

Article

# Chromium(III) Removal from Wastewater by Chitosan Flakes

Loris Pietrelli <sup>1,2,\*</sup>, Iolanda Francolini <sup>2</sup>, Antonella Piozzi <sup>2</sup>, Maria Sighicelli <sup>1</sup>,  
Ilaria Silvestro <sup>2</sup> and Marco Vocciante <sup>3</sup><sup>1</sup> ENEA, CR Casaccia, Via Anguillarese 301, 00060 Roma, Italy; maria.sighicelli@enea.it<sup>2</sup> Dipartimento di Chimica, Università di Roma Sapienza, 00185 Roma, Italy;

iolanda.francolini@uniroma1.it (I.F.); antonella.piozzi@uniroma1.it (A.P.); ilaria.silvestro@uniroma1.it (I.S.)

<sup>3</sup> Dipartimento di Chimica e Chimica Industriale, Università degli Studi di Genova, Via Dodecaneso 31, 16146 Genova, Italy; marco.vocciante@unige.it

\* Correspondence: loris.pietrelli@uniroma1.it

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**Featured Application:** The ability of chitosan as a low-cost and environmentally friendly Cr(III) adsorbent was studied to evaluate its potential application in the field of tannery wastewater treatment, in terms of the removal of chromium ions avoiding their conversion into Cr(VI), the compounds of which exert highly toxic and carcinogenic effects on biological systems.

**Abstract:** Chitosan is very effective in removing metal ions through their adsorption. A preliminary investigation of the adsorption of chromium(III) by chitosan was carried out by means of batch tests as a function of contact time, pH, ion competition, and initial chromium(III) concentration. The rate of adsorption was rather rapid ( $t_{1/2} < 18$  min) and influenced by the presence of other metal ions. The obtained data were tested using the Langmuir and Freundlich isotherm models and, based on  $R^2$  values, the former appeared better applicable than the latter. Chitosan was found to have an excellent loading capacity for chromium(III), namely 138.0 mg Cr per g of chitosan at pH = 3.8, but metal ions adsorption was strongly influenced by the pH. About 76% of the recovered chromium was then removed simply by washing the used chitosan with 0.1 M EDTA (Ethylenediaminetetraacetic acid) solution. This study demonstrates that chitosan has the potential to become an effective and low-cost agent for wastewater treatment (e.g., tannery waste) and in situ environmental remediation.

**Keywords:** chitosan; chromium; heavy metals; adsorption; kinetics; low-cost adsorbent; tannery; ion exchange

## 1. Introduction

The removal of heavy metal ions from aqueous solutions, both for pollution control and for the recovery of raw materials, has assumed increasing importance in recent years. Among the many metals potentially harmful to the environment and human health, chromium pollution is of considerable concern, as the metal is widely used in many industrial activities such as electroplating, leather tanning, nuclear power plants, and textile industries [1,2].

To address this problem, numerous separation techniques are currently available (e.g., adsorption, ion exchange, selective precipitation, nanofiltration, etc.), the selection of which, however, is far from trivial and deserves extensive attention to avoid a suboptimal choice or the failure of the reclamation activity [3]. In general, adsorption-based technologies have proven to be among the most viable alternatives proposed for the treatment of industrial wastewater contaminated by a wide variety of pollutants, both organic [4] and inorganic [5,6], due to the low processing and

instrumentation costs, the simplicity of operation, and the availability of different types of low-cost and environmentally friendly adsorbents. A wide range of materials, including activated carbon [7], metal oxides, carbon nanotubes, polymers, agricultural residues [8], and natural and modified clays [9], have been used successfully to adsorb heavy metals from aqueous solutions. This is even more interesting when it is possible to exploit low-cost adsorbent materials from industrial waste [10], with a double advantage for the environment, in line with operational guidelines such as the Circular Economy and the “near-zero discharge” of hazardous waste [11] established by the most recent European laws.

In this context, a very promising and inexpensive material is chitosan (poly- $\beta$ -(1 $\rightarrow$ 4)-2-amino-2-deoxy-D-glucose), a nitrogenous polysaccharide prepared from chitin by partially deacetylating its acetoamine groups using strong alkaline solutions at about 70 °C. Chitosan has a high potential for the adsorption of metal ions, since it has both amino and hydroxyl groups that can act as chelation sites for metal ions. One of the most interesting advantages of chitosan is its versatility, since the material can be easily physically modified to obtain different forms of polymers such as beads [12], membranes [13], or sponges [14] for different applications. Chitosan can also be easily chemically modified to increase its applications [15]. Recently, several critical reviews have been published on the many applications of chitosan as an environmentally friendly biomaterial [16], ranging from the medical field [17] to food technology [18] and environmental protection [19].

