

REMOVAL OF MALACHITE GREEN BY ADSORPTION ONTO ACTIVATED CARBON: MODELLING OF THE ADSORPTION KINETICS

ELIMINATION DU VERT DE MALACHITE PAR ADSORPTION SUR CHARBON ACTIF: MODELISATION DE LA CINETIQUE D'ADSORPTION

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RESUME

Dans ce travail, le charbon actif a été préparé à partir des résidus du fruit de Sapindus et utilisé comme adsorbant pour éliminer le vert de malachite contenu dans une solution aqueuse simulée. La cinétique d'adsorption du vert de malachite sur charbon actif a été étudiée en batch. Les effets du pH et du temps de contact ont été examinés. La quantité maximale adsorbée du vert de malachite a été obtenue à pH 6 (4.5 mg/g) et la quantité minimale a été obtenue à pH 2 (4.1 mg/g). Le temps d'équilibre apparent a été atteint après 120 min. Les conditions expérimentales optimales ont été déterminées. Dans le but de déterminer le modèle de cinétique d'adsorption le plus adéquat, les résultats expérimentaux ont été analysés en utilisant les modèles cinétiques pseudo-premier-ordre, pseudo-deuxième-ordre, pseudo-troisième-ordre, pseudo-premier-ordre, esquivel, et elovich. Le modèle non linéaire pseudo-premier-ordre est le plus adapté pour représenter les résultats expérimentaux. Cette présente étude montre que le charbon actif est un adsorbant moins onéreux, qui peut être utilisé, pour éliminer le vert de malachite de l'eau.

MOTS CLES: Charbon actif (AC), vert de malachite (VM), modélisation, linéaire, régression non-linéaire.

ABSTRACT

In this work, activated carbon was produced from residue of fruit of Sapindus and used for the application of adsorption removal of malachite green dye from simulated aqueous solution. Adsorption kinetics of malachite green onto activated carbon was studied in a batch system. The effects of pH and contact time were examined. The malachite green maximum adsorption occurred at pH 6 (4.5 mg/g) and the lowest adsorption occurred at pH 2 (4.1 mg/g). The apparent equilibrium was reached after 120 min. Optimal experimental conditions were determined. In order to determine the best-fit-adsorption kinetics, the experimental data were analyzed using pseudo-first-order, pseudo-second-order, pseudo-third-order, esquivel, and elovich models. Linear regressive and non-linear regressive methods were used to obtain the relative parameters. The statistical functions were estimated to find the suitable method which fit better the experimental data. Both methods were suitable to obtain the parameters. The non linear pseudo-first-order model was the best to fit the equilibrium data. The present work showed that activated carbon can be used as a low cost adsorbent for the malachite green removal from water.

KEYWORDS: Activated carbon (AC), malachite green (MG), Modelling, linear, non-linear regression.

1 INTRODUCTION

Malachite green (MG) is used in coloring paper, dyeing cottons, wools, silk, leather and coating for paper stock. The treatment of effluents containing such dyes is of great interest due to its harmful impacts on receiving waters (1). The best efficient method used for the quickly removal of dyes from the aqueous solution is the physical adsorption (2). Aromatic solutes showed slightly better adsorption than aliphatic solutes, due to the potential to form pi-pi bonds with the basal planes of activated carbon. No significant influence of solute charge or size was observed (3). This work aims to understand the potential of activated carbon (AC) to remove MG dye from simulated aqueous solution in batch mode. The adsorption kinetics efficiency of MG was investigated in order to optimize the experimental parameters such as contact time, and pH at an agitation speed of 300 rpm, initial adsorbent concentration of 5 mg/L and temperature of 25°C. The statistical functions were used to estimate the error deviations between experimental and theoretically predicted adsorption kinetics values, including Linear and non-linear method. The optimization procedure required a defined error function in order to evaluate the fit of equation to the experimental data.

