



Optimization of Process Parameters by Response Surface Methodology for Methylene Blue Removal Using Cellulose Dusts

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Abstract

This study was aimed to use Cellulose dusts (CD) produced in drying section of paper mills of paper making industry as a potential adsorbent to remove methylene blue (MB) dye from aqueous solution. The adsorbent was characterized by scanning electron microscopy and Fourier transform infrared spectrometer and X-ray Diffraction. The influences of the effective parameters including pH solution, adsorbent dosage, initial MB concentration, and contact time were optimized by CCD which stands for central composite design. The influence of these parameters on the adsorption capacity was analyzed using the batch process. The accuracy of the equation that is produced by CCD was affirmed by the variance analysis and also by calculating the correlation coefficient that connects the predicted and the empirical values of the percentage of removed MB dye. Maximum removal percentage of MB dye (98.05 %) which obtained at pH 9.84, adsorbent dosage 4.38 g L⁻¹, MB concentration 75.50 g L⁻¹ and time 208.13 min. Freundlich, Temkin, Harkins-Jura and Langmuir isotherms are used to analyze the empirical data. Results revealed that the data is in a satisfying agreement with the Freundlich isotherm (R²=0.99). Pseudo-first order, Pseudo-second-order, Elovich and Intraparticle diffusion models were used to fit the kinetic data and it is found out that MB dye's adsorption onto CD has a good agreement with the pseudo-second-order kinetic model. The results showed that CD can be an efficient and low-cost adsorbent for methylene blue adsorption.

Keywords: Cellulose Dusts; Adsorption; Optimization; Isotherm; Kinetics.

1. Introduction

Industrial activities generate large volume of wastewater containing hazardous inorganic (metals) or organic (microbes, dyes/pigments) species. Dye contaminated wastewater is one of the visual significant pollution generated from plastic, cosmetic, paper and pulp, textile, distilleries and tanneries industries [1].

Dyes usually have complex aromatic molecular structures which make them more stable and difficult to biodegrade. However, their disposal in water bodies exhibit adverse effect to the aquatic and human life by creating eutrophication, mutagenic, carcinogenic effects, dysfunction of the organs, also having effect on aquatic ecosystem by preventing sunlight from reaching into the stream [2, 3].

Azo dyes are the biggest type among several other chromogenic materials that are released by many industries [4, 5]. They are more than 70% of all commercial dyes which makes them to be undisputed major synthetic dye. They are characterized by one or more azo groups (N=N) as chromophore, with aromatic containing groups and other functional groups such as -OH and -SO₃H [6, 7]. Based on the chromophore group, 20–30 different groups of dyes are pollutant [8]. Methylene blue (MB) is one of major basic dyes, which is used for coloring purposes very much. This dye can cause permanent injury to the human eye, irritation to the gastrointestinal tract, and skin irritation. MB is a common pollutant

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material in textile wastewater and is frequently used in dye houses and textile industries. The MB's removal from aqueous solutions by applying substances that cost low (e.g. agricultural wastes) has been researched by others [9-11].

This dye as a thiazine cationic dye is selected as a model organic (heterocyclic aromatic) compound because it is known to adsorb more easily onto solids and in order to evaluate the behavior of sorbent for the removal of organic pollutants from its aqueous solutions [12].

There are numerous techniques available for dye removal from wastewater like membrane separation, [13, 14] degradation by micro-organisms and photochemical, [15] chemical oxidation, [16] anion exchange resins, [17] coagulation, [18] flocculation and adsorption [19].

Adsorption on the surface of the solids is getting increasing attraction in the field among other approaches since its price is lower and the activated adsorption significantly effects the dyes and pigments' removal and also other organic and inorganic pollution [20].

Processes involving adsorption are applied in different domains including medicine, environment, oil and gas, etc. for removing of pesticides, storage, metals, dyes, and oil spillage control [21]. For the removal of phenols, dyes, pesticides or other organic substances, the majority of systems apply carbon as an adsorbent because it has a remarkable ability in adsorption. Nonetheless, its common apply is limited because of its cost and its problem with regeneration [22].

Lately, various agricultural waste and residues are analyzed with the purpose of using them as adsorbent for removing dyes from waste waters due to fact that they cost low, their toxicity is low and they are easily available [23].

Cellulose dusts (CD) produced in drying section of paper mills has no economic use and value, and they are removed from the production cycle. Thus, these wastes can be applied as a low-cost and plentiful adsorbent for removing dyes from waste water.

Response surface methodology (RSM) is a sequential procedure and collection of mathematical and statistical techniques which is useful for mathematical modeling and examination of problems. During modeling, the first-order model is appropriate when we are at a point on the response surface that is distant from the optimum operating conditions and a little curvature in the system is seen [24, 25]. It leads the experiments quickly and efficiently along a path of enhancement toward the general vicinity of the optimum. After locating the region of the optimum a more elaborate second-order model, may be used to analyze and locate the optimum. Among the RSM designs, central composite design (CCD) is the most popular method due to its simple structure and good efficiency [26].

On the basis of primary data obtained experimentally, all the affecting variables including the dose of adsorbent, contact time, pH, and initial concentration of MB dye collectively have been optimized by statistical experimental design; RSM.

The objective of the present study was to investigate the optimum condition for maximum adsorption of MB dye on Cellulose dusts. The CCD model was used for precisely finding the role of individual process parameters and also optimization of process variables. Moreover, Cellulose dusts was characterized by using FTIR, SEM and XRD.

In this study, we studied the CD produced in drying section of paper mills as a potential adsorbent for the removal of dyes from wastewater. The objective of the present study was to investigate the optimum condition for maximum adsorption of MB dye on cellulose dusts. The research is classified to some section viz. (1) preparation of cellulose dust and defining its properties by XRD, FTIR and SEM; (2) optimization by CCD combined with RSM; (3) to specify how four independent parameters affect (Parameters including time, the dose of adsorbent, initial concentration of MB, initial pH) and also to understand their interactions on removing MB (%); and (4) to research about adsorption isotherms and kinetics model.

2. Experimental and Methods

2.1. Materials and Chemicals

Cellulose dusts was prepared from Mazandaran Wood and Paper Industries (MWPI) in Iran, Co., Ltd (Figure 1). Distilled water was used to wash it a few times and therefore purify it. It was then dried up to the point that there was no difference in its weight. Then it was applied as raw material.