Chromium can be found in the environment in the forms Cr(III) and Cr(VI), as its other oxidation states are not stable in aerated aqueous media [20]. The trivalent state is the most stable form in reducing conditions and is present as a cationic species ( $\text{Cr}(\text{OH})^{2+}$ ,  $\text{Cr}(\text{OH})_2^+$ ), with the first or second hydrolysis products dominating at pH values from 4 to 8. The low solubility of  $\text{Cr}(\text{OH})_3$  ( $\log k = -16.19$ ) considerably limits the concentration of Cr(III) for pH values above about 5.

Given its high danger to biological systems, many studies have focused on the removal of Cr(VI), while very few articles deal with the adsorption of Cr(III) by chitosan. Maruca et al. [21] reported the uptake of Cr(III) ions by chitosan flakes and the effect of  $\text{PO}_4^{3-}$  on the adsorption mechanism. Chui and collaborators [22] studied the removal of Cr(III) using a packed column filled with crab chitosan. More recently, Singh & Nagendran [23] reported comparative studies on the sorption of Cr(III) on chitin and chitosan in terms of a comparison between Langmuir and Freundlich isotherms. Overall, the papers often concern the application of chitosan membranes or beads, and not of chitosan flakes.

Considering that equilibrium analysis is the most important fundamental study required to evaluate the affinity of a sorbent, the ability of chitosan to remove chromium(III) by adsorption was studied in the present work to evaluate its potential application in the field of tannery wastewater treatment [1]. Numerous adsorption tests of chromium(III) on chitosan flakes were conducted to investigate the effects of contact time, pH, initial Cr(III) concentration, and, using real wastewater, ion competition. The thermodynamic behavior was assessed using the well-known Langmuir and Freundlich isotherm models.

## 2. Materials and Methods

### 2.1. Material and Reagents

Chitosan (molecular weight: 400 k, 66.9% <40 mesh, degree of deacetylation: 84–86%) was provided by Merck KGaA (Darmstadt, Germany) and used without further purification. Its surface area, estimated with the nitrogen adsorption method (BET, Quantachrome Nova 2200, Quantachrome Instruments, Boynton Beach, FL, U.S.), was equal to  $1.578 \text{ m}^2 \text{ g}^{-1}$ . The water content of this commercial chitosan, determined by thermogravimetric analysis (TGA1, Mettler Toledo, OH, U.S.), was 12.7%, while its decomposition temperature was 292.12 °C (Figure 1). In order to study the effect of particle size on the adsorption of metal ions, two granulometric fractions (<0.42 mm and >0.42 mm) were obtained from the starting material by using a sieve shaker. Analytical grade chemicals were supplied by Merck Co. (Kenilworth, NJ, U.S.); aqueous solutions were prepared dissolving  $\text{Cr}(\text{NO}_3)_3$  at different

concentrations in deionized water (Millipore Milli-Q, Merck KGaA, Darmstadt, Germany), and the initial pH was adjusted by adding a few drops of HNO<sub>3</sub> and NaOH solutions.

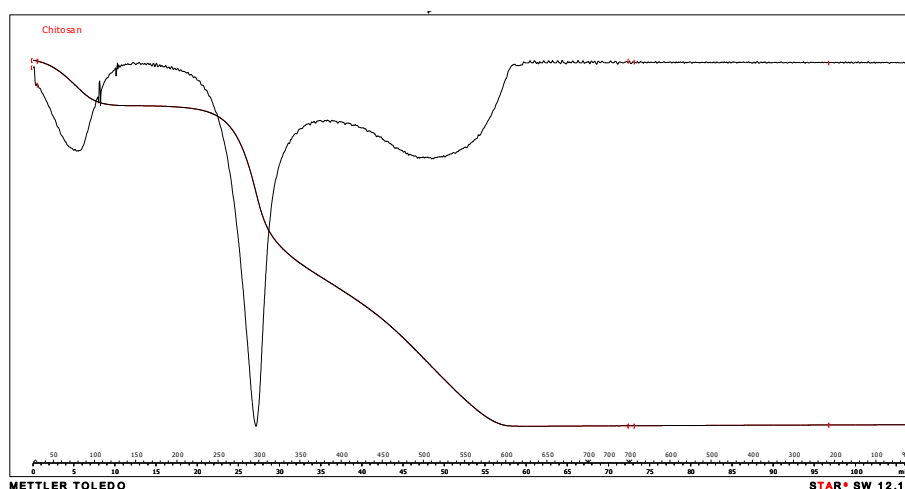


Figure 1. Thermogravimetric analysis (TGA) of the chitosan sample.

