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Leaching potential of metallic elements from contaminated soils under anoxia

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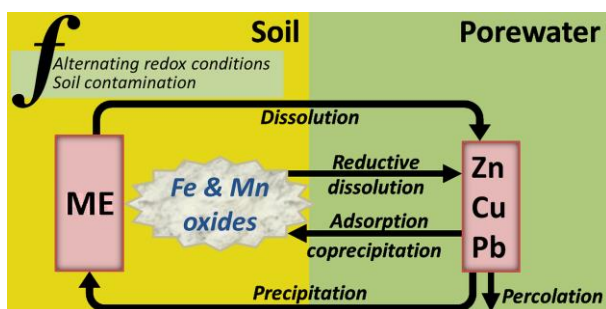
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Environmental Impact

The higher frequency and intensity of extreme rainfall events, as well as the variations in precipitation patterns and intensities as a result of climate change may cause important shifts in soil hydrological regime. This could expose contaminated soils to significant variations in redox potentials possibly influencing metallic element behaviour and their potential transfer to other environmental compartments. This study shows that alternating soil redox conditions may enhance the release and redistribution of metals in more labile fractions. In particular, the presence of redox-sensitive minerals, such as chalcopyrite, in soils affected by mine tailings, could release significant amounts of Cu due to the alternating redox conditions and therefore have important environmental implications.

Graphical abstract



The alternation of redox cycles may play a higher role in the release, leaching and redistribution of metallic elements from contaminated soils with respect to oxidizing and reducing conditions alone.

Abstract

Understanding metallic element (ME) behaviour in soils subjected to alternating redox conditions is of significant environmental importance, particularly for contaminated soils. Although variations in the hydrological status of soils may lead to the release of ME, redox-driven changes in ME dynamics are still not sufficiently understood. We studied the effects of alternating redox cycles

on the release, leaching and redistribution of Zn, Cu and Pb in metal mine-contaminated and non-contaminated soils by means of a column experiment. Although the release of Zn was promoted by the onset of reductive conditions, successive redox cycles favoured metal partitioning in less labile fractions limiting its further mobilization. The release of Cu in soil porewaters and redistribution in the solid phase towards more labile pools were strongly dependent on the alternation between oxidizing and reducing conditions. In contaminated soils, the presence of chalcopyrite could have determined the release of Cu under oxic conditions and its relative immobilization under subsequent anoxic conditions. The behaviour of Pb did not seem to be influenced by redox status, although higher concentrations in the column leachates with respect to soil porewaters suggested that alternating redox conditions could nonetheless result in substantial mobilization. This study provides evidence that the alternation of soil redox conditions may play a more important role in determining the release and leaching of metallic elements from soils with respect to reducing or oxidizing conditions considered separately.

Introduction

The negative impacts of climate change on human and environmental health have increased in many parts of the world, as weather-related hazards (floods, droughts and extreme temperatures) have become more frequent and more acute. The higher frequency and intensity of extreme rainfall events, as well as the variations in precipitation patterns and intensities - e.g. wetter winters with more frequent and more severe storm events¹ - may cause changes in soil hydrological regime. In extreme cases, oxic topsoils may become anoxic due to water saturation for a more or less prolonged period of time. At the onset of anaerobiosis a series of reactions occur that involve electron acceptors other than oxygen. When Mn and Fe are reduced, they become several orders of magnitude more soluble than their oxidized compounds and may rapidly release the elements that are associated to them to the soil solution.^{2,3}

As soils generally serve as a sink for numerous contaminants like pesticides, dioxins, agricultural chemicals, and metallic elements (ME), submergence may result in the release of these pollutants from the soil to other environmental compartments.⁴ While the effect of water saturation on the release of ME from riverine, agricultural and urban soils is well documented,⁵⁻⁹ less information is available about mine and smelter-affected areas.^{10,11}

Intermittent flooding/drying and the consequent reductive and oxidative conditions may affect the chemical speciation, bioavailability, toxicity and mobility of metals in the soil.^{2,3} It is generally

agreed that the release of ME into solution can be influenced by wet/dry cycles,^{12,13} but ME mobilization can also be counteracted by their stabilization due to other processes that take place under reducing conditions. Contradictory information regarding the effect of intermittent submergence on the behaviour of ME in pore water has been reported. In acid soils under anoxic conditions, ME can be reabsorbed on soil particles due to the increase in pH¹⁴ and, in the case of highly reduced soils, metals can be immobilized by precipitation as sulphides and carbonates.¹⁵ Soil redox oscillations may also increase the crystallinity of Fe oxides enhancing the fixation of some ME into their structure.¹⁶ Contrasting results were also obtained for the partitioning of ME between phases with variable lability, in flooded soils. While Kashem and Singh¹⁷ reported that the addition of organic matter (OM) favoured the partitioning of ME in more labile pools, Huang *et al.*¹⁸ showed that OM amendment of submerged soils caused a decrease in oxidizable, and Fe and Mn oxide-bound Cu leading to a relative increase in more stable forms.

This study aims at elucidating the influence of redox oscillations on the release and mobility of ME and the changes in their chemical forms in mine-impacted soils. As mining activities frequently take place in mountainous river valleys where the probability of flash floods is higher than in plains, it is necessary to understand ME behaviour in these soils in order to evaluate the possible effects of climate change on the environmental hazards posed by ME.

We hypothesised that soil metals released into solution under reducing conditions may be subsequently translocated during water drainage and potentially leached from the soil and/or redistributed in the more labile metal pools. To test this hypothesis we evaluated changes in Cu, Pb, and Zn concentrations in soil pore waters and leachates from soil columns subjected to a series of oxic/anoxic redox cycles. Metal distribution before and after the incubation, was evaluated by sequential extraction.

