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Molecular conductors as nanoparticles in the presence of long-chain alkyl imidazolium salts or amphiphilic molecules

Synthesis and thermoanalytical studies

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Abstract Nanoparticles of two molecule-based conductors, namely TTF·TCNQ and TTF[Ni(dmit)₂]₂, have been prepared in organic solution in the presence of ionic or nonionic species bearing a long-chain alkyl group, acting as growth-controlling agents. The size, morphology, and state of dispersion of the nanoparticles depended on the nature of the growth-controlling agent and the reaction temperature. In the presence of a long-chain alkyl-based ionic liquid at -50 °C, electron micrographs evidence that TTF·TCNQ nano-objects are frequently elongated, whereas TTF[Ni(dmit)₂]₂ nanoparticles are aggregated. In the presence of a neutral long-chain alkyl-based imine at room temperature, nanoparticles are spherical (mean diameter

<20 nm) and well dispersed. Vibration spectra evidence that the amounts of charge transfer for TTF·TCNQ and TTF[Ni(dmit)₂]₂ as nano-objects are very similar to those for the same phases as bulk materials. According to the thermoanalytical investigations, the prepared nanoparticles are stable thermally up to approximately 200 °C, and their decomposition is generally a multi-step process. Their heat treatment results in various sulfur-containing volatiles (CS₂, SO₂, H₂S); moreover, HCN is also detected in the case of nitrogen-containing molecules (TCNQ).

Keywords Tetrathiafulvalene · Metal dithiolene complexes · Nanoparticles · Ionic liquids · *N*-Octylfurfuryl-imine · Thermoanalytical measurements

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Introduction

Molecule-based conductors attracted a renewed interest since the 2000s, in particular as regards to their processing as thin films or nano-objects [1]. This represents an important asset for potential applications. TTF·TCNQ and TTF[Ni(dmit)₂]₂ (Fig. 1) belong to the family of donor acceptor molecule-based conductors [2]. They are among the most studied molecular metals because of (1) the relatively low cost of the starting compounds; (2) an easy synthesis; (3) their interesting physical properties. TTF·TCNQ is a typical one-dimensional compound in which a charge transfer does exist between the TTF molecule (electron donor) and the TCNQ molecule (electron acceptor). It exhibits a metal-like conductivity down to about 55 K [2]. TTF[Ni(dmit)₂]₂ exhibits a metallic behavior down to 3 K and was the first metallo-organic-based material undergoing a superconducting transition (at 1.6 K under application of a hydrostatic pressure of 7 kbar) [3].

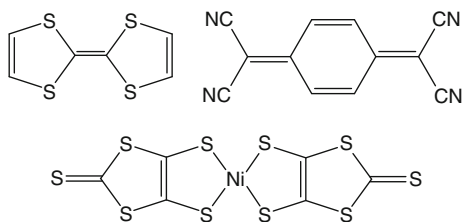


Fig. 1 Molecular formulas for TTF (*left*), TCNQ (*right*), and Ni(dmit)₂ (*bottom*)

Nanowires of molecule-based conductors have been widely described [4–7]. This nanowire-like morphology is not surprising for quasi-1-D molecular conductors. However, these compounds do not show a natural tendency to grow as roughly spherical objects. Ionic liquids containing the BMIM⁺ cation (BMIM⁺: 1-butyl-3-methylimidazolium, Fig. 2) are very good candidates to stabilize spherical metallic nanoparticles [8]. In that case, reactions are conducted in pure ionic liquid, acting as both the solvent and stabilizing agent. We have recently reported the preparation of TTF·TCNQ and TTF[Ni(dmit)₂]₂ nanoparticles in the presence of a mixture [BMIM][X]/acetonitrile [X⁻: BF₄⁻ or (CF₃SO₂)₂N⁻]. Observed mean diameters were in the 20–40 nm range [9–12]. In our case, the growth as spherical nano-objects in an ionic liquid/acetonitrile mixture could be explained by a self-organization of imidazolium salts as nano-domains in the solvent, i.e., CH₃CN [13]. Moreover, very small and well-dispersed metallic nanoparticles (diameter < 10 nm) were extensively prepared by adding neutral amphiphilic molecules, such as long alkyl chain amines or thiols, to the reaction medium [14]. The amphiphilic molecule, acting as a stabilizing agent, controlled the particle growth through coordination to the metal center. We have recently shown that octylamine could also stabilize 40 nm-TTF·TCNQ nanoparticles, in which TCNQ-OA (TCNQ whose a CN group has been substituted by an amino group) molecules were present at the particles surface and were responsible for their stabilization and their dispersibility in common organic solvents [15]. Furthermore, we have recently published the preparation of (TMTSF)₂·ClO₄ nanocrystals (TMTSF: tetramethyltetraselenafulvalene) using the electrocrystallization technique in the presence of *N*-octylfurfuryl-imine (Fig. 2), acting as a

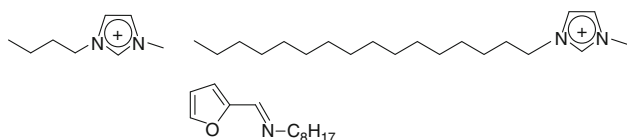


Fig. 2 Molecular formulas for BMIM⁺ (*left*), HDMIM⁺ (*right*), and *N*-octylfurfuryl imine (*bottom*)

growth-controlling agent [16]. Contrary to what observed in TTF·TCNQ/octylamine nanoparticles, the *N*-octylfurfuryl-imine molecule did not react with the TMTSF molecule but was presumably associated to the TMTSF through π -overlap via the furfuryl group during the growth, whereas the C₈ chain prevented nanocrystals from aggregation. High-resolution transmission electron micrographs also evidenced that nanocrystals (sizes in the 20–70 nm range) were actually made of aggregated individual 2–6 nm nanoparticles [16].

In order to obtain nanoparticles exhibiting mean diameters as small as 10 nm, we have evaluated, in the present paper, the use of (1) a bulky imidazolium cation, i.e., 1-hexadecyl-3-methylimidazolium (HDMIM⁺, Fig. 2), (2) the *N*-octylfurfuryl-imine amphiphilic molecule. The first part of this paper is devoted to the synthesis and spectral characterization of the molecule-based conducting nanoparticles, whereas the second part is dedicated to thermoanalytical studies of the TTF·TCNQ and TTF[Ni(dmit)₂]₂ nanoparticles.

