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Effects of Oxide Additives on the Microstructure of Surrogate Nuclear Fuels

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I. BACKGROUND **Motivation for Research**

An increasing demand for nuclear energy requires reliable and efficient fuels. Release of fission gases in fuels leads to reduced thermal conductivity in the fuel-cladding gap and thus fuel reliability and efficiency(Fig.1.)[1].

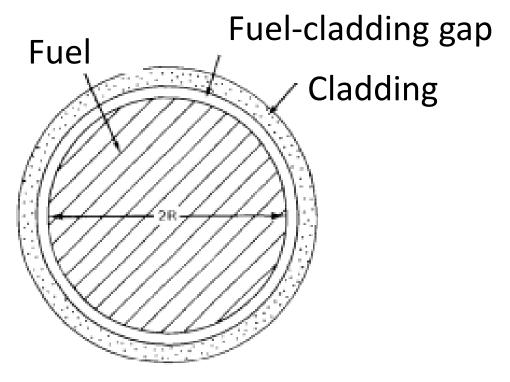


Fig.1. Cross sectional schematic of nuclear fuel-cladding system [2]. Larger grain sizes lead to improved fission product retention (Fig.2.). This study analyzes the impact of manganese dioxide (MnO_2) on the microstructure of cerium dioxide (CeO_2) , a surrogate for uranium dioxide (UO_2) .

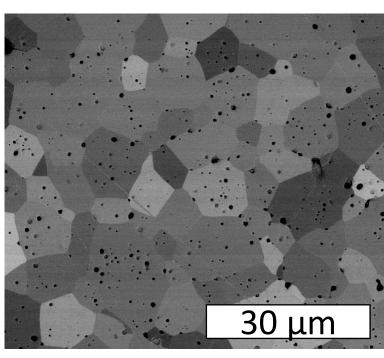


Fig.2. SEM micrograph of the microstructure of UO₂ with average grain size ~10 μ m [3].

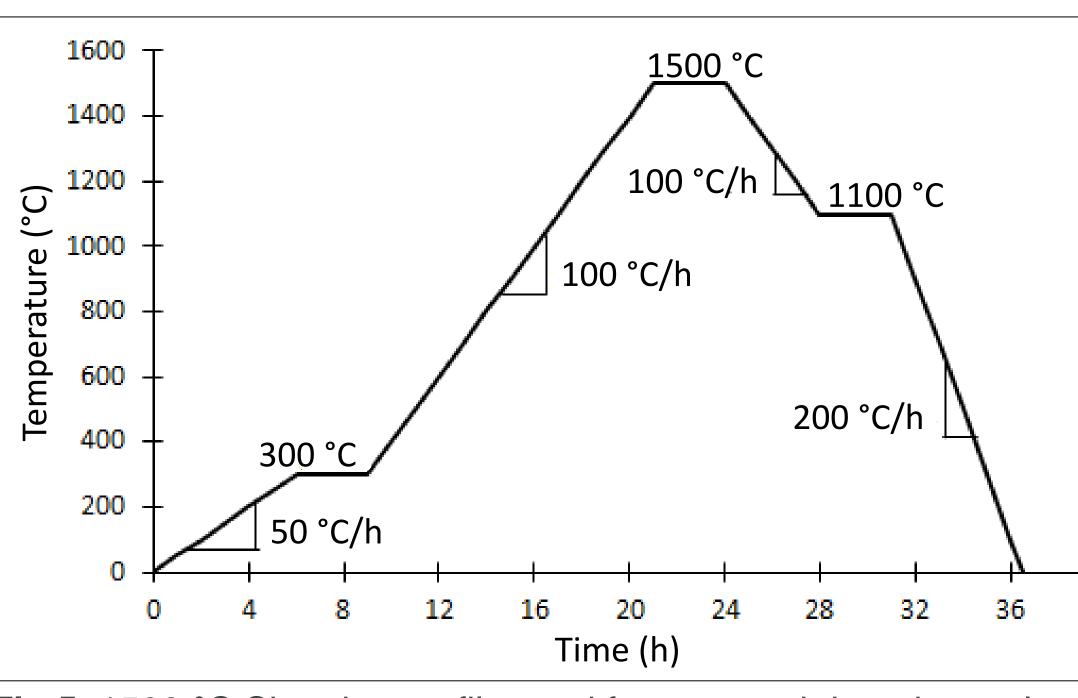
- A surrogate is used for UO₂ due to [4]:
 - Reduced radiation exposure
 - Decreased costs
 - Increased timeliness of experiments
- CeO₂ is used as a surrogate for UO_2 due to [4]:
 - Common cubic fluorite crystal structure
 - Similar melting temperature
 - Similar thermophysical properties

