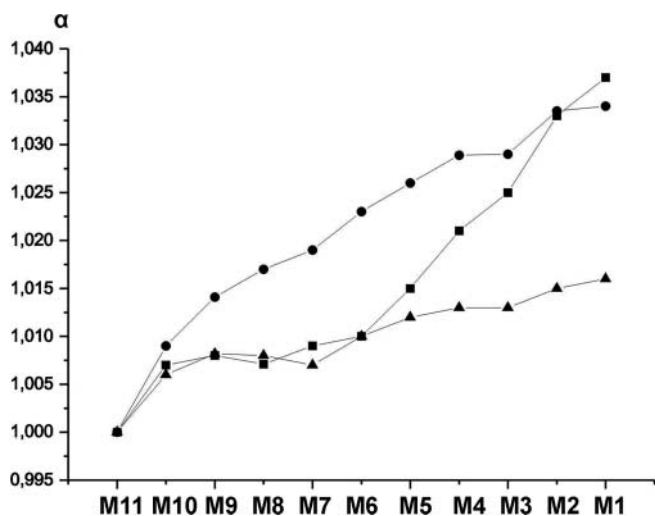


Fig. 5. Normalized lattice parameters ($a = a/a_{\text{Cl}}$ – ■, b/b_{Cl} – ▲, c/c_{Cl} – ●) in the crystal structures of the mendipite series $\text{Pb}_3\text{O}_2\text{Cl}_2\text{-Pb}_3\text{O}_2\text{Br}_2$.

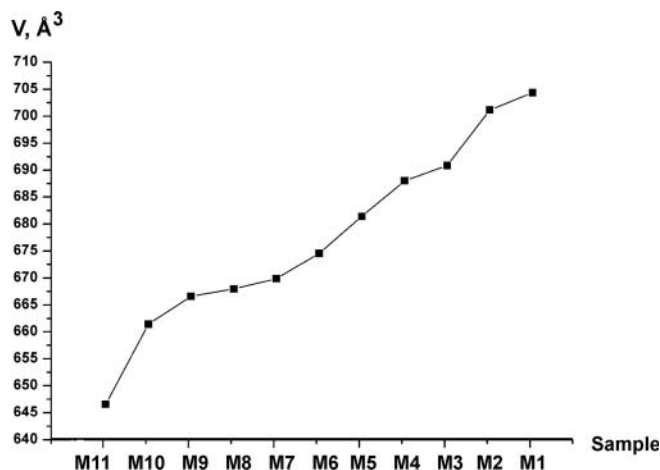


Fig. 6. Cell volume V , Å³ versus the batch number for mendipite-related $\text{Pb}_3\text{O}_2\text{Cl}_2\text{-Pb}_3\text{O}_2\text{Br}_2$ compounds.

is in agreement with the non-linear behavior of the *a* and *c* parameters (Fig. 5). The Pb-X bonds oriented along the *b* axis change only slightly.

The unit-cell parameters as determined by single crystal studies showed strong deviation from Vegard's rule (Vegard, Dale, 1928; Urusov, 1992), which can be clearly seen in Fig. 5. Note that the cell volume changes almost linearly (Fig. 6). Nonlinearity of the lattice parameters is caused by selective ordering of the halide anions over two crystallographically nonequivalent sites, X1 and X2 (Table 3; Fig. 7). Br atoms prefer the X2 position (Table 3), whereas Cl prefers the X1 site. Cl atoms are absent in the X2 site in M1–M3 samples. In M4–M6 samples, the X2 site is also predominantly occupied by Br atoms. Cl begins to prevail in this position only in the three last samples, M9–M11. On the contrary, the X1 site is preferred by Cl atoms. It is occupied exclusively by Cl in M10, M11 sample.

The φ angle between the faces of two adjacent OPb_4 tetrahedra (Fig. 8) was determined to analyze the influence of halogen atoms on the structure of the $[\text{O}_2\text{Pb}_3]^{2+}$ chain. From this analysis it is clear that different occupancy of the X1 site by Cl or Br atoms leads to the greatest φ angle changes. The X2 site does not influence the $[\text{O}_2\text{Pb}_3]^{2+}$

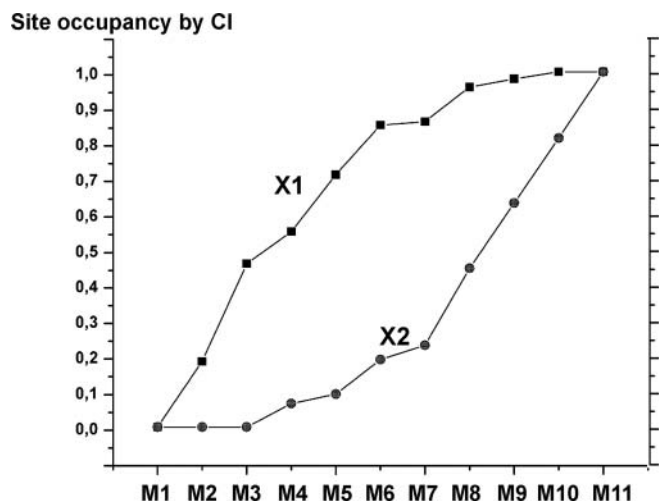


Fig. 7. The X1 (■) site and X2 (●) site occupancy by Cl atoms versus the batch number.

