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Self-Assembly Properties of Amphiphilic Iron(III) Spin Crossover Complexes in Water and at the Air–Water Interface

Paulo N. Martinho ^{1,†}, Irina A. Kühne ¹, Brendan Gildea ¹, George McKerr ², Barry O'Hagan ², Tia E. Keyes ³, Tibebe Lemma ³, Claudio Gandolfi ⁴, Martin Albrecht ^{1,4,‡} and Grace G. Morgan ^{1,*}

¹ School of Chemistry, University College Dublin (UCD), Belfield, Dublin, D4, Ireland; paulo.martinho@ucd.ie (P.N.M.); irina.kuhne@ucd.ie (I.A.K.); brendan.gildea@ucdconnect.ie (B.G.); martin.albrecht@dcb.unibe.ch (M.A.)

² School of Biomedical Sciences, University of Ulster, Cromore Road, Coleraine BT52 1SA, UK; g.mckerr@ulster.ac.uk (G.M.); bmg.ohagan@ulster.ac.uk (B.O.)

³ School of Chemical Sciences, Dublin City University, Dublin, D9, Ireland; tia.keyes@dcu.ie (T.E.K.); tlemma@gmail.com (T.L.)

⁴ Department of Chemistry, University of Fribourg, Ch. du Musée 9, CH-1700 Fribourg, Switzerland; claudio.gandolfi@unifr.ch

* Correspondence: grace.morgan@ucd.ie; Tel.: +353-1-716-2295

† Present Address: University of Lisboa, Faculty of Sciences, BioISI—Biosystems & Integrative Sciences Institute, Campo Grande, C8 bdg, 1749-016 Lisboa, Portugal and Department of Inorganic Chemistry, Faculty of Science, Charles University, 128 43 Prague 2, Czech Republic.

‡ Present Address: Department of Chemistry and Biochemistry, University of Bern, Freiestrasse 3, CH-3012 Bern, Switzerland.

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Abstract: The assembly properties of three known spin crossover iron(III) complexes 1–3, at the air–water interface, are reported. All three complexes are amphiphiles, each bearing a pair of C_n alkyl chains on the polyamino Schiff base salztrien ligand (*n* = 6, 12, or 18). Complex 1 is water-soluble but complexes 2 and 3 form Langmuir films, and attempts were made to transfer the film of the C₁₈ complex 3 to a glass surface. The nature of the assembly of more concentrated solutions of 3 in water was investigated by light scattering, cryo-SEM (scanning electron microscopy), and TEM (transmission electron microscopy), all of which indicated nanoparticle formation. Lyophilization of the assembly of complex 3 in water yielded a powder with a markedly different magnetic profile from the powder recovered from the initial synthesis, notably, the spin crossover was almost completely quenched, and the thermal behavior was predominantly low spin, suggesting that nanoparticle formation traps the system in one spin state.

Keywords: spin crossover; iron(III) amphiphile; micelle; Langmuir isotherm; Raman spectroscopy; lyophilization

1. Introduction

The continuing pressure to intensify hard drive bit densities has prompted research into molecular alternatives to the current methodology of using nanoislands of magnetic metal oxides, which are now starting to reach the superparamagnetic size limit [1,2]. Molecular switching phenomena which could be harnessed for this purpose include changes in redox state [3–5], spin state [6–8], conformation [9–11], or shape [12,13]. Spin state change is a particularly appealing vehicle for molecular switching, as it is often accompanied by significant hysteresis [14,15], and synthetic routes to spin transition systems tend to be relatively facile. In addition, there exist many external

perturbation methods to effect a spin state change including thermal [16–18], irradiation [19–21], magnetic field [22,23], pressure [24–26], or electric [27,28] inputs, and an equally wide range of detection methods, including measurement of magnetic properties [29], vibrational [21,30,31] and electronic [32–34] spectroscopic signatures, conductance [35], or surface plasmonic response [36]. This activity has intensified research into assembly protocols for spin crossover (SCO) complexes, and methods to prepare thin films [37,38], nanocrystals [39,40], and nanoparticles [41–46] have all been reported in recent years. Ligand design is important in preparing a SCO complex for a particular assembly method, and we have concentrated our design efforts into modification of the Fe^{III}salztrien SCO cation by attachment of one or more alkyl chains onto both phenoxy rings [47–49], or onto each of the secondary amine groups [50–52]. Such amphiphilic complexes show markedly different magnetic behavior to the parent complex in both the solid state and in solution, and these and related Mn(III) SCO systems [53], may be candidates for Langmuir–Blodgett (LB) [54,55] monolayer formation and transfer to a surface.

Generally, SCO properties of many alkylated Fe(II) complexes have also been reported [56–71], with fewer examples in Fe(III) [72], and these may show a variety of solution and solid state properties [67–71], including some with a wide thermal SCO hysteresis [67]. Clérac et al. have recently reported an Fe(II) SCO complex decorated with long alkyl chains with a wide thermal hysteresis around room temperature (RT), demonstrating that a spin transition and the presence of a monotropic phase leads to magnetic tristability at RT [73]. In crystalline phases, entropy-driven conformational isomerization of alkyl chains, such as *anti/gauche* changes or order/disorder transitions, can also affect magnetic properties and promote spin state switching, as strikingly illustrated in a Co(II) compound featuring a “reverse” spin transition [74].

We are interested in the potential for depositing SCO amphiphiles on surfaces using the LB technique [54,55], since it provides precise control of the monolayer thickness, as well as homogeneous deposition of the monolayer over large areas and the possibility to make multilayer structures with varying layer composition [75,76]. In 1988, Kahn et al. reported the first fabrication of thin films of SCO compounds using LB techniques [77], and Clérac et al. showed, in 2004, that the fabrication of LB films of triazole-based polymers which exhibit SCO behavior is possible [78]. The first magnetic characterization of the SCO in an LB film has been reported for a semi-fluorinated Fe(II) compound, which exhibits similar bulk and thin film magnetic properties [79]. In addition, light-induced excited spin state trapping (LIESST) has been observed for an iron(II) compound organized as an LB film, where the LIESST population and relaxation are faster for the LB film than for the powder [80]. Most of the recent studies on SCO amphiphiles concentrate on Fe(II)- and Co(II)-containing coordination centers [81–85]. Recently, Brooker et al. showed a close to ideal transfer ratio of Langmuir films onto a glass support in the first upstroke using an Fe(II) SCO compound [86], but in the subsequent downstroke, the first layer was quantitatively desorbed onto the air–water interface, suggesting only moderate adhesion of the complex to the support.

