

Thermal Properties of the Orthorhombic CaSnO_3 Perovskite under Pressure from *Ab initio* Quasi-Harmonic Calculations

J. Maul · I. M. G. Santos · J. R. Sambrano · A. Erba

November 9, 2015

Abstract Structural, elastic and thermodynamic properties of the orthorhombic CaSnO_3 perovskite are theoretically investigated at the *ab initio* level as a function of temperature and pressure. Harmonic and quasi-harmonic lattice dynamical calculations are performed with the CRYSTAL program, by explicitly accounting for thermal expansion effects and by exploring the effect of several DFT functionals. The anisotropic, directional elastic response of the system is characterized up to 20 GPa of pressure. The thermal lattice expansion and elastic bulk modulus are described at simultaneous temperatures up to 2000 K and pressures up to 20 GPa. The Gibbs free energy of formation of CaSnO_3 from CaO and SnO_2 as a function of temperature is also addressed by means of fully-converged phonon dispersion calculations on the three systems.

1 Introduction

Perovskites of the ABO_3 family are receiving special attention in recent years due to their peculiar physical and chemical properties, which lead to many possible applications [1–8]. Calcium stannate, CaSnO_3 , is known as an effective phosphorous host [9–20], for its potential as constituent of multifunctional optoelectronic devices [21], and for its unconventional luminescent behavior [22–24]. At ambient conditions, CaSnO_3 exhibits

a orthorhombic GdFeO_3 -type perovskite structure belonging to the $Pbnm$ space group, which is expected to be stable up to pressures of about 40 GPa and temperatures of about 2000 K [25]. Other phases have been reported (ilmenite-like, hexagonal, rhombohedral, tetragonal and cubic) [26–29], the orthorhombic one remaining the most studied [30–38]. Many fundamental properties of this material, however, still have to be reliably determined, which is the case of its thermal properties (its thermal structural and elastic features, its thermodynamic behavior, etc.). Furthermore, the evolution under increasing pressure of such thermal response properties is still not known, which would be of geochemical interest as CaSnO_3 is structurally and chemically closely related to CaSiO_3 and MgSiO_3 (two common constituents of the Earth mantle) [39,40,33,41]. In a recent study, we have investigated the effect of pressure on some vibrational spectroscopic features (Raman and IR spectra) of CaSnO_3 by means of *ab initio* calculations [42].

In the present study, a series of state-of-the-art *ab initio* techniques is used for the inclusion of temperature and pressure effects on several properties (structural, elastic, thermodynamic) of the orthorhombic phase of the CaSnO_3 perovskite. A recently developed fully-automated scheme, as implemented in a development version of the CRYSTAL14 program [43], is used for performing quasi-harmonic calculations [44,45], which allows to go beyond the standard harmonic approximation to the lattice potential and to effectively compute the directional thermal lattice expansion, temperature dependence of the bulk modulus, difference between constant-pressure and constant-volume specific heats, adiabatic bulk modulus, etc. [46–50] In this respect, it is worth mentioning that a recent experimental study by Redfern *et al.* [51] suggested CaSnO_3 to be highly quasi-

Laboratório de Combustíveis e Materiais, INCTMN-UFPB, Universidade Federal da Paraíba, CEP 58051-900, João Pessoa, PB, Brazil · Grupo de Modelagem e Simulação Molecular, INCTMN-UNESP, São Paulo State University, CEP 17033-360, Bauru, SP, Brazil · Dipartimento di Chimica, Università di Torino and NIS · Dipartimento di Chimica, Università di Torino and NIS, Nanostructured Interfaces and Surfaces, Centre of Excellence, Via Giuria 5, 10125 Torino, Italy
E-mail: : j.maul@quimica.ufpb.br

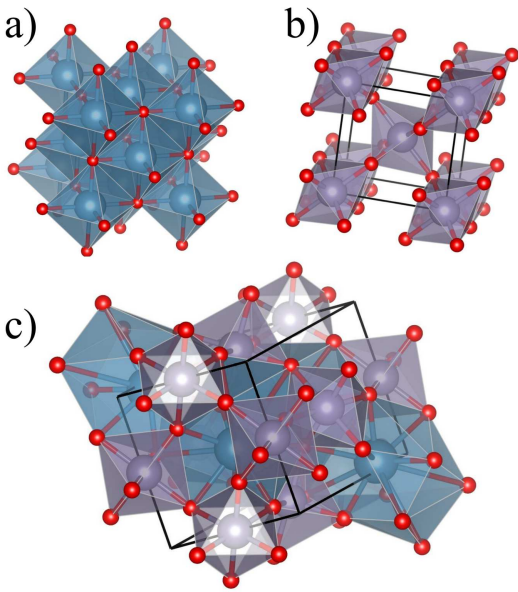


Fig. 1 (color online) Structure of a) cubic CaO ($Fm\bar{3}m$), b) tetragonal SnO₂ ($P4_2/mnm$) and c) orthorhombic CaSnO₃ ($Pbnm$). [59]

harmonic (without the need for an explicit anharmonic description of the system, then). The anisotropic elastic response of the system (in terms of elastic constants and directional seismic wave velocities) is investigated as a function of pressure [52–56]. Furthermore, the full phonon dispersion of the system is computed, which allows us to investigate other thermodynamic properties of the system and phonon density-of-states [57,58]. The Gibbs free energy of formation of CaSnO₃ from the two simple oxides CaO (cubic) and SnO₂ (tetragonal) is also evaluated as a function of temperature. The atomic structure of the three crystals is sketched in Figure 1.

The effect of the exchange-correlation functional of the density functional theory (DFT) on different properties (structural, thermal, elastic) is documented by considering different levels of approximation: from local-density, to generalized-gradient, to hybrids and range-separated hybrids.

2 Computational Methodology and Setup

All calculations are performed using a development version of the CRYSTAL14 program [43]. An all-electron atom-centered Gaussian-type-function (GTF) basis set is adopted, where Ca, Sn and O atoms are described by $8(s)6511(sp)21(d)$, $9(s)763111(sp)631(d)$ and $6(s)-2111(sp)1(d)$ contractions of primitive GTFs, respectively [42]. Seven different DFT functionals are adopted: SVWN [60,61], PBE [62], PBEsol [63], B3LYP [64,65], PBE0 [66], HSE06 [67] and HISS [68,69].

The tolerances for the evaluation of Coulomb and exchange infinite series are set to default values for all calculations. The shrinking factor is set to 8, which corresponds to a sampling over 125 independent \mathbf{k} -points within the first Brillouin zone in reciprocal space. The convergence on energy of the SCF procedure is set to the tight value of 10^{-10} hartree.

