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New Insight into the Influence of Rhamnolipid Bio-Surfactant on the Carbonate Rock/Water/Oil Interaction at Elevated Temperature

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Abstract: Tertiary recovery is directly dependent on the alteration in wettability and interfacial tension (IFT), hence releasing the trapped oil from rock pore spaces. Bio-surfactant water flooding to mobilise residual oil in reservoirs is a new and developing prospect that can be used more often in future due to its environmentally friendly nature and economic advantages. In this work, the impact of rhamnolipids as water soluble bio-surfactant solutions on the interfacial activities of saline water and the wettability of carbonate rock are studied at elevated temperature. The effectiveness of the bio-surfactant as a rock wettability modifier is analysed in the presence of different salinities, in particular SO_4^{2-} ions. The reason for the focus on SO_4^{2-} is its high affinity towards calcite surfaces, and hence its ability to intervene strongly on bio-surfactant performance. To achieve the objectives of this study, the oil-wet calcite samples at elevated temperature were put through a washing process that included bio-surfactant solutions in seawater at various concentrations of sodium sulphate ions, where the measurement of the contact angles of each sample after treatment and the IFT between the oil model and the washing solutions were taken. The obtained results illustrated that bio-surfactants (rhamnolipids) with incremental concentrations of SO_4^{2-} ions in sea water (up to three times higher than the original ion concentration) can lower the IFT, and assessed changing the rocks towards greater water-wettability. This study reveals that the alteration of SO_4^{2-} ions had a greater impact on the wettability alteration, whereas rhamnolipids were better at reducing the IFT between the oil phase and the aqueous phase. This study also looked at temperatures of 50 °C and 70 °C, which demonstrated undesired influences on the wettability and IFT. Bio-surfactants at high temperature showed less interfacial activity, thus indicating that rhamnolipids are not active at high temperatures, while the addition of SO_4^{2-} shows a continuous decrease in the contact angle and IFT measurements at high temperature.

Keywords: carbonate rock; bio-surfactant; rhamnolipids; wettability; contact angle; sulphate ions; elevated temperature; IFT

1. Introduction

With the world's energy demands growing, oil recovery has become a global priority. The development of oil extraction techniques has gradually changed and improved in order to increase oil recovery. Worldwide, oil production only recovers on average about a third of the oil originally present in reservoirs, which is the amount of oil recovered economically using conventional methods. The remaining two-thirds of residual oil in the reservoirs is the focus of Enhanced Oil Recovery (EOR) methods [1].

Microbial enhanced oil recovery (MEOR) represents the use of microbes in oil recovery, which is a tertiary EOR technique still under development. MEOR is widely appropriate in carbonate and

sandstone reservoirs for both light and heavy crude oil. MEOR has been proposed as an effective and cheap alternative for EOR, and involves utilising the metabolic products the microorganisms produce to aid the oil production. In general, there are seven types of bioproducts, which include biopolymers, bio-surfactants, biomass, solvents, acids, emulsifiers, and gases. It is noted that bio-surfactants are more attractive compared to other bioproducts due to their low toxicity, stability, and impact on the wettability alteration to increase enhanced oil recovery [2].

Bio-surfactants are amphiphilic, consisting of both hydrophobic and hydrophilic parts. Bio-surfactants have the potential to lower the surface and interfacial tension (IFT) by gathering at the interface between the two immiscible fluids. This reduces the repulsive forces and allows the two phases to interact and mix more easily; it also improves the mobility and solubility of the insoluble or hydrophobic organic compounds [3,4].

Rhamnolipids are a type of bio-surfactant that are produced mainly by *Pseudomonas aeruginosa*. It should be noted that because of their low toxicity, production from renewable sources, and antimicrobial (particularly antifungal) activity, the use of rhamnolipid bio-surfactants shows great promise for broad commercial application [5].

The concentration of a bio-surfactant is effective until the critical micelle concentration (CMC) is obtained. CMC is associated with micelle formation, which allows bio-surfactants to reduce the IFT and increase the solubility and bioavailability of a hydrophobic compound. CMC is known as the efficiency of the bio-surfactant, where less bio-surfactant is required to lower the IFT; an effective bio-surfactant will have a low CMC [6]. The remaining oil residual in reservoirs is generally located in restricted access regions where oil is trapped in pores by capillary pressure. Bio-surfactants can lower the IFT between oil/rock and water/oil and also alter the wettability to water-wet the system—this reduces the capillary forces, which allows water to move through the rock pores [7].

