



C-terminal Cysteines of CueR Act as Auxiliary Metal Site Ligands upon HgII Binding—A Mechanism To Prevent Transcriptional Activation by Divalent Metal Ions?

Balogh, Ria K.; Gyurcsik, Béla; Hunyadi-Gulyás, Éva; Schell, Juliana; Thulstrup, Peter W.; Jancsó, Attila; Hemmingsen, Lars Bo Stegeager

Published in:
Chemistry - A European Journal

DOI:
[10.1002/chem.201902940](https://doi.org/10.1002/chem.201902940)

Publication date:
2019

Document version
Publisher's PDF, also known as Version of record

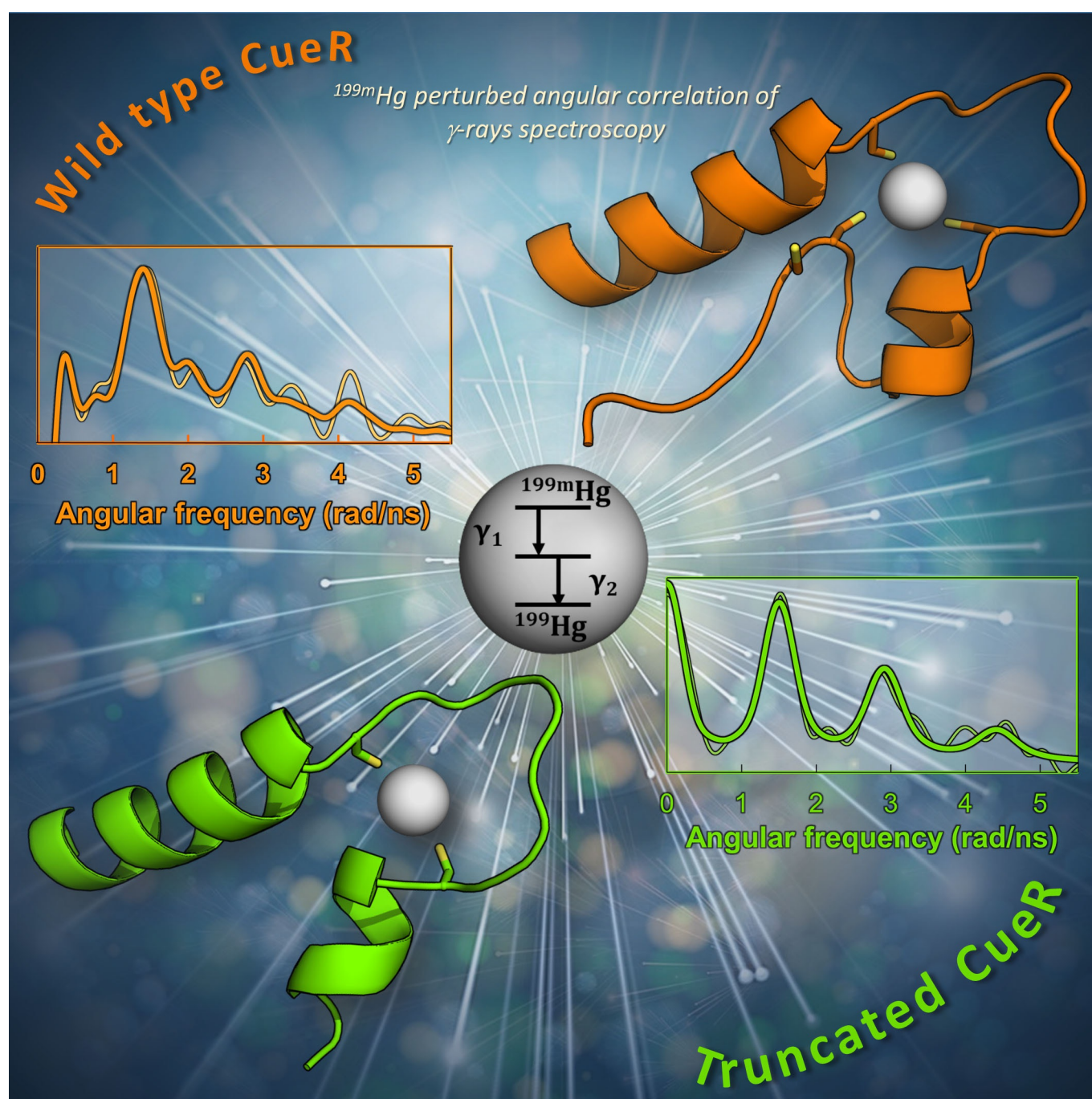
Document license:
[CC BY](#)

