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PAPER

New polyoxometalate-functionalized cellulosic fibre/silica hybrids for environmental applications

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Cellulosic fibre/silica hybrid materials functionalized with Keggin-type polyoxometalates ($[\text{PV}_2\text{Mo}_{10}\text{O}_{40}]^{5-}$, $[\text{PVMo}_{11}\text{O}_{40}]^{4-}$, $[\text{PMo}_{12}\text{O}_{40}]^{3-}$ or $[\text{PW}_{12}\text{O}_{40}]^{3-}$) were prepared by a sol–gel method at room temperature. The novel materials are composed of *ca.* 56 wt% of polysaccharides, *ca.* 37 wt% of propylamine-modified silica, 2 wt% of polyoxometalate, and 5% of hydration water. The silica network of these hybrids exhibits relatively high degree of condensation being distributed mainly on the surface of the cellulosic fibres as a dense film. The functionalization of silica with polyoxometalates *via* electrostatic interactions with protonated propylamino groups of modified silica was unambiguously confirmed. Despite their high silica content cellulose/silica hybrids retained basic cellulosic pulp properties—supramolecular and fibrous structure, porosity, relatively low density, *etc.* The novel bio-based material functionalized with 2% of $[\text{PVMo}_{11}\text{O}_{40}]^{4-}$ shows particularly high activity towards the oxidation of volatile organic compounds (VOCs) present in urban air thus anticipating future environmental applications.

Introduction

The combination of an inorganic phase with organic polymers at nano-, meso- or micro-scale affords organic/inorganic hybrid materials. These hybrids have received a great deal of attention in the last decade due to the unique characteristics obtained by the synergistic combination of inorganic and organic counterparts, with complementary properties.^{1–3} Furthermore, hybrids based on natural polymers are especially attractive because of the biodegradability and renewable properties of the latter that reduce end-of-life disposal problems. In particular, *cellulose* is the most abundant natural polymer with numerous technical applications from papermaking and plastics to biomedicine.⁴ Value-added applications of cellulosic materials are substantially extended upon appropriate chemical modification of native polymer (etherification, esterification, silylation, *etc.*)^{4–11} Among recent trends, cellulose derivatives of particular interest also include cellulose-based hybrids comprising an inorganic phase chemically linked or physically bound to cellulose.^{12,13}

Hybrid materials of cellulose or chemically modified cellulose with silica are the most widely reported cellulose-based hybrids.^{14–26} Silica, a bioactive and biocompatible inorganic material, can improve the thermal stability of cellulose, its lipophilic behavior,

and its affinity towards specific substrates.^{17–19,24–26} The incorporation of silica into a cellulosic matrix has been reported for hydroxypropyl cellulose¹⁴ and cellulose acetate,¹⁵ affording higher mechanical resistance of the ensuing materials. A silica based hybrid material with mechanical properties similar to those of cortical bones has been produced from cellulose acetate, which has bioadsorption properties.¹⁶

Recently, cellulose/silica hybrids with potentially attractive mechanical and thermal properties have been synthesized by a mild sol–gel method, using heteropolyacids as catalysts.¹⁸ Bleached kraft cellulosic pulp was used as cellulose source and tetraethoxysilane as the silica precursor. Silica was deposited on the fibres in the form of a thin film or as discrete particles attached to the cellulose surface predominantly by non-covalent bonds. Under optimized synthesis conditions roughly 40–60 wt% of silica was incorporated into the cellulosic material considerably diminishing its hydrophilicity and improving thermal stability. These were considered as prospective materials for thermal insulation or packing purposes.¹⁹

The surface modification of cellulose with pre-hydrolysed/condensed alkoxy silanes having variable functions (propylamine, alkyl, and others) is expected to influence the hydrophilicity/hydrophobicity balance and the electron donor/acceptor properties of the ensuing hybrid materials.^{21–25} Particularly, hydrophobization of cellulose has been reported by using alkyl-substituted silanes^{19,25} whereas the modification of silica with triazine derivatives has been reported to improve target fixation of reactive dyes in cellulose/silica hybrids.²⁶ The introduction of propylamino moieties on the surface of cellulose-silica hybrids, which can be achieved applying tetraethoxysilane and triethoxypropylamine

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silane in the sol–gel synthesis, could be useful to enable the electrostatic surface anchoring of polyoxometalates (POM).^{27–29} In this way, POMs with oxidative or acid catalytic properties,^{30–32} may be incorporated thus conferring catalytic properties to the cellulose-based hybrids.

To the best of our knowledge, cellulose fibres have never been reported as support material for POMs, although cellulose acetate was used as support for Keggin heteropolyacids ($\text{H}_3\text{PW}_{12}\text{O}_{40}$ and $\text{H}_3\text{PMo}_{12}\text{O}_{40}$).³³ Furthermore, Keggin polyoxometalate ($\text{H}_4\text{SiW}_{12}\text{O}_{40}$) nanotubes have been prepared by calcining layer-by-layer structured hybrid films coated cellulose acetate nanofibres.³⁴ In the present work, new cellulose/silica hybrid materials incorporating polyoxometalates, $[\text{PV}_2\text{Mo}_{10}\text{O}_{40}]^{5-}$, $[\text{PVMo}_{11}\text{O}_{40}]^{4-}$, $[\text{PMo}_{12}\text{O}_{40}]^{3-}$ and $[\text{PW}_{12}\text{O}_{40}]^{3-}$ as redox functionalities were synthesized and characterized. The catalytic activity of some of these hybrids for the oxidation of volatile organic compounds (VOCs) present in polluted urban air is revealed.

Experimental

Materials and methods

" $\text{H}_5[\text{PV}_2\text{Mo}_{10}\text{O}_{40}] \cdot 11\text{H}_2\text{O}$ " (**1**) and " $\text{H}_4[\text{PVMo}_{11}\text{O}_{40}] \cdot 10\text{H}_2\text{O}$ " (**2**) were prepared following known procedures.³⁵ It is noteworthy that, following the literature procedures, the heteropolyacid **1**, in addition to the several isomers of $[\text{PV}_2\text{Mo}_{10}\text{O}_{40}]^{5-}$, also contained some amount of α - $[\text{PVMo}_{11}\text{O}_{40}]^{4-}$ (25 mol%) and $[\text{PV}_3\text{Mo}_9\text{O}_{40}]^{6-}$ (8 mol%), as determined by ³¹P NMR spectroscopy in aqueous solutions. The heteropolyacid **2**, besides α - $[\text{PVMo}_{11}\text{O}_{40}]^{4-}$, also contained $[\text{PV}_2\text{Mo}_{10}\text{O}_{40}]^{5-}$ (15%). These are common chemical compositions for these heteropolyacids obtained by the ether-extraction method.^{35–37} Notwithstanding, since $[\text{PV}_2\text{Mo}_{10}\text{O}_{40}]^{5-}$ and α - $[\text{PVMo}_{11}\text{O}_{40}]^{4-}$ are by far the dominant polyoxoanions present in **1** and **2**, respectively, the latter compounds may be described by the formulas above in quotes as a matter of simplification. $\text{H}_3[\text{PMo}_{12}\text{O}_{40}] \cdot n\text{H}_2\text{O}$ (**3**), $\text{H}_3[\text{PW}_{12}\text{O}_{40}] \cdot n\text{H}_2\text{O}$ (**4**), tetraethoxysilane (TEOS, 98% w/w), 3-aminopropyltriethoxysilane (APTS, 95% w/w) and ethanol (99.5% w/w) were supplied by Sigma-Aldrich and were used as received. *Eucalyptus globulus* industrial bleached kraft pulp, used as cellulose fibre source, was supplied by Portugal (Portugal). This pulp (hereafter mentioned as cellulosic pulp or cellulose) contained about 85 wt% (on oven-dry pulp weight) of cellulose and 14 wt% of glucuronoxylan. Prior to its use, 1.5 g of air dried cellulosic pulp was disintegrated (swollen) in 150 mL of water for 48 h, filtered and then washed with two 50 mL portions of ethanol to remove the excess of water.