Materials and methods

All syntheses have been carried out under an argon atmosphere using freshly distilled and degassed solvents. TTF was purchased from Sigma, TCNQ from Fluka, and [HDMIM][CF₃SO₂]₂N from Io-li-tec. Other starting compounds were prepared following previously described procedures: (TTF)₃(BF₄)₂ [17], [(*n*-C₄H₉)₄N][Ni(dmit)₂] [18], and *N*-octylfurfuryl-imine [19].

TTF·TCNQ nanoparticles in the presence of [HDMIM][CF₃SO₂]₂N

A solution of 90 mg of TTF (0.44 mmol) and 780 mg of [HDMIM][CF₃SO₂]₂N (1.32 mmol) in 20 mL of acetonitrile was added dropwise to a solution of 90 mg of TCNQ (0.44 mmol) in 20 mL of acetonitrile at –50 °C. A fine black precipitate was obtained as soon as the TTF/[HDMIM][CF₃SO₂]₂N solution was added under stirring. Stirring was maintained over a period of 2 h at –50 °C. The suspension was then allowed to warm to room temperature. The black solid was filtered off, washed with 3 × 10 mL of acetonitrile, and finally dried under vacuum. The resulting black powder was air stable (yield: 94 %).

TTF[Ni(dmit)₂]₂ nanoparticles in the presence of [HDMIM][CF₃SO₂]₂N

A solution of 115 mg of [(*n*-C₄H₉)₄N][Ni(dmit)₂] (0.16 mmol) in 12 mL of acetone was added dropwise to a suspension of 60 mg of (TTF)₃(BF₄)₂ (0.08 mmol) and 146 mg of [HDMIM][CF₃SO₂]₂N (0.24 mmol) in 5 mL of acetonitrile at –50 °C. A fine black precipitate was

observed, and stirring was maintained over a period of 2 h at $-50\text{ }^{\circ}\text{C}$. The suspension was then allowed to warm to room temperature. The black solid was filtered off, washed with $3 \times 5\text{ mL}$ of acetonitrile, and finally dried under vacuum. The resulting black powder was air stable (yield: 60 %).

TTF[Ni(dmit)₂]₂ nanoparticles in the presence of *N*-octylfurfuryl-imine (chemical procedure)

A solution of 46 mg of $[(n\text{-C}_4\text{H}_9)_4\text{N}][\text{Ni}(\text{dmit})_2]$ (0.066 mmol) in 10 mL of acetone was added dropwise to a solution of 26 mg of $(\text{TTF})_3(\text{BF}_4)_2$ (0.033 mmol) and 23 μL of *N*-octylfurfuryl-imine (0.099 mmol) in 4 mL of acetonitrile at $25\text{ }^{\circ}\text{C}$. A fine black precipitate was observed, and stirring was maintained over a period of 6 h at $25\text{ }^{\circ}\text{C}$. The black solid was filtered off, washed with $3 \times 5\text{ mL}$ of acetonitrile, and finally dried under vacuum. The resulting black powder was air stable (yield: 55 %).

TTF[Ni(dmit)₂]₂ nanoparticles in the presence of *N*-octylfurfuryl-imine (electrochemical procedure)

The synthesis took place in a H-shaped electrocrystallization cell with a platinum wire in each compartment separated by a glass frit in the H cross piece. TTF (10 mg; 0.05 mmol), $[(n\text{-C}_4\text{H}_9)_4\text{N}][\text{Ni}(\text{dmit})_2]$ (70 mg; 0.10 mmol), 34 μL of *N*-octylfurfuryl-imine (0.150 mmol), and 12 mL acetonitrile were introduced in the anodic compartment. $[(n\text{-C}_4\text{H}_9)_4\text{N}][\text{Ni}(\text{dmit})_2]$ (10 mg; 0.014 mmol) and 12 mL acetonitrile were introduced in the cathodic compartment. Nanoparticles synthesis was conducted under galvanostatic conditions ($100\text{ }\mu\text{A cm}^{-2}$) during 2 days at room temperature. During electrolysis, the content of the cell was vigorously agitated with a magnetic stirrer. The black powder was collected from the anode, washed with $2 \times 5\text{ mL}$ of acetonitrile, and finally dried under vacuum. The resulting black powder was air stable (yield: 55 %).

Characterization methods

For transmission electron microscopy (TEM) observation, powder (0.5 mg) was dispersed in diethyl ether (2 mL) under slow stirring for 1 min. The TEM specimens were then prepared by evaporation of droplets of suspension deposited on carbon-supported copper grids. The experiments were performed on a JEOL Model JEM 1011 operating at 100 kV. Infrared spectra (in KBr matrix) were taken at room temperature on a PerkinElmer Spectrum GX spectrophotometer. Raman spectra were recorded at 80 K using a DILOR XY micro-Raman (785 nm laser line). The room temperature conductivity of the samples was measured on a compressed pellet by two-probe contacts using

