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Highly Stable Molybdenum Disulfide Protected Silicon Photocathodes for Photoelectrochemical Water Splitting

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Abstract

Developing materials, interfaces and devices with improved stability remains one of the key challenges in the field of photoelectrochemical water splitting. As a barrier to corrosion, molybdenum disulfide is a particularly attractive protection layer for photocathodes due to its inherent stability in acid, the low permeability of its basal planes, and the excellent hydrogen evolution reaction (HER) activity the MoS₂ edge. Here, we demonstrate a stable silicon photocathode containing a protecting layer consisting of molybdenum disulfide, molybdenum silicide, and silicon oxide which operates continuously for two months. We make comparisons between this system and another molybdenum sulfide – silicon photocathode embodiment, taking both systems to catastrophic failure during photoelectrochemical stability measurements and exploring mechanisms of degradation. X-ray photoelectron spectroscopy and transmission electron microscopy provide key insights into the origins of stability.

Introduction

Hydrogen is an important industrial chemical for ammonia synthesis and petroleum refining with annual production of over 65 million tons.¹ Sustainably produced hydrogen also offers a renewable alternative to fossil fuel based energy as an energy carrier for the transportation sector. Currently, 96 % of our hydrogen is produced from fossil fuels (mostly by steam methane reformation of natural gas) with CO₂ as a by-product. Photoelectrochemical (PEC) water splitting is an attractive and promising alternative route for the conversion of sunlight into hydrogen without direct production of CO₂.² However, for PEC to become economically viable, the efficiency and durability of electrodes must be improved.³

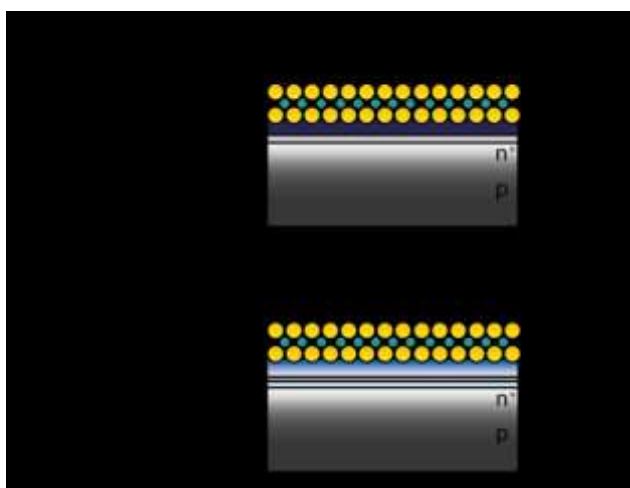
Silicon is a promising candidate material for use in PEC water splitting systems due to its excellent electronic properties, appropriate band gap of 1.1 eV, which is suitable for the small band gap electrode in a dual absorber device, and band alignment to drive the hydrogen evolution reaction (HER).⁴ However, bare silicon surfaces exhibit slow kinetics for hydrogen evolution and corrode readily in aqueous electrolytes, necessitating the use of strategies which speed the HER and slow corrosion. To date, several methods have been employed to stabilize semiconductor surfaces against corrosion including, thin film protection layers which can also enhance HER catalysis,⁵ the replacement of aqueous electrolyte with non-corrosive non-aqueous redox systems,^{6,7} and surface functionalization with organic molecules that prevent corrosion⁸. The simplest and most effective strategy to date is the use of conductive and electrochemically stable thin films that act as a physical barrier to corrosion.⁵ Numerous protecting and catalytically active materials have been employed with varying degrees of success including Pt,⁹ metal oxides,¹⁰⁻¹² phosphides,^{13,14} sulfides,^{15,16} and silicides⁹.

One such material that has shown excellent stability as a protection layer for PEC water splitting photocathodes is molybdenum disulfide (MoS₂). MoS₂ is known for its exceptional stability, HER activity in acid, and its low permeability, making it an excellent candidate as a protective catalyst for PEC applications.¹⁷ Indeed, MoS₂ protection layers have been utilized to extend the durability of gallium indium phosphide photocathodes for over 60 hrs¹⁸ and silicon photocathodes for over 100 hours.^{15,16}

Metal silicides, which are typically formed by sputtering a metal directly onto silicon, are also interesting candidates for protection layers on silicon due to their inherent stabilities,¹⁹ and the excellent interface between the silicon absorber and silicide layer. Silicides of platinum, for instance, have been investigated as protection layers for water splitting photoelectrodes

with excellent results; specifically, sputtered platinum silicide thin films show stable activity under continuous operation for 58 days (a total of 20 000 C.cm⁻² was passed) in 1 M hydrochloric acid.⁹ There have also been reports utilizing platinum, molybdenum, and tungsten silicides for the photoelectrochemical oxidation of iodide.²⁰ The longest lived photoanode for the photo-oxidation of iodide is a tungsten coated p⁺n-Si photoanodes that remained active for over 187 days of continuous illumination (>216 000 C.cm⁻²).²⁰ Interestingly, such silicide protection layers developed 20 years ago remain some of the longest lived photoelectrodes reported to date.