Figure 1. Photograph of cellulose dusts powder

Methylene blue (MB) (its chemical formula is: $C_{16}H_{18}ClN_3S \cdot 3H_2O$ and its molecular weight is equal to $319.86 \text{ g mol}^{-1}$; the analysis at wavelength 664 nm ; Figure 2 shows the chemical structure), NaOH (with 99% purity) and HCl (with 37% purity) were purchased from Merck (Germany). All solutions were prepared using distilled water. All chemicals were reagent grade and were used without further treatment.

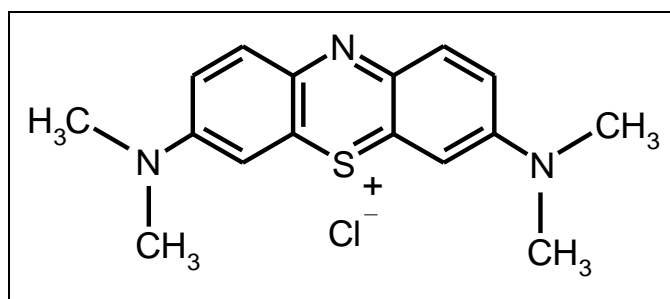


Figure 2. Chemical structure of methylene blue

2.2. Batch Adsorption Studies

MB colored solution was prepared and diluted with the desired concentrations ($20\text{-}100 \text{ mg L}^{-1}$). Experiments were conducted to investigate the effect of pH ($4\text{-}12$), adsorbent dose ($1\text{-}5 \text{ g L}^{-1}$), time ($30\text{-}270 \text{ min}$) and constant temperature on methylene blue adsorption by CD.

To measure moisture content, 1 g of this CD dried in an oven for 24 hours at a temperature of $103 \text{ }^\circ\text{C}$ and then the rest of the paper wastes stored in the bottle to use. In each experiment, a certain amount of adsorbent was added to a certain amount of dye solution and initial pH values were modified before by either HCl or NaOH, at 0.1 N .

The solution was mixed at different times and room temperature by orbital shaker (200 rpm). The solid phase was then separated by centrifugation (6000 rpm) through the liquid. Then, the dye concentration was measured at 664 nm maximum methylene blue wavelength using a UV-visible spectrophotometer (Cary- Eclipse).

The adsorption isotherms provide information about the distribution of dye molecules within the adsorbent at equilibrium. Generally, adsorption isotherm researches are done with the purpose of correlating the capacity of adsorption and the residual concentration of the adsorbate which exists in the aqueous solution. There are various isotherm equations that are accessible from the literature, four of them were chosen for the research: Harkins-Jura, Temkin, Langmuir and Freundlich Isotherm models. The dynamics of dyes adsorption onto cellulose dusts were analyzed using four different kinetic models: the pseudo first-order, pseudo-second-order, intraparticle diffusion and Elovich models [27-28].

2.3. Characterization of Adsorbent

The characterization of Cellulose dust was performed by employing scanning electron microscopy (SEM), X-ray diffraction (XRD) and Fourier transformed infrared spectroscopy (FTIR).

2.4. Measurement of Adsorption Efficiency

The standard calibration method was adopted by running a series of standard solutions of the dye. The adsorption efficiency was estimated for the various runs performed using CCD. Following equation was used to obtain the percentage of the dye that has been removed.

$$\%Removal = \frac{C_o - C_t}{C_o} \times 100 \tag{1}$$

Where, C_o represents the initial concentration in $mg L^{-1}$ and C_t represents the concentration in $mg L^{-1}$ remaining in the solution after time t . The adsorption capacity of the synthesized nanocomposites was calculated by the given equation.

$$\%q_e = \frac{(C_o - C_e) \times V}{W} \tag{2}$$

C_o stands for the initial concentrations of the dye in $mg L^{-1}$ and C_e represents the equilibrium concentrations of the dye in $mg L^{-1}$. V stands for the volume in liters of the solution and W represents the mass in grams of the adsorbent [40].

2.5. Central Composite Design

A standard response surface methodology (RSM) design known as central composite design (CCD) was used to study the parameter for adsorption of MB with Cellulose dusts. The CCD is apposite to be fit in a quadratic surface to get the Optimum Operating Parameters (OOP) and to also assess the influence of parameters and the interactions of them [29]. This design encompasses bi level factorial design with at least one central point in the experimental conditions and with axial points also termed as star points. To improve the accuracy, increased number of central points are often used that minimizes a great deal of experimental error. The orthogonality and rotatability is ensured by the value of α that can be calculated using the following equation.

$$\alpha = \sqrt[4]{2^f} \tag{3}$$

The required experimental runs are calculated using the following Equation.

$$N = 2^f + 2f + N_o \tag{4}$$

In above equation, f represents the number of variables to be studied, 2^f indicates factorial points, $2f$ denotes the axial points and N_o is the replicates of center points [30].

In the current studies, CCD of 4 variable levels was employed having 5 central points. Alpha value of 2.0 was selected for the design. The number of runs were evaluated to be 30. The CCD along with the corresponding responses is given in Table 1. Each experimental run was conducted and the responses were correlated with four variables using the given quadratic polynomial equation.

$$y = \beta_0 + \sum_{i=1}^4 \beta_i x_i + \sum_{i=1}^4 \sum_{j=1}^4 \beta_{ij} x_i x_j + \sum_{i=1}^4 \beta_{ii} x_i^2 \tag{5}$$

In the above equation, y represents the anticipated response that is the percentage removal, X_i denote the independent variables. β_0 , β_i , β_{ii} and β_{ij} represent the model constant, linear coefficient, quadratic coefficient and the interaction coefficient respectively. Response Surface Methodology is coupled with CCD to evaluate the significant factors and solve multivariate equations to obtain Optimum Operating Parameters. The first or second order polynomial equations are provided in accordance with the empirical responses and also their consequent correlation. The variance analysis (ANOVA) aids in the prominent contribution of experimental parameters and interactions of them. The response surface is described by the 3D graph that is used to predict the Optimum Operating Parameters based on their P-values and F-values [30].