²⁹Si and ³¹P solid-state magic angle spinning (MAS) NMR spectra and ¹³C–¹H solid-state MAS NMR with cross-polarization (¹³C CP/MAS NMR) spectra were recorded on a 9.4T Bruker Avance-400 spectrometer, as previously reported.³⁸ Chemical shifts in the ²⁹Si and ¹³C NMR spectra were referenced to TMS, using kaolinite and glycine, respectively, as external references. For the ³¹P NMR spectra, the chemical shifts are referenced to phosphoric acid (85%). A large number of scans were required to obtain ²⁹Si and ³¹P NMR spectra of acceptable quality due to the low content of silica and polyoxometalate in the hybrid materials.

The analyses of Mo, W, V and P were carried out by ICP spectrometry. For that, the hybrid materials were previously

digested in a mixture of HF/HNO₃/H₂O₂. Elemental analysis (C, H and N) was done using a Leco CHNS-932 apparatus. Thermogravimetric analyses were carried out between 30 and 700 °C under nitrogen at a heating rate of 5 °C min⁻¹ on a TGA-50 Shimadzu thermobalance. Infrared absorption spectra of samples in KBr pellets were recorded on a Mattson 7000 FTIR spectrometer. Diffuse reflectance spectra were registered on Jasco V-560 spectrophotometer, using MgO as reference.

Scanning electron microscopy (SEM) images were obtained on a FEG-SEM Hitachi S4100 microscope coupled with an energy dispersive X-ray spectrometer (EDS) and operating at 25 kV, using gold coated samples.

Mercury intrusion porosity was determined using the AutoPore IV 9500 instrument from Micromeritics. The pressure range varied between 0.5 psi and 30 000 psi, which enabled the measurement of pore diameters between 350 μm and 0.006 μm, respectively. The total porosity was calculated as the ratio of the total intruded volume and the sample bulk volume.

Wide angle X-ray diffraction data were collected using a Philips X'Pert diffractometer operating in the Bragg–Brentano configuration with Co-kα ($\lambda = 1.79 \text{ \AA}$) source at a current of 35 mA and an accelerating voltage of 40 kV. Data were collected by the step counting method (step 0.025° and time 1.0 s) in the 2θ range of 4–50°. Pellets of about 1 cm diameter and 1 mm thickness were prepared for the analysis by pressing the samples at 50 MPa during 3 min. The peak fitting was performed using the program Profile Fit (Philips Analytical X-ray Almelo, 1996, version 1.0).

Synthesis of cellulose/silica hybrid materials functionalized with polyoxometalates

A mixture of 37.5 mL of TEOS, 1.55 mL of APTS and 20.5 mL of ethanol was added to 10.2 mL of water acidified with 70 μL of concentrated nitric acid (65%) under mechanical stirring (200 rpm). Thereafter, while keeping the reaction flask in a water bath at room temperature (*ca.* 20 °C), 800 μL of concentrated nitric acid was added affording a pH \cong 1. The resulting solution was left reacting for 40 min, the stirring rate was increased to 350 rpm and ammonia (2.5% w/w, *ca.* 3.5 mL) was slowly added in 250 μL portions until the pH \cong 3. Subsequently the stirring rate was reduced to 200 rpm and the resultant sol was allowed to react for 10 min. To this formulation 1.5 g of previously swollen cellulosic pulp was added in small portions and the suspension left reacting for 5 min more. Thereafter, POM (**1** or **2**, 0.10 g) or (**3** or **4**, 0.30 g), dissolved in 2 mL of ethanol was added to the suspension and the reaction left running for further 15 min for POM **1** or **2** or 30 min for POM **3** or **4**. The final material was filtered on a glass G₃ filter and dried under vacuum at room temperature.

An unmodified cellulose/silica hybrid (free of propylamino moieties) was prepared by adding 53.0 mL of TEOS to 16 mL of water containing 60 μL of concentrated nitric acid (65%). The emulsion was stirred at 350 rpm for 1 h, at room temperature, and then 200 μL of ammonia (2.5% w/w) was added affording a pH \cong 3. After 5 min, the stirring rate was reduced to 200 rpm and 1.5 g of previously swollen cellulosic pulp was added. The suspension was left reacting for another 5 min and an aqueous solution of POM (**1**, 0.10 g in 2 mL) was added and allowed to

react for 15 min more under stirring. The solid material was filtered on a glass G₃ filter and dried under vacuum at room temperature.

Oxidation of volatile organic compounds (VOCs) present in urban air

Urban air-containing sample was collected in an urban zone near the North entrance of the University of Aveiro (Portugal), during the sunny day of 18 August 2009 from 16:00 to 17:00 h.

Teflon test tubes (20 cm × 0.4 cm i.d.) were filled with a known amount of the cellulose/aminopropylated silica hybrid materials functionalized and non-functionalized (control experiment) with POMs. Urban air was pumped through the Teflon tubes at a flow rate of 200 mL min⁻¹ and the C₅–C₁₁ VOCs in the effluent stream were collected in a stainless steel tube trap (15 cm × 6.0 mm o.d. × 5.2 mm i.d.) filled with 150 mg of Tenax-TA (60–80 mesh) and 150 mg of Carboxen B (60–80 mesh). The samples were desorbed and analysed by the thermal desorption/cryogenic concentration method, on a Crompack® CP 9000 gas chromatograph (GC) equipped with Chrompack Thermal Desorption–Cold Trap Injection (TCT) unit and a flame ionization detector (FID) as described previously.³⁹

Results and discussion

Synthesis of cellulose/silica hybrids functionalized with polyoxometalates

Novel cellulose-based hybrid materials containing silica modified with propylamino groups and small amount of polyoxometalate were prepared by sol–gel technique in an aqueous-ethanolic medium at room temperature (*ca.* 20 °C). Two main steps for the formation of the silica network in the cellulosic material were involved: (i) hydrolysis of appropriate amounts of TEOS and APTS (molar ratio 96/4) precursors at pH ≈ 1; (ii) condensation of the hydrolysed precursors affording a sol at pH ≈ 3 followed by the addition of bleached cellulosic pulp to the suspension in order to obtain the cellulose/modified silica material. The polyoxometalate is thereby incorporated into the hybrid material *via* electrostatic interaction with the protonated amino groups of the modified silica. It should be noted that, for the polyoxoanions [PMo₁₂O₄₀]³⁻ (PMo₁₂) and [PW₁₂O₄₀]³⁻ (PW₁₂) a higher amount of POM was added to the reaction medium, taking into account the lower charge of these polyoxoanions when compared to the [PVMo₁₁O₄₀]⁴⁻ (PVMo₁₁) and [PV₂Mo₁₀O₄₀]⁵⁻ (PV₂Mo₁₀) analogues.