Citation for published version (APA):
Balogh, R. K., Gyurcsik, B., Hunyadi-Gulyás, É., Schell, J., Thulstrup, P. W., Jancsó, A., & Hemmingsen, L. B. S. (2019). C-terminal Cysteines of CueR Act as Auxiliary Metal Site Ligands upon Hg^{II} Binding—A Mechanism To Prevent Transcriptional Activation by Divalent Metal Ions? *Chemistry - A European Journal*, 25(66), 15030–15035. <https://doi.org/10.1002/chem.201902940>

Bioinorganic Chemistry

C-terminal Cysteines of CueR Act as Auxiliary Metal Site Ligands upon Hg^{II} Binding—A Mechanism To Prevent Transcriptional Activation by Divalent Metal Ions?

Ria K. Balogh,^[a] Béla Gyurcsik,^[a] Éva Hunyadi-Gulyás,^[b] Juliana Schell,^[c] Peter W. Thulstrup,^[d] Lars Hemmingsen,^{*,[d]} and Attila Jancsó^{*,[a]}



affinity.^[11] However, according to model peptide studies^[12,13] and QM/MM calculations,^[14] Hg^{II} ions may be coordinated even more efficiently. Moreover, Hg^{II} is also able to bind to a CC sequence,^[15] and therefore coordination of Hg^{II} ion by the CCHH motif is also highly probable.

With the present work we aim to explore the role of the C-terminal CCHH motif with a particular focus on the binding of Hg^{II} to CueR. To achieve this, we studied the Hg^{II}-interaction of *E. coli* CueR and its truncated variant, lacking seven C-terminal residues (including the CCHH motif), Δ C7-CueR. The integrity of this variant was confirmed by CD spectroscopy and electrophoretic mobility shift assay, see Figure S3.

A series of ESI-MS spectra were recorded with the two protein variants, see Figures 2, S4 and S5. The disappearance of

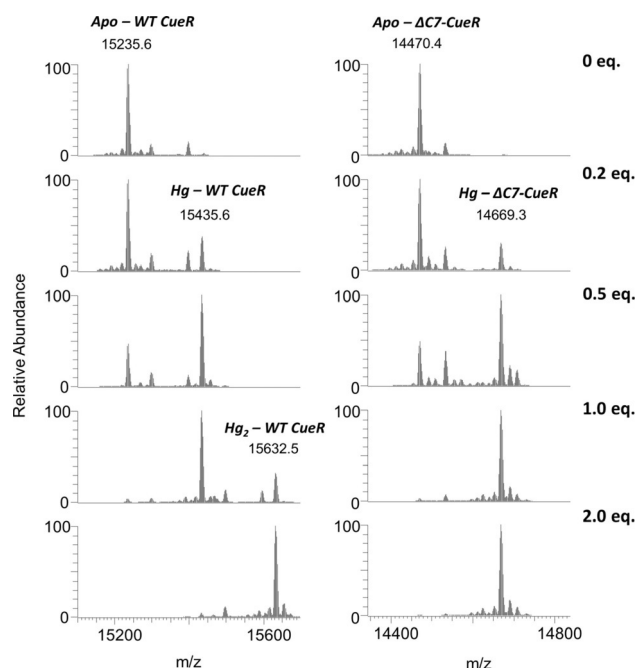


Figure 2. Deconvoluted native ESI-MS spectra of the WT and truncated CueR in the absence and presence of 0.2, 0.5, 1.0 and 2.0 equivalents of Hg^{II} ions. Individual samples contained 20 μ M protein in a 10 mM NH₄HCO₃ buffer, 0.5 mM TCEP, pH 7.5.