## 2.2. Adsorption Tests

Batch experiments (in triplicate) were carried out using 100 and 500 mg of adsorbent each time; chitosan was added to 50 mL of Cr solution in a conical flask. The stirring rate was set at 120 rpm for all adsorption/desorption tests using a temperature-controlled magnetic stirrer. Analysis of metal ions was carried out using an ICP-optical emission spectrometer (Inductively Coupled Plasma, Optima 2000 DV, Perkin Elmer, Waltham, MA, U.S.).

For experiments on the pH effect, the solutions were initially adjusted with aqueous solutions of acid or base (0.01 M HNO<sub>3</sub> and/or 0.01 M NaOH) to reach pH values between 0.5 and 5, and thus avoid the precipitation of Cr(OH)<sub>3</sub>. Isotherms were recorded during the execution of adsorption experiments with various initial metal concentrations ( $C_0 = 50\text{--}2000\text{ mg L}^{-1}$ ) at 20 °C. Kinetic tests were performed using 100 mg of chitosan flakes, 100 and 500 mg L<sup>-1</sup> as the initial metal concentration at 20 °C, and pH = 3.8 for fixed time intervals during adsorption ( $t = 0\text{--}24\text{ h}$ ). The effect of the granulometry of the flakes on the adsorption capacity was investigated using the two granulometric fractions of chitosan (<0.42 mm and >0.42 mm) under the same conditions. For all adsorption batch tests, a contact time of 120 min was set.

The equilibrium amount of metal in the solid phase, expressed as  $Q_e$  (mg g<sup>-1</sup>), was determined with reference to the mass balance equation:  $Q_e = (C_0 - C_e) \times (V/m)$ , where  $C_0$  and  $C_e$  (mg L<sup>-1</sup>) are the initial and equilibrium metal concentrations, respectively,  $V$  (L) is the volume of the aqueous solutions, and  $m$  (g) is the mass of the adsorbent.

Dynamic tests were performed using a glass column with an internal diameter of 0.9 cm and a bed high of 40 cm, filled with 2 g of chitosan. Tests were performed using both a real wastewater solution and a 500 mg L<sup>-1</sup> chromium(III) solution at pH = 3.5, imposing a flow rate of 23.6 mL h<sup>-1</sup> (1 BV h<sup>-1</sup>) by using a peristaltic pump.

The real wastewater, a tannery washing solution, had the following composition: pH = 3.2, Chemical Oxygen Demand, COD = 9.1 g L<sup>-1</sup>, Total Suspended Solids, TSS < 1 g L<sup>-1</sup>, Cr<sup>3+</sup> = 635 mg L<sup>-1</sup>, Na<sup>+</sup> = 1050 mg L<sup>-1</sup>, Mg<sup>2+</sup> = 760 mg L<sup>-1</sup>, Ca<sup>2+</sup> = 300 mg L<sup>-1</sup>, Zn<sup>2+</sup> = 115 mg L<sup>-1</sup>, Cd<sup>2+</sup> = 87 mg L<sup>-1</sup>, SO<sub>4</sub><sup>2-</sup> = 1.820 mg L<sup>-1</sup>, and Cl<sup>-</sup> = 818 mg L<sup>-1</sup>.

Desorption experiments were performed in batch mode ( $T = 20\text{ °C}$ ,  $t = 24\text{ h}$ ). In particular, after the end of the adsorption phase, the adsorbent material was separated from the supernatant using filtration membranes (0.22 μm). Then, desorption tests were performed using 50 mL solutions of H<sub>2</sub>SO<sub>4</sub> and EDTA as desorption reagents at a concentration of 0.1 M and 0.05 M, respectively.

The quantitative evaluation of desorption was carried out using desorption percentages calculated from the difference between the amount of metal loaded on the adsorbent after adsorption and the amount of metal in solution after desorption. To investigate the reuse capacity of the adsorbents, the above procedure was repeated 5 times under the same conditions (first adsorption and then desorption).