Al-Sulaimani et al. [8] found that the maximum production of residual oil went up by 50% by using bio-surfactant and chemical surfactant in a 50:50 mixing ratio. The study also looked at the influence of the bio-surfactant on the wettability of calcite surface by studying contact angle measurements. The results showed that a 0.25% treated (*w/v*) bio-surfactant solution changed to greater water-wetting as the angle decreased from 70.6 to 25.32°. Other studies [9–11] presented similar trends in the use of different bio-surfactants for the displacement of oil by water.

Wettability tests performed on carbonate reservoirs at elevated temperature have showed that increasing temperatures alter the wettability towards greater water-wetting, which was attributed to the de-attachment of oil-wet particles (i.e., calcium stearate) from the calcite surface, making the surface more water-wet [12–14]. In general, wettability and IFT are considered to be the main controlling parameters of fluid flow in carbonate rocks, particularly for high temperature reservoir conditions.

Water salinity has also been shown to have a significant impact via an increase of the capillary force, lowering of the IFT, and altering of the wettability of carbonate reservoirs [15,16]. Seawater is an example of brine water that contains different active ions such as calcium (Ca^{2+}), magnesium (Mg^{2+}), and sulphate (SO_4^{2-}) ions [17,18]. By injecting brine water into the rock fracture, SO_4^{2-} ions are adsorbed onto the positive calcite surface. This decreases the rock positive charges, which also reduces the electrostatic repulsion, thereby allowing cation ions to move closer to the surface. The presence of cation ions will bind with the negatively charged carboxylic acids and remove them from the rock surface. As the temperature increases, the concentration of Ca^{2+} and SO_4^{2-} near the surface increases, increasing the efficiency of the displacement of carboxylic groups. The removal of adsorbed carboxylic groups from the rock surface reduces the capillary forces, which improves oil recovery [19].

The main objective of this research was to assess the alteration in wettability of modified oil-wet calcium carbonate and IFT reduction under the influence of bio-surfactants. Moreover, this study investigated the effectiveness of bio-surfactants in the wetting alteration process as well as the fluid/fluid interfacial tension in the presence of most active sulphate ions at elevated temperature.

2. Materials and Methods

2.1. Chemicals, Solvents, and Bio-Surfactants

All the salt chemicals (CaCO_3 , NaCl , Na_2SO_4 , KCl , NaHCO_3 , KBr , MgF_2 , CaCl_2 , and SrCl_2), n-decane, stearic acid and calcite powder were supplied by Sigma-Aldrich. Stearic acid was used as a surface-active additive in n-decane to produce model oil resembling crude oil. Calcite powder with 99% purity was used to represent the carbonate rock reservoirs. Water soluble rhamnolipids, the bio-surfactant dominant form of dirhamnolipid, provided by the TeeGene Company (Teesside, UK), was added to the water solutions to study its application in EOR [20]. Different water compositions in the presence and absence of bio-surfactants were used as washing solutions to investigate their ability to wash out the adsorbed stearic acid from calcite surface. Table 1 presents the compositions of the washing solutions.

Table 1. Composition of washing solutions. Distilled water was considered as a reference and called WS_1 . WS_5 solution was made by adding 73 mg/L of the bio-surfactant to distilled water [20]. The base for the rest of the solutions was artificial sea water [21].

Additives to Distilled Water	WS_2 g/kg of Distilled Water	WS_3 g/kg of Distilled Water	WS_4 g/kg of Distilled Water	WS_5 g/kg of Distilled Water	WS_6 g/kg of Distilled Water	WS_7 g/kg of Distilled Water	WS_8 g/kg of Distilled Water
Cl^-	19.353	19.353	19.353	0	19.353	19.353	19.353
Na^+	10.765	10.765	10.765	0	10.765	10.765	10.765
SO_4^{2-}	2.711	5.422	8.133	0	2.711	5.422	8.133
Mg^{2+}	1.295	1.295	1.295	0	1.295	1.295	1.295
Ca^{2+}	0.414	0.414	0.414	0	0.414	0.414	0.414
K^+	0.387	0.387	0.387	0	0.387	0.387	0.387
HCO_3^-	0.142	0.142	0.142	0	0.142	0.142	0.142
Br^-	0.066	0.066	0.066	0	0.066	0.066	0.066
Sr^{2+}	0.008	0.008	0.008	0	0.008	0.008	0.008
H_3BO_3	0.026	0.026	0.026	0	0.026	0.026	0.026
F^-	0.001	0.001	0.001	0	0.001	0.001	0.001
Added							
Bio-Surfactant—Rhamnolipids	0	0	0	0.073 mL	0.073 mL	0.073 mL	0.073 mL