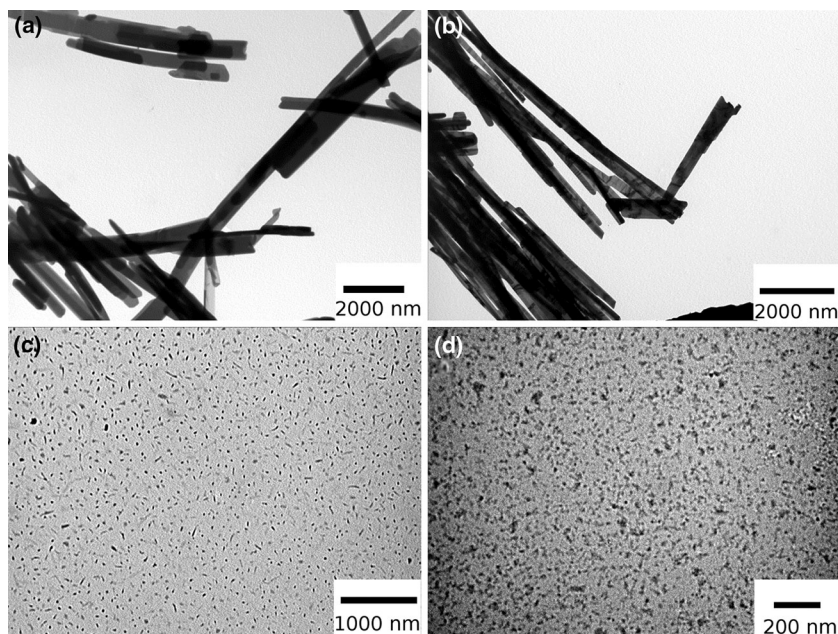
homemade equipment. The thermal analysis measurements were taken on a SETARAM Labsys Evo TG DSC thermal analyzer, in flowing high purity (6.0) Helium atmosphere (flow rate 90 mL min^{-1}), with a constant heating rate of 15 K min^{-1} , using standard 100 μL alumina crucibles. The weighed sample amounts were in the range of 3–5 mg, respectively. The measurements were taken in the temperature range 30–1000 $^{\circ}\text{C}$, and the reference crucible was empty (no ref. material used). The samples were analyzed “as received”. The results were processed using the thermoanalyzer’s Calisto Processing (v1.36) software. From every measurement, a previously recorded baseline was subtracted. In some cases on the heat flow curves, a Savitzky and Golay smoothing filter (number of averaged points 75–100) was applied, in order to reduce the baseline noise. In the above-mentioned smoothing, the “Peak” filter was used in order to better preserve the shape of the signal peaks. Parallel with the TG DSC measurement, the analysis of the evolved gases/decomposition products were carried out on a Pfeiffer Vacuum OmniStarTM Gas Analysis System coupled to the above-described TG. The gas splitters and transfer lines to the spectrometer were thermostated at $290\text{ }^{\circ}\text{C}$. The measurements were taken in SEM Bargraph Cycles acquisition mode, in which the total ion current (TIC), the analog bar graph spectra (for structure determination), and the separate ion current of each scanned individual mass (115 masses) was recorded. The scanned mass interval was 5–120 amu, with a scan speed of 50 ms amu^{-1} , and the spectrometer was operated in electron impact mode.

Results and discussion

Nanoparticles grown in the presence of 1-hexadecyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide or *N*-octylfurfuryl-imine: synthesis and spectral characterizations

The dropwise addition of a solution of TTF in $[\text{HDMIM}][(\text{CF}_3\text{SO}_2)_2\text{N}]/\text{acetonitrile}$ (1:3 molar eq. of $\text{HDMIM}^+/\text{TTF}$) into an acetonitrile solution of TCNQ at room temperature led to a black precipitate of $\text{TTF}\cdot\text{TCNQ}$. Transmission electron micrographs exclusively evidenced sticks (Fig. 3). Moreover, $\text{TTF}[\text{Ni}(\text{dmit})_2]_2$ as sticks was obtained by the dropwise addition of an acetone solution of $[(n\text{-C}_4\text{H}_9)_4\text{N}][\text{Ni}(\text{dmit})_2]$ on $(\text{TTF})_3(\text{BF}_4)_2$ dispersed in $[\text{HDMIM}][(\text{CF}_3\text{SO}_2)_2\text{N}]/\text{acetonitrile}$ (1:3 molar eq. of $\text{HDMIM}^+/\text{TTF}$, Fig. 3). Thus, at room temperature and for molar ratios explored, $[\text{HDMIM}][(\text{CF}_3\text{SO}_2)_2\text{N}]$ did not play its role of growth-controlling agent as it was the case for $[\text{BMIM}][\text{BF}_4]$ or $[\text{BMIM}][(\text{CF}_3\text{SO}_2)_2\text{N}]$ [9–11]. When the same reactions were carried out at $-50\text{ }^{\circ}\text{C}$, $\text{TTF}\cdot\text{TCNQ}$

Fig. 3 Electron micrographs for molecule based conductors grown in the presence of [HDMIM][(CF₃SO₂)₂N] (3 molar eq.): TTF·TCNQ (a 25 °C; c 50 °C) and TTF[Ni(dmit)₂]₂ (b 25 °C; d 50 °C)



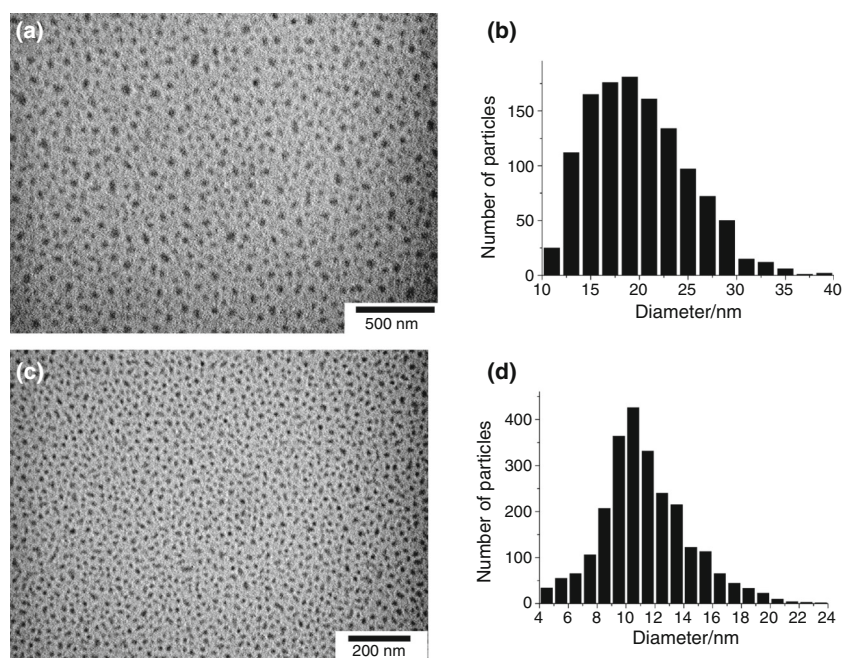
and TTF[Ni(dmit)₂]₂ were grown as nanoparticles (Fig. 3). For TTF·TCNQ, transmission electron micrographs showed a mixture of roughly spherical nanoparticles (diameters in the 15–40 nm range) and elongated nanoparticles (20–30 nm × 50–100 nm). For TTF[Ni(dmit)₂]₂, nanoparticles exhibited diameters in the 10–30 nm range but were frequently agglomerated. At low temperatures, i.e., –50 °C, the bulky [HDMIM][(CF₃SO₂)₂N] species could control the growth of the two molecular conductors as nano-objects, but these results were not quite satisfactory in terms of size, morphology, and state of dispersion. Indeed, as specified in the introduction, better results were obtained in the presence of the butyl-substituted imidazolium salt, i.e., [BMIM][(CF₃SO₂)₂N] [10, 11].