### 3. Results

#### 3.1. Adsorption Dynamics

The uptake of metal ions from the solution involves several steps, necessary for the transfer of the solute from the liquid phase to the specific sites within the chitosan particles (e.g., external diffusion and intraparticle diffusion).

In the case of chitosan, its chains have a large number of the  $-NH_2$  and  $-OH$  groups distributed throughout the structure, making the kinetic or mass transfer representation likely to be global. The  $-NH_2$  groups are the most important binding sites for metal ions [24], yet the hydroxyl groups can also contribute as coordinator groups, especially those in C-3 position [15,22]. To examine the adsorption mechanism of the metal ion of interest, two kinetic models were tested:

- i. the pseudo-first-order equation described by Lagergren [25], which can be rearranged to obtain a linear form as shown by Equation (1):

$$\text{Log}(q_e - q_t) = \text{Log}(q_e) - (k_1/2.303)t \quad (1)$$

- ii. a pseudo-second-order equation based on the equilibrium adsorption capacity, which can be expressed as in Equation (2):

$$t/q_t = (1/k_2q_e^2) + (1/q_e)t \quad (2)$$

In the above equations,  $q_e$  ( $\text{mg g}^{-1}$ ) represents the quantity of Cr(III) adsorbed when the system is at equilibrium,  $q_t$  ( $\text{mg g}^{-1}$ ) is the quantity of Cr(III) adsorbed at time  $t$ , and  $k_1$  ( $\text{min}^{-1}$ ) and  $k_2$  ( $\text{g mg}^{-1} \text{min}^{-1}$ ) are the rate constants of the pseudo-first and pseudo-second order kinetic models, respectively.

Given that Equations (1) and (2) are not able to provide information on the adsorption mechanism, the simplified intraparticle diffusion model [26] was also tested, being  $k_i$  ( $\text{g mg}^{-1} \text{min}^{-1}$ ) the rate constant of the model:

$$q_t = k_i t^{1/2} \quad (3)$$

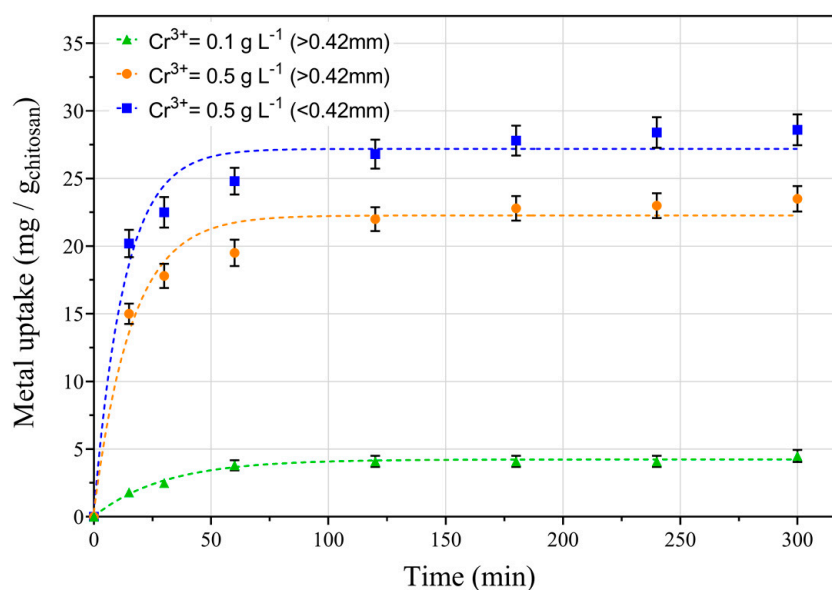
The validity of these models was assessed by analyzing the slopes and intercepts of  $\text{Log}(q_e - q_t)$  vs.  $t$ ,  $t/q_t$  vs.  $t$ , and  $q_t$  vs.  $t^{1/2}$  for each of the linearized equations.

The results obtained with different concentrations of chromium(III) are shown in Table 1 in terms of correlation coefficients ( $R^2$ ) as well as calculated and experimental adsorption capacity values.

**Table 1.** Values of the adsorption kinetic constants at  $T = 20$  °C,  $\text{pH} = 3.8$ ,  $C_0 = 0.5 \text{ g L}^{-1}$ .