Phonon dispersion calculations are performed with supercells (SC) of different SC sizes within the PBE level of theory. For CaO, two SCs of 216 and 512 atoms are used, for SnO₂, two SCs containing 384 and 750 atoms and for CaSnO₃, SCs with 160 and 540 atoms. Quasi-harmonic calculations on CaO and SnO₂ have been performed with SCs of 54 and 48 atoms, respectively.

2.1 Quasi-harmonic Properties

A fully-automated scheme for computing quasi-harmonic properties of crystals has recently been implemented in the CRYSTAL program, which relies on computing and fitting (with a cubic polynomial function) harmonic vibration frequencies at different volumes after having performed volume-constrained geometry optimizations [46–50]. Harmonic phonon frequencies are computed by diagonalizing the dynamical matrix following a “direct space” approach [57,46,70,71]. Constant-volume specific heat, $C_V(T)$, and entropy, $S(T)$, are estimated through standard statistical thermodynamics from harmonic phonon frequencies as computed at the equilibrium zero temperature, zero pressure volume. Quasi-harmonic properties are computed by considering a volume range extending from a -3% compression to a +6% expansion with respect to the equilibrium volume; four equidistant volumes are considered in this interval.

The *isotropic* thermal expansion coefficient, $\alpha_V(T)$, of the system is obtained by minimizing the isothermal Helmholtz free energy

$$F^{\text{QHA}}(T, V) = U_0^{\text{ZP}}(V) + k_B T \sum_{\mathbf{k}p} \left[\ln \left(1 - e^{-\frac{\hbar\omega_{\mathbf{k}p}(V)}{k_B T}} \right) \right], \quad (1)$$

with respect to volume at several temperatures, where k_B is Boltzmann’s constant and $U_0^{\text{ZP}}(V)$ is the zero-temperature internal energy of the crystal which includes the zero-point energy of the system: $E_0^{\text{ZP}}(V) = \sum_{\mathbf{k}p} \hbar\omega_{\mathbf{k}p}(V)/2$. The *anisotropic* thermal expansion can be computed as well [48].

The quasi-harmonic approximation (QHA) allows for combining pressure and temperature effects on structural and elastic properties of materials. By differentiating equation (1) with respect to the volume and

changing sign, the thermal pressure is obtained:

$$P(V; T) = -\frac{\partial F^{\text{QHA}}(V; T)}{\partial V}. \quad (2)$$

The description of the isothermal bulk modulus of the system at simultaneous high-temperatures and high-pressures, $K_T(P, T)$, can be obtained as an isothermal second derivative of equation (1) with respect to the volume and by exploiting relation (2):

$$K_T(P, T) = V(P, T) \left(\frac{\partial^2 F^{\text{QHA}}(V(P, T); T)}{\partial V(P, T)^2} \right)_T. \quad (3)$$

2.2 Pressure-constrained Structure Optimization

A fully analytical scheme, based on the stress tensor, is used for optimizing the crystal volume under a given external pressure [72]. The stress tensor $\boldsymbol{\sigma}$ is a symmetric second-rank tensor that can be computed in terms of analytical energy gradients with respect to lattice parameters:

$$\sigma_{ij} = \frac{1}{V} \frac{\partial E}{\partial \epsilon_{ij}} = \frac{1}{V} \sum_{k=1}^3 \frac{\partial E}{\partial a'_{ki}} a_{kj}, \quad (4)$$

with $\boldsymbol{\epsilon}$ second-rank symmetric pure strain tensor and $i, j, k = x, y, z$. In the expression above, a_{ij} are elements of a 3×3 matrix, \mathbf{A} , where Cartesian components of the three lattice vectors \mathbf{a}_1 , \mathbf{a}_2 and \mathbf{a}_3 are inserted by rows and V is the cell volume. When a distortion is applied to the cell, the lattice parameters transform as $a'_{ij} = \sum_{k=1}^3 (\delta_{jk} + \epsilon_{jk}) a_{ik}$, where δ_{jk} is the Kronecker delta. By adding an external hydrostatic “pre-stress” $\sigma_{ij}^{\text{pre}} = P\delta_{ij}$ to σ_{ij} and by inverting equation (4), one gets the expression for the constrained gradients

$$\frac{\partial H}{\partial a_{ij}} = \frac{\partial E}{\partial a_{ij}} + PV(\mathbf{A}^{-1})_{ji}. \quad (5)$$

With the inclusion of a hydrostatic pressure, the function to be minimized becomes the enthalpy $H = E + PV$ [73].

2.3 Elastic Tensor Calculation

If any finite pre-stress is absent, second-order elastic constants are simply defined as second energy density derivatives with respect to pairs of infinitesimal Eulerian strains:

$$C_{ijkl} = \frac{1}{V_0} \left(\frac{\partial^2 E}{\partial \epsilon_{ij} \partial \epsilon_{kl}} \right)_{\epsilon=0}. \quad (6)$$

An automated scheme for the calculation of the elastic tensor has been implemented in the CRYSTAL program [74, 52], that has been generalized also to low-dimensionality 1D and 2D systems [75]. Applications of this scheme cover a wide range of materials [76–84].

When a finite pre-stress $\boldsymbol{\sigma}^{\text{pre}}$ is applied in the form of a hydrostatic pressure P , within the frame of finite Eulerian strain, the elastic stiffness constants become [85–89]:

$$B_{ijkl} = C_{ijkl} + \frac{P}{2} (2\delta_{ij}\delta_{kl} - \delta_{il}\delta_{jk} - \delta_{ik}\delta_{jl}), \quad (7)$$

provided that the volume V_0 at which equation (6) is evaluated is replaced by the equilibrium volume $V(P)$ at pressure P . A fully automated implementation in the CRYSTAL program of the calculation of the stiffness tensor \mathbf{B} (and of $\mathbf{S} = \mathbf{B}^{-1}$, the compliance tensor) under pressure has recently been presented [53, 54]. A two-index representation of the elastic stiffness tensor is obtained ($B_{ijkl} \rightarrow B_{vu}$) by exploiting Voigt’s notation, according to which $v, u = 1, \dots, 6$ ($1 = xx, 2 = yy, 3 = zz, 4 = yz, 5 = xz, 6 = xy$) [90]. For the elastic constant calculation, four strained configurations are considered for each independent strain, with a dimensionless strain amplitude of 0.0075 (*i.e.* 0.75 %).