Metal dithiolene complexes have recently been evaluated as active components for the fabrication of organic thermoelectric generators [20]. Performances of these generators were expected to be enhanced with the active matter at the nanoscale and finely dispersed. It is the reason why we have evaluated the use of *N*-octylfurfuryl-imine for the growth of TTF[Ni(dmit)₂]₂ small particles either by a chemical or by an electrochemical route. The chemical route was similar to that described above, using (TTF)₃(BF₄)₂ and [(*n*-C₄H₉)₄N][Ni(dmit)₂] and simply replacing [HDMIM][(CF₃SO₂)₂N] by *N*-octylfurfuryl-imine (1–3 molar eq. of *N*-octylfurfuryl-imine/TTF). Transmission electron micrographs of the TTF[Ni(dmit)₂]₂ powder thus obtained evidenced well-dispersed nanoparticles with a mean diameter of about 20 nm (Fig. 4). These particles exhibited diameters similar to those obtained with [HDMIM][(CF₃SO₂)₂N] (see above). However, when using *N*-octylfurfuryl-imine, no agglomerates of particles were

observed. We have also performed the galvanostatic oxidation of TTF in the presence of [(*n*-C₄H₉)₄N][Ni(dmit)₂] (both reactant and supporting electrolyte) and *N*-octylfurfuryl-imine (3 molar eq. of *N*-octylfurfuryl-imine/TTF) under vigorous stirring. After a few hours, the platinum anode was covered by a black powder of TTF[Ni(dmit)₂]₂. The electrocrystallization was run over a period of 2 days to obtain a larger amount of product. Transmission electron micrographs showed very well-dispersed nanoparticles exhibiting a mean diameter of about 12 nm (Fig. 4). This result was the most satisfactory in terms of morphology (spherical particles), state of dispersion (no agglomerates), and smallness of the particles (the smallest ever published for TTF[Ni(dmit)₂]₂). The neutral amphiphilic molecule, *N*-octylfurfuryl-imine, allowed the preparation of small and well-dispersed spherical nanoparticles presumably due to an efficient π stacking with five-membered rings of both TTF and Ni(dmit)₂, the octyl chain attached to the nitrogen atom external to the cycle preventing the nanoparticles from agglomeration.

TTF·TCNQ and TTF[Ni(dmit)₂]₂ nanoparticles have been characterized by vibrational spectroscopy, i.e., infrared and Raman. Whatever the growth-controlling agent used and its molar amount, spectra did not evidence signals for [HDMIM][(CF₃SO₂)₂N] or *N*-octylfurfuryl-imine. Its presence in solution was essential to control the growth of the molecular material as nanoparticles but it was not adsorbed to the particles surface in the final material. Infrared and Raman spectra evidenced the presence of both the electron donor (TTF) and the electron acceptor (TCNQ or Ni(dmit)₂) within the nanopowders. Vibration spectra for TTF·TCNQ nanoparticles were shown in Fig. 5. They

Fig. 4 Electron micrographs and size histograms for TTF[Ni(dmit)₂]₂ grown in the presence of *N* octylfurfuryl imine (3 molar eq.): chemical route (a, b); electrochemical route (c, d)



were very similar to those previously described for TTF·TCNQ nanoparticles grown in the presence of [BMIM][BF₄] [9]. In particular, in the infrared spectrum, the nitrile stretching mode (ν_{CN}) at 2204 cm^{-1} allowed us to determine the amount of charge transfer from the TTF donor molecule to the acceptor TCNQ molecule. Using the linear correlation of ν_{CN} for TCNQ as a function of the degree of charge transfer, we obtained a value of 0.56, in relatively good agreement with that for single crystals, i.e., 0.59 [21]. Moreover, in the Raman spectrum, the C=C stretching mode in TCNQ located at 1418 cm^{-1} ($\nu_4 a_g$) gave a charge transfer of 0.55 [22]. Thus, according to infrared and Raman studies, the charge transfer in TTF·TCNQ as nanoparticles is rather similar to that on macroscopic single crystals.

Vibration spectra for TTF[Ni(dmit)₂]₂ nanoparticles were shown in Fig. 5. They were similar to those described for TTF[Ni(dmit)₂]₂ nanoparticles elaborated in the presence of [BMIM][(CF₃SO₂)₂N] [11] or TTF[Ni(dmit)₂]₂ nanowires electrodeposited on (001)-oriented silicon substrates [23, 24]. The infrared spectrum showed the CH ethylenic stretching vibration for TTF at 3087 cm^{-1} , and the characteristic doublet present in all compounds containing the M(dmit)₂ species (1071 and 1053 cm^{-1}) [25]. The Raman spectrum evidenced modes related to the Ni(dmit)₂ entity: 137, 343, 361, and 492 cm^{-1} [26]. Signals at 1430, 1471, and 1509 cm^{-1} were related to the TTF entity [27]. The more intense signal located at 1430 cm^{-1} was very sensitive to the amount of charge transfer between the electron donor TTF and the electron acceptor Ni(dmit)₂. From this Raman peak, the degree of charge