In this study, we investigate the photoelectrochemical activity and stability of two similar, but distinct protection schemes for silicon photocathodes, Scheme 1a and Scheme 1b. The first protection scheme (Si-2IL-MoS₂) has two interlayers (IL) between the surface MoS₂ and silicon composed of molybdenum silicide and molybdenum metal. The second protection scheme investigated has three IL (Si-3IL-MoS₂) identified as molybdenum oxide, molybdenum silicide and silicon dioxide. The photocathodes are amongst the most stable precious metal free electrodes prepared to date, with over 1538 h of continuous and stable photoelectrochemical hydrogen evolution,²¹ and over 64 000 C cm⁻² of charge passed for the Si-3IL-MoS₂ protection scheme. These non-precious metal photoelectrodes also show an excellent onset potential, reaching 0.5 mA cm⁻² at +0.31 V vs. RHE. This serves as a promising protection scheme for active and stable photoelectrochemical water splitting.



Scheme 1. Schematics detailing the catalyst and protection layers at the silicon interface in the synthesised (a) Si-2IL-MoS₂ and (b) Si-3IL-MoS₂ photocathodes.

Results and Discussions

The Si-2IL-MoS₂ photocathode. The Si-2IL-MoS₂ photocathode was prepared following the method published in ref 15 and detailed in Scheme 1a. The protection scheme composition was confirmed (ref 15) by cross-sectional TEM and found to contain two interlayers (IL) between a surface MoS₂ and the silicon substrate. The uppermost IL was a Mo metal layer. The second IL was identified as amorphous Mo_xSi.

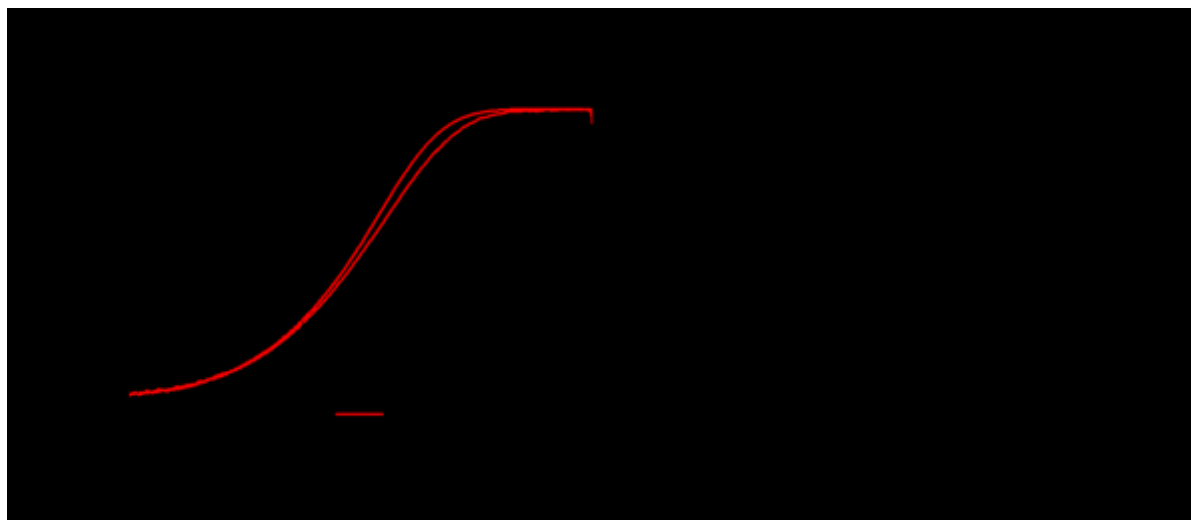


Figure 1. Photoelectrochemical stability measurements for the Si-2IL-MoS₂ photocathodes; Cyclic voltammograms, (a) were taken every 24 hours, here we show the CV taken on day 1 (black) and the last day of testing prior to failure (red). (b) Chronoamperometry measurements at E = 0 V vs. RHE.

To probe the stability and performance of the prepared photocathodes, the photoelectrodes were continuously illuminated under approximately AM1.5 until catastrophic failure (as measured by loss in photocurrent). The current density was measured throughout testing while holding the electrode at 0 V vs. RHE. A cyclic voltammogram was recorded once every 24 h to access changes in the electrodes current-voltage characteristics. The onset potential of the Si-2IL-MoS₂ photocathode was recorded at +0.28 V vs. RHE (defined as the potential required to reach 0.5 mA.cm⁻²), in line with our previous work on similarly prepared photoelectrodes (Fig. 1 a).¹⁵ Three repeat electrodes were prepared and their initial photoelectrochemical performance is compared in Fig. S2a. The stability measurements were conducted for 100 h previously.¹⁵ Here, we extend the stability and find the photocathode to be stable for 24 days (606 h) of continuous illumination (Fig. 1 b). The average photocurrent was 15.7 mA cm⁻² with > 34 000 C cm⁻² charge passed prior to failure.

