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Author(s)	Kent, Christopher; Rao Khandavilli, Udaya B. Rao; Alfarsi, Anas; Hanna-Brown, Melissa; McSweeney, Seán; Krüse, Jacob; Lawrence, Simon E.; Fitzpatrick, Dara
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Tracking Cocrystallization of Active Pharmaceutical Ingredients with Benzoic Acid Coformer using Broadband Acoustic Resonance Dissolution Spectroscopy (BARDS)

Christopher Kent[‡], Uday B. Rao Khandavilli[†] Anas Alfarsi[‡], Melissa Hanna-Brown[#], Seán McSweeney[¥], Jacob Krüse, Simon Lawrence^{‡†}, and Dara Fitzpatrick^{‡}*

[‡] School of Chemistry, Analytical and Biological Chemistry Research Facility, University College Cork, Cork, Ireland.

[†] Synthesis and Solid State Pharmaceutical Cluster, University College Cork, Ireland.

[#] Pfizer, Sandwich, Kent, United Kingdom.

[¥] BARDS Acoustic Science Labs, Bio-Innovation Centre, University College Cork

AUTHOR EMAIL ADDRESS: d.fitzpatrick@ucc.ie

Address: School of Chemistry, Analytical and Biological Chemistry Research Facility, University College Cork, Cork, Ireland

Email: d.fitzpatrick@ucc.ie. Tel.: +353 21 490 3299

ABSTRACT

This study investigates the use of Broadband Acoustic Resonance Dissolution Spectroscopy (BARDS) as a detection method for the formation of cocrystals. BARDS is a novel approach which uses reproducible changes in the compressibility of a solvent as a sample dissolves to characterize and differentiate between materials and in this case co-crystallization. Two cocrystal systems with a 1:1 stoichiometry were examined which used benzoic acid as a co-former with isonicotinamide, and with theophylline. Cocrystals were prepared using dry and wet milling for periods from 1 to 40 min and samples were analyzed using infrared spectroscopy, powder X-Ray diffraction and BARDS. Comparison of the BARDS data with the IR and PXRD data cross-validated the BARDS results. This study shows that BARDS can be used to rapidly assess the formation of these cocrystals at-line when

1 milling or as a relatively low cost tool in pre-formulation product development. The data can also be
2 used to gauge the unique entrained gas and gas volume generation of the co-crystal samples during
3 dissolution and their dissolution kinetics.
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9 INTRODUCTION

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11 Cocrystallization is the formation of a crystalline sample comprised of two or more compounds.¹ Over
12 the last two decades it has evolved into a mature science that is increasingly popular in the
13 pharmaceutical industry as a method to improve drug delivery through the formation of a new entity
14 with improved physiochemical properties.
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21 Solid state grinding is one method that has been used to synthesize cocrystals.² One of the main
22 advantages of this method is that solvent is not required, which means the solubility of the starting
23 materials is not an issue and doesn't require as much experimental development. The risk of forming
24 solvates is removed, and there is no solvent waste. There are some disadvantages as well; the solid-state
25 cocrystal cannot be analyzed using single crystal X-ray diffraction, meaning that alternative ways to
26 identify structural features, such as the hydrogen bond motifs present within the cocrystal, are required.
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28 The physical grinding of components together may cause a change to the degree of crystallinity and may
29 lead to changes in the physicochemical properties.
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40 Solution based cocrystallization methods involve evaporation or cooling and they have some clear
41 advantages over solid state grinding such as high yields of well-defined crystals with high purity that
42 can be structurally analyzed relatively easily.³ However, there are disadvantages to solution
43 crystallization, for example the formation of solvates that can be challenging to handle or have
44 regulatory issues. Another obstacle is the solubility of the starting materials with the solvent system;
45 both must dissolve into solution before cocrystallization may occur. The solvent can interfere with the
46 cocrystallization by being involved with the cocrystal intermolecular interactions, therefore forming part
47 of the crystal lattice and altering the design of the cocrystal. When a polymorphic cocrystal system is an
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option, the solution may lead to an unwanted kinetically controlled metastable polymorphs. The cost associated with the amount of solvent required to form the cocrystal can also be an issue.²

Infrared spectroscopy and powder x-ray diffraction are routinely used for the analysis of cocrystals, although other methods have been employed, for example terahertz spectroscopy.⁴ Broadband Acoustic Resonance Dissolution Spectroscopy (BARDS) is based on an acoustic phenomenon that has been long established,⁵ but has only recently been harnessed for analytical purposes.⁶

It is founded on reproducible changes in the compressibility of a solvent during dissolution of a sample. Entrained gas between and within particles creates gas evolution in conjunction with the reduction solubility of gas already in solution due to the dissolution of the solute. This leads to an increase in compressibility and a decrease in the speed of sound. The frequency change in induced resonant sound waves are used to capture changes in compressibility.

A more detailed and comprehensive outline of the underlying processes and equations involved in BARDS analysis is given by Fitzpatrick and coworkers (2012).⁶

In this study, BARDS was investigated to see whether it can be used to track cocrystallization. Cross-validation with the more routine methods for monitoring cocrystallization, namely IR spectroscopy and PXRD, was undertaken. Two 1:1 stoichiometry cocrystal systems were examined: benzoic acid with isonicotinamide,⁷ and benzoic acid with theophylline.⁸

Isonicotinamide (INA) has five known polymorphs^{9,10,11} and has been used in the pharmaceutical industry for its unique set of properties. Mixed salts of INA have a broad range of uses including being a widely used coformer.¹² It is the pyridyl nitrogen atom that acts as a hydrogen bond acceptor when mixed with hydrogen bond donors such as carboxylic acids. The bond between the carboxylic acid and pyridine is thought to be the most adaptable and most stable hydrogen bond and is the main type of bonding even in the presence of other good hydrogen bond donors.¹³

Theophylline (TP) is used to treat chronic obstructive pulmonary disease,¹⁴ and it is also used for its therapeutic effect for the respiratory disease, asthma.¹⁵ It has low solubility and poor bioavailability.

When theophylline forms a cocrystal with acids and bases, the solubility can be improved, but the

1 formation can only occur if the carbonyl groups and aromatic nitrogen atoms become involved in the
2 formation with the cocrystal former.¹⁶ Formulation of TP is difficult because it is known to change from
3 the anhydrate to a monohydrate form and back, depending on the relative humidity. The possibility of
4 crystalline hydrate formation makes it difficult to formulate a design that is consistent and reproducible
5 for this API.¹⁷