The cellulose/silica functionalized materials containing PVMo₁₁ and PV₂Mo₁₀ present a bright yellow colour, due to the presence of polyoxometalate, while those of PMo₁₂ and PW₁₂ were pale green and white, respectively (Fig. 1). The hybrid with PVMo₁₁ was found to be highly sensitive to the atmosphere as slow change to green colour could be observed upon brief exposure to laboratory air. This was due to the reduction of PV^VMo₁₁ to PV^{IV}Mo₁₁.⁴⁰ However, when the latter solid was immediately stored in an exsiccator colour change was not observed. The extensive washing of the new hybrid materials with water or ethanol did not reveal significant leaching of the polyoxometalates (less than 1% of the POM was extracted into the solvent).

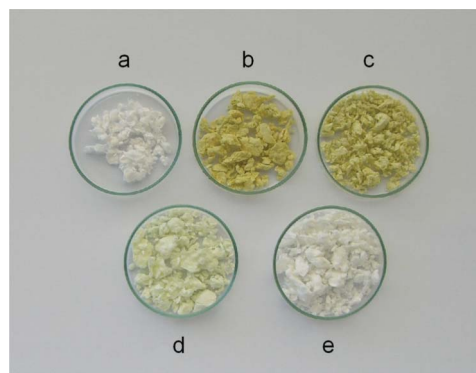


Fig. 1 Photograph of cellulosic pulp (a) and cellulose/silica hybrids functionalized with PV₂Mo₁₀ (b), PVMo₁₁ (c), PMo₁₂ (d) and PW₁₂ (e).

For comparison, a hybrid material was prepared without APTS, *i.e.*, using TEOS as the only precursor, and PV₂Mo₁₀ as polyoxometalate. Under these conditions, the material obtained was pale yellow with a very low amount of polyoxometalate (Table 1), which is easily leached. This confirms the role of the amino groups in promoting the chemical linkage between polyoxometalate and silica.^{38,41}

Characterization of polyoxometalate-functionalized cellulose/silica hybrids

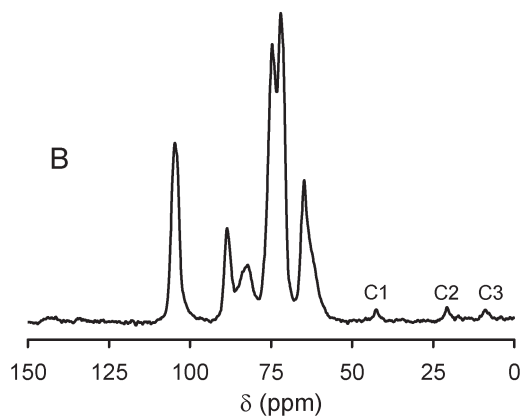
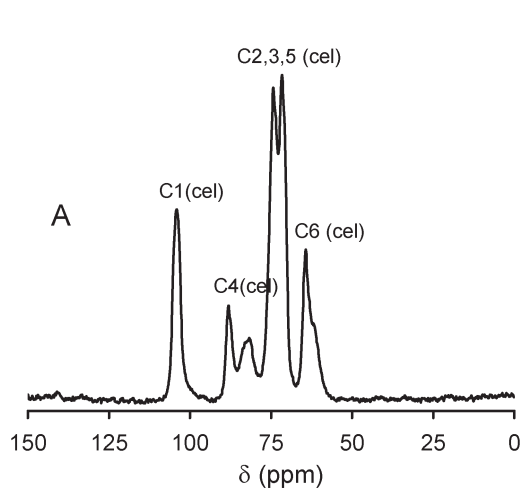
The novel materials were characterized by several spectroscopic and analytical techniques. Characterization by solid-state NMR provided important information regarding the structural features of the different components of the materials. ¹³C CP/MAS NMR spectrum (Fig. 2) showed six high intensity signals in the 50–120 ppm range, assigned to known carbon resonances of cellulose framework.⁴² The position and relative intensity of these resonances in the hybrids were similar to those observed in the spectrum of initial cellulosic fibres. Hence, this suggests that the chemical structure of cellulose was maintained in the hybrid materials and that, apparently, no covalent bonding between the cellulose and silica framework has been established. In addition, three signals of low intensity were observed at 42.5, 21 and 9 ppm (Fig. 2). These are due to the C1, C2 and C3 carbon resonances, respectively, of the propylamino moieties (–H₂N–C(1)H₂–C(2)H₂–C(3)H₂–Si).^{38,43} According to previous studies of Caravajal *et al.*⁴³ the C2 chemical shift at 21 ppm is an indication of protonation of the NH₂ group. Otherwise, for neutral amino groups the C2 resonance would appear at 27 ppm.⁴³ Thus, the ¹³C NMR results indicate the presence of propylamino moieties protonated at the NH₂ group. This result is expected because the reactions were performed in acidic conditions (the pH values of the reaction filtrates were about 1.5–2.0).

Information on the silica counterpart of the hybrids was provided by ²⁹Si NMR. ²⁹Si MAS NMR spectra of the hybrid materials showed, in the –80 to –120 ppm region, three signals at –92 ppm (8%), –101 ppm (50%) and –110 ppm (42%) (Fig. 3), corresponding to (SiO)₂Si(OH)₂ (Q²), (SiO)₃SiOH (Q³) and (SiO)₄Si (Q⁴) groups, respectively.^{41,43,44} According to the relative abundance of Qⁿ structures, the results here obtained indicate reasonably high degree of condensation of the bound

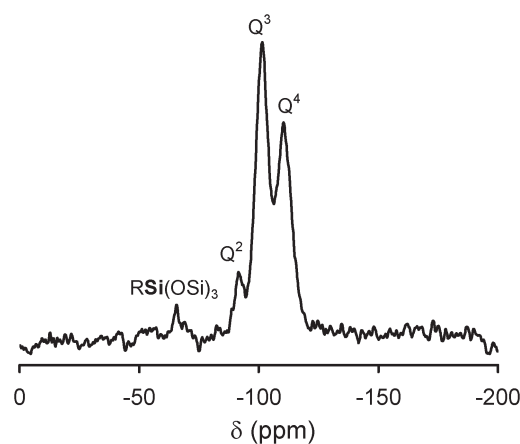
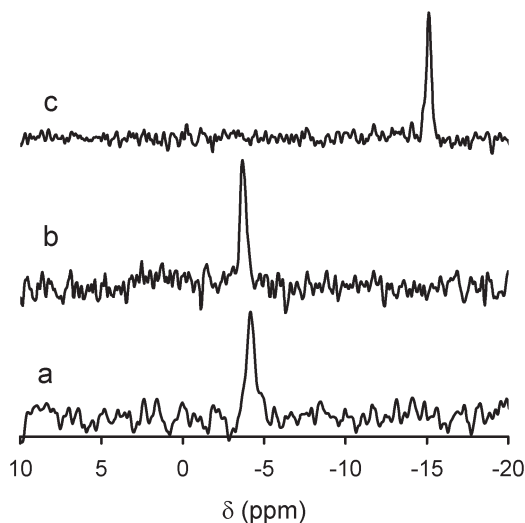
Table 1 Analyses results for cellulose/silica hybrids incorporating polyoxometalates

| Material | Mo,W (%) | V (%) | Mo/V (molar ratio) ^c | POM (%) ^d | Cellulose (%) ^e | SilicaNH ₃ (%) ^f | N (%) | Propylamine (%) ^g |
|---|----------|-------|---------------------------------|----------------------|----------------------------|--|-------|------------------------------|
| Cell/SiO ₂ (C ₃ H ₆ NH ₃)-POM ^a | | | | | | | | |
| PV ₂ Mo ₁₀ | 1.41 | 0.14 | 5 (5) | 2.5 | 54.3 | 37.0 | 1.12 | 2.5 |
| PVMo ₁₁ | 1.05 | 0.062 | 9 (11) | 1.8 | 55.0 | 37.7 | 1.08 | 2.4 |
| PMo ₁₂ | 0.65 | — | — | 1.0 | 58.1 | 35.4 | 1.02 | 2.2 |
| PW ₁₂ | 1.05 | — | — | 1.4 | 56.8 | 35.6 | 1.00 | 2.1 |
| Cell/SiO ₂ -POM ^b | | | | | | | | |
| PV ₂ Mo ₁₀ | 0.39 | 0.042 | 5 (5) | 0.7 | 50.6 | 43.9 ^h | — | — |