Amphiphilic iron(III) complexes, **1–3** [50–52,87], with C₆, C₁₂, and C₁₈ appended chains, all show solid state spin crossover, and we now report their Langmuir isotherm behavior and the results of attempts to transpose these to a glass surface. The nature of the assembly, **3** in aqueous solution, is also probed by scanning and transmission electron microscopy, and the magnetic properties of the powdered sample of **3** recovered by lyophilization are compared with those of the initial sample.

2. Results

Amphiphilic complexes **1–3**, Figure 1, were prepared as described elsewhere [50], and the potential for self-assembly as a Langmuir monolayer at the air–water interface was assessed for all three. An important criterion for monolayer formation is water insolubility, and the solubility of all three was initially assessed visually. The *bis*-C₆ complex **1** readily formed a purple solution in water, but the longer *bis*-C₁₂ and *bis*-C₁₈ analogues **2** and **3** were insoluble under ambient conditions. In line with these observations, complex **1** does not form a Langmuir film, Figure 1, presumably due to its solubility in water. Such behavior was observed previously for the Fe^{III}salztrien derivative with C₆ chains appended on the phenolate ring [47–49].

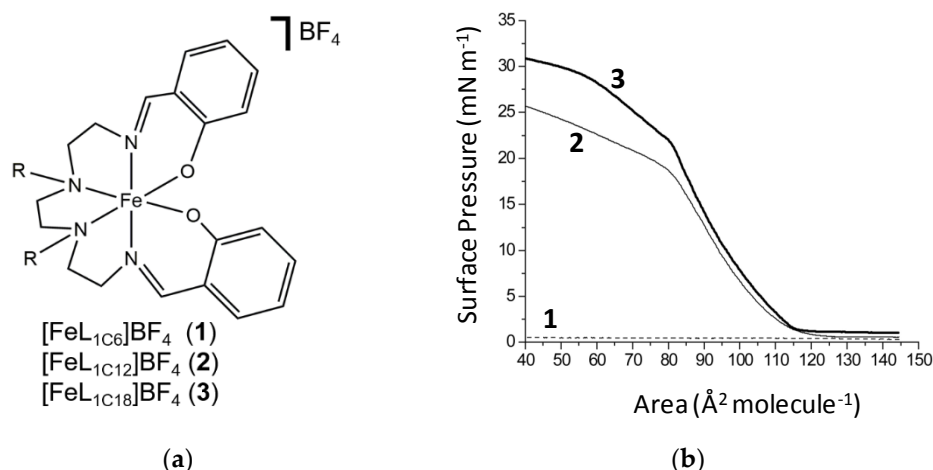


Figure 1. Schematic structure of 1–3 (a) and pressure–area isotherms for complexes 1–3 showing chain length dependence (b).

By contrast, compounds 2 and 3, which are water-insoluble, form Langmuir monolayers at the air–water interface, and the corresponding pressure–area isotherms show that the packing process is rather gradual, Figure 1. The area occupied by one surfactant in the Langmuir monolayer is estimated at 85 \AA^2 per molecule at the maximum packing density, given that this is the area/molecule at which the film collapses for both complexes. The accessible surface pressure of $20\text{--}25 \text{ mN m}^{-1}$ is, therefore, low, and the condensed phase for both compounds begins at 115 \AA^2 per molecule, and is sharply defined, in particular, for complex 3. Studies of the stability of the monolayers formed for both 2 and 3 were performed, Figure 2. From this, it can be concluded that both compounds form stable monolayers over a period longer than 90 min.

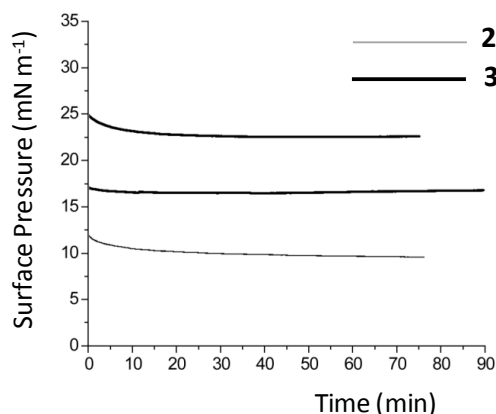


Figure 2. Time-dependence of film stability for complexes 2 and 3 at a fixed barrier position. The stability of complex 3 was measured at two different surface pressures.

Comparison of the behavior of compounds 1–3, which have alkyl chains appended to the NH backbone, with analogues where the alkyl chains are appended on the phenolate donors [48], allows us to conclude that similar maximum packing density is reached in both series. The pressure–area isotherms for *N*-functionalized complexes are less steep, indicating less ordered arrangements, possibly due to backfolding of the alkyl chains. Both types of complexes (*N*-functionalized and *O*-functionalized) do not show clear evidence of the transition from the condensed phase to the solid phase, before gradually collapsing. Lower collapse pressures for compounds 2 and 3, and the fact that the surface pressures start to rise at a larger area per molecule, might be attributed to the reduced amphiphilicity conferred by substitution on the NH backbone, rather than the phenolate ring. Besides

this, no differences in terms of time stability were observed between the two types of Fe(III) amphiphiles.

Attempts to transfer the C₁₈ amphiphilic compound **3** onto a glass support using the Langmuir–Blodgett (LB) technique were made using the vertical dipping method [54,55]. After holding the monolayer at a constant surface pressure of 4 mN/m, and allowing the monolayer to equilibrate for 10 min, the glass substrate was submerged. The dipper speed was maintained at 4 mm/min for upward and downward motion and, before each downstroke, the wafer was allowed to dry for 20 min in air. Interestingly, layers are only observed to be transferred to the glass surface upon upstroke, and the initial transfer ratio upon upstroke is high (0.9). In the downstroke, the layer is partially transferred back from the support to the air–water interface, as indicated by the negative transfer ratio and the third and fourth deposition cycles show a noticeable decrease in the transfer ratio, see Figure 3.