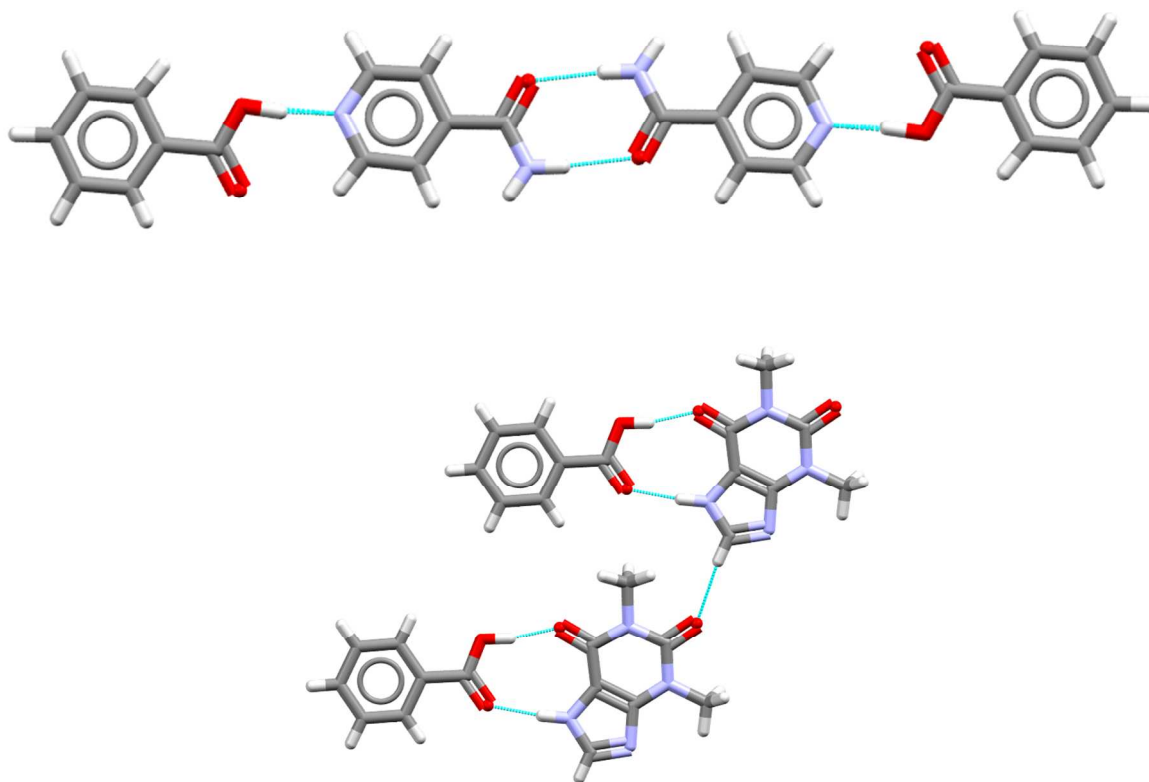
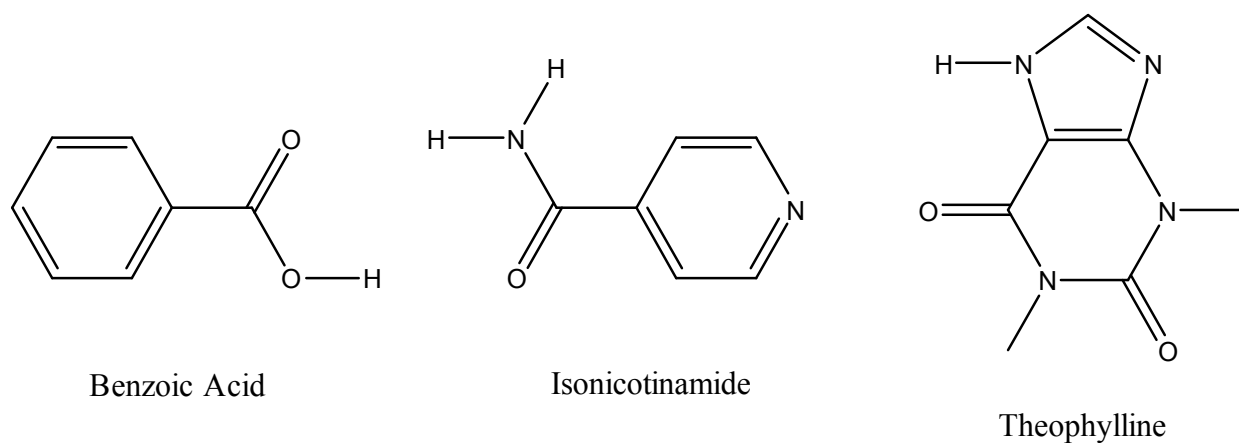


Figure 1: The compounds and co-crystal investigated in this study.

EXPERIMENTAL SECTION

1 **Materials** All compounds were purchased from Sigma Aldrich and used without any further
2 purification. Where indicated, distilled water was used.
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4 **Grinding Experiments.** Mechanical grinding experiments were conducted in a Retsch MM400 Mixer
5 mill, equipped with stainless steel 5 mL grinding jars and one 5 mm stainless steel grinding ball per jar.
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7 The mill was operated at a rate of 30 Hz for a variety of different times ranging from 1 min to 40 min.
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9 For the BA: INA cocrystal dry grinding was used, with a total sample mass of 0.75 g in a 1:1mole ratio
10 (mass 0.375 g of both was taken which is 0.003 moles of both). For the BA: TP cocrystal liquid assisted
11 grinding was used. For BA: TP a total mass of 0.75 g in a 1:1 mole ratio (mass of theophylline used was
12 0.45 g/ 0.0025 moles and mass of benzoic acid used was 0.3 g or 0.0025 moles) was used, with the
13 addition of 4 drops of methanol before milling. After milling, the sample was allowed to dry for 24 h.
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23 **Solution Crystallization.** A 1:1 molar ratio of cofomers were mixed together in the solid state,
24 dissolved and allowed to stand at room temperature until the solvent had completely evaporated,
25 typically 3-9 days.
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30 **Instrumentation** The BARDS instrumentation consists of a dissolution vessel made of borosilicate
31 glass. There is a microphone situated above the borosilicate glass vessel and within the dissolution
32 chamber. Figure 2 shows a schematic of the instrumentation. The dissolution vessel contains 25 mL of a
33 40: 60 (v/v) ethanol: water mixture. The vessel is placed off-center with the stirrer plate which allows
34 the magnetic follower to gently tap the inner glass wall. This mechanical energy results in an acoustic
35 resonance to occur in three different media, (i.e. the glass, solvent, and the air column above the
36 solvent). The microphone converts this induced acoustic resonance to a spectrum using a PC with a
37 sound card and specialized software. The resonances of the liquid vessel are recorded in a frequency
38 band of 0–20 kHz. The steady state resonance was recorded for thirty seconds at the beginning of each
39 run. At this point the sample was added automatically, from a weighing boat, using a tipper motor to 25
40 mL of solvent in the dissolution vessel. Each sample run lasts for five hundred seconds and each sample
41 was run in duplicate. Excel data files which contained the experimental data were used to obtain an
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average and spread of the duplicate results, for set points along the time axis of the spectra These results were transferred to another graphics program (Sigma Plot) which was used to compile the figures.



Figure 2: (A) Top view schematic diagram representing the contents of the dissolution chamber and (B) photograph of the BARDS dissolution chamber. Note the microphone is not in contact with the vessel. External view of the instrument (C) Tipper motor with a sample in a weighing boat ready for addition to the stirring solution below.

IR spectrums were recorded on a Bruker Tensor 37 in the range of 4000 to 500 cm^{-1} motif. PXRD analysis was carried out using a Bruker D2 Phaser over the 2θ range of 5-45° using a Lynxeye (1D mode) detector with a scan speed of 0.2 s for 797 steps. The Eva software was used to overlay the experimental data with the reference spectra from the CSD.¹⁸

RESULTS AND DISCUSSION

Table 1 shows the results of the grinding experiments for the two cocrystals for a variety of different grinding times. Analysis is based on the shift in the $\nu(\text{CO})$ for benzoic acid and PXRD analysis, see Supporting Information. In all cases it is the 1:1 cocrystal that is formed.^{7,8} Table 1 indicates that the co crystal for INA: BA formed completely after 15 min of dry grinding and the co crystal for TP: BA formed completely after 20 min of liquid assisted grinding.

