

Engineering a Highly Defective Stable UiO-66 with Tunable Lewis-Bronsted Acidity: The Role of the Hemilabile Linker

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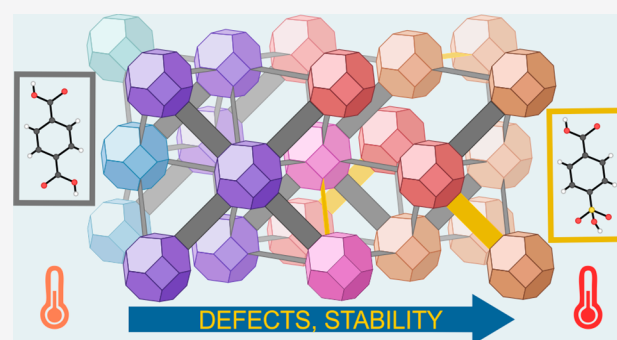
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ABSTRACT: The stability of metal–organic frameworks (MOFs) typically decreases with an increasing number of defects, limiting the number of defects that can be created and limiting catalytic and other applications. Herein, we use a hemilabile (HL) linker to create up to a maximum of six defects per cluster in UiO-66. We synthesized hemilabile UiO-66 (HL-UiO-66) using benzene dicarboxylate (BDC) as linker and 4-sulfonatobenzoate (PSBA) as the hemilabile linker. The PSBA acts not only as a modulator to create defects but also as a coligand that enhances the stability of the resulting defective framework. Furthermore, upon a postsynthetic treatment in H₂SO₄, the average number of defects increases to the optimum of six missing BDC linkers per cluster (three per formula unit), leaving the Zr-nodes on average sixfold coordinated. Remarkably, the thermal stability of the materials further increases upon this treatment. Periodic density functional theory calculations confirm that the hemilabile ligands strengthen this highly defective structure by several stabilizing interactions. Finally, the catalytic activity of the obtained materials is evaluated in the acid-catalyzed isomerization of α -pinene oxide. This reaction is particularly sensitive to the Bronsted or Lewis acid sites in the catalyst. In comparison to the pristine UiO-66, which mainly possesses Bronsted acid sites, the HL-UiO-66 and the postsynthetically treated HL-UiO-66 structures exhibited a higher Lewis acidity and an enhanced activity and selectivity. This is further explored by CD₃CN spectroscopic sorption experiments. We have shown that by tuning the number of defects in UiO-66 using PSBA as the hemilabile linker, one can achieve highly defective and stable MOFs and easily control the Bronsted to Lewis acid ratio in the materials and thus their catalytic activity and selectivity.



INTRODUCTION

UiO-66 is one of the best known metal–organic frameworks (MOFs) due to its outstanding stability, which is caused by the presence of high-valent metal cations creating clusters with high charge densities.¹ Moreover, UiO-66 is a prototypical MOF in defect engineering because it can bear a high number of defects without loss of structure.^{2–6} In general, the most common approach to create defects is the modulation synthesis approach, in which a large excess of the monocarboxylic acid is used, in combination with the dicarboxylic acid. The monotopic ligands reduce the speed of crystallization and occupy the metal coordination sites to generate defects. The most common modulators used are formic acid, acetic acid, and trifluoroacetic acid, all having a mono coordination mode.^{3,7–9} MOFs with hierarchical porosity have been synthesized using either monocarboxylic acids as modulators or by thermolysis of mixed linker MOFs.^{10–13} In catalysis, the defective UiO-66 (using trifluoroacetic acid as modulator) has already been shown to have a high reactivity in the Meerwein reduction of 4-*tert*-

butylcyclohexanone with isopropanol.¹⁴ Recently, we used L-proline as a chiral modulator for the synthesis of UiO-type MOF structures which showed an excellent reactivity in the diastereoselective aldol reaction.¹⁵ Nevertheless, in all cases, the thermal stability of the framework is reduced upon the generation of defects.^{3,16,17} Atzori et al. demonstrated that the maximum number of defects is 4.4 per cluster using benzoic acid in UiO-66.¹⁸ Bueken et al. reported a similar value of 4.3 missing linkers per cluster.¹⁹

Apart from monocarboxylic acid modulators, hemilabile ligands can be used in the synthesis of carboxylic MOFs (Figure 1). They have, next to the carboxylate group, a second coordinating functional group (e.g., sulfonate groups). The

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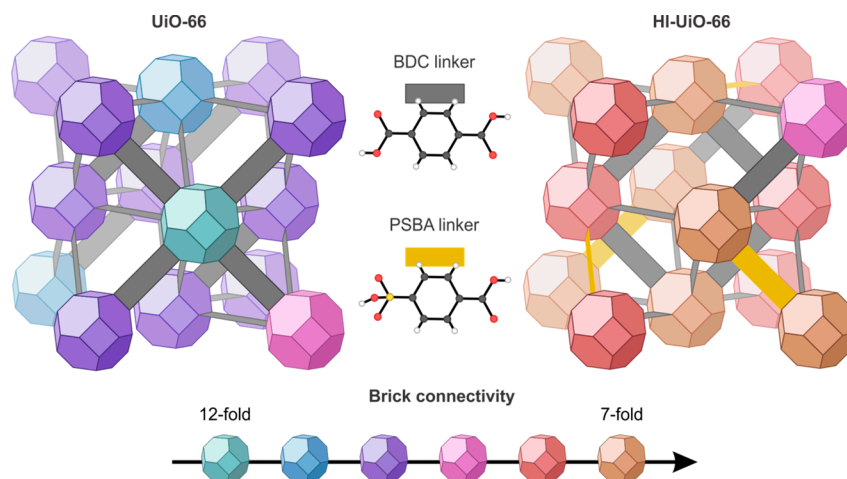


Figure 1. Schematic representation of the UiO-66 and HI-UiO-66 materials with possible configurations of the bricks that give rise to coordinatively unsaturated Zr sites.