Over the duration of the PEC stability measurements the light limiting saturation photocurrent of the Si-2IL-MoS₂ photocathode increased (Fig. 1) from approximately -16.5 mA cm⁻² to -19.0 mA cm⁻² (at -0.3 V vs. RHE). We speculate that this is due to a combination of the dissolution of molybdenum, and/or roughening of the electrode surface over the 25 days of continuous testing, leading to a thinning of the protection/catalytic layers resulting in a reduction of the parasitic light absorption. In addition to the increased photocurrent at around -0.3 V vs. RHE a slight increase in the onset potential was also measured for the photocathode. Specifically, the Si-2IL-MoS₂ photocathode increased from +0.28 V to +0.33 V vs. RHE. Increases in photoelectrode performance over the course of stability measurements have previously been observed for other MoS₂ protected photocathodes and have been ascribed to the gradual exposure of additional MoS₂ edge sites,^{15,18} which are known to be the active sites for HER.²² However, increasing the number of catalytic sites of the MoS₂ catalyst should only increase reaction kinetics leading to an increase in the slope of the current-voltage curve and, perhaps its saturation photocurrent value. Therefore, we ascribe the change in onset potential to a surface chemical composition modification such as the exposure of underlying layers, such as molybdenum or molybdenum silicide. On the 25th day of testing, failure of the Si-2IL-MoS₂ photocathode occurs during a chronoamperometry hold whereby the current drops from -16.4 mA cm⁻² to -1.9 mA cm⁻² over approximately 4.5 h. Complete failure (a reduction in photocurrent to < -0.1 mA cm⁻²) occurs over a 24 h period.

Despite the overall excellent photoelectrochemical stability of the photocathode, some fluctuation in the photocurrent onset, saturation density and chronoamperometry was observed over the course of the stability tests (Fig. 1). The fluctuations in current density can be attributed to two distinct phenomena. Firstly, the smaller amplitude (up to 1 mA cm⁻²), higher frequency noise which is attributed to the accumulation and removal of H₂ bubbles at the electrode surface. The second distinct noise in the data consists of significantly larger spikes in the current density (1 – 8 mA cm⁻²) that occur approximately every 24 h which we speculate could be due to formation of MoO₃ which is unstable at cathodic potentials in 0.5M sulfuric acid and/or contamination. Detailed discussion and explanation for the fluctuations is provided in the supporting information.

XPS characterisation was performed before and after photoelectrochemical experiments to identify the composition and chemical state of the two photocathodes (Fig. 2.). Peak fitting shows that the Mo 3d region contains peaks associated with Mo (228.2 and 231.6 eV) and Mo⁴⁺ (229.0 and 232.2 eV) only in addition to the S 2s region (226.2 eV). Quantification of the Mo

3d region reveals 20 % of the molybdenum is metallic, and 80 % is Mo (+4). In the silicon 2p region, no signal (within the detection limits) is detected. A doublet is observed in the sulfur 2p region at approximately 163 eV. Post catastrophic failure, both elemental (or zero valent) molybdenum and sulfur are below the detection limits for the XPS. Silver contamination was detected at the surface of the failed electrode. Evidently, the protection layer has been completely removed from the surface leading to catastrophic failure of the device.

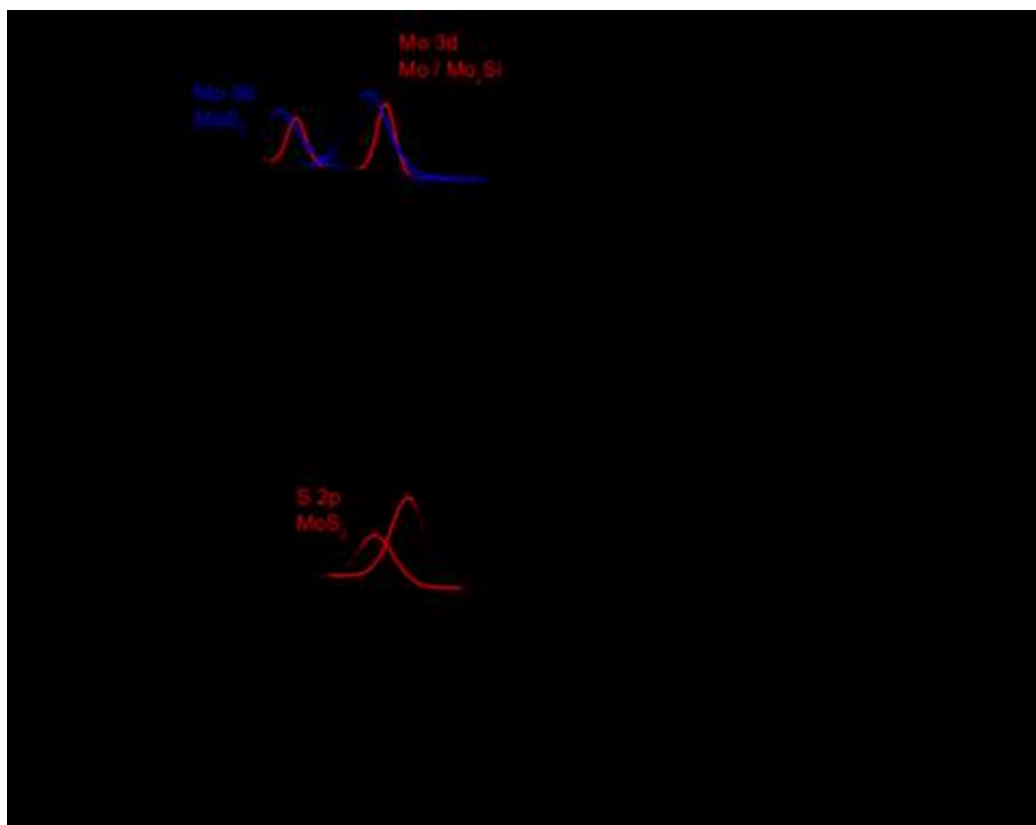


Figure 2. XPS data and fittings of the Si-2IL-MoS₂ photocathode for the (a) Mo 3d, (b) Si 2p and (c) S 2p regions. The as prepared photocathode spectra are shown in the upper plots, and post photoelectrochemical testing in the lower plots.

