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# Sorption of Phosphorus from Fertilizer Mixture

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## Abstract

Studying phosphorus (P) sorption behavior is among the prerequisites for P management in the crop fields. The work presented in this chapter described P sorption data when fertilizer mixture ( $\text{NH}_4\text{NO}_3$ ,  $\text{KH}_2\text{PO}_4$ , and KCl) was used to characterize sorption on soil. In addition to using fertilizer mixture, sorption experiments were also conducted using  $\text{KH}_2\text{PO}_4$  prepared in 0.01 M KCl, 0.005 M  $\text{CaCl}_2$ , and deionized water. The 24-h batch sorption experiments were conducted using a sandy soil to solution ratio of 1:2, and the equilibrium solution and sorbed data were described using Freundlich isotherm. Sorption kinetics experiments were conducted using times, 4, 8, 12, and 24 h. The Freundlich isotherm constant and sorbed P kinetics data for 0.005 M  $\text{CaCl}_2$  were significantly greater ( $p < 0.05$ ) than for 0.01 M KCl and/or fertilizer mixture. The Freundlich isotherm constant and sorbed P kinetics data for deionized water were significantly lower ( $p < 0.05$ ) than for 0.01 M KCl and/or fertilizer mixture. There was no significant difference in Freundlich isotherm constant and sorbed P kinetics data for 0.01 M KCl and fertilizer mixture. The sorption data showed the importance of using the fertilizer mix applied to the field when conducting sorption experiments.

**Keywords:** fertilizer mixture, isotherm, sorption coefficient, sorption kinetics

## 1. Introduction

Phosphorus (P) is applied with different nutrients to crop fields. Examples of field crops that need fertilizer mixture are shown in **Table 1**. Varying nutrient combinations can significantly affect the interactions of P with soil due to varying ionic strength and pH [1–7]. For example ionic strength was positively correlated to P sorption [2]. The specific affinity and the valence of the cation on the soil exchange site were also associated to P sorption capacity [7]. Supporting electrolytes are used for conducting P sorption experiments assuming representation of the true chemistry of the field solutions without necessarily considering the varying fertilizer mix applied to the soil. **Table 2** shows examples of supporting electrolytes that were used to characterize P sorption in the past studies. In this chapter, it was hypothesized that P sorption isotherm constants and kinetics data for fertilizer mix were significantly different from supporting electrolytes commonly used for conducting P sorption.

Sorption isotherms are used to describe relationships between sorbed and solution P in a given sorption experiment at constant temperature and act as indicators

Field crop	Fertilizer mixture distribution
Sugarcane	200 kg N, 50 kg P <sub>2</sub> O <sub>5</sub> , and 200 kg K <sub>2</sub> O per acre
Canola spring type	160 lb N, 30 lb P <sub>2</sub> O <sub>5</sub> , and 40 lb K <sub>2</sub> O per acre
Canola winter type	175 lb N, 30 lb P <sub>2</sub> O <sub>5</sub> , and 40 lb K <sub>2</sub> O per acre
Corn (for grain) dryland	120 lb N, 20 lb P <sub>2</sub> O <sub>5</sub> , and 20 lb K <sub>2</sub> O per acre
Corn (for grain) irrigated	180 lb N, 70 lb P <sub>2</sub> O <sub>5</sub> , and 70 lb K <sub>2</sub> O per acre
Cotton (1500 lb yield goal)	105 lb N, 140 lb P <sub>2</sub> O <sub>5</sub> , and 80 lb K <sub>2</sub> O per acre
Grain sorghum	80 lb N, 80 lb P <sub>2</sub> O <sub>5</sub> , and 80 lb K <sub>2</sub> O per acre
Peanuts	0 lb N, 80 lb P <sub>2</sub> O <sub>5</sub> , and 80 lb K <sub>2</sub> O per acre
Small grain-barley	100 lb N, 80 lb P <sub>2</sub> O <sub>5</sub> , and 80 lb K <sub>2</sub> O per acre
Small grain-oats	105 lb N, 80 lb P <sub>2</sub> O <sub>5</sub> , and 80 lb K <sub>2</sub> O per acre
Small grain-cover crop	60 lb N, 80 lb P <sub>2</sub> O <sub>5</sub> , and 80 lb K <sub>2</sub> O per acre
Small grain-wheat	120 lb N, 80 lb P <sub>2</sub> O <sub>5</sub> , and 80 lb K <sub>2</sub> O per acre
Small grain silage	160 lb N, 100 lb P <sub>2</sub> O <sub>5</sub> , and 160 lb K <sub>2</sub> O per acre
Sorghum silage	150 lb N, 80 lb P <sub>2</sub> O <sub>5</sub> , and 160 lb K <sub>2</sub> O per acre
Soybeans	0 lb N, 70 lb P <sub>2</sub> O <sub>5</sub> , and 100 lb K <sub>2</sub> O per acre
Sunflower	80 lb N, 80 lb P <sub>2</sub> O <sub>5</sub> , and 80 lb K <sub>2</sub> O per acre
Sweet sorghum	80 lb N, 80 lb P <sub>2</sub> O <sub>5</sub> , and 80 lb K <sub>2</sub> O per acre
Tobacco	50 lb N, 100 lb P <sub>2</sub> O <sub>5</sub> , and 180 lb K <sub>2</sub> O per acre
Kenaf	175 lb N, 100 lb P <sub>2</sub> O <sub>5</sub> , and 100 lb K <sub>2</sub> O per acre
Truffles	50 lb N, 80 lb P <sub>2</sub> O <sub>5</sub> , and 80 lb K <sub>2</sub> O per acre