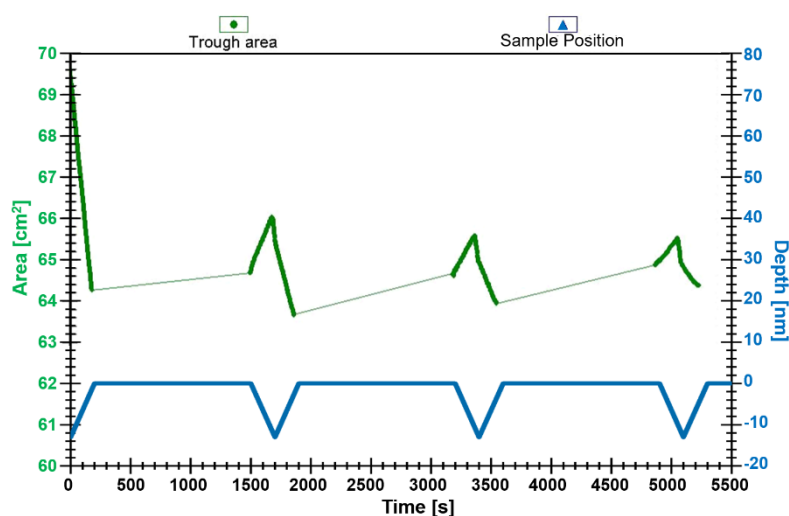


Figure 3. Plot of transfer ratio for attempted transfer and deposition of complex **3** onto a glass support at 4 mN m⁻¹. Emersion and immersion are indicated by a solid line. The transferred layer was allowed to dry for 20 min after each emersion.

In subsequent cycles, the material deposited upon upstroke is completely removed upon downstroke. The profile obtained for the attempts at surface transfer and deposition of C₁₈ amphiphile, compound **3**, suggests that only a Langmuir–Blodgett monolayer can be fabricated, rather than multilayers. The monolayer showed good stability upon air drying, and attempts were made to record the variable temperature Raman spectra of the deposited layer. However, the absence of any signal suggests that the sample size on the glass support is too small.

Initial solubility tests of the *bis*-C₁₂ and *bis*-C₁₈ complexes, **2** and **3**, indicated that both were insoluble in water. Suspensions of excess solid of each in water were then subjected to prolonged sonication to encourage micellar formation with the aim of recovering the solution particles by lyophilization for magnetic characterization. Sonication had no effect on complex **2**, but yielded a purple suspension in the case of **3**, which, after filtering, yielded a transparent purple solution, which was further investigated by light scattering and electron microscopy. The maximum concentration of the assembly of **3** that could be achieved at room temperature was 0.57 mM (this was estimated by evaporating the solvent and weighing the residue). At this concentration, dynamic light scattering (DLS) indicated an average particle size of 350 nm with a polydispersity of 0.22; see Figure 4.

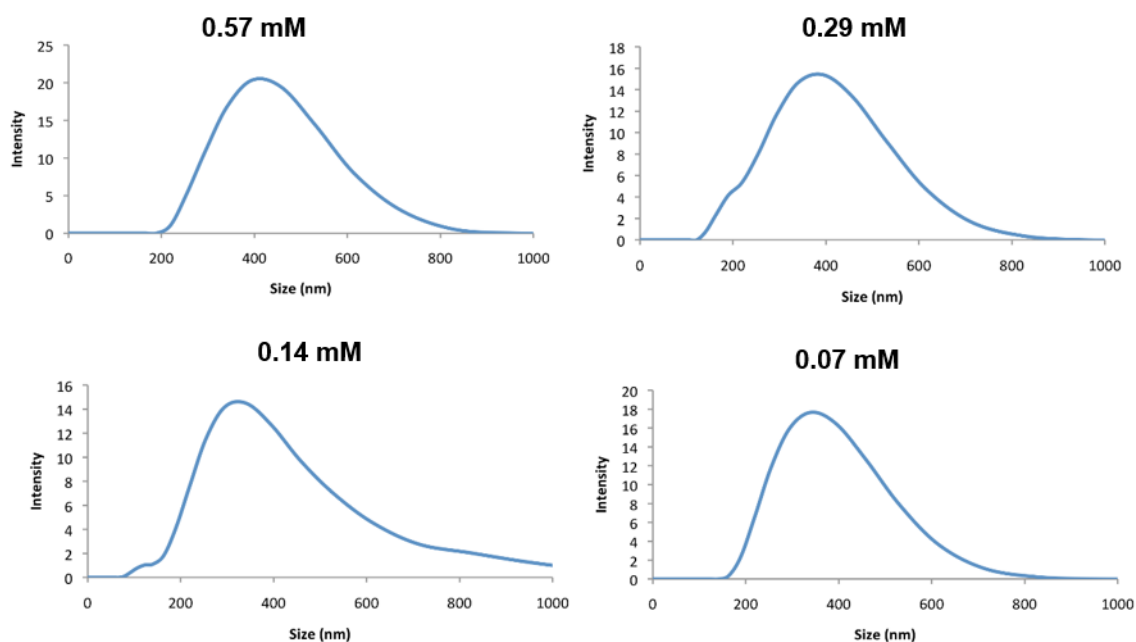


Figure 4. Plot of particle size distribution of complex **3** in water at indicated concentrations measured by DLS.

Stepwise dilution up to one-eighth of this concentration caused a relatively small variation in particle size (Figure 4 and Table 1), suggesting only a small concentration dependence but a relatively high polydispersity.

Table 1. Particle size of **3** in water measured by dynamic light scattering.

Concentration (mol L ⁻¹)	Diameter (nm)	Polydispersity
5.70×10^{-4}	400	0.310
	330	0.235
	320	0.102
2.85×10^{-4}	300	0.364
	270	0.201
	270	0.378
1.43×10^{-4}	320	0.333
	250	0.051
	280	0.318
7.13×10^{-5}	300	0.283
	320	0.165
	290	0.218

Cryo-SEM images, recorded on a further diluted 3.0×10^{-5} M frozen solution assembly of **3**, reveal a range of particles sizes, Figure 5a-c. The majority of particles are of the order of 300 nm (0.3 microns, green line in Figure 5c) but can grow to around 800 nm (0.8 microns, orange line in Figure 5c) which matches the particle sizes obtained by DLS, Figure 4. Larger particles are also discernible, suggesting that the nano-sized particles aggregate into larger particles on the micron scale, Figure 5a. At higher magnification, these are revealed to be comprised of smaller polydispersive oblong-shaped particles ranging from around 0.5–2 microns in width, Figure 5c.