2.2. Sample Preparations

2.2.1. Modification of Water-Wet Calcite to Oil-Wet Calcite

In order to produce an oil-wet calcite surface resembling oil reservoir conditions, calcite rock was treated with dissolved polar components in the model oil. To this end, first 500 mL of n-decane was added to 1.42 g of stearic acid to make a 0.01 M concentration. Afterwards, 80 mL of oil model was mixed with 20 g of calcite powder within a conical flask. Then, the conical flask was placed into a hot plate stirrer at a speed of 50 rpm/min stirring for 24 h to collect the filtered liquid. Finally, calcite powder was transferred into a beaker oven at a temperature of 40 °C for 24 h to dry and solidify.

2.2.2. Wash Solution Preparation

To determine the effect of active ions such as SO_4^{2-} as well as the bio-surfactant (rhamnolipids) on oil-wet calcite, different forms of wash solutions (WS) were made to wash the modified calcite, which is listed in Table 2, with distilled water as a reference WS_1 . The compositions are presented in Table 1.

Table 2. Wash solution label description.

Wash Solution Label	Wash Solution Name
WS ₁	Distilled water
WS ₂	Seawater (composition found in Kester, et al., 1967 [21])
WS ₃	Seawater with 2X SO ₄ ²⁻
WS ₄	Seawater with 3X SO ₄ ²⁻
WS ₅	Distilled water + Bio-surfactant
WS ₆	Seawater + Bio-surfactant
WS ₇	Seawater with 2X SO ₄ ²⁻ + Bio-surfactant
WS ₈	Seawater with 3X SO ₄ ²⁻ + Bio-surfactant

2.2.3. Washing Process of Modified Calcite

In order to determine the effect of the saline solution and the bio-surfactant on the calcite surface, a washing process of the modified calcite was performed. First, 4 g of modified calcite and 20 mL WS₁ were mixed into a beaker. Then, the beaker was placed on a hot plate stirrer at a speed of 50 rpm/min stirring for 48 h to collect the filtered liquid. Finally, the beaker was placed in a vacuum oven at a temperature of 40 °C for 24 h for the solid to completely dry. All these processes were done for the remaining of WS₂–WS₈ solution.

2.3. Contact Angle Measurement

To measure the contact angle, the calcite powder was crushed into fine granular forms and then pressed into small disks that provided the samples with a flat and smooth surface for accurate contact angle measurement when the water droplet could come into contact with the powder surface. Finally, the disk was placed on a Kruss DSA 100 Goniometer analyser (KRÜSS GmbH, Hamburg, Germany) with distilled water released on the surface to measure the contact angle at room temperature of modified calcite sample and samples treated (washed) with three different temperatures. Table 3 presents a list of selected samples with different wash solutions at three different temperatures for contact angle measurement.

Table 3. Selected treated samples with different wash solutions at three different temperatures for contact angles.

Wash Solution	Temperature	Label
WS ₁	25 °C	K ₁
WS ₂	25 °C	K ₂
WS ₅	25 °C	K ₃
WS ₆	25 °C	K ₄
WS ₃	25 °C	K ₅
WS ₄	25 °C	K ₆
WS ₁	50 °C 70 °C	K ₇ K ₈
WS ₂	50 °C 70 °C	K ₉ K ₁₀
WS ₅	50 °C 70 °C	K ₁₁ K ₁₂
WS ₇	25 °C 50 °C 70 °C	K ₁₃ K ₁₄ K ₁₅
WS ₈	25 °C 50 °C 70 °C	K ₁₆ K ₁₇ K ₁₈

2.4. Interfacial Tension Measurement

To carry out the required experiments for the IFT measurements, the Du Nouy ring method using a Kruss Digital Tensiometer Model K9 (KRÜSS GmbH, Hamburg, Germany) was utilised. First, 10 mL of model oil and 10 mL of wash solution were mixed into a beaker and then heated with a hot plate stirrer. After that, the mixtures of the model oil and washing solution were transferred into special glassware to measure the IFT (Table 4).

Table 4. Mixture label description.