The Si-3IL-MoS₂ Photocathode. The synthesis of the Si-3IL-MoS₂ photocathode is presented in Scheme 1b. In brief, p-type silicon was n⁺ doped, cleaned and coated with a 3.6 nm layer of molybdenum by sputtering. The electrodes were subsequently annealed in a H₂ atmosphere at 550 °C (2 h), followed by a second anneal step in 10 % H₂S in H₂ at 250 °C (1 h).

Cross sectional transmission electron microscope (TEM) images of the as prepared sample is shown in Fig. 3. evidencing the formation of a ~3.7 nm silicon oxide (SiO₂) layer directly above the silicon, and a ~10 nm layer of molybdenum/molybdenum oxide (Mo/MoO_x) above the SiO₂. There is also evidence of an additional thin (1 – 2 nm) layer between the MoO_x and

SiO₂ which is likely molybdenum silicide (MoSi_x). Platinum is the uppermost layer, deposited during cross sectional TEM preparation as a protection layer for the FIB ion beam milling.

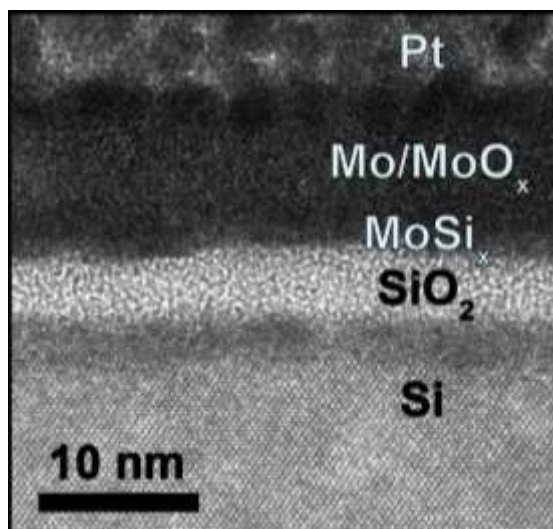


Figure 3. Transmission electron micrograph (TEM) of the as prepared molybdenum protected Si-3IL-MoS₂ photocathode. Platinum is deposited during cross sectional TEM preparation as a protection layer for the FIB ion beam milling so is not present during the photoelectrochemical measurements.

The overall thickness of the Si-3IL-MoS₂ protection scheme (~14 nm) is greater than the Si-2IL-MoS₂ protection scheme. This is due to a thicker MoO_x layer, in combination with an additional layer of SiO₂. Interestingly, the distinct basal plane *d* spacing (0.62 nm) of MoS₂ are absent in the Si-3IL-MoS₂ TEM (Fig. 3) indicating a non-continuous film of MoS₂. The Si-3IL-MoS₂ photocathode synthesis requires a pre-reduction in a hydrogen atmosphere, which has been shown to limit the growth of MoS₂,²³ however, XPS (Fig. 4) verifies the growth of MoS₂ at the surface of the photocathode.