In the current research work, time, pH, adsorbent dosage, and MB concentration were taken as the 4 factors. Each of the factors was studied with 5 levels. Table 1 shows the coded and uncoded levels of these independent variables. 30 runs were performed and the removal of MB was taken as the responses (Table 2).

Table 1. Central composite design for four factors with their levels and observed responses

Factors	Units	Levels			- α	+ α
		Low (-1)	Central (0)	High (+1)		
(A) pH	-	6	8	10	4	12
(B) adsorbent dosage	$g L^{-1}$	2.2	3.3	4.4	1.1	5.5
(C) MB Concentration	$mg L^{-1}$	40	60	80	20	100
(D) time	min	90	150	210	30	270

3. Result and Discussion

3.1. Characterization of Adsorbent

The SEM micrographs of the CD before MB adsorption (Figure 3a) indicate regular morphology with an external, it reveals the surface texture and porosity of adsorbent. After MB adsorption (Figure 3b), the pores are absent on the surface of MB loaded CD, this change in morphology is due to accumulation of MB on CD surface.

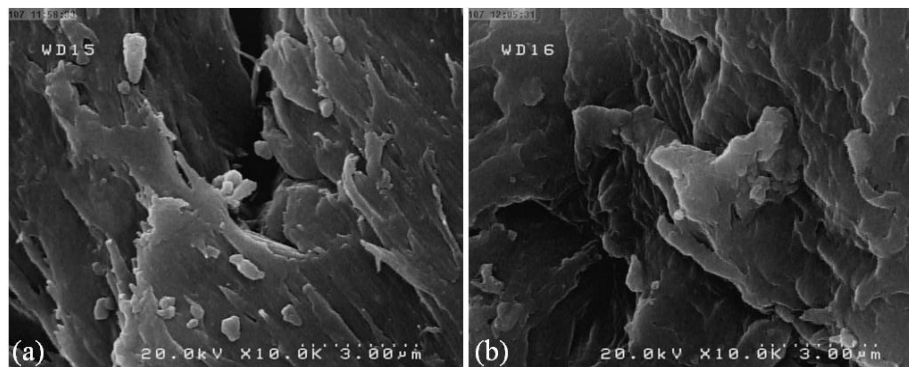


Figure 3. SEM analysis of (a) unloaded CD (b) CD loaded with MB.

X-ray diffraction studies of CD were performed to determine the structure of the adsorbent. The X-ray diffractogram of the CD is shown in Figure 4. The XRD diffractogram indicated that the prominent peaks appeared at 2θ positions of 15.82, 22.30, 29.50 and 36.03. The corresponding values of the Miller indices at these peaks were observed to be (101), (002) and (040). These are the two peaks corresponding two different phases in cellulose (I and II) [31]. The CDs were analyzed to contain orthorhombic structure as illustrated in JCPDS-Section 15. The crystal size of cellulose based on the Debye–Scherrer formula was found to be around 56 nm in accordance with FWHM which stands for full width at half-maximum of the (002) peak [32].

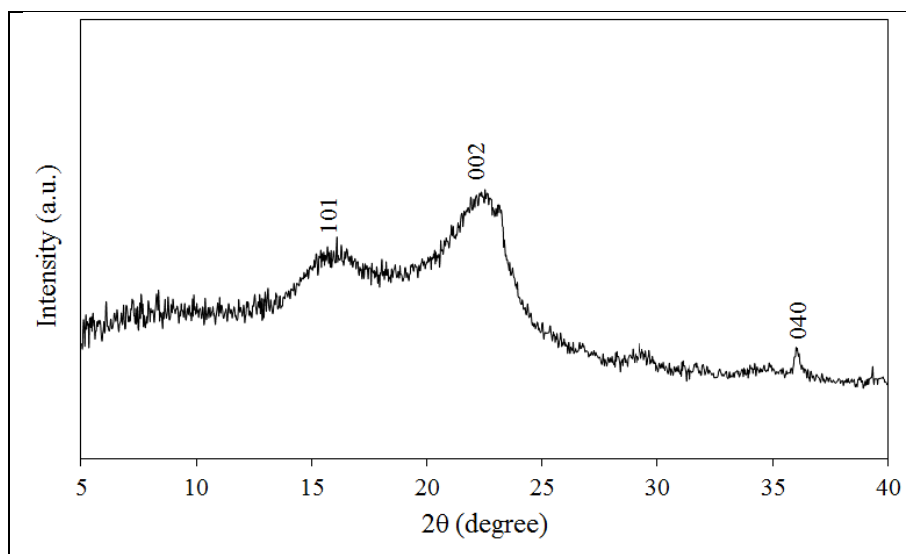


Figure 4. The XRD pattern of CD

FTIR Spectrum of CD before and after adsorption of MB in the range of 650–4000 cm^{-1} (Figure 5) exhibits characteristic peaks appeared at 3300 cm^{-1} because of O–H group (carboxylic acids, phenols and alcohols) on the surface of adsorbent as in cellulose, pectin and lignin [33, 34]. While the peak at 2900 cm^{-1} representing Carbon-Hydrogen distortion in in biomass structure (cellulose, hemicellulose, lignin) [35].

The bands in region 1595 cm^{-1} and 1504 cm^{-1} are assigned to C=C aromatic stretching of lignin [36]. The bands at 1230 cm^{-1} and 1025 cm^{-1} corresponds to C-O stretching vibration of carboxylic acids and alcohols. The band at 896 cm^{-1} can be due the bending vibration of C-H [37, 38].

After adsorption of MB, we note that the peaks intensity of O-H (3300 cm^{-1}), C-O (1230 cm^{-1} , 1025 cm^{-1}) and C-H (2900 cm^{-1}) decreased, these shifts indicate that an interaction was occurred between MB and these groups.