Mixtures Label	Oil Mixture	Wash Solution	Temperature
M ₁	N-decane	WS ₁	25 °C
M ₂	Dissolved steric acid + n-decane	WS ₁	25 °C
M ₃	Dissolved steric acid + n-decane	WS ₂	25 °C
M ₄	Dissolved steric acid + n-decane	WS ₃	25 °C
M ₅	Dissolved steric acid + n-decane	WS ₄	25 °C
M ₆	Dissolved steric acid + n-decane	WS ₅	25 °C
M ₇	Dissolved steric acid + n-decane	WS ₆	25 °C
M ₈	Dissolved steric acid + n-decane	WS ₇	25 °C
M ₉	Dissolved steric acid + n-decane	WS ₈	25 °C
M ₁₀	Dissolved steric acid + n-decane	WS ₅	50 °C
M ₁₁	Dissolved steric acid + n-decane	WS ₇	50 °C
M ₁₂	Dissolved steric acid + n-decane	WS ₈	50 °C
M ₁₃	Dissolved steric acid + n-decane	WS ₅	70 °C
M ₁₄	Dissolved steric acid + n-decane	WS ₇	70 °C
M ₁₅	Dissolved steric acid + n-decane	WS ₈	70 °C

3. Results and Discussion

3.1. Contact Angle Measurement

Three water droplets were used and repeated on different sections of the disk surface with at least three readings taken for each droplet, which produced an accurate average contact angle for each sample. Table 5 shows the contact angle of a modified sample before any additive are used.

Table 5. Contact angle measurement on modified calcite before contact with any washing solutions.

#Attempt	Modified Calcite			Average Reading Contact Angle (°)	Wetting System
	Reading 1 (°)	Reading 2 (°)	Reading 3 (°)		
Water Droplet 1	108.6	108.3	109.4	108.8	Oil-Wet
Water Droplet 2	109.8	110.1	110.0	110.0	Oil-Wet
Water Droplet 3	107.1	106.9	107.4	107.1	Oil-Wet
Overall Average Water Droplet Contact Angle (°)				108.6	Oil-Wet

In this work, IFT measurement between the oil model and washing solution containing bio-surfactant was conducted to determine the CMC of the bio-surfactant (rhamnolipids) in distilled water and seawater where the average CMC was approximately 73 mg/L. This bio-surfactant concentration in the washing solution was used under different conditions to assess the impact.

Tables 6 and 7 show the contact angle results after washing with WS₁ and WS₅. The addition of bio-surfactant into washing solutions reduced the contact angle from 108.6° to 73.4°, altering the surface to weakly water-wet. However, washing the oil-wet calcite with distilled water did not show a significant impact, as seen with WS₁, where the contact angle reduced from 108.6° to 94°, altering the surface to neutral-wet.

Table 6. Contact angle measurements on a calcite surface washed by WS₁.

WS ₁					
#Attempt	Reading 1 (°)	Reading 2 (°)	Reading 3 (°)	Average Reading Contact Angle (°)	Wetting System
Water Droplet 1	91.7	91.9	92.1	91.9	Neutral-Wet
Water Droplet 2	97.8	98.4	97.2	97.8	Neutral-Wet
Water Droplet 3	92.6	92.3	91.9	92.3	Neutral-Wet
Overall Average Water Droplet Contact Angle (°)				94	Neutral-Wet

Table 7. Contact angle measurements on a calcite surface washed by WS₅.

WS ₅					
#Attempt	Reading 1 (°)	Reading 2 (°)	Reading 3 (°)	Average Reading Contact Angle (°)	Wetting System
Water Droplet 1	70.5	69.3	69.9	69.9	Slightly Water-Wet
Water Droplet 2	76.5	76.0	75.5	76.0	Slightly Water-Wet
Water Droplet 3	74.2	73.7	74.6	74.2	Slightly Water-Wet
Overall Average Water Droplet Contact Angle (°)				73.4	Slightly Water-Wet

Changing the washing solution from distilled water to artificial seawater (WS₂) had a significant impact on the wettability of the modified calcite, where the measured average contact angle was reduced from 108.6° to 61.1°. It is worth mentioning that the addition of bio-surfactant to WS₂ did not show a significant impact on wettability, with the contact angle only reduced to 58.4°. This means that the presence of salt ions in the water hinders the effect of the bio-surfactant. The summary of all contact angle measurements with their corresponding labels are provided in Table 8.

Table 8. Measured contact angles for all treated calcite surfaces.