1st Order Kinetic Model			2nd Order Kinetic Model			Intraparticle Diffusion Model	
$k_1$ ( $\text{min}^{-1}$ )	$q_{e, cal}$ ( $\text{mg g}^{-1}$ )	$R^2$	$k_2$ ( $\text{g mg}^{-1} \text{min}^{-1}$ )	$q_{e, cal}$ ( $\text{mg g}^{-1}$ )	$R^2$	$k_i$ ( $\text{g mg}^{-1} \text{min}^{-1}$ )	$R^2$
$1.07 \times 10^{-2}$	9.32	0.7838	$5.07 \times 10^{-3}$	23.89	0.9987	1.711	0.9199

The correlation coefficient  $R^2$  for the pseudo-second-order adsorption model was the highest and, in fact, its estimate of the equilibrium adsorption capacity  $q_{e, cal}$  was quite close to the experimental  $q_t$  values (23–28.5  $\text{mg g}^{-1}$  as shown in Figure 2 for the two grain-sizes). These results suggest that a second-order mechanism is predominant and that the overall Cr(III) adsorption rate is controlled by a chemisorption process.



**Figure 2.** Effect of the grain size of the flakes on the adsorption of Cr(III) on chitosan (100 mg), at pH = 3.8, T = 20 °C, and C<sub>0</sub> = 100 and 500 mg L<sup>-1</sup> of Cr(III).

### 3.2. Grain Size Effect

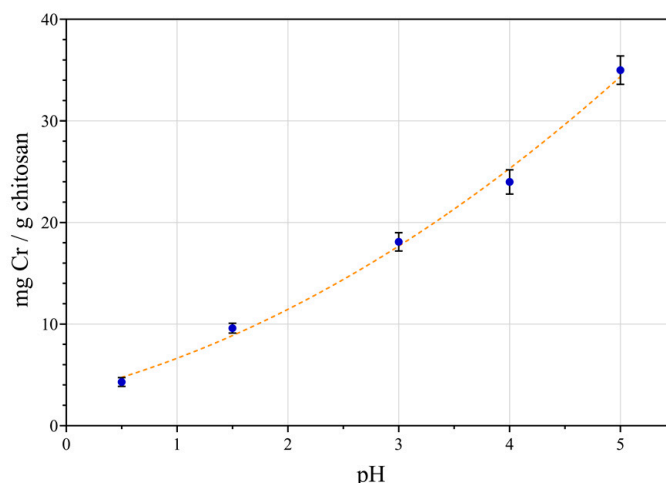
Figure 2 shows the effect of the grain size of the flakes on the adsorption capacity at pH = 3.8 and T = 20 °C. It can be observed that the metal uptake was higher on particles with a small size (<0.42 mm). This is likely due to the higher surface area exposed by these particles, which favors the removal of Cr(III) from the solution in the initial stages of the adsorption process. This phenomenon, previously reported for the adsorption on chitin [21], chitosan [21,27,28], and Neem sawdust [29], was further improved by the ability of metal ions to penetrate into the internal structure of chitosan.

Figure 2 also confirms that the adsorption process was rather rapid, with  $t_{1/2} < 18$  min and the maximum adsorption obtained in about 120 min. The reference time for the subsequent equilibrium tests was thus set at 120 min, an adequate compromise between accuracy and speed in the execution of experimental tests.

### 3.3. Effect of pH

Figure 3 shows the effect of pH on Cr(III) adsorption on chitosan. Notably, the pH of the solution strongly affected the adsorption of metal ions, with the latter increasing with the pH of the solution. Under acidic conditions, the amino groups (R-NH<sub>3</sub><sup>+</sup>) and the hydroxyl groups (R-OH<sub>2</sub><sup>+</sup>) are protonated and the molecule is a sort of polycation, with a reduced number of binding sites available for the adsorption of Cr(III); according to [30], the pK<sub>a</sub> of the amine groups is 6.3. In addition, the positive surface charge may hinder the adsorption of metal ions. On the contrary, a high pH will favor their adsorption since the nitrogen free electron doublet is responsible for cations coordination. Considering that  $K_{ps} = [Cr^{3+}] \times [OH]^{-3} = 6.7 \times 10^{-31}$ , chromium(III) hydroxide begins to precipitate at pH ≈ 6.5; for pH values higher than 3.8, there is a significant reduction of the Cr(III) fraction, with formation of Cr(OH)<sup>2+</sup> and Cr(OH)<sub>2</sub><sup>+</sup> hydrolyzed complex species [20]. The result is an increase in chromium adsorption due mainly to hydrolyzed forms. Therefore, the adsorption of metal ions is mainly due to the electrostatic interactions between counter ions.

