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# 8-hydroxyquinoline-modified clay incorporated in an epoxy coating for the corrosion protection of carbon steel

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## ABSTRACT

In the present work, a well-known corrosion inhibitor (8-hydroxyquinoline (8HQ)) was inserted within the montmorillonite platelets (8HQ-MMT) and the modified clay was incorporated (3 wt.%) into a solvent-free epoxy coating which was afterwards deposited on carbon steel substrates. First, the inhibitive action of 8HQ was investigated by electrochemical methods carried out on a bare carbon steel rotating disk electrode in a 0.1 M NaCl solution. Then, electrochemical impedance measurements were performed to assess the effect of the 8HQ-MMT in the epoxy coating for the corrosion protection. The results were compared with a reference sample constituted by the epoxy coating containing an ammonium quaternary salt-modified clay. It was shown that the two coatings presented good barrier properties. Dry and wet adherence measurements revealed an improvement of the adherence when the 8HQ-MMT was incorporated into the coating by comparison with the reference sample. It was concluded that the 8HQ mainly had an effect at the metal/coating interface but its concentration was too low to afford significant corrosion protection of the carbon steel.

## 1. Introduction

Since the beginning of the 2000s, clay based composite coatings have attracted considerable attention as effective and environmental friendly materials [1–6]. The clays can be incorporated into polymer matrices including epoxy or polyurethane and they offer a significant improvement in various properties, *i.e.* mechanical [7–10], thermal stability [11,12], flame retardant [13] and oxygen and moisture barrier [14–16] in comparison with bulk polymer. Among the clay family, sodium montmorillonite (MMT) is commonly applied in polymer composites. Their specific layered structure makes intercalation/exfoliation of the clay possible and leads to a high cation exchange capacity and to a high surface area, explaining the use of clays as reinforcing fillers for polymers [17–19]. However, due to their hydrophilic property, the clays dispersion in an organic phase is generally limited and the MMT platelets tend to aggregate; hence, to obtain both a good compatibility with the organic matrix and the expansion of the stacked layers, the clay is often modified with hydrophobic compounds, such as alkyl (or benzyl) ammonium cations, which can be easily exchanged with organic onium ions on the gallery surfaces [8]. Other treatments, *i.e.* phosphonium surfactants, organo silane,

hyperbranched polymers have been developed in academic researches for various applications of polymer nanocomposites [20–23].

For the corrosion protection, MMT was also used as nano containers to introduce inhibitors into the polymer matrix. The aim is to avoid inhibitors interaction with the matrix, particularly for organic compounds. Recently, Ghazi et al. have intercalated benzimidazole and zinc cations in the MMT galleries for the corrosion protection of mild steel by an epoxy ester coating. The modified clay gave a promising active properties for the corrosion protection by the formation of insoluble complexes in artificial defects [24]. Sari et al. used MMT and graphene oxide nanoclays mixtures. The clay functionalization with aminosilane and 1,4-butanediol diglycidyl ether significantly improved the protection performance of an epoxy coating [25]. In previous studies, our group have also used MMT to incorporate organic inhibitors, such as amino trimethylphosphonic acid or indole 3 butyric acid, into a solvent epoxy resin for corrosion protection of carbon steels [26–28]. The presence of modified MMT enhanced the barrier properties of the epoxy resin and improved the corrosion protection thanks to the presence of the organic molecules at the metal/coating interface.

In the present work, a natural MMT was modified by 8 hydroxyquinoline (8HQ) and then, the modified MMT (8HQ MMT) was

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incorporated in an epoxy resin for the corrosion protection of a carbon steel. By comparison with our previous studies, the solvent epoxy resin was replaced by a solvent free system [26–28]. 8HQ is known for its chelating properties on different metals and as corrosion inhibitor of copper [29], of magnesium [30], of aluminium [31,32] and of carbon steel [33]. An alkyl quaternary ammonium salt modified MMT was compared with the 8HQ MMT. This organo modified clay is well known to have good dispersion, improving the barrier properties of organic coatings [34,35]. The comparison between the two coatings allowed the barrier effect, due to the modified MMT dispersion into the epoxy matrix, to be differentiated from the role of the inhibitor at the metal/coating interface.

## 2. Experimental

### 2.1. Materials

8 hydroxyquinoline (8HQ), was purchased from Alfa Aesar (purity = 99%) and was used as received. In neutral aqueous solution, the 8HQ concentration was chosen near its solubility limit ( $3 \times 10^{-3}$  M). A  $\text{Fe}(\text{8HQ})_2$  chelate was prepared according to the following protocol [36]: the 8HQ, at its solubility limit, was added to a  $\text{Fe}_2\text{SO}_4$  solution (5%) and the solution was stirred until a precipitate was formed. Then, the precipitate was filtered, rinsed with water and ethanol and dried.

Natural montmorillonite (Cloisite® 116) and modified montmorillonite (Cloisite® 30B) were supplied by BYK Southern Clay product. The modified montmorillonite has been treated with an alkyl quaternary ammonium salt. Cloisite® 116 and Cloisite® 30B were henceforth called “MMT” and “AA MMT”, respectively.

To characterize the inhibitive efficiency of 8HQ, a XC35 carbon steel rod of  $1 \text{ cm}^2$  cross sectional area was used as working electrode. Its composition in percent weight was  $C = 0.35$ ,  $Mn = 0.65$ ,  $Si = 0.25$ ,  $P = 0.035$ ,  $S = 0.035$  and  $Fe$  to 100. A heat shrinkable sheath left only the tip of the carbon steel cylinder in contact with the solution. The carbon steel samples were abraded with successive SiC papers down to grade 1200 and cleaned in ethanol in an ultrasonic bath and finally dried in warm air. For the coatings, carbon steel sheets ( $150 \times 100 \times 2 \text{ mm}$ ) were used as substrates. They were polished with abrasive papers from 80 to 600 grade and cleaned with ethanol.