^a Cellulose/propylamine-modified silica materials incorporating different polyoxometalates. ^b Hybrid material prepared without modified silica, *i.e.*, experiment performed with no APTS. ^c In parenthesis, the value expected for the corresponding molybdovanadophosphate anion. ^d Based on the Mo (or W) analysis. ^e Based on the thermogravimetric analysis. ^f Obtained by difference (100-solvent (%)-cellulose (%)-POM (%)). ^g Based on the nitrogen elemental analysis and charge balance equations. ^h Amount of unmodified silica.

**Fig. 2** ¹³C CP/MAS NMR spectra of initial cellulosic pulp (A) and of cellulose/silica hybrid functionalized with PVMo₁₁ (B).

silica frameworks (this being, however, incomplete), being similar to those reported for the cellulose/silica hybrid materials obtained by heteropolyacid catalysed sol-gel process¹⁸ and for silica materials containing POM clusters.⁴⁴ In the ²⁹Si MAS NMR spectra of the hybrids (Fig. 3), in addition to the Qⁿ peaks, a signal centred at *ca.* -66 ppm was also observed. This signal assigned to RSi(OSi)₃ (R = CH₂-CH₂-CH₂-NH₃⁺)^{43,44} confirms the linkage of propylamino moieties at the silica surface. No peak was observed at -45 ppm, which is the ²⁹Si

**Fig. 3** ²⁹Si MAS NMR spectrum of cellulose/silica hybrid functionalized with PV₂Mo₁₀.**Fig. 4** ³¹P MAS NMR spectra of cellulose/silica hybrids functionalized with PVMo₁₁ (a), PMo₁₂ (b) and PW₁₂ (c). Spectrum of PV₂Mo₁₀ hybrid (not shown) was very similar to the spectrum of PVMo₁₁ hybrid (a).

chemical shift that may arise from the ²⁹Si resonance of non-reacted APTS physically sorbed at the silica surface.⁴³

³¹P MAS NMR spectra showed single signals at -4.1 ppm for both PV₂Mo₁₀- and PVMo₁₁-incorporated cellulose/silica materials (Fig. 4). Similar results were obtained after immobilization

of the same polyoxometalates on propylamine-modified silica support,³⁸ thus confirming the presence of the molybdovanadophosphate anions in the hybrid materials. As was found for other materials,³⁸ no clear distinction between the two polyoxometalates, $[\text{PV}_2\text{Mo}_{10}\text{O}_{40}]^{5-}$ and $[\text{PVMo}_{11}\text{O}_{40}]^{4-}$, when immobilized in the cellulose/silica support could be made by this technique. In the case of $[\text{PV}_2\text{Mo}_{10}\text{O}_{40}]^{5-}$, which appears in solution mainly as a mixture of multiple positional isomers,³⁶ a possible structural rearrangement of $[\text{PV}_2\text{Mo}_{10}\text{O}_{40}]^{5-}$ to the most stable isomer/isomers accompanied by electrostatic interaction with the support material took place.³⁸ Cellulose/silica materials containing $[\text{PMo}_{12}\text{O}_{40}]^{3-}$ and $[\text{PW}_{12}\text{O}_{40}]^{3-}$ exhibited ^{31}P NMR signals at -3.7 ppm and -15.2 ppm, respectively (Fig. 4). This is in agreement with reported values for the corresponding heteropolyacids, for composites involving these Keggin heteropolyanions in cellulose esters³³ and with ^{31}P NMR data for related materials,⁴⁵ thus confirming the presence of the different Keggin species.

The Fourier transform infrared (FTIR) absorption spectra of the novel cellulose-based hybrid materials were identical to each other, *i.e.*, no differences due to the presence of different polyoxometalates could be found. Thus, the absorption bands due to the Keggin-type polyoxometalate structure, which usually appear in the region of wavenumbers lower than 1100 cm^{-1} ,⁴⁵ were masked by those of the cellulosic matrix and functionalized silica. This fact is not surprising taking into consideration the very low amount of POM in the hybrids (< 3 wt%, Table 1). However, notable differences were found when the spectra of the hybrids were compared to that of initial cellulosic pulp (Fig. 5). The FTIR spectrum of cellulosic pulp presents several characteristic bands, the most intense being between 1200 – 1000 cm^{-1} (mainly due to the C–O–C and C–O–H stretching vibrations⁴⁶). For the hybrids, additional bands appeared at *ca.* 450 cm^{-1} and 795 cm^{-1} (Fig. 5, marked with *) assigned to δ (O–Si–O) and ν_{s} (Si–O–Si), respectively.⁴⁴ The band expected at about 1080 cm^{-1} (ν_{as} (Si–O–Si)) was overlapped with the bands of cellulose in the same spectral region. In addition, a sharp band at 1385 cm^{-1} arose in the spectra of the hybrids, indicating the presence of nitrate anions.

Thermograms of hybrids showed three main steps of weight loss (Fig. 6): a first step up to about $120\text{ }^\circ\text{C}$ corresponding to

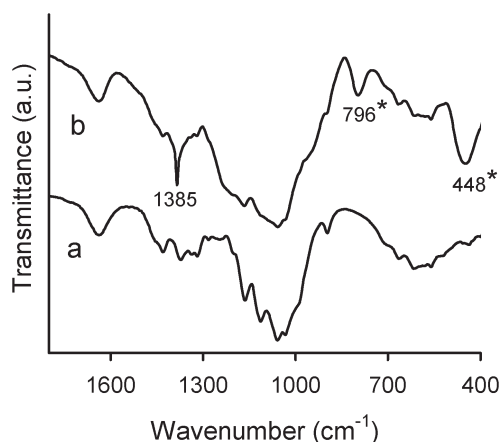


Fig. 5 FTIR spectra of cellulose pulp (a) and cellulose/silica hybrid functionalized with $\text{PV}_2\text{Mo}_{10}$ (b).

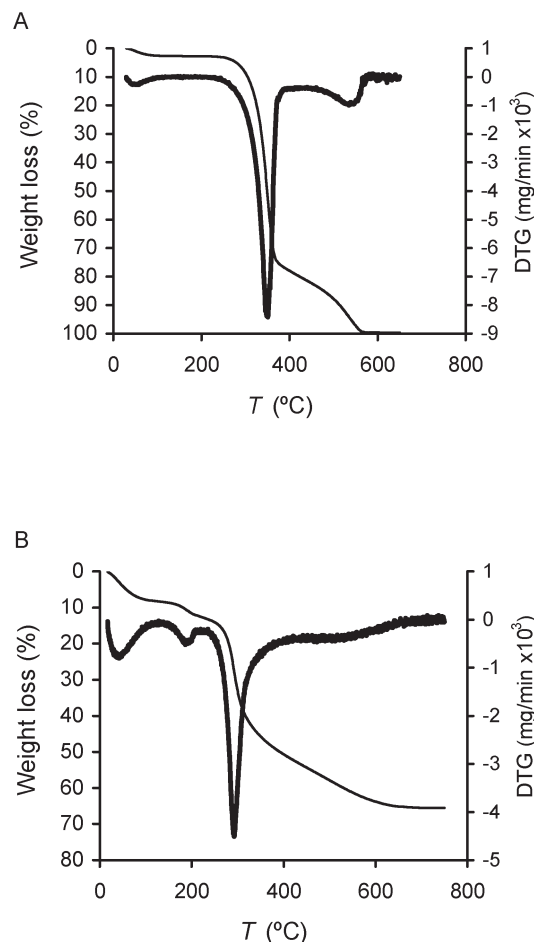


Fig. 6 Thermogravimetric curve (thin line) and corresponding derivative curve (DTG, thick line) of cellulose pulp (A) and cellulose/silica hybrid functionalized with $\text{PV}_2\text{Mo}_{10}$ (B).