**Table 1.**  
*Field crop and fertilizer mixture distributions.*

of field P retention potential [24–27]. Sorption coefficient is among the coefficients described in the isotherms that is used to model P movement in the field [28]. Phosphorus sorption capacity has also been used as an important management tool in many crop fields [29]. Therefore, there is a need to identify the appropriate chemistry of the field solutions before conducting P sorption experiments and modeling P movement in the crop fields. Sorption kinetics data trends were reported to provide clues on the mechanisms of sorption reactions [30]; appropriate solution chemistry should also be carefully chosen for the sorption kinetics experiments. It was also hypothesized that the isotherms that describe sorption data from fertilizer mixture are different from isotherms that describe data from typical laboratory supporting electrolytes.

The importance of laboratory P sorption and kinetics data in modeling and understanding of the P dynamics in crop fields has been documented [31, 32]. The P sorption characteristics help to properly calibrate theoretical models that aim at mimicking field processes [31, 32]. Accurate laboratory sorption data collected using true field solution chemistry will therefore improve models as predictive tools for P movement. The objective of the study was to determine the differences in P sorption behavior for P in fertilizer mixture (N, P, and K) prepared in deionized water and in P fertilizer (KH<sub>2</sub>PO<sub>4</sub>) prepared in 0.01 MKCl, 0.005 M CaCl<sub>2</sub>, and deionized water.

Electrolyte	Isotherm	Soil	Reference
0.1 M CaCl <sub>2</sub>	ND	- Typic Argiudolls - Typic Hapludolls - Entic Haplustolls - Abruptic Argiudolls - Petrocalcic Paleudolls	[8]
0.01 M CaCl <sub>2</sub>	- Langmuir - Freundlich - Temkin	- Vertisol - Rhizospheric soil - Haploboroll - Hapludoll - Eutrochrept - Haplaquept - Brunic Arenosols (dystric) - Haplic Regosol (dystric) - Sandy mixed Humic Dystrochrept - Very fine, mixed, semiactive, Oxyaquic Haplocryoll - Fine, illitic, frigid Typic Haplaquept - Coarse-loamy, mixed, mesic Oxyaquic, Eutrochrept	[1, 9–14]
0.001 M CaCl <sub>2</sub>	- Freundlich	- Haplustalf - Orthent - Tropaquept	[15]
0.01 M KCl	-Langmuir -Linear	- Sandy, siliceous, hyperthermic Aeric Alaquods - Alaquods and Alorthods - Loamy, siliceous, subactive, thermic Arenic Paleudults - Fine-loamy, siliceous, subactive, thermic Aquic Paleudults - Fine, mixed, semiactive, thermic Typic Umbraquults - Fine-loamy, mixed, semiactive, acid, thermic Histic Humaquepts	[16–19]
0.02 M KCl	- Langmuir	- Loamy, siliceous, hyperthermic Arenic Glossaqualf	[20]
0.05 M KCl	- Langmuir - Freundlich - Linear	- Quartzipsamments - Paleudults - Loamy-skeletal, carbonatic hyperthermic Lithic Udorthents - Loamy, carbonatic, hyperthermic, shallow Typic Fluvaquents - Loamy, skeletal, carbonatic, hyper thermic, Lithic Udorthents	[4, 21, 22]
0.1 M NaNO <sub>3</sub>	- Langmuir	- Sandy, siliceous, hyperthermic Ultic Alaquod	[23]
0.1 M NaCl	ND	- Aquic or Oxyaquic Haplocryods	[3]
Deionized water	ND	- Aquic or Oxyaquic Haplocryods	[3]