2 EXPERIMENTAL PART

MG (4-(4-(dimethylamino) alpha-phenylbenzylidene)-2,5-cyclohexadien-1-ylidene) dimethylammonium chloride, $C_{23}H_{25}ClN_2$, $M_w = 364$ g/mol, Figure 1) used in the present study was purchased from Merck (Germany) and was selected from the list of dyes normally used in Algeria. The sieved residue of fruit of *Sapindus* was washed with distilled water to remove any residues or impurities. Subsequently, it was dried in an oven for 12 hours at 80°C. The material was pyrolyzed in a fluidized bed furnace at different temperature range. The pyrolysis process was undergone at temperature 300, 400 and 500°C for half an hour. Then, the material produced was discharged from the first cyclone of the fluidized bed furnace. This fast pyrolysis method produced variety of material at different temperature. Activation of the material was done by using steam average flowrate 300cc/min at temperature 800°C in a muffle furnace for 1 and 2 hours. The activated carbon was then ground and dried in an oven at 100°C for overnight.

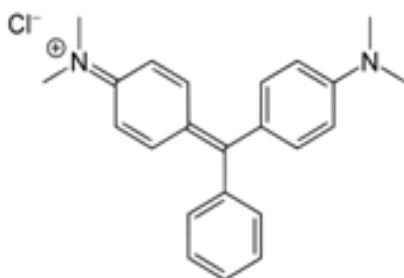


Figure 01: Structure of malachite green (4)

Adsorption kinetics of MG onto AC was studied in a batch system. The effects of pH and equilibrium time were examined. The adsorption parameters were optimized. In each experiment pre weighed amount of adsorbent (50 mg) was added to 50 mL of dye solution (5 mg/L) taken in a 250 mL of conical flask and 0.1 M NaOH or 0.1 M HCl were added to adjust the pH value. This solution was agitated at 300 rpm and centrifugated. The MG concentration in solution was determined at $\lambda_{max} = 620$ nm by spectrophotometer UV-1700 PHARMA SPEC SHIMADZU. The amount of MG adsorbed amount per mass unit of adsorbent at time t, q (mg/g), (Eq. (1)) was calculated as:

$$q = (C_0 - C) \frac{V}{M} \quad (1)$$

Where C_0 is the initial MG concentration (mg/L), C is the dye concentration at time t, V is the solution volume (L) and M is the adsorbent mass (g) (5). The effect of pH was conducted by mixing 1 g of adsorbent with 1 L of MG synthetic solution of 5 mg/L. The solution pH was varied from 2 to 12, by adding 0.1M NaOH or 0.1M HCl solutions. The suspension was shaken for 24 h at 25 °C. Kinetic experiments were performed by mixing 50 mL of dye solution (5 mg/L) with 50 mg (0.05 g) of adsorbent. The initial pH for each dye solution was set at 6. The suspensions were kept under agitation during 24 hours. MG concentrations in the supernatants were calculated and allowed to determine the amount adsorbed of dyes onto AC. The experiments were realized against time (5, 10, 15, 20, 25, 30, 40, 50, 60, 90, 120, 150, 180, 240 and 300 min).

3 RESULTS AND DISCUSSION

To study the effect of every parameter, it is necessary to fix the values of others. The elimination of pollutant from simulated aqueous solution by adsorption is extremely influenced by the medium of the solution which affects the nature of the adsorbent surface charge, the ionization extent, the aqueous adsorbate species speciation and the adsorption rate. The adsorptive process through functional groups dissociation on the adsorbate and adsorbent were affected by a pH change (6). The adsorption of MG increases with the increase of pH of the solution. Figure 2 shows the effect of the pH on the adsorption capacity of MG onto AC at various initial solution pH in the range 2–12 under the following conditions: initial dye concentration of 5 mg/L and AC dose of 1 g. From Figure 2, we noticed that the pH of the medium affect strongly the kinetic of MG fixation, showing an increase of the adsorption capacity with the pH, from 4.1 to 4.5 mg/g for pH increasing from 2 to 6.