the signals of the apo-form in the presence of 1.0 equivalent of Hg^{II} implies that Hg^{II} ions display high affinity to both proteins. The spectra obtained at twofold Hg^{II}-excess per protein clearly demonstrate the availability of two binding sites for Hg^{II} ions in the Wild-type (WT) CueR. These are most likely the metal ion binding loop formed by C112 and C120, and the C-terminal CCHH motif. Participation of the latter CCHH sequence motif in Hg^{II} binding is supported by the lack of signals corresponding to a Hg₂- Δ C7-CueR complex, even at twofold Hg^{II}-excess over the truncated protein. Both the Hg-CueR and Hg₂-CueR species are observed at 1.0 equivalent Hg^{II}, suggesting that there is no significant difference in the Hg^{II}-binding affinities of the two sites.

¹⁹⁹Hg-perturbed angular correlation (PAC) spectroscopy^[12,13,16–20] was used to elucidate the metal site structures and

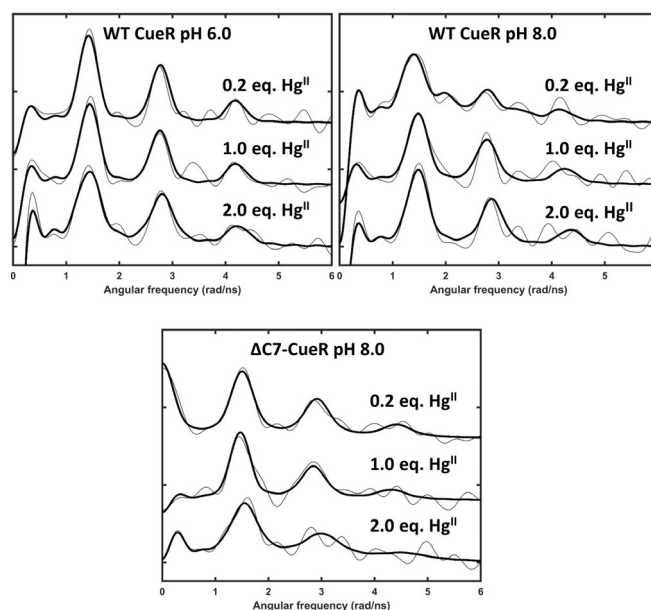


Figure 3. Experimental (grey) and fitted (black) ¹⁹⁹Hg PAC spectra of WT and truncated CueR in the presence of DNA with 0.2, 1.0 and 2.0 equivalents of Hg^{II}. Top left: WT at pH 6.0; top right: WT at pH 8.0 $c_{\text{WT CueR}} = 12 \mu\text{M}$, 0.5 equiv. DNA, and bottom: Δ C7-CueR at pH 8.0 $c_{\Delta\text{C7-CueR}} = 8.4 \mu\text{M}$, 0.5 equiv. DNA.

dynamics at the nanosecond timescale, see Figure 3 and Supporting Information Figure S6. At pH 6.0 and Hg^{II}:CueR of 0.2 and 1.0, the signals agree well with a HgS₂ coordination geometry, that is, coordination of Hg^{II} by two cysteinates.^[18] This is also the case at Hg^{II}:CueR of 2.0, although a slightly larger line-width is observed, in particular for the first peak at around 1.4 rad ns⁻¹. This line broadening presumably reflects the occupation of two HgS₂ sites, and it can originate either from minor differences in structure of the two sites, or from metal site dynamics at the nanosecond time scale becoming more pronounced upon binding of the second Hg^{II} (Figure 3).