Table 1: The effect of different grinding times on the synthesis of the cocrystals.

coformer	coformer	Time (min)	Method	Products
BA	INA	1	Dry grinding	Both cofomers
BA	INA	2	Dry grinding	Both cofomers
BA	INA	5	Dry grinding	Both cofomers
BA	INA	10	Dry grinding	Mix of cofomers and Co Crystal
BA	INA	15	Dry grinding	Co Crystal
BA	INA	20	Dry grinding	Co Crystal
BA	INA	40	Dry grinding	Co Crystal
BA	TP	5	Liquid assisted grinding	Both cofomers
BA	TP	10	Liquid assisted grinding	Both cofomers
BA	TP	15	Liquid assisted grinding	Both cofomers
BA	TP	20	Liquid assisted grinding	Co Crystal

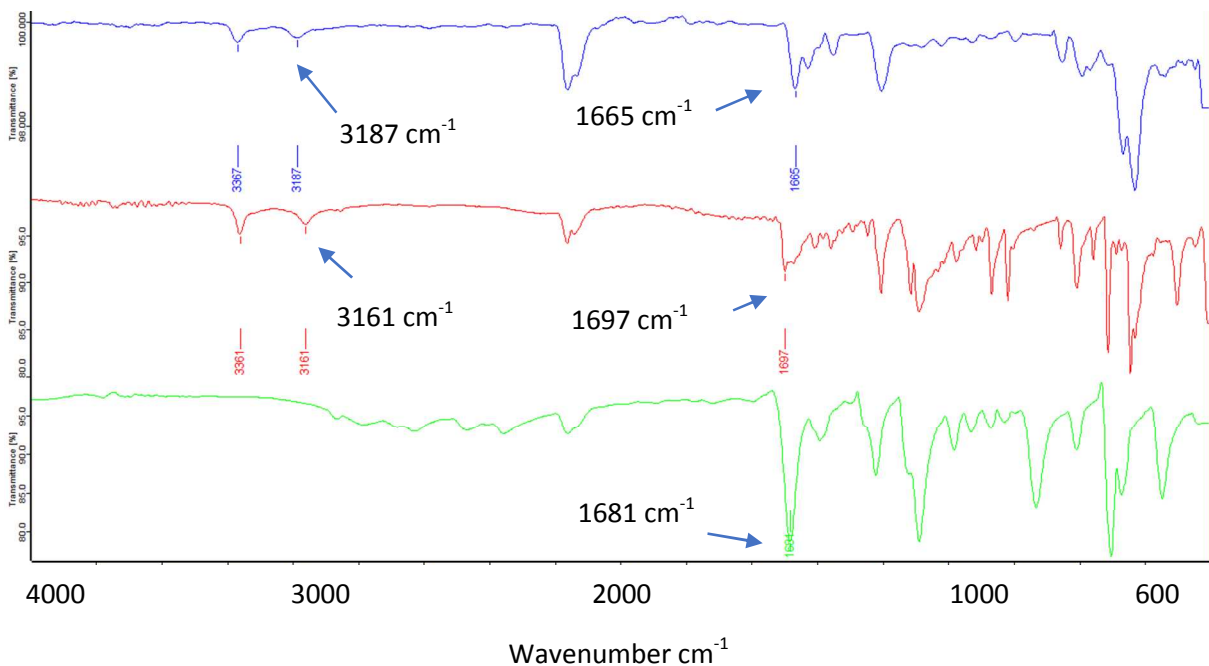


Figure 3: IR analysis for the formation of a 1:1 cocrystal of INA: BA after 15 min of milling (red spectrum) with BA (green spectrum), and INA (blue spectrum). Note the presence of key absorbances in the red spectrum at 3161 cm^{-1} and 1697 cm^{-1}

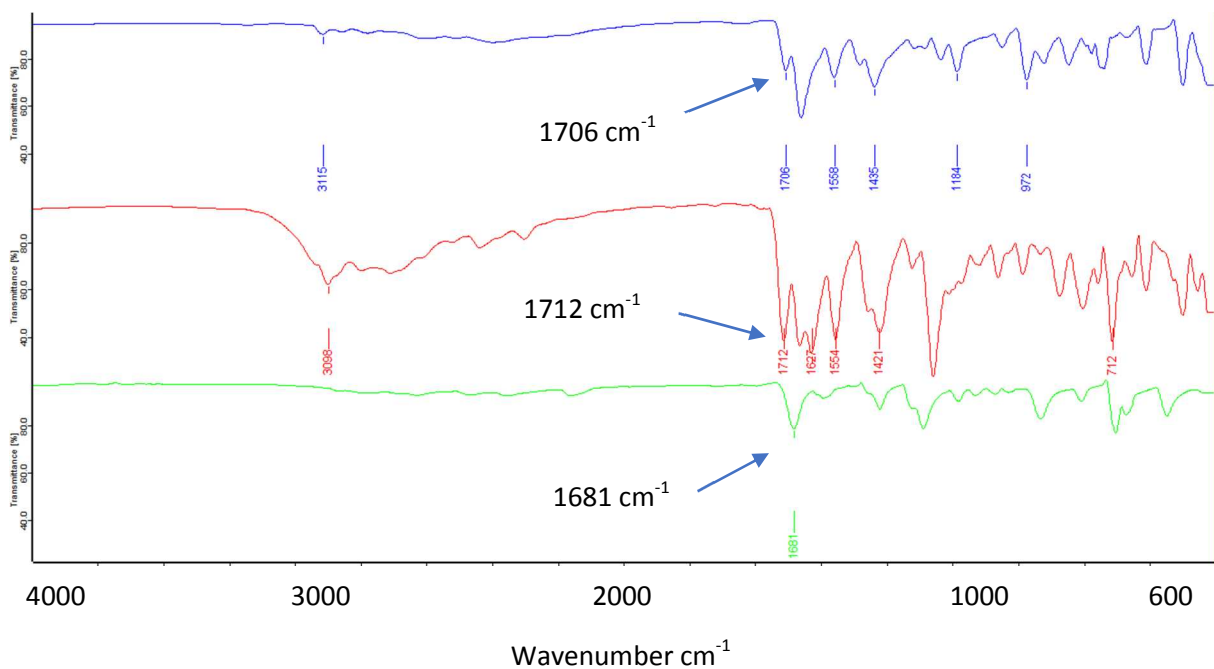


Figure 4: IR analysis for the formation of the 1:1 cocrystal of TP: BA (red spectrum) with BA (green spectrum), and INA (blue spectrum).

In this study, BARDS will be assessed for its ability to track cocrystallization. When assessing a 1:1 blended mixture of API and coformer for cocrystallization, the mixture (300 mg) is added to the solvent.

The process of dissolution introduces gas into the solvent, which changes its compressibility. Each component contributes to the response and a unique acoustic profile for the blend occurs. This is the principle behind tracking the formation of a cocrystal by solution based, solid state and liquid assisted grinding. The dissolution of the mixture changes the speed of sound in the dissolution vessel. The formation of air bubbles from the mixture increases the compressibility of the solution. The frequency is lowered due to the speed of sound being altered to a lower velocity and this is picked up by a microphone within the dissolution chamber.