5–8% weight loss, essentially due to moisture release; a second step between 150 and $230\text{ }^\circ\text{C}$ (4.5–5.5% weight loss), which is tentatively attributed to the decomposition of the protonated propylamino moieties and nitrate anions (not observed in the hybrid without APTS); a third step accounting for more than 50% weight loss, beginning at about $250\text{ }^\circ\text{C}$ and finishing at 600 – $700\text{ }^\circ\text{C}$, with maximum decay rate at *ca.* $295\text{ }^\circ\text{C}$ for the hybrids with $\text{PV}_2\text{Mo}_{10}$ and PVMo_{11} and at $310\text{ }^\circ\text{C}$ and $340\text{ }^\circ\text{C}$ for those with PMo_{12} and PW_{12} , respectively. It is noteworthy that cellulosic pulp presents its maximum decay rate at *ca.* $350\text{ }^\circ\text{C}$ (Fig. 6). Thus, it seems, regarding the last decomposition step that silica($\text{C}_3\text{H}_6\text{NH}_3^+$)-POM promoted, in general, the thermal decay of cellulose. Cellulosic pulp gave almost no char after the thermal degradation up to $800\text{ }^\circ\text{C}$ (less than 1%) suggesting the presence of cellulose with relatively high crystallinity in the starting material.⁴⁷ Accordingly, the thermogravimetric data enabled the approximate quantification of cellulose content in the cellulose/silica functionalized hybrids, which is of 54–58% (Table 1).

Based on the Mo (or W) elemental analysis the amount of polyoxometalate attached to the novel cellulose/silica functionalized materials was found to be of 1.0–2.5 wt% (Table 1). Moreover, the Mo/V molar ratio indicates that the $\text{PV}_2\text{Mo}_{10}$ and

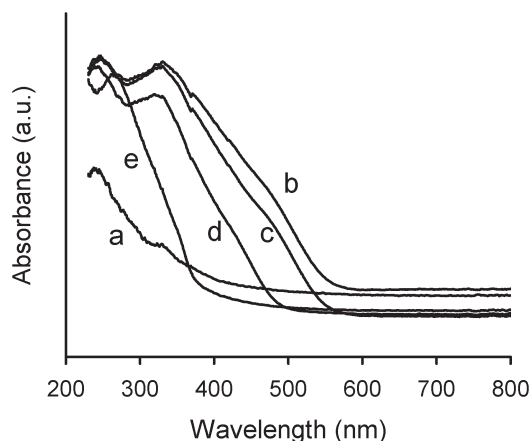


Fig. 7 Diffuse reflectance spectra of initial cellulosic pulp (a) and cellulose/silica hybrids functionalized with PV_2Mo_{10} (b), $PVMo_{11}$ (c), PMo_{12} (d) and PW_{12} (e).

$PVMo_{11}$ anions did not degrade under the experimental conditions used in the synthesis. On the other hand, the calculated amount of propylamine-modified silica was found to be of 35–38 wt% (Table 1). Based on the nitrogen elemental analysis and charge balance equations (*i.e.*, considering that the charge of protonated propylamine is balanced by the charge of nitrate anions and polyoxoanions) the amount of propylamino moieties was estimated to be of 2.1–2.5 wt% (Table 1). The amount of nitrate anions was estimated to be 2.2–2.3 wt%.

The cellulose/silica functionalized materials containing PV_2Mo_{10} , $PVMo_{11}$ or PMo_{12} showed significant absorption in

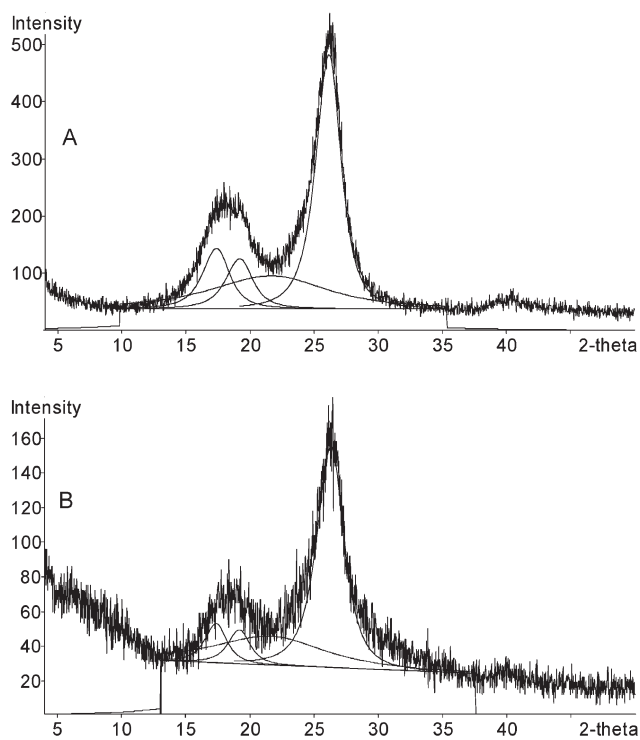


Fig. 8 X-ray diffractograms of pellets of initial cellulosic pulp (A) and cellulose/silica hybrid functionalized with PV_2Mo_{10} (B) showing fitted peaks from the crystalline reflections and the amorphous halo.

the visible region of the diffuse reflectance spectra, with a characteristic shoulder at *ca.* 470 nm (PV_2Mo_{10} and $PVMo_{11}$ hybrids) and *ca.* 420 nm (PMo_{12} hybrid) (Fig. 7). The absorption of visible wavelength confirms the presence of polyoxometalates,⁴⁵ since both initial cellulosic pulp (Fig. 7) and cellulose/modified silica hybrid without polyoxometalate do not absorb in the visible region (*i.e.*, both samples are white). In the UV region, two bands centred at 320–330 nm and 240–250 nm were detected. These UV bands are mainly due to the presence of polyoxometalates, with some contribution of other components of the hybrids (cellulose and propylamine-modified silica). For the PW_{12} hybrid (white material) only a single band centred at 265 nm was observed (Fig. 7).