As previously reported [21], the final pH values of the equilibrated solutions were higher as the Cr(III) concentration became smaller (Table 2). This is probably due to the fact that Cr ions are Lewis acids; therefore, the lower the concentration, the higher the pH (Figure 3). Moreover, the chromium adsorption capacity increases by increasing the metal ions concentration (see Section 3.4), which causes a greater competition with H<sup>+</sup> protons.



**Figure 3.** Effect of pH (initial value) on chromium(III) adsorption on chitosan ( $C_0 = 500 \text{ mg L}^{-1}$ ).

**Table 2.** pH variation vs. Cr(III) concentration ( $\text{pH}_{\text{in}} = 3.8$ ).

Cr(III) (ppm)	pH
50	5.16
100	5.02
200	4.98
400	4.67
500	4.47

As already stated, it is accepted that chitosan amino groups are the main reactive sites for metal ions and that hydroxyl groups (in particular in C-3 position) may contribute to sorption. Metal sorption may involve different mechanisms (chelating, electrostatic attraction) depending on the pH, the solution, and the metal (concentration, speciation, etc.). Protonation of the amino groups at acidic pH increases the adsorption of anionic species, while cations interactions increase with the pH due to deprotonation of amino/hydroxyl groups. Moreover, increasing the ions competition, due to increase in metal concentration, the pH decreases. The fraction of free (accessible) amine groups is the key parameter.

### 3.4. Adsorption Isotherms

To determine the maximum adsorption capacity of Cr(III) on chitosan, a study was carried out on the adsorption isotherm by comparing the most common models; in particular, data were analyzed using the Langmuir and Freundlich equations:

$$q_e = Q^\circ k_L C_{eq} / (1 + k_L C_{eq}) \quad \text{or, linearized} \quad 1/q_e = 1/(Q^\circ k_L) (1/C_{eq}) + 1/Q^\circ \quad (4)$$

$$q_e = k_F C_{eq}^{1/n} \quad \text{or, linearized} \quad \text{Log}(q_e) = \text{Log}(k_F) + 1/n \text{Log}(C_{eq}) \quad (5)$$

where  $q_e$  ( $\text{mg g}^{-1}$ ) is the amount of Cr(III) on the solid phase at equilibrium, and  $C_{eq}$  ( $\text{mg L}^{-1}$ ) is the equilibrium concentration of Cr(III) in the aqueous phase. According to Langmuir's equation,  $Q^\circ$  ( $\text{mg g}^{-1}$ ) is the amount of Cr(III) required for a complete coverage of available adsorption sites, while  $k_L$  is an empirical coefficient related to the affinity of adsorption sites for the adsorbed species. With reference to Freundlich's equation,  $k_F$  and  $n$  are empirical constants representing the adsorption capacity and adsorption intensity, respectively; all parameters can be estimated through the intercepts and slopes of the linearized forms of isotherm equations.



The essential characteristics of the Langmuir equation can be expressed in terms of a dimensionless separation factor  $R_L$ , which has been defined in [27] as:

