

# Experimental parameter uncertainty in proton exchange membrane fuel cell modeling. Part I: Scatter in material parameterization

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## Abstract

Ever since modeling has become a mature part of proton exchange membrane fuel cell (PEMFC) research and development, it has been plagued by significant uncertainty lying in the detailed knowledge of material properties required. Experimental data published on several transport coefficients are scattered over orders of magnitude, even for the most extensively studied materials such as Nafion membranes, for instance. For PEMFC performance models to become predictive, high-quality input data is essential. In this bipartite paper series, we determine the most critical transport parameters for which accurate experimental characterization is required in order to enable performance prediction with sufficient confidence from small to large current densities. In the first part, a macro-homogeneous two-phase membrane-electrode assembly model is furnished with a comprehensive set of material parameterizations from the experimental and modeling literature. The computational model is applied to demonstrate the large spread in performance prediction resulting from experimentally measured or validated material parameterizations alone. The result of this is a ranking list of material properties, sorted by induced spread in the fuel cell performance curve. The three most influential parameters in this list stem from membrane properties: The Fickian diffusivity of dissolved water, the protonic conductivity and the electro-osmotic drag coefficient.

*Keywords:* polymer electrolyte membrane, fuel cell, model, parameterization, uncertainty analysis, experimental characterization

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## 1. Introduction

Fuel cell researchers that consider numerical modeling for their work are confronted with the difficult question which material properties to measure and plug into the models in order to obtain reliable simulation results. With finite financial and time budgets at hand, effort put into the characterization of membrane-electrode assembly (MEA) components must be prioritized. Is it worthwhile to determine the tortuosity of gas diffusion pathways through the porous layers with high resolution imaging, or should the focus be on precise measurements of water transport through the proton exchange membrane (PEM)? How will uncertainties in these parameters affect the predictive power of a detailed fuel cell model? MEA engineers face similar problems: For the efficient development of improved multifunctional materials for PEM fuel cells (PEMFCs), it is critical to know what quantity to optimize

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3 for. Will a thinner membrane yield significant improvements in fuel cell performance, or is it more productive to look  
4 into reducing contact resistances? Even when the MEA setup is fixed and known in detail, modelers are confronted  
5 with the problem of picking the right material parameterization among the many that have been proposed in the  
6 open literature. Published data on the most critical transport processes occurring in PEM fuel cells (*NB* for the  
7 very same materials) are sometimes scattered significantly, which raises the question which to adopt.  
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10 Recently, we have demonstrated how difficult it can be to reliably predict fuel cell performance with numerical  
11 modeling under the present spread of experimental material data in the literature [1]. The water diffusivity in the  
12 membrane [2, 3], electro-osmotic drag [4], protonic conductivity of the membrane [5] and the phase change rates of  
13 water [6] are some out of several properties that are required for state-of-the-art two-phase PEMFC models and for  
14 which measured data disagree largely; further ones are discussed in this paper.  
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17 Not many studies systematically address the variation in performance prediction of fuel cell models arising from  
18 uncertainty in their parameterization. Typically, a single specific constitutive *parameterization* is implemented for  
19 each material property, consisting of a functional relationship and a set of fit *parameters*. Only the latter have  
20 been varied to estimate the model sensitivity to them (e.g., [7, 8, 9]). Uncertainty resulting from the choice of the  
21 complete constitutive relationships, on the other hand, is largely unexplored. In Part I of this bipartite paper series,  
22 we intend to fill this gap.  
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28 In order to get a quantitative picture of the uncertainty in the performance prediction of state-of-the-art PEMFC  
29 models due to disagreement or scatter in the experimental literature data on fuel cell materials and processes, a model  
30 is needed that is both easily modifiable and numerically efficient. Here, we build upon our previously developed  
31 macro-homogeneous, steady-state, two-phase model of a five-layer MEA [10]. Being built for easy parameter  
32 substitution and fast execution of many different simulations (e.g., with different constitutive material properties),  
33 this one-dimensional model offers a **good** middle ground between model complexity, extensibility and computational  
34 efficiency. It models the coupled charge, heat and mass transport processes in through-plane direction within a  
35 differential PEMFC. The constitutive parameterization of the model is discussed in Sec. 2, reviewing the literature  
36 on several essential material properties. With this comprehensive database of proposed material parameterizations  
37 at hand, the scatter in fuel cell performance resulting from it will be presented in Sec. 3. This allows us to conclude  
38 with a sorted list of MEA properties which cause uncertainty in the resulting polarization curve and for which more  
39 reliable experimental characterization is therefore required. In Part II [11], we will extend this survey by carrying  
40 out extensive forward uncertainty propagation analyses to get a more rigorous insight into the relative importance  
41 of PEMFC material parameterizations.  
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## 52 **2. Constitutive model parameterization**

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54 For a mathematical description of the model, the reader is referred to Sec. S1 of the supplementary material, and  
55 a list of symbols is provided in Sec. S2. In the following, we assume that the readers have familiarized themselves  
56 with these.  
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Table 1: Electrochemical model parameters.

Parameter	Value	Source
$a_A$	$14 \text{ m}_{\text{Pt}}^2 \text{ cm}^{-3}$	[14]
$a_C$	$28 \text{ m}_{\text{Pt}}^2 \text{ cm}^{-3}$	[14]
$E_A$	$16 \text{ kJ mol}^{-1}$	[12]
$E_C$	$67 \text{ kJ mol}^{-1}$	[13]
$j_{A,C}^{0,\text{ref}}$	$0.54 \text{ A cm}_{\text{Pt}}^{-2}$	[12]
$j_C^{0,\text{ref}}$	$2.47 \times 10^{-8} \text{ A cm}_{\text{Pt}}^{-2}$	[13]
$\alpha_A, \tilde{\alpha}_A$	0.5	[12]
$\alpha_C, \tilde{\alpha}_C$	1	[13]
$\delta_A$	0	[12]
$\delta_C$	0.54	[13]
$\Delta H^{\text{ref}}$	$-285.83 \text{ kJ mol}^{-1}$	[15]
$\Delta S_A^{\text{ref}}$	$0.104 \text{ J mol}^{-1} \text{ K}^{-1}$	[16]
$\Delta S_C^{\text{ref}}$	$-163.3 \text{ J mol}^{-1} \text{ K}^{-1}$	[16]

### 2.1. Electrochemical reaction

The most common approach to model the reaction kinetics of PEM fuel cells is the Butler–Volmer equation

$$J_{A,C} = J_{A,C}^0 \left( \exp \left[ \frac{\alpha_{A,C} F \eta_{A,C}}{RT} \right] - \exp \left[ -\frac{\tilde{\alpha}_{A,C} F \eta_{A,C}}{RT} \right] \right) \quad (1)$$

where  $J_{A,C}^0 = j_{A,C}^0 a_{A,C} (1 - s)$  is the product of exchange current density, reactive surface area density and a correction factor  $(1 - s)$  for site blockage by liquid water.  $\alpha_{A,C}$  ( $\tilde{\alpha}_{A,C}$ ) are the forward (backward) half-reaction transfer coefficients. Measurement data for the hydrogen oxidation reaction suggests that they sum up to unity in the anode, at least for moderate current densities [12]. For the oxygen reduction reaction, on the other hand, there is no consensus in the literature on whether  $\alpha_C = \tilde{\alpha}_C = 1$  holds, or whether there is a doubling of Tafel slope at intermediate voltages ( $\alpha_C = 1$  to  $\alpha_C = 0.5$ ) [13]. We use the former. The exchange current densities can be written as [13, 12]

$$j_{A,C}^0 = j_{A,C}^{0,\text{ref}} \left( \frac{p_{A,C}}{P^{\text{ref}}} \right)^{\delta_{A,C}} \exp \left[ \frac{E_{A,C}}{R} \left( \frac{1}{T^{\text{ref}}} - \frac{1}{T} \right) \right] \quad (2)$$

where  $j_{A,C}^{0,\text{ref}}$  are the exchange current densities at reference conditions ( $P^{\text{ref}} = 1 \text{ atm}$  and  $T^{\text{ref}} = 80^\circ\text{C}$ ),  $\delta_{A,C}$  the kinetic reaction orders,  $E_{A,C}$  the half-reaction activation energies, and finally,  $p_A = p_{\text{H}_2}$  and  $p_C = p_{\text{O}_2}$  are the reactant gas partial pressures in the two electrodes. The electrochemical parameters of the model are summarized in Tab. 1.

### 2.2. Electronic and ionic conductivities

Constant values are used for the effective electronic conductivities  $\sigma_e^{\text{CL}} = 390 \text{ S m}^{-1}$  (for a catalyst layer with ionomer volume fraction  $\epsilon_i = 0.3$ ) [17] and  $\sigma_e^{\text{GDL}} = 450 \text{ S m}^{-1}$  (for a SGL 28 AA compressed by 1 MPa) [18].

The protonic conductivity of PFSA membranes and the ionomer phase in the catalyst layers has been the subject of more than 200 papers, with the majority focusing primarily on Nafion with equivalent weight (EW) 1100 [3]. For this reason, it is the first MEA material property that we pay particular attention to here, to examine the uncertainty in PEMFC performance prediction arising from scatter in experimental data on this material property. Being a strong function of the state of ionomer hydration,  $\sigma_p$  dominates the total ohmic losses in the cell in regions where the membrane is relatively dry. In Tab. 2, we summarize published parameterizations for Nafion 1100 EW,

Table 2: Review of protonic conductivities in vapor-equilibrated Nafion membranes. Below the largest  $\lambda$  at which  $\sigma_p(\lambda) = 0$ , the conductivity is set to vanish, which is omitted here for brevity. RT is short for room temperature.

Publication	Protonic conductivity [ $\text{S m}^{-1}$ ]	Temperature	Activation energy [ $\text{kJ mol}^{-1}$ ]	Membrane
Hsu et al., 1980 [19]	$16(f_w(\lambda) - 0.1)^{1.5}$	RT		EW 1050–1500
Springer et al., 1991 [20]	$0.5139 \max\{1, \lambda\} - 0.326$	30 °C	10.54	N117
Morris & Sun, 1993 [21]	$12.5(f_w(\lambda) - 0.06)^{1.95}$	23–100 °C		N117
Sone et al., 1996 [22]	$-0.145 + 1.57a - 4.55a^2 + 8.86a^3$	80 °C	1.3	N117
Eikerling et al., 1998 [23]	$0.07 + 7(f_w(\lambda)/f_w(22) - 0.1)$			N117
Thampan et al., 2000 [24]	$\sigma_0(\lambda, T)(f_w(\lambda) - f_w(1.9))^{1.5}$			EW 1100
Costamagna, 2001 [25]	$0.58\lambda - 0.5$	30 °C	10.54 (from [20])	N117
Edmondson & Fontanella, 2002 [26]	$27.2(f_w(\lambda) - 0.03)^{1.38}$	RT		N117
Kulikovsky, 2003 [27]	$0.5738\lambda - 0.7192$	80 °C		
Weber & Newman, 2004 [28]	$50(f_w(\lambda) - 0.06)^{1.5}$	30 °C	15	N117
Meier & Eigenberger, 2004 [29]	$0.46\lambda - 0.25$	25 °C	9.894	N117
Yang et al., 2004 [30]	$1.3 \times 10^{-5} \exp[14a^{0.2}]$	80–140 °C		N115
Choi et al., 2005 [31]	$\sigma_\Sigma(a, \lambda, T) + \sigma_G(a, \lambda, T) + \sigma_E(a, \lambda, T)$			
Fimrite et al., 2005 [32]	$\sigma_0(\lambda, T)(f_w(\lambda) - f_w(1.65))^{1.5}$			EW 1100
Hwang et al., 2009 [33]	$\begin{cases} 0.75(\lambda - 2.3) & \lambda < 5 \\ 0.41(\lambda - 5) + 3 & \lambda \geq 5 \end{cases}$	30 °C	10.54 (from [20])	N117
Maldonado et al., 2012 [34]	$-2.91 + 23.61a - 46.09a^2 + 40.98a^3$	80 °C	$13.9a^2 - 8.87a + 11.8$	N115
Zhao et al., 2012 [35]	$77(f_w(\lambda) - 0.1)^2$	80 °C		EW 1100