Materials and methods

Soil Sampling

Contaminated soil samples were collected from the river valley located in Bălan, Harghita county, Romania (Fig. 1). Mixed Cu-Pb-Zn deposits have been mined in this area and during the long period of mining activity, the atmosphere, waters and soils have been contaminated with ME.¹⁹ Soils were sampled from areas chosen at increasing distance from the main contamination source, the preparation plant where Cu flotation had been carried out, to compare different degrees of contamination. Each sample was composed of a series of subsamples collected over an area of 400

m² and adequately homogenized. The soil subsamples of a highly contaminated area (HCN) were collected approximately 150 m uphill from the contamination source, while a less contaminated area (LCN) was sampled about 5 km downhill from the preparation plant. A non-contaminated soil (NCN), developing on a similar bedrock, was sampled from an area near the town of Izvorul Mureşului, approximately 8 km SW of the source of contamination. This sample was taken as reference soil representing background values for ME concentrations. All samples were air-dried, gently crushed to pass through a 2 mm sieve. A small portion of each sample was further ground to < 0.5 mm for pseudo-total metal analysis.

General Soil Properties

Physical and chemical soil properties were determined using standard methods.²⁰ Soil pH was measured potentiometrically in 1M KCl using a 1:2.5 soil:solution ratio. Particle-size analysis of the soil material (< 2 mm) was conducted by the sieve-pipette method after dispersion with sodium hexametaphosphate. Total C and N contents of the bulk soil were measured on ground subsamples (< 0.5 mm) by dry combustion (NA2100, CE Instruments, Milan, Italy). Since the soil material was free of carbonates, total C was considered as organic C. Pseudo-total metal content was determined by microwave-assisted *aqua regia* (HCl:HNO₃, 3:1 solution) digestion (MILESTONE Start D). Metallic element extraction with dithionite-citrate-bicarbonate (DCB) and ammonium oxalate (AO) were performed as described by Mehra and Jackson.²¹ Although these extractions were originally intended for the determination of Fe oxides, they can nevertheless be used to evaluate metal contents associated with oxides and OM (e.g. adsorbed, co-precipitated, complexed) and their potential release from soils under anoxic conditions.^{22,23} All extracts were analysed in duplicate for Cu, Fe, and Mn concentration by flame atomic absorption spectrometry (FAAS, Perkin-Elmer 400).

Column Experiment

Each soil was packed (in duplicate) into PVC columns 70 cm in height and 7 cm in diameter as sketched in Fig. 2. Soil columns were exposed to a series of four 7-day anoxic-oxic cycles. Each cycle was composed of 5 days anoxic conditions during which the soils were oversaturated with a 10mM CaCl₂ + 3mM lactose solution, followed by column drainage for 2 days. Prior to column leaching (i.e. on the 5th day of each anoxic period), soil pore waters were extracted by Rhizon samplers installed at a depth of 10 and 30 cm (UR and LR, respectively) from the soil surface, while soil leachates were collected from the bottom of the column during drainage. More details of the column setup and soil packaging have been provided elsewhere.²⁴ Total metal (Cu, Pb, and Zn)

contents in the solutions were analyzed by FAAS.

After four cycles of flooding and draining the soils were carefully extracted from the columns, and three 10 cm soil sections sampled at depths corresponding to the upper (U) and lower (L) Rhizons, and bottom of the column (B). All soils were dried at 40°C prior to further analysis.

Metal Fractionation

Metal fractionation on all soil samples was performed in triplicate using the modified BCR sequential extraction procedure as described by Rauret *et al.*²⁵ This method was used to extract and quantify metals that are contained in fractions generally described as exchangeable (F1), bound to Fe–Mn oxides (F2), bound to organic matter (F3) and residual (F4). Total Cu, Pb and Zn in each extract was determined by FAAS. Recoveries were calculated as the ratio of the sum of ME contents in each BCR fraction to pseudo-total ME concentrations determined on the bulk soil. In all soils, Zn, Cu and Pb recoveries ranged between 83-103, 74-101 and 75-101% of the pseudo-total concentrations in the initial soils, respectively. The influence of experimental conditions on the partitioning of ME was evaluated by comparing with ME distribution in the initial soils (I) not subjected to the redox cycling.

Results and discussion

Soil Properties

Soil texture, pH, C/N ratio, and pseudo-total Cu, Pb, and Zn concentrations are reported in Table 1. The soils are acid with a rather coarse texture that could favour ME loss from the soil to the water table. In the studied area, the main sources of Cu, Pb, and Zn are chalcopyrite (CuFeS₂), galena (PbS), and sphalerite ((Zn,Fe)S), respectively.²⁶ Therefore, due to past exploration and exploitation works, elements from these minerals were released into the environment.²⁷ The soils sampled from the mining areas (HCN and LCN) had relatively higher amounts of all metals, in particular Cu, with respect to the soil from the reference area (NCN). The total concentration of Cu, Pb, and Zn (Table 1), however, was not indicative of significant contamination. The ME contents in the contaminated soils exceeded the Romanian legislative recommended concentrations (Order no. 756/1997), but generally did not exceed the threshold alert or intervention values.