concept of the use of hemilabile ligands was introduced by Morris et al.²⁰ For example, the hemilabile MOF, denoted as Cu-SIP-3, prepared by Xiao et al. showed an ultrasensitive low pressure nitric oxide adsorption upon the reversible phase transformation induced by the change in the coordination of sulfonate group when the water is removed.²¹ McHugh et al. also demonstrated that a copper paddlewheel MOF exhibited good hydrolytic stability due to the presence of a hemilabile linker.²² Until now, hemilabile linkers have only been used to explore the selective gas adsorption, separation, dielectric, and other physical properties of MOFs.^{20,23,24}

Herein, we report the synthesis and characterization of hemilabile UiO-66 (HI-UiO-66) using benzene-1,4-dicarboxylic acid (BDC) and 4-sulfobenzoic acid potassium salt (PSBA) in a 1:1 ratio to obtain a highly defective UiO-66 (Figure 1). Because of the inherent weak coordination mode of the sulfonate group, a simple postsynthetic treatment in H_2SO_4 was carried out, which resulted in a total of six BDC missing linkers per cluster (HI-UiO-66- SO_4). This is the theoretical limit to still maintain a three-dimensional UiO-66 network.^{3,19} Importantly, the thermal stabilities of the three MOFs follow the order UiO-66 < HI-UiO-66 < HI-UiO-66- SO_4 , which confirms that the material with the highest number of defects is also the most stable one. The catalytic properties of the obtained materials were evaluated in the isomerization of α -pinene oxide, as this reaction allows for discriminating the role of Lewis and Brønsted acid sites. In comparison to the pristine UiO-66 as Brønsted acid catalyst, which showed a moderate activity of 40% conversion and low selectivity (37%) at 0.5 h, full conversion and an enhanced selectivity (71%) is observed for the HI-UiO-66- SO_4 materials as Lewis acid catalyst.

EXPERIMENTAL SECTION

Synthesis of the Hemilabile UiO-66. The synthesis procedure is based on a modified procedure of Biswas et al.²⁵ $\text{ZrO}_2\text{Cl}_2 \cdot 8\text{H}_2\text{O}$ (1 g, 3.1 mmol), terephthalic acid (515 mg, 3.1 mmol) and 4-sulfobenzoic acid potassium salt (745 mg, 3.1 mmol) were dissolved in a mixture of formic acid (12 mL, 310 mmol) and dimethylacetamide (30 mL) in an ultrasonic bath for 30 min. The resulting mixture was placed in a Teflon-lined autoclave at 150 °C for 24 h and subsequently cooled to room temperature. The solid was collected by filtration and washed with dimethylformamide (DMF) and methanol and subsequently stirred in DMF at 70 °C (24 h) and methanol under ambient

conditions (24 h) to remove unreacted linker, modulator, and DMF. The final purified product was dried under vacuum at 65 °C.

Postsynthetic Modification (HI-UiO-66- SO_4). HI-UiO-66- SO_4 : 150 mg catalyst was treated in 20 mL of H_2SO_4 (0.05 M). The suspension of the solid was stirred for 24 h at room temperature. The resulting solid was filtered off and thoroughly washed multiple times with water until neutral pH was obtained and consequently stirred with methanol for 24 h. The precipitate was filtered and dried under vacuum at 65 °C.

Catalytic Procedure. In a typical catalytic test, 50 mg of catalyst (preactivated for 16 h at 150 °C under vacuum) was suspended in 5 mL of toluene (anhydrous) in a 25 mL round-bottom flask. Hereafter, 130 mg of α -pinene oxide (0.86 mmol) and 146 mg of dodecane (0.86 mmol) as an internal standard were added. This mixture was magnetically stirred at 70 °C. Aliquots were taken from the reaction mixture by syringe at indicated time intervals, filtered with a nylon filter to remove catalyst particles, and diluted with toluene. The products were separated by GC and identified by GC-MS. During a recycling experiment, the catalyst was centrifuged by filtration, washed with toluene and methanol and dried under vacuum at 150 °C and was then reused. For the filtration tests, the catalyst was filtered off, and the supernatant was reacted further in another vial.

Computational Methodology. All calculations on the two-brick unit cell of studied UiO-66-type materials were performed using the periodic Vienna Ab Initio Simulation Package (VASP code) with the density functional theory (DFT) method to represent the crystal environment.^{26–30} We applied the projector augmented wave (PAW) approach,³¹ and the Brillouin zone was sampled by the Γ -point at PBE level of theory³² including Grimme-D3 dispersion corrections.^{33,34} The kinetic energy cutoff for the plane waves was 700 eV; the convergence threshold for the electronic self-consistent field (SCF) calculations was set to 10^{-8} eV, and a Gaussian smearing of 0.025 eV was included to improve convergence.

Each equilibrium structure was obtained by performing a fixed-volume geometry optimization at the optimal unit cell volume, which was determined by fitting a Rose–Vinet equation of state³⁵ to the electronic energies of seven optimized structures at different unit cell volumes. All geometries were optimized via the conjugate gradient method until the energy difference between subsequent steps became smaller than 10^{-7} eV. Hessians were calculated for the optimized structures via a finite difference approach. The corresponding normal modes and vibrational frequencies were obtained after a normal-mode analysis via TAMkin.³⁶

RESULTS AND DISCUSSION

PSBA as a Hemilabile Linker to Create Defects. Synthesis of UiO-66 with hemilabile linkers was carried out

by varying the ratio of PSBA to BDC (0.2:1, 1:1, and 10:1) and the number of defects was characterized by proton nuclear magnetic resonance (^1H NMR) of digested MOFs and thermogravimetric (TGA) analysis (Table S1). The material obtained using the 0.2:1 ratio has neither PSBA in the structure nor any extra defect corresponding to pristine UiO-66 (Figures S1, S3). Using the 10:1 ratio, the obtained material contained a high relative amount of PSBA versus BDC (0.4:1, Figure S2). However, after washing with sulfuric acid, a lower number of defects was obtained (5 defects per cluster, Figure S4). This is probably due to excess of PSBA in the structure, which was removed during the postsynthetic treatment. We noticed a 40% mass loss after the acid treatment. We found the material obtained by using a 1:1 ratio was the best to retain the number of defects even after acid treatment, which is discussed in detail.