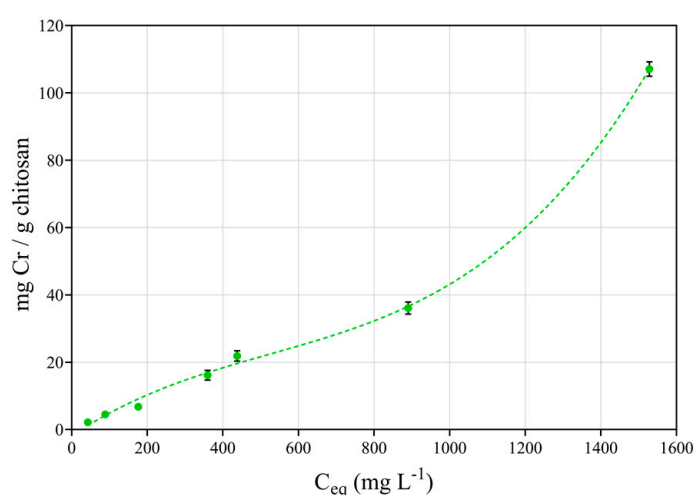
$$R_L = 1/[1 + (k_L C_0)] \quad (6)$$

where  $C_0$  is the highest initial Cr(III) ion concentration ( $\text{mg L}^{-1}$ ).  $R_L$  is related to the shape of the isotherm: the adsorption is unfavorable if  $R_L > 1$ , favorable if  $0 < R_L < 1$ , irreversible if  $R_L = 0$ , and linear if  $R_L = 1$ . All the estimated isotherm parameters are reported in Table 3. The  $R_L$  value confirms the affinity between chitosan and chromium ions, and the adsorption equilibrium data correlate well with the Langmuir isotherm equation, with a maximum adsorption capacity estimated at  $138.04 \text{ mg g}^{-1}$ . This implies a monolayer interaction of chromium on the adsorbent [31]. Ngah and colleagues [23] found that  $Q^\circ$  was  $30.03 \text{ mg g}^{-1}$  using Cr(III) in the range  $4\text{--}14 \text{ mg L}^{-1}$  and identified the Langmuir isotherm as the best model for the adsorption on cross-linked chitosan. However, the chromium concentrations considered were much lower than those investigated in the present work, and the number of amino groups available for ions coordination in cross-linked chitosan is limited, which obviously results in a lower adsorption capacity. Eiden and colleagues [32] found that  $Q^\circ$  was  $62 \text{ mg g}^{-1}$  for chitosan flakes at  $\text{pH} = 4$ , but no information regarding chitosan characteristics was provided (especially regarding its degree of deacetylation).

**Table 3.** Langmuir's and Freundlich's isotherm parameters at  $20^\circ\text{C}$ ,  $C_0 = 0.5 \text{ g L}^{-1}$ , and  $\text{pH} = 3.8$ .

Langmuir			Freundlich			
$Q^\circ (\text{mg g}^{-1})$	$k_L$	$R^2$	$R_L$	$k_F (\text{mg g}^{-1})$	$1/n$	$R^2$
138.04	$3.7 \times 10^{-4}$	0.9925	0.575	0.035	1.061	0.9694

Indeed, the basis of the high adsorption capacity found in our study lies precisely in the fact that the investigated chromium concentrations are much higher than those reported in other works (see Figure 2, which shows that adsorption on chitosan increases with the concentration of the target species). Although the different experimental conditions make it difficult to compare the results obtained, it can be observed that at lower concentrations, the adsorption is in line with the values found in the literature (Figures 2 and 4).



**Figure 4.** Adsorption isotherm at  $20^\circ\text{C}$ ,  $C_0 = 50\text{--}2000 \text{ mg L}^{-1}$ , and  $\text{pH} = 3.8$ .

### 3.5. Desorption

After Cr(III) adsorption, the chitosan flakes were washed thoroughly with deionized water and treated with the desorption agents. Desorption tests were performed using  $\text{H}_2\text{SO}_4$  and EDTA as

desorption reagents; the chromium desorption efficiency of chitosan flakes for different washing solutions is reported in Table 4.

**Table 4.** Desorption efficiency (%) for each adsorption cycle.

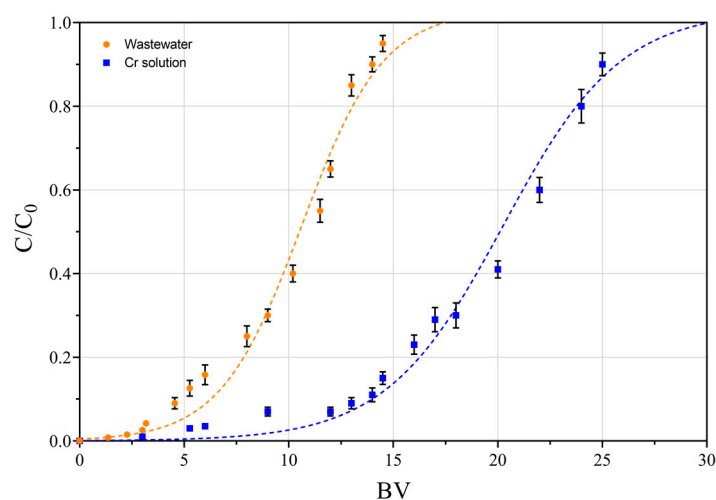
Reagent	1st Cycle	2nd Cycle	3rd Cycle	4th Cycle	5th Cycle
EDTA 0.05 M	53.3	-	-	-	-
EDTA 0.1 M	76.5	73.4	70.6	71.7	72.1
H <sub>2</sub> SO <sub>4</sub> 0.05 M	37.8	-	-	-	-
H <sub>2</sub> SO <sub>4</sub> 0.1 M	45.3	46.5	40.7	43.1	39.7

As reported in [23], EDTA is an efficient desorption agent: being a hexadentate chelating agent, it is capable to form a strong complex with Cr(III) ions. Considering the pH effect on the chromium adsorption, sulfuric acid (such as other acidic media) was also considered as a reagent potentially able to remove Cr(III) ions from the chitosan polymer.