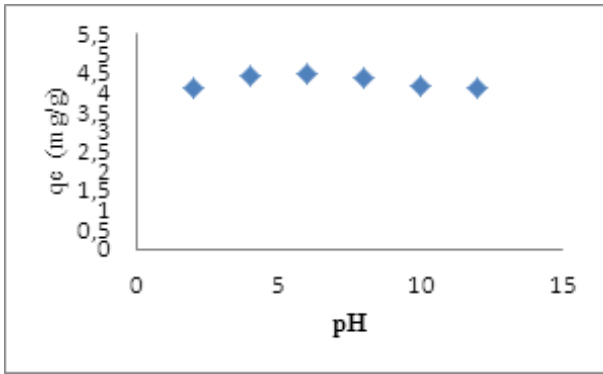


Figure 02: Effect of the initial pH on the AC equilibrium adsorption capacity

From this study, it is obvious that in the basic medium, the negatively charged species tends dominating and the surface began to acquire a negative charge. In this case the adsorbent surface is negatively charged. The MG adsorption increased due to the increasing of electrostatic attractions between the negative charge of AC particles and the positive charge of MG species.

Figure 3 highlights the adsorption kinetics of MG onto AC. In the light of the result, the synthetic sample could be divided in three zones: (i) 0-30 min, which indicated the fast adsorption of MG, suggesting rapid external diffusion and surface adsorption; (ii) 30-60 min, showed a gradual equilibrium, and (iii) 60-300 min, indicated the plateau of the equilibrium state. The adsorption was rapid at the initial stage of the contact, but it gradually slowed down until the equilibrium. The fast adsorption at the initial stage can be attributed to the fact that a large number of surface sites are available for adsorption. After a lapse of time, the remaining surface sites are difficult to be occupied because of the repulsion between the solute molecules of the solid and bulk phases make it took to long time to reach equilibrium. Adsorption is a complex process whereby it is influenced by several parameters related to adsorbent and to the physicochemical conditions under which the process is carried out (7). In order to understand the mechanism of the adsorption process, the following equations: pseudo-first order (Lagergren Model) (1), pseudo-second order (8), esquivel (9), pseudo-third order (10), and elovich (11) were selected to fit the experimental kinetic data. Equations of these models were illustrated in Table 1.

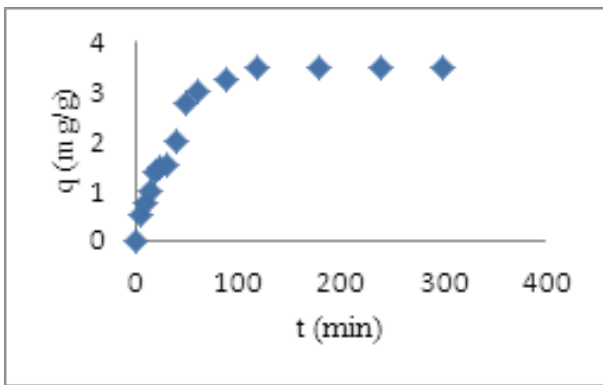


Figure 03: MG adsorption kinetic on AC

Table 01: Adsorption kinetics models and their linear and non linear forms

Applied model	Non Linear form	Linear form	Reference
Pseudo-first order			
Pseudo-first order (type1)	$q = q_e(1 - e^{-k_1 t})$	$\log(q_e - q) - \log(q_e) = -\frac{k_1 t}{2.303}$	(12)
Pseudo-first order (type2)		$\ln(q_e - q) - \ln(q_e) = -k_1 t$	(13)
Pseudo-first order (type3)	$C = C_0 e^{-k_1 t}$	$\ln\left(\frac{C}{C_0}\right) = -k_1 t$	(14)
Pseudo-first order (type4) Helfrich		$\ln\left[1 - \frac{C_0 - C}{C_0 - C_e}\right] = -k_1 t$	(15)
Pseudo-second order			
Pseudo-second order type 1		$\frac{t}{q} = \frac{1}{k_2 q_e^2} + \frac{t}{q_e}$	(16)