II. EXPERIMENTAL Materials Synthesis

- CeO_2 lattice (Fig.3. and Fig.4.).
- profile in Fig. 5.



Fig.3. Milled 500 ppm MnO₂doped CeO₂ powder. Pure, 1000, and 2500 ppm samples were also fabricated.



1200 °C and 1400 °C profiles were also tested.

Characterization Techniques

- incorporation analysis
 - instrumentation shifting.
- microstructural analysis
- analysis

st, M.W.D. Cooper, J. Spino, J.A. Turnbull, P. Van Uffelen, C.T. Walker, Fission gas release from UO2 nuclear fuel: A REFERENCES review, J. Nucl. Mater. 513 (2019) 310-345. [2] D.R. Olander, Fundamental Aspects of Nuclear Reactor Fuel Elements, National Technical Information Service. (1976) 1-[3] Z. Hiezl, D.I. Hambley, C. Padovani, W.E. Lee, Processing and microstructural characterisation of a UO2-based ceramic for disposal studies on spent AGR fuel, J. Nucl. Mater. 456 (2015) 74–84. https://doi.org/10.1016/j.jnucmat.2014.09.002.

microstructure of surrogate nuclear fuels Riley C. Winters^{1,2}, Adrianna E. Lupercio^{1,2}, Cayden Doyle^{1,2}, Andrew T. Nelson³, Brian J. Jaques^{1,2}

• As-received CeO₂ and MnO₂ powders were processed in a high energy planetary ball mill (HEPBM) to mix, reduce particle sizes, and to incorporate Mn⁺ into the

• Milled powder was pressed using a dual action die into right cylinder pellets at 150 MPa then sintered with the

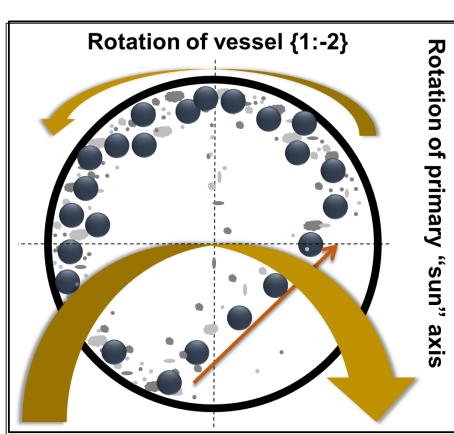


Fig.4. HEPBM motion schematic.

Fig.5. 1500 °C Sintering profile used for pure and doped samples.