NMR studies were performed to verify the incorporation of PSBA and to determine the experimental molar ratio of PSBA and BDC in the obtained frameworks after digestion of the samples (Table 1). From these NMR results it was observed

Table 1. Composition and Properties of the UiO-66, HI-UiO-66, and HI-UiO-66-SO₄ Materials

sample	initial PSBA:BDC ^a in reaction mixture	observed PSBA:BDC ^b in the MOF	number of defects per cluster ^c
UiO-66			3.4
HI-UiO-66	1:1	0.20:1	4.4
HI-UiO-66-SO ₄		0.11:1	6.0

^aTheoretical ratio of PSBA and BDC. ^bExperimentally obtained ratio determined by means of NMR. ^cNumber of defects based on TGA and NMR results as described by Shearer et al.³

that the molar ratio PSBA/BDC in the structure was maintained even after heating the sample up to 400 °C for 24h (Figure S5). This is in contrast to the results observed for monocarboxylic acid based modulators for which a lower decoordination temperature (325 °C for trifluoroacetate) was obtained.¹⁴ This suggests that the hemilabile linkers coordinate with both their carboxylate side and their sulfonate side, stabilizing the hemilabile linkers in the structure.

The infrared (IR) spectra show the characteristic bands at 1117 ($\nu_{\text{as}}\text{SO}_3$) and 1035 ($\nu_{\text{s}}\text{SO}_3$) and 1008 ($\nu_{\text{s}}\text{SO}_3$) cm^{-1} that can be assigned to the asymmetric and symmetric stretching frequencies of sulfonate, respectively, with a slight shift from 1102 ($\nu_{\text{as}}\text{SO}_3$), 1027 ($\nu_{\text{as}}\text{SO}_3$), and 1002 ($\nu_{\text{as}}\text{SO}_3$) cm^{-1} of PSBA due to the coordination with the Zr-nodes (Figure S6).^{37–39} The same applies for the peak at 740 cm^{-1} , shifted from 760 cm^{-1} , attributed to the bent C–S stretching vibration. These results confirm the coordination of both the carboxylic acid groups of BDC and the sulfonate group of PSBA to Zr(IV).

To gain further insight on the binding of the sulfonate groups, X-ray photoelectron spectroscopy (XPS) measurements were carried out and analyzed by fitting the respective elements with their binding energy (Figure 2). For PSBA (K-salt), the deconvoluted S 2p peaks at 167.9 and 169.1 eV were attributed to phen-SO₃[−], coordinated to potassium, while the peaks at 168.7 and 169.9 eV were assigned to hydrolyzed species phen-SO₃H.^{40,41} The HI-UiO-66 shows mainly the bands at 168.1 and 169.3 eV (green asterisk), matching with coordinated phen-SO₃[−] in PSBA linker and corroborating the

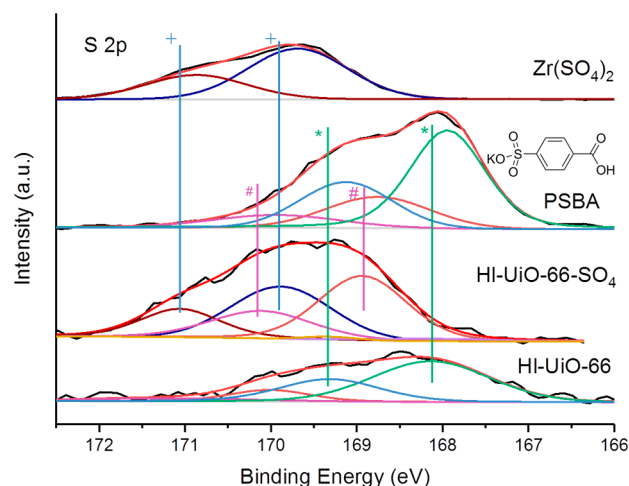


Figure 2. Deconvoluted XPS spectra of HI-UiO-66, HI-UiO-66-SO₄, PSBA, and Zr(SO₄)₂, highlighting S 2p element.

coordination of PSBA to zirconium sites and some traces of (free dangling) protonated PSBA linkers.

The crystallinity of the obtained solids was verified by means of powder X-ray diffraction (PXRD) (Figure 3a). An obvious broad diffraction in the 2θ range between 3 and 7° was observed in HI-UiO-66, originating from the cluster defects.^{3,42} To calculate the number of defects, we performed TGA

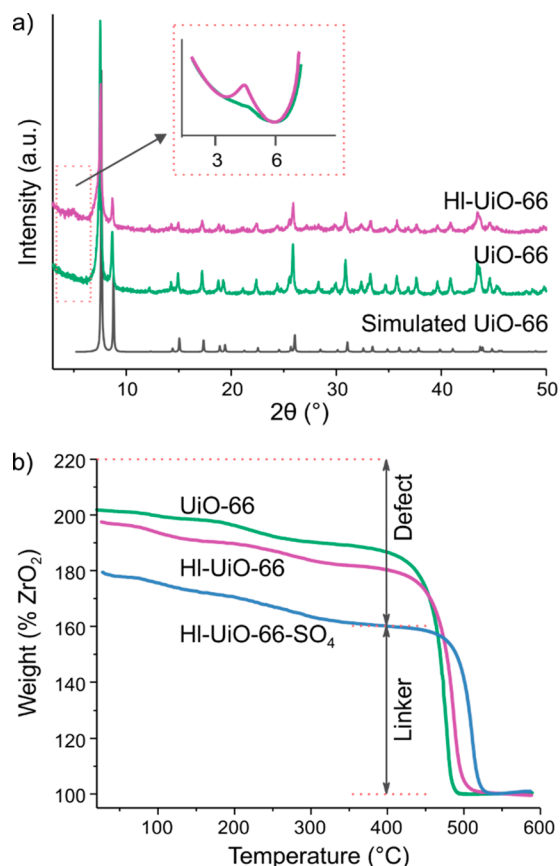


Figure 3. (a) PXRD pattern of the pristine UiO-66 and the HI-UiO-66 samples. The inset shows the overlap of PXRD patterns from 3 to 7° and (b) TGA results obtained of UiO-66, HI-UiO-66, and HI-UiO-66 washed with H₂SO₄ solution.

analysis, which is one of the most general and efficient methods to study defects in MOFs as the number of defects can be calculated based on the weight loss (Figure 3b).^{2,3,13,38,41} As can be seen from Table 1, on average, the HI-UiO-66 contains 4.4 defects per cluster. This number is significantly higher than the number of defects present in the pristine UiO-66 using only BDC as linker. This indicates that the PSBA does not only act as a bidentate linker, but also as a modular to create extra defects.