which can be (and have been) plugged into numerical models. Where an Arrhenius expression is used to account for temperature dependence, the reported activation energies are also given. Hsu et al. [19] originally carried a result over from percolation theory to express the protonic conductivity by a shifted power law, such that one can write

$$\sigma_p = M_i \sigma_0(T) \max\{f_w - f_0, 0\}^\beta \quad (3)$$

where

$$M_i = \frac{\epsilon_i}{\tau_i^2} \quad (4)$$

is the microstructure factor of the ionomer [36], with  $\epsilon_i = \tau_i = 1$  in the PEM and  $\epsilon_i = 0.3$ ,  $\tau_i = 1.4$  [37] in the catalyst layers (CLs). The water volume fraction in the hydrated ionomer is given by

$$f_w = \frac{\lambda V_w}{\lambda V_w + V_m}. \quad (5)$$

Later models by Thampan et al. [24] and Fimrite et al. [32] have extended the percolation-based conductivity by additionally expressing the prefactor  $\sigma_0(T)$  as a function of  $\lambda$ . Springer et al. [20], on the other hand, proposed a piecewise linear law in  $\lambda$  in their seminal modeling work – a simple correlation that is still widely used today. The phenomenological parameterizations used by Sone et al. [22], Yang et al. [30] and Maldonado et al. [34] use the water vapor activity (i.e., relative humidity)  $a$  to fit the observed ionic conductivity. To implement these into our model, the water activity in the bulk membrane is calculated by inverting the sorption isotherm  $\lambda(a)$  (see Sec. 2.7). Yang’s relationship is excluded from our following analysis though, as it predicts far larger conductivities than all others and is considered to be an outlier for this reason.

[FIGURE 1]

Figure 1: Proposed parameterizations for protonic conductivity as a function of water content in Nafion membranes. Solid lines denote explicit functional relationships  $\sigma_p(\lambda)$ , dashed lines are parametric curves  $\{\lambda(a), \sigma_p(a)\}$ , dotted lines are purely empirical fits which are piecewise linear in  $\lambda$ . RT is short for room temperature.

These proposed parameterizations of  $\sigma_p$  are juxtaposed in Fig. 1 to show that they vary considerably, although some spread may certainly partially be the result of different measurement temperatures. Despite characterizing the same class of PFSA membranes, the correlations by Hsu et al. [19], Morris & Sun [21], and Zhao et al. [35] predict far lower conductivity than the others. As noted by Weber et al. [28, 3], the percolation model fits most measured data best. Their coefficients ( $f_0 = 0.06$ ,  $\beta = 1.5$ ,  $\sigma_0 = 50 \text{ S m}^{-1} \times \exp[E_\sigma/R(1/T^{\text{ref}} - 1/T)]$ ,  $E_\sigma = 15 \text{ kJ mol}^{-1}$ ,  $T^{\text{ref}} = 30^\circ\text{C}$ ) are used here as the baseline parameterization of ionic conductivity.

### 2.3. Thermal conductivity

The through-plane thermal conductivity  $k$  of the individual layers plays a subordinate role in MEA modeling, as will be shown in Part II. We therefore restrict the discussion on  $k$  to a single baseline parameterization for each layer type. In humidified Nafion membranes, it can be approximated by a linear interpolation between the thermal conductivities of water and the dry membrane [38]:

$$k^{\text{PEM}} = f_w k_w + (1 - f_w) k_0^{\text{PEM}} \quad (6)$$

where

$$k_0^{\text{PEM}} = (0.451 - 0.286\bar{T}) \frac{\text{W}}{\text{m K}} \quad (7)$$

is a linear approximation to the thermal conductivity of dry Nafion, and

$$k_w = \left(1.6630\bar{T}^{-1.15} - 1.7781\bar{T}^{-3.4} + 1.1567\bar{T}^{-6.0} - 0.432115\bar{T}^{-7.6}\right) \frac{\text{W}}{\text{m K}} \quad (8)$$

is the internationally recommended correlation of the thermal conductivity of liquid water at 1 bar up to  $110^\circ\text{C}$  with  $\bar{T} = T/300 \text{ K}$  [39]. A detailed review of available experimental data on the thermal conductivity of the remaining porous media can be found in [40]. Here, the Maxwell–Eucken equation [41]

$$k = k_s \frac{2k_s + k_f - 2(k_s - k_f)\epsilon_p}{2k_s + k_f + (k_s - k_f)\epsilon_p} \quad (9)$$

is used, where  $k_s$  is the theoretical conductivity of the solid bulk material,  $k_f$  the conductivity of the fluid filling the pore space, and  $\epsilon_p$  denotes the pore volume fraction. In order to obtain the bulk conductivity  $k_s$  from effective conductivity measurements on real MEA materials, one can invert Eq. 9 using  $k_f = 0$  or  $k_f \approx 0.003 \text{ W m}^{-1} \text{ K}^{-1}$ , depending on whether the experiment was conducted in vacuum or air. For vacuum, this yields

$$k_s = k_0 \frac{2 + \epsilon_p}{2(1 - \epsilon_p)} \quad (10)$$

where  $k_0$  is the effectively measured thermal conductivity of the dry porous layer. Alhazmi et al. [42, 43] have conducted the most comprehensive measurements of thermal conductivity of dry gas diffusion layers (GDLs),  $k_0^{\text{GDL}}$ , as a function of temperature, clamping pressure and polytetrafluoroethylene content under vacuum conditions. We have fitted the following functional relationship to their data for the SGL 10 series, assuming that the correlations with temperature  $T$  and clamping pressure  $P_{\text{cl}}$  are independent:

$$k_0^{\text{GDL}} = (0.776 - 0.430\bar{T}) \left( \frac{P_{\text{cl}}}{P^{\text{ref}}} \right)^{0.21} \frac{\text{W}}{\text{m K}} \quad (11)$$

where  $P^{\text{ref}} = 1$  bar. A power law was chosen for the pressure dependence because contact resistivities follow the same relationship (see Sec. 2.13). For the CLs a constant value is used for the dry thermal conductivity,  $k_0^{\text{CL}} = 0.22 \text{ W m}^{-1} \text{ K}^{-1}$  [44], which lies approximately in the middle of the reported range of values in the literature. Humidity dependence is added through Eqs. 9 and 12 in the CLs just like in the GDLs.

For the fluid conductivity  $k_{\text{f}}$ , we assume that liquid water and the gas mixture form transport channels in through-plane direction along which heat is transported in parallel:

$$k_{\text{f}} = s k_{\text{w}} + (1 - s) k_{\text{g}}. \quad (12)$$

This choice is motivated by measurements on humidified GDLs [45, 46], which revealed an overall increase in effective thermal conductivity of as much as 50% at  $s = 0.25$  and even more at higher saturations.

The gas phase conductivity  $k_{\text{g}}$  depends on the gas composition. We model it as a linear combination of the conductivities of the individual gas components, with the mole fraction as weights and species conductivities  $k_X$  from [47, 48]:

$$k_{\text{g}} = \sum_X y_X k_X, \quad X = \text{H}_2, \text{O}_2, \text{H}_2\text{O}, \text{N}_2. \quad (13)$$

#### 2.4. Water diffusivity in the ionomer

Water diffusion within and through the membrane has been a topic of extensive research over the past decades and has been thoroughly reviewed in the works of Kusoglu & Weber [2, 3]. The numerous studies carried out to measure and parameterize the water diffusivity vary in the transport mode considered (driven by a pressure gradient, concentration gradient, temperature gradient, or chemical potential gradient), in the experimental technique (NMR, QENS, conductivity), in the types of diffusion coefficients being measured, in the studied regime (Fickian vs. non-Fickian, steady-state vs. transient), and even in the data separation and interpretation (e.g., correcting for measurement device resistance and interfacial resistance or not). As a result, the reported diffusion coefficients of dissolved water in Nafion are scattered over one to two orders of magnitude, which calls for a quantitative analysis of the modeling uncertainty associated with this material property.

As shown in Fig. 2a,c, the intradiffusivity  $D_{\text{i}}$  is a monotonically increasing function of  $\lambda$  [3], whereas most studies agree that the Fickian diffusivity  $D_{\text{F}}$  exhibits a pronounced peak around  $\lambda \approx 2 - 4$  (Fig. 2b,d). This local maximum in  $D_{\text{F}}$  stems from the *Darken factor*, which relates the two diffusivities [20]:

$$D_{\text{F}} = \frac{\partial \ln a}{\partial \ln \lambda} D_{\text{i}} = \frac{\lambda}{a} \left( \frac{\partial \lambda}{\partial a} \right)^{-1} D_{\text{i}} \quad (14)$$

[FIGURE 2]

Figure 2: Water diffusivity in Nafion membranes as a function of water content. (a) Experimentally measured intradiffusivity. (b) Experimentally measured Fickean diffusivity. Only direct measurements that don't rely on Eq. 14 are shown. (c) Parametric expressions for intradiffusivity. (d) Parametric expressions for Fickean diffusivity. In (c) and (d), solid lines represent explicit functions  $D(\lambda)$ , dashed lines are parametric curves  $\{\lambda(a), D(a)\}$  as a function of activity  $a$  and dotted lines are converted from intradiffusivity using the Darken factor (from Springer's isotherm [20] where none is given in the respective publication). All ordinates reach up to the self-diffusivity of water at 25 °C,  $2.3 \times 10^{-5} \text{ cm}^2 \text{ s}^{-1}$  [70, 71].

Table 3: Review of water diffusion coefficients in vapor-equilibrated Nafion membranes.