Pseudo-second order type 2		$\frac{1}{q} = \frac{1}{k_2 q_e^2 t} + \frac{1}{q_e}$	(17)
Pseudo-second order type 3	$q = q_e \left[1 - \frac{1}{1 + k_2 t} \right]$	$\left(\frac{1}{q} - \frac{1}{q_e} \right) q_e^2 = \frac{1}{k_2 t}$	(18)
Pseudo-second order type 4		$q = q_e - \left(\frac{1}{k_2 q_e} \right) \frac{q}{t}$	(19)
Pseudo-second order type 5		$(q - q_e) q_e = - \left(\frac{1}{k_2} \right) \frac{q}{t}$	(20)
Pseudo-second order type 6		$\frac{q}{t} = k_2 q_e^2 - k_2 q q_e$	(21)
Pseudo-second order type 7		$\frac{1}{q_e - q} - \frac{1}{q_e} = k_2 t$	(22)
Pseudo-second order type 8		$\frac{1}{t} = k_2 q_e^2 \left(\frac{1}{q} \right) - k_2 q_e$	(23)
Pseudo-second order type 9		$\frac{q_e}{q_e - q} - 1 = k_2 t$	(24)
Pseudo-second order type 10		$\frac{\theta}{1 - \theta} = k_2 t$	(25)
Pseudo-second order type 11		$C - \frac{1}{C_0} = k_2 t$	(26)
Pseudo-second order type 12		$\frac{1}{C} - \frac{1}{C_0} = k_2 t$	(27)
Pseudo-second order type 13		$\frac{1}{C_0 - C} = k_2 t + b$	(28)
Esquivel model (type 1)	$q = q_e \left(\frac{t}{t + K_E} \right)$	$\frac{1}{q} = \left(\frac{K_E}{q_e} \right) \frac{1}{t} + \frac{1}{q_e}$	(9)
Esquivel model (type 2)	$q = q_e \left(\frac{t}{t + K_E} \right)$	$\left(\frac{1}{q} - \frac{1}{q_e} \right) q_e = K_E \frac{1}{t}$	(9)
Pseudo-third order	$q = q_e \left[1 - \frac{1}{(1 + 2k_3 t)^{1/2}} \right]$	$\frac{1}{q^2} = \frac{1}{q_e^2} + k_3 t$	(10)
Elovich (type 1)	$\frac{dq}{dt} = k_4 \exp(-k_5 q)$	$q = k_5 \ln(k_5 k_4) + k_5 \ln(t)$	(11)
Elovich (Roginsky-Zeldovich) (type 2)	$\frac{dq}{dt} = k_7 \exp(k_6 q)$	$q = (1/k_6) \ln(k_6 k_7) + (1/k_6) \ln(t)$	(28)

Where k_1 is pseudo-first order rate constant (min^{-1}), k_2 is pseudo-second order rate constant ($\text{g}/(\text{mg min})$), k_3 is pseudo-third order rate constant ($\text{g}^2/(\text{mg}^2 \text{min})$), K_E is esquivel rate constant (min), k_4 is elovich rate constant

($\text{mg}/(\text{g min})$), k_5 is extent of surface coverage and activation energy of the process (g/mg), k_6 extent of surface coverage and activation energy of the process (g/mg), k_7 elovich rate constant ($\text{mg}/(\text{g min})$), q_e is amount of

adsorption at equilibrium (mg/g), and θ dimensionless parameter ($=q/q_e$). A non-linear and linear fitting procedure

using Excel and Origin software were used respectively. The constants of all models were given in Table 2.