3 Results and Discussion

3.1 Quasi-harmonic Description

The thermal expansion of CaSnO₃, as theoretically determined by means of quasi-harmonic lattice dynamical calculations is reported in Figure 2. The lower panel of the figure reports the equilibrium cell volume as a function of temperature, up to 1000 K, as obtained with seven different DFT functionals and as compared with available experimental data (empty symbols) [51, 91]. We can clearly see that while different functionals provide rather different descriptions of the absolute value of the equilibrium volume, they all describe a similar trend as it comes to the temperature dependence. As expected, the LDA functional underestimates the cell volume. Two functionals (the generalized-gradient PBE and the global hybrid B3LYP) significantly overestimate the volume while the other functionals provide a reasonable description of the equilibrium structure. The two experimental datasets reported in the figure show some large discrepancies between each other as regards the lattice expansion. Data by Chen *et al.* [91], which measure a lower expansion, are found to better agree with present quasi-harmonic simulations than those by Redfern *et al.* [51], where a much larger thermal expansion is reported.

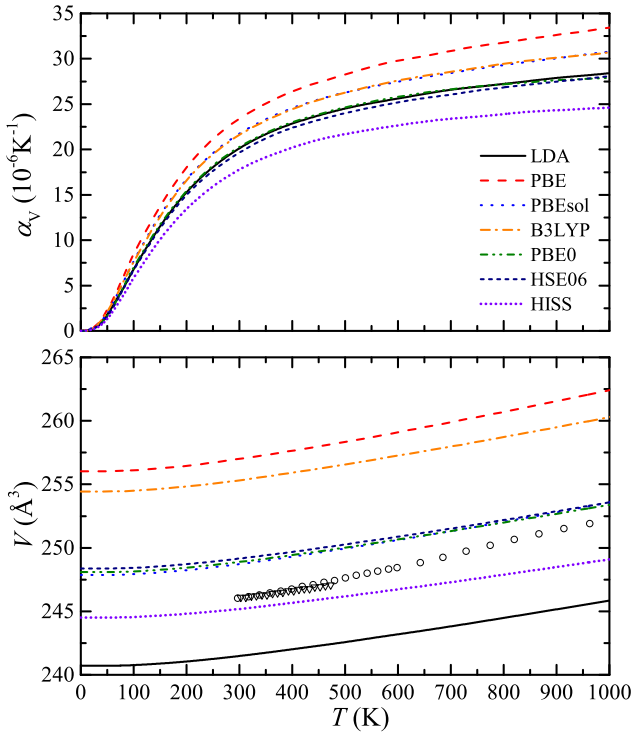


Fig. 2 (color online) Thermal expansion coefficient (upper panel) and temperature dependent cell volume (lower panel) of CaSnO_3 up to 1000 K, as computed with seven different DFT functionals. Empty circles and triangles are experimental data from Redfern *et al.* [51] and from Chen *et al.* [91], respectively.

In order to better highlight the different description of thermal expansion given by different functionals, we may refer to the upper panel of Figure 2, where the volumetric thermal expansion coefficient $\alpha_V(T) = 1/V(T)[\partial V(T)/\partial T]$ is reported. Despite all functionals provide a similar description of the lattice thermal expansion of CaSnO_3 , some fine differences can be noticed: the HISS functional shows the smallest thermal expansion while the PBE one the largest (followed by PBEsol and B3LYP, which give an almost identical description of α_V). The other three adopted functionals provide a rather similar description with each other.

The quasi-harmonic approximation represents an effective theoretical tool for describing structural and average elastic properties of solids at simultaneous high pressure and high temperature, that is at extreme conditions which are difficult to be explored experimentally. In Figure 3, we report the thermal evolution, up to 2000 K, of the thermal expansion coefficient, of the isothermal bulk modulus and of the three independent lattice parameters a , b and c , at five different pressures up to 20 GPa, as obtained with the PBE functional. From a computational point of view, the lattice dynamics of the system has been evaluated at 15 volumes from

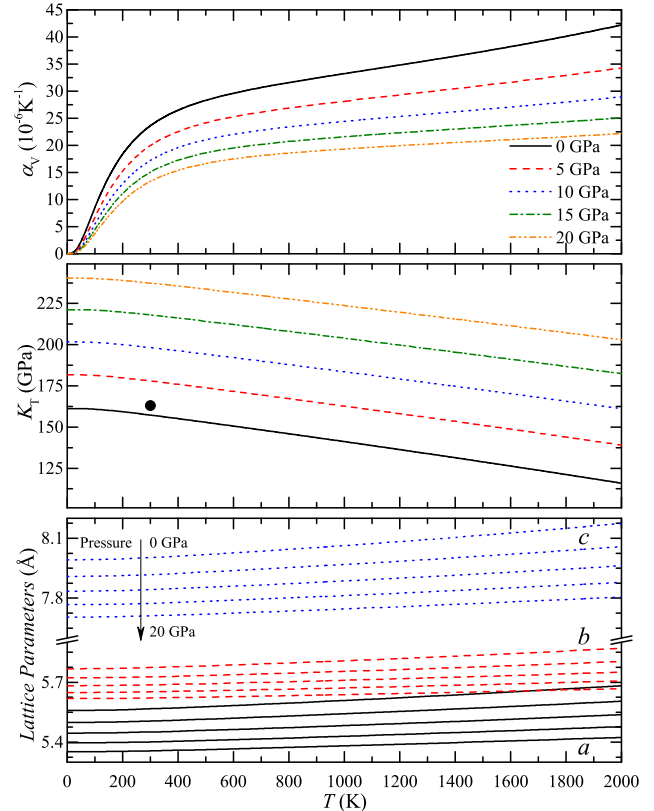


Fig. 3 (color online) Thermal expansion coefficient (upper panel), isothermal bulk modulus (middle panel) and lattice parameters (lower panel) of CaSnO_3 as a function of temperature, at 5 different pressures from 0 to 20 GPa. In the middle panel, the full circle represents the experimental value for the isothermal bulk modulus at room conditions [33]. All data are obtained with the PBE functional.

a -10% compression to a 10% expansion of the equilibrium zero-pressure, zero-temperature volume. From the upper panel of the figure, we can see that the thermal expansion coefficient behaves quite regularly up to 2000 K (with a tiny deviation from linearity), thus confirming the high quasi-harmonic character of CaSnO_3 . As pressure increases, the thermal expansion obviously decreases and the coefficient becomes even more regular and linear at high temperatures, as expected given that pressure tends to suppress explicit anharmonic terms in the lattice potential. The evolution of the isothermal bulk modulus as a function of pressure and temperature is quantitatively given in the middle panel of the figure, where the experimental value at room conditions is also given as a full circle [33]. In the lower panel of the figure, the directional, anisotropic, thermal expansion of orthorhombic CaSnO_3 is documented. The a and c lattice parameters are found to expand more than the b one by about a factor of 1.2.