Label	Contact Angle (°)	Wetting System
K ₁	94.0	Neutral-wet
K ₂	73.3	Water-wet
K ₃	61.1	Water-wet
K ₄	58.4	Water-wet
K ₅	54.6	Water-wet
K ₆	41.3	Water-wet
K ₇	91.1	Neutral-wet
K ₈	100.3	Slightly oil-wet
K ₉	75.9	Water-wet
K ₁₀	85.7	Slightly water-wet
K ₁₁	91.8	Neutral-wet
K ₁₂	102.4	Slightly Oil-wet
K ₁₃	52.0	Water-wet
K ₁₄	62.6	Water-wet
K ₁₅	81.7	Slightly water-wet
K ₁₆	35.5	Water-wet
K ₁₇	41.9	Water-wet
K ₁₈	49.0	Water-wet

3.1.1. The Effect of the Bio-Surfactant and Salinity on the Contact Angle at 25 °C

The obtained data confirm that rhamnolipids altered the wetting state from oil-wet to water-wet. Figure 1 illustrates the effectiveness of the added rhamnolipids with different salt concentrations at a contact angle of 25 °C. It can be also seen from Figure 1 that the modification of the salt composition with affective ions such as SO₄²⁻ caused a further reduction in the contact angle, with the contact

angle reduced to 54.6°. The addition of rhamnolipids to modified seawater with three times SO₄²⁻ converted the calcite surface to water-wet with a measured contact angle of 41.3°. Interestingly, the rhamnolipids had a small impact on the contact angle when the salt ions took part in a wettability alteration. This could be due to the rhamnolipids reaching the CMC, therefore no additional reduction was demonstrated in the obtained results.

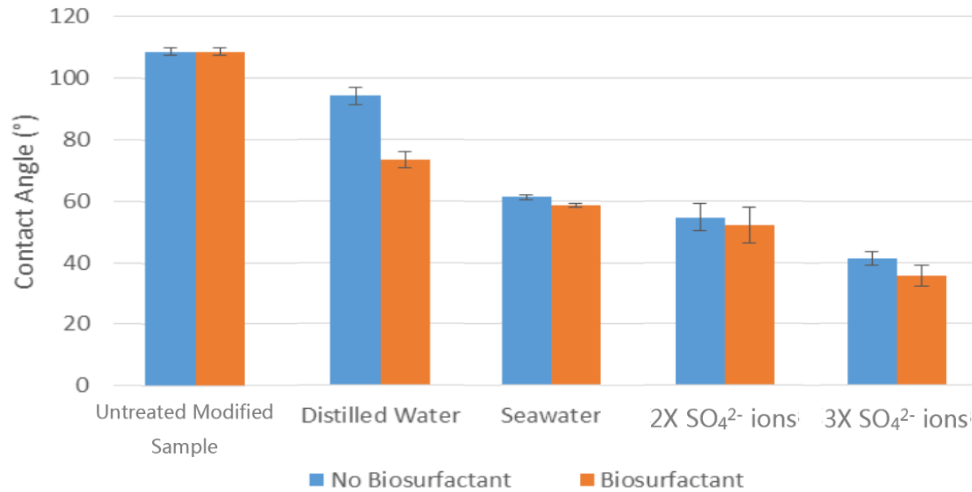


Figure 1. Contact angle measurements for the untreated modified calcite sample and samples treated with different water solutions in the presence and absence of bio-surfactant at 25 °C.

3.1.2. Effect of Elevated Temperature on Bio-Surfactant Behaviour

In this section, the effect of temperature on the effectiveness of the bio-surfactant in changing the wetting state of the modified calcite was investigated. The temperature was increased during the washing process to simulate high temperature reservoirs, obtaining the results displayed in Figure 2. In general, it can be seen from Figure 2 that the increase in temperature has a negative impact on the contact angle measurements when the system remains oil-wet.