### 2.2. Modification of the clay by the 8HQ (8HQ MMT)

For  $\text{pH} < 5$ , the 8HQ molecules are protonated ( $\text{C}_9\text{H}_7\text{ONH}^+$ ). As the MMT is a cationic exchanger, the  $\text{Na}^+$  and/or the  $\text{Ca}^{2+}$  ions in the MMT interlayers can be replaced by the protonated 8HQ. 8HQ MMT was prepared according to the following procedure [26–28]: the 8HQ (7.0 g) was dissolved in a HCl solution ( $\text{pH} = 3$ ). Then, this solution was added drop wise in a suspension of MMT (3.0 g) in distilled water and the mixture was stirred for 24 h at  $70^\circ\text{C}$ . The obtained 8HQ MMT (green precipitate) was collected by centrifugation (speed of 6000 rpm) and washed with water until no chloride was detected in the filtrate using a 0.1 M  $\text{AgNO}_3$  solution. Finally, the 8HQ MMT was dried at  $80^\circ\text{C}$  in a vacuum oven for 48 h.

### 2.3. Coating samples

The coating was a solvent free epoxy resin prepared from diglycidylbisphenol A (Epon 828, from Hexion) as base and a low viscosity modified cycloaliphatic polyamine (Ancamine 2735, from Air Products and Chemicals) as hardener. The AA MMT and the 8HQ MMT were incorporated into the epoxy coating at a concentration of 3 wt.%. In previous works, it was shown that the best corrosion protection was obtained for low clay contents ( $\leq 5\%$ ) [26,37]. The modified MMT were progressively added into the hardener by ultrasonification. Then, a magnetic stirring was performed for 24 h to obtain a homogeneous

mixture. The pre polymer mixtures were applied on the XC35 carbon steel sheets by spin coating with a speed of 600 rpm during 90 s. After drying at room temperature for 24 h, the coatings were  $30 \pm 1 \mu\text{m}$  thick (measured by Minitest 600 Erichen digital meter).

### 2.4. Analytical characterizations

The crystalline structure of the MMT, the AA MMT and the 8HQ MMT were characterized by x ray diffraction (XRD, Siemens diffractometer D5000) with  $\text{Cu K}\alpha$  radiation ( $\lambda = 1.54 \text{ \AA}$ ).

FTIR spectra of the 8HQ, the 8HQ Fe chelate and the black products removed from the carbon steel electrode surface after 20 h of immersion in a 0.1 M NaCl solution containing 8HQ were recorded using KBr pellets with a Nexus 670 Nicolet spectrometer equipped with DTGS KBr detector (resolution of  $4 \text{ cm}^{-1}$  in the  $400 \text{ cm}^{-1}$ – $4000 \text{ cm}^{-1}$  region).

Thermogravimetric analysis (TGA) was used to determine thermo degradation of the 8HQ, of the MMT and of the 8HQ MMT. The weight loss curves were obtained using a Setaram thermogravimetric analyser. Experiments were performed in air atmosphere over a temperature range from  $20^\circ\text{C}$  to  $1000^\circ\text{C}$  with a  $10^\circ\text{C}/\text{min}$  heating ramp.

The leaching experiments of 8HQ from 8HQ MMT were performed in a 0.5 M NaCl solution. 0.05 g of 8HQ MMT was introduced in 100 mL of 0.5 M NaCl solution. The suspension was stirred all along the duration test. A few millilitres of the suspension were periodically removed (from 1 h to 24 h) and filtered. The released 8HQ in the NaCl solution was determined by UV vis spectroscopy using a GBC Cintra 40 spectrometer at  $\lambda_{\text{max}} = 240 \text{ nm}$ . Calibration curve was first obtained for a series of standard solutions (from 1 mg/L to 10 mg/L of 8HQ). Release experiments were performed in triplicate.

### 2.5. Electrochemical characterizations

First, the inhibitive efficiency of the 8HQ was determined by classical electrochemical measurements with a standard three electrode cell, in which the rod of XC35 carbon steel constituted the working electrode. A platinum grid and a saturated calomel electrode (SCE) were used as counter electrode and reference electrode, respectively. Polarisation curves were obtained using a Solartron 1287 electrochemical interface. Anodic and cathodic polarisation curves were carried out at a scan rate of  $1 \text{ mV s}^{-1}$ . Impedance measurements were performed using a Solartron 1287 electrochemical interface connected to a Solartron 1250 frequency response analyser. Impedance diagrams were obtained under potentiostatic regulation, at the corrosion potential, over a frequency range of 65 kHz–10 mHz with 8 points per decade, using a 10 mV peak to peak sinusoidal voltage.

The performance of the coatings was assessed by electrochemical impedance spectroscopy (EIS) in a conventional three electrode cell, realized by fixing a cylindrical Plexiglas tube on top of the coated sample, and filled with a 0.5 M NaCl solution. The surface area of the working electrode was of  $28 \text{ cm}^2$ . An SCE and a Pt sheet were used as reference and counter electrode, respectively. A VMP3 BioLogic (Science Instruments) was used to measure the impedance of the coated samples. Measurements were performed under potentiostatic conditions, at the corrosion potential ( $E_{\text{corr}}$ ) over a frequency range of 100 kHz–10 mHz with six points per decade using 30 mV peak to peak sinusoidal voltage. Electrochemical measurements were obtained in triplicate.