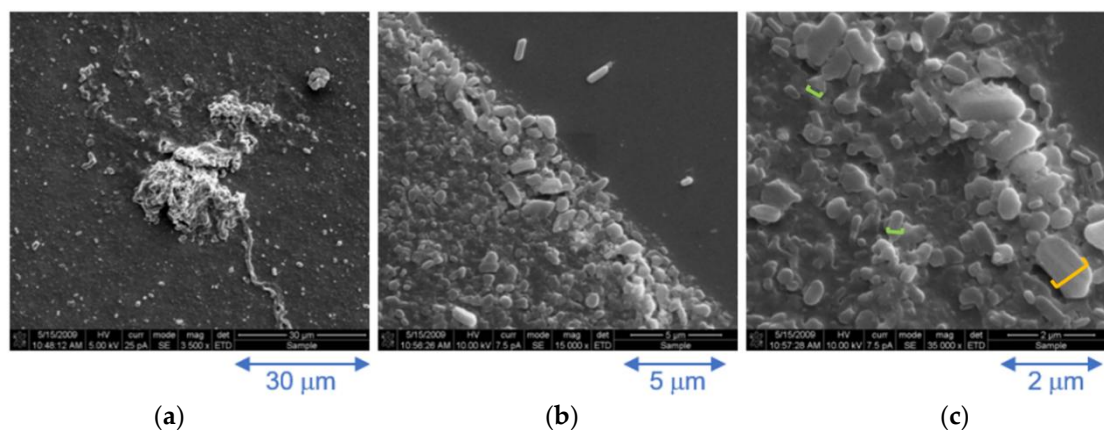


Figure 5. Cryo-SEM images of a 0.03 mM aqueous solution of complex **3** at increasing magnification; (a) $\times 3500$, (b) $\times 15,000$ and (c) $\times 35,000$.

Freshly prepared powdered samples of complexes **1–3** all show a gradual spin crossover, Figure 6 [50,52] and, given the interest in preparing nanoparticles of SCO complexes, it was of interest to investigate if the nanosized micellar particles of **3** could be harvested by freeze-drying, in order to probe their magnetic characteristics.

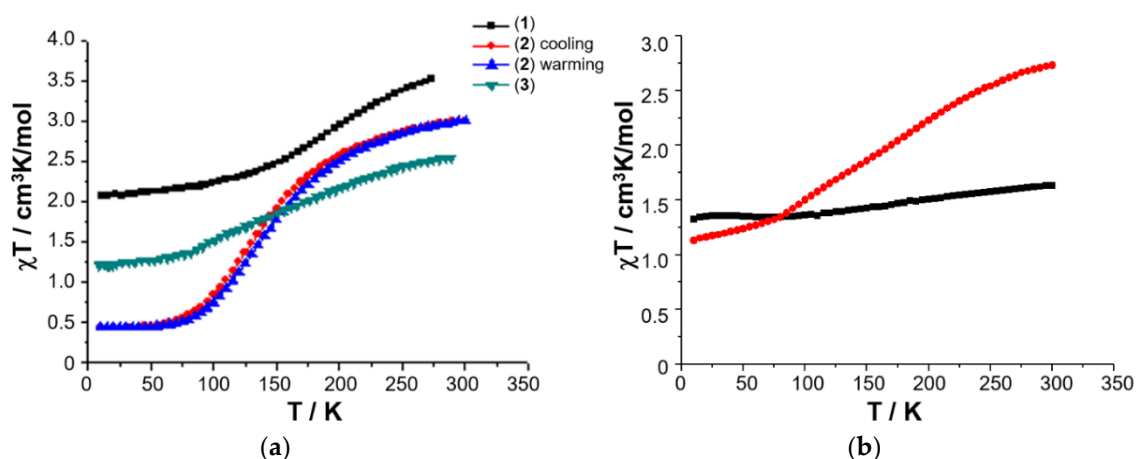


Figure 6. Temperature-dependence of the χT product for **1–3** (a), and variable temperature $\chi_m T$ of solid sample of $[\text{FeL}_{1\text{C}18}]\text{BF}_4, \mathbf{3}$ (red circles), and freeze-dried sample of $[\text{FeL}_{1\text{C}18}]\text{BF}_4, \mathbf{3}$ (black squares), extracted from freeze-dried aqueous assembly of **3** (b). Magnetic data originally published in [50,52].

A 0.03 mM assembly of **3** in water was therefore prepared using sonication and heating, and the resulting solution gravity-filtered to remove all traces of insoluble material as before. The resulting clear solution was then frozen in a Dewar of liquid nitrogen, before being placed under vacuum for 24 h. The recovered dry powder was collected, and its magnetic susceptibility recorded on a SQUID magnetometer. The solid obtained from lyophilization of the aqueous solution of **3** shows a completely different magnetic profile to that of the original solid, Figure 7b, with an almost temperature independent $\chi_m T$ where the high spin (HS) fraction falls from 29% at RT down to 21% at 10 K (assuming $\chi_m T = 0.50 \text{ cm}^3 \text{ mol}^{-1} \text{ K}$ for low spin (LS) Fe(III) with a small orbital contribution and spin only $\chi_m T = 4.38 \text{ cm}^3 \text{ mol}^{-1} \text{ K}$ for HS Fe(III)). Elemental analysis reveals that the composition of the solid sample of lyophilized **3** is different from the freshly prepared anhydrous powder as revealed by the difference between the experimental and calculated values: Exp. (Calc.): C 62.38 (67.26); H 9.40 (9.68); N 3.11 (5.60). The challenges in effectively pumping off all the water mean that the freeze-dried sample may be hydrated and would, therefore, be expected to have a different magnetic profile.

An investigation into differences in the vibrational modes of all three compounds was made by Raman spectroscopy using an excitation wavelength of 785 nm (4 mW) in a temperature range

between 79 K to 295 K. The variable temperature Raman spectra of the C₆ alkylated compound (**1**) show six main modes which change upon cooling from 295 K to 79 K; (Figure 7, indicated by *). The most temperature-responsive Raman modes are identified as 266, 271, 611, 1335, 1599, and 1619 cm⁻¹ at room temperature, and can be associated with the high spin state. These vibrational modes remain practically unchanged down to 223 K. Further cooling to 198 K shows a change in some modes: the mode at 1619 cm⁻¹, attributed primarily to phenyl ring stretch, decreases in intensity, the mode at 1335 cm⁻¹ splits into two distinct vibrational modes, and the mode at 611 cm⁻¹ starts to decrease in relative intensity, giving rise to a new mode at around 620 cm⁻¹. Finally, the broad Raman features at 266 and 271 cm⁻¹ start to resolve, showing that the population of LS species is increasing. The Raman spectrum at 79 K shows significant shifts of between 5 and 20 cm⁻¹, and the key features now appear at 259, 279, 613, 623, 1338, 1348, and 1600 cm⁻¹. It is known from the magnetic profile of compound **1** that, at this temperature, there is a mixture of HS and LS species, and the Raman spectrum at this temperature also confirms this observation, with retention of features from the higher temperature spectrum.