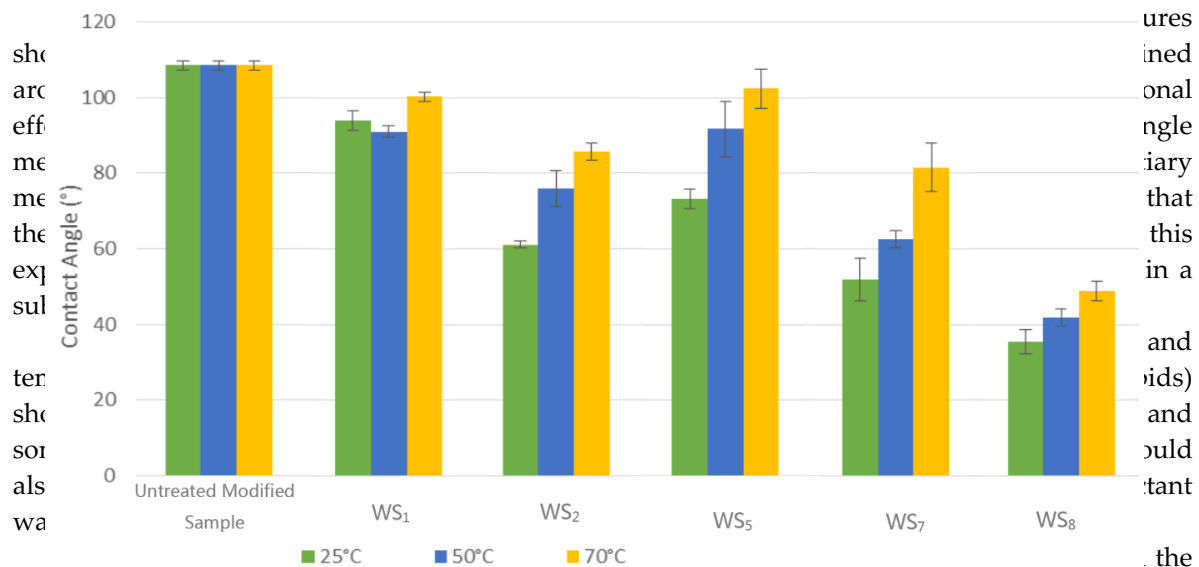


Figure 2. Contact angle measurement at elevated temperature for the untreated modified calcite sample and samples treated with different water solutions.

Industrial observation of the washing process using distilled water at different temperatures showed that the contact angle measurement at 70 °C was negligible, as the measurement remained in the oil-wet state. To understand the interaction of different ions and bio-surfactants in washing solutions, the interfacial activities of all the selected solutions were measured. Table 9 shows the reference point of pure n-decane with distilled water and the reduction in IFT when stearic acid was dissolved in n-decane.

Table 9. IFT measurements for reference solutions.

around 91.1° to 100.3°. However, distilled water mixed with bio-surfactant emphasised the additional effect on the wettability. Furthermore, other washing solutions showed a high contact angle measurement at 70 °C, which is inconsistent with the thermal recovery theory used in the tertiary method using external heat to raise the temperature and recover oil residuals. It has been stated that the thermal recovery temperature range can go above 350 °C [22]. Only increasing this to 70 °C in this experiment means that we cannot fully assume that a contact angle above 70 °C will result in a subsequent trend of the contact angle remaining oil-wet as the temperature increases.

According to Pacwa-Płociniczak et al. [6], bio-surfactants are active at extreme pH, salinity, and temperature levels. In contrast, based on Figure 2, it can be seen that the bio-surfactant (rhamnolipids) shows less effectiveness at high temperatures, this pattern might only occur with rhamnolipids and some other specific bio-surfactants. The minor change to water-wet in terms of the wettability could also be a result of the bio-surfactant being degradable, while at high temperatures the bio-surfactant was broken down, therefore not affecting wettability.

Additionally, as it has been shown in Figure 2, the increase in SO_4^{2-} concentration altered the wettability. In fact, by increasing the SO_4^{2-} concentration, a drop pattern in the contact angle measurement was observed. Looking at different concentrations of SO_4^{2-} at increasing temperatures showed a different consistent outline of wettability remaining for oil-water. Strand et al. [23], conducted experiments on the effect of different concentrations of Ca^{2+} and SO_4^{2-} brine composition mixtures using the oil-wet chalk core flooding method at high temperatures. Their results showed that SO_4^{2-} absorption onto chalk surfaces increased as the temperature and concentration of Ca^{2+} increased, which is similar to the pattern shown in Figure 2.

3.2. Interfacial Tension Measurement

To understand the interaction of different ions and bio-surfactants in washing solutions, the interfacial activities of all the selected solutions were measured. Table 9 shows the reference point of pure n-decane with distilled water and the reduction in IFT when stearic acid was dissolved in n-decane.

Table 9. IFT measurements for reference solutions.

Mixture Label	Reading 1 (mN/m)	Reading 2 (mN/m)	Reading 3 (mN/m)	Average Reading IFT (mN/m)
M ₁	27.1	27.0	27.1	27.1
M ₂	9.3	9.6	10.0	9.6

Similarly to the IFT measurement for the reference case, the IFT for all selected solutions labelled in Table 4 were measured and the results are provided in Table 10.

Table 10. Measured IFT between the model oil and all washing solutions.