Publication	Water diffusivity $D_i$ or $D_F$ [ $10^{-6} \text{ cm}^2 \text{ s}^{-1}$ ]	Temperature	Activation energy	Membrane
Springer et al., 1991 [20]	$D_F = 2.563 - 0.33\lambda + 0.0264\lambda^2 - 0.000671\lambda^3$ for $\lambda > 4$	30 °C	20.1 kJ mol <sup>-1</sup>	N117
Fuller, 1992 [63]	$D_i = 3.5 \times 10^4 \lambda / 14$	$\infty$	20.3 kJ mol <sup>-1</sup>	N117
Motupally et al., 2000 [64, 49]	$D_i = 0.631(-0.501 + \lambda - 0.0209\lambda^2)$	30 °C		N117
	$D_F = \begin{cases} 3100\lambda(\exp[0.28\lambda] - 1) & \lambda < 3 \\ 417\lambda(161 \exp[-\lambda] + 1) & \lambda \geq 3 \end{cases}$	$\infty$	20.3 kJ mol <sup>-1</sup>	N115
Ye & LeVan, 2003 [68]	$D_F = 6.76\tilde{p}^{1.5} / (5.9673 - 8.9472\tilde{p} + 4.0622\tilde{p}^2)$ , $\tilde{p} = p_{\text{H}_2\text{O}} / \text{kPa}$	23.5–25 °C		
Kulikovsky, 2003 [27]	$D_F = 4.1(\lambda/25)^{0.15}(1 + \tanh[(\lambda - 2.5)/1.4])$	80 °C		N117,125
Weber & Newman, 2004 [28]	$D_i = 18f_w$	30 °C	20 kJ mol <sup>-1</sup>	N112,115,117
Ge et al., 2005 [65]	$D_i = 27.2f_w$	30 °C	20.1 kJ mol <sup>-1</sup>	N112,115,117
Zhao et al., 2011 [58]	$D_F$ : piecewise exponential in $\lambda$ , interpolation in $T$	23–70 °C		N1110
Myles et al., 2011 [66]	$D_i = 6.0667a + 3.1333a^2$ $D_i = 7.2167a + 3.4833a^2$	50 °C 60 °C		N117 N117
Mittelsteadt & Staser, 2011 [62]	$D_F = \begin{cases} 732 \exp[0.12\lambda] + 5.41 \exp[1.44\lambda] & \lambda < 4 \\ 1.58 \times 10^{11} \exp[-4.66\lambda] + 1450 \exp[0.04\lambda] & \lambda \geq 4 \end{cases}$	$\infty$	20.3 kJ mol <sup>-1</sup>	N115,117 NR211,212
Caulk et al., 2012 [69]	$D_F = 0.032 \text{ mol cm}^{-3} \text{ bar}^{-1} \times \exp[3.4a] P_{\text{sat}} V_m (\partial a / \partial \lambda)$	90 °C	22 kJ mol <sup>-1</sup>	
Lokkiluoto & Gasik, 2013 [67]	Phenomenological expressions for $D_F(\lambda)$ and $D_i(\lambda)$	30 °C		N117
Vetter & Schumacher, 2018 [10, 62]	$D_F = \frac{3.842\lambda^3 - 32.03\lambda^2 + 67.74\lambda}{\lambda^3 - 2.115\lambda^2 - 33.013\lambda + 103.37}$	80 °C	20 kJ mol <sup>-1</sup>	N115,117 NR211,212

where  $\lambda(a)$  is the vapor uptake isotherm (see Sec. 2.7). Since  $\lambda(a)$  is relatively flat at  $\lambda \approx 2 - 4$ , the Fickean diffusivity  $D_F$  peaks there. However, the existence of a local maximum in  $D_F$  is still a subject of controversy [59, 69, 72].

$D_F$  can either be measured directly or calculated from  $D_i$  using Eq. 14. A compilation of published correlations between either of them and the level of hydration (via  $a$ ,  $\lambda$  or  $f_w$ ) is given in Tab. 3 and Fig. 2c,d. Given  $D_F$ , the effective diffusivity required for the water flux equation (supplementary Eq. S9) can be calculated as

$$D_\lambda = M_i D_F \quad (15)$$

with  $M_i$  from Eq. 4. We have recently identified the parametric expression given in the last row of Tab. 3 as the most plausible and convenient for numerical modeling of water transport within Nafion membranes [10]. It is used as the baseline parameterization also for the present study.

For the temperature dependence, the Arrhenius equation is assumed to hold almost unanimously in experimental and modeling studies, even though a Speedy–Angell power law fits experimental data better [71]. The research

[FIGURE 3]

Figure 3: Activation energy of the water diffusion coefficient in Nafion membranes as a function of membrane hydration.

Table 4: Review of electro-osmotic drag coefficients in vapor-equilibrated Nafion membranes.

Publication	Electro-osmotic drag coefficient $\xi$	Temperature	Membrane
Springer et al., 1991 [20]	$2.5\lambda/22$	30 °C	N117
Fuller & Newman, 1992 [79]	$-BCa \exp[-Ca]$ , $B = -3.7206$ , $C = 1.339$	25–37.5 °C	N117
Fuller, 1992 [63]	$((0.35\lambda)^{-4} + 1.47^{-4})^{-1/4}$	25–37.5 °C	N117
Eikerling et al., 1998 [23]	$1.2 + 1.3(f_w(\lambda)/f_w(22))^5$		N117
Dutta et al., 2001 [80]	$0.0029\lambda^2 + 0.05\lambda$		
Kulikovsky, 2003 [27, 81]	$\max\{1, 0.117\lambda - 0.0544\}$	80 °C	
Weber & Newman, 2004 [28]	$\min\{1, \lambda\}$	30 °C	
Meier & Eigenberger, 2004 [29]	$1 + 0.028\lambda + 0.0026\lambda^2$	25 °C	
Ge et al., 2006 [82]	polynomial in $\lambda$ , linear interpolation in $T$	30–80 °C	N117
Lokkiluoto & Gasik, 2013 [67]	$\sqrt{\lambda}/2$		N117

[FIGURE 4]

Figure 4: Electro-osmotic drag coefficient in Nafion membranes as a function of water content. (a) Experimental data. Open symbols denote measurements on vapor-equilibrated, closed symbols on liquid-equilibrated (and subsequently dried) membranes. (b) Proposed parameterizations. Solid lines denote explicit functional relationships  $\xi(\lambda)$ , the dashed line is a parametric curve  $\{\lambda(a), \xi(a)\}$ .

group of Eisenberg [76, 73] was the first to estimate the activation energy  $E_d \approx 20 \text{ kJ mol}^{-1}$  for a Nafion membrane immersed in liquid water, which is close to the value for water self-diffusion ( $19.2 \text{ kJ mol}^{-1}$  [75]). This value has subsequently been used by almost all researchers to fit their experimental data and to model  $D_F(\lambda, T)$ , neglecting that it might depend on the mode and level of membrane hydration. Later measurements have brought up a variety of other values: 12–16 [56], 22 [69], 24 [57] and  $28 \text{ kJ mol}^{-1}$  [58]. Gong et al. [51], Kreuer et al. [54] and Guillermo et al. [74] investigated the dependency of  $E_d$  on hydration. Their findings are juxtaposed in Fig. 3. A polynomial least-squares fit to the available data yields

$$E_d = (38.0f_w^2 - 47.9f_w + 29.2) \text{ kJ mol}^{-1}, \quad (16)$$

which is the activation energy implemented in the present model to adapt the rational polynomial in  $\lambda$  from Tab. 3 to arbitrary temperatures in the baseline parameterization.

For a Nafion membrane of type NR211, which is considered for the baseline simulation,  $\rho_m = 1.97 \text{ g cm}^{-3}$  [77] and  $m_m = 1020 \text{ g mol}^{-1}$  [78], such that the equivalent volume is given by  $V_m = m_m/\rho_m \approx 517.8 \text{ cm}^3 \text{ mol}^{-1}$ .

### 2.5. Electro-osmosis

Electro-osmotic drag in Nafion has been another subject of controversial debate owing to the complexity of coupled ion/water transport and the difficulty in measuring it [4, 3]. Springer et al. [20] proposed a linear approximation for the electro-osmotic drag coefficient

$$\xi = \xi_1 \lambda / \lambda_1 \quad (17)$$

based on a single data point for Nafion 117 immersed in liquid water:  $(\xi_1, \lambda_1) = (2.5, 22)$ . This relationship is still widely used in MEA modeling. Later measurements [83] have led researchers to conclude that vapor-equilibrated PFSA membranes are more appropriately characterized by  $\xi \equiv 1$ .

Fig. 4 and Tab. 4 provide an overview over published measurements and parameterizations of the electro-



osmotic drag coefficient, from which it becomes clear that a conclusive reliable correlation  $\xi(\lambda, T)$  is still missing. Care must be taken when interpreting Fig. 4a, because some of these measurements were performed on liquid-equilibrated membranes with different degrees of drying to obtain mid-range water contents. These data might not be representative for vapor sorption. Agreement has not even been found on the general trend of the drag coefficient, with data suggesting an increasing, invariant, or even decreasing value of  $\xi$  with increasing membrane hydration.

A frequently used class of parameterizations are piecewise linear functions [27, 98, 99], which take the form

$$\xi = \begin{cases} \lambda & 0 \leq \lambda < 1 \\ 1 & 1 \leq \lambda \leq \lambda_v(1) \\ 1 + (\xi_1 - 1) \frac{\lambda - \lambda_v(1)}{\lambda_1 - \lambda_v(1)} & \lambda_v(1) < \lambda \leq \lambda_1 \end{cases} \quad (18)$$

when taking the theoretical upper bound  $\xi \leq \lambda$  [79] into account. Here,  $\lambda_1$  is the water content for a liquid-equilibrated membrane,  $\lambda_v(1)$  the water content for a vapor-equilibrated membrane at unit vapor activity (see Sec. 2.7), and  $\xi_1$  the drag coefficient for a liquid-equilibrated membrane. Eq. 18 is a simple attempt at taking Schroeder's paradox ( $\lambda_1 \gg \lambda_v(1)$ , see Sec. 2.7) into account in a global parameterization  $\xi(\lambda)$ . Weber & Newman [28] suggested to use the Arrhenius equation

$$\xi_1 = 2.55 \exp \left[ \frac{E_\xi}{R} \left( \frac{1}{T^{\text{ref}}} - \frac{1}{T} \right) \right] \quad (19)$$

with  $T^{\text{ref}} = 30^\circ\text{C}$  and an activation energy of  $E_\xi = 4 \text{ kJ mol}^{-1}$  to model the temperature dependence of  $\xi_1$ .

In summary, Springer's linear law (Eq. 17) matches the widely scattered experimental data best. It is therefore used here as the baseline parameterization, together with Eq. 19 for temperature dependence.

## 2.6. Thermo-osmosis

Thermo-osmotic transport of dissolved water is an entropic effect and occurs from cold to hot regions in hydrophilic membranes, i.e., in direction of the positive temperature gradient [100, 101]. The thermo-osmotic transport coefficient  $D_T$  is thus negative for typical PFSA membranes [102, 103]. Kim & Mench [103] studied the temperature dependence of thermo-osmosis in liquid-equilibrated Nafion 112, Flemion SH50 and Gore-Select membranes and reported that the Arrhenius equation holds with an activation energy that is indistinguishable from that of concentration gradient-driven diffusion. This suggests that the transport mechanism for thermo-osmosis might be the same as for diffusion, albeit with different driving force. We therefore fitted their measured values of the thermo-osmotic transport coefficient in N112 using  $T^{\text{ref}} = 80^\circ\text{C}$  and the activation energy  $E_d$  from Eq. 16 and obtained

$$D_T(T) = -7.2 \times 10^{-7} \text{ mol m}^{-1} \text{ s}^{-1} \text{ K}^{-1} \exp \left[ \frac{E_d}{R} \left( \frac{1}{T^{\text{ref}}} - \frac{1}{T} \right) \right]. \quad (20)$$

Measurements of  $D_T$  for *vapor-equilibrated* PFSA membranes are still missing in the literature. The magnitude of  $D_T$  in various anion exchange membranes has been reported to increase with growing water content [104, 105], and if the transport mechanism of thermo-osmosis is indeed similar to diffusion, it appears natural to assume that

[FIGURE 5]

Figure 5: Vapor sorption isotherms for Nafion membranes. Symbols denote data from experimental measurements, lines represent fitted curves at  $T = 30^\circ\text{C}$  (solid) and  $T = 80^\circ\text{C}$  (dashed).