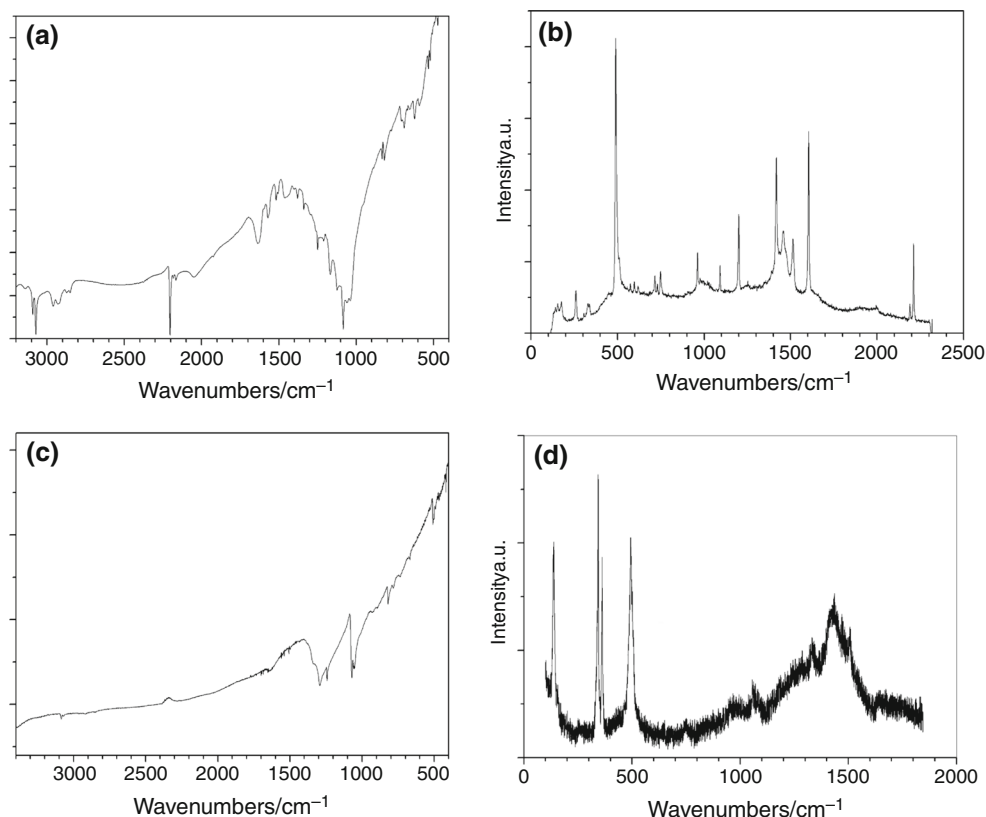
transfer was evaluated to 0.86 [28], in relatively good agreement with that obtained from band structure calculations, i.e., 0.80 [3].

Room temperature conductivity measurements on TTF·TCNQ and TTF[Ni(dmit)₂]₂ nanoparticle powders gave values in the 0.1 S cm^{-1} range. They were of the same order of magnitude than those reported for TTF·TCNQ and TTF[Ni(dmit)₂]₂ nanoparticle powders prepared in the presence of [BMIM][BF₄] or [BMIM][(CF₃SO₂)₂N] [9–11]. They were obviously lower than those on single crystals [23] due to resistive boundaries between the particles.

Nanoparticles grown in the presence of 1-hexadecyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide or *N*-octylfurfuryl-imine: thermoanalytical studies

Besides its wide application area, thermal analysis is an indispensable tool for the evaluation of thermal stability, degradation of organic molecular materials too. Muraoka et al. [29] have investigated the temperature dependence of the ac heat capacity of a molecule-based superconductor κ -(BEDT-TTF)₂Cu(NCS)₂ under external pressures and with magnetic fields, while Ishikawa et al. [30] performed thermodynamic investigations by relaxation calorimetric technique on the organic superconductor κ -(BEDT-TTF)₂Ag(CN)₂H₂O. Bhattacharjee et al. [31] have used thermogravimetric analysis to study the thermal degradation of a bimetallic oxalate ligand based molecular magnetic material $\{N(n\text{-C}_4\text{H}_9)_4[\text{Fe}^{\text{II}}\text{Fe}^{\text{III}}(\text{C}_2\text{O}_4)_3]\}_\infty$. They

Fig. 5 Infrared and Raman spectra for TTF·TCNQ (a, b) and for TTF[Ni(dmit)₂]₂ nanoparticles (c, d)



have found that the decomposition takes place in multiple steps, the chemical products and reaction pathways were established using TG measurement and supplemented by the IR and powder XRD studies. In our study, we have also used simultaneous thermogravimetry differential scanning calorimetry (TG DSC) and mass spectrometric evolved gas analysis (MS EGA) for the determination of decomposition intervals, steps and volatile degradation products on TTF·TCNQ and TTF[Ni(dmit)₂]₂ molecule-based conductors. In view of the forthcoming integration of these nanomaterials in future thermoelectric generators, it is of key interest to study their thermoanalytical properties. This has been performed for both TTF·TCNQ and TTF[Ni(dmit)₂]₂ nanoparticles, and these latter being the most promising for thermoelectric applications as mentioned above.

As shown in Fig. 6, the results of thermoanalytical investigations for TTF·TCNQ sticks grown at 25 °C are depicted. The green curve is the mass variation (TG), and the blue curve is the heat flow, while the violet curve corresponds to the derivative thermogravimetric curve (DTG). On the TG and DTG curves there can be seen that the sample decomposes in five more or less differentiable steps (the limits of these steps are not shown in the Fig. 6). The first step is between 64.3 and 184.9 °C, with a mass loss of 4.4 %. This is due to the evaporation of the

physically adsorbed water (obtained from the mass spectrometric evolved gas analysis measurements). The next mass loss step of 9.9 % is more evident and can be found in the temperature region of 187.3 and 274 °C; this appears as a sharp peak on the DTG curve. This is also accompanied by a sharp exothermic peak on the heat flow curve (onset point 232.9 °C, peak maxima 235.8 °C and heat of decomposition 160 J g⁻¹), which is due to a fast degradation reaction. Between 273.9 and 437.6 °C, an additional amount of 13.2 %, while in the 440.5 and 709.4 °C region, 16.7 % is lost. In the last step, the speed of the mass loss (on the DTG curve) decreases between 712.5 and 1013.7 °C 13.2 % of the starting material is lost. The total mass decrease in the investigated temperature region is 57.97 %.

As shown in Fig. 7, the results of thermoanalytical investigations for TTF·TCNQ nanoparticles grown at -50 °C are presented. By comparing the shape and run of all three curves with the ones presented on Fig. 6, it can be seen that they are very similar. Even the mass loss intervals determined in the previous measurement are in good accordance. The temperature intervals and the corresponding mass losses are the following: step 1 74.7 and 184.7 °C, mass loss 3 %, step 2 189.2 and 260.2 °C, mass loss 7.16 %, step 3 262.6 and 432 °C, mass loss 13.7 %, step 4 435.7 and 684.4 °C, mass loss 13.6 %, step 5 712.5 and 1013.7 °C, mass loss 13.2 %.