It should be noted that heteropolyacids of PV_2Mo_{10} , $PVMo_{11}$ and PMo_{12} present broad bands in the UV-Vis spectra with absorption maximum at 400–420 nm (O→Mo) and the corresponding heteropolyacid with PW_{12} presents two bands in the UV region with maxima near 330 nm and 260 nm (O→W).^{40,45,48} Therefore, in comparison to the Keggin heteropolyacids, a blue shifting of the O→Mo and O→W charge transfer bands occurred after POM immobilization in the cellulose-based hybrids. The different chemical environments of the polyoxoanions in the heteropolyacids, where a complex network of water molecules and hydrated protons surrounds the

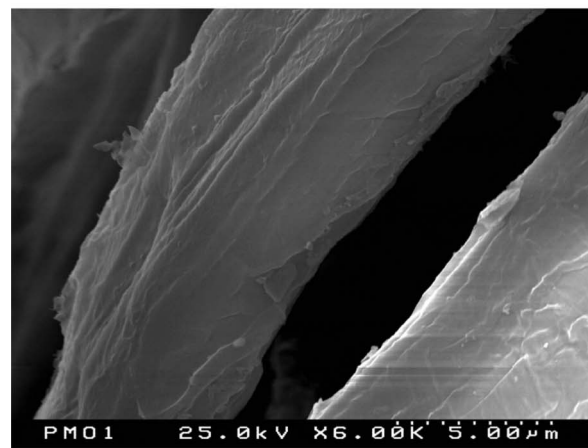


Fig. 9 Typical SEM images of cellulose/silica hybrids functionalized with POM.

polyoxoanion⁴⁹ and in the hybrids, in which the polyoxoanions are held in the silica support by electrostatic interactions with the protonated propylamino moieties, are responsible for the occurrence of the charge transfer bands at different energy. Similar behaviour occurred for PVMo_{11} immobilized on the silica surface modified with cationic quaternary ammonium moieties, $[(\text{CH}_2)_3\text{N}(\text{CH}_3)_3]^+$.⁵⁰

Wide angle X-ray diffractograms (Co- $k\alpha$ radiation) of cellulose/silica functionalized materials containing POM presented the non-resolved reflections from the 101 and $10\bar{1}$ planes around 18° and the reflection from the 002 plane at about 26° (Fig. 8). Therefore, the basic polymorph of cellulose I,^{51,52} present in the cellulose substrate (Fig. 8), was preserved in the hybrid materials. The degree of crystallinity was calculated by a procedure⁵³ based on the integrated areas of the peaks from crystalline cellulose and amorphous halo as shown in Fig. 8. The value of crystallinity for the initial cellulosic matrix was 71.5% while for the cellulose/silica materials functionalized with POM degrees of crystallinity in the range of 70–72% were obtained. Since no significant decrease of the crystallinity of cellulose occurred in the hybrids, when compared to the initial cellulosic pulp, this suggests that the synthesis procedure did not significantly solvate the cellulose crystallites.

Table 2 Mercury intrusion porosimetry results for the cellulose/silica hybrids functionalized with polyoxometalates

| Material | Porosity (%) | Skeletal density (g cm^{-3}) |
|---|--------------|---|
| Cellulosic pulp | 83 | 1.04 |
| Cell/SiO ₂ (C ₃ H ₆ NH ₃)-PV ₂ Mo ₁₀ | 78 | 1.18 |
| Cell/SiO ₂ (C ₃ H ₆ NH ₃)-PVMo ₁₁ | 76 | 1.25 |
| Cell/SiO ₂ (C ₃ H ₆ NH ₃)-PMo ₁₂ | 80 | 1.30 |

SEM analysis of the novel materials showed that functionalized silica is coating cellulosic fibres predominantly in the form of a dense film (Fig. 9). The micrographs did not reveal discrete silica domains on the fibres surface or isolated silica between the fibres. EDS analyses confirmed the presence of C, O, and Si at the fibres surface as well as minor amounts of Mo. Overall, SEM-EDS analyses revealed, in general, fairly homogeneous coverage of cellulosic fibres with silica, although some regions might be only partially coated with silica as could be observed in high resolution micrographs (Fig. 9).

Mercury intrusion porosimetry results (Table 2) show that all cellulose-based materials presented in this work have very high porosity (>75%). The decrease of porosity in the hybrids, when compared to that of the initial cellulosic pulp, can be a consequence of the filling of some fibre pores by silica and

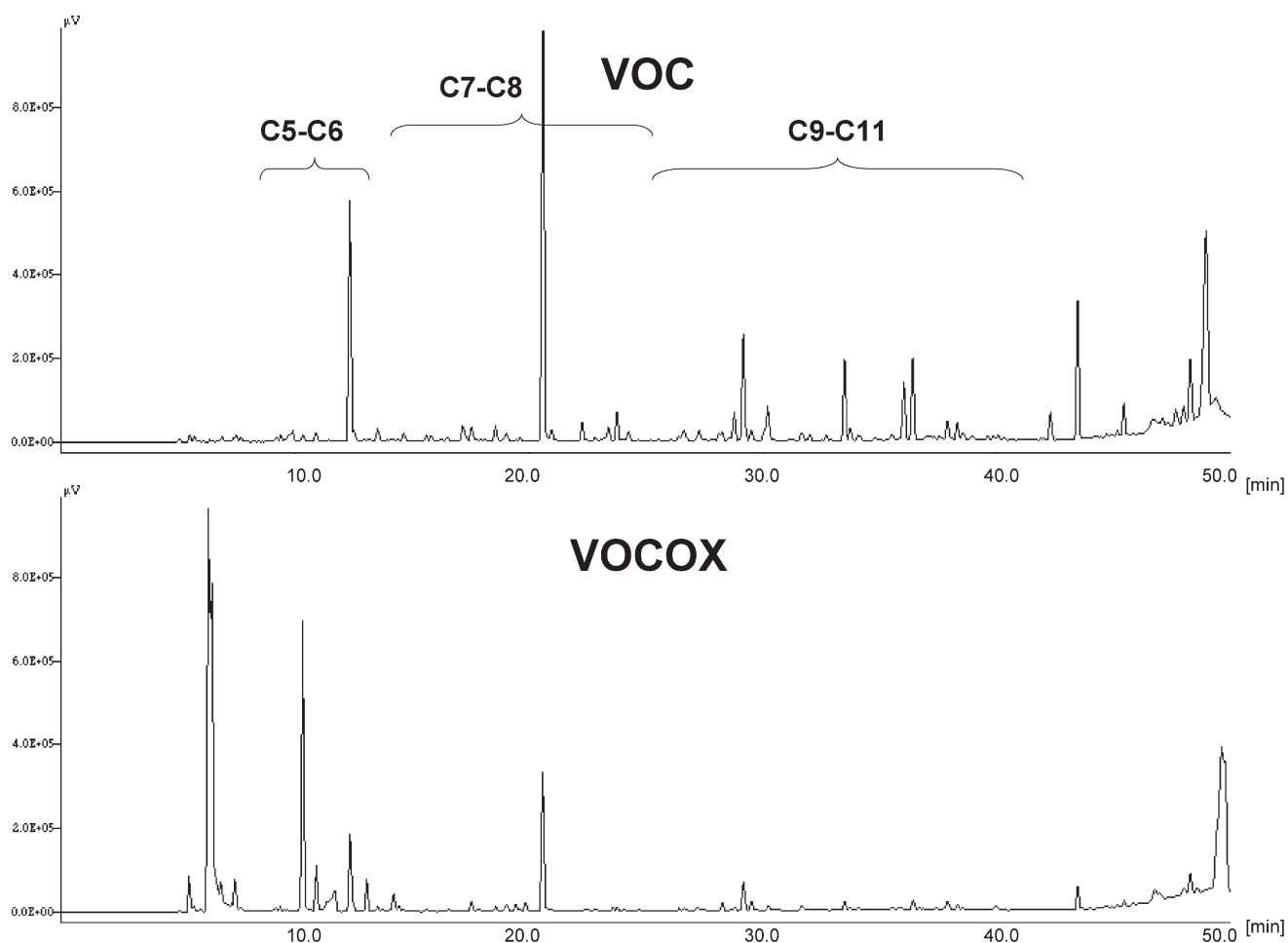


Fig. 10 Normalised GC/FID chromatograms of VOCs in polluted urban air before (VOC) and after contact with PVMo_{11} -functionalized cellulose/silica hybrid (VOCOX).