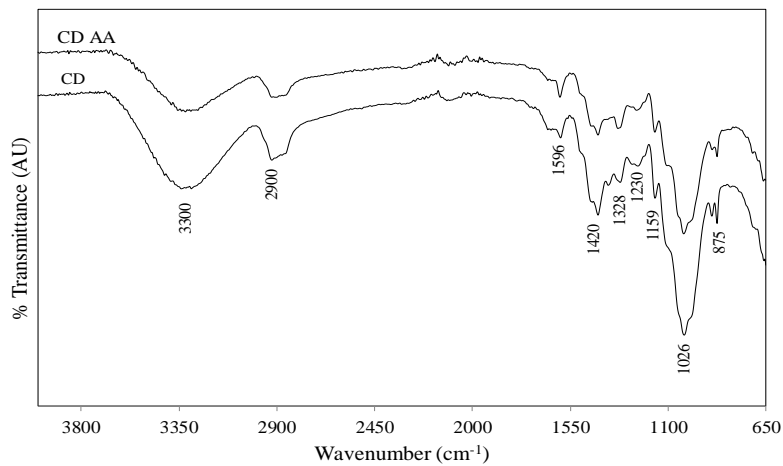


Figure 5. The FTIR spectrums of CD before and after adsorption (CD AA) of MB

3.2. Response Surface Methodology

The influences of different parameters such as pH (A), MB initial concentration (B), adsorbent dosage (C) and time (D) on MB removal percentage by CD were investigated and optimized. Optimization of MB removal percentage using RSM and CCD was employed to evaluate the interaction among the significant variables and also to determine their optimal values.

Table 2 represents the regression results of the response surface from CCD. Statistical analysis was carried out by fitting experimental data to a general model equation given in Equation 5. An empirical relationship articulated by a second-order polynomial equation with interaction provisos was fitted between the experimental results obtained on the basis of CCD model and the input functional variables. The last equation which is provided for MB dye in terms of actual (natural) parameters is applied which is the form of the transformed model equation in terms of coded and actual factors (Equations 6 and 7):

Final equation in terms of actual factors for MB removal

$$\begin{aligned}
 \text{Removal (\%)} = & 0.5533 + 0.0139 X_A + 0.0552 X_B + 0.00285 X_C + 0.00106 X_D - 4.823 \\
 & \times 10^{-4} X_A X_B + 7.813 \times 10^{-6} X_A X_C - 1.615 \times 10^{-5} X_A X_D - 9.943 \\
 & \times 10^{-5} X_B X_C + 7.670 \times 10^{-5} X_B X_D - 2.604 \times 10^{-6} X_C X_D - 3.411 \\
 & \times 10^{-4} X_A^2 - 0.00598 X_B^2 - 1.9349 \times 10^{-5} X_C^2 - 2.636 \times 10^{-6} X_D^2
 \end{aligned} \tag{6}$$

Final equation in terms of coded factors for MB removal

$$\begin{aligned}
 \text{Removal (\%)} = & 0.95 + 9.875 \times 10^{-3} X_A + 0.019 X_B + 4.375 \times 10^{-3} X_C + 0.023 X_D \\
 & - 1.062 \times 10^{-3} X_A X_B + 3.125 \times 10^{-4} X_A X_C - 1.937 \times 10^{-3} X_A X_D - 2.187 \\
 & \times 10^{-3} X_B X_C + 5.063 \times 10^{-3} X_B X_D - 3.125 \times 10^{-4} X_C X_D - 1.365 \\
 & \times 10^{-3} X_A^2 - 7.240 \times 10^{-3} X_B^2 - 7.740 \times 10^{-3} X_C^2 - 9.490 \times 10^{-3} X_D^2
 \end{aligned} \tag{7}$$

Table 2. Design matrix for the central composite designs

Run No.	Experimental factors				MB Removal (%)
	A	B	C	D	
1	6	4.4	80	210	0.965
2	8	3.3	60	150	0.953
3	10	2.2	80	90	0.91
4	8	3.3	20	150	0.895
5	12	3.3	60	150	0.97
6	8	3.3	100	150	0.939
7	10	4.4	80	90	0.92
8	6	4.4	40	90	0.91
9	8	1.1	60	150	0.87
10	8	5.5	60	150	0.968
11	8	3.3	60	150	0.945
12	8	3.3	60	150	0.95
13	10	4.4	40	90	0.925

14	10	4.4	80	210	0.972
15	6	4.4	80	90	0.896
16	4	3.3	60	150	0.915
17	8	3.3	60	150	0.949
18	6	4.4	40	210	0.959
19	10	4.4	40	210	0.968
20	10	2.2	80	210	0.927
21	6	2.2	80	90	0.885
22	6	2.2	40	90	0.875
23	6	2.2	40	210	0.913
24	8	3.3	60	270	0.96
25	8	3.3	60	150	0.945
26	8	3.3	60	30	0.86
27	10	2.2	40	210	0.935
28	6	2.2	80	210	0.917
29	10	2.2	40	90	0.89
30	8	3.3	60	150	0.949

The statistical significance of quadratic model at 95% confidence level was tested by ANOVA results, which are capable of modeling to fit the response parameters. The variables that are presented in ANOVA (Table 3) were the error terms, the main effects and the interaction effects. The amounts of F and p were applied to stand for the significance of these parameters in the quadratic model. SS which stands for the sum of squares and DF which stands for the degrees of freedom and MS which stands for the mean squares were computed and in accordance with that the amount of F was gained by dividing the amount of MS which is related to each term and the amount of MS that is gained by dividing SS by DF.

In the quadratic model, the F value of 20.516 indicated that the model was statistically significant. The p value of less than 0.0001 confirm high significancy of model, while its accuracy was confirmed by its high determination of coefficient R^2 . In the present study, the value of $R^2 = 0.9504$ indicated that 4.96 % of the total variables were not explained by the model. The amount of the adjusted determination coefficient (R^2_{adj}) was 0.9040 has reasonable agreement with observed R^2 .