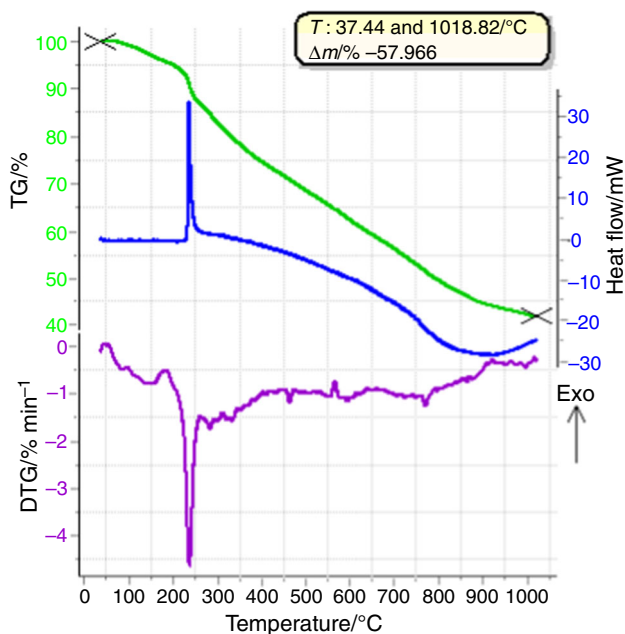


Fig. 6 TG DTG DSC trace of TTF·TCNQ molecule based conductors grown in the presence of [HDMIM][(CF₃SO₂)₂N] (3 molar eq.) at 25 °C

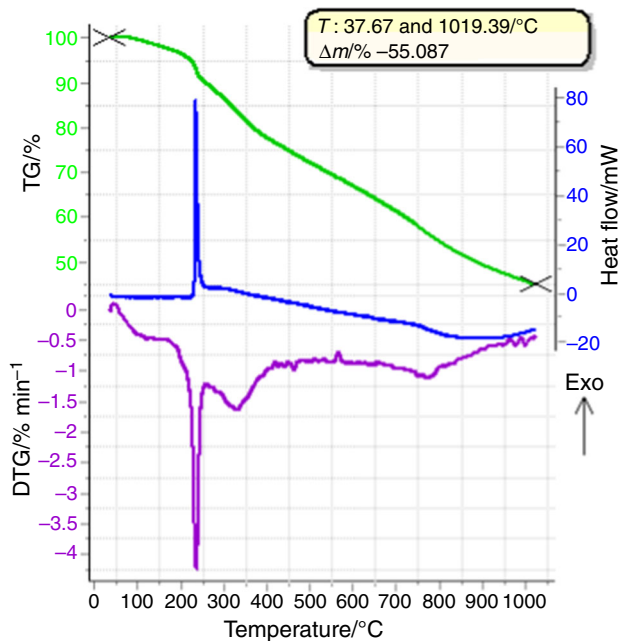


Fig. 7 TG DTG DSC trace of TTF·TCNQ molecule based conductors grown in the presence of [HDMIM][(CF₃SO₂)₂N] (3 molar eq.) at 50 °C

and step 5 687.2 and 1017.4 °C, mass loss 16.9 %. In this case, the total mass loss is 55.1 %. The parameters of the exothermic peak, which accompanies the mass loss in the second step, are the following: onset point 233.4 °C, peak maxima 234.2 °C and heat of decomposition 166.2 J g⁻¹. Due to the similarity of the two set of results (Figs. 6, 7),

we can draw the conclusion that the preparation method and the size and shape of the obtained particles (i.e., sticks and nanopowder) do not influence the thermal behavior of both materials. Moreover, the investigated materials are thermally stable up to 180 190 °C, which is in good agreement with the results of de Caro et al. [15].

Parallel with the thermoanalytical measurement, the analysis of the evolved gases/decomposition products was carried out. As shown in Fig. 8, the ion currents (i.e., concentration variation) of some selected fragments/molecules are plotted against temperature. Between 200 and 250 °C, a sharp rise in the intensity of all curves can be observed. This corresponds to the second mass loss step in Fig. 7. By plotting the analog spectra of the evolved volatiles at 236 °C (Fig. 9), where the speed of mass loss is the highest (this corresponds to the minima of the DTG curve), the formula of various fragments/molecules can be deduced; thus, the products of the thermal degradation can be identified. As shown in Fig. 9, m/z 78 amu corresponds to the C₆H₆ formula, which is one of the degradation products of TCNQ, while to the m/z 76 amu corresponds to the C₆H₄ and/or CS₂ formulae, the former resulting from TCNQ and the latter from the breakage of the TTF. Another very characteristic decomposition products can also be observed, as the CS (m/z 44 amu, daughter ion of CS₂), CH₂CN (m/z 40 amu), HCN, and CN (m/z 27 amu and m/z 26 amu, respectively).

By following the course of the various fragments identified, a rough degradation mechanism can be proposed. In the 200 and 250 °C interval (Fig. 8), the intensity of all fragments/molecules is rising, which means that the