The spectrum recorded with 0.2 equivalent of Hg^{II} per CueR at pH 8.0 is more complex than at pH 6.0. Qualitatively, the first peak is shifted to slightly lower frequency and exhibits considerable broadening, and the second peak (ca. 2.8 rad ns⁻¹) is significantly attenuated, to the extent that it barely rises above the noise level. A reliable analysis of the data requires the inclusion of two nuclear quadrupole interactions (NQIs). One of these NQIs is very similar to that observed in the spectra at pH 6.0, most likely reflecting a HgS₂ structure. The other NQI has a higher asymmetry parameter and a lower frequency, see Table S1, indicating a higher coordination number than 2. The lower frequency agrees well with an ideal trigonal planar HgS₃ structure, but the relatively high asymmetry parameter rules out this possibility. However, in the simple angular overlap model (AOM),^[21] a T-shaped HgS₃ coordination geometry gives the same frequency as a trigonal planar structure, but an asymmetry parameter of 1. Thus, a HgS₃ structure in between trigonal planar and T-shaped, with the third ligand in a slightly longer Hg–S distance seems to be a plausible structural interpretation of the low frequency signal. It is also

possible that the PAC data reflect a trigonal planar HgS_2N structure, with a histidine coordinating, as this would give an asymmetry parameter different from zero. However, this seems less likely, given the thiophilicity of Hg^{II} , and the UV absorption data, *vide infra*. Finally, it is conceivable that the spectrum reflects intermediate (nanosecond) exchange between HgS_2 and HgS_3 structures. Notice that this entails a flip of principal axis of the electric field gradient tensor, which has V_{zz} along the axis of HgS_2 but perpendicular to the HgS_3 plane, and therefore the asymmetry parameter will depend on the dynamics in a non-trivial manner. It cannot be excluded that the data recorded at 1.0 equivalent of Hg^{II} also contain signals reflecting both of these species, but the reduced chi-square does not improve significantly upon including a second NQL. Consequently, we have only included the high frequency NQL (HgS_2) in the analysis. For the experiment with 2.0 equivalents of Hg^{II} the signal may be satisfactorily fitted with just one (high frequency) NQL, presumably reflecting HgS_2 structure for both Hg^{II} bound to CueR (Figure 3).

Most interestingly, the $^{199\text{m}}\text{Hg}$ PAC spectrum recorded at pH 8.0 with 0.2 equivalents Hg^{II} for $\Delta\text{C7-CueR}$ exhibits a signal reflecting only HgS_2 structure (Figure 3). The fact that the $\Delta\text{C7-CueR}$ Hg^{II} site exhibits a HgS_2 structure strongly supports the interpretation presented above for the WT CueR: if HgS_3 is formed by occupation of the functional site, a third thiolate is recruited from the CCHH motif, or vice versa, Hg^{II} binds to the CCHH motif and recruits one of the cysteines from the functional binding site. With 2.0 equivalents of Hg^{II} per $\Delta\text{C7-CueR}$ at pH 8.0, the signal changes as compared to experiments with ≤ 1 equivalent Hg^{II} , presumably because the functional metal site is filled, and the additional Hg^{II} accommodates a coordination geometry other than linear HgS_2 due to weak or non-specific Hg^{II} adducts. This agrees well with the ESI-MS data, where no $\text{Hg}_2\text{-}\Delta\text{C7-CueR}$ was observed. Thus it is likely that the signal includes more than one NQL. Surprisingly, the signal shifts to slightly higher frequency, which is difficult to account for, except if a positive charge appears in the equatorial plane of HgS_2 , *vide infra*.

To further characterize the metal site coordination geometries, we applied UV absorption spectroscopy (Figure 4). Hg^{II} -thiolate complexes possess characteristic charge transfer (CT) bands in the region of 230–300 nm. Moreover, features of the absorption spectrum reflect the coordination geometry of the complexes. Using $\text{Hg}(\text{SET})_2$ and $[\text{Et}_4\text{N}][\text{Hg}(\text{SBut})_3]$ model compounds, the UV-absorption spectra of linearly and trigonal planar coordinated Hg^{II} , respectively, were characterized.^[24] Linearly coordinated Hg^{II} -thiolate species display a transition at around 230 nm.^[22] The increase of the coordination number shifts the absorption bands towards longer wavelengths.^[23,25] The spectrum of a trigonal Hg^{II} -thiolate complex has a characteristic absorption maximum at 245 nm with a distinct shoulder at around 290 nm.^[22] Qualitatively, the absorption difference spectra at sub-equimolar $\text{Hg}^{\text{II}}:\text{WT CueR}$ ratios exhibit a characteristic absorption at around 290 nm reflecting the presence of HgS_3 structure (Figure 4), in agreement with the PAC data, *vide supra*. The PAC data indicate 40% HgS_3 and 60% HgS_2 at 0.2 equivalents Hg^{II} . We used the recorded spectrum with