Table 3. Analysis of variance (ANOVA) for the response surface quadratic model

Source	Sum of Squares	DF	Mean Square	F value	P value
Model	0.0289	14	0.002066	20.516	< 0.0001
A-pH	0.00234	1	0.002340	23.243	0.0002
B- adsorbent dosage	0.00878	1	0.008779	87.179	< 0.0001
C-concentration of MB	0.000459	1	0.000459	4.5621	0.0496
D-time	0.012376	1	0.012376	122.91	< 0.0001
AB	1.806×10^{-5}	1	1.806×10^{-5}	0.1794	0.6779
AC	1.563×10^{-6}	1	1.563×10^{-6}	0.0155	0.9025
AD	6.006×10^{-5}	1	6.006×10^{-5}	0.5965	0.4519
BC	7.656×10^{-5}	1	7.656×10^{-5}	0.7603	0.397
BD	4.101×10^{-4}	1	4.101×10^{-4}	4.0723	0.0618
CD	1.563×10^{-6}	1	1.563×10^{-6}	0.0155	0.9025
A ²	5.107×10^{-5}	1	5.107×10^{-5}	0.5072	0.4873
B ²	0.00144	1	0.00144	14.277	0.0018
C ²	0.00164	1	0.00164	16.317	0.0011
D ²	0.00247	1	0.00247	24.529	0.0002
Residual	0.00151	15	0.0001	-	-
Lack of Fit	0.001463	10	0.0001463	15.399	0.0038
Pure Error	4.75×10^{-5}	5	9.5×10^{-6}	-	-

Cor Total	0.03043	29	-	-	-
Std. dev.	0.0100	-	R-squared	-	0.9504
C.V. %	1.0815	-	Adjusted R-squared	-	0.9040
Mean	0.9278	-	Predicted R-squared	-	0.7209
PRESS	0.0085	-	Adequate Precision	-	15.996

The lack of fit term is actually non-significant and shows that quadratic model is valid for explanation of experimental work. Table 3 represents that PRESS which stands for the prediction error sum of squares, gives a fine residual scaling and also reveals that sufficient precision (signal to noise ratio) must be higher than 4 to construct desirable. Thus, in the quadratic model for MB's removal process, the ratio of 15.996 indicate an adequate signal for the model to be used to navigate the design space [39, 40].

Figure 6a shows the predicted versus actual graph for MB removal percentage, which indicated actual values were distributed relatively close to the straight line and denote. This indicated that the quadratic model was needed for estimating the percentage of the optimized removal of MB under empirical factors.

As shown in Figure 6b (normal probability plot), the normality assumption was satisfied as the residual plot approximated along a straight line. The plot of studentized residuals versus run number was tested and shown in Figure 6c. The residuals were distributed in a random manner near ± 3.0 , which show high efficiency of model for well presentation explanation of real data.

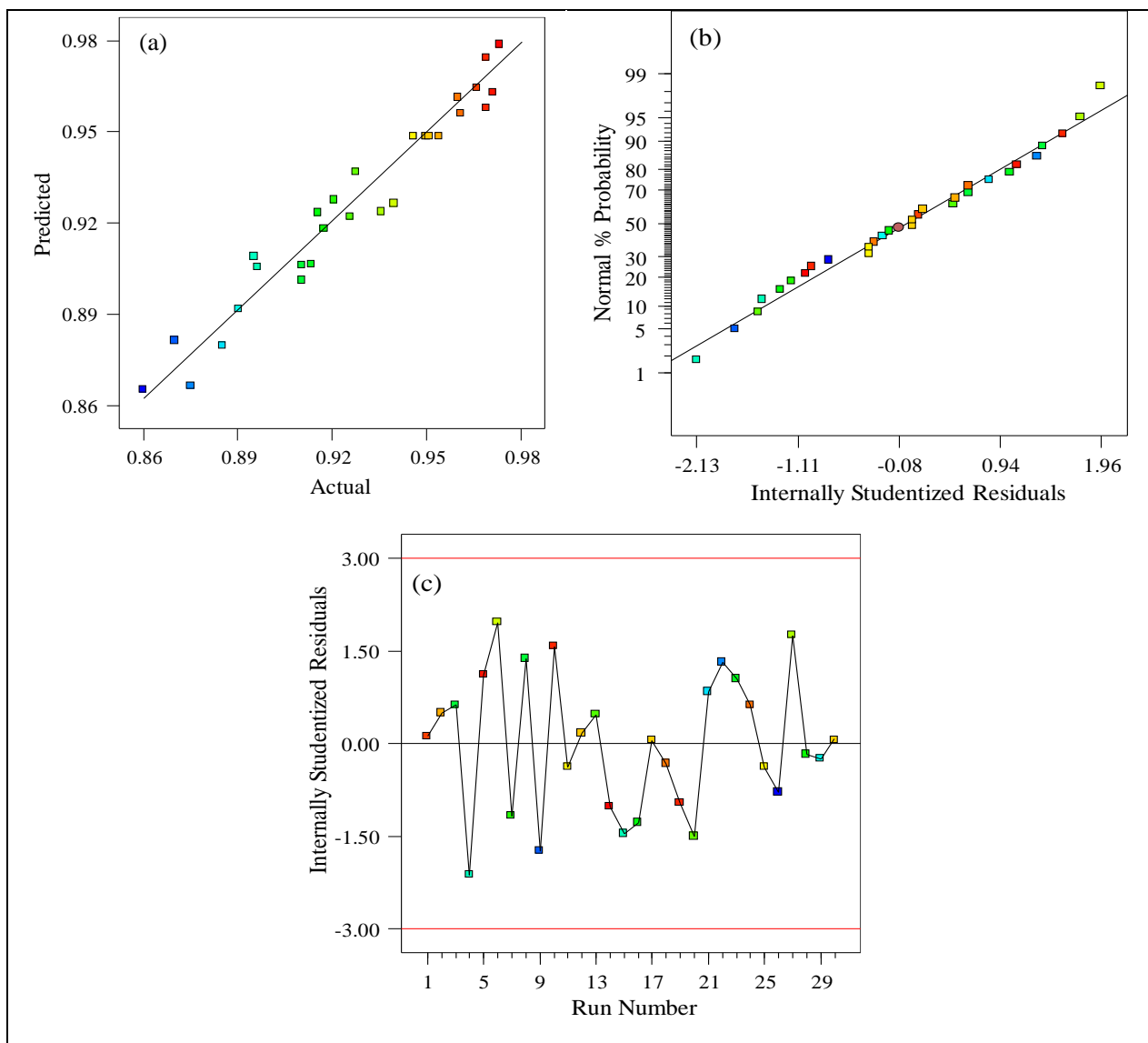


Figure 6. Predicted versus experimental data (a) Normal probability plot of residuals; (b) Internally studentized residual and run; (c) Plots for removing MB

3.2.1. Effect of Process Variables on MB Removal

Based on the data from Table 3, the mean effect plots are displayed in Figure 7. At the time that the response is at its biggest possible amount, the system in these experiments is optimized. It can be observed that pH, adsorbent dosage, concentration of MB and time have positive impact on the MB removal. According to Table 4, adsorbent dosage is has the most influence as an individual parameter on MB removal ($F = 87.18, p < 0.0001$).