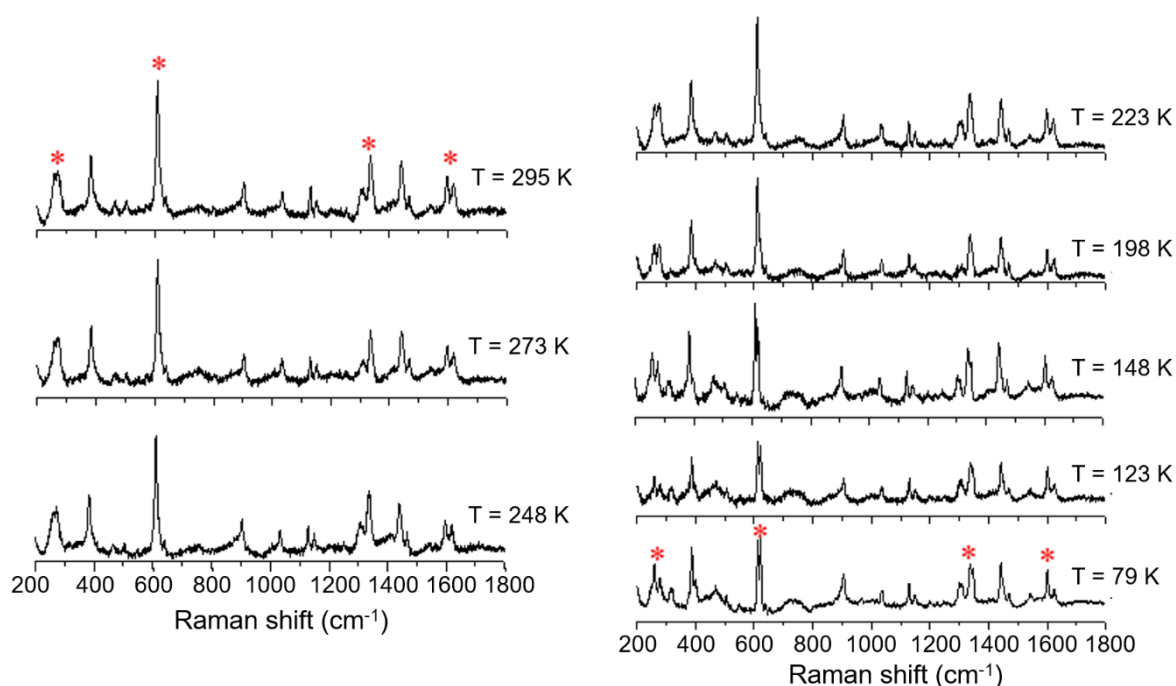


Figure 7. Variable temperature Raman spectra of [FeL₁C₆]BF₄, **1**, using laser excitation at 785 nm in the temperature range of 79 K to 295 K.

The magnetic plot of the C₁₈ amphiphile, compound **3**, shows that it has a very gradual thermal spin transition, which is incomplete at both ends of the temperature range, Figure 6. Raman spectra of this complex also show six main modes which change upon cooling from 295 K to 79 K, which is highlighted in Figure 8. As observed before, at RT, the main modes which undergo a change with temperature are at 268, 272, 612, 625, 1338, 1598, and 1626 cm⁻¹. These vibrational modes change gradually over the measured range, which is in good agreement with the very gradual change the magnetic plot of this compound presents. At 79 K, compound **3** shows vibrational modes at 319, 612, 624, 1346, and 1605 cm⁻¹. The major change observed with temperature was for the vibrational modes at 268 and 272 cm⁻¹, which merge, and a new vibrational mode at 319 cm⁻¹ appears. This spectral region is expected to be dominated by metal–ligand deformation modes that are expected to be profoundly affected by the metal spin state transition. It should be also noted that the vibrational mode, attributed to Fe(III)–O stretch around 623 cm⁻¹, which is characteristic of the LS state, *vide infra*, never disappears, which is in good agreement with the magnetic result for compound **3**.