Metallic elements are known to be associated to Fe and Mn oxides by binding or coprecipitation.^{28,29} This has been especially observed in Fe and Mn nodules which form in wetland soils under alternating redox conditions.^{30,31} Therefore, DCB and AO soil extracts do not

only estimate the contents of crystalline and amorphous Fe/Mn oxides in the soils, respectively²¹ but can give an appraisal of the amount of oxide-associated ME.²³

Table 2 reports the amounts of ME presumably associated with Fe/Mn oxides and potentially subject to release during their reductive dissolution. A significant proportion of total Cu present in contaminated soils ($\approx 54\%$) was extractable in AO, suggesting that Cu is predominantly bound to short-range aluminosilicates and organo-mineral complexes.³² In contrast, only 35% of total Cu was AO-extractable in NCN suggesting that the potential for metal mobilization due to reductive dissolution of amorphous Fe and Mn oxides or OM oxidation under redox conditions is lower in non-contaminated soils. The higher contents of AO- with respect to DCB-extractable Cu in all soils suggest that this element is not entirely associated to the mineral phases solubilized by reductive dissolution. However, we cannot exclude that some Cu might be present in mineral forms that are readily mobilized at the pH of AO extraction (pH 3).

Most of the Zn ($\approx 52\%$) and some of the Pb (11-20%) contained in HCN and LCN was extractable in DCB, though the proportion of these metals also extractable in AO was limited. In fact, only about 9-16% of Zn was AO-extractable while the concentration of Pb was below detection limit. In NCN the proportion of Zn in the DCB extract was slightly greater with respect to the contaminated soils, thus suggesting the distribution of ME in different pools may be more indicative of their potential release from soil subjected to alternating redox conditions, with respect to total concentrations. These results also suggest that Zn may be mainly associated with crystalline oxides of Fe and Mn as reported by Van Laer *et al.*³³ and Tack *et al.*³⁴

Column Experiments

The mobility and transport of ME in soils was assessed as a function of alternating redox cycles by sampling and analysing soil pore waters and leachates. With successive redox cycles mean pH values of soil leachates tended to increase in HCN, LCN, and NCN from 6.2, 5.8, and 6.9 respectively, to 7.1, indicative of the typical increase in soil pH under reductive conditions. The column setup used did not allow for measurement of redox potentials during the experiment. However, the increase in Fe and Mn concentrations in soil porewaters as a result of the reductive dissolution of Fe and Mn (hydr)oxides was used as a proxy for reducing conditions, as previously discussed in Balint *et al.*²⁴ Moreover, we evaluated the change in soil Eh under anoxic conditions in a preliminary batch experiment. Under non-limiting electron donor conditions (lactose addition), Eh rapidly decreased to values of -150, -189, and -216 mV within 2 days of anoxic incubation for HCN, LCN, and NCN respectively.

Significant amounts of Zn were readily released in solution during the first redox cycle and continued to increase reaching highest concentrations during the second reductive phase for HCN and LCN (Fig. 3). The significant amounts of Zn associated with Fe and Mn oxyhydroxides (Table 2) suggest that the increasing trend in porewater Zn concentrations under reducing conditions may be related to the reductive dissolution of these metal oxides. This is supported by our previous findings on the reductive dissolution of Mn and Fe oxyhydroxides from these soils²⁴ that showed a similar increase in porewater Mn followed by Fe solubilisation, in line with their different reduction potentials. In fact, the principal control mechanism for Zn in wetland soils is considered to be dependent on the hydrous Fe and Mn oxides.³¹ Van Laer *et al.*³³ also attributed the mobilization of Zn in waterlogged soils to the reductive dissolution of Fe and Mn oxides which may have occluded or sorbed ME. While some authors have shown that Zn is substantially fixed by soils rich in Fe oxides,³⁵ Sipos *et al.*³⁶ found that Zn is less immobilized by soil mineral phases with respect to Cu and Pb. In our experiment, after an initial increase, soil porewater Zn concentrations tended to decrease with each subsequent redox cycles (Fig. 3). Successive reductive dissolution of Fe and Mn oxyhydroxides could have resulted in the exposure of new sorption sites on the surface of soil minerals which may act as a sink for Zn.¹³ Coprecipitation of Zn with newly formed oxides under oxic conditions could also have limited the subsequent release of this ME during anoxic conditions. Similar results were obtained by Contin *et al.*¹⁶ who observed the stabilization of Zn and other ME in soils amended with Fe-rich waste subjected to similar redox alternations. The aging of Zn-containing coprecipitates could have further contributed to Zn stabilization as suggested by the findings of Davranche and Bollinger.²⁸ These processes could also explain the relatively low and decreasing concentrations of Zn in the oxic soil leachates with successive redox cycles (Fig. 3). Although NCN showed similar trends with respect to HCN and LCN, the extent of Zn release was less pronounced even though this soil had similar amounts of Zn associated with the reducible Fe and Mn oxides (Table 2). This could be due to the higher clay and silt content in NCN with respect to HCN and LCN (Table 1), which might have led to a more pronounced metal sorption under reducing conditions.