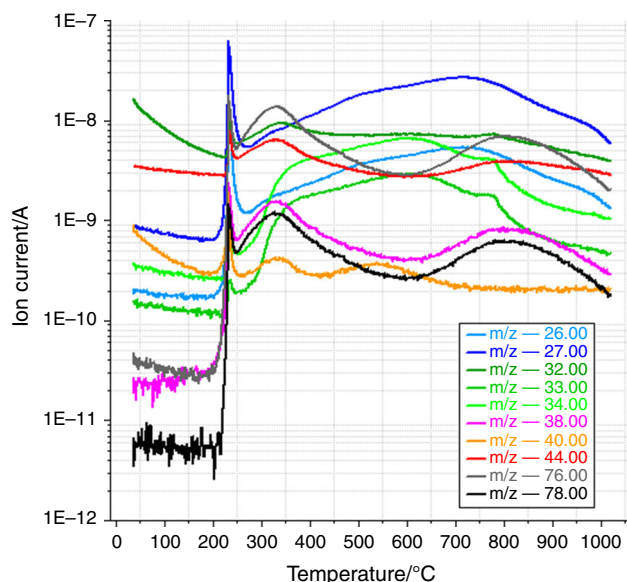
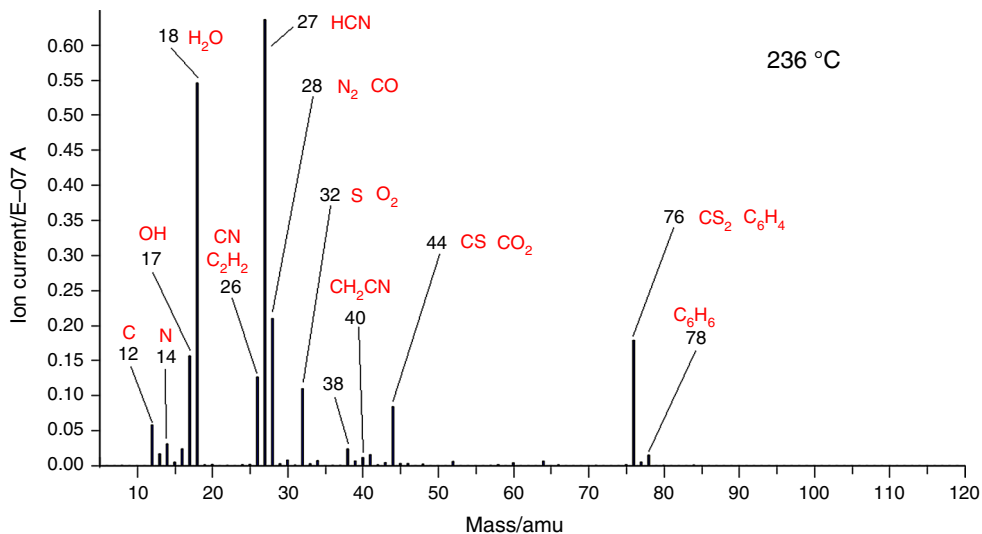


Fig. 8 Ion currents (concentration variation) of some selected fragments plotted against temperature

Fig. 9 Analog mass spectra of the evolved volatiles (at 236 °C) of TTF-TCNQ molecule based conductors grown in the presence of [HDMIM][(CF₃SO₂)₂N] (3 molar eq.) at 50 °C



TTF-TCNQ structure collapsed and in some manner degraded thermally, from now on a continuous degradation occurs, with release of various characteristic degradation products. During the third mass loss step (between 262.6 and 432 °C), the degradation products of both TTF and TCNQ appear, while in the last mass loss step the volatiles resulting from TCNQ are dominating (m/z 78 amu and m/z 76 amu).

The evolution maxima of HCN and CN (m/z 27 amu and m/z 26 amu), resulting from TCNQ, is roughly around 710 °C. The presence of degradation products (e.g., fluorinated compounds), which could result from

[HDMIM][(CF₃SO₂)₂N], were not identified, which agrees with the results obtained from the vibrational spectroscopy measurements.

As shown in Fig. 10, the TG DTG DSC trace of TTF[Ni(dmit)₂]₂ sticks grown at 25 °C are shown, while in Fig. 11 for those particles, which were grown at -50 °C. Again, both materials behave in the same way from thermal point of view. In both cases, based on TG and DTG curves, five mass loss steps can be differentiated, from which the second and third step is well defined. The second, third, and the last step is accompanied by a small, broad endotherm. The corresponding mass losses for each

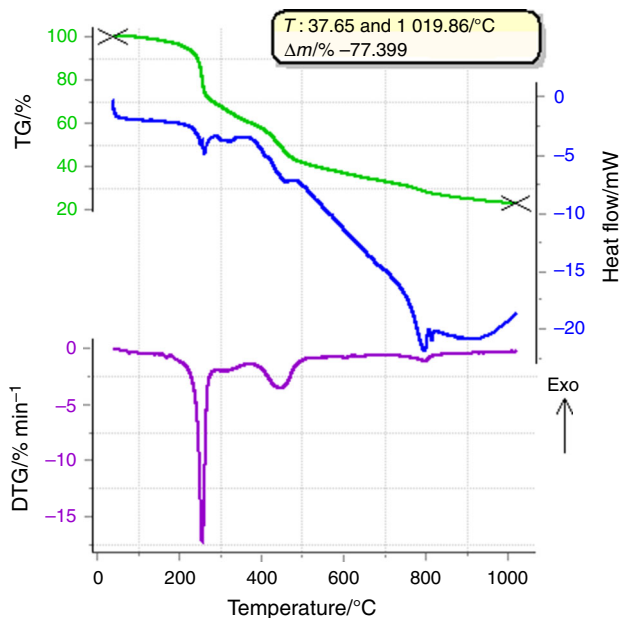


Fig. 10 TG DTG DSC trace of TTF[Ni(dmit)₂]₂ molecule based conductors grown in the presence of [HDMIM][(CF₃SO₂)₂N] (3 molar eq.) at 25 °C