$D_T \rightarrow 0$  as  $\lambda \rightarrow 0$ . Therefore, in order not to overestimate the effect of thermo-osmosis much in the model for vapor-equilibrated membranes, a linear approximation between zero and the reported coefficient for liquid-equilibrated Nafion is used here, analogous to Springer’s linear interpolation of the electro-osmotic drag coefficient in Eq. 17:

$$D_T(\lambda, T) = \frac{\lambda}{\lambda_l} D_T(T). \quad (21)$$

### 2.7. Membrane hydration

In the parameterization of the equilibrium water content of the ionomer (see supplementary Eq. S22), we account for simultaneous partial contact of the ionomer with liquid water and water vapor as well as for Schroeder’s paradox by writing

$$\lambda_{\text{eq}} = s\lambda_l + (1 - s)\lambda_v \quad (22)$$

where  $\lambda_l$  ( $\lambda_v$ ) denotes the hydration number when the membrane is liquid-equilibrated (vapor-equilibrated). The vapor sorption isotherm  $\lambda_v(a)$  has been the subject of a vast number of experimental and theoretical studies. A chronological listing of proposed explicit and implicit functional relationships is given in Tab. 5. Historically, polynomial fits to experimental data (of degree 3 or higher in  $a$ ) have been popular in PEMFC modeling, starting with Springer’s and Hinatsu’s measurements at 30 and 80 °C, respectively. Temperature dependence can be introduced by linearly interpolating between these two polynomials. In Fig. 5a, the different polynomials are plotted together with a selection of experimental data points from the open literature, showing that there is an uncertainty band of around 2 in width over the entire activity range, with a tendency to widen toward saturation. The water vapor activity is calculated as

$$a = \frac{y_{\text{H}_2\text{O}}}{y_{\text{sat}}} = \frac{p_{\text{H}_2\text{O}}}{P_{\text{sat}}} \quad (23)$$

using [124]

$$P_{\text{sat}} = P_c \exp\left[\frac{T_c}{T} \left(-7.8595\hat{T} + 1.8441\hat{T}^{1.5} - 11.787\hat{T}^3 + 22.681\hat{T}^{3.5} - 15.962\hat{T}^4 + 1.8012\hat{T}^{7.5}\right)\right] \quad (24)$$

with  $\hat{T} = 1 - T/T_c$ , where  $P_c = 22.064$  MPa and  $T_c = 647.096$  K are the critical pressure and temperature of water, respectively.

Futerko & Hsing have initiated the semi-empirical modeling of vapor sorption in Nafion. Several of the developed models are based on modified Flory–Huggins solution theory [107, 111, 129], or multilayer adsorption theory, in its finite-layer Brunauer–Emmett–Teller (BET) form [24, 130, 112], in the infinite-layer limit [131], as a superposition of different sorption modes [122], or extended by elastic swelling [109, 115]. Other swelling models use thermal, chemical and elastic equilibrium assumptions [117, 119]. Many of these sorption models are based upon an additive decomposition of the total water content into a chemically bound (subscript b in Tab. 5) or clustered (c) part and a free (f) part.

Table 5: Review of water vapor sorption isotherms of Nafion membranes. Only parameterizations which apply to the full vapor activity range  $a \in [0, 1]$  are included.

Publication	Equilibrium water content $\lambda_v$	Temperature	Membrane
Springer et al., 1991 [20]	$0.043 + 17.81a - 39.85a^2 + 36.0a^3$	30 °C	N117
Hinatsu et al., 1994 [106]	$0.300 + 10.8a - 16.0a^2 + 14.1a^3$	80 °C	N117,125
Futerko & Hsing, 1999 [107]	numerical solution of the implicit equation $(1 - f_{mb}) \exp[(1 - 1/r)f_{mb} + \chi f_{mb}^2] = a$ with $f_{mb} = (r + \lambda_b)/(r + \lambda_v)$ , $\lambda_b = Ka/(1 + Ka)$ , $r = V_m/V_w$ , $\chi = 1.936 - 2.18 \text{ kJ mol}^{-1}/RT$ , $K = 0.0256 \exp[22.4 \text{ kJ mol}^{-1}/RT]$		N117
Thampan et al., 2000 [24]	$\lambda_v = \lambda_{BET} = \lambda_m Ka/(1 - a) \times (1 - a^n - na^n(1 - a))/(Ka(1 - a^n) + 1 - a)$ with $\lambda_m = 1.8$ , $K = 150$ , $n = 13.5$	25–30 °C	N117
Meyers & Newman, 2002 [108]	$\lambda_v = \lambda_2(1 + \exp[0.3 - \lambda_2])$ where $\lambda_2$ solves the coupled implicit equations		
and	$\begin{cases} \lambda_3 \exp[f_1 \lambda_3 + f_2 \lambda_2] = K_1(1 - \lambda_3)(\lambda_2 - \lambda_3) \\ K_2(\lambda_2 - \lambda_3) \exp[f_2 \lambda_3 + f_3 \lambda_2] = a \end{cases}$		
Weber & Newman, 2004 [28]	with $K_1 = 100$ , $K_2 = 0.217 \exp[1 \text{ kJ mol}^{-1}/R \times (1/303.15 \text{ K} - 1/T)]$ , $f_1 = 2(m_{22} - 2m_{31} - 2m_{23})/m_m$ , $f_2 = 2(m_{23} - m_{22})/m_m$ , $f_3 = 2m_{22}/m_m$ , $m_{22} = -41.7 \text{ g mol}^{-1}$ , $m_{23} = -52.0 \text{ g mol}^{-1}$ , $m_{31} = -3721.6 \text{ g mol}^{-1}$		N117
Kulikovsky, 2003 [27]	$0.3 + 6a(1 - \tanh[a - 0.5]) + 3.9\sqrt{a}(1 + \tanh[(a - 0.89)/0.23])$	80 °C	N117,125
Choi & Datta, 2003 [109]	numerical solution of the implicit equation $(\lambda_v - \lambda_b)/(1 + \lambda_v - \lambda_b) = a \exp[-V_w P/RT]$ with $\lambda_b = \lambda_{BET}$ , $\lambda_m = 1.8$ , $K = 100$ , $n = 5$ , $P = \kappa f_w - a_p \sigma \cos \theta / f_w$ , $a_p = 210 \text{ m}^2 \text{ cm}^{-3}$ , $\kappa = 183 \text{ atm}$ , $\sigma = 72.1 \text{ mN m}^{-1}$ , $\theta = (116 - 7.15a + 28.4a^2 - 39.3a^3)^\circ$ (fit to data from [110])		N117
Meier & Eigenberger, 2004 [29]	$17.81a - 39.85a^2 + 35a^3$		
Choi et al., 2005 [111]	numerical solution of the implicit equation $(1 - f_{mb}) \exp[(1 - 1/r)f_{mb} + \chi f_{mb}^2] = a \exp[-V_w P/RT]$ with $f_{mb} = (r + \lambda_b)/(r + \lambda_v)$ , $\lambda_b = \lambda_{BET}$ , $\lambda_m = 1.8$ , $K = 1000$ , $n = 5$ , $P = 2E(f_m^{1/3} - f_w^{1/3})/9 - a_p \sigma \cos \theta / f_w$ , $r = V_m/V_w$ , $f_m = 1 - f_w$ , $a_p = 210 \text{ m}^2 \text{ cm}^{-3}$ , $\sigma = 72.1 \text{ mN m}^{-1}$ , $\theta = 98^\circ$		EW 1100
Takata et al., 2007 [112]	$V_m/m_w \times A_L B_L p_{H_2O}/(1 + A_L p_{H_2O}) \times (1 + (n - 1)(A_C p_{H_2O})^{n-1})$ with $A_L = 1.53 \times 10^{-10} \text{ Pa}^{-1} \times \exp[39 \text{ kJ mol}^{-1}/RT]$ , $B_L = 0.160 \text{ g cm}^{-3}$ , $n = 5.15$ $A_C = 2.40 \times 10^{-12} \text{ Pa}^{-1} \times \exp[46 \text{ kJ mol}^{-1}/RT]$	10–80 °C	N117
Costamagna et al., 2008 [113]	$\lambda_m K ka/(1 - ka)/(1 + (K - 1)ka)$ with $\lambda_m = 2.3$ , $K = 70$ , $k = 0.7$	20 °C	N117
Ochi et al., 2009 [114]	$0.8486 + 24.594a - 112.7a^2 + 300a^3 - 358.78a^4 + 162.77a^5$	30–90 °C	N117
Kusoglu et al., 2009 [115]	same as Choi & Datta, but with $P = E(1 - [1 + \kappa(f_m^{-1/3} - 1)][1 - f_c^{1/2}]/[1 - f_{c,dry}^{1/2}])$ , $f_c = f_w + f_m f_{c,dry}$ , $E/\text{MPa} = (1000 - T/0.4 \text{ K})/(1 - f_{c,dry}^{1/2})$ , $f_m = 1 - f_w$ , $f_{c,dry} = V_{SO_3}/V_m$ , $V_{SO_3} = 40.94 \text{ cm}^3 \text{ mol}^{-1}$ , $\kappa = 5.6$	25–85 °C	EW 1100
Mittelsteadt & Liu, 2010 [116]	$(1 + 0.2325a^2(T/^\circ\text{C} - 30)/30)(13.41a - 18.92a^2 + 14.22a^3)$		
Myles et al., 2011 [66]	$16.0674a - 32.3781a^2 + 28.4170a^3$ $15.0395a - 28.3372a^2 + 24.4519a^3$	50 °C 60 °C	N117 N117
Eikerling & Berg, 2011 [117]	$3.0a^{0.2} + 11.0a^4$	30 °C	N117
Li et al., 2013 [118]	$\lambda_v = \lambda_{DM} = \lambda_m ka/(1 - ka) + \lambda_m(K - 1)ka/(1 + (K - 1)ka)$ with $\lambda_m = 3.1$ , $K = 11.4$ , $k = 0.80$ $\lambda_m = 3.3$ , $K = 3.8$ , $k = 0.79$ $\lambda_m = 3.1$ , $K = 9.5$ , $k = 0.75$	25 °C 20 °C 50 °C	N117 N117 N112
Kreuer, 2013 [119, 120]	numerical solution of the implicit equation $\lambda_v = \lambda_f + \sum_{i=1}^5 \prod_{j=1}^i \theta_j$ with $\lambda_f = 1/(PV_w/RT - \ln a)$ , $\theta_j = \exp[\Delta H_j/RT - 1/\lambda_f]/(1 + \exp[\Delta H_j/RT - 1/\lambda_f])$ , $P = E'(\lambda_v V_w/V_m)^{1/3}$		N117
Didierjean et al., 2015 [121]	$0.165 + 13.86a - 24.51a^2 + 23.01a^3$	25 °C	N117
Shi et al., 2016 [122]	$\lambda_v = \lambda_{DM}$ with $\lambda_m = 2.671$ , $K = 7.269$ , $k = 0.7677$	25 °C	NR212
Morin et al., 2017 [123]	$0.053056 + 41.1263a - 180.83a^2 + 406.89a^3 - 381.59a^4 + 69.385a^5 + 62.335a^6$	25 °C	N117

Choi et al. [111] proposed a model in which they combine Flory–Huggins solution theory with elastic polymer swelling, using BET theory for the strongly bound water molecules. However, their description of the Flory–Huggins interaction parameter  $\chi$  isn't explicit enough to allow us to reproduce their general sorption isotherm. Kreuer [119] proposed a thermodynamic sorption model with account for internal elastic pressure, which, after correction of the model equations [120], underestimates the membrane hydration due to the treatment of Nafion as an elastomer. These models are therefore excluded from our following uncertainty analysis.