### 2.6. Adherence and salt spray tests

Pull off test was performed for the dry coatings by Positest Automatic Adhesion Tester following ASTM D 4541. Cross cut adherence test was carried out according to ASTM D3359 method B. The adherence was evaluated using cross hatch cutter (6 blades 1 mm width). A tape with an adhesive strength of 9.5 N per 25 mm (ISO 2409) was immediately applied to the cross cut coating surface. The tape was

removed by a quick pull. The cross cut test was performed before and after 6 h of immersion in distilled water after carefully removing the water excess with tissue paper.

Salt spray test was performed according to ASTM B117 using a Q FOGCCT 600 chamber. Scribes were made on the samples before exposure to the salt spray. The specimens were placed in the salt spray chamber (angle of 20°). A solution of 5% NaCl was sprayed on the samples ( $T = 35\text{ }^{\circ}\text{C}$ ). The spraying was maintained for the duration of the test. The samples were observed at the end of the test (240 h).

### 3. Results

First, the inhibitive efficiency of the 8HQ was assessed from the electrochemical measurements. Then, the modifications of the MMT by the 8HQ were investigated by XRD and TGA measurements. The protective properties of the solvent free epoxy coating containing the AA MMT or the 8HQ MMT were monitored by EIS measurements during exposure to the aggressive solution. The dielectric constant of the coatings ( $\epsilon_w$ ) in wet conditions were determined by using the complex capacitance plots. Finally, adherence and salt spray tests were realised.

#### 3.1. Efficiency of the 8HQ in aqueous solution

First, the electrode surface was observed by optical microscopy before immersion and after 20 h of immersion in the electrolyte with and without 8HQ. The photographs are shown in Fig. 1. In Fig. 1b and c, the shape of hydrodynamic patterns are observed due to the electrode rotation rate. Without inhibitor (Fig. 1b), the electrode surface was covered by brown corrosion products. Similar photographs, taken *in situ*, for the corrosion of pure iron in neutral chloride solutions, have shown that with increasing immersion time (20 h), the corrosion led to the formation of porous layers which progressively decreased the active surface area of the bare metal [38]. You et al. have demonstrated, by an in depth analysis of electrochemical and electrohydrodynamical impedance measurements that the cathodic oxygen reduction occurs through two diffusion processes: one was essentially representative of the mass transport in liquid phase and the second one corresponded to the diffusion across the porous layers of corrosion products. They concluded that the degree of coupling between them was dependent on the immersion time and on the electrode rotation rate [38]. In Fig. 1c, black powdered products are present on the carbon steel surface after immersion in the solution containing the 8HQ. The black products were removed from the steel surface and analysed by FTIR spectroscopy. The FTIR spectrum of the black products was compared to those of the 8HQ and of the synthesized chelate  $\text{Fe}(\text{8HQ})_2$  (Fig. 2). The characteristic bands of the spectra are reported in Table 1. For the synthesized chelate and the black products, the spectra present the characteristic bands of the 8HQ molecule: C–O, C–N and C = N at  $1107\text{ cm}^{-1}$ ,  $1277\text{ cm}^{-1}$  and  $1495\text{ cm}^{-1}$ , respectively but the bands are shifted to lower wave number by comparison with the 8HQ spectrum. This shift can be attributed to the formation of the Fe–O bond between ferrous ions and the 8HQ molecule (Fig. 3) which is confirmed by the appearance of a weak band at  $523\text{ cm}^{-1}$  for the chelate and the black products on the FTIR spectra [39,40]. Thus, it can be concluded that in the NaCl solution in the presence of the 8HQ, a chelate was formed on the carbon steel surface by reaction between the 8HQ and the  $\text{Fe}^{2+}$  ions.

The polarisation curves of the carbon steel, obtained with and without 8HQ after 20 h of immersion at the corrosion potential ( $E_{\text{corr}}$ ), are presented in Fig. 4. In the presence of 8HQ, a slight shift of  $E_{\text{corr}}$  towards more positive value and a decrease of both the anodic and cathodic current densities are observed. The polarisation curves indicate that 8HQ is a mixed inhibitor acting both on the anodic and cathodic processes.

Fig. 5 reports the impedance diagrams (Nyquist and Bode coordinates) obtained after three immersion times at  $E_{\text{corr}}$  with and without inhibitor. Without inhibitor, the impedance values are low and

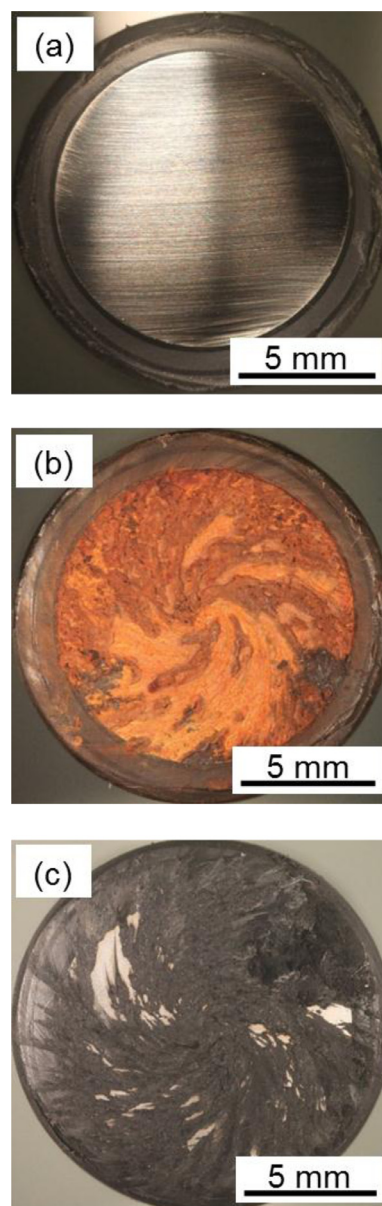


Fig. 1. Optical micrographs of the carbon steel surface: (a) before immersion and after 20 h of immersion in a 0.1 M NaCl solution: (b) without inhibitor and (c) with 8HQ ( $3 \times 10^{-3}\text{ M}$ ). Electrode rotation rate of 500 rpm.