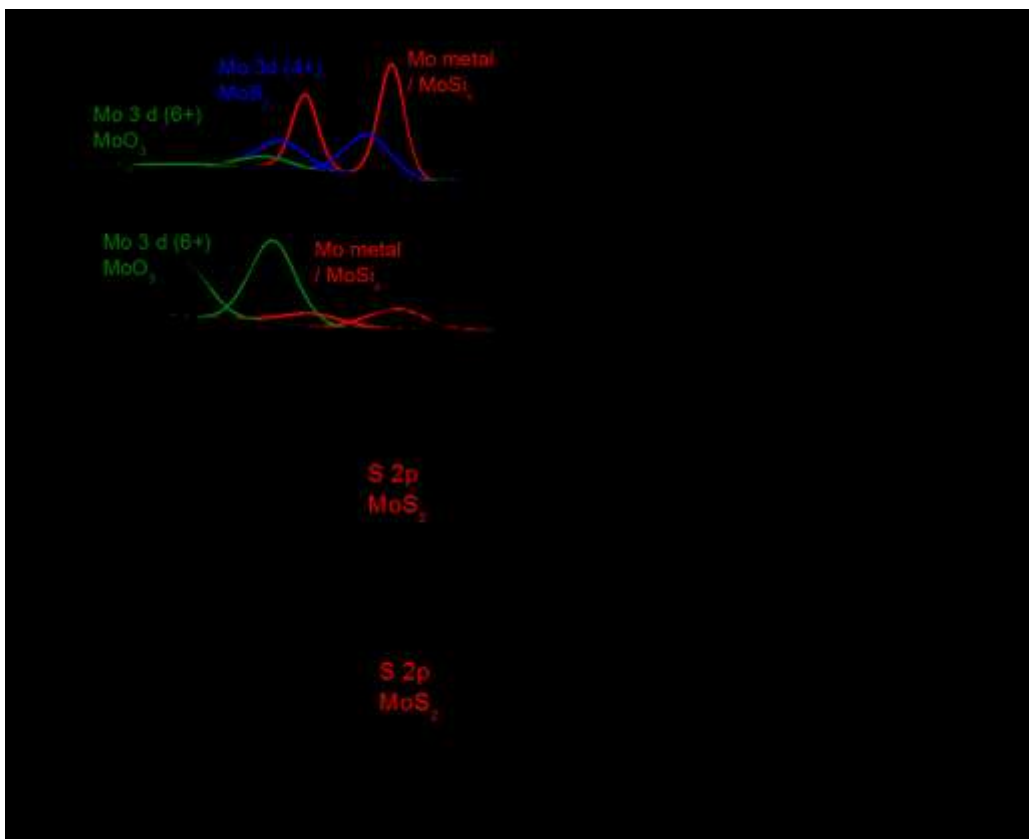


Figure 4. XPS data and fittings of the Si-3IL-MoS₂ photocathode for the (a) Mo 3d, (b) Si 2p and (c) S 2p regions. The as prepared photocathode spectra are shown in the upper plots, and post photoelectrochemical testing in the lower plots.

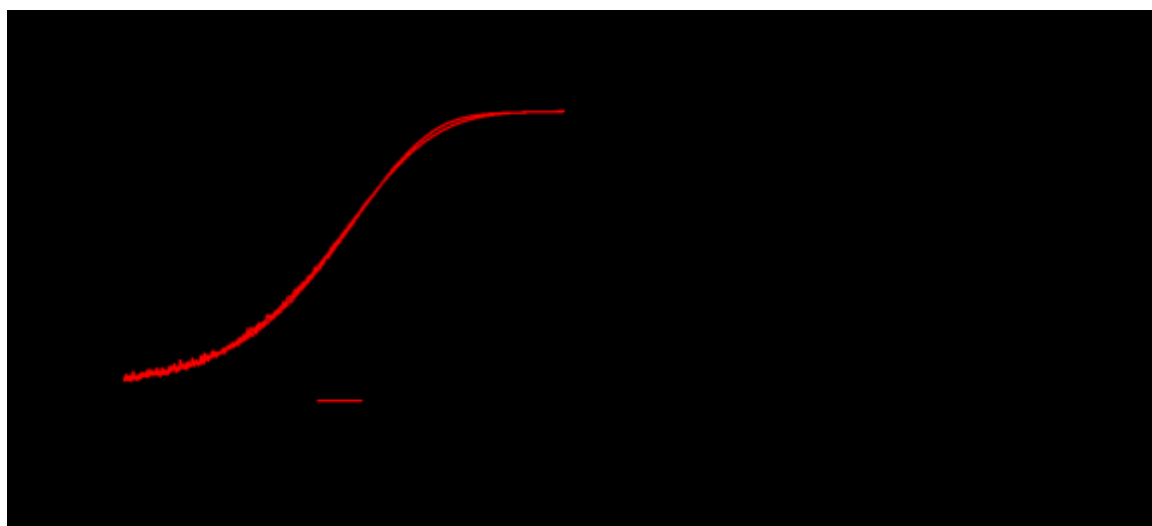


Figure 5. Photoelectrochemical stability measurements for the Si-3IL-MoS₂ photocathodes; Cyclic voltammograms, (a) were taken every 24 hours, here we show the CV taken on day 1 (black) and the last day of testing prior to failure (red). (b) Chronoamperometry measurements at E = 0 V vs. RHE.

The cyclic voltammograms (Fig. 5 a) reveal an excellent onset potential of +0.31 V (defined here as the potential required to achieve -0.5 mA cm^{-2})¹⁷ and achieves -10 mA cm^{-2} at +0.09 V vs. RHE on both day 1 and day 62. Similar photoelectrochemical performance was recorded for repeat samples as shown in Fig. S2b. The electrodes were found to be active with an average saturation photocurrent density of 11.1 mA cm^{-2} for over 64 days (1538 h) under continuous illumination and PEC testing (Fig. 5 b), equivalent to 320 calendar days of H_2 production under AM 1.5, sunny conditions with a solar capacity factor of 20 %.^{1,24} Bubble formation at the electrode surface was observed throughout the testing period passing $> 64\,000 \text{ C cm}^{-2}$ prior to failure which is equivalent to $0.66 \text{ g of H}_2 \text{ cm}^{-2}$ (7.39 L of H_2 at STP). We observe similar noise/fluctuations in the chronoamperometry measurements as in that recorded for the Si-2IL-MoS₂ photocathode (see supporting information for detailed explanations). For the Si-3IL-MoS₂ photocathode, the photocurrent increased from -15.3 mA cm^{-2} to -18.5 mA cm^{-2} over the course of the durability measurement. Analogous to the Si-2IL-MoS₂ photocathode, we speculate that this is due to a combination of the dissolution of molybdenum, and/or roughening of the electrode surface over the 64 days of continuous testing leading to a thinning of the protection/catalytic layers resulting in a reduction of the parasitic light absorption.