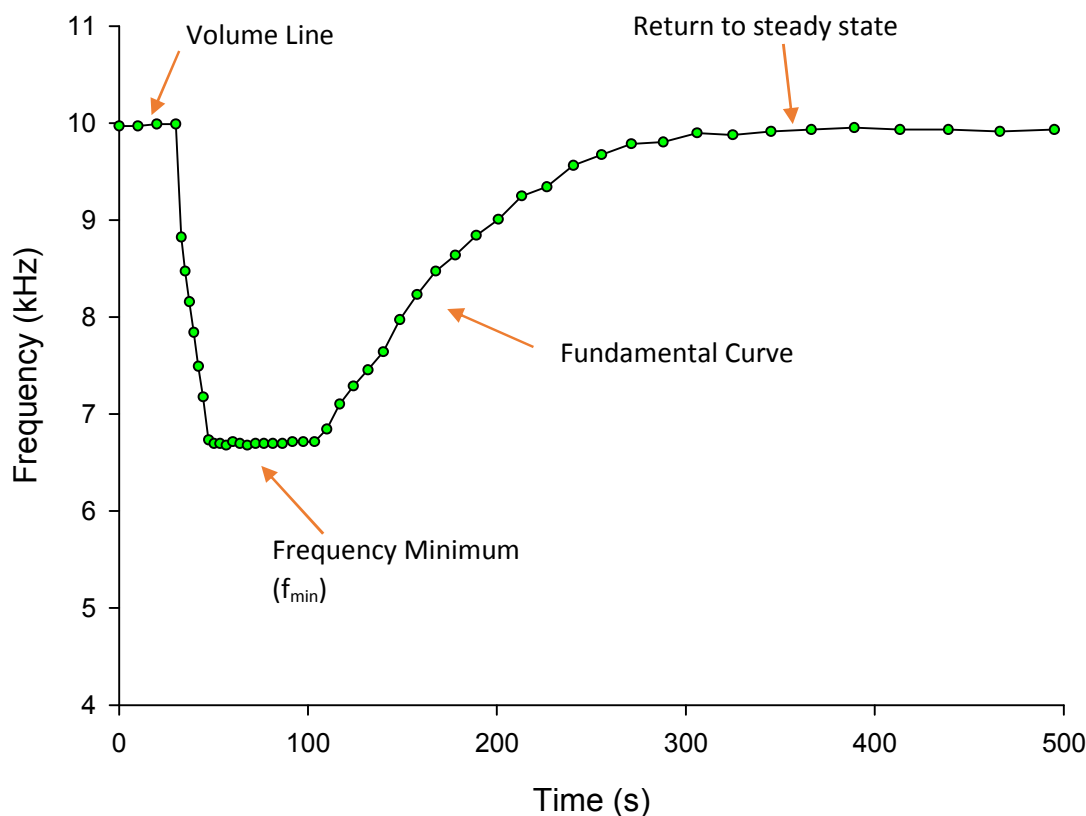


Figure 5: Sigma plot of a BARDS spectrum of benzoic acid dissolved in 25 mL of a 40:60 ethanol: water mixture. Note the steady state resonances for 30 s before the addition of the compound.

The steady state resonance was recorded for thirty seconds at the beginning of each run. At this point the sample was added automatically, from a weighing boat, to 25 mL of solvent in the dissolution

1 vessel. Each sample run lasts for five hundred seconds. Excel data is then transferred to the Sigma Plot
2 program which was used to compile the Bards overlays in Figures 6-9.
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4 The frequency minimum (f_{\min}) of the BARDS response in Figure 5 represents an equilibrium between
5 the rate of gas formation and the rate of liberation of gas from the surface of the solvent. Figure 5 shows
6 the BARDS spectrum of Benzoic acid, a well-known coformer used in the formation of both cocrystals
7 in this study. The borosilicate vessel is induced to resonate by the magnetic stirrer as shown in Figure 2.
8 The initial thirty seconds of the spectra represent just the resonance of the vessel containing 25 mL of
9 solvent. The sample is subsequently added which results in a change in the resonant frequency. These
10 changes are hereafter known as the fundamental curve.
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12 After the initial PXRD and IR analysis of the synthesized cocrystals, the next step was to choose a
13 dissolution solvent for BARDS analysis. Deionized water was used initially, however the solubility of
14 the BA was found to be poor and the milled 1:1 mixture of the INA: BA caused the mixture to clump
15 together. Analysis of this sample mixture was unsuccessful due to poor wetting and the sample would
16 sit on the surface of the deionized water after one minute of dry grinding. Therefore, no change to the
17 steady state frequency was observed and hence no acoustic response was obtained. An organic modifier
18 (ethanol) was used to improve the solubility of BA and to achieve a response from the 1:1 mixture. A
19 40:60 (v/v) ethanol: water mixture was found to be optimum and was used for analysis and this yielded
20 interpretable acoustic responses (Figures 5-9).
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22 The sample mass placed into the milling jar is an important parameter in terms of formation time of
23 cocrystal. The sample mass used throughout this study is 0.75 g. Other important parameters during
24 milling are the frequency used on the Retsch MM400 mixing mill, the size of the milling balls, the size
25 of the milling jar and whether a molecular lubricant is added or not. It is well understood that adding a
26 molecular lubricant such as methanol (i.e. liquid assisted grinding) will increase the reaction rate and
27 increase yield ¹⁹.
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42 Baseline Data Analysis

The PXRD spectra in Figure 6 (A) and (B) shows the overlay of the experimental data in red for INA and BA with the reference spectra in black from the Cambridge Structural Database (CSD).