We performed Ar-adsorption for all samples. The logarithmic isotherms are shown in Figure 4. Careful

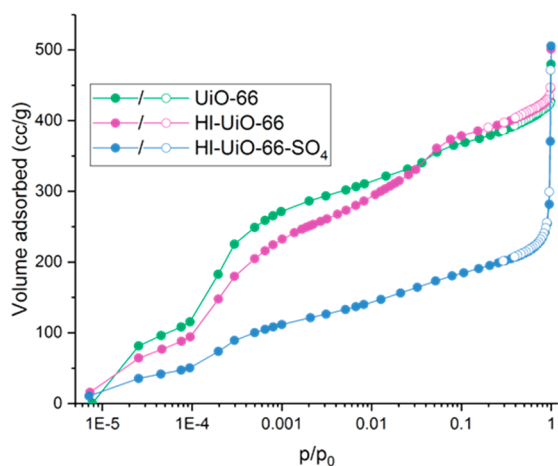


Figure 4. Logarithmic Ar-isotherms for the UiO-66, HI-UiO-66, and HI-UiO-66-SO₄.

inspection of these pristine isotherms shows that the UiO-66 isotherm has two major condensation steps, one at very low pressures (around $1.10^{-5} p/p_0$) and one at a slightly higher pressure (around $1.10^{-4} p/p_0$). There is a small increase in uptake as well around $5.10^{-2} p/p_0$. These three pore size regions are assigned to respectively the two cages in UiO-66 (0.8 and 1.1 nm) and larger micropores (around 1.5–2.0 nm), due to defects. The HI-UiO-66 shows a decrease in Ar-uptake in the two smaller pore size areas but a very noticeable and significant increase in Ar-uptake around 1.5–2.0 nm. This strongly suggests that the number of larger micropores has increased in the HI-UiO-66. The pore size distribution, based upon a silica/zeolite kernel, is shown in the Supporting Information (Figures S14–S16).

The energy dispersive X-ray spectroscopy (EDS) mapping of HI-UiO-66 confirmed that the S species are uniformly distributed inside MOFs (Figure S9), confirming that the PSBA linkers are homogeneously distributed in the materials.

Postsynthetic Modification with Sulfuric Acid and a Remarkable Increase in Stability. Sulfonates have a coordination strength to the Zr₆ clusters that is weaker than that of carboxylates.²³ Therefore, it is comparatively easier to replace the weaker coordinating ligands by a stronger one. In this regard, HI-UiO-66 was immersed in an acid solution to partially remove the PSBA linker. TGA measurements show that after immersing the samples in an H₂SO₄ solution, the average number of defects increased from 4.4 to 6.0 defects per cluster (Figure 3b). Six defects per cluster is considered the maximum of defects to still maintain a stable 3D UiO-66 framework.^{3,19} As shown in Table 1 and in Figure S11, this observation is accompanied with a reduction in the amount of PSBA and BDC. The relative amount of PSBA that was

removed during the acid treatment was higher than the relative amount of BDC linker that was removed upon the treatment (50 and 17%, respectively). Notably, although the number of defects increased after the postsynthetic treatment, the crystalline structure is still preserved (Figure S12). Nevertheless, the Langmuir surface area of HI-UiO-66 decreased after the treatment in the H₂SO₄ solution (Figures 4, S13). HI-UiO-66-SO₄ has six missing linkers per cluster, which corresponds to a sixfold structure with more defects than the eightfold *re*o UiO-66. The *re*o UiO-66 is a structure that possesses ordered missing clusters.⁴³

When looking again at the Ar-isotherms (Figure 4), we see that for the HI-UiO-66-SO₄, a large interparticle porosity has emerged, indicating that at least some of the particles have reduced into fines. This is corroborated by scanning electron microscopy measurements and a statistical analysis of the size of the particles (Figure S7). From the pore size distributions, it is clear that all micropores reduce significantly; also, the surface area and the pore volume reduce strongly to 928 m²g⁻¹ and 213 cm³g⁻¹.

The deconvoluted S 2p XPS spectra were analyzed to confirm the presence of sulfonate and sulfate species in HI-UiO-66-SO₄ (Figure 2). Two peaks at 168.9 and 170.2 eV (purple hash) in HI-UiO-66-SO₄ are attributed to the “freely dangling” phen-SO₃H of the PSBA linker, whereas the peaks at 169.9 eV and 171.0 (blue plus) in the UiO-66-SO₄ are indicative for these sulfonate groups, coordinated to Zr-sites. This indicates that the PSBA decoordinates after H₂SO₄ treatment. Analysis of the deconvoluted bands gives a rough indication of the distribution of the surface species. For the HI-UiO-66-SO₄, approximately 40% of the sulfur is present as sulfate (blue plus), and 60% is present as hemilabile linker (purple hash). Further corroboration is found in the oxygen region of the XPS spectra (Figure S19). The O 1s region has a new peak at 534.4 eV in HI-UiO-66 compared with UiO-66. This peak is attributed to Zr–O(SO₃). After washing with H₂SO₄, this peak mainly disappears, and another peak is found at 532.7 eV, attributed to O 1s of SO₄²⁻.