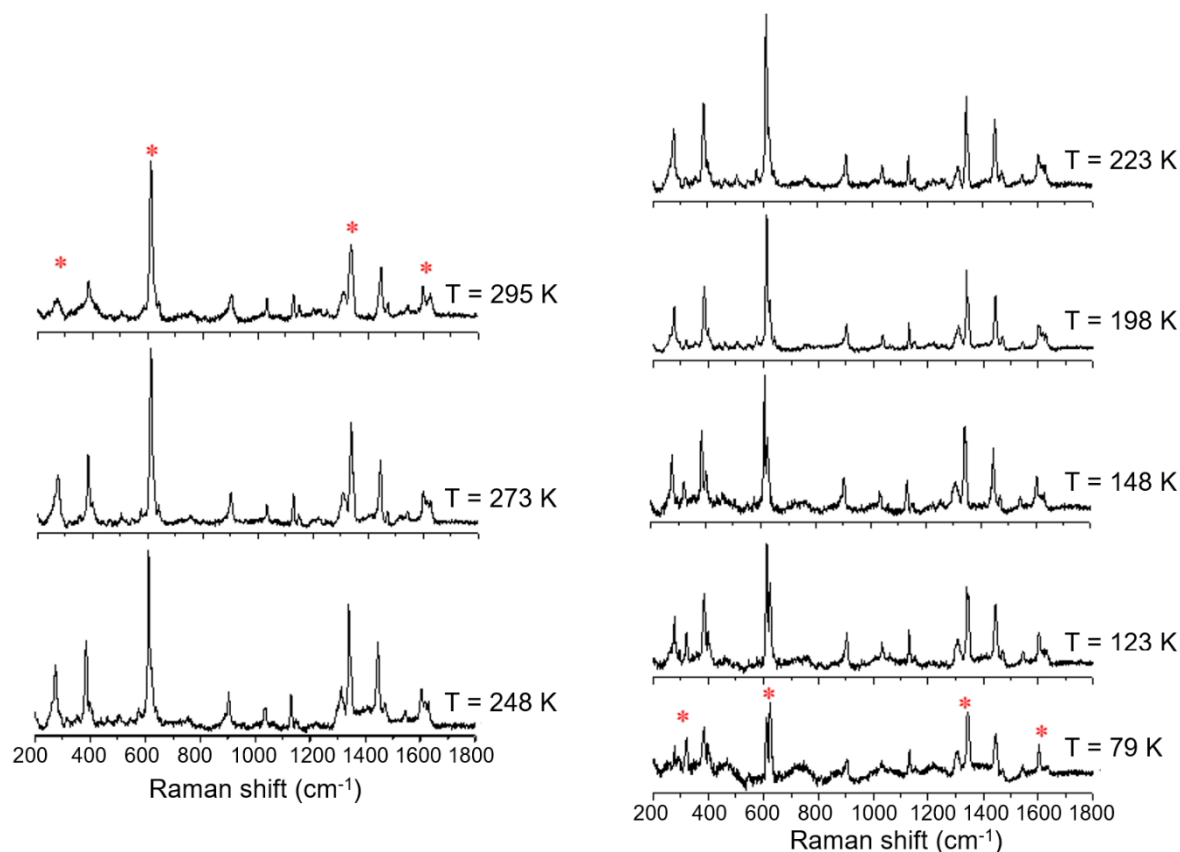


Figure 8. Variable temperature Raman spectra of $[\text{FeL}_{1\text{C}18}]\text{BF}_4$, **3**, using laser excitation at 785 nm in the temperature range of 79 K to 295 K.

Salts of $[\text{Fe}(\text{sal})_2\text{trien}]^+$ are well studied and well characterized for HS, LS, and SCO electronic configurations, and reports of spectroscopic behavior in solution have been published [88,89]. This cation is known to have two distinct absorbance bands: one corresponding to the HS electronic state (≈ 500 nm), and a second one corresponding to the LS electronic state (≈ 650 nm). Resonance Raman (rRaman) for the HS state using laser excitation at 532 nm, and for the LS state at 633 nm, can be used to maximize the enhancement achieved by using laser excitation with energy close to the metal-to-ligand charge transfer (MLCT) band. Besides this enhancement of the main modes, less predominant modes can be unobservable when using rRaman, whereas Raman may highlight these modes and should also be used for spectroscopic characterization. We have used Raman spectroscopy with a laser excitation at 785 nm, and rRaman with a laser excitation at 532 nm, which is in resonance with the high-spin UV-vis maximum absorbance band. Raman and rRaman spectra of compounds **1–3** were recorded using these two different laser lines (LL), one at 532 nm (green laser) and a second at 785 nm (red laser), and both were determined at RT and at 79 K using a 50 \times magnification lens. The Raman spectra of **2** were reported elsewhere before [52], but will be included for completeness. The sample was dispersed on a glass slide, and the variable temperature Raman and rRaman, which are shown in Figures 9–11 for compound **1–3**. Spectra obtained using 532 nm excitation laser produces enhanced rRaman spectra, but do not permit the observation of thermal spin transition. On the other hand, the spectra obtained using the out-of-resonance 785 nm excitation laser lose the enhancement but permit the observation of the thermal spin transition between HS and LS electronic states.

Variable temperature rRaman spectra of **1** using 532 nm excitation show characteristic features for HS Fe(III), which remain unchanged upon cooling, see Figure 9 and Table 2. When the laser line was changed from 532 nm to 785 nm, the appearance of a new mode at 623 cm⁻¹ (which corresponds to $\nu(\text{Fe-O})$) was observed upon cooling to 79 K i.e., indicative of a spin transition. The main $\nu(\text{Fe-O})$ HS mode at 608 cm⁻¹ also becomes less intense due to the increase in LS population on cooling. From comparison of the results obtained for both Raman and rRaman experiments, it was clear that while using rRaman produces an enhancement of the vibrational modes, some very important structural information might be less noticeable, or even undetectable, which does not happen if non-resonance Raman is used to characterize and follow thermal spin changes.

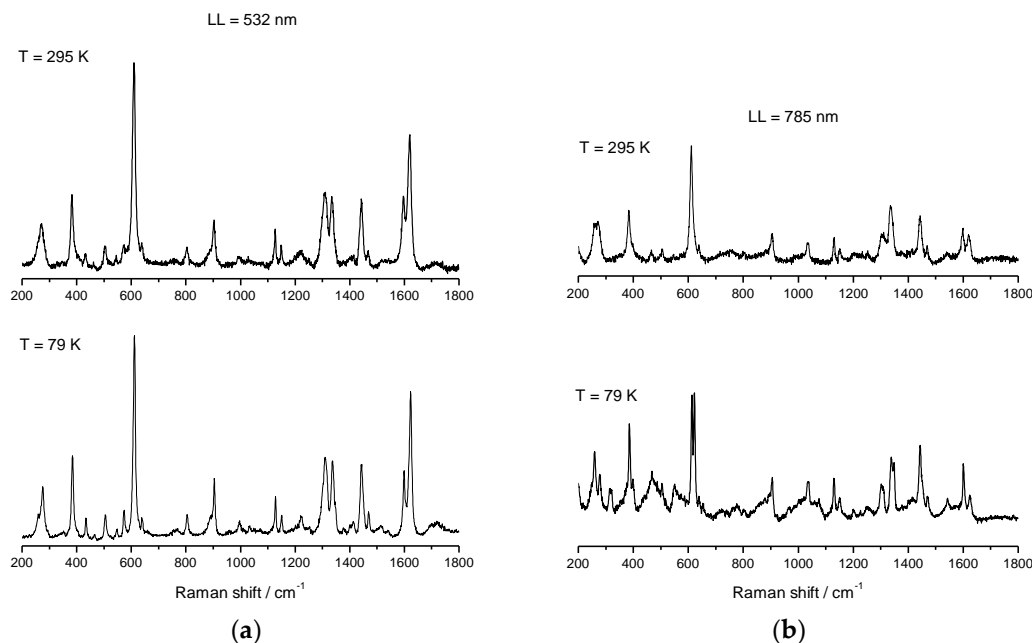


Figure 9. Variable temperature rRaman spectra of $[\text{FeL}_{1\text{C}6}]\text{BF}_4$, **1**, using laser ex 532 nm (a) and variable temperature Raman spectra of $[\text{FeL}_{1\text{C}6}]\text{BF}_4$, **1**, using laser excitation of 785 nm (b).