As for Zn, the release of Cu into soil solution was favoured by the onset of reductive conditions, with highest porewater concentrations observed during the first redox cycle in contaminated and, to a lesser extent, non-contaminated soils (Fig. 3). The limited release of Cu in NCN could be attributed to lower total (Table 1) and oxide-associated (Table 2) contents with respect to contaminated soils. With successive redox cycles porewater Cu concentrations tended to decrease

steadily reaching values close to or below detection limits in the fourth reductive phase. This may be due to the decreasing Cu solubility with increasing pH, as generally observed in acidic soils under reducing conditions. Moreover, recent studies by Biasioli *et al.*,¹² Fulda *et al.*³⁷ and Matocha *et al.*³⁸ have shown that, under anoxic conditions, Cu(II) can be reduced to Cu(I) or Cu(0) and subsequently stabilized as insoluble Cu(I)-humic acids complexes or precipitated as Cu₂O. Huang *et al.*¹⁸ presented similar results for OM-amended paddy soils. They attributed the decrease in Cu release to the degradation of dissolved organic carbon (DOC) with which Cu is usually complexed, or to the formation of less soluble Cu minerals. Moreover, in our experiment, the decrease in Cu release from soils during anoxic periods might have been due to the addition of lactose which served as the main electron donor and diminished the possibility of oxidation of organometallic complexes present in the soil. The behaviour of Cu in the contaminated soils could also be explained by the presence of chalcopyrite (CuFeS₂) and other Cu-bearing minerals from mine tailings. In oxic soils, the surface of chalcopyrite is generally covered with a layer of Cu-rich sulphides, sulphates and oxides (Cu₂S, CuS, CuSO₄ and Cu₂O). The dissolution of Cu-containing salts present in the outer layer of this mineral could have contributed to the relatively high Cu concentrations in porewaters upon initial flooding. However, the contrasting increase in leachate Cu concentrations with successive cycles suggests that other mechanisms, apart from dissolution, predominate during the oxic phases. In fact, significant Cu leaching might be due to the oxidation of chalcopyrite under oxic conditions as reviewed by Li *et al.*³⁹ The absence of mine tailings in non-contaminated soils would explain why this increasing trend was not observed in NCN.

Although the release of Pb under anoxic conditions was probably related to the reductive dissolution of Fe and Mn oxides, no clear trend was observed with successive redox cycles in all soils analysed. This was probably due to relatively low amounts of Fe and Mn oxide-associated Pb in the soils (Table 2). Moreover, the favourable readsorption of Pb on soil mineral constituents at near-neutral pH²⁹ could have further limited its release with successive redox cycles. However, contrasting results on the release of Pb from flooded soils have been reported. While Ma and Dong⁴⁰ showed that Pb concentrations in pore water of a highly contaminated soil decreased with incubation time, Grybos *et al.*⁴¹ observed that Pb release in an organic, wetland soil increased both under anoxic and oxic conditions. Higher Pb concentrations were generally observed in the soil leachates with respect to the porewaters throughout the experiment, suggesting that this metal is nonetheless potentially lost from soils subjected to alternating redox conditions.

BCR Sequential Extraction

It is widely accepted that the mobility, transport and bioavailability of ME in soils does not only depend on their total contents but also on their partitioning between more or less labile pools. The distribution of ME in operationally-defined fractions of decreasing lability (F1-F4) resulted in the partitioning of total metal contents between soil phases generally referred to as exchangeable, reducible, oxidizable and residual pools, respectively. In general, it appears that metal partitioning was affected by the reduction-oxidation cycles. Although Zn release into solution during the column experiments was mainly attributed to the reductive dissolution of Fe and Mn oxides, only a small decrease in the amount of Zn in the reducible fraction (F2) with respect to the initial soils was observed after the successive redox cycles (Table 3 and Fig. 4). Similarly, Zn in the exchangeable fraction (F1) tended to decrease with redox cycling, while little differences were observed for Zn contents in the oxidizable fraction (F3; Table 3). The concentrations of Zn released into solution during the reductive phases (Fig. 3) were however between 1 and 2 orders of magnitude smaller than Zn contents in the exchangeable and reducible fractions (Table 3). Nonetheless, in all soils alternating redox cycling resulted in an increase in the proportion of Zn in the residual fraction, at the expense of the exchangeable pool (F4; Fig. 4) possibly corroborating the decreasing trend in Zn leached from the columns with successive redox cycles. These results suggest that alternating redox conditions may favour the redistribution of Zn from exchangeable and reducible forms to less labile, more retained forms. A similar decrease in exchangeable and carbonate-associated fractions of Zn, and an increase in the oxidizable fraction was reported for soils incubated under a wetting/drying regime.⁴²

After incubation, Cu in the contaminated soils presented a relative increase in the proportion found in exchangeable forms, particularly in the lower sections of the column (Fig 4), in line with our previous hypothesis that soluble Cu salts from mine tailings may play an important role in determining Cu mobility. This is in agreement with the steady increase in leached Cu concentrations with successive redox cycling (Fig. 3). Alternating redox conditions did not seem to have an effect on Cu redistribution between different fractions in non-contaminated soils where Cu was mainly distributed in the residual (F4) and, to a minor extent, reducible (F2) fractions. However, in contaminated soils, redox cycling led to lower metal contents in reducible and oxidizable fractions (Table 3), with respect to the initial soils. These results are in disagreement with the results obtained by Han *et al.*⁴² and Zheng and Zhang⁹ who report a redistribution of Cu from more labile to more stable forms in ME-spiked soils subjected to wet/dry conditions.

Similarly, Businelli *et al.*⁴³ observed a redistribution of metals with time from organically-bound to the residual fractions, particularly for Cu and Zn, in a compost amended landfill covering soil. Moreover, the results obtained by McLaren and Crawford⁴⁴ who developed a similar fractionation method suggested that the distribution of Cu between soil constituents might be influenced by the association with free Mn (hydr)oxides, as indicated by the DCB and AO extractions (Table 2).

In both contaminated and non-contaminated soils, most of the Pb was found in reducible and residual fractions in approximately equal proportions (Fig. 4), while Pb contents in exchangeable and oxidizable pools were below detection limits (Table 3). In all cases, no important differences were observed in the distribution of Pb between these fractions as a function of alternating redox conditions. Pueyo *et al.*⁴⁵ also found Pb partitioned mainly in the reducible and residual fractions with little change in their distribution as a function of time. The significant Pb contents in the labile reducible fraction (about 40-50% of the total content) suggest that important amounts of this metal can be potentially released into the environment from soils subject to alternating redox conditions. However, the amounts of Pb released into solution during the successive redox cycles (Fig. 3) were rather limited (only about 1% of the total Pb content) suggesting that Pb lability may not necessarily be affected by redox cycles.