the diagrams are little modified during the test. In contrast, in the presence of the 8HQ, the impedance increases when the immersion time increases from 2 h to 20 h, probably due to a higher surface coverage by the 8HQ Fe chelate. In addition, it can be seen that the shape of the diagrams is progressively modified and a second time constant appears in the high medium frequency range. The presence of this time constant might be attributed to the progressive formation of the 8HQ Fe chelate on the steel surface and can explain the anodic and cathodic character of the 8HQ (Fig. 4). In the present work, equivalent circuits were not used to extract impedance parameters because diffusion processes cannot be simply modelled by resistances and capacitances. For the sake of clarity, the polarisation resistance values ( $R_p$ ) were graphically determined by extrapolation on the real axis at low frequency (Nyquist plots) and used to evaluate the inhibitor efficiency (IE) according to the following equation:

$$\text{IE} = \frac{R_p^i - R_p^w}{R_p^i} \times 100 \quad (1)$$



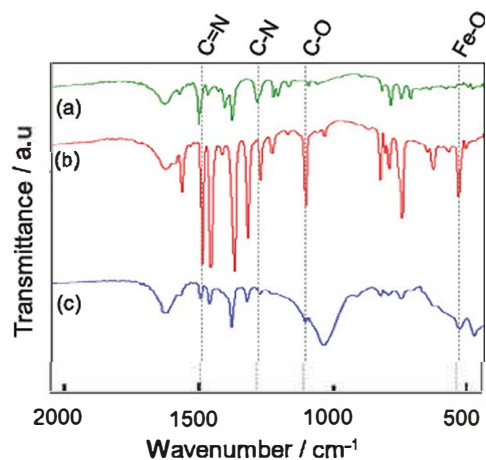


Fig. 2. FTIR spectra of: (a) 8HQ, (b) synthesized 8HQ-Fe chelate and (c) black products removed from the carbon steel surface.

Table 1

Characteristic bands of FTIR spectra obtained for 8HQ, synthesized 8HQ-Fe chelate and black products taken from the carbon steel electrode surface.

Bond	8HQ (cm <sup>-1</sup> )	8HQ-Fe chelate (cm <sup>-1</sup> )	Black products on the electrode (cm <sup>-1</sup> )
Fe O		523	523
C O	1097	1107	1108
C N	1288	1277	1277
C=N	1508	1495	1496

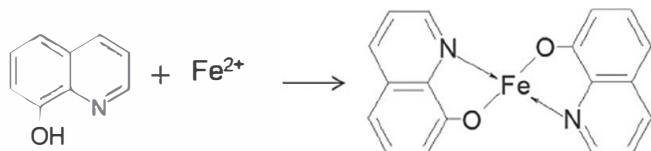


Fig. 3. Schematic illustration of the chelate formation between 8HQ and ferrous ions.

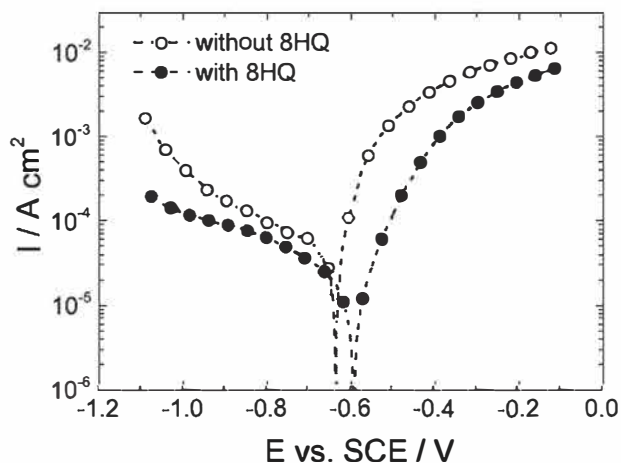


Fig. 4. Polarization curves obtained for the carbon steel electrode after 20 h of immersion in a 0.1 M NaCl solution without inhibitor and with 8HQ ( $3 \times 10^{-3}$  M).

Electrode rotation rate of 500 rpm.

Where  $R_p^i$  and  $R_p^w$  are the polarisation resistance values in the presence or in the absence of the inhibitor, respectively. The  $R_p$  and IE values are reported in Table 2. The efficiency reached 80% after 20 h of immersion which corresponds to a moderate corrosion inhibition efficiency.