Mixtures Label	IFT (mN/m)
M ₁	27.1
M ₂	9.6
M ₃	5.9
M ₄	5.4
M ₅	2.9
M ₆	4.6
M ₇	4.0
M ₈	0.5
M ₉	0.5
M ₁₀	2.7
M ₁₁	1.0
M ₁₂	1.0
M ₁₃	4.2
M ₁₄	0.7
M ₁₅	0.7

was reduced to 5.9 mN/m and 4 mN/m, respectively.

The biggest reduction in IFT was observed by increasing the SO_4^{2-} ions concentration. It should be noted that there was a huge difference between the IFT results with and without addition of the bio-surfactant in presence of SO_4^{2-} ions in water solutions. In fact, by adding bio-surfactant to 2X SO_4^{2-} ions, there was a significant decline of 4.4 mN/m in IFT from 5.9 to 0.5 mN/m. Similarly, it was noticeable that there was a gradual decrease in IFT for 3X SO_4^{2-} ions and bio-surfactant solution.

The IFT has a direct impact on the displacement processes and multiphase flow in porous media. Isehinwa and Olubukola [24] studied the effect of salinity on the IFT oil/brine system of five different Niger Delta reservoirs. Their results indicated a strong relationship between IFT and salinity. Figure 3 illustrates the effect of salinity on the IFT. The results showed that as the salinity increased, the IFT decreased. The highest reduction of IFT was observed when SO_4^{2-} ions added to the water solutions. As its concentrations were increased [25].

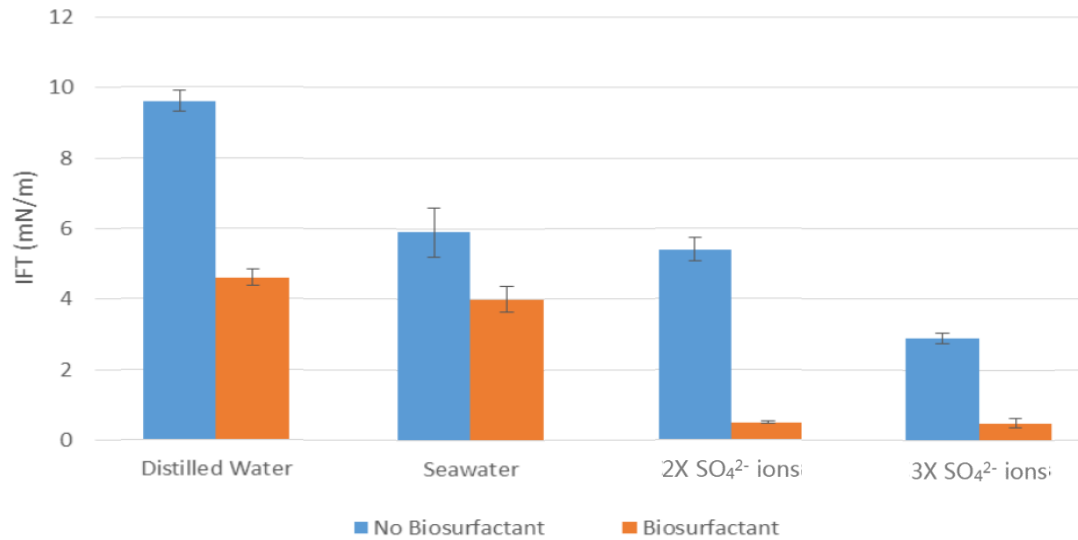


Figure 3. Measured IFT between model oil and different washing solutions at 25 °C.

Figure 3. Measured IFT between model oil and different washing solutions at 25 °C.

In terms of seawater with and without a bio-surfactant effect, the results showed that the IFT was reduced to 5.9 mN/m and 4 mN/m, respectively.

3.2.2. Effect of High Temperatures on the Interfacial Activities of Bio-Surfactants

The biggest reduction in IFT was observed by increasing the SO_4^{2-} ions concentration. It should be noted that there was a huge difference between the IFT results with and without addition of the bio-surfactant on IFT at high temperature. In terms of distilled water, first the IFT was decreased to 1.9 mN/m from 25 °C to 50 °C. Then, by increasing the temperature to 70 °C, an increase in the IFT was observed, showing that the mixture of distilled water and bio-surfactant did not significantly affect the reduction of IFT at high temperatures. However, Abouseoud et al. [26] investigated the bio-surfactant produced by *Pseudomonas fluorescens* on the surface tension, and their results suggested that when the temperature increased to 100 °C, there was no significant effect on the bio-surfactant's performance and it was still able to reduce surface tension. Varadavenkatesan and Murty [27] observed similar results.