In the case of adsorbent dosage and time an obvious increment in MB removal is observed by increasing the adsorbent dosage and time. Furthermore, the MB removal was almost unchanged by increasing the MB concentration. Considering the pH, MB removal was enhanced by increasing this factor. At low high level of X_A, X_B and low level of X_C with increasing time to upper level, MB removal went higher.

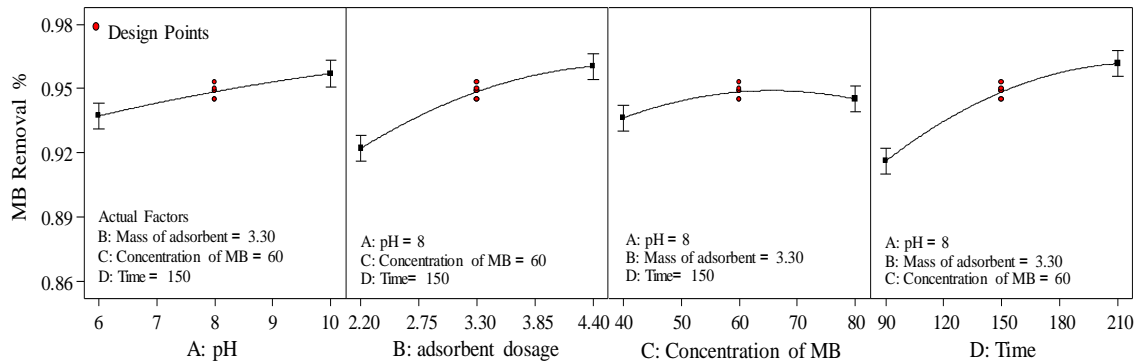


Figure 7. Major effects of each parameter on MB removal

3.2.2. 3D Response Surface Plot

3D response surface plots for the measured responses were constructed to better find out about the impacts of independent parameters and interactions of them on the dependent ones. The interaction between the corresponding parameters was negligible when a circular 3D response of the surfaces was found. Further, elliptical or saddle natures of the plot indicate that the interaction between the corresponding variables was significant [41]. Furthermore, parallel lines suggest that there was no significant interaction. In 3D response plots, the upper horizontal axis represents the dependent variable. Response surface methodology (RSM) was applied by considering all the significant interactions among variables in the CCD to find the most important effects. The 3D graphs correspond to MB removal percentage, which describe their variation against the parameters that were plotted at optimum values of others factors and their curvatures is proportional with presence of significance interaction among the variables.

Figure 8 demonstrates that the time and dose of the adsorbent highly affect the MB removal percentage and its bigger amount at higher time especially at high amount of adsorbent was found. Effect of higher dosage of adsorbent probably due to larger reactive surface area and functional group which are sufficient for complexation and binding MB.

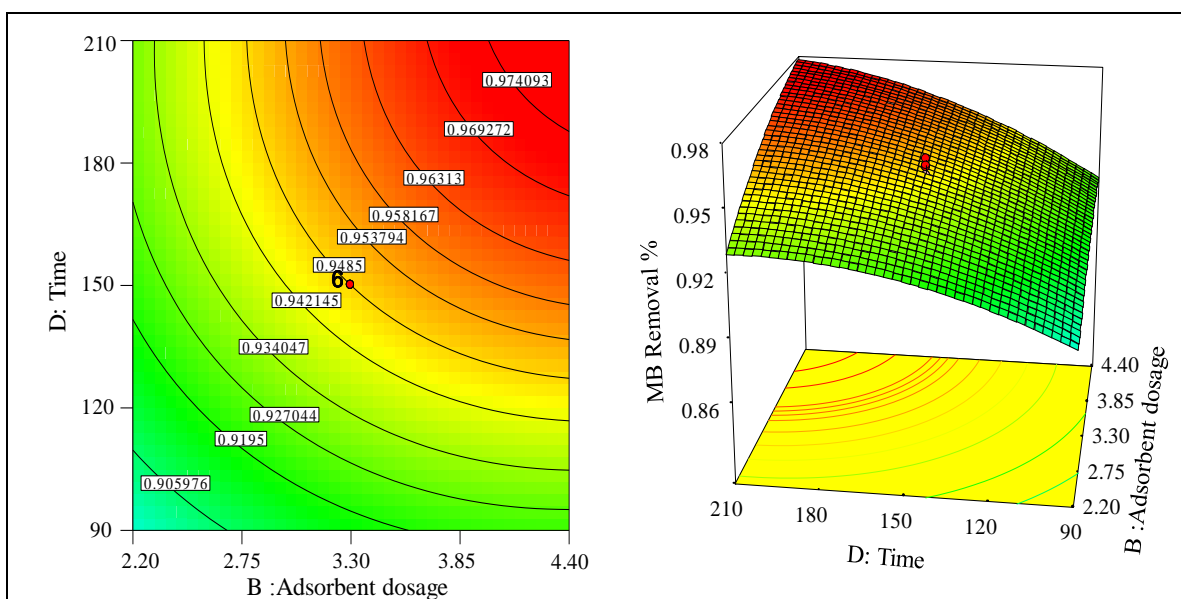


Figure 8. Response surface (left) and contour (right) plots for the combined effects of adsorbent dosage and time on the MB removal

Figure 9 shows the effect of time and initial MB concentration on MB removal show that removal rate at beginning of the process is rapid because of high available surface area and vacant site of adsorbent, while their subsequent saturation is proportional with reduction in signal. Although the actual amount of MB at higher initial MB concentration, probably due to saturation of reactive sites or increasing the repulsive force between the adsorbed MB and bulk molecules lead to reduction in the percentage of removal [42]. The fact that the removal percentage of MB is greatly dependent to its initial concentration reveals the great impact and influence of diffusion and adsorption site on removal percentage. Curve portion in 3D plot Figure 8 suggested that significant removal of dye was found to have almost good interactive influence of these two variables.