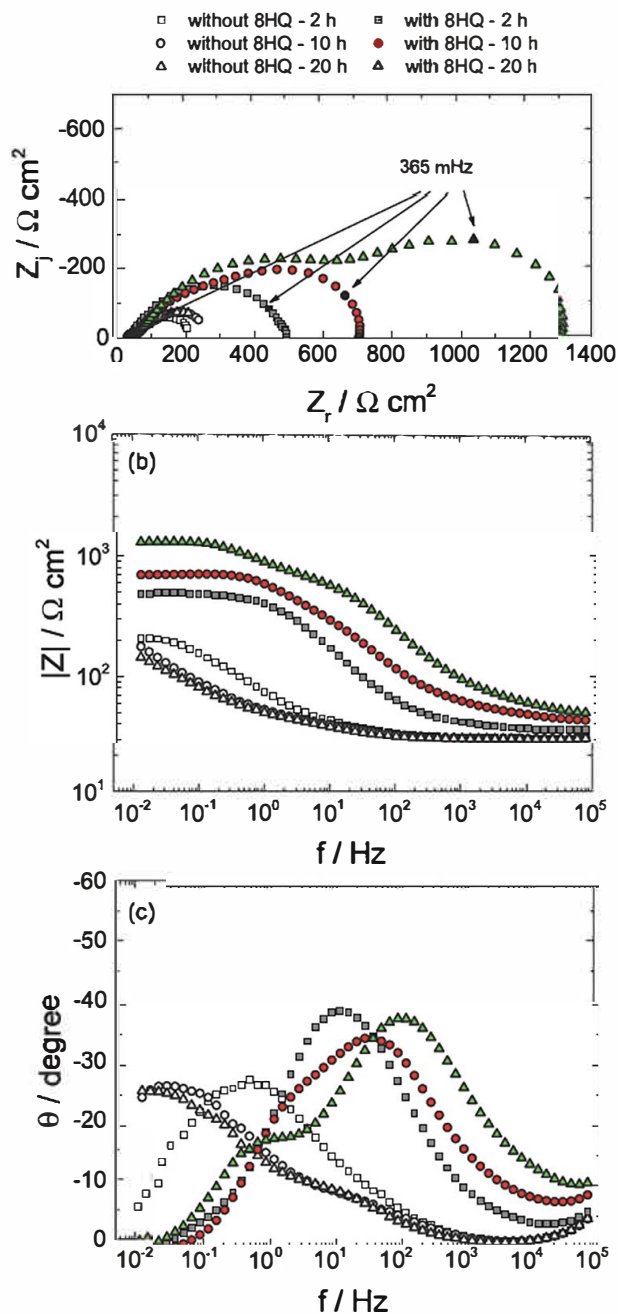


Fig. 5. Electrochemical impedance diagrams obtained at  $E_{corr}$  for the carbon steel electrode for various immersion times in a 0.1 M NaCl solution with and without 8HQ ( $3 \times 10^{-3}$  M): (a) Nyquist plot, (b) impedance modulus, (c) phase angle. Electrode rotation rate of 500 rpm.

Table 2

$R_p$  values extracted from the impedance diagrams in Fig. 5 and corresponding efficiency values.

Immersion time	without 8HQ $R_p^w$ ( $\Omega \text{ cm}^2$ )	with 8HQ $R_p^i$ ( $\Omega \text{ cm}^2$ )	Inhibitor efficiency IE (%)
2 h	206 ± 11	511 ± 24	59 ± 4
10 h	267 ± 12	741 ± 50	64 ± 4
20 h	270 ± 8	1425 ± 141	81 ± 2

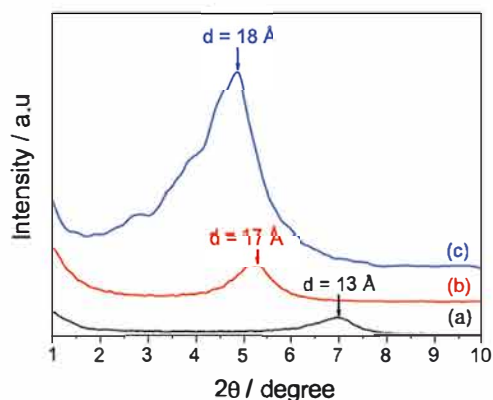


Fig. 6. X-ray diffraction spectra of: (a) MMT, (b) 8HQ-MMT and (c) AA-MMT.

### 3.2. Characterization of 8 HQ modified clay

Fig. 6 displays XRD patterns of the MMT, the AA MMT and the 8HQ MMT. The peak corresponding to the  $d_{001}$  plane appears at  $2\theta = 7.0^\circ$ ,  $5.2^\circ$  and  $4.8^\circ$  for the MMT, the 8HQ MMT and the AA MMT, respectively. The  $d_{001}$  peak corresponds to the  $d$  spacing of the lamellar structure of the MMT. The  $d$  spacing was calculated from Bragg's law:

$$\lambda = 2d \sin \theta \quad (2)$$

where  $\lambda$  is the wavelength of the x ray radiation,  $\theta$  is the glancing angle and  $d$  is the interplanar spacing of the clay layers. The calculated  $d$  spacing was 13 Å, 17 Å and 18 Å for the MMT, the 8HQ MMT and the AA MMT, respectively. The increase of the  $d_{001}$  spacing for the 8HQ MMT, compared to the MMT, confirmed the cations exchange by the 8HQ.

Fig. 7 shows TGA thermograms for the 8HQ, the MMT and the 8HQ MMT. Between  $40^\circ\text{C}$  and  $100^\circ\text{C}$ , the 10% mass loss for the MMT (Fig. 7b) correspond to the evaporation of adsorbed water. In the temperature range of  $400^\circ\text{C}$  to  $700^\circ\text{C}$ , an additional weight loss of about 8% can be observed, attributed to the dehydration of the remaining water bonded in the crystal lattice [41]. For the 8HQ MMT (Fig. 7c), four mass losses are identified. The first one, associated to

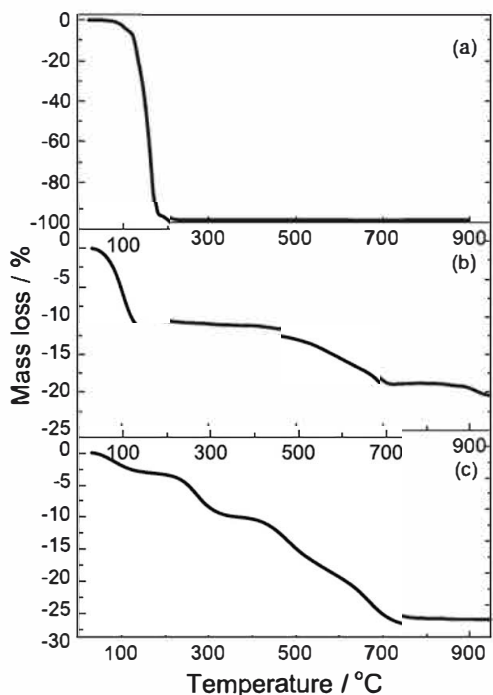


Fig. 7. TGA thermograms of: (a) 8HQ, (b) MMT, and (c) 8HQ-MMT.