POM. The high porosity of these materials is an important characteristic when considering their possible catalytic applications for solid-gas or solid-liquid heterogeneous oxidation reactions. In addition, the apparent *skeletal* density of the materials increased from 1.0 g cm^{-3} in cellulosic pulp to about $1.2\text{--}1.3 \text{ g cm}^{-3}$ in the hybrids, mainly due to the silica coating on the fibres.

Oxidation of urban air VOCs using polyoxometalate-functionalized cellulose/silica hybrids

The POM-functionalized cellulose/silica hybrids were tested for urban air VOCs removal, at *ca.* $25 \text{ }^\circ\text{C}$, in order to evaluate the oxidative catalytic potential of these new materials in environmental applications. Polluted urban air was pumped through Teflon tubes filled with cellulose/silica hybrids without or with POMs (PVMo_{11} and $\text{PV}_2\text{Mo}_{10}$) and the treated air was analysed by GC/FID. Both catalysts showed reactivity with VOCs as revealed by change of colour of the starting hybrid materials. In fact, the yellow colour of POM-functionalized hybrids changed to green colour indicating the reduction of vanadium atoms in POMs ($\text{V}^{\text{V}} \rightarrow \text{V}^{\text{IV}}$) upon VOCs oxidation. Qualitative analysis of GC chromatograms (see example depicted in Fig. 10 for air treated with PVMo_{11} -functionalized cellulose/silica hybrid) reveals the oxidation of the majority of $\text{C}_5\text{--C}_{11}$ VOCs towards low molecular weight degradation products (peak shift to low retention times) or oxygenated products (peak shift to high retention times). It is noteworthy that both catalysts could be regenerated by passing purified air (Air K) through the tubes, as revealed by colour change of hybrid materials (from green to yellow, $\text{V}^{\text{IV}} \rightarrow \text{V}^{\text{V}}$), being, however, the re-oxidation slower for the PVMo_{11} hybrid. Hence, aforementioned POMs incorporated in cellulose/silica hybrids may be suggested as effective and regenerable catalysts for VOCs oxidation. This opens new perspectives towards functional filtering materials for indoor and outdoor air purification. Further work on mechanistic studies and oxidation of particular air pollutants with cellulose/silica hybrids functionalized with POM is in progress.

Conclusions

The preparation of cellulose/silica hybrid materials functionalized with Keggin-type polyoxometalates employing sol-gel synthesis under mild conditions has been developed. The new bio-based materials are composed by *ca.* 37 wt% of propylamine-modified silica dispersed over cellulosic fibres (*ca.* 56 wt%) and a small amount (2%) of Keggin-type polyoxoanions attached electrostatically to the protonated amino groups (Fig. 11). Taking into account the larger excess of propylamino moieties to POM clusters (*e.g.*: estimated as 40/1 for $[\text{PVMo}_{11}\text{O}_{40}]^{4-}$) only a small fraction of the protonated amino groups in modified silica should be involved in the interaction with POM cluster, with the major part being probably attracted to other counter-anions, namely nitrate anions. By analogy with other cellulose/silica hybrid materials obtained using similar procedure,¹⁸ it is presumed that the silica counterpart is bound to cellulosic fibres by hydrogen bonds between cellulose hydroxyl and siloxane groups (Fig. 11), since no covalent bonding (C–O–Si) between cellulose and silica has been detected. This covalent bonding, if it occurs, is of less importance and a subject of discussion.

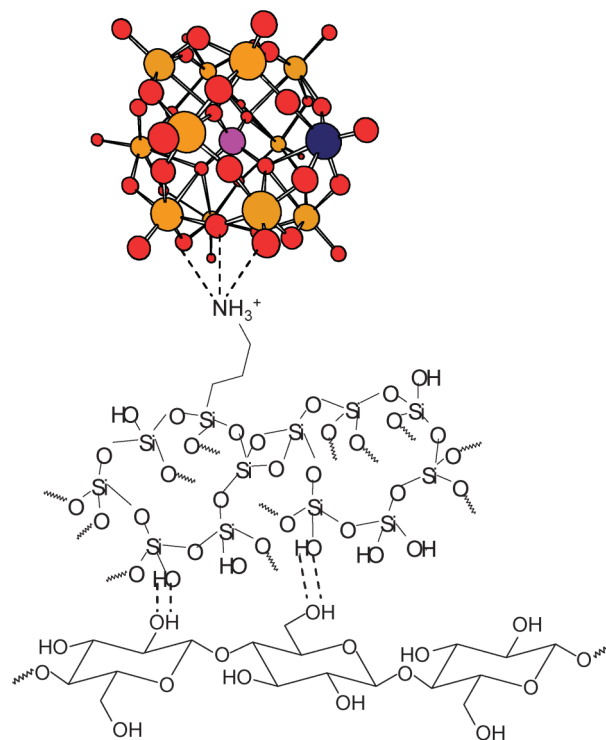


Fig. 11 Schematic representation of the main plausible chemical interactions that occur in the novel cellulose/silica hybrids functionalized with POM (silica is bound to cellulose essentially by hydrogen bonding though minor frequency covalent bonding cannot be completely excluded). In the ball and stick representation of the Keggin structure of POM the P, Mo (or W) and O atoms are depicted as pink, orange and red spheres, respectively. The substituting heteroatom (vanadium) is shown in blue.

Tshabalala and co-workers,²¹ based on X-ray photoelectron spectroscopy results, proposed the covalent bonding of cellulose with siloxane groups in wood coated with polysiloxane by sol-gel process. However, in other studies of related cellulose/silica hybrid materials, the covalent bonding between cellulose and silica was not confirmed.^{17,18,20,22,25,26}

Most of the original cellulose properties, namely its fibrous structure, crystallinity, high porosity, and low density, were preserved in the hybrid materials. Silica, that served as support for POMs, was intended to protect cellulose against “oxidative ageing” while avoiding the direct contact between POM and cellulose. However, this purpose was not completely fulfilled since some decrease in the thermal stability of cellulose was observed for the POM-functionalized hybrids (except for that containing $[\text{PW}_{12}\text{O}_{40}]^{3-}$). Keggin-type POMs used in this study, $[\text{PVMo}_{11}\text{O}_{40}]^{4-}$ and $[\text{PV}_2\text{Mo}_{10}\text{O}_{40}]^{5-}$, in the form of cellulose/silica hybrids manifested, in general, redox behaviour comparable to that normally observed in liquid phase, *i.e.*, POMs retained their catalytic properties upon immobilization. The preliminary results for VOCs oxidation in polluted urban air demonstrate the potential of these materials in gas-solid heterogeneous oxidative catalysis for eventual air purification purposes. Systematic studies on anthropogenic indoor and outdoor VOCs oxidation with the new POM-functionalized cellulose/silica hybrid materials are under way.