Conclusions

The behaviour of Cu, Pb and Zn under alternating soil redox conditions revealed that metal contaminants can be susceptible to release or redistribution in forms which are more or less prone to subsequent leaching. We hypothesized that in redox dynamic environments the potential mobilization of ME from contaminated soils may pose a serious environmental hazard. We observed that the alternation of soil redox conditions may play a more important role in determining ME mobilization with respect to reducing or oxidizing conditions considered separately. The release and redistribution of Zn in pools with reduced lability seemed to be controlled by the reductive dissolution and oxidative reprecipitation of Fe and Mn oxides. Copper dynamics, on the other hand, seemed to be influenced by the presence of redox-sensitive Cu-bearing minerals in contaminated soils. The mobility of Cu was probably determined by the dissolution of chalcopyrite under oxidizing conditions and relative stability under anoxia. Although the behaviour of Pb did not seem to follow a clear trend with redox cycling, the higher concentrations in the leachates with respect to soil porewaters suggested that alternating redox conditions could have favoured Pb mobility. In conclusion, this study pointed out the potential

hazards posed by temporary changes in redox status of mine-tailing contaminated soils, and highlighted the risk of Cu leaching from Cu-bearing minerals due to alternating reductive and oxidative conditions.

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Table 1. Main physico-chemical properties of the studied soils.

	HCN	LCN	NCN
pH _{KCl}	5.0	5.1	5.1
C _{org} (g kg ⁻¹)	37	39	35
N _{tot} (g kg ⁻¹)	3	4	3
Fe (g kg ⁻¹)	42	44	23
Mn (mg kg ⁻¹)	1487	1140	1134
Cu (mg kg ⁻¹)	102	92	18
Pb (mg kg ⁻¹)	46	37	33
Zn (mg kg ⁻¹)	148	130	111
Clay (g kg ⁻¹)	29	43	83
Silt (g kg ⁻¹)	277	354	480
Texture (ISSS)	sandy loam	loam	silt loam

Table 2. Metal concentrations in DCB and AO soil extracts of contaminated (HCN, LCN) and non-contaminated (NCN) soils. Values represent mean metal concentrations \pm standard deviation and values in parenthesis represent metal concentrations expressed as a percentage of pseudo-total contents reported in Table 1.

	HCN		LCN		NCN	
	DCB ^a	AO ^b	DCB	AO	DCB	AO
Fe (g kg ⁻¹)	11.5 \pm 0.3 (27)	5.0 \pm 0.2 (12)	12.3 \pm 0.2 (28)	5.7 \pm 0.2 (13)	11.7 \pm 0.4 (52)	3.2 \pm 0.1 (14)
Mn (mg kg ⁻¹)	652 \pm 15 (44)	519 \pm 13 (35)	794 \pm 9 (70)	533 \pm 11 (47)	957 \pm 54 (84)	685 \pm 50 (60)
Cu (mg kg ⁻¹)	22 \pm 1 (22)	56 \pm 3 (55)	16 \pm 1 (18)	49 \pm 1 (53)	1 \pm 0 (5)	6 \pm 0 (35)
Pb (mg kg ⁻¹)	9 \pm 1 (20)	b.d.l. ^c	4 \pm 1 (11)	b.d.l.	b.d.l.	b.d.l.
Zn (mg kg ⁻¹)	78 \pm 1 (53)	23 \pm 2 (16)	67 \pm 2 (52)	12 \pm 1 (9)	68 \pm 4 (61)	12 \pm 0 (11)

^a Dithionite-citrate-bicarbonate. ^b Ammonium oxalate. ^c Below detection limit.

Table 3. Distribution of metals among fractions of the BCR sequential extraction of contaminated (HCN, LCN) and non-contaminated (NCN) soils.

Values represent mean metal concentrations \pm standard deviation.