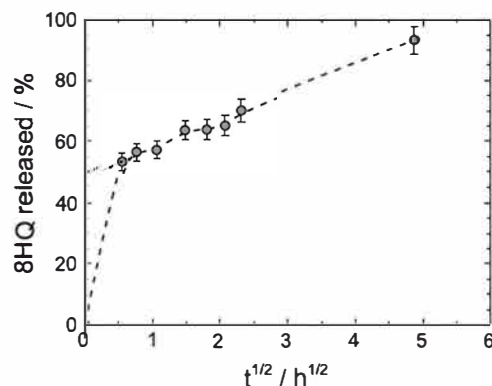


Fig. 8. 8HQ release from 8HQ-MMT in a 0.5 M NaCl solution.

water evaporation is similar to that observed for the MMT (Fig. 7b) but the mass loss is lower (3%). In the  $200-700^\circ\text{C}$  region, three mass losses are seen but the two last ones are not well separated. The mass loss beginning at  $220^\circ\text{C}$  can be associated to the degradation of the 8HQ which is weakly adsorbed on the MMT. The one beginning at  $420^\circ\text{C}$  is attributed to the degradation of the 8HQ inserted in the clay gallery by the ions exchange reaction [42]. It can be seen that the 8HQ degradation occurs around  $170^\circ\text{C}$  (Fig. 7a) while for the 8HQ MMT it started at  $220^\circ\text{C}$ . This shift can be explained by the barrier effect of the clay platelets which slowed down the volatile compounds diffusion [42]. Then, from  $500^\circ\text{C}$  to  $700^\circ\text{C}$ , the mass loss is similar to that observed for the MMT (Fig. 7b). A precise deconvolution of these two last steps was not possible but from the comparison of the two thermograms (Fig. 7b and c), the 8HQ content in the clay platelets is evaluated at about 16 wt. %.

Fig. 8 shows the release of the 8HQ from the 8HQ MMT during 24 h in a 0.5 M NaCl solution. The data are plotted as a function of the square root of time. A linear dependence is observed between 15 min and 24 h indicating that the release of 8HQ was mainly controlled by the inhibitor diffusion within the MMT platelets. After 24 h of immersion, the 8HQ release reached about 100%. The extrapolation of the straight line to  $t = 0$  gives a value of 50% of released 8HQ. From these two observations, it can be assumed that the rapid 8HQ release at the beginning of immersion would be due to the presence of the 8HQ molecules weakly adsorbed on the MMT, in agreement with the TGA thermogram (Fig. 7c). After 24 h of immersion, all the 8HQ molecules, incorporated in the MMT platelets, were released into the NaCl solution.

### 3.3. Corrosion protection performance of the coatings

Fig. 9 shows the impedance diagrams obtained for the epoxy coatings containing the AA MMT and the 8HQ MMT for different exposure times to the NaCl solution. For the coating with the AA MMT, the impedance diagrams are modified during immersion. The modulus is high at the beginning of immersion (about  $4 \cdot 10^8 \Omega \text{cm}^2$ ), then it decreases for longer exposure times and slightly increases after 21 days of immersion (Fig. 9a). After 10 days, two time constants are visible: the first one in the high frequency domain represents the properties of the coating and the second one in the low frequency region corresponds to the reaction occurring on the carbon steel surface at the bottom of the pores of the coating [43]. This result indicates a degradation of the steel/coating interface and the slight increase of the modulus at low frequency ( $8 \cdot 10^6 \Omega \text{cm}^2$  after 13 days to  $1 \cdot 10^7 \Omega \text{cm}^2$  after 21 days of immersion) might be linked to the formation of corrosion products at the bottom of the pores. For the epoxy coating containing the 8HQ MMT, independently of the immersion time, the diagrams (impedance modulus and phase angle) remain unchanged and only one time constant is observed (Fig. 9b and b'). This result shows that the barrier properties