We made a surprising observation when we performed thermal stability tests of the three MOF materials. TGA experiments show that the thermal stability further increases from UiO-66, HI-UiO-66, to HI-UiO-66-SO₄ (Figure 3b). To confirm the result, the variable-temperature powder X-ray diffraction (VTXRD) was performed. As can be seen from Figure S21, these VTXRD results correspond very well to the results obtained from the TGA measurements. A higher thermal stability was observed with increasing numbers of defects which resulted in the following order: UiO-66 (450 °C) < HI-UiO-66 (480 °C) < HI-UiO-66-SO₄ (515 °C). Compared with HI-UiO-66 and UiO-66, the sulfonate group of hemilabile ligands makes the HI-UiO-66-SO₄ more stable in spite of the high amount of defects. The results correspond to the work of Muesmann et al. concluding that Cu(BDS) (BDS = *p*-benzenedisulfonate) showed stability higher than that of Cu(BDC).⁴⁴ The enhanced stability of HI-UiO-66-SO₄ suggests that the sulfate groups play an important role in the stabilization of this defective structure. This phenomenon was also observed in pristine UiO-66 (Figure S17).

To obtain a better insight into the structural stability of the modified material, periodic DFT calculations were carried out on the UiO-66 and HI-UiO-66 materials. For UiO-66, the calculations were performed on the two bricks structure with isolated missing linker defects denoted as type 6 in the work of

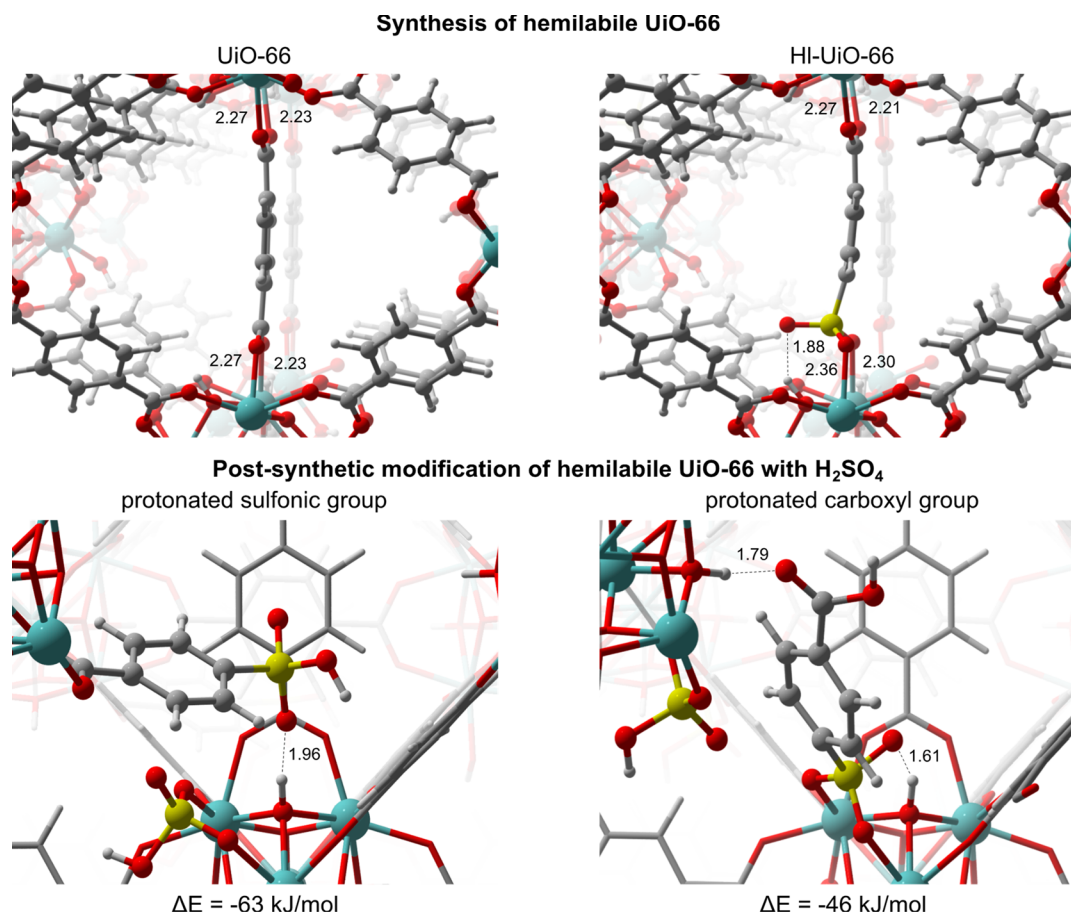


Figure 5. Top: Representation of UiO-66 with one missing linker defect and HI-UiO-66 with one missing linker defect and one PSBA linker. Bottom: Creation of defects in HI-UiO-66 upon postsynthetic treatment with H₂SO₄. Representation of geometrically optimized HI-UiO-66-SO₄ structures in which the PSBA linker is in the dangling state due to protonation of (a) sulfonic and (b) carboxyl group while the HSO₄⁻ is adsorbed on the defective site.

Rogge et al.⁴⁵ Each defect site was capped with one chemisorbed water molecule as was indicated by IR studies (Figure S20). The HI-UiO-66 was constructed accordingly by substituting one out of ten BDC linkers in the defective cluster with a PSBA linker (Figure 5, top). The stability of the materials was calculated based on the difference in electronic energy, following the equation: $\Delta E = (E_{\text{HI-UiO-66}} + E_{\text{BDC}}) - (E_{\text{UiO-66}} + E_{\text{PSBA}})$. The difference in electronic energy between the HI-UiO-66 with one PSBA linker and the pristine UiO-66 indicates that the modified material is around 27 kJ/mol more stable than UiO-66. Even though the bond between the sulfonate group and Zr atoms is expected to be weaker, the sulfonate group is stabilized by additional hydrogen bonding interaction with the μ_3 -OH group and one of the three oxygen atoms from the linker (Figure 5, top). The H-O distance has “moderate strength”, according to the classification of the hydrogen bond of Steiner et al., as it is situated in the range of 1.5–2.2 Å.^{46,47} This additional interaction slightly alters the position of the modified linker. The PSBA linker is bent and therefore the interaction of sulfonate group with Zr is weaker which is also seen by the longer Zr-O bond distance compared to the BDC linker. In the case of a straight configuration, the oxygen atom would be too close to the hydrogen atom causing repulsion. It is due to the bending of the linker that an optimal hydrogen bond distance can be obtained.