The $\nu(\text{Fe-O})$ vibrational mode of **2** using 532 nm excitation shows characteristic features of HS Fe(III) at 615 cm⁻¹ and 618 cm⁻¹, see Figure 10, which can be assigned to 68% HS character at RT. On cooling to 79 K, a single feature is observable at 618 cm⁻¹, see Table 2. In the Raman experiment with 785 nm laser excitation, a more distinct separation between modes at 613 and 623 cm⁻¹ at 79 K could be observed, although the complete appearance of the mode at 623 cm⁻¹ at low temperature and the complete disappearance of the mode at 613 cm⁻¹ at low temperature was not observed. The appearance of the new mode at 623 cm⁻¹ at 79 K was attributed to the LS signature which, at 79 K, was estimated at 95% from the magnetic data. The absence of the mode at 1627 cm⁻¹, corresponding to $\nu(\text{C=N})$ at 79 K, also indicates a thermal spin transition.

Variable temperature rRaman spectra of **3** at 295 K with 532 nm excitation show two features at 611 and 614 cm⁻¹, Figure 11, which can be attributed to a mixture of both HS and LS Fe(III) electronic states with 53% HS character at this temperature; see Table 2. On cooling to 79 K, no significant change was observed, and the same vibrational modes were identified, revealing, once again, that rRaman enhances the signals obtained but can, sometimes, omit relevant information. In the rRaman experiment with 785 nm excitation, a more distinct separation between the modes at 613 and 623 cm⁻¹ could be observed, as seen in Figure 9, and the absence of the $\nu(\text{C=N})$ mode at 1627 cm⁻¹, upon cooling to 79 K, indicates a thermal spin transition. It should be noted from the magnetic measurement that, at both ends of temperature range, the complete population of both HS and LS was never attained and, therefore, Raman modes characteristic of both HS and LS should be present at RT and 79 K.

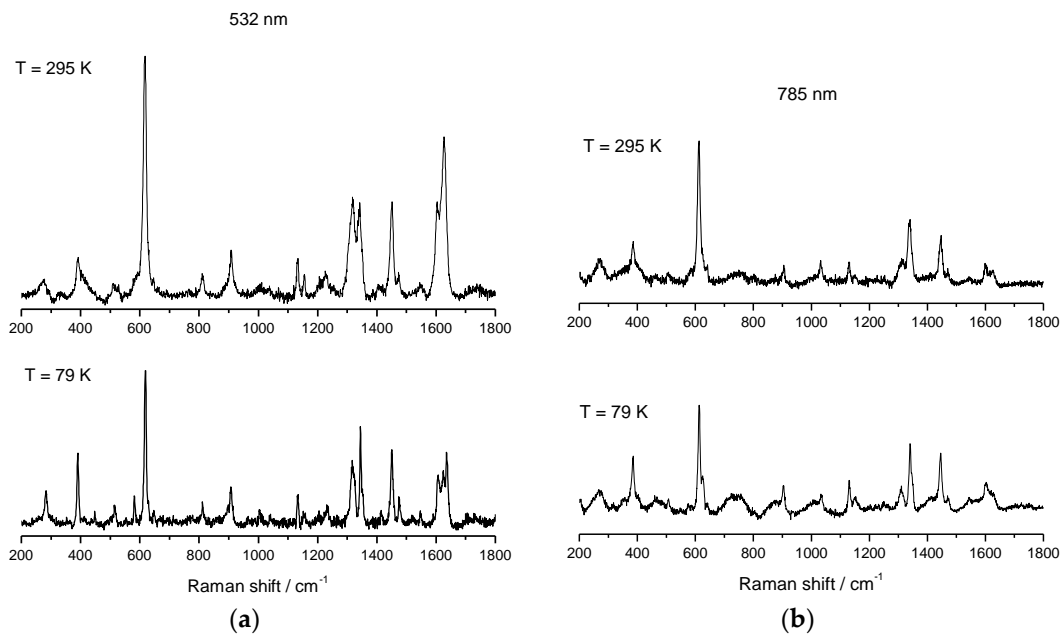


Figure 10. Variable temperature rRaman spectra of $[\text{FeL}_{1\text{C}12}]\text{BF}_4$, **2**, using laser excitation of 532 nm (a) and variable temperature Raman spectra of $[\text{FeL}_{1\text{C}12}]\text{BF}_4$, **2**, using laser excitation of 785 nm (b) [52].

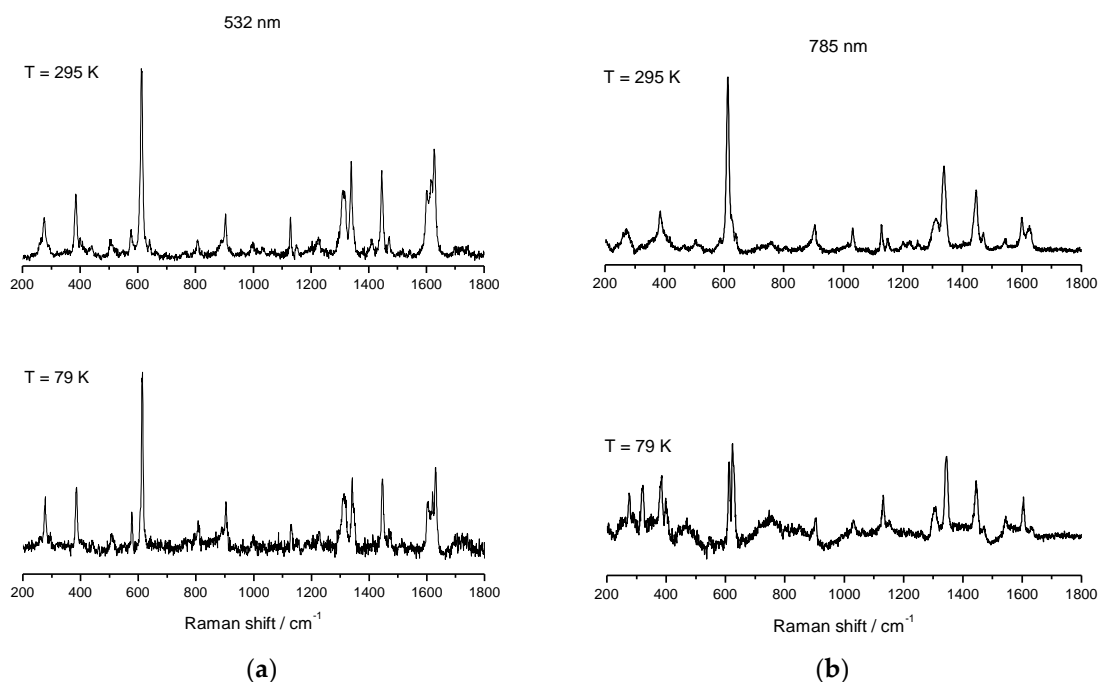


Figure 11. Variable temperature rRaman spectra of $[\text{FeL}_{1\text{C}18}]\text{BF}_4$, **(3)**, using laser excitation of 532 nm (a) and variable temperature Raman spectra of $[\text{FeL}_{1\text{C}18}]\text{BF}_4$, **(3)**, using laser excitation 785 nm (b).