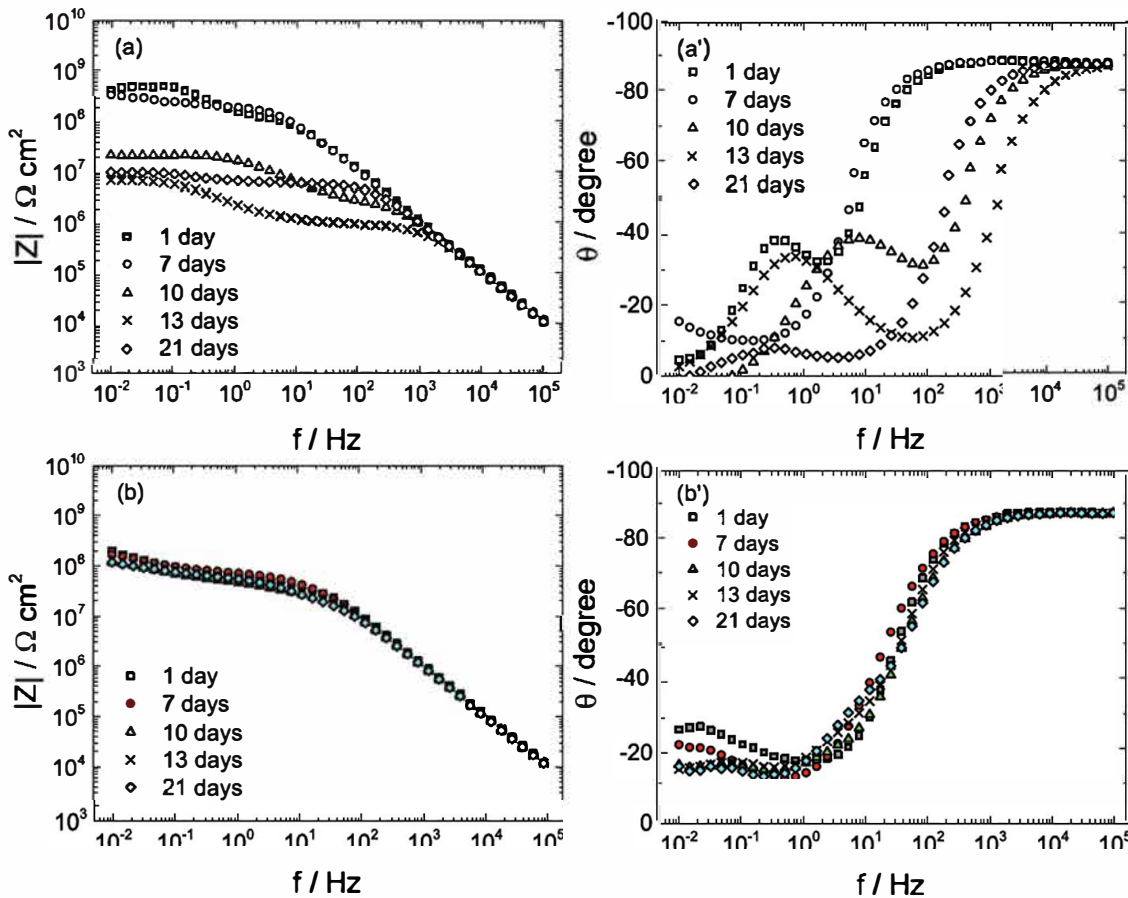


Fig. 9. Electrochemical impedance diagrams obtained for different immersion times in a 0.5 M NaCl solution for the carbon steel covered by the epoxy coatings containing: (a, a') AA-MMT and (b, b') 8HQ-MMT.

are maintained during exposure to the NaCl solution without any degradation of the substrate. For the two coatings, the impedance modulus for short immersion are relatively similar (about  $2.4 \times 10^8 \Omega \text{ cm}^2$ ) showing similar barrier properties.

From the impedance data, the complex capacitance diagrams can be obtained to determine the dielectric constant of the epoxy coatings as a function of the immersion time in the NaCl solution [44]. Fig. 10 shows, as an example, the complex capacitance plots obtained for the coating containing 8HQ MMT. When the immersion time increases, the coating capacitance values slightly increase from  $12.2 \text{ nF cm}^2$  after 1 day to

$13.3 \text{ nF cm}^2$  after 21 days. Then, the dielectric constant ( $\epsilon_w$ ) was calculated from the usual relationship expressing the capacitance of a plane capacitor, i.e.,

$$\epsilon_w = \frac{C_w \delta}{\epsilon_0} \quad (3)$$

where  $\epsilon_w$ ,  $\epsilon_0$ , and  $\delta$  are the dielectric constant of the coating under wet condition, the vacuum permittivity and the thickness of the coating, respectively. The calculated  $\epsilon_w$  values were about  $4.5 \pm 1$  and they little depend on the immersion time (variations are lower than the uncertainty). The permittivity reported in literature for other dry epoxy coatings is in the range of 3–5 [44,45]. Thus, in the present work, the measured permittivity in wet condition ( $\epsilon_w$ ) indicated a low water content in the coatings linked to the good barrier properties of the two epoxy coatings.

### 3.5. Adherence results and salt spray test

The adherence results from the pull off and the cross cut tests are summarized in Table 3. From the pull off test, in dry condition, it can be seen that the strength was two times higher for the coating containing 8HQ MMT compared to the coating containing AA MMT. However, the cross cut test did not allow the two coatings adherence to be differentiated in dry condition and the two coatings have the same classification (5B). In contrast, in wet condition, the epoxy film containing AA MMT was completely removed from the steel surface (0B) while only few coating fragments were removed for the sample containing 8HQ MMT (4B). These results indicate that the adherence of the as prepared coatings containing 8HQ MMT was higher and the adhesion properties were maintained after exposure to the NaCl solution. 8HQ acts as an