XPS characterisation was performed before and after photoelectrochemical experiments to identify the composition and chemical state of the Si-3IL-MoS₂ photocathode (Fig. 4.). XPS peak fitting of the freshly prepared photocathode shows that the Mo 3d region contains peaks associated with metallic Mo (228.0 and 231.2 eV) Mo⁴⁺ (228.9 and 232.1 eV) and Mo⁶⁺ (232.6 and 235.8 eV) in addition to the S 2s peak (226 eV). Quantification of the relative areas of each molybdenum oxidation states reveal the electrode contains 55% Mo and/or MoSi_x, 35% MoS₂ and/or MoO₂, 10% MoO₃. The “as prepared” photocathode silicon 2p signal has a rather high silicon oxide signal. This can be accredited to the formation of a silicon oxide layer, which is approximately 3.7 nm thick according to the cross sectional TEM (Fig. 3).²⁵ Sulfur species are also detected by XPS in the “as prepared” electrode. Prior to PEC testing the XPS peaks at 226 eV and 162 eV were assigned to S²⁻ 2s and 2p_{3/2} respectively.¹⁵

The failure of the Si-3IL-MoS₂ photocathode occurs after 1538 h of continuous PEC testing. The failure is catastrophic and occurs during a chronoamperometry measurement where the photocurrent drops abruptly from $\sim 10 \text{ mA cm}^{-2}$ to -0.2 mA cm^{-2} over 5 h. After the failure, XPS characterisation reveals a drastic transformation in the chemical oxidation states of the elements present at the surface of the photocathode. Similarly to the Si-2IL-MoS₂ electrode,

silver contamination was detected at the surface of the failed electrode. Significantly, the proportion of oxidized molybdenum (Mo (+6)) increased from 10 % to 75 %. Post failure, S²⁻ can still be observed in the sulfur 2p region indicating that some MoS₂ likely is still present. However, the decrease in 2 Mo⁴⁺ relative to Mo⁶⁺ indicates that some of the MoO₂ and MoS₂ has been removed from the electrode surface. The presence of an additional S 2p doublet at ~168.6 eV post failure can be assigned as the oxidized sulfur species SO₄²⁻ which is attributed to the usage of sulfuric acid as electrolyte throughout the PEC testing. The Si 2p signal post-failure consists of both oxidised SiO₂ (103.2 eV) and either elemental or silicide Si (99.8 eV), difficult to discern due to the small (0.2 eV) difference between their binding energies. The XPS signal for Si/MoSi_x is significantly enhanced post PEC testing suggesting that the molybdenum based protection layer is damaged, and thinner/discontinuous in places.

To further probe the failure mechanism of the Si-3IL-MoS₂ photocathode, scanning electron microscopy (SEM) and cross sectional TEM were used. SEM reveals a highly-roughened surface with both pits (10 – 40 nm) and agglomerates (up to 100 nm) scattered across the surface (Fig. 6 a). The cross sectional TEM images of the same sample (Fig. 6 b and Fig. S1) reveal significant roughening across the surface with pits up to 30 nm deep, as well as the addition of particles (10 – 20 nm). Behind the cavity, silicon lattice planes remain visible suggesting the cavity is only a few nm wide. Scanning TEM energy dispersive x-ray spectroscopy (STEM-EDS) elemental maps (Fig. 6.c-f) show that the pits extend through the protective molybdenum based layers, into the silicon and that silicon is etched away. The particles that decorate the surface were found to contain predominantly molybdenum, and in some cases silver (contaminant). Discussion regarding the origin of the contaminant is provided in the supporting information. In light of the cross sectional TEM (Fig. 6 b) whereby large pits and a broken protection layer is observed we speculate that such morphological changes are likely the eventual cause of electrode failure. Given that these failures occur during chronoamperometry testing we conclude that failure is not as a direct consequence of the cyclic voltammetry sweeps.