Experimentally, a much larger factor of about 4.3 was suggested by Redfern *et al.* [51], which seems way

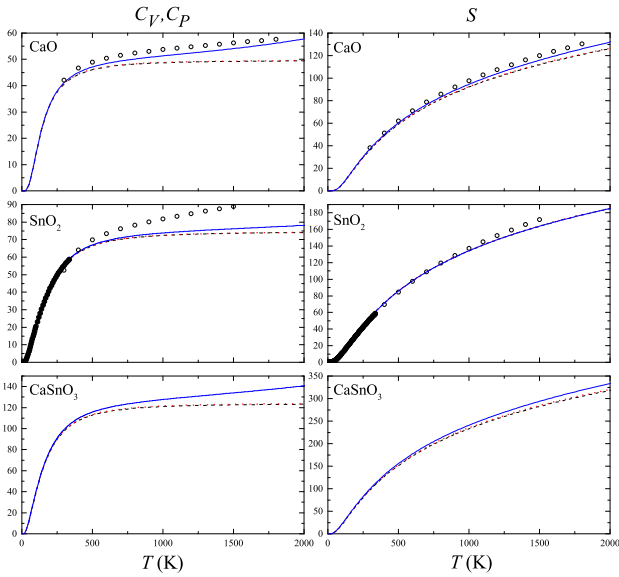
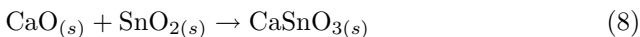


Fig. 4 (color online) Heat capacity (left panel) and entropy (right panel) of CaSnO₃, CaO and SnO₂. All data in J mol⁻¹K⁻¹. In the left panels, dashed and dotted lines refer to C_V and continuous lines to C_P , in the right ones, to S_V and S_P , respectively. Full circles are experimental data by Gurevich *et al.* [92] while empty ones by Robie [93].

too large for a system such as CaSnO₃, which is not very structurally anisotropic. This large experimental anisotropy of the thermal expansion is reflected in the large volumetric thermal expansion reported in Figure 2 (see empty circles).

3.2 Phonon Dispersion and Thermodynamics

Thermodynamic properties require a good description of phonon dispersion in order to be accurately described at the *ab initio* level. In this section, we report converged harmonic (constant-volume specific heat C_V and entropy S_V) and quasi-harmonic (constant-pressure specific heat C_P and entropy S_P) thermodynamic properties of CaSnO₃. Furthermore, thermodynamic functions of both CaO and SnO₂ are studied, with the purpose in mind to determine the formation free energy of the following solid state reaction [28,94]:



These thermodynamic functions are reported in Figure 4 for the three systems up to 2000 K of temperature. Harmonic properties are obtained with the largest super-cells considered (see Section 2 for details), which ensure full convergence as regards the sampling of phonon dispersion within the Brillouin zone. The quasi-harmonic constant-pressure specific heat is obtained by summing on top of the harmonic C_V , the term $C_P(T) - C_V(T) =$

Table 1 Gibbs free energy of formation of CaSnO₃ from the CaO and SnO₂ oxides, as determined with the PBE functional.

T (K)	ΔG (kJ mol ⁻¹)
0	-58.65
300	-59.98
1000	-65.32
2000	-73.19

$\alpha_V^2(T)K_T(T)V(T)T$, as derived from the quasi-harmonic quantities given in Figures 2 and 3. Experimental data from Gurevich *et al.* [92] (full circles) and from Robie [93] (empty circles) are also reported, when available. We can see that our calculations provide a satisfactory description of the entropy (a key ingredient to the study of the formation free energy), apart from a slight underestimation at very high temperatures. Specific heat of CaO is also reasonably described whereas for SnO₂ we get a poor description of the constant-pressure specific heat, due to a large underestimation of the thermal expansion (the quasi-harmonic approximation seems to fail rather dramatically for this system). Absolute values of C_V and C_P at room temperature (in units of J/(mol K)) are: 40.77 and 41.18 for CaO; 55.92 and 56.06 for SnO₂, and 98.24 and 99.26 for CaSnO₃, respectively.

The formation Gibbs’s free energy of the solid state reaction (8) is given in Table 1 at few different temperatures. The CaSnO₃ perovskite is found to be thermodynamically more stable than the two isolated oxides at all temperatures, which is consistent with the experimentally observed “irreversibility” of CaSnO₃ once formed from both oxides or molten salts [28,94,34,95,96,16,97].

From the full phonon dispersion of a system, the phonon density-of-state (PDOS) can be determined, which is defined by the equation [58]:

$$g(\omega) = \frac{1}{V_{\text{BZ}}} \int_{\text{BZ}} \sum_{p=1}^{3N} \delta(\omega_{\mathbf{k}p} - \omega) d\mathbf{k}, \quad (9)$$

where V_{BZ} is the volume of the Brillouin zone and the integration is performed over it. From equation (9), it is seen that the PDOS is normalized to $3N$, being N the number of atoms per cell ($\int g(\omega) d\omega = 3N$). The total PDOS can be partitioned into atomic contributions $g(\omega) = \sum_a g_a(\omega)x_a$ where the sum runs over the atomic species of the system, x_a is the fraction of atomic

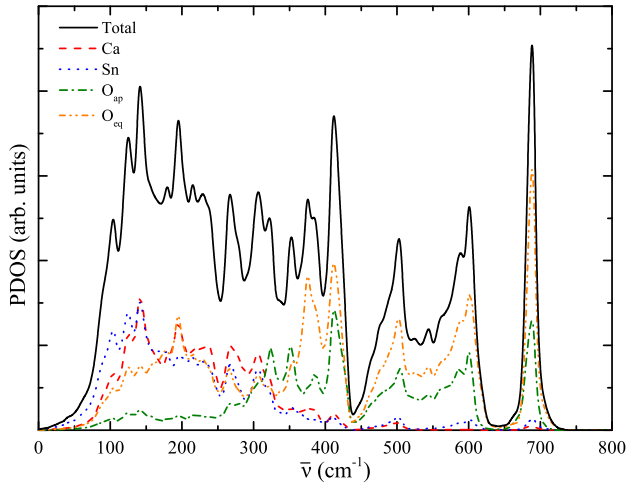


Fig. 5 (color online) Total and atomic phonon density-of-states of CaSnO_3 .

species a , and

$$g_a(\omega) = \frac{1}{n_{\mathbf{k}}} \sum_{p,\mathbf{k}} |\mathbf{e}_{p,\mathbf{k};a}|^2 \delta(\omega_{\mathbf{k}p} - \omega), \quad (10)$$

where $\mathbf{e}_{p,\mathbf{k}}$ are the eigenvectors of the dynamical matrices $W^{\mathbf{k}}$ and the integral in equation (9) has been replaced by the sum over the $n_{\mathbf{k}}$ sampled \mathbf{k} -points within the first Brillouin zone. Atomic PDOSs are useful tools for discussing which atoms are mostly involved in phonons of given spectral regions. Figure 5 reports the total PDOS (black continuous line) as well as the atomic PDOSs of Ca, Sn and the two symmetry-independent oxygen atoms, apical O_{ap} and equatorial O_{eq} . The PDOS of CaSnO_3 clearly shows three distinct regions: a broad one ranging from 0 up to about 440 cm^{-1} , a second one from about 450 to about 625 cm^{-1} and a third sharp one at about 700 cm^{-1} . The optical vibrational spectrum above approximately 350 cm^{-1} is almost entirely dominated by motions of the oxygen atoms (see yellow and green lines). At lower wave-numbers, all atoms are involved in the lattice vibrations, Ca and Sn atoms becoming dominant in very soft collective phonon modes.