To further understand the hemilabile nature of the HI-UiO-66 material, the stability of the PSBA linker protonated by H₂SO₄ was investigated. In the acidic solution the PSBA linker can be attacked and protonated either on the sulfonate or carboxyl group which results in the creation of defective site with two adjacent Zr atoms. Experimental observations indicate that during the postsynthetic treatment with H₂SO₄ defects creation occurs on the sulfonate acid group side due to its weaker coordination to the Zr metal center. Once the linker is protonated by the acid it is in the dangling, metastable state and the HSO₄⁻ anion adsorbs on the coordinatively unsaturated Zr centers (Figure 5, bottom). The dangling linker is always stabilized by hydrogen bond interaction of moderate strength with the neighboring μ_3 -OH group. The stability of the material with dangling, protonated PSBA linker was calculated by the following expression: $\Delta E = (E_{\text{HI-UiO-66-H}_2\text{SO}_4}) - (E_{\text{HI-UiO-66}} + E_{\text{H}_2\text{SO}_4})$, in which $E_{\text{HI-UiO-66-H}_2\text{SO}_4}$ stands for either protonated carboxyl or sulfonate group of PSBA. The electronic structure calculations show that the protonated sulfonate group with adsorbed HSO₄⁻ anion is around 17 kJ/mol more stable than the protonated carboxyl group. This indicates that upon postsynthetic treatment the defects creation is initiated by the cleavage of Zr-sulfonic acid group bond. The sulfonate group is a weaker base than the carboxyl group and therefore it is the best leaving group. These calculations confirmed that the

creation of defects in HI-UiO-66 occurs preferentially by removal of PSBA linkers. Second, to gain computational insight in the origin of the increased stability of HI-UiO-66 after treatment, we calculated the free energies of defective and nondefective structures of HI-UiO-66 within the harmonic approximation. In contrast to the electronic energy, the free energy includes contributions due to the normal modes such as vibrational entropy. In this way, temperature effects could be estimated. We have found that, at 300 K, defective structures have an entropy that is about 100 J/mol/K higher than nondefective structures. Even when considering the approximation of our approach, this large entropy difference is a clear indication that defective structures are entropically favored above nondefective structures at high temperatures.

Lewis and Brønsted Sites, Catalytic Activity. The defects in UiO-66 have been widely studied but their role of acidity is not yet always clear.^{2,48} De Vos and coworkers reported an approach using trifluoroacetic acid as a modulator to create defects as Lewis acid sites. The Lewis acid catalyzed Meerwein–Ponndorf–Verley reaction of 4-tert-butylcyclohexanone (TCH) and isopropanol was used as a probe reaction, and the nature and amount of the acid sites was measured by IR-monitored CD₃CN chemisorption.¹⁴ Hereafter, the defects as Lewis acid sites have been explored frequently.² The defects as Brønsted acid sites were evaluated by potentiometric acid–base titration and then correlated with the Brønsted acid catalyzed epoxide ring-opening reaction with alcohol.^{49,50} Recent discoveries about UiO-66 have shown that after defect engineering the material possesses not only Lewis acid sites but also a significant number of Brønsted sites which induce a very dynamic acidity.^{51–53} This property of the Brønsted sites in UiO-66 plays a crucial role in the reactions where proton rearrangement can occur.^{54–56} It is critical to identify the roles of these defects in catalysis, either Lewis sites or Brønsted sites or Lewis sites and Brønsted sites together. Several studies characterized the defects using Fourier transform infrared (FT-IR) or FT-IR with CO or titration.^{14,49,57,58} Herein, the isomerization of α -pinene oxide as probe reaction was examined to identify and verify the roles of defect under *in situ* catalytic condition.

α -Pinene oxide is a very sensitive substrate which rearranges easily under the influence of a Lewis or Brønsted acid based catalyst. One of the formed products during the Lewis-acid-catalyzed rearrangement of α -pinene oxide is campholenic aldehyde (CA), an intermediate in the production of sandalwood fragrance and santalol.^{59,60} The other major products produced during this acid-catalyzed isomerization are summarized in Scheme 1. The selectivity during the isomerization of α -pinene depends highly on the type of acid sites present in the reaction medium. This makes the reaction a powerful probe to determine the types of acid sites present in the catalysts.⁶¹ Brønsted acids behave differently than Lewis acids for the formation of CA. More specifically, a high selectivity toward CA (up to 85%) was obtained using catalysts with only Lewis acid sites, whereas for the Brønsted acid catalyzed reaction this selectivity is lower than 55% and a significant amount of trans-carveol is formed.⁶¹

The catalytic activity of HI-UiO-66 was examined for the isomerization of α -pinene oxide in toluene at a reaction temperature of 70 °C. Prior to the catalytic tests, the materials were activated under vacuum at 150 °C for 16 h. As can be seen from Figure 6, the pristine UiO-66 exhibited a conversion of α -pinene oxide of 54% after a reaction time of 3 h, whereas

Scheme 1. Formation of Different Products during the Isomerization of α -Pinene Oxide