Table 2. Representative Raman shifts and IR stretches of compounds 1–3.

Bond Type	rRaman Shift (cm^{-1})		Raman Shift (cm^{-1})		Literature [88]		IR (cm^{-1})
	LL = 532 nm		LL = 785 nm		HS	LS	
complex 1	T = 295 K	T = 79 K	T = 295 K	T = 79 K	HS	LS	-
$\nu(\text{Fe-O})$	609	611	611	613,623	612	622	602
$\nu(\text{C-O})$	1311	1310	1309	1308	1310	1310	1302
$\nu(\text{C=C})$	1596	1598	1599	1600	1600	1600	
$\nu(\text{C=N})$	1620	1622	1619	1625	1630	absent	1620

complex 2	T = 295 K	T = 79 K	T = 295 K	T = 79 K	HS	LS	-
$\nu(\text{Fe-O})$	615,618	618	613	613,623	612	622	602
$\nu(\text{C-O})$	1317	1317	1317	1310	1310	1310	1302
$\nu(\text{C=C})$	1603	1607	1602	1605	1600	1600	-
$\nu(\text{C=N})$	1627	1635	1627	absent	1630	absent	1620
complex 3	T = 295 K	T = 79 K	T = 295 K	T = 79 K	HS	LS	-
$\nu(\text{Fe-O})$	611,614	612,614	612,625	613,623	612	622	602
$\nu(\text{C-O})$	1308	1307	1309	1310	1310	1310	1302
$\nu(\text{C=C})$	1599	1601	1600	1605	1600	1600	-
$\nu(\text{C=N})$	1627	1630	1627	absent	1630	absent	1620

3. Discussion

We have investigated the behavior of three known SCO iron(III) complexes which host hydrophobic alkyl chains varying from C_6 to C_{18} . The ability of these compounds to form monolayers at the air–water interface was investigated, and we have found that the C_6 compound **1** does not form monolayers at the air–water interface, due to partial solubility of the compound in water. On the other hand, the C_{12} and C_{18} amphiphilic compounds **2** and **3** do form stable monolayers at the air–water interface over a long period of time (>90 min). Besides forming stable monolayers at the air–water interface, the Langmuir isotherms of both compounds do not show a steep increase in surface pressure with compression, indicating that the arrangements at the air–water interface are not very ordered, possibly due to backfolding of the alkyl chains. Attempts to transfer the monolayer formed at the air–water interface by the C_{18} amphiphilic compound **3**, onto a glass support, resulted in the transfer of only one layer. This indicates that it is only possible to fabricate a Langmuir–Blodgett monolayer of compound **3**. Characterization of the transferred LB monolayer by Raman spectroscopy was not possible, due to insufficient material being deposited on the glass surface. We have also observed that the longer amphiphilic compound **3** forms micellar aggregates in aqueous solution, and that lyophilization of micellar aggregates yields a hydrated product with a magnetic profile that is almost temperature independent, with a predominance of the LS electronic state.

4. Materials and Methods

Pressure–area isotherms and time stability were measured at 25 °C on a KSV MiniMicro Langmuir–Blodgett trough (Biolin Scientific, Espoo, Finland), with a surface area between 1700 and 8700 mm². Water was purified with a Barnstead Nanopure system (Thermo Fisher Scientific, UK), and its resistivity was measured to be higher than 18 M Ω cm. Chloroform (puriss. p.a. $\geq 99.8\%$, Fluka, was used as spreading solvent. Typically drops of the surfactant solution (20 μL , 0.50 mM) were deposited using a microsyringe on the water subphase. After letting the solvent evaporate for 30 min, the barriers were compressed at 6 mm min⁻¹ (3 cm² min⁻¹), and the surface pressure was monitored using a platinum Wilhelmy plate. Variable temperature magnetic susceptibility measurements were performed on a Magnetic Property Measurement System (MPMS XL-7) (Quantum Design, CA, USA) sample with Superconducting Quantum Interference Device (SQUID). Dynamic light scattering was measured using a Malvern PCS-4700 instrument (Malvern Panalytical, Malvern, UK) equipped with a 256-channel correlator. The 488.0 nm line of a Coherent Innova-70 Ar ion laser was used as the incident beam with 250 mW laser power, and the employed scattering angles, θ , were 90°. The temperature was maintained at 25 °C with an external circulator. Data analysis was performed according to standard procedures and interpreted through a cumulant expansion of the field autocorrelation function to the second order. Transmission electron microscopy (TEM) was carried out using a JEOL 2000FX 200 kV instrument (JEOL, Tokyo, Japan) instrument equipped with a LaB₆ filament, angle PGT EDXS detector and Gatan slow-scan and wide-angle TV-rate cameras. This instrument has a high-tilt pole piece with a point resolution of 0.3 nm and $\pm 60^\circ$ double-tilt capability using the Gatan low-background double-tilt holder. Cryo-SEM measurements were completed using an FEI Dual-beam Cryo-focused Ion Beam Scanning Electron Microscope. The Raman spectra were acquired on a confocal Labram HR microscope from Horiba coupled to a 532 nm diode (5 mW) or

785 nm solid state laser source (4 mW) using a 100× (N.A. = 0.90) Olympus objective. The Raman band of a silicon wafer at 520 cm⁻¹ and Rayleigh line were used to calibrate the spectrometer, and the accuracy of the spectral measurement was estimated to be better than 1 cm⁻¹. Measurements at low temperature were carried out in a temperature-controlled stage (THMS 600 from Linkam Scientific Instruments (Linkam Scientific Instruments, Tadworth, UK)). The Raman spectra were recorded over the temperature range 295–77 K. The acquisition time for all spectra was set to 6 s, unless otherwise stated. Spectra were background corrected using a 4-point polynomial generated using LabSpec software (Horiba Scientific, Kyoto, Japan).

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