3.3 Elastic Properties

This section is devoted to the discussion of the anisotropic elastic response of CaSnO_3 , at ambient conditions and under pressure. The fourth rank elastic tensor of the orthorhombic CaSnO_3 perovskite has 9 symmetry-independent elastic constants C_{vu} , which are reported in Table 2 as computed with several DFT functionals. All functionals describe similarly the anisotropy of the elastic response, given that the ratios among the elastic constants are similar in all cases. On the contrary, absolute

Table 2 Single-crystal elastic constants C_{vu} CaSnO_3 as computed with seven DFT functionals at 0 K. The VRH average bulk modulus \bar{K} at 0 K and the isothermal K_T and adiabatic K_S bulk moduli at 300 K of temperature are also reported. All data in GPa.

	LDA	PBE	PBEsol	B3LYP	PBE0	HSE06	HISS	Exp. [33]
C_{11}	284.6	260.9	269.1	288.2	295.9	293.7	318.3	
C_{22}	351.1	299.5	328.0	325.0	343.8	341.2	366.6	
C_{33}	317.0	281.1	300.0	305.9	324.0	320.6	349.1	
C_{44}	100.1	90.5	95.8	99.3	104.6	104.0	112.9	
C_{55}	77.8	74.3	76.1	83.9	86.8	86.1	94.4	
C_{66}	103.0	90.6	97.7	97.7	105.9	105.1	114.2	
C_{12}	150.7	124.6	136.0	127.9	138.7	139.9	147.2	
C_{13}	121.2	100.6	110.9	107.3	114.3	113.1	121.0	
C_{23}	120.7	100.7	108.9	106.3	111.8	113.0	118.4	
\bar{K}	192.3	165.6	178.1	177.7	187.8	187.1	200.5	
K_S^{300}	185.6	159.4	172.1	171.0	180.4	180.6	194.9	167.2
K_T^{300}	183.9	157.7	170.4	169.3	178.7	179.0	193.4	162.6

values differ rather significantly. Single-crystal elastic constants of CaSnO_3 have not yet been experimentally measured so that the performance of the different functionals can be evaluated just on average elastic properties. Following the Voigt-Reuss-Hill (VRH) averaging scheme [98], average elastic properties of isotropic polycrystalline aggregates can be derived starting from computed elastic constants. The average bulk modulus \bar{K} obtained in this way is given in the table. Computed elastic properties refer to 0 K and neglect even zero-point motion effects. From the quasi-harmonic calculations discussed before, however, the isothermal bulk modulus K_T can be computed as a function of temperature. Some experimental techniques for measuring the bulk modulus of a crystal involve elastic waves and are characterized by very short time-scales that prevent the system from reaching a thermal equilibrium; in these cases, an adiabatic bulk modulus, K_S , is measured instead. Adiabatic and isothermal bulk moduli do coincide with each other at zero temperature only, K_S always being larger than K_T at any finite temperature. The quasi-harmonic approximation also offers a way to compute the adiabatic bulk modulus from the isothermal one, given that:

$$K_S = K_T + \frac{\alpha_V^2 V T K_T^2}{C_V} = K_T \times \frac{C_P}{C_V}, \quad (11)$$

where the dependence of all quantities on temperature is just omitted for clarity sake. In Table 2, both the isothermal and adiabatic bulk moduli are reported at room temperature and compared with available experimental data. From that comparison, we see that the HISS functional provides the most rigid description of the system, followed by LDA and both PBE0 and HSE06. On the other hand, PBE, PBEsol and B3LYP provide a softer description of the system, which is seen

Table 3 Elastic stiffness constants B_{vu} (in GPa) of CaSnO₃ under pressure up to 20 GPa, as computed with the PBE functional. Average elastic properties are also reported such as bulk modulus \overline{K} , shear modulus \overline{G} , Young modulus \overline{E} (in GPa), Poisson's ratio σ (dimensionless) and longitudinal \overline{v}_p and transverse \overline{v}_s seismic wave velocities (in km/s).

Pressure	0 GPa	5 GPa	10 GPa	15 GPa	20 GPa
B_{11}	260.9	276.1	292.2	308.9	322.3
B_{22}	299.5	335.1	363.0	390.7	418.2
B_{33}	281.1	308.6	332.6	356.2	380.2
B_{44}	90.5	96.4	101.5	105.6	109.8
B_{55}	74.3	77.4	79.9	81.8	83.7
B_{66}	90.6	98.0	103.4	108.5	114.2
B_{12}	124.6	143.4	166.0	185.8	206.8
B_{13}	100.6	118.3	134.5	150.4	165.8
B_{23}	100.7	115.9	133.1	148.2	163.4
\overline{K}	165.6	185.5	205.2	223.8	242.1
\overline{G}	85.0	89.8	92.9	96.0	98.8
\overline{E}	217.7	231.9	242.1	251.9	260.9
σ	0.281	0.292	0.303	0.312	0.320
\overline{v}_p	7.181	7.407	7.594	7.764	7.924
\overline{v}_s	3.960	4.015	4.040	4.063	4.087

to be much closer to the experimental elastic rigidity of CaSnO₃.