*ND, not determined.*

**Table 2.**  
 Supporting electrolytes, soils, and sorption isotherms for the literature studies.

## 2. Sorption experiments and trends in sorption data

### 2.1 Determination of soil properties

The soil samples used for sorption experiments were air dried, passed through 2-mm sieve and first analyzed for pH, total carbon, oxalate extractable iron, oxalate extractable aluminum, and exchangeable calcium. Particle size distribution of the

soil samples was also determined. A soil to water solution ratio of 1:2 was prepared, and the soil pH was measured using a standardized pH meter (model: AR15; manufacturer: Fisher Scientific) [33]. The combustion method with the element analyzer (Carbo-Erba NA 2500 instrument (Model: NA 2500; manufacturer: CE instruments, Italy) was used to measure total carbon. The inductively coupled plasma (ICP) (model: Optima 700 DV; manufacturer: Perkin Elmer) was used to analyze for oxalate iron and aluminum after extraction with oxalate solution [34]. Exchangeable calcium was also analyzed using ICP after extraction with 0.2 M  $\text{NH}_4\text{Cl}$  [35]. Particle size distribution was determined by hydrometer method [35].

## 2.2 Determination of sorption isotherms

An example of P sorption experiment that involved using  $\text{KH}_2\text{PO}_4$  fertilizer prepared in 0.01 M KCl, 0.005 M  $\text{CaCl}_2$ , and deionized water and fertilizer mixture ( $\text{NH}_4\text{NO}_3$ ,  $\text{KH}_2\text{PO}_4$ , and KCl) prepared in deionized water was used for this study. The fertilizer rates, 50 kg  $\text{P}_2\text{O}_5 \text{ ha}^{-1}$ , 200 kg N  $\text{ha}^{-1}$ , and 200 kg  $\text{K}_2\text{O} \text{ ha}^{-1}$ , applied to sugarcane fields were used to prepare the fertilizer mixture. Potassium chloride and calcium chloride were used because they are commonly used to conduct P sorption experiments with the assumption that the solutions' ionic strength and pH are close to those of the crop fields. The concentration, 0.005 M for  $\text{CaCl}_2$  and 0.01 M for KCl were used to attain the equivalences of  $\text{Ca}^{2+}$  and  $\text{K}^+$ . Deionized water was used because irrigation water is used to provide the necessary plant moisture.

The two sandy soils, Margate (sandy, siliceous, hyperthermic Mollic Psammaquents) and Immokalee (sandy, siliceous, hyperthermic Arenic Alaquods) used for the experiment contribute the most to the greatest percentage of soils used for sugarcane production in Southwestern Florida. Five soil samples of each of the soil horizons, A and Bh for Immokalee soil and A and Bw for Margate soil, were used to represent the varying soil properties (e.g., total carbon, iron, and aluminum). The soil samples were sampled from two sugarcane fields each of 12 ha located in Hendry County, southwestern Florida (26.75° N, 80.93° W).

The initial P concentrations ( $C_0$ ) used for the experiment ranged from 8 to 60  $\text{mg L}^{-1}$ . The soil to solution ratio of 1:2 (10 g of soil and 20 mL of solution) was used and equilibrium solution concentration was analyzed after 24 h of shaking. Blanks where soil was shaken with only 0.01 M KCl, 0.005 M  $\text{CaCl}_2$ , and deionized water were also included in the experiment, and the blank equilibrium concentrations were subtracted from the treatment sample equilibrium concentrations. The experiments were conducted at room temperature (25°C). Before analyzing the solution concentrations, soil solutions were centrifuged at 5000 rpm for about 20 min and filtered using 42 Whatman filter. The spectrophotometer (HACH DR/4000U) was used to analyze solution P at a detection wavelength of 880 nm.