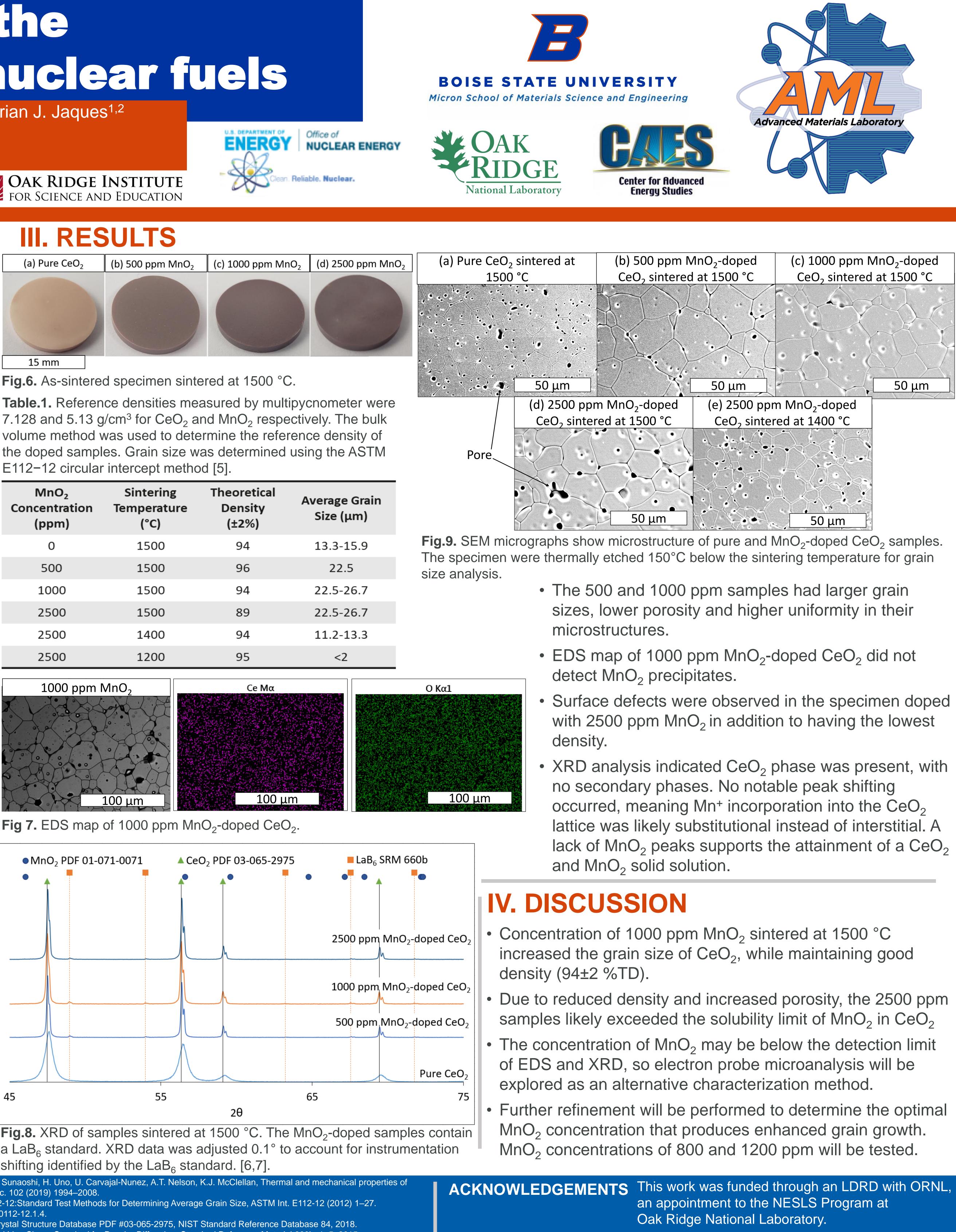
• X-ray diffraction (XRD) for phase purity and dopant

• A lanthanum hexaboride (LaB₆) standard was used in the MnO₂-doped samples to identify

Scanning electron microscopy (SEM) for

Energy dispersive spectroscopy (EDS) for chemical

CeO2, J. Am. Ceram. Soc. 102 (2019) 1994–2008. [5] ASTM Standard, E112-12:Standard Test Methods for Determining Average Grain Size, ASTM Int. E112-12 (2012) 1–27. https://doi.org/10.1520/E0112-12.1.4. [6] FIZ/NIST Inorganic Crystal Structure Database PDF #03-065-2975, NIST Standard Reference Database 84, 2018. [7] NIST Line Position and Line Shape Standard for Powder Diffraction, Standard Reference Material 660b, 1–5, 2015.



MnO ₂ Concentration (ppm)	Sinte Temper (°C
0	150
500	150
1000	150
2500	150
2500	140
2500	120