In Fig. 5b the sorption models are plotted at 30 °C (where possible) to reveal their temperature-independent scatter, which is of an extent comparable to the polynomial correlations. Models that include a temperature dependence are also plotted in Fig. 5c at two different temperatures, highlighting that the effect of temperature is still far from understood. Four out of six models predict an increase in water uptake at high activity with rising temperature, whereas the other two predict the opposite. The experimental data these models have been validated on, as well as the proposed polynomial expressions, likewise disagree on the effect of temperature.

Thampan et al. [24] first recognized that finite-layer BET theory can be used to fit experimental vapor uptake data very well. Although Costamagna et al. [113] and Li et al. [118] later demonstrated that the Guggenheim–Anderson–de Boer (GAB) equation and a dual-mode (DM) model work just as well, the BET isotherm presents a good compromise between physical interpretation, high quality of fit, and suitability for model implementation (being an explicit relationship between  $a$  and  $\lambda_v$ , unlike the more complex implicit sorption models):

$$\lambda_v = \lambda_m \frac{Ka}{1-a} \frac{1 - (n+1)a^n + na^{n+1}}{1 + (K-1)a - Ka^{n+1}} \quad (25)$$

where  $K$  denotes the ratio of the absorption equilibrium constant of the first layer to that of the subsequent layers, determining the shape of  $\lambda_v$  at low relative humidity.  $n$  is the number of adsorbed layers, governing the increase of water uptake at high relative humidity. The water loading at monolayer coverage  $\lambda_m$  can be estimated by [24]

$$\lambda_m = \frac{a_p^{\text{PEM}}}{A_w} \frac{V_m}{N_A} \quad (26)$$

where the area occupied by each adsorbed water molecule on the pore surface is approximately given by [132]

$$A_w = \sqrt{3} \left( \frac{V_w}{2N_A} \right)^{2/3}. \quad (27)$$

$N_A$  is the Avogadro constant and  $a_p^{\text{PEM}} = 210 \text{ m}^2 \text{ cm}^{-3}$  the pore surface area per unit volume of the membrane [133]. Fitting Eq. 25 to published experiments on Nafion 117, NR211 and NR212 at temperatures between 20 °C and 30 °C [20, 21, 59, 127, 112, 78, 128] yields  $K = 92$  and  $n = 12.8$ , which is close to the finding of Thampan et al. [24]. We use this BET isotherm for the baseline model parameterization. It also provides convenient oversaturation behavior, as depicted in Fig. 5d: As  $a \rightarrow \infty$ ,  $\lambda_v \rightarrow \lambda_m n \approx 23$ , which happens to coincide with the reported hydration number for liquid-equilibrated Nafion membranes. We therefore use  $\lambda_l = \lambda_m n$  in Eqs. 17, 18, 21 and 22.

## 2.8. Gas diffusivity

We describe the reduced gas diffusivities through the tortuous pore space by the microstructure factor

$$M_p = \frac{\epsilon_p}{\tau_p^2} (1-s)^\phi \quad (28)$$

with compression-dependent porosity  $\epsilon_p$  and pore tortuosity  $\tau_p$  as detailed in Sec. 2.13. For the saturation exponent  $\phi$  we use  $\phi^{\text{GDL}} = 3$  [134] and  $\phi^{\text{CL}} = 1.5$  [135]. The following dry radii are used for the Knudsen diffusivity in supplementary Eq. S15:  $r_{p,\text{dry}}^{\text{GDL}} = 15 \mu\text{m}$  for a SGL 24 BC [136] and  $r_{p,\text{dry}}^{\text{CL}} = 20 \text{nm}$  for a Nafion/carbon black CL [137].

## 2.9. Liquid water transport

The effective liquid water transport coefficient  $D_s$  (see supplementary Eq. S20) depends on several material properties. In particular the functional relationship between saturation  $s$  and capillary pressure  $p_c$  is a topic of extensive research with a large variety of fitting functions that have been proposed [138]. It strongly depends not only on the wettability and compression of the porous medium, but also on the exact kind of water transport process (primary injection, withdrawal, further injections). We found the overall impact of  $D_s$  on fuel cell performance to be mostly small, though, as will be shown in Part II of this series. Therefore, we focus on a single parameter set here rather than screening the literature for different parameterizations. A study on the effects of artificially altered capillary pressure–saturation relationships can be found in [139]. The van Genuchten law, which Gostick et al. [140] have found to apply to many common GDLs, is used here:

$$\frac{\partial p_c}{\partial s} = \frac{p_b}{lm} \left( s_w^{-1/m} - 1 \right)^{1/l-1} s_w^{-1/m-1}, \quad s_w = \frac{1-s}{1-s_{\text{im}}} \quad (29)$$

with the following parameters for the secondary water injection curve of compressed SGL carbon paper:  $m = 0.6$ ,  $l = 100$ , a breakthrough pressure of  $p_b = 1.07 \text{bar}$ , and an immobile saturation of  $s_{\text{im}} = 0.08$  [140]. Mualem’s model is used for the relative hydraulic permeability, reading [141]

$$K_{\text{rel}} = (1-s_w)^2 \left( 1 - s_w^{1/m} \right)^{2m} + 10^{-6} \quad (30)$$

where the small offset serves to bypass numerical difficulties under dry conditions, i.e., to avoid that  $K_{\text{rel}} \rightarrow 0$  as  $s \rightarrow s_{\text{im}}$ . We note, however, that the relative permeability in partially saturated GDLs is an active area of research, and alternate expressions such as S-shaped functions were recently proposed [142]. For the absolute permeability of the GDLs, the semi-heuristic Carman–Kozeny equation for fibrous porous media is employed, as it proved to work well for carbon paper [143]:

$$K_{\text{abs}}^{\text{GDL}} = \frac{\epsilon_p^3 d_f^2}{16k_K (1-\epsilon_p)^2} \quad (31)$$

where  $d_f = 8.0 \mu\text{m}$  and  $k_K = 4.54$  are the fiber diameter and the Kozeny constant for a SGL 24 BA [143]. A constant value of  $K_{\text{abs}}^{\text{CL}} = 0.1 \mu\text{m}^2$  [144] is assumed for the CLs. Finally, to complete the definition of the effective liquid water transport coefficient  $D_s$  (supplementary Eq. S20), the internationally recommended correlation of the dynamic viscosity of liquid water at 1 bar up to 110 °C is used [145]:

$$\mu = \left( 280.68 \overline{T}^{-1.9} + 511.45 \overline{T}^{-7.7} + 61.131 \overline{T}^{-19.6} + 0.45903 \overline{T}^{-40} \right) \mu\text{Pa s} \quad (32)$$

Table 6: Review of mass transfer coefficients for water vapor absorption/desorption in Nafion membranes.

Publication	Mass transfer coefficients [ $10^{-3} \text{ cm s}^{-1}$ ]	Temperature	Activation energy	Membrane
Rivin et al., 2001 [59]	1031	20 or 32 °C		N117
Ye & LeVan, 2003 [68]	$\sim p_{\text{H}_2\text{O}}^{1.5}$	23.5–25 °C		
Berg et al., 2004 [146]	0.57	70–80 °C		N112
Ge et al., 2005 [65]	$k_a = 1.14f_w$ $k_d = 4.59f_w$	30 °C	20 kJ mol <sup>-1</sup>	N112,115,117
Satterfield et al., 2008 [147]	$k_d = 0.14\text{--}0.29$	70 °C	25–31 kJ mol <sup>-1</sup>	N112,115,1110,1123
Monroe et al., 2008 [148]	630	50 °C		N112,115,117
Hallinan & Elabd, 2009 [149]	20	30 °C		N117
Adachi et al., 2010 [150, 151]	0.45	70 °C		N112,115,117, DE2021CS
He et al., 2011* [152]	$1.85 \max\{0, \lambda - \lambda_0\}^{1.25}$ , $\lambda_0 = 3.17$	25 °C		NR212
Tabuchi et al., 2011 [61]	$k_d = 0.2\text{--}1$	30 °C		N1110
Kongkanand, 2011 [151]	$k_a = 0.0184a^2 + 0.0586a + 0.129$ $k_d = 0.256 a^2 + 0.148 a + 0.191$	80 °C	28.1 kJ mol <sup>-1</sup> 29.7 kJ mol <sup>-1</sup>	DE2020
Kusoglu & Weber, 2012* [2, 153]	$0.68 \max\{0, \lambda - \lambda_0\}^{1.6}$	25–80 °C		NR211,212
Didierjean et al., 2015* [121]	$400V_m P_{\text{sat}}/RT$	25 °C		N117

\*To be divided by the sorption slope  $\partial\lambda_v/\partial a$  when used in supplementary Eq. S22.

[FIGURE 6]

Figure 6: Experimentally determined interfacial mass transfer coefficients for Nafion membranes as a function of water content. Individual symbols represent measurement data and lines show fitted functions. Data from [151, 2] were converted from relative humidity to water content using Eq. 25.

with  $\bar{T} = T/300 \text{ K}$ . In the molar volume of liquid water  $V_w = m_w/\rho_w$ ,  $m_w = 18.015 \text{ g mol}^{-1}$  is the molar mass of water, and  $\rho_w$  its mass density given at standard atmospheric pressure by [124]

$$\rho_w = \rho_c \left( 1 + 1.9927\hat{T}^{1/3} + 1.0997\hat{T}^{2/3} - 0.51084\hat{T}^{5/3} - 1.7549\hat{T}^{16/3} - 45.517\hat{T}^{43/3} - 674694\hat{T}^{110/3} \right) \quad (33)$$

with  $\hat{T}$  and  $\rho_c$  like in Eq. 24.

### 2.10. Vapor sorption

Numerous experiments have been carried out to measure the interfacial mass transfer coefficients  $k_a$  and  $k_d$  (supplementary Eq. S22), as listed in Tab. 6. They differ in the experimental procedure (absorption, desorption, permeation, liquid vs. vapor basins) and the driving force considered (concentration or hydration number gradient vs. activity gradient). Measurements that are based upon a water activity difference between the ionomer ( $a_i$ ) and vapor phases ( $a$ ) are marked with an asterisk in Tab. 6. These coefficients can be approximately converted to  $k_a$  and  $k_d$  for use in supplementary Eq. S22 by dividing them by the local slope of the sorption isotherm,  $\partial\lambda_v/\partial a$  [121], because for small deviations from equilibrium,

$$(\lambda - \lambda_v) \approx \frac{\partial\lambda_v}{\partial a}(a_i - a). \quad (34)$$

Some of the more systematic studies that include the moisture dependence of the water transport resistance at the ionomer–gas interface are compared in Fig. 6. Ye & LeVan [68] reported an approximate power-law correlation with the partial pressure of water vapor. Since it is unclear how their expression translates to supplementary Eq. S22

Table 7: Review of evaporation/condensation rates.