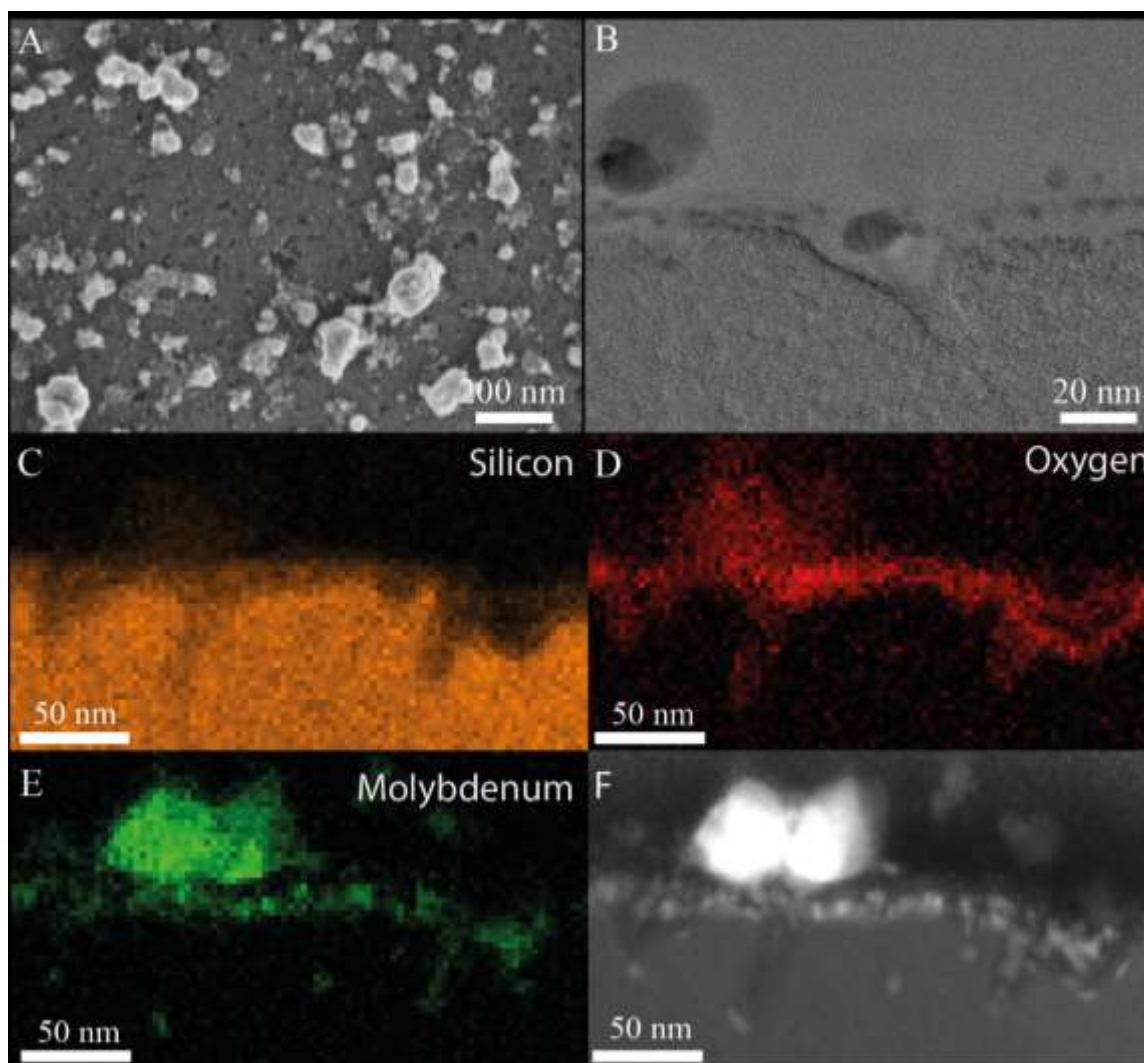


Figure 6. (A) Scanning electron micrograph (SEM) and (B) cross-sectional transmission electron micrograph (TEM) of the Si-3IL-MoS₂ photocathode post failure. Both pits and agglomerates are clearly visible at the surface of the electrode. STEMEDS mapping of silicon (C), oxygen (D) and molybdenum (E) of the sample after failure (64 days of testing). (F) The STEM micrograph of the same region is also shown.

Table 1 summarises the photoelectrochemical properties and durability of the two protection schemes investigated in this work. We speculate that the overall thickness of the Si-3IL-MoS₂ protection scheme, relative to that of the Si-2IL-MoS₂, in combination with both the ~3.7 nm SiO₂ and 1 – 2 nm MoSi_x layers provide additional chemical stability and thus contribute to the longevity of the devices enhancing the lifetime both with regards to time and total charge passed. We highlight the importance of future studies in this field that can uncover the specific role of individual layers (both chemical composition and thickness). Through a series of systematically designed and synthesised silicon photocathodes we plan to uncover the specific role of the individual layers (MoO_x/SiO₂/MoSi_x) in future experiments. Impurities at the

electrode surface are one of the many possible contributing factors to the failure of photoelectrodes. Impurities can be introduced into the system during various stages of the synthesis/testing, e.g. sputter deposition, annealing in the tube furnace, from electrolyte, water and the reference electrode. Furthermore, prior to deployment of industrial scale photoelectrochemical plants investigation of such sources and roles of impurities on device performance and stability will be of great importance.

Table 1. Summary of photoelectrochemical performance of the two photocathodes.

Photocathode	Onset Potential (V vs. RHE)	Photocurrent Saturation (mA cm⁻²)	Stability (h)	Charge passed (C cm⁻²)	Volume of H₂ (STP L)
Si-2IL-MoS ₂	+0.28	16.5 – 19.0	606	>34 000	3.98
Si-3IL-MoS ₂	+0.31	15.3 – 18.5	1638	>64 000	7.39

Conclusion

In conclusion, we report over 1500 h stability and over 64 000 C cm⁻² charge passed for a precious metal free molybdenum sulfide catalysed and molybdenum silicide protected silicon photocathode. Based on the high activity and durability of the molybdenum disulfide catalyst, we maintain an excellent onset potential for >62 days of continuous hydrogen production. Our investigation of two similar, but distinct protection schemes on silicon photocathodes provide evidence that molybdenum disulfide coupled to molybdenum oxide, molybdenum silicide layers and silicon dioxide creates an extremely effective barrier to corrosion. Catastrophic electrode failure reveals that some molybdenum remains present though over 75 % of the molybdenum is oxidised. Microscopy of the electrode post-mortem unveils a surface exhibiting pits and agglomerates/particles. Our findings demonstrate significant photoelectrode longevity and provide insights into degradation mechanisms that can enable strategies for future synthesis of highly durable PEC water splitting electrodes.