Table 02: A Linear and non-linear Pseudo-first order, Pseudo-second order, Esquivel, Pseudo-third order and Elovich kinetics constants related to the adsorption of MG onto AC

Model	Linear Method	Non-linear Method
	Pseudo-first order (type 1)	Non-linear Pseudo-first order (type 1)
q_e	3.501	3.600
K_1	0.001	0.023
R^2	0.843	0.983
Equation	$\log(q_e - q) - \log(q_e) = -0.015 * t$	$q = 3.600 * (1 - \exp(-0.023 * t))$
	Pseudo-first order (type 2)	
q_e	3.501	
K_1	0.034	
R^2	0.843	
Equation	$\ln(q_e - q) - \ln(q_e) = -0.034 * t$	
	Pseudo-first order (type 3)	Non-linear Pseudo-first order (type 2)
C_0	5	5
K_1	0.005	0.011
R^2	0.419	0.773
Equation	$\ln(C/C_0) = -0.005 * t$	$C = 5 * \exp(-0.011 * t)$
	Pseudo-first order (type 4)	
C_0	5	
K_1	0.006	
R^2	0.457	
Equation	$\ln(1 - ((C_0 - C)/(C_0 - C_e)) = -0.006 * t$	
	Pseudo-second order (type 1)	
q_e	4.081	4.297
K_2	0.007	0.025
R^2	0.986	0.963
Equation	$(t/q) = 0.245 * t + 8.959$	$q = 4.297 * (1 - (1/(1 + 0.025 * t)))$
	Pseudo-second order (type 2)	
q_e	3.876	
K_2	335.449	
R^2	0.974	
Equation	$1/q = 9.303 * (1/t) + 0.258$	
	Pseudo-second order (type 3)	
q_e	3.610	
K_2	0.008	
R^2	0.974	
Equation	$((1/q) - (1/q_e)) * (q_e^2) = 118.9 * (1/t)$	
	Pseudo-second order (type 4)	
q_e	4.279	
K_2	0.006	
R^2	0.821	
Equation	$q = -41.28 * (q/t) + 4.272$	
	Pseudo-second order (type 5)	
q_e	4.210	
K_2	0.001	
R^2	0.821	
Equation	$(q - q_e) * q_e = -169.000 * (q/t)$	
	Pseudo-second order (type 6)	
q_e	4.895	
K_2	0.004	
R^2	0.821	
Equation	$(q/t) = -0.019 * q + 0.093$	
	Pseudo-second order (type 7)	
q_e	3.69	

K_2	0.021	
R^2	0.843	
Equation	$(1/(qe-q))-(1/qe) = 0.021*t$	
Pseudo-second order (type 8)		
q_e	4.000	
K_2	0.006	
R^2	0.974	
Equation	$1/t=0.104*(1/q)-0.026$	
Pseudo-second order (type 9)		
q_e	3.690	
K_2	0.078	
R^2	0.843	
Equation	$(qe/(qe-q))-1 = 0.078*t$	
Pseudo-second order (type 10)		
q_e	3.690	
K_2	0.078	
R^2	0.843	
Equation	$(\theta/(1-\theta)) = 0.078*t$	
Pseudo-second order (type 11)		
C_0	5.000	
K_2	0.009	
R^2	-4.210	
Equation	$(C-(1/C_0)) = 0.009*t$	
Pseudo-second order (type 12)		
C_0	5.000	
K_2	0.002	
R^2	0.655	
Equation	$((1/C)-(1/C_0)) = 0.002*t$	
Pseudo-second order (type 13)		
C_0	5.000	
K_2	0.003	
R^2	0.334	
Equation	$(1/(C_0-C)) = -0.003*t+0.910$	
Esquivel Model (type 1)		
q_e	3.876	4.296
K_E	36.058	39.642
R^2	0.974	0.963
Equation	$1/q = 9.303*(1/t)+0.258$	$q=2.145*(t/(t+7.322))$
Esquivel Model (type 2)		
q_e	3.610	
K_E	32.940	
R^2	0.974	
Equation	$((1/q)-(1/q_e))*q_e = 32.940*(1/t)$	
Pseudo-third order (type 1)		
q_e	0.957	
K_3	-0.005	
R^2	0.202	
Equation	$(1/q^2) = -0.005*t + 1.092$	
Elovich (type 1)		
K_4	0.300	
K_5	0.897	
R^2	0.937	
Equation	$q = 0.897*\ln(t) - 1.178$	
Elovich (type 2)		
K_4	0.241	
K_5	1.114	
R^2	0.937	
Equation	$q = 0.897*\ln(t) - 1.178$	