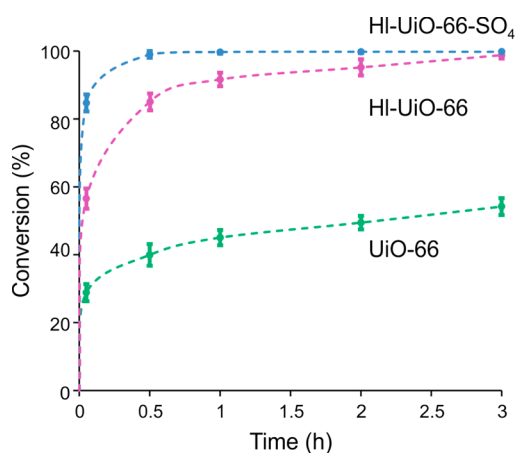
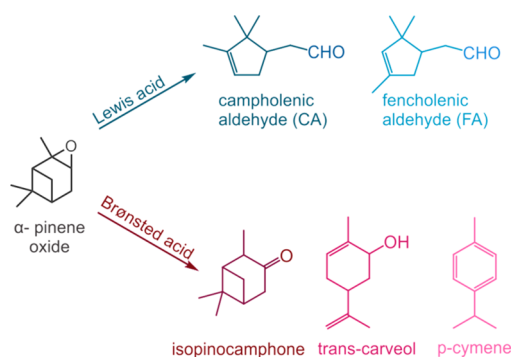


Figure 6. Conversion vs time for the isomerization of α -pinene oxide over UiO-66 and HI-UiO-66 and HI-UiO-66-SO₄.

for HI-UiO-66, full conversion was obtained under the same reaction conditions. Remarkably, for HI-UiO-66 and HI-UiO-66-SO₄, 85 and 99% conversion was achieved respectively within only 30 min of reaction time (Figure 6). Therefore, the catalytic activity follows the order: HI-UiO-66-SO₄ > HI-UiO-66 > UiO-66, corroborating that the number of defects influences the rate of the isomerization reactions.

As can be seen from Figure 7, the HI-UiO-66 and HI-UiO-66-SO₄ exhibit a higher selectivity (>55%) to CA than the pristine UiO-66 material (47%). The observed reactivity of HI-UiO-66 and HI-UiO-66-SO₄ are comparable to that of other Lewis acid based catalysts (such as Cu-BTC) while the reactivity of UiO-66 is comparable to that of Brønsted acid catalysts such as Dowex 50Wx4-100.⁵⁹ To understand the selectivity in these three Zr-MOFs, it is necessary to know the nature of the active sites, whether they reveal Lewis or Brønsted acidity.

The acid character of all three MOFs was assessed by the adsorption of acetonitrile-d₃ (CD₃CN) followed by FT-IR. Figure 8 (left) shows the difference in IR spectra of adsorbed CD₃CN on the evacuated MOFs (normalized to 10 mg/cm²), in which three ν (CN) vibrational bands can be observed at 2299, 2272, and 2260 cm⁻¹. These bands can be attributed to the adsorption of CD₃CN on Lewis and Brønsted acid sites and to physisorbed CD₃CN, respectively.^{14,62} Lewis acid sites in UiO-66 are associated with the dehydroxylation process of the inorganic cluster and are present in all the studied MOFs. However, small differences in the position of the band and in its intensity are observed in the spectra. These bands appear on

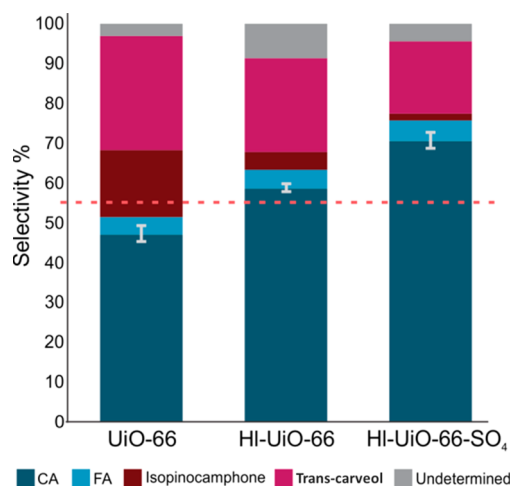


Figure 7. Product selectivity for the isomerization of α -pinene oxide at 100% conversion over UiO-66 and HI-UiO-66 and HI-UiO-66 washed with H_2SO_4 solution. The dotted line indicates the selectivity of 55%.

the modified UiO-66 samples at a slightly higher wavenumber (2299 cm^{-1}) than on the spectrum of the pristine UiO-66 (2296 cm^{-1}), which suggests a higher Lewis acid strength.⁶³ Furthermore, the intensity of such band decreases in the order $\text{HI-UiO-66-SO}_4 > \text{HI-UiO-66} > \text{UiO-66}$.

In Figure 8 (right), the normalized band area of all the studied MOFs shows this clear trend, which is a direct indication of the enhancement in the number of defects that can be created by the use of hemilabile ligands and further treatment with H_2SO_4 . These results are in agreement with the catalytic activity test and the higher selectivity to CA exhibited by the samples HI-UiO-66-SO₄ and HI-UiO-66. An opposite trend is obtained when the band at 2272 cm^{-1} is analyzed. This band is related to the amount of Brønsted acid sites in the structure that can originate from the $\mu_3\text{-OH}$ groups of the inorganic cluster.⁶³ Whereas this band is only present as a shoulder on the spectra of the samples HI-UiO-66-SO₄ and HI-UiO-66, it is quite prominent in the spectrum of the pure UiO-66, which indicates a higher relative amount of Brønsted acid sites on the latter. Such trend is also easily visualized in Figure 8 (right) and the results also correlate with the catalytic activity

tests where UiO-66 behaved more like a Brønsted acid, forming CA with a selectivity of only 47%.