Sorbed P ( $S$ ) was equal to  $V/M (C_0 - C_e)$  where  $V$ ,  $M$ ,  $C_0$ , and  $C_e$  are volume of solution, mass of soil, initial solution P concentration, and equilibrium P concentration, respectively. Sorption data for all the supporting electrolytes were fitted to Freundlich isotherm [sorbed ( $S$ ) versus equilibrium solution ( $C$ ) concentration]. The Freundlich sorption isotherm is represented by Eq. (1).

$$S = K_f C^N \quad (1)$$

where  $S$  is the amount of P sorbed ( $\text{mg kg}^{-1}$ ),  $C$  is the solution P concentration ( $\text{mg L}^{-1}$ ),  $K_f$  is the Freundlich sorption coefficient ( $\text{L}^N \text{ kg}^{-1} \text{ mg}^{1-N}$ ), and  $N$  is an empirical constant. The coefficients for fertilizer mixture were compared

to coefficients for 0.01 M KCl, 0.005 M CaCl<sub>2</sub>, and deionized water. The paired t-test was used to identify significant differences in Freundlich sorption coefficients.

### 2.3 Sorption kinetics experiments

The sorption of P has been assumed as a kinetic process [15, 28]. Three initial concentrations (C<sub>0</sub>), 19, 29, and 38 mg L<sup>-1</sup> were used for conducting sorption kinetics experiment. The A horizon of Immokalee soil was used for sorption kinetics experiment, and the soil to solution ratio of 1:2 (10 g of soil and 20 mL of solution) was used. Solution concentrations were analyzed after 4, 8, 12, and 24 h. The paired t-test was used to identify significant differences in relative concentrations (C/C<sub>0</sub>) and sorbed concentrations between fertilizer mixture and supporting electrolytes. R-software was used for statistical analyses. Graphs of relative concentrations (C/C<sub>0</sub>) and sorbed concentrations (S) versus time were plotted to show the data trends over a 24 h period.

### 2.4 Selected soil properties

The average percent sand, pH (1:2 soil:water volume), total carbon, oxalate iron, oxalate aluminum, and exchangeable Ca for A horizon of Immokalee soil were 97.0%, 6.8, 15.2 g kg<sup>-1</sup>, 234.0 mg kg<sup>-1</sup>, 280.4 mg kg<sup>-1</sup>, and 3.6 cmolc kg<sup>-1</sup>, respectively. The average percent sand, pH (1:2 soil:water volume), total carbon, oxalate iron, oxalate aluminum, and exchangeable Ca for A horizon of Margate soil were 97.5.0%, 8.3, 11.2 g kg<sup>-1</sup>, 661.1 mg kg<sup>-1</sup>, 307.3 mg kg<sup>-1</sup>, and 6.6 cmolc kg<sup>-1</sup>, respectively. The average percent sand, pH (1:2 soil:water volume), total carbon, oxalate iron, oxalate aluminum, and exchangeable Ca for Bh were 87.5%, 6.8, 39.7 g kg<sup>-1</sup>, 114.4 mg kg<sup>-1</sup>, 305.0 mg kg<sup>-1</sup>, and 5.4 cmolc kg<sup>-1</sup>, respectively. The average percent sand, pH (1:2 soil:water volume), total carbon, oxalate iron, oxalate aluminum, and exchangeable Ca for Bw were 97.2%, 8.4, 3.9 g kg<sup>-1</sup>, 149.0 mg kg<sup>-1</sup>, 89.0 mg kg<sup>-1</sup>, and 2.0 cmolc kg<sup>-1</sup>, respectively.

### 2.5 Changes of sorption isotherm coefficients with supporting electrolytes

Although all sorption data fitted Freundlich isotherms with R<sup>2</sup> values greater than 0.9, the Freundlich coefficients varied with the type of supporting electrolytes (Table 3). For both 0.01 M KCl and fertilizer mixture, the Freundlich isotherm constant was significantly lower ( $p < 0.05$ ) than for 0.005 M CaCl<sub>2</sub> and significantly greater ( $p < 0.05$ ) than for deionized water (Table 4). Although the same equivalence was used for K<sup>+</sup> and Ca<sup>2+</sup>, sorption was greater for 0.005 M CaCl<sub>2</sub> than