BCR step	HCN				LCN				NCN			
	I ^a	U ^b	L ^c	B ^d	I	U	L	B	I	U	L	B
<i>Zn (mg kg⁻¹)</i>												
F1	11.9 \pm 3.3	4.4 \pm 0.6	5.4 \pm 0.9	3.7 \pm 0.4	3.9 \pm 0.6	2.0 \pm 0.4	2.0 \pm 0.2	1.8 \pm 0.1	8.4 \pm 0.7	6.7 \pm 0.3	6.7 \pm 0.4	7.1 \pm 0.2
F2	18.5 \pm 1.9	15.2 \pm 0.6	18.2 \pm 0.8	15.8 \pm 0.9	10.5 \pm 0.9	8.5 \pm 1.1	8.7 \pm 1.7	7.6 \pm 0.8	9.6 \pm 0.4	8.7 \pm 0.5	9.3 \pm 1.4	8.6 \pm 0.7
F3	13.6 \pm 0.8	13.9 \pm 1.3	15.1 \pm 2.6	16.6 \pm 3.6	9.2 \pm 0.3	9.1 \pm 0.7	8.4 \pm 1.2	8.3 \pm 1.0	10.9 \pm 0.3	11.0 \pm 0.4	10.8 \pm 0.9	10.4 \pm 0.3
F4	108.7 \pm 12.0	117.4 \pm 7.2	109.4 \pm 2.9	113.7 \pm 7.5	94.8 \pm 7.8	91.4 \pm 1.7	90.9 \pm 2.1	90.2 \pm 2.4	73.0 \pm 1.0	79.3 \pm 5.3	78.8 \pm 1.9	78.0 \pm 5.3
Sum	152.7 \pm 18.0	150.9 \pm 9.7	148.1 \pm 7.2	149.8 \pm 12.4	118.4 \pm 9.6	111.0 \pm 3.9	110.0 \pm 5.2	107.9 \pm 4.3	101.9 \pm 2.4	105.7 \pm 6.5	105.6 \pm 4.6	104.1 \pm 5.3
<i>Cu (mg kg⁻¹)</i>												
F1	4.2 \pm 0.3	6.0 \pm 0.2	8.8 \pm 0.3	6.7 \pm 0.5	2.4 \pm 0.3	3.2 \pm 0.5	5.8 \pm 0.8	5.8 \pm 0.6	b.d.l. ^e	b.d.l.	b.d.l.	b.d.l.
F2	41.3 \pm 0.8	35.3 \pm 0.5	34.6 \pm 0.3	33.4 \pm 1.2	29.7 \pm 0.5	24.4 \pm 1.1	25.0 \pm 0.4	24.8 \pm 0.5	1.5 \pm 0.2	1.5 \pm 0.2	1.5 \pm 0.1	1.7 \pm 0.2
F3	17.2 \pm 0.5	14.9 \pm 1.5	13.5 \pm 1.1	13.9 \pm 1.7	8.8 \pm 0.5	7.2 \pm 1.2	5.6 \pm 0.6	5.4 \pm 0.3	b.d.l.	b.d.l.	b.d.l.	b.d.l.
F4	40.3 \pm 1.4	47.1 \pm 2.4	40.3 \pm 1.5	44.4 \pm 2.4	43.4 \pm 2.7	42.1 \pm 1.5	38.6 \pm 0.9	37.4 \pm 2.7	11.7 \pm 0.0	12.5 \pm 0.3	12.1 \pm 0.6	12.3 \pm 0.5
Sum	103.0 \pm 3.0	103.3 \pm 4.6	97.2 \pm 3.2	98.4 \pm 5.8	84.3 \pm 4.0	76.9 \pm 4.3	75.0 \pm 2.7	73.4 \pm 4.1	13.2 \pm 0.2	14.0 \pm 0.5	13.6 \pm 0.7	14.0 \pm 0.7
<i>Pb (mg kg⁻¹)</i>												
F1	b.d.l.	b.d.l.	b.d.l.	b.d.l.	b.d.l.	b.d.l.	b.d.l.	b.d.l.	b.d.l.	b.d.l.	b.d.l.	b.d.l.
F2	22.2 \pm 1.9	22.4 \pm 1.3	23.3 \pm 1.4	23.2 \pm 1.4	12.4 \pm 0.2	10.9 \pm 1.6	11.3 \pm 1.8	11.5 \pm 0.4	11.5 \pm 1.3	14.4 \pm 2.8	17.3 \pm 2.6	13.7 \pm 1.7
F3	b.d.l.	b.d.l.	b.d.l.	b.d.l.	b.d.l.	b.d.l.	b.d.l.	b.d.l.	b.d.l.	b.d.l.	b.d.l.	b.d.l.
F4	21.6 \pm 1.9	22.5 \pm 1.8	22.1 \pm 1.4	23.3 \pm 1.1	22.8 \pm 5.7	16.6 \pm 0.9	16.1 \pm 1.0	16.1 \pm 2.6	16.5 \pm 0.4	14.9 \pm 2.0	15.6 \pm 2.7	18.2 \pm 2.4
Sum	43.8 \pm 3.8	44.9 \pm 3.1	45.4 \pm 2.8	46.5 \pm 2.5	35.2 \pm 5.9	27.5 \pm 2.5	27.4 \pm 2.8	27.6 \pm 3.0	28.0 \pm 1.7	29.3 \pm 4.8	32.9 \pm 5.3	31.9 \pm 4.1

^a Initial soil. ^b Upper soil layer. ^c Lower soil layer. ^d Bottom soil layer. ^e Below detection limit.

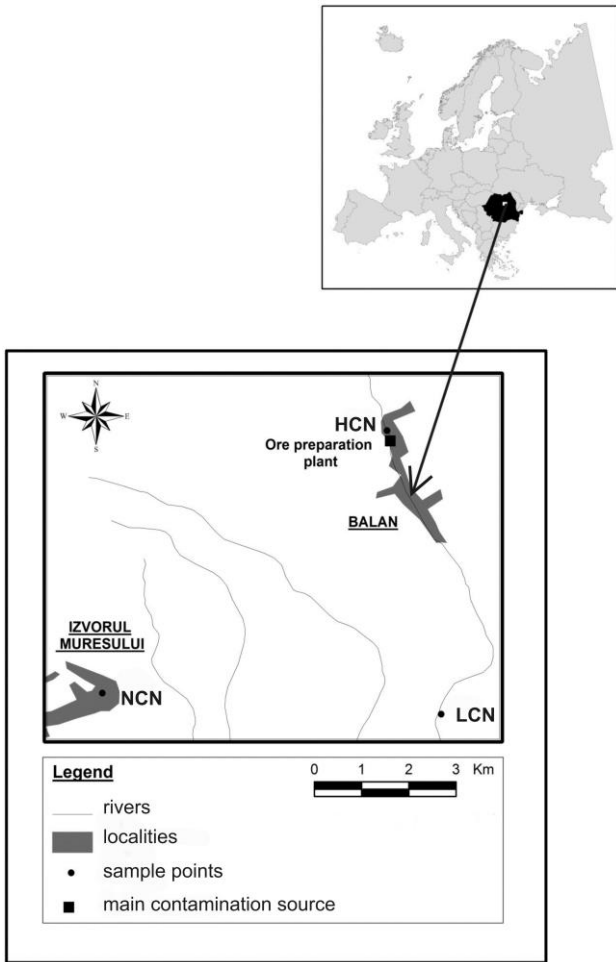


Fig. 1. Schematic map of the studied area. HCN – highly contaminated soil; LCN – low contaminated soil; NCN – non-contaminated soil.