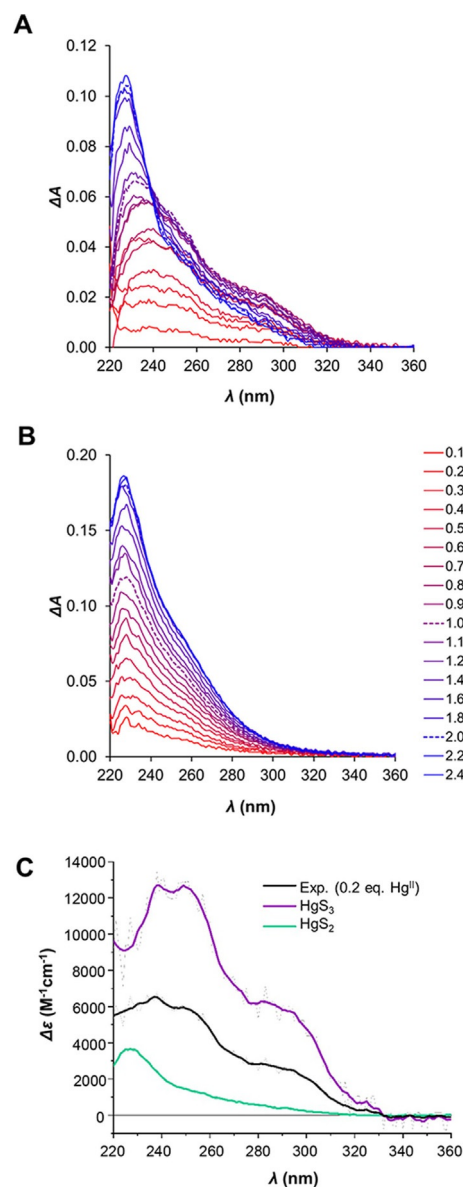


Figure 4. UV absorption difference spectra of WT CueR (A) and $\Delta\text{C7-CueR}$ (B) titrated with Hg^{II} ions (0.1–2.4 equivalents). Spectra recorded in the presence of 1.0 and 2.0 equivalents of Hg^{II} are shown with dashed lines. pH 7.5, $c_{\text{WT CueR}} = 14 \mu\text{M}$, $c_{\Delta\text{C7-CueR}} = 12 \mu\text{M}$. (C) Estimated molar absorbance for the HgS_2 and HgS_3 species derived from the WT CueR UV absorption spectrum recorded with 2.0 equiv. Hg^{II} and 0.2 equiv. Hg^{II} combined with the relative population of the two species derived from $^{199\text{m}}\text{Hg}$ PAC data, see the text for details.

2.0 equivalents Hg^{II} per WT CueR (Figure 4A) to determine the molar absorption of the HgS_2 species (green curve in Figure 4C). Next, we predicted the pure HgS_3 molar absorption spectrum (Figure 4C, purple curve) by assuming that the experimentally determined spectrum is given by 0.6 HgS_2 + 0.4 HgS_3 . The UV absorption spectra derived in this manner for HgS_2 and HgS_3 agree well with those reported in the literature,^[23] strongly supporting the interpretation of the PAC data presented above. We present molar absorption data at selected wavelength values in Table 1. The UV absorption spectra recorded for $\Delta\text{C7-CueR}$ exclusively exhibit the signature of HgS_2

Sample	λ [nm]	ε [cm ⁻¹ mM ⁻¹]
Hg-MerR ^[22]	242	19.8
	260	14.6
	290	6.45
1:2 Hg/L16C ^[23]	240	2.70
1:3 Hg/L16C ^[23]	247	16.8
	265	10.6
	295	5.00
CueR (HgS ₃)	240	12.6
	250	12.6
	290	5.82
CueR (HgS ₂)	230	3.52

structure, corroborating the interpretation of other experimental data. Surprisingly, the absorbance for ΔC7-CueR continues to increase beyond 1.0 equivalent Hg^{II} and saturates at ca. 2:1 Hg^{II}:ΔC7-CueR, indicating that the truncated protein can accommodate two Hg^{II} ions in a HgS₂ coordination environment. This may be realized if a dinuclear Hg₂S₂ site is formed with the two thiolates as bridging ligands. Interestingly, this agrees with the unexpectedly high frequency observed by PAC spectroscopy, which can be explained by the presence of a positive charge (the second Hg^{II}) in the Hg₂S₂ structure, *vide supra*. The fact that the species with two Hg^{II} bound per CueR monomer is not observed in ESI-MS implies that the binding of the second Hg^{II} is relatively weak.