A perfect UiO-66 contains four Brønsted acidic sites ($\mu_3\text{-OH}$) per cluster and no Lewis acidic sites. Yang et al. showed that one defect site in UiO-66 results in one Lewis (Zr vacancy) and one Brønsted (Zr-OH) acid site (Figure S20, Scheme 1).^{57,64,65} But, upon thermal treatment, the two Brønsted sites (Zr-OH and $\mu_3\text{-OH}$) in this defective UiO-66 will dehydroxylate annihilating both Brønsted sites and forming a Zr-O-Zr linkage. As a result, a dehydroxylated UiO-66 with x defects contains $4 - x$ Brønsted acid sites and x Lewis acid sites. In our case, the UiO-66 contains 1.7 defects resulting in 2.3 Brønsted acidic sites and 1.7 Lewis acidic sites per cluster with a ratio of Brønsted to Lewis acidic sites of 1.35 (2.3/1.7). Similarly, HI-UiO-66 contains 1.8 Brønsted acidic sites and 2.2 Lewis acidic sites per cluster with a ratio of 0.82 and HI-UiO-66-SO₄ contains 1.0 Brønsted acidic sites and 3.0 Lewis acidic sites per cluster with a ratio of 0.33. It is clear that upon increasing the number of defects, the amount of Lewis acidic sites increases, resulting in the enhanced selectivity to CA. Since HI-UiO-66-SO₄ has the highest number of defects, and the highest ratio of Lewis sites to Brønsted sites, it has the highest selectivity toward CA in comparison to the other two catalysts. These observations are consistent with the results obtained from CD_3CN adsorption analysis. The Brønsted acid sites are dominant in UiO-66 with a low defect density while the Lewis acid sites are prominent in HI-UiO-66-SO₄ with a high defect density.

To compare the reactivity of these Zr-MOF materials with the other heterogeneous MOF based catalysts and zeolites, we summarize the reactivity and selectivity of some published catalysts in Table S3. From this table, HI-UiO-66-SO₄ shows the best activity and selectivity in comparison to the other reported MOF catalysts. In comparison to $\text{Cu}_3(\text{BTC})_2$,^{59,66} which only has Lewis acid sites, a higher catalytic activity but a lower selectivity was obtained. However, this catalyst is not stable and cannot be recycled. Compared to the best homogeneous Lewis-acid catalyst (ZnBr_2),⁶⁷ the HI-UiO-66-SO₄ exhibits a similar selectivity but with a higher turnover.

In addition, a hot filtration test after 3 min of reaction revealed that after the removal of the catalyst no further conversion was noted, confirming the heterogeneous nature of the catalyst (Figure S23). Nevertheless, during the recycling experiments, the activity and selectivity decreased slightly

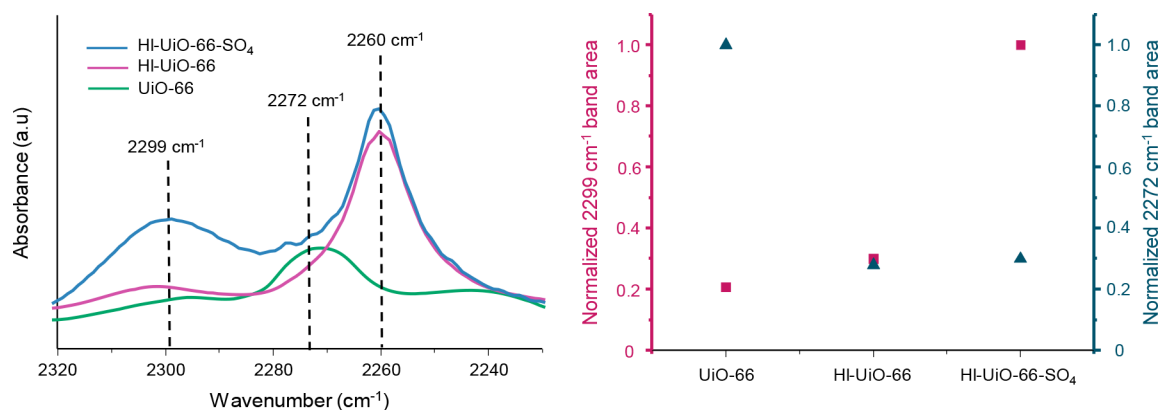


Figure 8. Left: Difference IR spectra of adsorbed CD_3CN (normalized to $10\text{ mg}/\text{cm}^2$). CD_3CN was adsorbed at room temperature for 1 h and desorbed under vacuum for 1 h. Right: Amount of Brønsted (triangles) and Lewis (squares) acid sites detected by CD_3CN , expressed as normalized area of the bands at 2272 and 2299 cm^{-1} , respectively.

during each run (Figure S24). This decrease might be due to the blocking of the pores during the consecutive runs, as the surface area slightly decreased during each run (Figure S26). PXRD measurements showed that the structure remained intact (Figure S25).

CONCLUSIONS

In summary, we presented a stable yet highly defective Zr-MOF, obtained using the hemilabile linker strategy. PSBA not only acts as a modulator for the MOF synthesis to increase the number of defects but also as coligand to stabilize the defective structure. The HI-UiO-66 is one of the first reported frameworks that further stabilizes (instead of the typical destabilization) as the number of defects increases. A simple postsynthetic treatment with H₂SO₄ even further enhances the stability while the number of defects sites increases again. This HI-UiO-66-SO₄ showed the optimum number of defects (while maintaining structural integrity) and showed the highest reactivity for the acid-catalyzed isomerization of α -pinene oxide with a high selectivity for campholenic aldehyde. We also found that the increase in the number of defects increases the Lewis acid characteristics and hence favors the campholenic aldehyde formation. By controlling the number of defects, we can tune the acid behavior of the catalysts, as a low number of defects leads to a Bronsted dominated catalyst, and vice versa. The nature and concentration of the defective sites were further elucidated by IR-monitored CD₃CN sorption. The use of a hemilabile linker, either as mixed linker or as the isolated linker, is an interesting concept in defect engineering of MOFs. Due to the different possibilities in coordination ability of the two different groups in a hemilabile linker, defects can be generated and hence the stability can be affected. Meanwhile, the easy postsynthetic treatment can be useful to adapt and adjust the final structure of the pristine materials for potential applications.

ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge at <https://pubs.acs.org/doi/10.1021/jacs.9b13070>.

Detailed catalysis results and data analysis results including PXRD, VTXRD, TGA-DSC, dissolution/¹H NMR, nitrogen adsorption IR, catalysis results, and the modeling data (PDF)

Structures (ZIP)

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Notes

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