Despite being efficient in terms of chitosan regeneration, EDTA persists in municipal wastewater treatments, making its use in technical applications potentially unwelcome. Although future research may investigate other chelating agents that are biodegradable and more suitable for wastewater treatment plants, the current approach has considered the possibility of treating the spent solution to recover EDTA by precipitation in acidic media as a precautionary measure.

### 3.6. Dynamic Tests

The effect of competing ions can be observed from the breakthrough curves shown in Figure 5. The inhibition effect of anions such as chlorides and sulfates has been reported for Cr(VI) [33]; in addition, it has been reported that chitosan forms complexes with transition metal ions, but not with complexes with alkali and alkali earth metal ions due to the absence of *d* and *f* unsaturated orbitals [34,35]. Therefore, the presence of Na, Mg, and Ca does not reduce the chromium adsorption. In contrast, the removal of chromium is influenced by the presence of zinc and cadmium due to their affinity with chitosan [36]. When real wastewater was considered, chitosan was able to absorb 104.7 mg g<sup>-1</sup> of chromium, while 127.5 mg g<sup>-1</sup> were adsorbed using a chromium solution.



**Figure 5.** Breakthrough tests performed with real wastewater solution (pH = 3.2) and 500 mg L<sup>-1</sup> chromium(III) solution (at pH = 3.5), using a flow rate of 1 BV h<sup>-1</sup> (BV = 23.6 mL).

The breakthrough curve should be symmetrical and sigmoid in shape; however, a deviation from a symmetrical S-shape and not very steep curves can be observed in Figure 5. The packed column was probably not very homogeneous (chitosan flakes seem to be inappropriate to be used in a packed



column) and the flow-rate was too fast, causing physical non-equilibrium processes and high mass transfer zone [37].

#### 4. Conclusions

The present study indicated the suitability of chitosan in applications aimed at removing chromium(III) ions from aqueous solutions. Chitosan is a low-cost reagent and its utilization is novel, non-toxic, and environmentally compatible. Although its effectiveness has been proven in the adsorption of several metal ions, such as Cu, Ag, Cd, Zn, Mo, V, Pb, and Cr(VI), Cr(III) has been poorly investigated. Moreover, among the usable forms of chitosan, flakes are the least studied. In this work, the adsorption of Cr(III) on chitosan flakes was investigated considering concentrations (up to 2000 mg L<sup>-1</sup>) that are much higher than those reported in the few studies present in literature. This made it possible to highlight a much higher adsorption capacity of chitosan (up to 138.45 mg g<sup>-1</sup>), well beyond the values observed so far.

The pH significantly influences the adsorption capacity of the biopolymer, the latter increasing with the pH of the solution. The adsorption process is rather rapid: it was found that 2 h of contact time are sufficient to reach about 95% of the adsorption equilibrium; the Langmuir equation provided the best fit over the entire concentration range, thus suggesting a monolayer interaction of chromium on the adsorbent.

The adsorption capacity for Cr(III) in dynamic tests was found to be penalized by the competition between ions, and particularly influenced by the presence of zinc and cadmium due to their strong affinity with chitosan. The recovery of Cr(III) ions (and the consequent regeneration of chitosan) can be obtained by desorption with a 0.1 M EDTA solution, with the latter being recoverable by precipitation in an acidic medium.

Since chitosan is basically a low-cost and environmentally friendly material, its use as it is produced (i.e., flakes) is desirable, especially considering the greater theoretical adsorbent capacity due to a major availability of both amino and hydroxyl groups, but involves some critical issues. In particular, since the shape and size characteristics of commercial chitosan flakes introduce hydrodynamic limitations such as column clogging, a batch reactor is probably more appropriate as a experimental setup. Further investigations are therefore underway to quantify the effectiveness of alternative operating solutions.

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