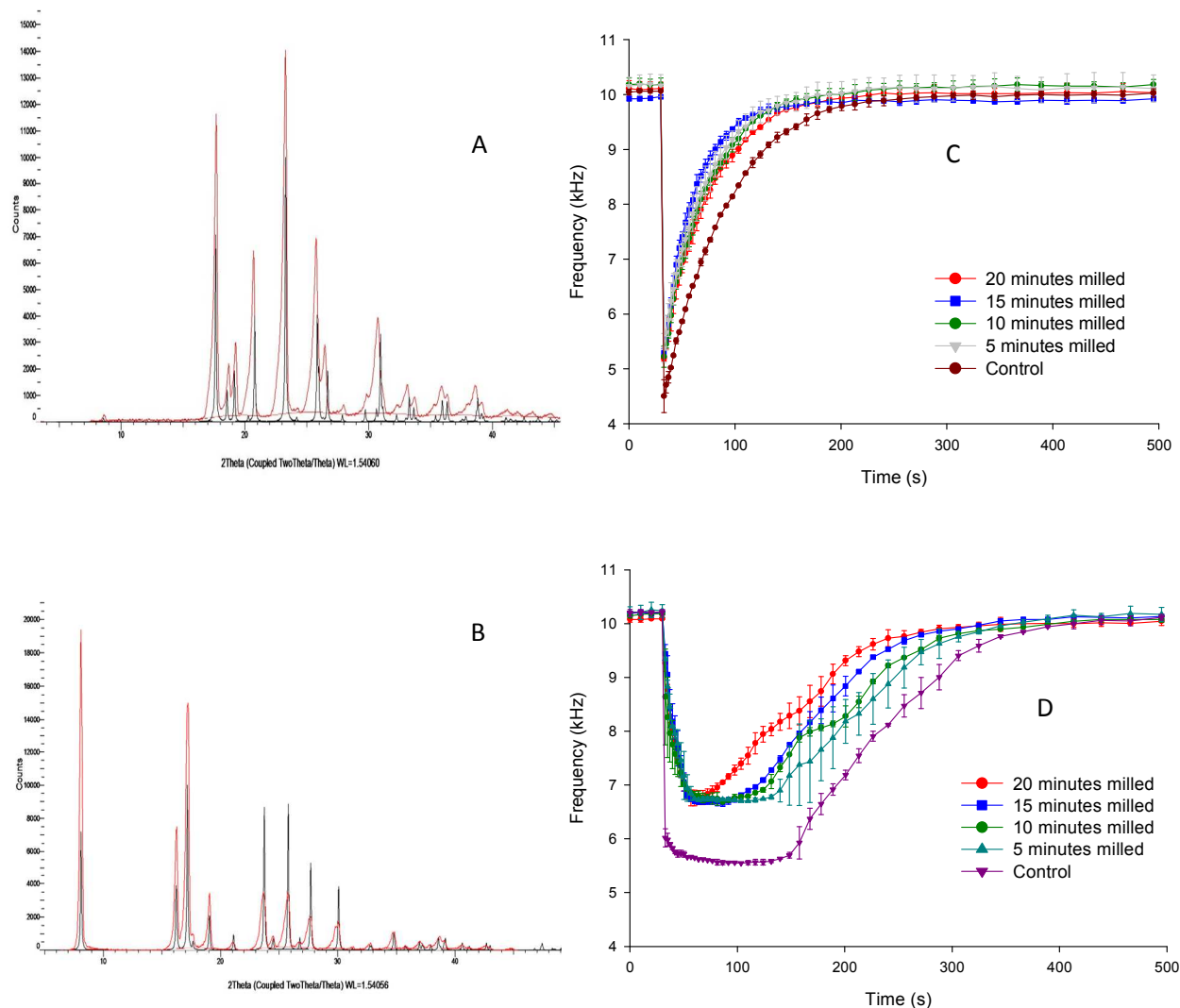


Figure 6: PXRD spectra of (A) INA, and (B) BA (20 min) compared to their corresponding Cambridge Structural Database spectra. The corresponding BARDS spectra (C) and (D) are shown, respectively. Note: the milling process has little effect on the BARDS response for INA, whereas BA changes significantly with increased milling time. Note: the PXRD analysis yielded the same result throughout the milling process, this indicates that the crystal structure for both BA and INA remained the same. Also, the control in (C) and (D) and in all figures are pre milled samples.

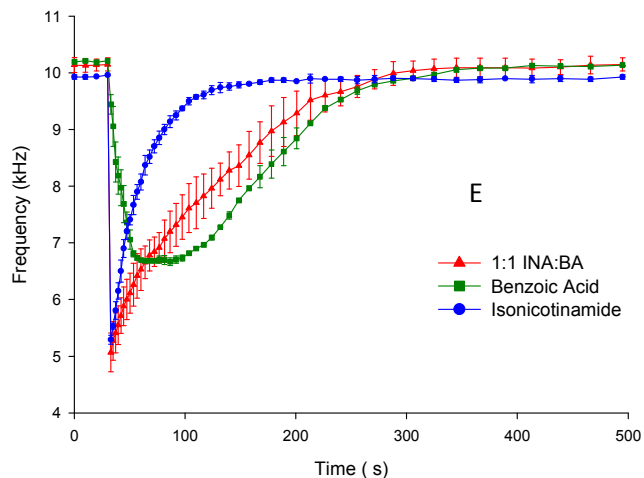
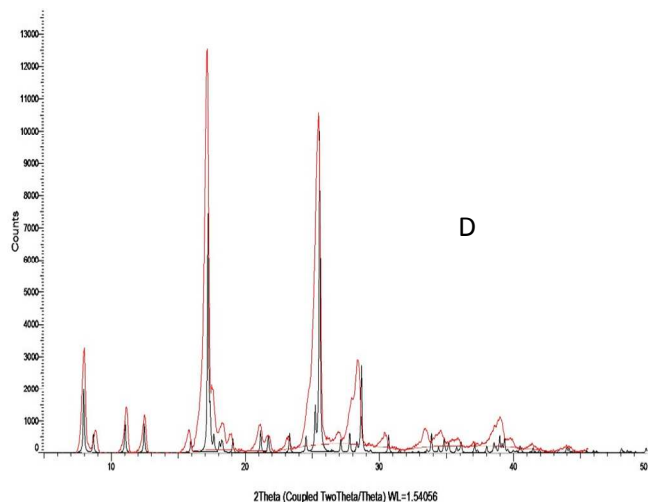
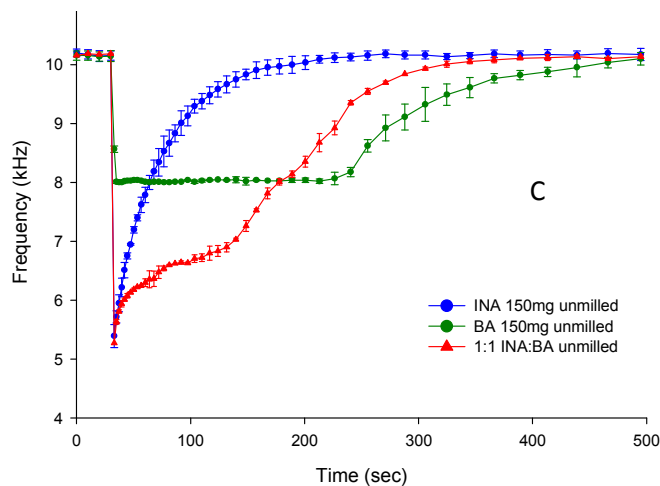
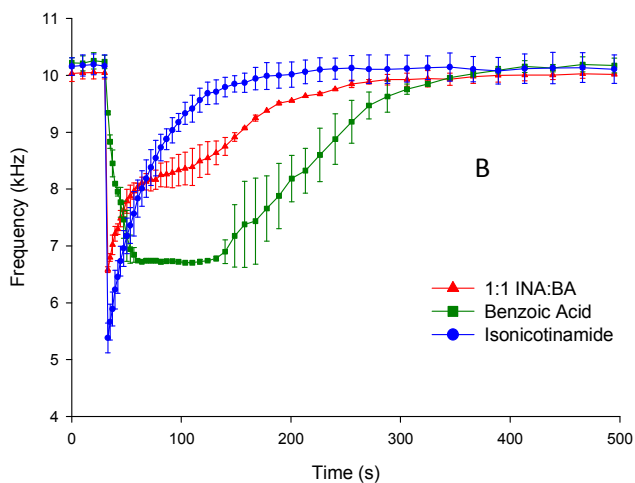
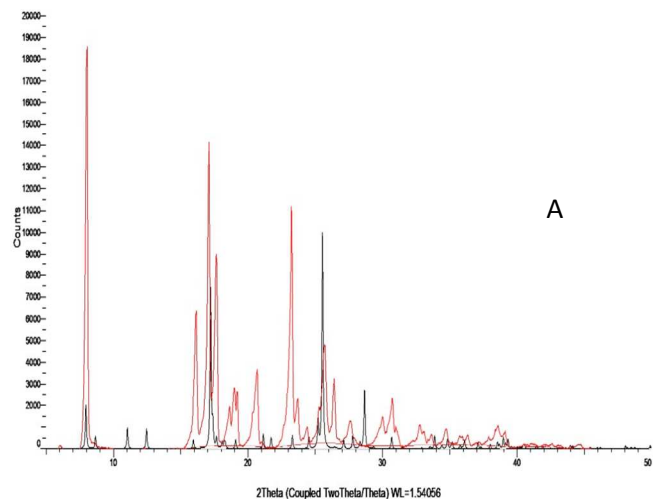
The PXRD pattern suggests there is no change in the crystalline form of either irrespective of the milling time. This means that any changes in the BARDS spectra are caused by other morphological