Soil	Horizon	0.01 M KCl	0.005 M CaCl <sub>2</sub>	Deionized water	Fertilizer mixture
Immokalee soil	A	S = 4.8 C <sup>0.5</sup>	S = 13.5 C <sup>0.4</sup>	S = 2.7C <sup>0.8</sup>	S = 4.6C <sup>0.5</sup>
Margate soil	A	S = 7.2 C <sup>0.6</sup>	S = 24.6 C <sup>0.3</sup>	S = 5.4C <sup>0.5</sup>	S = 7.3 C <sup>0.6</sup>
Immokalee soil	Bh	S = 19.1C <sup>0.6</sup>	S = 79.0C <sup>0.2</sup>	S = 13.1C <sup>0.8</sup>	S = 21.1C <sup>0.6</sup>
Margate soil	Bw	S = 10.3C <sup>0.3</sup>	S = 28.0 C <sup>0.2</sup>	S = 6.7C <sup>0.4</sup>	S = 9.5C <sup>0.3</sup>

**Table 3.**  
 Average sorption isotherms of five replicates showing variabilities in Freundlich sorption coefficients.

Comparisons	A-Immokalee	A-Margate	Bh	Bw
Fertilizer mixture versus deionized water	S	S	S	S
Fertilizer mixture versus 0.005 M CaCl <sub>2</sub>	S	S	S	S
Fertilizer mixture versus 0.01 M KCl	NS	NS	NS	NS

*NS, no significant difference ( $\alpha = 0.05$ ); S, significantly greater or lower ( $\alpha = 0.05$ ).*

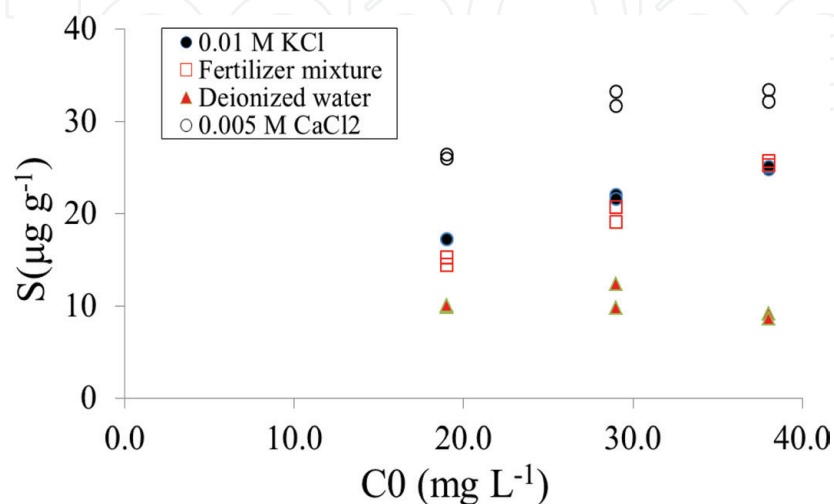
**Table 4.**

*Comparisons of Freundlich sorption coefficients and sorbed phosphorus concentrations.*