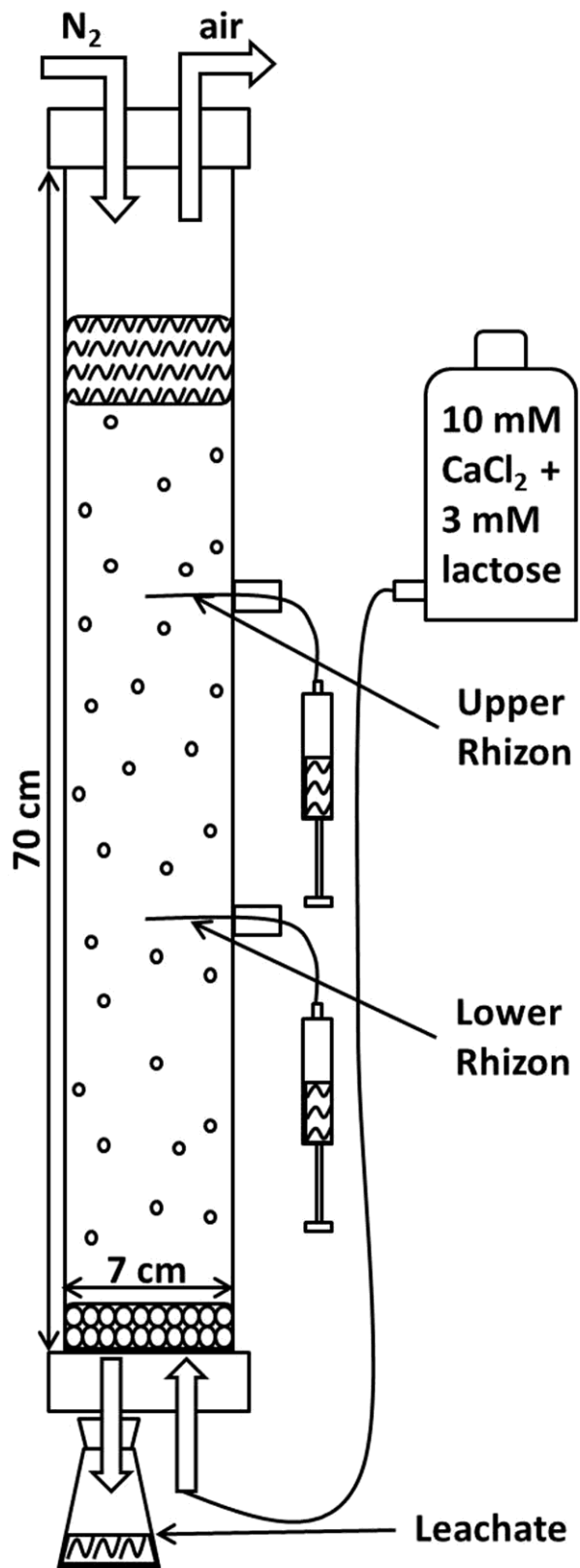


Fig. 2. Sketch of the column experimental setup. Adapted from Balint *et al.*²⁴

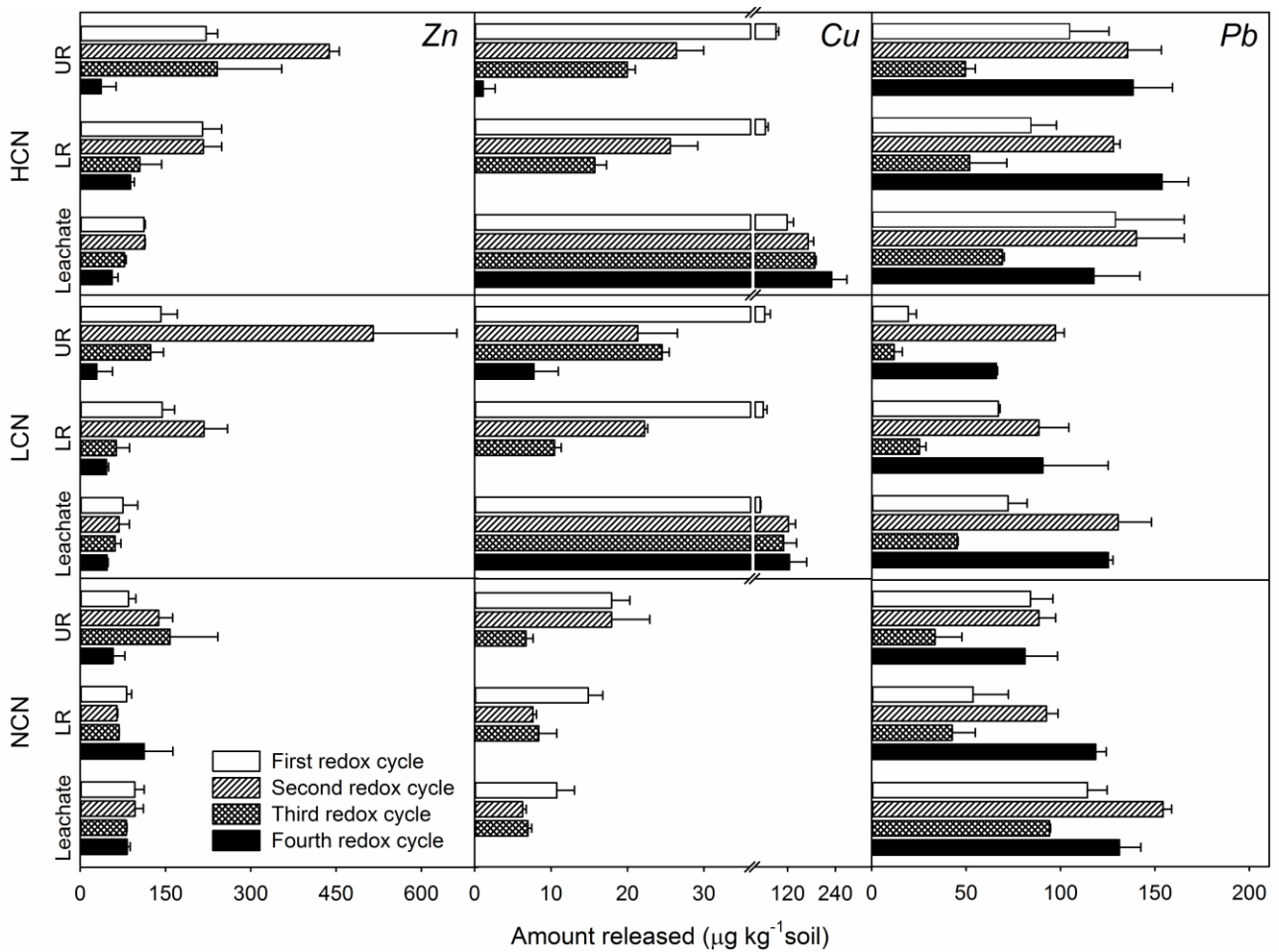


Fig. 3. Amount of Zn, Cu and Pb and released in solution, as extracted by Rhizon samplers and leached through the column in the four redox cycles. UR, LR – soil solution sampled with the upper and lower Rhizon, respectively. Leachate – soil solution drained from the column. HCN, LCN, and NCN – highly, low, and non-contaminated soil, respectively. Note the break in the scale of Cu. Error bars represent standard deviations.

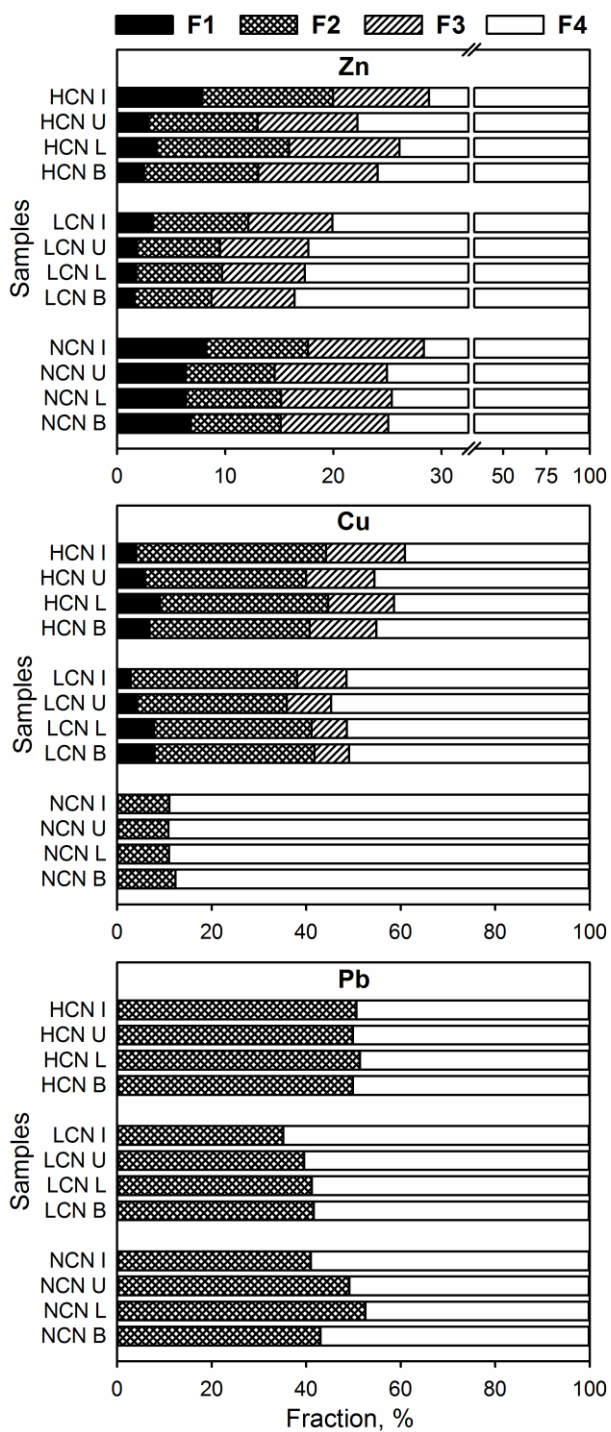


Fig. 4. BCR metal fractionation performed on initial (I) and column-incubated soil (U, L, B – upper, lower, and bottom layers, respectively) after 4 redox cycles. HCN, LCN, and NCN – highly, low, and non-contaminated soil, respectively. Note the break in the fraction scale of Zn.