The optimization procedure required a defined error function in order to evaluate the fit of equation to the experimental data. In this part, the best-fitting equation is determined using the well-known special functions to

calculate the error deviation between experimental and predicted data. The mathematical equations of these error functions were illustrated in Table 3.

Table 03: Mathematical equations of error functions

Error functions	Equations	References
ARED	$ARED = \frac{100}{n} \sum_{i=1}^n \left \frac{q_{\text{exp}} - q_{\text{calc}}}{q_{\text{exp}}} \right _i$	(29)
ARE	$ARE = \frac{\sum q_{\text{calc}} - q_{\text{exp}} }{n}$	(30)
SAE = EABS	$SAE = EABS = \sum_{i=1}^n q_{\text{exp}} - q_{\text{calc}} $	(31)
ARS	$ARS = \sqrt{\frac{\sum \left[\frac{(q_{\text{exp}} - q_{\text{calc}})}{q_{\text{exp}}} \right]^2}{(n-1)}}$	(32)
MPSD	$MPSD = 100 \sqrt{\frac{\sum \left[\frac{(q_{\text{exp}} - q_{\text{calc}})}{q_{\text{exp}}} \right]^2}{n-p}}$	(33)
$\Delta q(\%) = 100 * ARS$	$\Delta q(\%) = 100 \sqrt{\frac{\sum \left[\frac{(q_{\text{exp}} - q_{\text{calc}})}{q_{\text{exp}}} \right]^2}{(n-1)}}$	(34)
SSE	$SSE = \sum (q_{\text{calc}} - q_{\text{exp}})^2$	(35)
MPSED	$MPSED = \sqrt{\frac{\sum \left[\frac{(q_{\text{exp}} - q_{\text{calc}})}{q_{\text{exp}}} \right]^2}{(n-p)}}$	(36)
HYBRID	$HYBRID = \frac{1}{(n-P)} \sum_{i=1}^n \left \frac{q_{\text{exp}} - q_{\text{calc}}}{q_{\text{exp}}} \right _i$	(37)

Where n is the number of experimental data points, q_{calc} is the predicted (calculated) quantity of MG adsorbed onto AC, q_{exp} is the experimental data, p is the number of parameters in each kinetic model, ARED is the average relative error deviation (dimensionless parameter), ARE is average relative error (dimensionless parameter, ARS is the average relative standard error (dimensionless parameter), HYBRID is the hybrid fractional error function

(dimensionless parameter), MPSD Marquardt's is the percent standard deviation (dimensionless parameter), MPSED Marquardt's is the percent standard deviation (dimensionless parameter), SAE=EABS is the sum of absolute error (mg/g), SSE is the sum of the squares of the errors (mg/g)², and $\Delta q(\%)$ is the normalized standard deviation (mg/g). The constants of all error analysis are represented in Table 4.