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References

- 1 A. P. Wight and M. E. Davis, *Chem. Rev.*, 2002, **102**, 3589.
- 2 F. Hoffmann, M. Cornelius, J. Morell and M. Fröba, *Angew. Chem. Int. Edit.*, 2006, **45**, 3216.
- 3 C. Sanchez, P. Belleville, M. Popall and L. Nicole, *Chem. Soc. Rev.*, 2011, **40**, 696.
- 4 D. Klemm, B. Heublein, H.-P. Fink and A. Bohn, *Angew. Chem. Int. Edit.*, 2005, **44**, 3358.
- 5 T. Heinze and T. Liebert, *Prog. Polym. Sci.*, 2001, **26**, 1689.
- 6 J. Kim, S. Yun and Z. Ounaies, *Macromolecules*, 2006, **39**, 4202.
- 7 M. L. Hassan, C. N. Moorefield and G. R. Newkome, *Macromol. Rapid Commun.*, 2004, **25**, 1999.
- 8 H. Tonami, H. Uyama and S. Kobayashi, *Macromolecules*, 2004, **37**, 7901.
- 9 U. Bora, K. Kannan and P. Nahar, *J. Membr. Sci.*, 2005, **250**, 215.
- 10 C. Liu, R. Sun, A. Zhang, M. Qin, J. Ren and X. Wang, *J. Agric. Food Chem.*, 2007, **55**, 2399.
- 11 S. Boufi, M. R. Vilar, V. Parra, A. M. Ferraria and A. M. B. Rego, *Langmuir*, 2008, **24**, 7309.
- 12 S. Yano, K. Iwata and K. Kurita, *Mat. Sci. Eng. C-Bio S*, 1998, **6**, 75.
- 13 W. A. Daoud, J. H. Xin and Y. Zhang, *Surf. Sci.*, 2005, **599**, 69.
- 14 S. Yano, *Polymer*, 1994, **35**, 5565.
- 15 R. A. Zoppi and M. C. Gonçalves, *J. Appl. Polym. Sci.*, 2002, **84**, 2196.
- 16 K. Tanaka and H. Kozuka, *J. Mater. Sci.*, 2005, **40**, 5199.
- 17 R. S. Gill, M. Marquez and G. Larsen, *Micropor. Mesopor. Mat.*, 2005, **85**, 129.
- 18 S. Sequeira, D. V. Evtuguin, I. Portugal and A. P. Esculcas, *Mat. Sci. Eng. C-Bio S*, 2007, **27**, 172.
- 19 S. Sequeira, D. V. Evtuguin and I. Portugal, *Polym. Compos.*, 2009, **30**, 1275.
- 20 H. S. Barud, R. M. N. Assunção, M. A. U. Martines, J. Dexpert-Ghys, R. F. C. Marques, Y. Messaddeq and S. J. L. Ribeiro, *J. Sol-Gel Sci. Technol.*, 2008, **46**, 363.
- 21 M. A. Tshabalala, P. Kingshott, M. R. Vanlandingham and D. Plackett, *J. Appl. Polym. Sci.*, 2003, **88**, 2828.
- 22 M. Abdelmouleh, S. Boufi, M. N. Belgacem, A. Dufresne and A. Gandini, *J. Appl. Polym. Sci.*, 2005, **98**, 974.
- 23 B. Ding, C. Li, Y. Hotta, J. Kim, O. Kuwaki and S. Shiratori, *Nanotechnology*, 2006, **17**, 4332.
- 24 A. Hou, Y. Shi and Y. Yu, *Carbohydr. Polym.*, 2009, **77**, 201.
- 25 X. Chen, Y. Liu, H. Lu, H. Yang, X. Zhou and J. H. Xin, *Cellulose*, 2010, **17**, 1103.
- 26 J. Alongi, M. Ciobanu and G. Malucelli, *Cellulose*, 2011, **18**, 167.
- 27 M. T. Pope and A. Muller, *Angew. Chem. Int. Ed. Engl.*, 1991, **30**, 34.
- 28 M. T. Pope and A. Muller (ed.). *Polyoxometalate Chemistry: From Topology Via Self Assembly to Applications*, Kluwer, Dordrecht, 2001.
- 29 D. L. Long, E. Burkholder and L. Cronin, *Chem. Soc. Rev.*, 2007, **36**, 105.
- 30 I. V. Kozhevnikov, *Chem. Rev.*, 1998, **98**, 171.
- 31 N. Mizuno and M. Misono, *Chem. Rev.*, 1998, **98**, 199.
- 32 A. R. Gaspar, J. A. F. Gamelas, D. V. Evtuguin and C. Pascoal Neto, *Green Chem.*, 2007, **9**, 717.
- 33 F. de C. Oliveira, J. Schneider, A. Siervo, R. Landers, A. M. G. Plepis, J. Pireaux and U. P. Rodrigues-Filho, *Surf. Interface Anal.*, 2002, **34**, 580.
- 34 B. Ding, J. Gong, J. Kim and S. Shiratori, *Nanotechnology*, 2005, **16**, 785.
- 35 G. A. Tsigdinos and C. J. Hallada, *Inorg. Chem.*, 1968, **7**, 437.
- 36 L. Pettersson, I. Andersson, A. Selling and J. H. Grate, *Inorg. Chem.*, 1994, **33**, 982.
- 37 K. Nomiya, K. Yagishita, Y. Nemoto and T. Kamataki, *J. Mol. Catal. A: Chem.*, 1997, **126**, 43.
- 38 J. A. F. Gamelas, D. V. Evtuguin and A. P. Esculcas, *Transition Met. Chem.*, 2007, **32**, 1061.
- 39 M. G. Evtugina, T. Nunes, C. Pio and C. S. Costa, *Atmos. Environ.*, 2006, **40**, 6277.
- 40 M. T. Pope, *Heteropoly and Isopoly Oxometalates*, Springer-Verlag, Berlin, 1983.
- 41 B. J. S. Johnson and A. Stein, *Inorg. Chem.*, 2001, **40**, 801.
- 42 K. Wickholm, P. T. Larsson and T. Iversen, *Carbohydr. Res.*, 1998, **312**, 123.
- 43 G. S. Caravajal, D. E. Leyden, G. R. Quinting and G. E. Maciel, *Anal. Chem.*, 1988, **60**, 1776.
- 44 R. C. Schroden, C. F. Blanford, B. J. Melde, B. J. S. Johnson and A. Stein, *Chem. Mater.*, 2001, **13**, 1074.
- 45 J. A. F. Gamelas, F. M. Santos, V. Felix, A. M. V. Cavaleiro, E. de Matos Gomes, M. Belsley and M. G. B. Drew, *Dalton Trans.*, 2006, 1197.
- 46 M. Kacurakova, A. C. Smith, M. J. Gidley and R. H. Wilson, *Carbohydr. Res.*, 2002, **337**, 1145.
- 47 H. Haykiri-Acma, S. Yaman and S. Kucukbayrak, *Fuel Process. Technol.*, 2010, **91**, 759.
- 48 J. A. F. Gamelas, A. M. V. Cavaleiro, E. de Matos Gomes, M. Belsley and E. Herdtweck, *Polyhedron*, 2002, **21**, 2537.
- 49 G. M. Brown, M. R. Noe-Spirlet, W. R. Busing and H. A. Levy, *Acta Crystallogr., Sect. B: Struct. Crystallogr. Cryst. Chem.*, 1977, **33**, 1038.
- 50 C. N. Kato, A. Tanabe, S. Negishi, K. Goto and K. Nomiya, *Chem. Lett.*, 2005, **34**, 238.
- 51 A. C. O' Sullivan, *Cellulose*, 1997, **4**, 173.
- 52 D. Fengel and G. Wegener, *Wood Chemistry, Ultrastructure, Reactions*, Walter de Gruyter, Berlin, 1984.
- 53 P. Ricou, E. Pinel and N. Juhasz, *JCPDS—International Centre for Diffraction Data, Advances in X-ray Analysis*, 2005, **48**, 170.