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The highest reduction of IFT was observed when SO_4^{2-} ions added to the water solutions. As its concentrations were increased [25].

Figure 4 illustrates the effect of distilled water and different concentrations of SO_4^{2-} ions with bio-surfactant on IFT at high temperature. In terms of distilled water, first the IFT was decreased to 1.9 mN/m from 25 °C to 50 °C. Then, by increasing the temperature to 70 °C, an increase in the IFT was observed, showing that the mixture of distilled water and bio-surfactant did not significantly affect the reduction of IFT at high temperatures. However, Abouseoud et al. [26] investigated the bio-surfactant produced by *Pseudomonas fluorescens* on the surface tension, and their results suggested that when the temperature increased to 100 °C, there was no significant effect on the bio-surfactant's performance and it was still able to reduce surface tension. Varadavenkatesan and Murty [27] observed similar results.

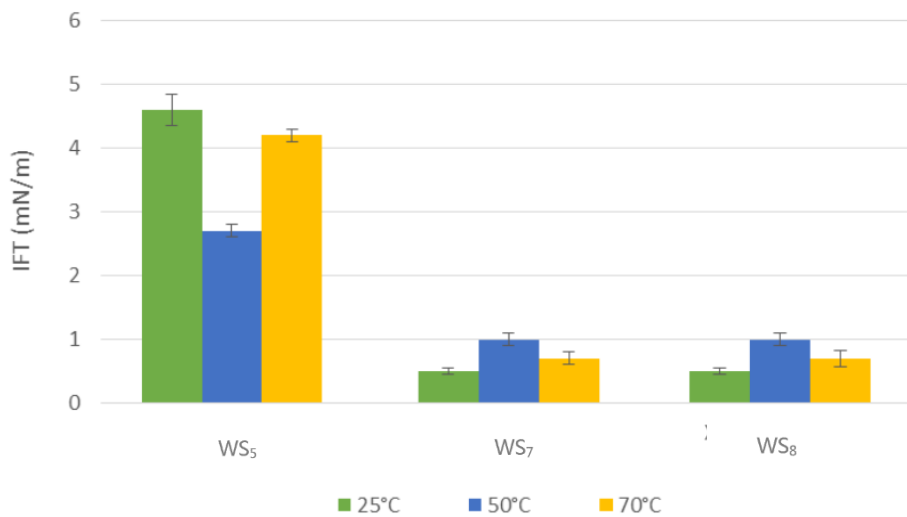


Figure 4. Measured IFT between model oil and washing solutions containing bio-surfactant at elevated temperature.

4. Conclusions The highest reduction of IFT was observed when SO_4^{2-} ions added to the water solutions. As is shown in Figure 4, a similar pattern was observed for different concentration of SO_4^{2-} on IFT from 25 to 70 °C. Many research studies have been carried out to find cheap, effective, and environmentally friendly methods to recover the remaining oil residual in the reservoirs. The objective of this research was to obtain more detailed information on the mechanism and effects of a bio-surfactant (rhamnolipids) on interfacial activities of oil/water and wettability of calcium carbonate as a substitute for carbonate rock in presence of salt and at elevated temperature. The results showed that rhamnolipids is capable to alter the wettability of carbonate rocks and reduce the IFT between oil and water. Rhamnolipids successfully improved the modified oil-wet rock to water-wet rock, and the observation of the IFT showed the existence of a large decrease even under the disturbance of different brine waters. The increase of SO_4^{2-} showed a positive impact on the rocks wettability and IFT, which also helps to lower the IFT and modify the rocks more towards water-wet. The resultant diagrams showed that presence of salt had a greater effect on the contact angle reduction, whereas rhamnolipids were better at reducing the IFT.

The results showed that increasing temperature conditions had undesired effects on the performance of the bio-surfactant. Even though addition of biosurfactant to water solutions under higher temperature showed the rock surface remaining oil-wet, but the salinity were not affected as the results still demonstrates a reduction in the contact angle and IFT measurement after increase in temperature.

The results showed that rhamnolipids is capable to alter the wettability of carbonate rocks and reduce the IFT between oil and water. Rhamnolipids successfully improved the modified oil-wet rock to water-wet rock, and the observation of the IFT showed the existence of a large decrease even under the disturbance of different brine waters. The increase of SO_4^{2-} showed a positive impact on the rocks wettability and IFT, which also helps to lower the IFT and modify the rocks more towards water-wet. The resultant diagrams showed that presence of salt had a greater effect on the contact angle reduction, whereas rhamnolipids were better at reducing the IFT.

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