Table 04: Error deviation data related to the MG adsorption onto AC using most commonly used functions

Error functions	ARED	SAE = EABS	MPSED	SSE	HYBRID	ARE	ARS	$\Delta q(\%) = 100 * ARS$	MPSD
Linear Pseudo-first order type 1	88,296	27,01	0,957	62,345	1,03	0,883	0,92	91,985	139,372
Linear Pseudo-first order type 2	16,785	3,238	0,262	1,577	0,196	0,168	0,252	25,18	30,21

Non Pseudo-first order type 1	Linear									
		7,418	1,878	0,104	0,372	0,086	0,074	0,1	10,035	12,491
Linear second type 1	Pseudo-order	53,541	15,202	0,599	19,59	0,624	0,535	0,575	57,557	81,729
Linear second type 2	Pseudo-order	150,535	22,657	2,595	53,576	1,756	1,505	2,493	249,307	221,741
Linear second type 3	Pseudo-order	52,495	15,165	0,582	19,46	0,612	0,525	0,559	55,964	77,503
Linear second type 4	Pseudo-order	56,451	16,019	0,631	21,709	0,658	0,564	0,607	60,67	86,129
Linear second type 5	Pseudo-order	85,274	25,928	0,926	57,19	0,995	0,853	0,889	88,957	134,008
Linear second type 6	Pseudo-order	61,762	17,55	0,69	26,031	0,72	0,617	0,663	66,314	94,32
Linear pseudo-second type 7	order	18,704	6,069	0,221	3,961	0,218	0,187	0,213	21,277	33,97
Linear pseudo-second type 8	order	55,365	15,806	0,617	21,105	0,646	0,553	0,593	59,318	84,577
Linear pseudo-second type 9	order	41,342	6,62	0,644	5,594	0,482	0,413	0,619	61,912	63,684
Linear pseudo-second type 10	order	41,342	6,62	0,644	5,594	0,482	0,413	0,619	61,912	63,684
Non pseudo-second order type 1	linear	10,03	2,856	0,126	0,804	0,117	0,1	0,121	12,109	17,097
linear esquivel type 1		150,043	22,601	2,584	53,275	1,75	1,5	2,482	248,245	220,95
linear esquivel type 2		132,837	18,876	2,359	42,237	1,55	1,328	2,267	226,707	200,085
Non esquivel	linear	10,199	2,881	0,128	0,803	0,119	0,102	0,123	12,355	17,256
Linear elovich model type 1		15,884	3,891	0,208	1,353	0,185	0,159	0,2	19,979	23,516

Linear elovich									
model type 2	15,884	3,891	0,208	1,353	0,185	0,159	0,2	19,979	23,516

Adsorption kinetic data are the basic requirements for the design of adsorption systems. In order to optimize the design of a specific sorbate/sorbent system to remove MG from aqueous solution, it is important to establish the most appropriate correlation for the experimental kinetic data. Applicability of some statistical tools to predict optimum adsorption kinetic of MG onto AC after linear regression analysis showed that the highest R^2 value and the lowest ARE, ARE, SAE, ARS, MPSD, Δq , SSE, MSPED, and HYBRID values could be suitable and meaningful tools to predict the best-fitting equation models. The best fitting is determined based on the use of these functions to calculate the error deviation between experimental and predicted equilibrium adsorption kinetic data, after linear analysis. Hence, according to Table 4, it seems that the non linear pseudo-first order model was the most suitable models to satisfactorily describe the studied adsorption phenomenon. Therefore, based on these mentioned results, the best useful error estimation statistical tools should point out the non linear pseudo-first order model followed by non linear pseudo-second order and non linear esquivel as the best-fitting models.

4 CONCLUSION

AC was used for the MG adsorption in simulated aqueous solution. In batch mode, the adsorption was highly dependent on various operating parameters, such as: contact time, and pH. The obtained results allowed to establish the following optimal conditions: 120 min time contact and pH 6 leading to 70 % MG removal obtained at home temperature. The adsorption kinetic of MG onto AC can be better fitted by non linear pseudo-first-order model as compared to the non-linear pseudo-second-order model, linear pseudo-second-order model, linear pseudo first order, pseudo third order, elovich, and esquivel models. On the whole, the experimental results showed that AC is suitable adsorbent for the removal of MG dye.

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