In Figure 5, we present model structures which agree with all the experimental data presented in this work. At pH 8.0

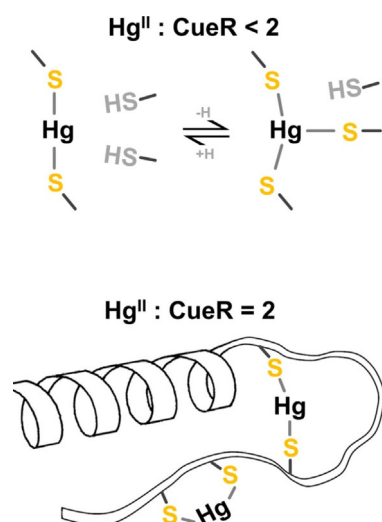


Figure 5. Model structures of Hg^{II} bound to WT CueR at pH 8.0. Binding of Hg^{II} to CueR gives rise to an equilibrium between HgS₂ and HgS₃ when Hg^{II}:CueR < 2, and to pure HgS₂ coordination upon addition of 2 Hg^{II} ions per protein monomer. This can only be accounted for if the CCHH C-terminal motif participates in the coordination of Hg^{II}, see the text for details.

with 0.2 and 1.0 equivalent Hg^{II}, two species co-exist, most likely the linear HgS₂ and a HgS₃ structure with the equatorial Hg-S bond being longer than the other two. Such structures have also been observed in small, Hg^{II} containing inorganic compounds.^[26] The increased availability of deprotonated cysteines with increasing pH agrees well with this change in speciation observed from pH 6.0 to pH 8.0, that is, a change from HgS₂ towards HgS₃ coordination mode, and a similar trend has been observed for *de novo* designed proteins by Irazzo et al.^[18] The additional thiolate is most likely recruited from the CCHH motif, or vice versa, and may thus prevent the docking of the C-terminal helix into the hydrophobic pocket, and consequently inhibit activation of transcription. The net negative charge of HgS₃ may be stabilized due to the presence of lysine or arginine in the C-terminal fragment of CueR in almost all the organisms listed in Figure 1. That is, we hypothesize that the CCHH motif is not involved in the function of CueR when sensing the monovalent coinage metals, but it does take part in binding of divalent metal ions, a mechanism that would account for the selectivity of CueR.

It may seem intriguing that with 1.0 equivalent Hg^{II} both the PAC and UV absorption spectra differ significantly from those recorded with 0.2 equivalent Hg^{II}. However, a simple probabilistic model qualitatively accounts for this change, assuming that the two sites are independent (i.e. distributing Hg^{II} randomly among the 4 metal sites of a protein dimer), and that population of two adjacent sites (the functional site and the C-terminal site) leads to formation of HgS₂, because there are no more cysteines locally available to form HgS₃, see the Supporting Information for details. This very simple interpretation is to some extent supported by the ESI-MS data, which display population of the Hg₂-CueR species when Hg^{II} and CueR are present in equimolar amounts. Obviously, the model is too simple because formation of HgS₃ requires that cysteines from both metal binding sites are involved, but the alternative, that is, that one binding site (either the functional site or the C-terminal site) binds Hg^{II} with significantly higher affinity than the other, does not agree with the spectroscopic data, because this would imply that the HgS₂/HgS₃ ratio should be the same at 0.2 and at 1.0 equivalent, nor with the ESI-MS data, which indicate the presence of Hg₂-CueR already at 1.0 equivalent Hg^{II}. At 2.0 equivalents Hg^{II}, of course, there is no more possibility to form HgS₃, because the protein is saturated with Hg^{II} in HgS₂ structures. Similar geometrical rearrangement was observed in metallothioneins (by UV absorption) upon saturating the protein by the metal ion in a titration with Hg^{II}.^[27,28] The function of the CCHH motif has also been studied by Stoyanov and Brown, using an *in vivo* assay to monitor the CueR controlled transcription.^[4] The double mutation of histidine (H131N/H132N) or cysteine residues (C129S/C130S) and truncation from G128 in *E. coli* CueR resulted in an only slightly altered induction of the transcription by cognate metal ions. Although experimental data were not presented, Stoyanov and Brown indicated that the selectivity of reaction with other, unspecified metal ions was not affected. To further explore this issue, a series of *in vitro* and *in vivo* transcriptional assays should be conducted.