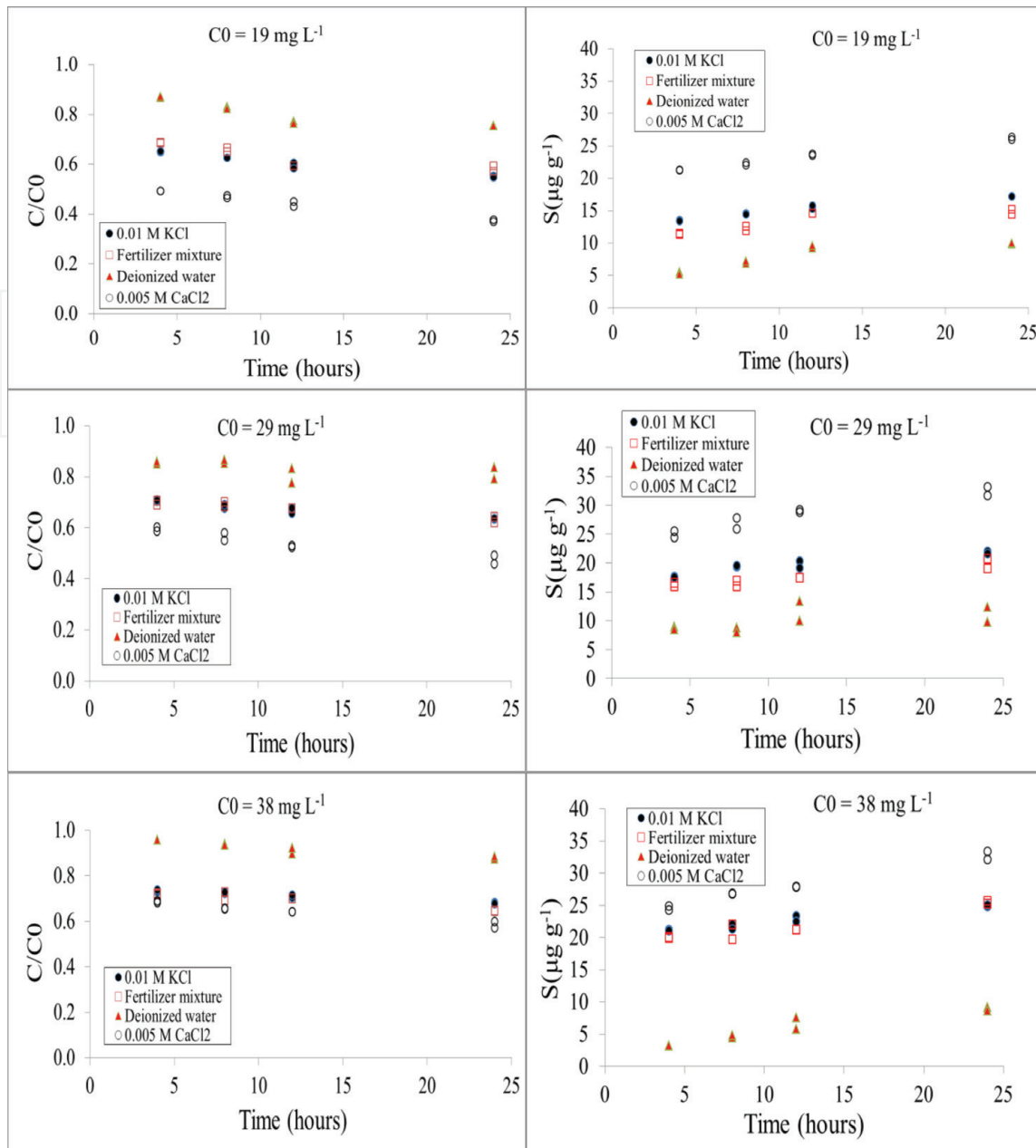
for 0.01 M KCl and fertilizer mixture likely due to the influence of charge (+2) on Ca<sup>2+</sup> that reduce the electrostatic repulsion effect between phosphate and the soil surface. A similar trend (Ca<sup>2+</sup> > K<sup>+</sup>) for P sorption was identified in other sorption studies [36]. Phosphorus sorption was also reported to increase with increase in background electrolyte concentration [37]. The sorption characteristics of P (KH<sub>2</sub>PO<sub>4</sub>) prepared in deionized water was lower than for 0.01 M KCl and fertilizer mixture probably due to significantly lower K<sup>+</sup> concentration that contributes less to ionic strength than in the latter two. The greater Freundlich coefficients for Bh were greater than A horizon because of the greater total carbon, oxalate iron, and oxalate aluminum that enhance greater P sorption. Soils with greater free aluminum and iron were associated with greater P sorption by different researchers [38–40].

## 2.6 Trends in sorption kinetics data for different supporting electrolytes

**Figure 1** shows the sorbed P concentrations as a function of initial concentrations for a 24 h time step with no significant difference in sorbed concentrations for 0.01 M KCl and fertilizer mixture. **Figure 2** shows the trends in relative solution concentrations (C/C<sub>0</sub>) and sorbed concentrations (P). While the relative solution concentrations decreased over time, sorbed concentrations increased over time. For both 0.01 M KCl and fertilizer mixture, sorbed P kinetics data were significantly lower ( $p < 0.05$ ) than for 0.005 M CaCl<sub>2</sub>, and significantly greater ( $p < 0.05$ ) than for deionized water (**Table 4**). Sorption was fast for the first hours due to the presence of high P affinity sorption sites on the exchange sites and gradual for the following hours (**Figure 2**). A fast P sorption first phase followed by a steady phase was also documented in other studies [2, 15, 41].

**Figure 1.**

*Sorbed P (S) concentration as a function of initial concentrations (C<sub>0</sub>) for A horizon of Immokalee soil.*



**Figure 2.**  
 Relative solution ( $C/C_0$ ) and sorbed ( $S$ ) concentration as a function of time for A horizon of Immokalee soil.

### 3. Summary/conclusions

The results presented in this chapter suggest that if another nutrient is applied with P in the field, the P sorption behavior should be studied with the applied fertilizer mix, and P prepared in recommended supporting electrolyte as well. The sorption characterization with the two scenarios will help in identifying the appropriate sorption characteristics (sorption isotherm coefficients and kinetics constants) used for predicting P movement and P management options.