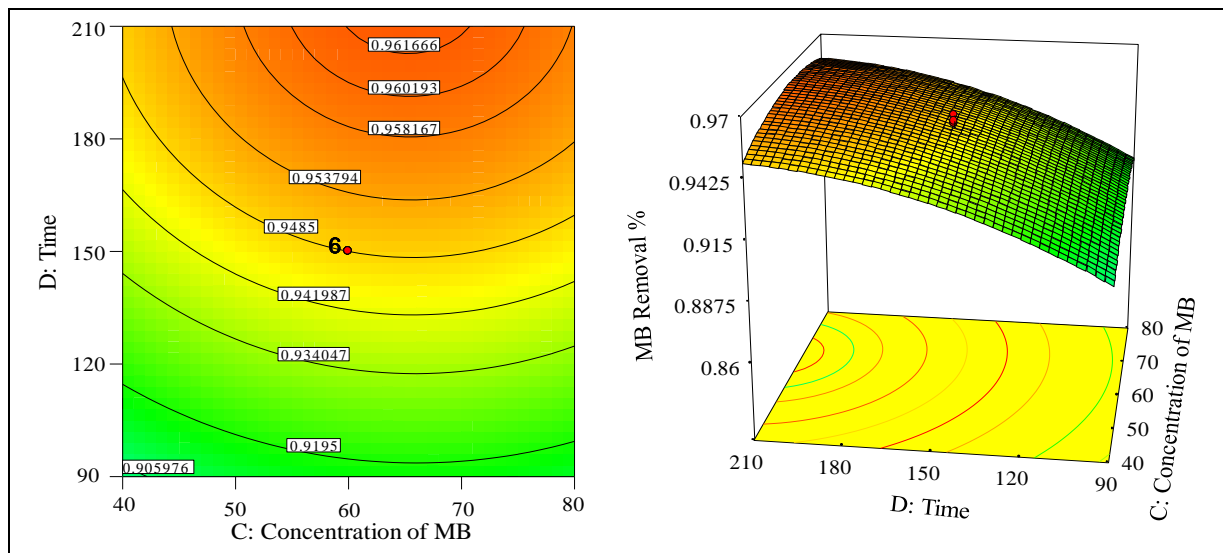


Figure 9. The response surface plots (left) and their corresponding contour plots (right) showing effect of time and concentration of MB for the removal of MB

3.2.3. Model Validation for MB Dye

The validation experiment was performed by several sets of different combination of variables (pH, adsorbent dosage, MB concentration and time) each within their experimental ranges. The model predicted value and experimental value were used to determine the validity of model (Table 4). The fact that the estimated results properly agree with the empirical amounts is approved by these results.

Table 4. Comparing the experimental and predicted values.

Run No.	pH	adsorbent dosage (g L ⁻¹)	MB concentration (mg L ⁻¹)	Time (min)	MB removal	
					Experimental ^a	Predicted
1	7	0.2	30	80	0.815±0.026	0.772
2	8.5	0.35	50	140	0.862±0.074	0.839
3	10.5	0.4	75	200	0.936±0.053	0.863

^a Mean ± standard deviation.

On the basis of best fitted equation, the model optimization of variables was done to get the maximum removal percentage. 17 optimized results were obtained and it was found that the maximum optimized conditions for MB dye (pH 9.84, adsorbent dosage 4.38 g L⁻¹, MB concentration 75.50 g L⁻¹ and time 208.13 min) which showed 98.05% removal of MB.

3. Adsorption Kinetics

The research regarding kinetic adsorption was done in order to describe the dynamics of the process of adsorption in connection with the rate constant's order. Various kinetic models are accessible in order to find out the function of the adsorbent and also to analyze the controlling mechanism of the process of adsorption. In the research the pseudo-first order, pseudo-second order, Elovich and intraparticle diffusion models have been tested for the adsorption of MB onto the adsorbent.

The kinetic parameters of pseudo-first order and pseudo second order were determined from the respective plots at 100 mg L⁻¹ initial concentration of the MB (Table 5). On careful analysis data illustrated that adsorption process agrees with pseudo second order model with R² value near unity (R² around 0.999). The experimental q_e of the pseudo first order model did not follow the calculated value, which show failure and inability of this model for fitting data.

Also, more accurate values of q_e are found in pseudo second order model which specifies that the process of adsorption is restrained by chemisorption. Chemisorption is the share of electrons or the exchange of them among the dye molecules and the active sites on the adsorbents [43, 44].

In similar studies pseudo second order was reported as the appropriate kinetic model in the adsorption of reactive dye on magnetic composite of activated carbon and iron oxide and in removal of Direct Blue color-06 using activated carbon [45, 46].

In order to model the mass transfer resistance that is related to the adsorption which is under research, the Intraparticle diffusion and Elovich models were applied. The amounts of Intraparticle diffusion rate constant (K_{id}) is computed from the slope of the curve. The amounts of R^2 , K_{id} and intercept are presented in Table 5. The amounts of R^2 for this diffusion model were 0.991 that also shows the probability of in mass transfer and adsorption rate. The intercepts propose an opinion about the thickness of the boundary layer. The larger the intercept is equal to greater value of boundary layer effect. The Elovich constants can be calculated from the plots of q_t versus $\ln t$ and regarding this matter the small amounts of R^2 (the correlation coefficient) verifies that the model is not suitable for MB's adsorption by the adsorbent (Table 5).