Publication	Rate expressions	Coefficients	Rate values*
Nguyen & White, 1993 <sup>†</sup> [156]	$\gamma_c = k_c$ $\gamma_e = k_c$	$k_c = 1 \text{ s}^{-1}$	$\gamma_c = 1 \text{ s}^{-1}$ $\gamma_e = 1 \text{ s}^{-1}$
Nguyen, 1999 [157]	$\gamma_c = k_c \epsilon_p (1 - s)$ $\gamma_e = k_e \epsilon_p s RT / V_w$	$k_c = 100 \text{ s}^{-1}$ [157] $k_e = 100 \text{ atm}^{-1} \text{ s}^{-1}$ [157] $k_c = 100 \text{ s}^{-1}$ [158] $k_e = 1 \text{ atm}^{-1} \text{ s}^{-1}$ [158]	$\gamma_c = 56 \text{ s}^{-1}$ $\gamma_e = 22\,000 \text{ s}^{-1}$ $\gamma_c = 56 \text{ s}^{-1}$ $\gamma_e = 220 \text{ s}^{-1}$
He et al., 2000 [159]	$\gamma_c = k_c \epsilon_p (1 - s) y_{\text{H}_2\text{O}}$ $\gamma_e = k_e \epsilon_p s RT / V_w$	$k_c = 100 \text{ s}^{-1}$ [159] $k_e = 100 \text{ atm}^{-1} \text{ s}^{-1}$ [159] $k_c = 5000 \text{ s}^{-1}$ [160] $k_e = 10^{-4} \text{ Pa}^{-1} \text{ s}^{-1}$ [160] $k_c = 100 \text{ s}^{-1}$ [161] $k_e = 5 \text{ atm}^{-1} \text{ s}^{-1}$ [161]	$\gamma_c = 14 \text{ s}^{-1}$ $\gamma_e = 22\,000 \text{ s}^{-1}$ $\gamma_c = 700 \text{ s}^{-1}$ $\gamma_e = 1200 \text{ s}^{-1}$ $\gamma_c = 14 \text{ s}^{-1}$ $\gamma_e = 1100 \text{ s}^{-1}$
Natarajan & Nguyen, 2001 [162]	$\gamma_c = k_c \epsilon_p (1 - s) RT y_{\text{H}_2\text{O}}$ $\gamma_e = k_e \epsilon_p s RT / V_w$	$k_c = \text{N/A}$ $k_e = \text{N/A}$	$\gamma_c = \text{N/A}$ $\gamma_e = \text{N/A}$
Nam & Kaviani, 2003 [163]	$\gamma_c = \Gamma_m a_{\text{lg}} \sqrt{RT/2\pi m_w}$ $\gamma_e = \Gamma_m a_{\text{lg}} \sqrt{RT/2\pi m_w}$	$\Gamma_m = 0.006$ $a_{\text{lg}} = 1000 \text{ m}^{-1}$	$\gamma_c = 970 \text{ s}^{-1}$ $\gamma_e = 970 \text{ s}^{-1}$
Weber et al., 2004 [164]	$\gamma_c = k_m a_{\text{lg}} RT$ $\gamma_e = k_m a_{\text{lg}} RT$	$k_m a_{\text{lg}} = 100 \text{ mol bar}^{-1} \text{ cm}^{-3} \text{ s}^{-1}$	$\gamma_c = 2\,900\,000 \text{ s}^{-1}$ $\gamma_e = 2\,900\,000 \text{ s}^{-1}$
Birgersson et al., 2005 [165]	$\gamma_c = k_c \epsilon_p$ $\gamma_e = k_e s RT / m_w$	$k_c = 100 \text{ s}^{-1}$ $k_e = 100 \text{ s m}^{-2}$	$\gamma_c = 70 \text{ s}^{-1}$ $\gamma_e = 3\,300\,000 \text{ s}^{-1}$
Eikerling et al., 2006 <sup>‡</sup> [166]	$\gamma_e = k_e RT \xi_{\text{lg}} / N_A L^{\text{CL}}$	$k_e = 1.4 \times 10^{18} \text{ atm}^{-1} \text{ cm}^{-2} \text{ s}^{-1}$	$\gamma_e = 13\,000 \text{ s}^{-1}$
Wu et al., 2009 [167]	$\gamma_c = \Gamma_m \Gamma_s a_p (1 - s) \sqrt{RT/2\pi m_w}$ $\gamma_e = \Gamma_m \Gamma_s a_p s \sqrt{RT/2\pi m_w}$	$\Gamma_m = 0.006$ $\Gamma_s \lesssim 0.2$	$\gamma_c = 1\,500\,000 \text{ s}^{-1}$ $\gamma_e = 390\,000 \text{ s}^{-1}$

\*Evaluated at  $T = 80 \text{ }^\circ\text{C}$ ,  $\epsilon_p = 0.7$ ,  $s = 0.2$ ,  $y_{\text{H}_2\text{O}} = 0.25$ ,  $\Gamma_s = 0.1$ ,  $a_p = 20 \text{ m}^2 \text{ cm}^{-3}$ ,  $L^{\text{CL}} = 10 \text{ }\mu\text{m}$ ,  $\xi_{\text{lg}} = 200$ .

<sup>†</sup>Evaporation/condensation in gas channel.

<sup>‡</sup>Evaporation model for cathode catalyst layer only.

though, we exclude this early result from the following analysis. Ge et al. [65] proposed to take  $k_a$  and  $k_d$  proportional to the ionomer's water volume fraction  $f_w$  analogous to the intra-diffusion coefficient of water in bulk Nafion as proposed by Weber & Newman [28] (cf. Tab. 3). They also adopted the activation energy  $E_{\text{ad}} = 20 \text{ kJ mol}^{-1}$  from a diffusion measurement by Yeo & Eisenberg [76]. Later measurements of the activation energy of sorption have yielded larger values,  $E_{\text{ad}} \approx 29 \text{ kJ mol}^{-1}$  [154, 147, 151]. He et al. [152] and Kusoglu et al. [2] proposed power laws in  $\lambda$  to express the mass transfer coefficients, based on data that possibly stem from the same measurement carried out at the Lawrence Berkeley National Laboratory, but assuming different residual hydration numbers  $\lambda_0$  (3.17 vs. 0 in Fig. 6). Converting their measured data points from activity-driven to hydration number-driven using Eq. 34 yields an approximately constant mass transfer coefficient  $k_{a,d} \approx 10^{-3} \text{ cm s}^{-1}$  independent of  $\lambda$  (data not shown). Kongkanand [151] used a polynomial in water activity to fit the coefficients.

Although there is disagreement on whether the interfacial mass transfer grows with increasing  $\lambda$  [68, 65, 151] or not [152, 2], the studies that differentiate between absorption and desorption agree that the former is substantially slower than the latter [65, 151, 121]. To date, the true dependence on moisture is essentially an open problem. Most researchers report the mass transfer coefficients in the range  $10^{-4}$ – $10^{-2} \text{ cm s}^{-1}$  in the relevant temperature range of 50–90 °C. This is also the range numerically examined in early models by Okada [95, 155]. With these considerations in mind, we adopt Ge's correlation for the baseline simulation.

[FIGURE 7]

Figure 7: Review of evaporation rates (open symbols) and condensation rates (closed symbols). Symbols are half-filled if both are equal. Colored symbols denote modified parameters used in conjunction with earlier established rate expressions (black symbols) as listed in Tab. 7. For comparison, a recent series of *ex-situ* and *in-situ* evaporation measurements in SGL 24 BA [168, 169] and Toray TGP-H-060 [170] were recast into the form of supplementary Eq. S23 and shown as heart symbols.

2.11. Evaporation and condensation

Supplementary Eq. S23 is the commonly employed out-of-equilibrium way to account for liquid–vapor phase change in macro-homogeneous PEMFC modeling with explicit representation of both phases. A comprehensive comparison of different expressions for the rates  $\gamma_{e,c}$  developed in modeling is given in Tab. 7. In the final column, they are evaluated for a typical state of fuel cell operation, demonstrating that the effective evaporation and condensation rates used in published two-phase models vary over as much as five orders of magnitude. Fig. 7 shows the timeline of when these rates came about in the literature, revealing no sign of convergence nor even a trend.

A feature shared by many models is that the condensation (evaporation) rate is assumed to be proportional to  $1 - s$  ( $s$ ) to account for the change of available phase boundary, as first proposed by Nguyen [157]. He et al. [159] proposed that condensation be proportional to the molar water vapor fraction  $y_{\text{H}_2\text{O}}$ , but more recently developed models have abandoned this assumption. Nam & Kaviani [163] suggested to use the Hertz–Knudsen equation from kinetic gas theory to model both directions of the phase change, assuming continuity of the temperature at the phase boundary and equal rates. Weber et al. [164] used a simpler expression, also based on equal rates for evaporation and condensation. These formulas explicitly include the liquid–gas interfacial area density  $a_{\text{lg}}$ , which is unknown a priori, making even an estimation of the rates difficult. Wu et al. [167] estimate it to be  $a_{\text{lg}} = \Gamma_s a_p (1 - s)$  for condensation and  $a_{\text{lg}} = \Gamma_s a_p s$  for evaporation with an interfacial area accommodation coefficient  $\Gamma_s \lesssim 0.2$ .  $a_p$  denotes the average pore surface area density of the porous domain. A recent experimental study [168] has shown that  $a_{\text{lg}} \sim s$  and that there is no clear correlation between the evaporation rate per interfacial area and  $s$ , suggesting that the evaporation rate should indeed increase linearly with  $s$ .

It is known from numerous experimental measurements that the condensation coefficient of water is roughly a decade larger than the evaporation coefficient, and that both decrease with increasing pressure and temperature [6]. The proposed coefficient  $\Gamma_m = 0.006$  in Nam’s and Wu’s models is consistent with this experimental data at atmospheric pressure for condensation, but about a decade too large for evaporation. For this reason we implement the Hertz–Knudsen law in the form proposed by Wu et al. (final row in Tab. 7) as the baseline parameterization of phase change with  $\Gamma_s = 0.1$ ,  $\Gamma_m = 0.006$  for condensation, but  $\Gamma_m = 0.0005$  for evaporation, analogous to [10]. We also use  $s_{\text{nw}} = 1 - s_w$  in place of  $s$  to prevent  $s_{\text{nw}}$  from becoming negative during evaporation. The specific pore surface area  $a_p$  is modeled as a function of compression as discussed in Sec. 2.13.

It should be noted, however, that the general validity of the Hertz–Knudsen equation was recently questioned and that it was hypothesized that evaporation/condensation be governed by a balance of tiny pressure differences and momentum flux instead [171]. The Hertz–Knudsen rates used here are among the higher ones used in modeling works (cf. Fig. 7), quickly yielding saturated vapor where liquid water is present. In this quasi-equilibrium regime, the exact values of the rates become irrelevant [172, 173], and any functional dependencies of the rate expressions on  $s$ ,  $T$  etc. can essentially be dropped. This is in agreement with the reported observation that evaporation in



partially saturated GDLs is fast enough to be transport-limited (diffusion-limited) rather than area- or kinetics-limited [168, 169, 170].