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## References

- [1] Ahmed MF, Kennedy IR, Choudhury ATMA, Kecske ML, Deaker R. Phosphorus adsorption in some Australian soils and influence of bacteria on the desorption of phosphorus. *Communications in Soil Science and Plant Analysis*. 2008;**39**:1269-1294
- [2] Wang S, Jin X, Bu Q, Zhou X, Wu F. Effect of particle size, organic matter and ionic strength of phosphate sorption in different trophic lake sediments. *Journal of Hazardous Materials*. 2006:95-105
- [3] Giesler R, Anderson T, Lovgren L, Persson P. Phosphate sorption in aluminum- and iron-rich humus soils. *Soil Science Society of America Journal*. 2005;**69**(1):77-86
- [4] Violante A, Pigna M. Competitive sorption of arsenate and phosphate on different clay minerals and soils. *Soil Science Society of America Journal*. 2002;**66**(6):1788-1796
- [5] Curtin D, Syers JK, Bolan NS. Phosphate sorption by soil in relation to exchangeable cation composition and pH. *Australian Journal of Soil Research*. 1992;**31**:137-149
- [6] Froelich PN. Kinetic control of dissolved phosphate in natural rivers and estuaries: A primer on the phosphate buffer mechanism. *Limnology and Oceanography*. 1988;**33**(4 pt 2):649-668
- [7] Barrow NJ, Bowden JW, Posner AM, Quirk JP. Describing the effects of electrolyte on adsorption of phosphate by a variable charge surface. *Australian Journal of Soil Research*. 1980;**18**:395-404
- [8] Rubio G, Cabello MJ, Boem FHG, Munaro E. Estimating available soil phosphorus increases after phosphorus additions in Mollisols. *Soil Science Society of America Journal*. 2008;**72**:1721-1727
- [9] Börling K, Otabbong E, Barberis E. Phosphorus sorption in relation to soil properties in some cultivated Swedish soils. *Nutrient Cycling in Agroecosystems*. 2001;**59**:39-46
- [10] Debicka M, Kocowicz A, Weber J, Jamroz E. Organic matter effects on phosphorus sorption in sandy soils. *Archives of Agronomy and Soil Science*. 2015. DOI: 10.1080/03650340.2015.1083981
- [11] Daly K, Jeffrey D, Tunney H. The effect of soil type on phosphorus sorption capacity and desorption dynamics in Irish grassland soils. *Soil Use and Management*. 2001;**17**:12-20
- [12] Djodjic F, Borling K, Bergstrom L. Phosphorus leaching in relation to soil type and soil phosphorus content. *Journal of Environmental Quality*. 2004;**33**(2):678-684
- [13] Maguire RO, Foy RH, Bailey JS, Sims JT. Estimation of the phosphorus sorption capacity of acidic soils in Ireland. *European Journal of Soil Science*. 2001;**52**:479-487
- [14] McDowell R, Condron L. Influence of soil constituents on soil phosphorus sorption and desorption. *Communications in Soil Science and Plant Analysis*. 2001;**32**(15, 16):2531-2547
- [15] Agbenin JO, Tiessen H. Phosphorus sorption at field capacity and soil ionic strength: Kinetics and transformation. *Soil Science Society of America Journal*. 1995;**59**:998-1005
- [16] Reddy KR, O'Connor GA, Gale PM. Phosphorus sorption capacities of wetlands soils and stream

sediments impacted by dairy effluent. *Journal of Environmental Quality*. 1998;**27**:438-447

[17] Villapando RR, Graetz DA. Phosphorus sorption and desorption properties of the spodic horizon from selected Florida Spodosols. *Soil Science Society of America Journal*. 2001;**65**:331-339

[18] Kang J, Hesterberg D, Osmond DL. Soil organic matter effects on phosphorus sorption: A path analysis. *Soil Science Society of America Journal*. 2009;**73**(2):360-366

[19] Ige DV, Akinremi OO, Flatten DN. Direct and indirect effects of soil properties on phosphorus retention capacity. *Soil Science Society of America Journal*. 2007;**71**(1):95-100

[20] He ZL, Alva AK, Li YC, Calvert DV, Banks DJ. Sorption-desorption and solution concentration of phosphorus in a fertilized sandy soil. *Journal of Environmental Quality*. 1999;**28**:1804-1810