Table 5. Kinetic parameters removal of MB

Model	Plot	Parameters	Value	
First-order kinetic	Plot the values of $\text{Log}(q_e - q_t)$ versus t to give a linear relationship from which k_1 and q_e can be determined from the slope and intercept, respectively	$\text{Log}(q_e - q_t) = \text{Log}q_e - \frac{k_1 t}{2.303}$	K_1 (min^{-1})	9.212×10^{-3}
			$q_{e(\text{calc})}$ ($\text{mg} \cdot \text{g}^{-1}$)	12.142
			$q_{e(\text{exp})}$ ($\text{mg} \cdot \text{g}^{-1}$)	24.1875
			R^2	0.9186
Pseudo-second-order kinetic	Plot the values of t/q_t versus t to give a linear relationship from which q_e and k_2 can be determined from the slope and intercept, respectively	$\frac{t}{q_t} = \frac{1}{k_2 q_e^2} + \frac{t}{q_e}$	k_2 ($\text{g} \cdot \text{mg}^{-1} \cdot \text{min}^{-1}$)	5.033×10^{-3}
			$q_{e(\text{calc})}$ ($\text{mg} \cdot \text{g}^{-1}$)	23.226
			$q_{e(\text{exp})}$ ($\text{mg} \cdot \text{g}^{-1}$)	24.1875
			R^2	0.9997
Intraparticle diffusion	The values of K_{id} and C were calculated from the slope and intercept of the plot of q_t versus $t^{0.5}$, respectively	$q_t = k_{id} t^{0.5} + C$	K_{id} ($\text{mg} \cdot \text{g}^{-1} \cdot \text{min}^{-0.5}$)	0.2269
			C	11.099
			R^2	0.9753
Elovich	Plot the values of q_t versus $\ln t$ to give a linear relationship from which b and a can be determined from the slope and intercept, respectively	$q_t = \frac{1}{b} \ln(ab) + \frac{1}{b} \ln t$	a ($\text{mg} \cdot \text{g}^{-1} \cdot \text{min}^{-1}$)	767.287
			b ($\text{g} \cdot \text{mg}^{-1}$)	0.8237
			R^2	0.9829

4. Adsorption Equilibrium Study

The adsorption isotherm studies were conducted the adsorption system (MB dye solution and Cellulose dusts) at the optimum conditions (initial MB concentration: 20–100 $\text{mg} \cdot \text{L}^{-1}$, pH of 11, adsorbent mass 4.4 $\text{g} \cdot \text{L}^{-1}$, time of 210 min at 25 °C). Passing of time makes the system to finally reach equilibrium. After reaching that state, the mixture was centrifuged in order to distinct the supernatant and spent adsorbent. Various isotherms regarding adsorption (Langmuir, Freundlich, Temkin and Harkins-Jura models) were applied to investigate the empirical equilibrium data of the adsorption of dye. Their constant factors and the correlation coefficient (R^2) were computed which are presented in Table 6.

The amounts of K_L which represents the Langmuir adsorption constant ($\text{L} \cdot \text{mg}^{-1}$) and q_0 which stands for theoretical maximum adsorption capacity ($\text{mg} \cdot \text{g}^{-1}$) were attained from the intercept and slope of the plot of C_e/q_e vs. C_e , respectively.

In fit best in freundlich isotherms model, this implies that the adsorption heat of all molecules drops linearly with the rise in the adsorbent surface occupancy. It also proposes that the properties of adsorptions be determined by a homogenous distribution of binding energies up to the time that a threshold binding energy is gained [47]. Same results were attained in the adsorbtion of methylene blue on to date palm leaves [48]. The parameters of Freundlich isotherm model like k_F ($(\text{mg} \cdot \text{g}^{-1})$) ($(\text{mg} \cdot \text{L}^{-1})$) and n (which represents the adsorption's intensity and capacity) were computed from the intercept and slope of the linear plot of $\ln q_e$ vs $\ln C_e$ respectively.

The heat of the adsorption and the adsorbent adsorbate interaction were evaluated by using temkin isotherm model. Parameters of the model are as follows: B represents the temkin constant corresponding to the adsorption's heat (J mol^{-1}), T stands for the absolute temperature (K), R denotes the universal gas constant ($8.314 \text{ J mol}^{-1} \text{ K}^{-1}$) and K_T is the equilibrium binding constant (L mg^{-1}).

Harkins Jura Isotherm model has the assumption of the heterogeneous surface and explains the multilayer adsorption of the adsorbate onto the adsorbent's surface. The values of the respective model constants were obtained from the slope and intercept of the plot between $\log C_e$ and $1/q_e^2$ as presented in Table 6.

The experimental results show that the Harkins-Jura Isotherm model was well followed by CD adsorbent for removal of MB and the R^2 value was observed to be 0.857 as shown in Table 6.

Table 6. Isotherm constant parameters and correlation coefficients calculated for removal of MB

Isotherm	Equation	Parameters	Value
Langmuir	$\frac{C_e}{q_e} = \frac{C_e}{q_o} + \frac{1}{K_L q_o}$ A plot C_e/q_e versus C_e should indicate a straight line of slope $1/q_o$ and an intercept of $1/(K_L q_o)$	K_L	0.04185
		q_o	153.85
		R^2	0.4174
Freundlich	$\text{Log} q_e = \text{Log} K_F + \frac{1}{n} \text{Log} C_e$ The values of K_F and $1/n$ were determined from the intercept and slope of linear plot of $\ln q_e$ versus $\ln C_e$, respectively	K_F	6.285
		n	1.102
		R^2	0.993
Temkin	$q_e = B \ln A + B \ln C_e$ Values of B and A were calculated from the plot of q_e against $\ln C_e$	A	1.827
		B	10.39
		R^2	0.916
Harkins-Jura	$\frac{1}{q_e^2} = \frac{B_H}{A} - \frac{1}{A} \text{Log} C_e$ Values of A and B_H were calculated from the plot of $1/q_e^2$ against $\text{Log} C_e$	A	17.857
		B_H	0.543
		R^2	0.857

5. Conclusion

Cellulose Dusts is a low-cost and efficient adsorbent for the removal of methylene blue from aqueous solutions. The CD properties was determined by FTIR, SEM and XRD. RSM based CCD was carried out for optimization of variable factors like the duration of contact, agitation speed, the dose of adsorbent, pH and MB initial concentration. The optimum conditions for MB removal were MB concentration of 75.50 g L^{-1} , pH 9.84, adsorbent dosage 4.38 g L^{-1} and time 208.13 min, where 98.05% of MB removal could be obtained from the proposed model. The data which was obtained from the equilibrium state was satisfactorily in agreement with Freundlich isotherm equations ($R^2=0.99$). It was concluded that the adsorption rate was close agreement with the pseudo-second-order model and had an acceptable correlation coefficient. Intraparticle diffusion model's results proposed that the methylene blue's adsorption on CD is a process consisting several steps which includes sorption on the external surface and also diffusion inside the adsorbent. Our results taken together show that CD an interesting low-cost adsorbent candidate for the removal of methylene blue from aqueous solutions

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