### 2.12. Latent heat

The molar latent heat of water condensation  $H_{ec}$  used in supplementary Eq. S6 can be parameterized as

$$H_{ec} = 52.51 \exp \left[ 0.261\tilde{T} - 0.044\tilde{T}^2 - 0.0044\tilde{T}^3 \right] \text{ kJ mol}^{-1} \quad (35)$$

with  $\tilde{T} = \ln [1 - T/T_c]$ . This is a global least-squares fit to the tabulated data from [174]. The latent heat of water sorption can be written as

$$H_{ad} = H_{ec} + H_{mix} \quad (36)$$

where the mixing enthalpy  $H_{mix}$ , which is often assumed constant or even neglected in published models, is a function of membrane hydration [175, 176] and temperature [177] for Nafion. Since a reliable parameterization for  $H_{mix}$  is absent from the literature, a new one is proposed here, based on the measurements on Nafion 115 by Wadsö & Jannasch [177]. Their data indicate that  $H_{mix}$  essentially decays exponentially in  $\lambda$ , with a temperature-dependent deviation in the very dry regime ( $\lambda < 3$ ). We fitted the phenomenological approximation

$$H_{mix} = (a_1 \exp [-b_1\lambda] + a_2\lambda \exp [-b_2\lambda^2]) \text{ kJ mol}^{-1} \quad (37)$$

with temperature-dependent coefficients to their data in the range  $40^\circ\text{C} \leq T \leq 100^\circ\text{C}$  and  $0.05 \leq \lambda \leq 5$  and obtained least squared residuals for

$$\begin{aligned} a_1 &= -107.5\bar{T}^2 + 253.9\bar{T} - 138.7 \\ a_2 &= 106.8\bar{T} - 102.4 \\ b_1 &= 2.006\bar{T}^2 - 4.365\bar{T} + 2.931 \\ b_2 &= 108.7\bar{T}^2 - 262.8\bar{T} + 159.5 \end{aligned} \quad (38)$$

where  $\bar{T} = T/300 \text{ K}$ .

### 2.13. Compression and contact resistivities

To account for the effects of clamping pressure on the computational domain, the layer thicknesses are modeled as a function of pressure. Given the compressive strain  $\varepsilon_c$ , one can write

$$L = L_0(1 - \varepsilon_c) \quad (39)$$

where  $L_0$  denotes the thickness of the uncompressed layer. We model SGL 24 GDLs on both sides of the MEA, for which  $L_0^{\text{GDL}} = 190 \mu\text{m}$  [178]. The relationship between strain and applied clamping pressure  $P_{cl}$  for these GDLs is [179]

$$\varepsilon_c^{\text{GDL}} = -0.0083 \left( \frac{P_{cl}}{1 \text{ MPa}} \right)^2 + 0.0911 \left( \frac{P_{cl}}{1 \text{ MPa}} \right). \quad (40)$$

Table 8: Contact resistance parameters.

Interface	ECR			TCR		
	$R_0$ [m $\Omega$ cm $^2$ ]	$\zeta$ [-]	Ref.	$R_0$ [K cm $^2$ W $^{-1}$ ]	$\zeta$ [-]	Ref.
CL/GDL	29.4	0.89	[184]	1.56	0.71	[38]
GDL/BP	3.34	0.53	[185]	2.89	0.64	[186]

The two catalyst layers are also compressed under applied pressure, but reliable measurement data is rare. We fitted the following two-parameter function to the compressive strain of the CL reported by Burheim et al. [180]:

$$\varepsilon_c^{\text{CL}} = 0.422 \left( 1 - \exp \left[ -\frac{P_{\text{cl}}}{0.970 \text{ MPa}} \right] \right) \quad (41)$$

Together with Eq. 39, this equation is used to express  $L^{\text{CL}}$  as a function of applied pressure. For the baseline simulation,  $L_0^{\text{CL}} = 10 \mu\text{m}$  and  $P_{\text{cl}} = 1 \text{ MPa}$  are used. Nafion NR211 with constant thickness  $L_0^{\text{PEM}} = 25.4 \mu\text{m}$  [77] is chosen as the membrane, assuming for simplicity that swelling and compression cancel one another ( $\varepsilon_c^{\text{PEM}} = 0$ ).

Kumbur et al. [179] measured the average specific pore surface  $a_p$  at three different compaction pressures from 0 to 1.4 MPa for a few GDLs from SGL Group with MPLs, finding a moderate increase with increasing pressure. Asymptotically, though, it is clear that at very large applied pressures, the pore surface area density must come down again. We therefore fitted a quadratic polynomial to their data for SGL 24 BC:

$$\frac{a_p^{\text{GDL}}}{1 \text{ m}^2 \text{ cm}^{-3}} = -1.96 \left( \frac{P_{\text{cl}}}{1 \text{ MPa}} \right)^2 + 8.18 \left( \frac{P_{\text{cl}}}{1 \text{ MPa}} \right) + 23.4. \quad (42)$$

As no data seems to be available in the literature for CLs, we set  $a_p^{\text{CL}} = a_p^{\text{GDL}}$ .

Kumbur et al. also determined the compression dependence of porosity for SGL carbon paper (with MPL) in the pressure range up to 4 MPa and proposed the relationship

$$\epsilon_p^{\text{GDL}} = \epsilon_{p,0}^{\text{GDL}} \left( \frac{0.9}{1 + \varepsilon_c^{\text{GDL}}} + 0.1 \right) \quad (43)$$

where  $\epsilon_{p,0}^{\text{GDL}} = 0.75$  [179] is the porosity of the uncompressed GDL. While Eq. 41 determines how much CLs are compressed, it is unclear how this changes their porosity. We assume that only the pore space of the CLs is compressed, which yields

$$\epsilon_p^{\text{CL}} = \frac{\epsilon_{p,0}^{\text{CL}} - \varepsilon_c^{\text{CL}}}{1 - \varepsilon_c^{\text{CL}}} \quad (44)$$

where  $\epsilon_{p,0}^{\text{CL}}$  is the porosity of the uncompressed CL, taken as 0.4 [181]. The relationship between through-plane pore tortuosity and applied clamping pressure is commonly expressed indirectly, with tortuosity as a function of porosity. For the present model, a parabolic least-squares fit to experimental data for SGL 24 DA by Flückiger et al. [182] is used,

$$\tau_p^{\text{GDL}} = -17.3 (\epsilon_p^{\text{GDL}})^2 + 18.8 \epsilon_p^{\text{GDL}} - 1.72, \quad (45)$$

whereas the tortuosity of the CLs is assumed constant ( $\tau_p^{\text{CL}} = 1.5$  [183, 37]) due to the apparent absence of published measurement data on its compression dependence.

To complete the constitutive parameterization of our model, the electrical and thermal contact resistivities  $R_e$

Table 9: Reference operating conditions.

Symbol	Explanation	Value
$P_A$	Gas pressure in anode gas channel	1.5 bar
$P_C$	Gas pressure in cathode gas channel	1.5 bar
$RH_A$	Relative humidity in anode gas channel	100%
$RH_C$	Relative humidity in cathode gas channel	100%
$T_A$	Temperature of anode bipolar plate	80 °C
$T_C$	Temperature of cathode bipolar plate	80 °C
$\alpha_{O_2}$	Oxygen mole fraction in dry oxidant gas	21%

and  $R_T$  in supplementary Eq. S24 remain to be specified for the different MEA interfaces. As recently highlighted [187], they usually follow power laws of the form

$$R = R_0 \left( \frac{P_{cl}}{1 \text{ MPa}} \right)^{-\zeta} \quad (46)$$

where the coefficients  $R_0$  and  $\zeta$  vary with the materials considered – a result that was also found theoretically for contacting fractal surfaces, for which the exponent is  $\zeta \in [0.5, 1]$  depending on the degree of material plasticity and surface roughness [188]. Numerical simulations based on contact mechanics have confirmed this relationship [189, 190] for electrical contact resistance (ECR) at the GDL/BP interface. Thermal and electrical contact resistances are implemented in the present MEA model using the values listed in Tab. 8.

Finally, we note that through-plane thermal contact resistance (TCR) is significantly reduced by the presence of liquid water (e.g., [46]). Since the data in the literature is too scattered though to allow for a reliable parameterization with respect to  $s$ , the dependency of the contact resistivities on moisture is neglected here. This aspect requires additional experimental clarification.

### 3. Results

In the model parameterization in Sec. 2, special attention was paid to six of the most critical and controversial transport parameters for Nafion-based MEAs, and their openly available constitutive relationships were reviewed in Tabs. 2 to 7. We now employ our macro-homogeneous steady-state MEA model to determine the degree of uncertainty in the fuel cell performance prediction associated with these six material parameterizations. Implemented in COMSOL Multiphysics, the model is numerically solved with the finite element method. 50 elements with quadratic Lagrangian shape functions are used per MEA layer (totaling in 250 finite elements for the entire MEA), and the damped Newton method is used with a relative error tolerance of  $10^{-5}$  to solve the coupled nonlinear differential equations simultaneously. Tab. 9 lists the operating conditions at which all simulations shown in this first part are carried out. They were chosen to approximately represent a differential section of a PEMFC operated in an automotive scenario.

To quantify the discrepancy between the different model outputs, we selected the maximum reachable current density  $I_{\max}$  and the peak power density  $P_{\max}$  as key figures. Note that these are only two out of an entire range of possible performance indicators.

[FIGURE 8]

Figure 8: Impact of selected parameterizations on predicted fuel cell performance at reference operating conditions. Solid lines represent voltage (left axes), dashed lines represent power density (right axes). The baseline parameterization is shown with slightly thicker black lines.

### 3.1. Protonic conductivity

Fig. 8 shows the resulting polarization and power density curves predicted by the model when each of the six major parameters are substituted while all others are retained at the baseline as described in Sec. 2. As can be recognized from Fig. 8a, employing different expressions for the ionic conductivity  $\sigma_p$  of the Nafion membrane leads to enormous scatter. From the experimental scatter of  $\sigma_p$  alone, the total bandwidth of obtained values is  $1.10 \text{ A cm}^{-2}$  for  $I_{\max}$  and  $0.52 \text{ W cm}^{-2}$  for  $P_{\max}$ , making the protonic conductivity of the ionomer a material property of very large uncertainty.

The parameterizations by Hsu et al. [19], Morris & Sun [21] and Zhao et al. [35], which were shown to predict generally low conductivities in Fig. 1, yield current densities that are roughly a factor of two lower than the others, over the entire range of cell voltages. Yet, if these three are dismissed as possible measurement outliers, the remaining performance variation is still appreciable (up to about  $0.4 \text{ A cm}^{-2}$ ). To reach large current densities, the ionomer needs to offer high conductivity at low water content, because electro-osmotic drag dries out the anode side of the membrane (cf. [10]). It is for this reason that the conductivity data by Maldonado et al. [34] yields high performance at intermediate voltages, followed by a sudden drop. The parameterizations which suggest higher  $\sigma_p$  at low  $\lambda$  values (in particular, those by Springer et al. [20], Costamagna [25] and Meier & Eigenberger [29]) outperform others at large current densities. It must be stressed that even small differences in the conductivity expression that may appear marginal at first can have a great impact on fuel cell models and that detailed knowledge of  $\sigma_p(\lambda)$  at low  $\lambda$  is critical.

### 3.2. Water diffusivity in the ionomer

Fig. 8b shows the influence of the water diffusion coefficient  $D_\lambda$ . Generally, higher diffusivities result in stronger back diffusion, which helps keeping the anode humidified at large current densities, but the effect of nonlinear features in  $D_\lambda$  can be subtle. Ye & LeVan's data [68] yields the poorest cell performance due to the very low diffusivity in the dry and wet regimes, despite being the largest in between (Fig. 2d). The correlation proposed by Lokkiluoto & Gasik [67], which globally predicts low diffusivity, also yields a low limiting current density. To maintain high performance at large currents, good water diffusivity under very dry conditions is crucial, similar to our conclusion for the ionic conductivity above. Fuller's [63], Kulikovsky's [27], Zhao's [58] and Caulk's [69] expressions yield the lowest membrane resistivities because their back diffusion coefficients remain the largest toward low  $\lambda$  values (cf. Fig. 2). Our own fit [10] to measurement data by Mittestadt & Staser [62] (with Arrhenius correction using Eq. 16) lies somewhat in the middle of the overall scatter of polarization curves. With  $1.34 \text{ A cm}^{-2}$  for  $I_{\max}$  and  $0.52 \text{ W cm}^{-2}$  for  $P_{\max}$ , the uncertainty associated with  $D_\lambda$  is even bigger than with  $\sigma_p$ .