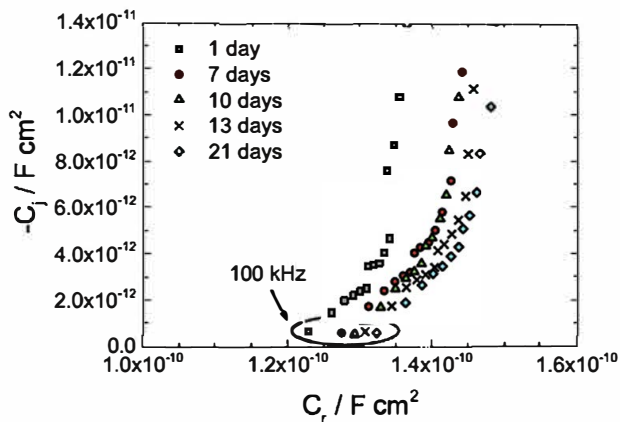


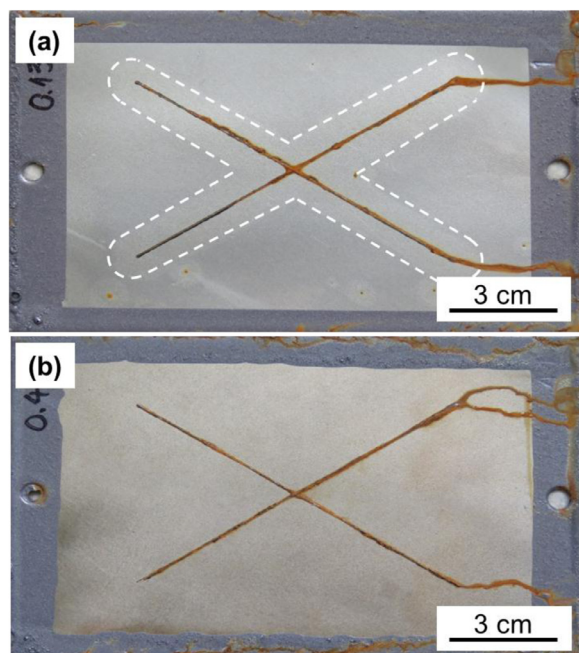
Fig. 10. Complex-capacitance plots corresponding to EIS spectra (Fig. 9b and b') obtained as a function of immersion time in a 0.5 M NaCl solution for the coating containing 8HQ-MMT (the axis are not orthonormal to better visualize the data).



**Table 3**

Adherence tests results for the AA-MMT and the 8HQ-MMT epoxy coatings in dry and wet conditions.

	Dry condition	Cross-cut classification	Wet condition	Cross-cut classification
	Pull-off strength (MPa)		Pull-off strength (MPa)	
AA-MMT	1.3 ± 0.2	5B	–	0B
8HQ-MMT	2.6 ± 0.1	5B	–	4B

**Fig. 11.** Photographs of the carbon steel coated with the epoxy coatings containing: (a) the AA-MMT and (b) the 8HQ-MMT after 240 h of exposure to the salt spray test.

Dotted line represents the delaminated surface area.

adhesion promoter, probably due to its affinity with both the substrate and the epoxy matrix. The adherence improvement in the presence of 8HQ MMT is in agreement with recent results reported on polyepoxy powder coatings containing halloysite or mesoporous silica loaded with 8HQ [46].

The salt spray test was used to evaluate the corrosion performance of the epoxy coatings containing AA MMT or 8HQ MMT. The photographs of the samples after 240 h of salt spray exposure are shown in Fig. 11. For the coating containing AA MMT, the delamination is significantly expanded all around the scratch (about 1 cm). For the coating containing 8HQ MMT, no delamination is observed and corroborates the adherence improvement previously shown (Table 3). In the scratch, the corrosion is similar for both samples.

EIS measurements (both global and local) were performed on scratched coatings to reveal the inhibitive role of the 8HQ on the scratch (results not shown). Whatever the exposure time to the aggressive solution, no difference between the samples containing the AA MMT and the 8HQ MMT was observed and it was concluded that there was no clear corrosion inhibition due to the 8HQ in the scratches.

#### 4. Discussion

From the impedance results (Figs. 9 and 10), it was shown that MMT treated with a linear hydrocarbon (AA MMT) or heterocyclic compound (8HQ MMT) and incorporated into a solvent free epoxy coating provided good barrier properties. This can be attributed to the good dispersion of the organically modified MMT in the epoxy coating, which slowed down the diffusion of water and aggressive species towards the

metal substrate. However, it can be assumed that the barrier properties would not be sufficient for long term corrosion protection of the carbon steel. From the salt spray test (Fig. 11) and from global and local impedance measurements on scratched coatings (results not shown), the inhibitive effect of 8HQ MMT incorporated into the epoxy coating was not clearly shown; corrosion products were always visible in the scratches. This could be attributed to the low 8HQ content in the coating (3 wt.% of 8HQ MMT incorporated in the coating corresponds approximately to 0.5 wt.% of 8HQ, as calculated from Fig. 7), which would not be enough to impede the corrosion process, even if all the 8HQ molecules were released from the MMT platelets (Fig. 8).

A higher adherence value in dry condition (pull off strength) was obtained for the coating containing 8HQ MMT (Table 3). In addition, the cross cut test, in wet conditions, confirmed the better adhesion of the coating containing the 8HQ MMT compared to the samples containing the AA MMT. From the impedance diagrams (Fig. 9b), it was seen that the 8HQ MMT coating performance was maintained during immersion and no delamination was observed during salt spray test (Fig. 11). These two results were attributed to the adherence improvement in the presence of the 8HQ MMT.

#### 5. Conclusions

This study focused on the influence of an organic inhibitor (8HQ) incorporated in the MMT platelets to improve the protection of a carbon steel by a solvent free epoxy coating. The 8HQ MMT was compared with an alkyl quaternary ammonium (AA MMT). The impedance results showed the good barrier properties of the epoxy coatings containing the organo modified MMT. The adherence test in dry and wet conditions combined with the salt spray test clearly highlighted the beneficial role of the 8HQ at the steel/coating interface (improvement of the adhesion). However, the inhibitive role of the molecule was not observed which might be attributed to the low content of the active molecule in the coating.

This study underlined the limitation to use containers, such as MMT, to incorporate organic inhibitors into coatings because the inhibitor concentration would not be sufficient to afford good corrosion protection. On the other hand, the containers cannot be introduced at higher contents due to their detrimental effects on the barrier and the mechanical properties [26,37].

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