The effect of an applied hydrostatic pressure on the elastic properties of CaSnO₃ is documented in Table 3, where elastic stiffness constants B_{vu} (computed as described in Section 2.3) are reported along with other average elastic properties, such as the average bulk \overline{K} and shear \overline{G} moduli obtained within the VRH scheme. From the bulk modulus and the average shear modulus, Young's modulus \overline{E} and Poisson's ratio σ can be defined as well:

$$\overline{E} = \frac{9\overline{K}\overline{G}}{3\overline{K} + \overline{G}} \quad \text{and} \quad \sigma = \frac{3\overline{K} - 2\overline{G}}{2(3\overline{K} + \overline{G})}. \quad (12)$$

Average values of transverse (shear) and longitudinal seismic wave velocities, for an isotropic polycrystalline aggregate, can also be computed from the elastic constants and the density ρ of the crystal as:

$$\overline{v}_s = \sqrt{\frac{\overline{G}}{\rho}} \quad \text{and} \quad \overline{v}_p = \sqrt{\frac{\overline{K} + 4/3\overline{G}}{\rho}}. \quad (13)$$

As expected, all elastic constants do increase as pressure increases. The diagonal elastic constants B_{22} and B_{33} are the ones that increase faster with pressure with slopes of 5.9 and 4.9, respectively, followed by B_{12} with a slope of 4.1. Shear constants display a small dependence on pressure, B_{55} being the less affected by pressure with a slope of just 0.5. Transverse seismic wave velocities are not much affected by pressure (with a

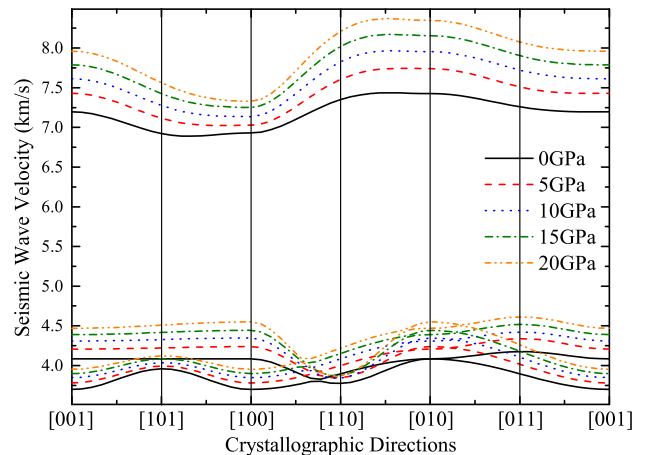


Fig. 6 (color online) Directional seismic wave velocities of CaSnO₃ under pressures up to 20 GPa, as computed with the PBE functional.

derivative of 0.006 km/s/GPa), at variance with longitudinal ones (with 0.037 km/s/GPa); these values are in qualitative agreement with experimental determinations by Schneider *et al.* [99] of 0.012 and 0.052 km/s/GPa, respectively.

The propagation of elastic waves in a crystal is an anisotropic physical phenomenon, which, once the fourth-rank elastic tensor has been determined, can be fully-characterized by solving Christoffel's equation [100, 101]. Directional seismic wave velocities (a quasi-longitudinal v_p one and two quasi-transverse v_{s1} and v_{s2} ones) of CaSnO₃, as computed at different pressures in the present study, are reported in Figure 6. It is seen that the directions of minimum propagation velocity of longitudinal waves are between the [101] and [001] ones, whereas along directions from [110] and [010] longitudinal waves propagate with the largest velocity. Apart from an obvious increase in the average propagation velocity, the effect of pressure on the anisotropic propagation of elastic waves in calcium stannate is found to be relatively small, with just a slight increase of the anisotropy of longitudinal waves.

4 Conclusions

State-of-the-art techniques for the *ab initio* investigation of solids have been applied to predict structural, elastic and thermodynamic properties of the orthorhombic CaSnO₃ perovskite at high temperatures and pressures, which are not yet experimentally known. Harmonic and quasi-harmonic calculations, as combined with advanced algorithms for the calculation of anisotropic elastic properties, do indeed constitute a powerful tool for the quantum-chemical description of solids

at P and T conditions of real interest in fields such as material science and geochemistry.

The effect on computed properties of the DFT functional is explicitly discussed. Functionals belonging to four families of the DFT (local-density, generalized-gradient, hybrid and range-separated hybrid approximations) are considered. If large differences on absolute values of computed properties are found, most properties are seen to be almost independent of the adopted functional as regards their behavior as a function of temperature or pressure. The generalized-gradient PBE and PBEsol, and the global hybrid B3LYP functionals are found to provide the best description of the elastic features of the system.

Acknowledgements J.M. acknowledges the Brazilian scholarship program “Ciência sem Fronteiras” (Process Number 248425/2013-7/SWE). Furthermore, we are grateful for the programs PROCAD 2013/Proc. 88881.068492/2014-01 and FAPESP (2013/19289-0, 2013/07296-2) and also to the Center for Scientific Computing of the São Paulo State University (GridUNESP).