1 parameters, e.g., particle size or entrained gas. The corresponding BARDS spectra for INA for selected
2 milling times is shown in Figure 6 (C). Note that the spectra change slightly with increased milling time
3 and that there is a decrease in the time taken (ΔT) for the profiles to reach the steady state from the f_{\min}
4 value. The lack of change indicates that the milling process has little effect on the dissolution of the INA
5 samples. The f_{\min} correspond to an equilibrium between the liberation of gas in solution and the exiting
6 of gas at the surface. It also represents the maximum compressibility of the solution. The PXRD
7 spectrum, in Figure 6 (B), shows the overlay of the experimental data in red for BA with the reference
8 spectrum in black from the CSD. The same PXRD pattern was obtained for each minute of milling,
9 which indicates that the same polymorphic structure of BA was obtained for each minute of milling.
10 The corresponding BARDS spectra for selected milling times are shown beside figure 6 (D). Notice that
11 the spectra change dramatically with increased milling time. This indicates that the milling process has a
12 greater effect on the dissolution of the BA samples. There is an expected increase in dissolution rate of
13 the samples with increasing milling time due to the increase in the overall surface area per unit weight,
14 this is observed at 20 min of milling in the BARDS spectrum in figure 6 (D). The control sample has a
15 frequency minimum of 5.8 kHz. This increases to 6.8 kHz after 5 min of dry grinding. Notice also that
16 there is a plateau within each spectrum up to the 20 min time point of dry grinding. The equilibrium
17 time is reducing as the milling time increases. This means that the BA particles are releasing less gas
18 trapped within and between particles post milling. Baseline data for each co-former is now considered
19 sufficient to proceed with the analysis of the cocrystallisation.
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46 Cocrystallisation Analysis

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48 The post milling (5 min and 15 min) PXRD and BARDS data for 0.75 g INA:BA are shown in Figure
49 7. Figure 7 (A) shows there is little correlation between the coformer samples and the cocrystal
50 reference PXRD spectrum after 5 min of milling. The corresponding BARDS spectrum 7 (B) shows that
51 the red profile associated with the same sample has characteristics of both the INA and the BA spectra,
52 not unlike the mixture of the unmilled material shown in Figure 7 (C red profile) However, after 15 min
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of milling, Figure 7 (D) shows there is significant correlation between the sample spectrum and the PXRD reference. Also, the BA features of the BARDS spectrum are now absent in Figure 7 (E). Therefore, the data indicates that it takes 15 to 20 min to form the INA:BA cocrystal with a sample mass of 0.75g.



1 **Figure 7:** (A) PXRD data after 5 min of solid state grinding for a 1:1 mixture of INA: BA. (B) Corresponding
2 BARDS spectra for (A). Each profile represents 300 mg of material. Note the red profile showing characteristics
3 of BA similar to the unmilled mixture shown by the red profile in (C). (D) PXRD data after 15 min of solid state
4 grinding for a 1:1 mixture of INA: BA. Note: the PXRD spectra match between the experimental and reference
5 indicating the formation of the 1:1 cocrystal. (E) The corresponding BARDS spectra (red) showing the absence of
6 BA characteristics.
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13 The BARDS spectral data corresponds to the PXRD analysis. Following grinding of the 1:1 mixture
14 components together for 5 min, the characteristic plateau of the BA constituent is evident in Figure 7
15 (B) whereas, in Figure 7 (E) the formation of a smooth curve and the disappearance of this plateau
16 indicates the formation of the newly formed cocrystal.
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23 The formation of the alternate model cocrystal (1:1 TP: BA) was not achieved after 20 min of solid
24 state grinding. It has been reported that the theophylline cocrystal can be formed in a 10 mL milling jar
25 with two to three 10 mm steel balls using the dry grinding method.^{8,19} However, the liquid assisted
26 grinding method was used in this study to help in the formation of the cocrystal by adding four drops of
27 methanol to the total load of 0.75 g. Figure 8 (A) shows the PXRD overlay of the milled samples and
28 the reference spectra for the co-crystal after five min of liquid assisted grinding. The sample spectra
29 shows the characteristics of both unbound TP and BA. The corresponding BARDS spectra in 8 (B) also
30 show the characteristics of both compounds with the presence of a plateau ~ 8 kHz (red profile). This
31 suggests that the structure of the 1:1 cocrystal of TP: BA has not formed with a plateau of un-milled BA
32 similar to the red profile in Figure 8 (C). However, PXRD data in Figure 8 (D) shows evidence that 1:1
33 cocrystal has formed after 20 minutes of milling.
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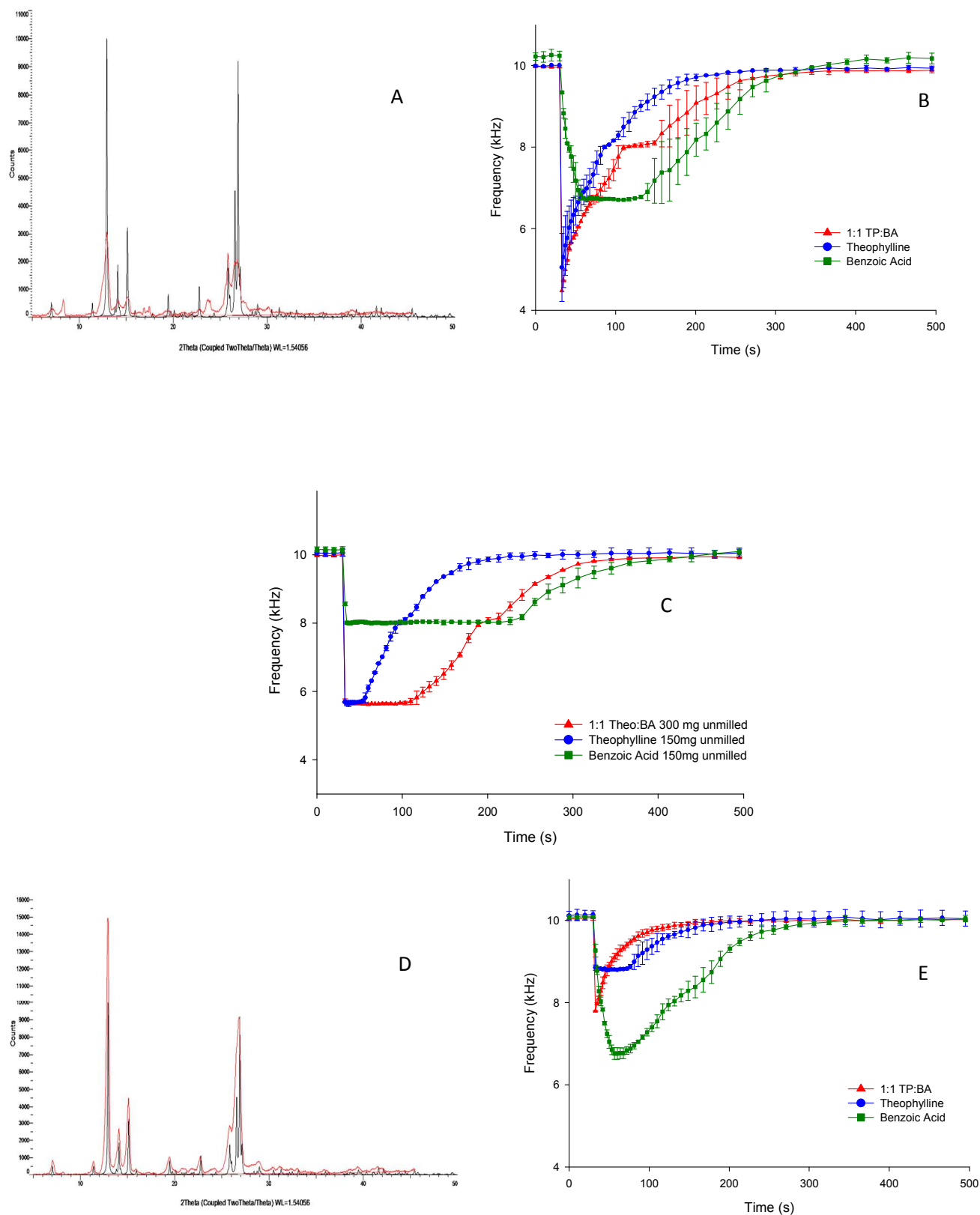