[21] Zhou M, Li Y. Phosphorus-sorption characteristics of calcareous soils and limestone from the southern Everglades and adjacent farmlands. *Soil Science Society of America Journal*. 2001;**65**:1404-1412

[22] Rhue RD, Harris WG, Nair VD. A retardation-based model for phosphorus transport in sandy soil. *Soil Science*. 2006;**171**:293-304

[23] Bhatti JS, Comerford NB, Johnston CT. Influence of oxalate and soil organic matter on sorption and desorption of phosphate onto a spodic horizon. *Soil Science Society of America Journal*. 1998;**62**:1089-1095

[24] Chen Z, Xing B, McGill WB. A unified sorption variable for environmental applications of the Freundlich equation. *Journal*

*of Environmental Quality*. 1999;**28**(5):1422-1128

[25] Idris OAA, Ahmed HS. Phosphorus sorption capacity as a guide for phosphorus availability of selected Sudanese soil series. *African Crop Science Journal*. 2012:59-65

[26] Travis SC, Etnier EL. A survey of sorption relationships for reactive solutes in soil. *Journal of Environmental Quality*. 1981;**10**(1):8-17

[27] Richardson CJ. Mechanisms controlling phosphorus retention capacity in freshwater wetlands. *American Association for the Advancement of Science*. 1985;**228**(4706):1424-1427

[28] Šimůnek J, van Genuchten MT. Modeling nonequilibrium flow and transport processes using HYDRUS. *Vadose Zone Journal*. 2008;**7**(2):782-797

[29] Beauchemin S, Simard RR. Soil phosphorus saturation degree: Review of some indices and their suitability for P management in Québec, Canada. *Canadian Journal of Soil Science*. 1999:615-628

[30] Wang Y, Shen Z, Niu J, Liu R. Adsorption of phosphorus on sediments from the three gorges reservoir (China) and the relation with sediment compositions. *Journal of Hazardous Materials*. 2009;**162**:92-98

[31] Simunek J, Sejna J, van Genuchten MT. *The HYDRUS-2D Software Package for Simulating the Two and One Dimensional Movement of Water, Heat, and Multiple Solutes in Variably Saturated Media*. USDA, Riverside, CA: US Salinity Laboratory Agricultural Research Service; 1999

[32] Hassan G, Reneau RB, Hagedorn C, Jantrania AR. Modeling effluent distribution and nitrate transport

through an on-site wastewater system.  
*Journal of Environmental Quality*.  
2008;**37**:1937-1948

sediment phosphorus in a tropical  
reservoir. *Journal of Environmental  
Engineering*. 2000:993-998

[33] Sparks DL. Soil Science Society of  
America Series (SSSA) 5: Methods of  
Soil Analysis. 1996

[34] McKeague JA, Day JH. Dithionite  
and oxalate extractable Fe and Al as  
aids in differentiating various classes of  
soils. *Canadian Journal of Soil Science*.  
1966;**46**:13-22

[35] Soil Survey Staff. Keys to Soil  
Taxonomy. Washington, DC: U.S. Gov.  
Print. Office; 1996

[36] Pardo MT, Guadalix ME, Garcia-  
Gonzalez MT. Effect of pH and  
background electrolyte on P sorption  
by variable charge soils. *Geoderma*.  
1992;**54**:275-284

[37] Antoniadis V, Koliniati R, Efstratiou  
E, Golia E, Petropoulos S. Effect of soils  
with varying degree of weathering and  
pH values on phosphorus sorption.  
*Catena*. 2016;**139**:214-219

[38] Gerard F. Clay minerals, iron/  
aluminum oxides, and their  
contribution to phosphate sorption in  
soils—A myth revisited. *Geoderma*.  
2016;**262**:213-226

[39] Kisinyo PO, Othieno CO, Gudu  
SO, Okalebo JR, Opala PA, Maghanga  
JK, Ng'etich WK, Agalo JJ, Opile RW,  
Kisinyo JA, Ogola BO. Phosphorus  
sorption and lime requirements of  
maize growing acid soils of Kenya.  
*Sustainable Agriculture Research*.  
2013;**2**(2):116-123

[40] Wolde Z, Haile W. Phosphorus  
sorption isotherms and external  
phosphorus requirements of some soils  
of Southern Ethiopia. *African Crop  
Science Journal*. 2015;**23**(2):89-99

[41] Appan A, Wang H. Sorption  
isotherms and kinetics of