Methods

Synthesis of n⁺p-Si-Mo. For both photocathode syntheses, p-type (0.1-0.9 ohm) wafers (WRS) were cleaned (standard SC1 (5:1:1 H₂O:H₂O₂:NH₄OH at 50 °C for 10 minutes), HF for 30 s, and then SC2 (H₂O:H₂O₂:HCl at 50 °C for 10 minutes) etches) and doped with POCl₃ in a tube furnace at 900 °C for 10 mins. Following a further standard wafer clean (SC1 and SC2) and SiO₂ BOE etch, molybdenum metal (3.6 nm) was then sputtered onto the wafer with <10 min between the SiO₂ etch and pump down in the sputtering system, thus minimising the re-growth of native SiO₂. Post Mo sputtering, the sample was cleaved into approximately 0.2 cm² sized pieces which were subsequently treated as described below.

Synthesis of the Si-2IL-MoS₂ Photocathode. The Si-2IL-MoS₂ photocathode was prepared following the method published in Ref 15 and detailed in Scheme 1a. In brief, the n⁺p-Si-Mo wafer was sulfidized in an H₂S environment at 250 °C for 45 mins.

Synthesis of the Si-3IL-MoS₂ Photocathode. The synthesis of the Si-3IL-MoS₂ photocathode is presented in Scheme 1b. In brief, the prepared n⁺p-Si-Mo were annealed in a H₂ atmosphere at 550 °C (ramp rate 5 degree/min, 2 h dwell), followed by a second anneal step in 10 % H₂S in H₂ at 250 °C (ramp rate 5 degree/min, 1 h dwell). 99.999 % UHP (ultra high purity) hydrogen gas was used for all the syntheses plumbed into the tube furnace using swage lock stainless steel pipping.

Electrode Mounting. Electrical back contacts were made to the photocathode by diamond scribing InGa eutectic onto the surface, followed by a thin layer of carbon paint (Ted Pella, DAG-T-502) and finally, an insulated conductive wire. The back side of the electrode was fully encapsulated/isolated from electrolyte with two layers of Loctite® E-120 hydrosol epoxy. A third layer of Loctite® EA 9462 epoxy was subsequently applied and left for 24 h drying prior to PEC testing. The total illuminated electrode surface area was 0.24 cm².

Photoelectrochemical Measurements. To probe the stability of the prepared photocathodes, the electrodes were continuously illuminated under approximately AM1.5 until catastrophic failure (as measured by a loss in photocurrent). The current density was continuously measured while holding the electrode at 0 V vs. RHE, with the exception of a cyclic voltammogram recorded once every 24 h. Photoelectrochemical (PEC) testing was conducted in custom designed H-cells with a quartz window and in a three electrode configuration. All cell components (cell, gas dispersion tube, magnetic stir bar, quartz windows etc) were cleaned with aqua regia prior to PEC measurements. 0.5 M sulfuric acid electrolyte (99.9% purity,

Sigma Aldrich) with an H₂ purge and continuous stirring was used throughout all testing. A mercury/mercury sulfate reference electrode was used as a reference electrode and an iridium wire as a counter electrode which was separated from the photocathode electrode by a Nafion membrane. The reference electrode was calibrated with two Pt wires in H₂SO₄ before and after PEC testing to confirm stability over the 64 days of stability testing. The sample was illuminated with an Abet Solar Simulator (model number 11002) xenon lamp. The incident illumination was calibrated to match AM1.5 solar spectrum. Given the longevity of the PEC testing (64 days), the light source was calibrated both before and after PEC testing using an Ocean Optics Jaz EL 200-XR1 spectrometer and the CV data adjusted for the measured loss in solar intensity.

Physical Characterisation. XPS analysis was performed with a Phi Versaprobe 1. SEM microscopy was made with FEI Magellan 400 XHR Scanning Electron Microscope (SEM). Two thin lamella samples were prepared using focused ion beam (FIB, FEI Helios nanolab 600i) lift-out technique to perform cross-section TEM imaging. Samples were thinned by 30kV Ga ion beam and final cleaning process was done at 3kV accelerating voltage to polish of ion beam induced amorphous materials. Surface roughness and elemental distribution was analysed by high resolution TEM images and scanning TEM energy dispersive x-ray spectroscopy (STEM-EDS) operated at 80kV (TEM, FEI Titan environmental TEM 80-300).

Supporting Information

Additional discussion regarding the stability data, as well as HRTEM of the failed Si-3IL-MoS₂ is provided in the supporting information. This material is available free of charge via the Internet at <http://pubs.acs.org>.

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