In summary, we have demonstrated that up to two Hg^{II} ions bind with high affinity to WT CueR, one at the functional (C112 and C120) metal binding site, and the other at the C-terminal CCHH motif. Moreover, under conditions where the protein is not saturated by Hg^{II}, a higher coordination number (presumably Hg₃) is observed for WT CueR but not for ΔC7-CueR, indicating that side chains from the CCHH motif may be recruited as auxiliary ligands at the functional metal site (or vice versa). This implies a mechanism where the specificity of CueR for monovalent coinage metal ions and against divalent metal ions is achieved by coordination to divalent metal ions by the CCHH motif, preventing the docking of the C-terminal helix into the hydrophobic pocket,^[5] and consequently inhibiting activation of transcription. Indeed, the CCHH motif provides a selection of ligands that may participate in coordination of both soft and intermediate metal ions. As the findings presented here on Hg^{II} do represent a special case, the generalization to other divalent metal ions should be considered carefully.

Acknowledgements

We acknowledge the financial support received from the Federal Ministry of Education and Research (BMBF) through grant 05K16PGA and from the European Union's Horizon 2020 Framework research and innovation program under grant agreement no. 654002 (ENSAR2). We further thank J.G. Correia (C2TN-DECN-IST-UL) and project CERN-FIS-PAR-0005–2017 FCT-Portugal for technical assistance during the beam time and Jens-Christian Navarro Poulsen and Morten J. Bjerrum (University of Copenhagen) for their support during expression and purification of the ΔC7-CueR variant. We thank ISOLDE/CERN for beam time, EURONS and NICE for financial support. Financial support from the Hungarian National Research, Development and Innovation Office (GINOP-2.3.2-15-2016-00038, GINOP-2.3.2-15-2016-0001 and K 16/120130) is also acknowledged.

Conflict of interest

The authors declare no conflict of interest.

Keywords: coordination modes · CueR metalloregulatory protein · mercury · metal ion selectivity · perturbed angular correlation (PAC) spectroscopy

[1] J. V. Stoyanov, J. L. Hobman, N. L. Brown, *Mol. Microbiol.* **2001**, *39*, 502–512.