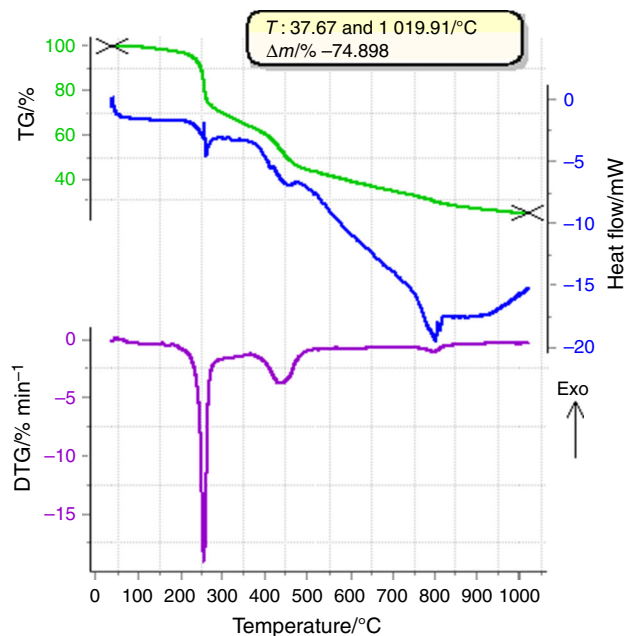
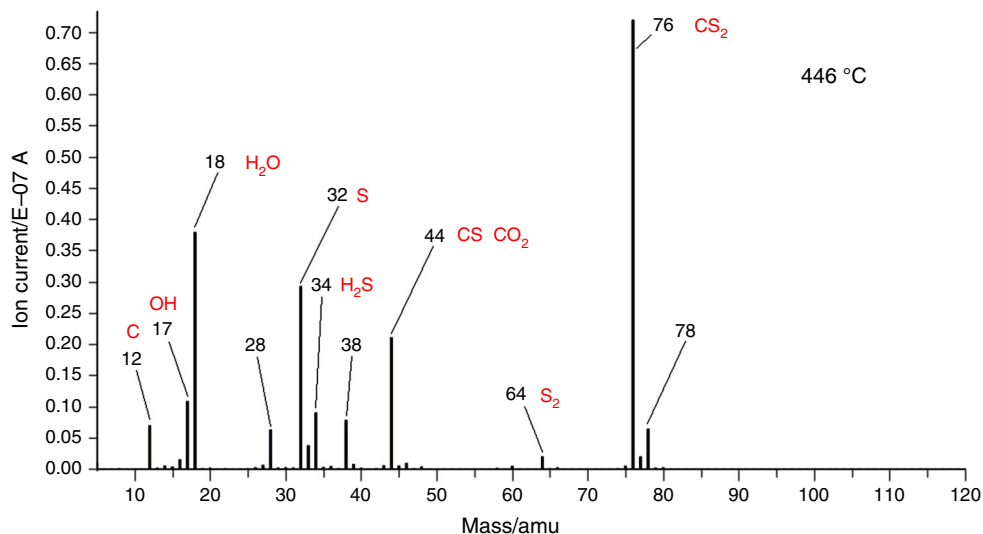


Fig. 11 TG DTG DSC trace of TTF[Ni(dmit)₂]₂ molecule based conductors grown in the presence of [HDMIM][(CF₃SO₂)₂N] (3 molar eq.) at 50 °C

Fig. 12 Analog mass spectra of the evolved volatiles (at 446 °C) of TTF[Ni(dmit)₂]₂ molecule based conductors grown in the presence of [HDMIM][[(CF₃SO₂)₂N] (3 molar eq.) at 50 °C



step are listed below: in Fig. 10, step 1 74.2 and 194.1 °C, mass loss 2.7 %, step 2 198.9 and 330.7 °C, mass loss 30.3 %, step 3 331.9 and 528.9 °C, mass loss 23.9 %, step 4 530.5 and 735.1 °C, mass loss 9.1 %, and step 5 738 and 1014.2 °C, mass loss 8.4 %. The total mass loss is 74.9 %.

From Fig. 11, step 1 61.4 and 195.3 °C, mass loss 3.8 %, step 2 198.2 and 332.3 °C, mass loss 32.4 %, step 3 334.6 and 529.7 °C, mass loss 23.6 %, step 4 532.8 and 730.1 °C, mass loss 8.1 % and step 5 732.7 and 1016.7 °C, mass loss 8.9 %. The total mass loss is 77.4 %.

It can be concluded that the major part of the sample is lost in the second and third step (approx. 55 %).

By evaluating the mass spectrometric evolved gas analysis results of the gaseous products released at 446 °C (Fig. 12), one can see that the major decomposition products are various sulfur compounds, such as carbonyl sulfide CS₂ (m/z 76 amu, parent ion and base peak), hydrogen sulfide H₂S (m/z 34 amu) and diatomic sulfur molecule (m/z 64 amu).

By following the course of the various fragments identified (Fig. 13), it can be stated that the slow decomposition of the sample begins at quite low temperatures (125 °C), with evolution of CS₂ (m/z 76 amu), so the TTF[Ni(dmit)₂]₂ is less stable thermally. Between 290 and 390 °C, a strong evolution of H₂S (m/z 34 amu) takes place.

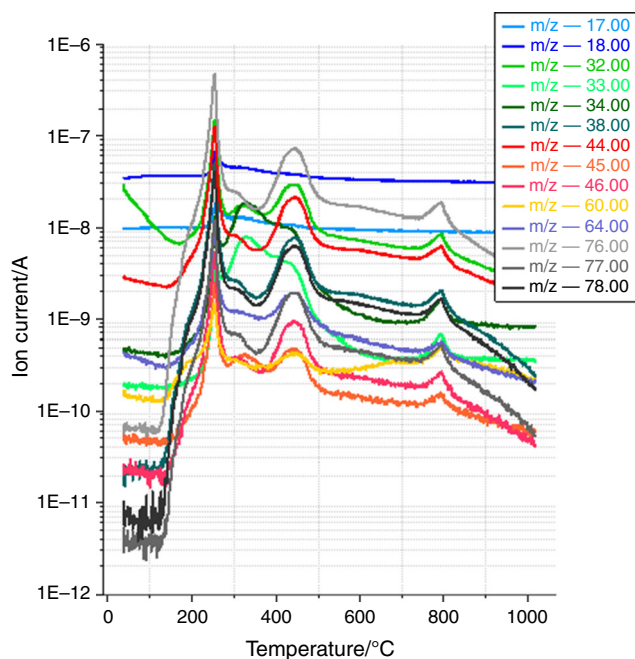


Fig. 13 Ion currents (concentration variation) of some selected fragments plotted against temperature

Conclusions

In conclusion, we have prepared nanoparticles of the two well-known molecular conductors, TTF·TCNQ and TTF[Ni(dmit)₂]₂. Working in long-chain alkyl imidazolium salts/acetonitrile or acetone mixtures at low temperatures, particles of sizes in the 10–50 nm range were grown. However, the more interesting results in terms of state of dispersion and size control were obtained in the presence of the neutral amphiphilic molecule, namely *N*-octylfurfuryl-imine. Spherical nanoparticles as small as 5–15 nm were grown. This is remarkable for quasi-one-dimensional systems whose natural tendency is to grow as elongated needles or wires. According to spectral, thermoanalytical, and conductivity data, nanoparticle powders had a similar behavior than those for bulk materials. We are currently investigating the use of other solvents and

liquid surfactants (in various relative amounts) to grow and organize the conducting nanoparticles on conversion coatings for instance. In another hand, we consider the synthesis of composite materials by dispersing the nanoparticles in a polymer matrix. The resulting materials will undergo an integration in thermoelectric generators in order to assess their ability to convert heat energy into electrical energy.

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