Figure 8: (A) PXRD data and corresponding BARDS data (B) after 5 min of liquid assisted grinding for a 1:1 mixture of TP: BA. Each profile in (B) represents 300 mg of material. (C) BARDS spectra of un-milled

Theophylline and Benzoic Acid and a mixture of both.(D) PXRD data after 20 min of liquid assisted grinding for a 1:1 mixture of TP: BA and (E) the corresponding BARDS data.

The BARDS data in Figure 8 (E) also corresponds to the PXRD analysis in Figure 8 (D). The formation of a smooth curve and the disappearance of the BA plateau indicates the formation of the newly formed cocrystal.

Figure 9 (A and B) shows a clear trend is occurring in the formation of both cocrystals for INA: BA and TP: BA after 15 and 20 min of dry grinding and liquid assisted grinding, respectively. The formation of both 1:1 cocrystal forms is highlighted by the red (TP:BA) and the red and light blue (INA:BA) coloured data points. IR data for samples at all time points are available in the supplementary information.

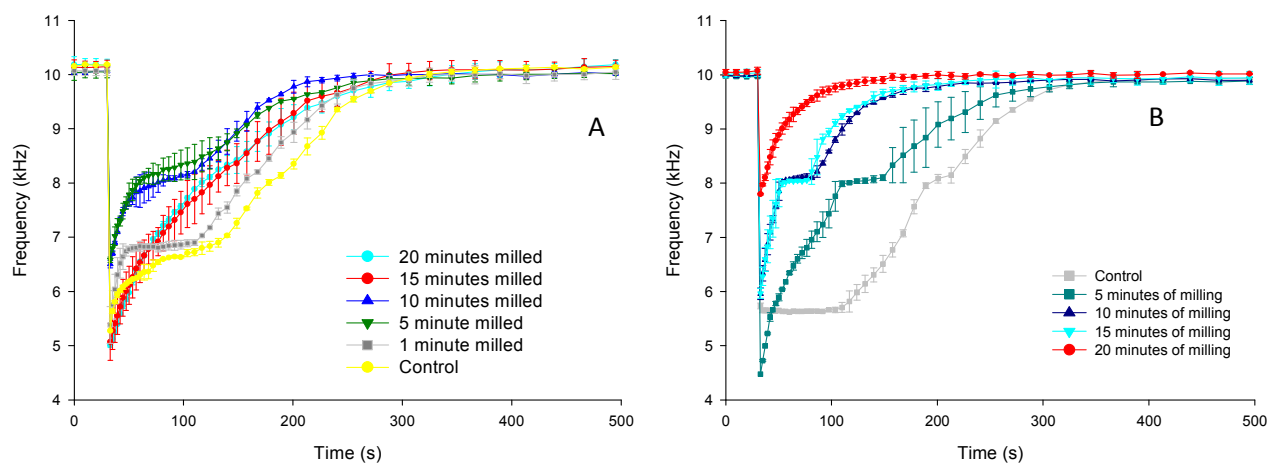


Figure 9: (A) An overlay for selected milling times of 1:1 INA: BA and (B) A similar overlay for 1:1 TP: BA.

Note: The BARDS data after 15 and 20 min of grinding indicates the formation of cocrystal as there is a direct overlay. This is also evident between the PXRD data between the experimental and reference spectra in Figures 7 (D) and 8 (D).

BA and INA still retain their characteristic spectrum in their crystallized form using the solvent evaporation method (Figure 10). The plateau is still present at 8 kHz with the crystal sample of BA, Figure 10 (A), and the INA has its characteristic smooth dissolution pattern in the form of a curve. The cocrystal PXRD pattern is not evident and there is a slight plateau at 8 kHz in the BARDS profile which indicates that benzoic acid is still present in its unbound form. The reason for this is due to INA being

less soluble in ethanol compared to BA. This could lead to the INA precipitating before binding to the BA.

This shows the importance of choosing the right solvent system in the formation of cocrystals using this method. When the crystals are ground down using a mortar and pestle the formation of a fine powder causes the dissolution of the crystal samples to occur at an accelerated rate leading to the return to the steady state frequency before the 100s mark figure 10 (B). This makes discrimination more difficult using BARDS.