- [2] A. Changela, K. Chen, Y. Xue, J. Holschen, C. E. Outten, T. V. Halloran, A. Mondragón, *Science* **2003**, *301*, 1383.
- [3] K. Chen, S. Yuldasheva, J. E. Penner-Hahn, T. V. O'Halloran, *J. Am. Chem. Soc.* **2003**, *125*, 12088–12089.
- [4] J. V. Stoyanov, N. L. Brown, *J. Biol. Chem.* **2003**, *278*, 1407–1410.
- [5] S. J. Phillips, M. Canalizo-Hernandez, I. Yildirim, G. C. Schatz, A. Mondragón, T. V. O'Halloran, *Science* **2015**, *349*, 877–881.
- [6] M. V. Golyanskiy, W. A. Gunderson, M. P. Hendrich, S. M. Cohen, *Biochemistry* **2006**, *45*, 15359–15372.
- [7] J. S. Cavet, W. Meng, M. A. Pennella, R. J. Appelhoff, D. P. Giedroc, N. J. Robinson, *J. Biol. Chem.* **2002**, *277*, 38441–38448.
- [8] Z. Ma, D. M. Cowart, R. A. Scott, D. P. Giedroc, *Biochemistry* **2009**, *48*, 3325–3334.
- [9] C. M. Phillips, E. R. Schreiter, Y. Guo, S. C. Wang, D. B. Zamble, C. L. Drennan, *Biochemistry* **2008**, *47*, 1938–1946.
- [10] D. C. Bebout, *Encyclopedia of Inorganic and Bioinorganic Chemistry* **2011**.
- [11] E. Mesterházy, B. Boff, C. Lebrun, P. Delangle, A. Jancsó, *Inorg. Chim. Acta* **2018**, *472*, 192–198.
- [12] D. Szunyogh, H. Szokolai, P. W. Thulstrup, F. H. Larsen, B. Gyurcsik, N. J. Christensen, M. Stachura, L. Hemmingsen, A. Jancsó, *Angew. Chem. Int. Ed.* **2015**, *54*, 15756–15761; *Angew. Chem.* **2015**, *127*, 15982–15987.
- [13] D. Szunyogh, B. Gyurcsik, F. H. Larsen, M. Stachura, P. W. Thulstrup, L. Hemmingsen, A. Jancsó, *Dalton Trans.* **2015**, *44*, 12576–12588.
- [14] L. Rao, Q. Cui, X. Xu, *J. Am. Chem. Soc.* **2010**, *132*, 18092–18102.
- [15] T. M. DeSilva, G. Veglia, F. Porcelli, A. M. Prantner, S. J. Opella, *Biopolymers* **2002**, *64*, 189–197.
- [16] S. Chakraborty, S. Pallada, J. T. Pedersen, A. Jancsó, J. G. Correia, L. Hemmingsen, *Acc. Chem. Res.* **2017**, *50*, 2225–2232.
- [17] A. Jancsó, J. G. Correia, A. Gottberg, J. Schell, M. Stachura, D. Szunyogh, S. Pallada, D. C. Lupascu, M. Kowalska, L. Hemmingsen, *J. Phys. G* **2017**, *44*, 064003.
- [18] O. Iranzo, P. W. Thulstrup, S.-b. Ryu, L. Hemmingsen, V. L. Pecoraro, *Chem. Eur. J.* **2007**, *13*, 9178–9190.
- [19] L. Hemmingsen, K. N. Sas, E. Danielsen, *Chem. Rev.* **2004**, *104*, 4027–4062.
- [20] T. Butz, S. Saibene, T. Fraenzke, M. Weber, *Nuclear Instruments and Methods in Physics Research Section A: Accelerators, Spectrometers, Detectors and Associated Equipment* **1989**, *284*, 417–421.
- [21] R. Bauer, S. J. Jensen, B. Schmidt-Nielsen, *Hyperfine Interact.* **1988**, *39*, 203–234.
- [22] S. P. Watton, J. G. Wright, F. M. MacDonnell, J. W. Bryson, M. Sabat, T. V. O'Halloran, *J. Am. Chem. Soc.* **1990**, *112*, 2824–2826.
- [23] G. R. Dieckmann, D. K. McRorie, D. L. Tierney, L. M. Utschig, C. P. Singer, T. V. O'Halloran, J. E. Penner-Hahn, W. F. DeGrado, V. L. Pecoraro, *J. Am. Chem. Soc.* **1997**, *119*, 6195–6196.
- [24] J. G. Wright, M. J. Natan, F. M. MacDonnell, D. M. Ralston, T. V. O'Halloran, *Progress in Inorg. Chem.* **1990**, *38*, 323–412.
- [25] F. Jalilvand, B. O. Leung, M. Izadifard, E. Damian, *Inorg. Chem.* **2006**, *45*, 66–73.
- [26] A. Manceau, K. L. Nagy, *Dalton Trans.* **2008**, 1421–1425.
- [27] E. Freisinger, *Inorg. Chim. Acta* **2007**, *360*, 369–380.
- [28] O. Schicht, E. Freisinger, *Inorg. Chim. Acta* **2009**, *362*, 714–724.

Manuscript received: June 26, 2019

Accepted manuscript online: July 31, 2019

Version of record online: October 15, 2019