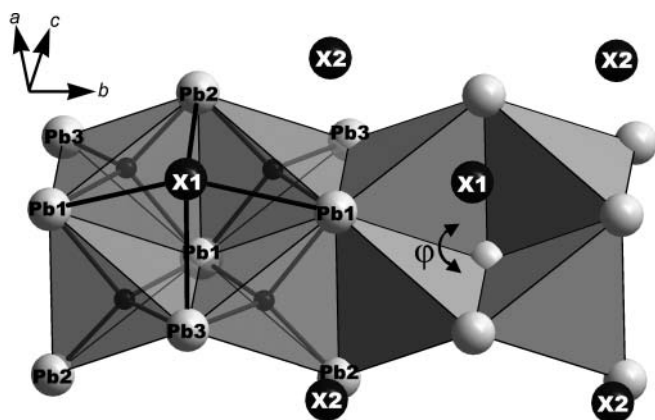


Fig. 8. The $[O_2Pb_3]^{2+}$ chain in the crystal structures of the mendipite series $Pb_3O_2Cl_2$ – $Pb_3O_2Br_2$. φ is the angle between the $Pb_2Pb_1Pb_1$ and $Pb_3Pb_1Pb_1$ faces of the two adjacent tetrahedra.

chain geometry that much. With the decrease in the X1 site occupancy (Table 3) by the Br atoms, the φ angle increases (Table 6, Fig. 9). Starting from the M7 composition, the φ angle begins to decrease, reaching $71.102(5)^\circ$ in M11 that contains Cl atoms only, thus “closing” or “pulling together” the square unit formed by the four adjacent tetrahedra of the $[O_2Pb_3]^{2+}$ chain. These observations may be interpreted as a conformation of the $[O_2Pb_3]^{2+}$ double chains in the crystal structures of the mendipite series compounds. The $[O_2Pb_3]^{2+}$ double chains obviously adapt their geometry to the size of the X atom. Br or Cl ions in the X1 site are attached to the square unit as shown in Fig. 8 and “stretching” or “pulling together” this unit, accordingly. This mechanism resembles the scheme of adaptation of XA_2 ($X=O, N; A = \text{metal}$) single anion-centered tetrahedral chains to the large halide ions in the crystal structures of some Ln nitro- and oxyhalides suggested by Krivovichev and Filatov (1998). From the M7 sample, the X1 site becomes essentially chlorine-rich (Table 3), which leads to the “closing” of the square unit formed by the four adjacent OPb_4 tetrahedra. Occupancy of the X1 site by bromine atoms becomes smaller than 0.15 (Table 3).

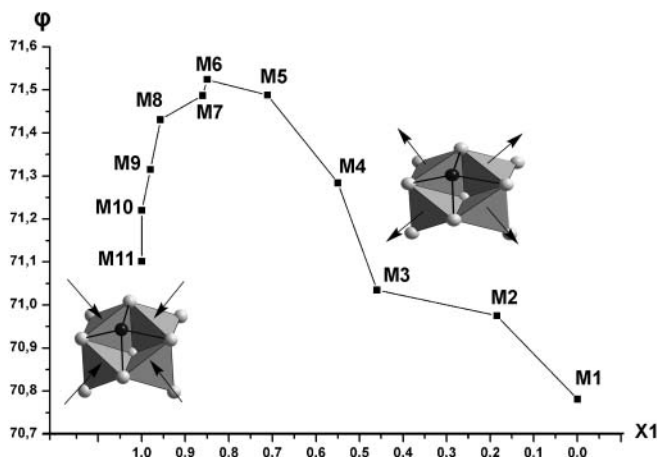


Fig. 9. The φ angle versus the X1 site occupancy in the crystal structures of the mendipite series $Pb_3O_2Cl_2$ – $Pb_3O_2Br_2$. The φ angle increases by the reduction of X1 site occupancy by the Br atoms, thus “opening” the square unit formed by the four adjacent tetrahedra in a $[O_2Pb_3]^{2+}$ chain. Starting from the M7 synthesis the value of φ angle again decreases, thus “pulling together” the unit. (See the text for details).

Thus, since the M7 composition, the crystal structure becomes less “strained”, which leads to the decrease in the φ angle.

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