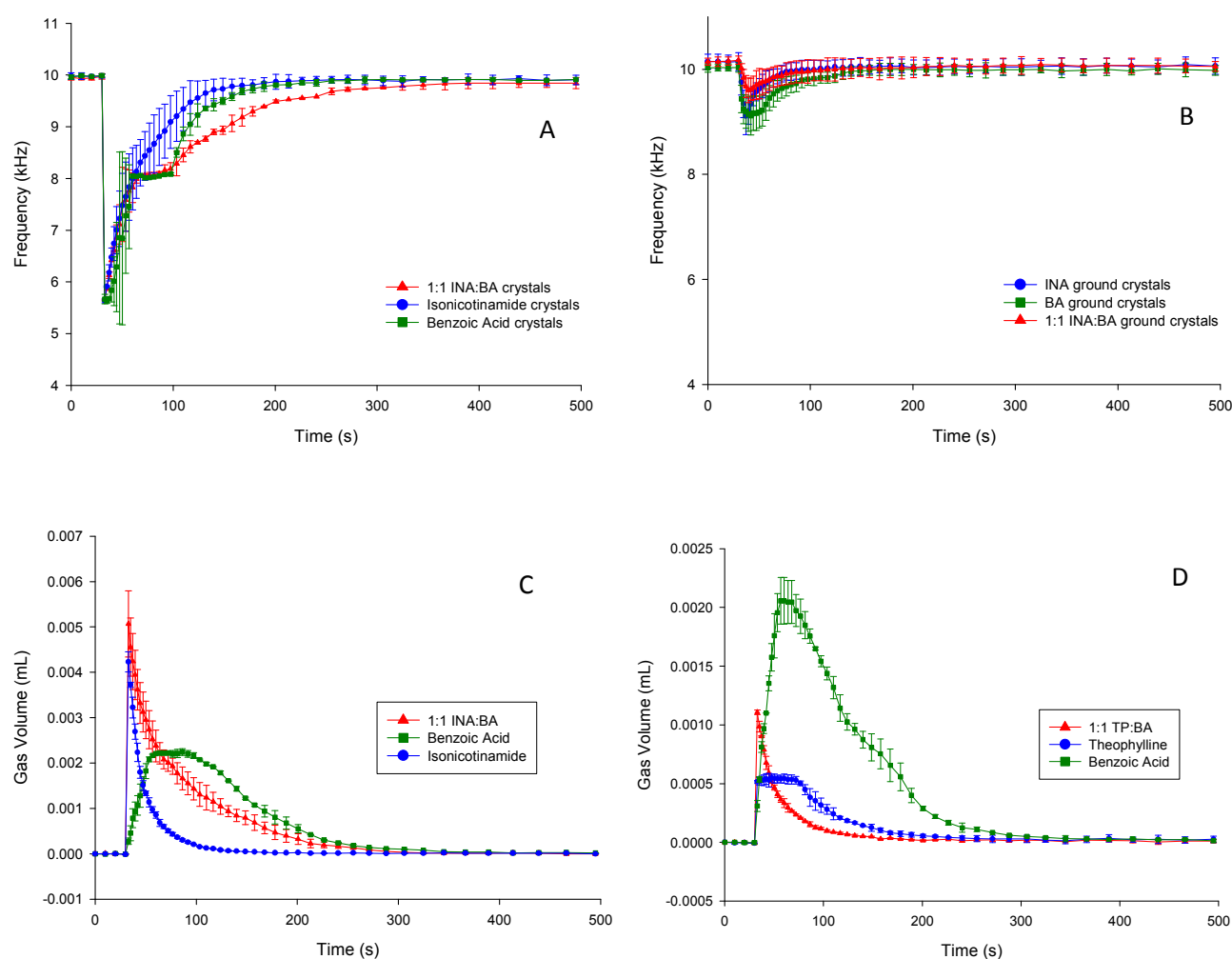


Figure 10: Comparisons of BARDS spectrum for the analysis of 300mg of crystallized 1:1 mixture of INA: BA. Note the formation of a smooth curve indicates the formation of the cocrystal for both. (C) Gas Volume plot after 15 min of solid state grinding for a 1:1 mixture of INA: BA and both cofomers. (D) Gas Volume plot after 20 min of liquid assisted grinding for a 1:1 mixture of TP: BA and both cofomers.

1 Finally, the data in Figures 7 (D) and 8 (D) have been parsed through the Crawford equation (ref 6,
2 equation 4) to produce gas volume time course plots for both cocrystal formations. The data gives
3 useful information on how gas is released, in real time, from the cocrystals during dissolution. Note that
4 the release of gas is intermediate between the two cofomers of each system. This is unique data which
5 IR or PXRD data does not yield. It provides data for kinetic analysis of dissolution rates and entrained
6 gas within the samples.
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16 DISCUSSION

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18 BARDS analysis of the 1:1 mixture at set time periods through the milling process shows the
19 changes in the acoustic response in the formation of both cocrystal systems, see figure 9 (A) and (B).
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21 The disappearance of the characteristic BA plateau and the formation of a smooth BARDS response
22 indicates the formation of both 1:1 cocrystals.
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28 The changes in compressibility of the solvent are evident during dissolution of the 1:1 mixture
29 through the changes of frequency of induced acoustic resonance with the magnetic stirrer. Steady state
30 frequency occurs at 10 kHz and decreases with the addition of the sample mixture after 30s. The
31 frequency subsequently rises back to the steady state frequency after the sample dissolves in the solvent.
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33 BARDS data in Figures 6 - 10 demonstrate that all samples yield significantly different acoustic
34 dissolution profiles and are highly reproducible. Different components of the blends produce intrinsic
35 spectra as shown in Figures 6 (A and B). A binary mixture results in frequency profiles which are
36 intermediate to those of the components when analyzed separately⁶.
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46 One of the strengths of BARDS is that a control sample (blend) can be made and tested. If a
47 50:50 binary mixture is made then samples from the blender should match the control spectra if
48 blending or milling has been carried out correctly. Small changes in mass of one component must be
49 made up by another component as a fixed sample mass is used for a BARDS experiment. This leads to
50 intrinsic sensitivity for the technique.
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CONCLUSION

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3 It has been shown that BARDS can track the formation of 1:1 cocrystal forms of INA: BA and TP:
4 BA which occurs after 15 and 20 min of solid state and liquid assisted grinding, respectively. BARDS
5 analysis has the potential for usage at-line to track co-crystal formation where BA is the conformer. It is
6 likely other cocrystals may be tracked in a similar way. The sample turnaround time is just a few
7 minutes with minimal sample preparation and method development. The gas volume time courses
8 during dissolution also provide unique insights into the solvation of the cocrystals and starting materials.
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12 The next step in the process of analyzing cocrystals using BARDS is to move to pilot scale analysis.
13 This would allow for sampling from the same batch rather than making up the samples individually and
14 testing them consecutively. Also, to use different conformers, other than benzoic acid, to identify if they
15 can also be tracked during cocrystal formation. Dissolution rates and kinetics will also be investigated.
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AUTHOR INFORMATION

Corresponding Author

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32 Email: d.fitzpatrick@ucc.ie. Tel.: +353 21 490 3299
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Notes

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37 Seán McSweeney and Dara Fitzpatrick are directors of BARDS Acoustic Science Labs.
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45 project.
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Supporting Information:

Infrared spectra for mixtures of API and co-former obtained at various milling times are available in the supporting information.

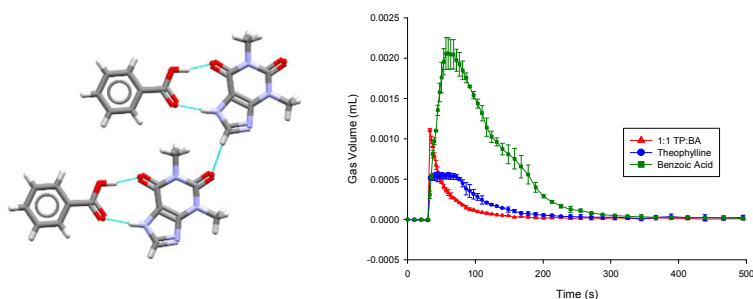
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Tracking Cocrystallization of Active Pharmaceutical Ingredients with Benzoic Acid Coformer using Broadband Acoustic Resonance Dissolution Spectroscopy (BARDS)

Christopher Kent, Uday B. Rao Khandavilli, Anas Alfarsi, Melissa Hanna-Brown, Seán McSweeney, Jacob Krüse, Simon Lawrence, and Dara Fitzpatrick.

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Co-crystals of Theophylline and Benzoic Acid produce unique gas evolution profiles during dissolution which can be used to screen for formation and purity.



Broadband Acoustic Resonance Dissolution Spectroscopy (BARDS) is a rapid method for characterising powders, blends and formulations. In this study, it is shown how dissolution of co-crystals produce reproducible gas evolution profiles which are distinct from the starting materials. This new approach has potential as a complimentary test method to confirm the presence of pure co-crystals or mixtures thereof.