References

1. M.A. Peña, J.L.G. Fierro, *Chem. Rev.* **101**(7), 1981 (2001)
2. S. Royer, D. Duprez, F. Can, X. Courtois, C. Batiot-Dupeyrat, S. Laassiri, H. Alamdari, *Chem. Rev.* **114**(20), 10292 (2014)
3. A. Kubacka, M. Fernández-García, G. Colón, *Chem. Rev.* **112**(3), 1555 (2012)
4. R. Thalinger, A.K. Opitz, S. Kogler, M. Heggen, D. Stroppa, D. Schmidmair, R. Tappert, J. Fleig, B. Klötzer, S. Penner, *J. Phys. Chem. C* **119**(21), 11739 (2015)
5. L. An, H. Onishi, *ACS Catal.* **5**(6), 3196 (2015)
6. E.N. Armstrong, K.L. Duncan, E.D. Wachsman, *Phys. Chem. Chem. Phys.* **15**, 2298 (2013)
7. Y. Yang, J. Íñiguez, A.J. Mao, L. Bellaiche, *Phys. Rev. Lett.* **112**, 057202 (2014)
8. M. Misono, in *Heterogeneous Catalysis of Mixed Oxides Perovskite and Heteropoly Catalysts, Studies in Surface Science and Catalysis*, vol. 176, ed. by M. Misono (Elsevier, 2013), pp. 67 – 95
9. Z. Liu, Y. Liu, *Mater. Chem. Phys.* **93**(1), 129 (2005)
10. B. Lei, B. Li, H. Zhang, W. Li, *Opt. Mater.* **29**(11), 1491 (2007)
11. B. Lei, B. Li, H. Zhang, L. Zhang, Y. Cong, W. Li, *J. Electrochem. Soc.* **154**(7), H623 (2007)
12. Z. Fu, W. Li, S. Du, H.K. Yang, J.H. Jeong, *J. Electrochem. Soc.* **156**(10), J308 (2009)
13. K. Ueda, Y. Shimizu, *Thin Solid Films* **518**(11), 3063 (2010)
14. X.Y. Chen, C. Ma, S.P. Bao, H.Y. Zhang, *J. Alloys Compd.* **497**(12), 354 (2010)
15. X.L. Pang, C.H. Jia, G.Q. Li, W.F. Zhang, *Opt. Mater.* **34**(1), 234 (2011)
16. T. Nakamura, M. Shima, M. Yasukawa, K. Ueda, *J. Sol-Gel Sci. Technol.* **61**(2), 362 (2012)
17. Z. Liang, J. Zhang, J. Sun, X. Li, L. Cheng, H. Zhong, S. Fu, Y. Tian, B. Chen, *Physica B* **412**(0), 36 (2013)
18. J. Zhang, B. Chen, Z. Liang, X. Li, J. Sun, L. Cheng, H. Zhong, *J. Alloy. Compd.* **612**(0), 204 (2014)
19. S.D. Kim, K.S. Hwang, S. Hwangbo, *Electron. Mater.* **9**(4), 405 (2013)
20. K.S. Hwang, Y.S. Jeoni, S. Hwangbo, J.T. Kim, *Ceram. Int.* **39**(7), 8555 (2013)
21. K. Ueda, T. Maeda, K. Nakayashiki, K. Goto, Y. Nakachi, H. Takashima, K. Nomura, K. Kajihara, H. Hosono, *Appl. Phys. Express* **1**(1), 015003 (2008)
22. Y. Karabulut, M. Ayvacikli, A. Canimoglu, J. Garcia Guinea, Z. Kotan, E. Ekdal, O. Akyuz, N. Can, *Spectroscopy Letters* **47**(8), 630 (2014)
23. V. Orsi Gordo, Y. Tuncer Arslanli, A. Canimoglu, M. Ayvacikli, Y. Galvão Gobato, M. Henini, N. Can, *Appl. Radiat. Isotopes* **99**(0), 69 (2015)
24. A. Canimoglu, J. Garcia-Guinea, Y. Karabulut, M. Ayvacikli, A. Jorge, N. Can, *Appl. Radiat. Isotopes* **99**(0), 138 (2015)
25. S. Tateno, K. Hirose, N. Sata, Y. Ohishi, *Phys. Earth Planet. Inter.* **181**(12), 54 (2010)
26. B. Durand, H. Loiseau, *J. Appl. Crystallogr.* **11**(4), 289 (1978)
27. R. Ropp, in *Encyclopedia of the Alkaline Earth Compounds*, ed. by R. Ropp (Elsevier, Amsterdam, 2013), pp. 351 – 480
28. W.W. Coffeen, *J. Am. Ceram. Soc.* **36**(7), 207 (1953)
29. G. Hautier, A. Miglio, D. Waroquiers, G.M. Rignanese, X. Gonze, *Chem. Mater.* **26**(19), 5447 (2014)
30. H. Mizoguchi, H.W. Eng, P.M. Woodward, *Inorg. Chem.* **43**(5), 1667 (2004)
31. M. Tarrida, H. Larguem, M. Madon, *Phys. Chem. Miner.* **36**(7), 403 (2009)
32. A. Yangthaisong, *Chin. Phys. Lett.* **30**(7), 077101 (2013)
33. J. Kung, R.J. Angel, N.L. Ross, *Phys. Chem. Miner.* **28**(1), 35 (2001)
34. J. Kung, Y.J. Lin, C.M. Lin, *J. Chem. Phys.* **135**(22), 224507 (2011)
35. P. McMillan, N. Ross, *Phys. Chem. Miner.* **16**(1), 21 (1988)
36. A. Vegas, M. Vallet-Regí, J.M. González-Calbet, M.A. Alario-Franco, *Acta Crystallogr., Sect. B: Struct. Sci* **42**(2), 167 (1986)
37. D. Cherrad, D. Maouche, M. Boudissa, L. Reffas, M. and Louail, M. Maamache, K. Haddadi, Y. Medkour, *Physica B* **429**, 95 (2013)
38. E. Moreira, C.A. Barboza, E.L. Albuquerque, U.L. Fulco, J.M. Henriques, A.I. Araiújo, *J. Phys. Chem. Solids* **77**(0), 85 (2015)
39. R. Caracas, R.E. Cohen, *Geophys. Res. Lett.* **32**(6), L06303 (2005)
40. O. Kuskov, R. Galimzyanov, in *Chemistry and Physics of Terrestrial Planets, Advances in Physical Geochemistry*, vol. 6, ed. by S. Saxena (Springer New York, 1986), pp. 310–361
41. R.C. Liebermann, L.E.A. Jones, A.E. Ringwood, *Phys. Earth Planet. Inter.* **14**(2), 165 (1977)
42. J. Maul, A. Erba, I.M.G. Santos, J.R. Sambrano, R. Dovesi, *J. Chem. Phys.* **142**(1), 014505 (2015)
43. R. Dovesi, R. Orlando, A. Erba, C.M. Zicovich-Wilson, B. Civalleri, S. Casassa, L. Maschio, M. Ferrabone, M. De La Pierre, Ph. D’Arco, Y. Noël, M. Causá, M. Rérat, B. Kirtman, *Int. J. Quantum Chem.* **114**, 1287 (2014)
44. R.E. Allen, F.W. De Wette, *Phys. Rev.* **179**, 873 (1969)
45. L.L. Boyer, *Phys. Rev. Lett.* **42**, 584 (1979)

46. A. Erba, *J. Chem. Phys.* **141**, 124115 (2014)
47. A. Erba, M. Shahrokhi, R. Moradian, R. Dovesi, *J. Chem. Phys.* **142**, 044114 (2015)
48. A. Erba, J. Maul, R. Demichelis, R. Dovesi, *Phys. Chem. Chem. Phys.* **17**, 11670 (2015)
49. A. Erba, J. Maul, M. De La Pierre, R. Dovesi, *J. Chem. Phys.* **142**, 204502 (2015)
50. A. Erba, J. Maul, M. Itou, R. Dovesi, Y. Sakurai, *Phys. Rev. Lett.* **115**, 117402 (2015)
51. S.A.T. Redfern, C. Chen, J. Kung, O. Chaix-Pluchery, J. Kreisel, E.K.H. Salje, *J. Phys.: Condens. Matter* **23**(42), 425401 (2011)
52. A. Erba, A. Mahmoud, R. Orlando, R. Dovesi, *Phys. Chem. Miner.* **41**, 151 (2014)
53. A. Erba, A. Mahmoud, D. Belmonte, R. Dovesi, *J. Chem. Phys.* **140**, 124703 (2014)
54. A. Mahmoud, A. Erba, K. Doll, R. Dovesi, *J. Chem. Phys.* **140**, 234703 (2014)
55. V. Lacivita, A. Erba, R. Dovesi, Ph. D'Arco, *Phys. Chem. Chem. Phys.* **16**, 15331 (2014)
56. A. Erba, A.M. Navarrete-López, V. Lacivita, P. D'Arco, C.M. Zicovich-Wilson, *Phys. Chem. Chem. Phys.* **17**, 2660 (2015)
57. A. Erba, M. Ferrabone, R. Orlando, R. Dovesi, *J. Comput. Chem.* **34**, 346 (2013)
58. J. Baima, M. Ferrabone, R. Orlando, A. Erba, R. Dovesi, *Phys. Chem. Minerals* (2015). DOI 10.1007/s00269-015-0781-6
59. K. Momma, F. Izumi, *J. Appl. Crystallogr.* **44**(6), 1272 (2011)
60. P.A.M. Dirac, *Proc. Cambridge Phil. Soc.* **26**, 376 (1930)
61. S.H. Vosko, L. Wilk, M. Nusair, *Can. J. Phys.* **58**, 1200 (1980)
62. J.P. Perdew, K. Burke, M. Ernzerhof, *Phys. Rev. Lett.* **77**, 3865 (1996)
63. J.P. Perdew, A. Ruzsinszky, G.I. Csonka, O.A. Vydrov, G.E. Scuseria, L.A. Constantin, X. Zhou, K. Burke, *Phys. Rev. Lett.* **100**, 136406 (2008)
64. A.D. Becke, *J. Chem. Phys.* **98**, 5648 (1993)
65. C. Lee, W. Yang, R.G. Parr, *Phys. Rev. B* **37**, 785 (1988)
66. C. Adamo, V. Barone, *J. Chem. Phys.* **110**, 6158 (1999)
67. A. V. Krukau, O. A. Vydrov, A. F. Izmaylov and G. E. Scuseria, *J. Chem. Phys.* **125**(22), 224106 (2006)
68. T. M. Henderson, A. F. Izmaylov, G. E. Scuseria and A. Savin, *J. Chem. Phys.* **127**(22), 221103 (2007)
69. T. M. Henderson, A. F. Izmaylov, G. E. Scuseria and A. Savin, *J. Chem. Theory Comput.* **4**(8), 1254 (2008)
70. K. Parlinski, Z.Q. Li, Y. Kawazoe, *Phys. Rev. Lett.* **78**, 4063 (1997)
71. A. Togo, F. Oba, I. Tanaka, *Phys. Rev. B* **78**, 134106 (2008)
72. K. Doll, *Molecular Physics* **108**(3-4), 223 (2010)
73. I. Souza, J. Martins, *Phys. Rev. B* **55**, 8733 (1997)
74. W.F. Perger, J. Criswell, B. Civalleri, R. Dovesi, *Comput. Phys. Commun.* **180**, 1753 (2009)
75. A. Erba, M. Ferrabone, J. Baima, R. Orlando, M. Rérat, R. Dovesi, *J. Chem. Phys.* **138**, 054906 (2013)
76. A. Erba, K.E. El-Kelany, M. Ferrero, I. Baraille, M. Rérat, *Phys. Rev. B* **88**, 035102 (2013)
77. A. Mahmoud, A. Erba, K.E. El-Kelany, M. Rérat, R. Orlando, *Phys. Rev. B* **89**, 045103 (2014)
78. A. Erba, R. Dovesi, *Phys. Rev. B* **88**, 045121 (2013)
79. K.E. El-Kelany, A. Erba, P. Carbonnière, M. Rérat, *J. Phys.: Cond. Matter* **26**, 205401 (2014)
80. J. Baima, A. Erba, R. Orlando, M. Rérat, R. Dovesi, *J. Phys. Chem. C* **117**, 12864 (2013)
81. V. Lacivita, A. Erba, Y. Noël, R. Orlando, Ph. D'Arco, R. Dovesi, *J. Chem. Phys.* **138**, 214706 (2013)
82. J.C. Tan, B. Civalleri, A. Erba, E. Albanese, *CrystEngComm* **17**, 375 (2015)
83. K.E. El-Kelany, P. Carbonnière, A. Erba, M. Rérat, *J. Phys. Chem. C* **119**, 8966 (2015)
84. A. Erba, M.T. Ruggiero, T.M. Korter, R. Dovesi, *J. Chem. Phys.* **143**, 144504 (2015)
85. B.B. Karki, G.J. Ackland, J. Crain, *J. Phys.: Cond. Matter* **9**(41), 8579 (1997)
86. J. Wang, J. Li, S. Yip, S. Phillpot, D. Wolf, *Phys. Rev. B* **52**, 12627 (1995)
87. B.B. Karki, L. Stixrude, R.M. Wentzcovitch, *Rev. Geophys.* **39**, 507 (2001)
88. D.C. Wallace, *Thermodynamics of Crystals* (Wiley, New York, USA, 1972)
89. D.C. Wallace, *Rev. Mod. Phys.* **37**, 57 (1965)
90. J.F. Nye, *Physical properties of crystals* (Oxford University Press, Oxford, 1957)
91. C.J. Chen, J. Kung, C.M. Lin, M. Zhang, S.A.T. Redfern, ArXiv e-prints (2010)
92. V.M. Gurevich, K.S. Gavrichev, V.B. Polyakov, R.N. Clayton, S.D. Mineev, G. Hu, V.E. Gorbunov, L.N. Golushina, *Thermochim. Acta* **421**(12), 179 (2004)
93. R.A. Robie, B.S. Hemingway, J.R. Fisher, Thermodynamic properties of minerals and related substances at 298.15 k and 1 bar (105 pascals) pressure and at higher temperatures. Tech. Rep. B - 1452, United States Geological Survey (1978)
94. Y. Tanaka, *Bull. Chem. Soc. Jpn.* **17**(2), 70 (1942)
95. Z. Lu, J. Liu, Y. Tang, Y. Li, *Inorg. Chem. Commun.* **7**(6), 731 (2004)
96. J. Zhao, N.L. Ross, R.J. Angel, *Phys. Chem. Miner.* **31**(5), 299 (2004)
97. S. Shojaei, S.A. Hassanzadeh-Tabrizi, M. Ghashang, *Ceram. Int.* **40**(7, Part A), 9609 (2014)
98. R. Hill, *J. Mech. Phys. Solids* **11**, 357 (1963)
99. B.W. Schneider, W. Liu, B. Li, *High Pressure Res.* **28**(3), 397 (2008)
100. M.J.P. Musgrave, *Crystal Acoustics* (Holden-Day, San Francisco, California, 1970)
101. B.A. Auld, *Acoustic Fields and Waves in Solids* (Krieger Publishing Company, Malabar, Florida, 1973)