### 3.3. Electro-osmosis

Next, we turn our attention to the electro-osmotic drag coefficient  $\xi$ . Generally, the fewer water molecules are dragged along with each traversing proton, the more evenly the membrane remains humidified across its thickness,

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2  
3 which in turn reduces the ohmic resistivity of the membrane due to the monotonicity of  $\sigma_p$  in  $\lambda$ . It thus comes with  
4 no surprise that the large variation in measurement data on  $\xi$  translates to wide scatter in the resulting polarization  
5 curves. As Fig. 8c shows, the empirical constitutive expressions by Springer et al. [20], Dutta et al. [80] and Ge et  
6 al. [82] yield the best performance, since they predict the lowest  $\xi$  at medium to dry conditions (cf. Fig. 4b). It so  
7 happens, though, that this is also the regime where the reported experiments diverge the most (cf. Fig. 4a). When  
8 the parameterizations of Fuller [79, 63], Eikerling [23], Meier & Eigenberger [29] or Lokkiluoto & Gasik [67] are  
9 employed, the fuel cell model **cannot reach large** current densities, because the anode dries out more. In summary,  
10 the electro-osmotic drag coefficient in Nafion is a model parameter with large uncertainty.  
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### 16 3.4. Membrane hydration

17  
18 A natural implication of the monotonic increase of membrane conductivity with increasing hydration is that  
19 higher sorption isotherms yield better performance prediction. This effect is apparent in Fig. 8d. Takata’s [112] and  
20 Costamagna’s [113] vapor sorption models are among those which yield the smallest  $\lambda_v$  values in the low and high  
21 activity regimes, respectively (cf. Fig. 5). They therefore result in the steepest decline of the polarization curve in the  
22 ohmic region. Morin’s data [123] and Kusoglu’s model, on the other hand, both suggest good membrane hydration  
23 over the entire activity range, resulting in a polarization curve that extends toward higher current densities. With  
24 a total spread of  $0.60 \text{ A cm}^{-2}$  for the tested parameterizations, the water vapor uptake of the membrane is a  
25 significant source of modeling uncertainty. Given that it depends also on the membrane’s hygro-thermal history  
26 [191, 34] (which has been ignored in our present analysis), this highlights that detailed experimental characterization  
27 of the membrane is required to make PEMFC models predictive.  
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### 35 3.5. Vapor sorption kinetics

36  
37 The next effect in the line is the sorption of water vapor at the ionomer–gas interface. With the four hydration-  
38 dependent expressions for the mass transfer coefficients plotted in Fig. 6, we estimate the modeling uncertainty  
39 originating from measurement data on  $k_{a,d}$ , setting  $k_a = k_d$  for the two parameterizations that do not distinguish  
40 between absorption and desorption [152, 2]. Fig. 8e shows the resulting polarization curves, which start to separate  
41 no earlier than at intermediate cell voltages. He’s expression [152] yields the most limited cell performance due to  
42 the fast-dropping mass transfer coefficients toward low hydration numbers. The slight kink near 0.6 V is due to the  
43 kink in  $k_{a,d}$  at  $\lambda_0 = 3.17$ , which is used here with the intention of adopting He’s expression without modification, in  
44 spite of the absence of residual hydration in the employed sorption isotherm (Eq. 25). Konkanand’s data [151] yields  
45 the largest current densities, for their mass transfer coefficients remain positive (still allowing for moderate vapor  
46 absorption) when the electro-osmotic drag dries out the anode side of the membrane. Once again, we find that the  
47 constitutive material behavior at very low water content has a major impact in the predicted cell performance and  
48 therefore needs to be known with high accuracy – higher than currently available in the literature.  
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### 57 3.6. Evaporation and condensation

58  
59 Finally, the effect of adopting different expressions for the evaporation and condensation rates is shown in Fig. 8f.  
60 Wu’s Hertz–Knudsen expression [167] is used as the baseline parameterization with lowered evaporation coefficient  
61 (see Sec. 2.11), which results in higher liquid water saturation, and consequently, lower limiting current density.  
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63  
64  
65

Table 10: Variation in current-voltage characteristics due to scatter in selected parameterizations at reference operating conditions. Parameters are sorted by uncertainty in decreasing order.

Parameters	$I_{\max}$ [A cm <sup>-2</sup> ]		$P_{\max}$ [W cm <sup>-2</sup> ]	
	std. dev.	total spread	std. dev.	total spread
$D_{\lambda}$	0.43	1.34	0.16	0.52
$\sigma_p$	0.34	1.10	0.15	0.52
$\xi$	0.32	0.91	0.07	0.21
$\lambda_v$	0.19	0.60	0.06	0.22
$k_a, k_d$	0.15	0.36	0.05	0.11
$\gamma_c, \gamma_e$	0.04	0.10	0.03	0.06

With a difference of just 31 mA cm<sup>-2</sup>, the effect is relatively small, though. Expressions which give fast evaporation and slow condensation, such as [157, 159, 165] (cf. Fig. 7), reduce the liquid water saturation further, in particular in the CCL, allowing for better oxygen access to the catalyst sites. In total, though, the scatter induced by the different phase change rates is much smaller than for the five ionomer properties discussed above, because in the examined parameter range, evaporation is fast enough to yield a relative gas humidity of  $\approx 100\%$  across almost the entire MEA where liquid water is present.

### 3.7. Summary

The results of our uncertainty analysis are summarized in Tab. 10, where in addition to the total spread, also the standard deviations for the limiting current density and peak power density are given. **The standard deviation measures the mean expected uncertainty for picking material parameterizations from the literature at random, whereas the total spread demonstrates how dramatic the uncertainty can be in a worst-case scenario.** This corroborates that the same ranking list is obtained when individual parameterizations, which may be deemed measurements outliers, are given less weight or omitted from the analysis. Evidently, getting the water transport inside the membrane and across its interfaces correct is the key to predictive macro-homogeneous MEA modeling for PEMFCs. Most uncertainty stems from the ionic conductivity as well as from the two major water transport effects in the ionomer: back diffusion and electro-osmosis. **These must become known more precisely than is presently the case. Our results show how unreliable PEMFC models can be in the large current density regime and how easy it is to alter the polarization curve by favoring certain material parameterizations in the open literature over others.**

## 4. Conclusion

With the present work, we have found a quantitative answer to the questions raised in the introductory section: When characterizing a MEA with the intention of extracting material properties for PEMFC modeling, special attention should be paid to the membrane properties, in particular the Fickian diffusivity of dissolved water, the protonic conductivity and the electro-osmotic drag coefficient. At least under the present modeling assumptions and operating conditions, these are the top three traits of Nafion which cause the most uncertainty in the predicted fuel cell performance, based on the available data in the open literature. For high accuracy at large current densities, their functional dependency on the state of membrane hydration is critical, in particular in the dry regime (low  $\lambda$ ). This finding is relevant not only for fuel cell experimentalists, but also for modelers who are faced with the task of selecting appropriate MEA material parameterizations for their calculations.

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3 There are certainly important performance-limiting in-plane transport effects (e.g., liquid water accumulation in  
4 the flow channels) that a 1D through-plane model cannot account for. The reported effect of individual transport  
5 parameters is therefore expected to change in higher-dimensional models. However, through-plane transport pro-  
6 cesses need to be understood, and uncertainties in this direction need to be reduced, since higher-dimensional models  
7 also contain the uncertainties from through-plane transport, in addition to those of the perpendicular direction(s).  
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11 With the present scatter in the available data records on Nafion properties, predictive performance prediction  
12 is very difficult. Measurement errors should be quantified more routinely in experimental studies than they have  
13 been in the past. Further research is required to improve the understanding of the transport processes of hydrogen  
14 ions and water molecules across the electrolyte membrane, not only under varying water content, but also to clarify  
15 the temperature dependence of water uptake and electro-osmotic drag.  
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19 In Part I, we have studied only *local* parameter sensitivity, i.e., at fixed operating conditions and with all  
20 material parameterizations but one fixed. Given the highly nonlinear nature of PEMFCs, the picture might change  
21 in different scenarios. So far, we have focused on uncertainty in the extrema of polarization characteristics, which  
22 is induced by scatter in the proposed parameterizations for six of the most critical MEA properties. In Part II, we  
23 perform a *global* parameter sensitivity analysis to gain insight into the general model response to changes in the  
24 MEA parameters, not only limited to cell performance, but also regarding the predicted heat and water balance.  
25 **Linking the scatter in material parameterizations found here with an extensive uncertainty propagation analysis**  
26 **in Part II reveals just how critical better experimental characterization of the ionomer is for predictive PEMFC**  
27 **modeling.**  
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41  
42  
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## 44 References

- 45  
46  
47 [1] R. Vetter and J. O. Schumacher, in *34<sup>th</sup> PSI Electrochemistry Symposium* (Villigen, Switzerland, 2018),  
48 <https://doi.org/10.21256/zhaw-3621>.  
49  
50  
51 [2] A. Kusoglu and A. Z. Weber, in *Polymers for Energy Storage and Delivery: Polyelectrolytes for Batteries and*  
52 *Fuel Cells*, edited by K. A. Page, C. L. Soles, and J. Runt (American Chemical Society, 2012), vol. 1096, pp.  
53 175–199.  
54  
55  
56 [3] A. Kusoglu and A. Z. Weber, *Chem. Rev.* **117**, 987–1104 (2017).  
57  
58  
59 [4] W. Dai, H. Wang, X.-Z. Yuan, J. J. Martin, D. Yang, J. Qiao, and J. Ma, *Int. J. Hydrog. Energy* **34**, 9461–9478  
60 (2009), ISSN 03603199.  
61  
62  
63  
64  
65

- 1  
2  
3 [5] L. Liu, W. Chen, and Y. Li, *J. Membr. Sci.* **504**, 1–9 (2016).  
4  
5 [6] R. Marek and J. Straub, *Int. J. Heat Mass Transf.* **44**, 39–53 (2001).  
6  
7 [7] C. Min, X. Liu, Y. He, B. Yin, W. Jiang, and W. Tao, *J. Power Sources* **160**, 359–373 (2006).  
8  
9 [8] D. Zhao, F. Gao, P. Massonnat, M. Dou, and A. Miraoui, *IEEE Trans. Energy Convers.* **30**, 1008–1018 (2015).  
10  
11 [9] B. Laoun, M. W. Naceur, A. Khellaf, and A. M. Kannan, *Int. J. Hydrog. Energy* **41**, 9521–9528 (2016).  
12  
13 [10] R. Vetter and J. O. Schumacher, *Comput. Phys. Commun.* **234**, 223–234 (2019).  
14  
15 [11] R. Vetter and J. O. Schumacher, *J. Power Sources* (2019), in press,  
16  
17 <https://doi.org/10.1016/j.jpowsour.2019.04.057>.  
18  
19 [12] K. C. Neyerlin, W. Gu, J. Jorne, and H. A. Gasteiger, *J. Electrochem. Soc.* **154**, B631–B635 (2007).  
20  
21 [13] K. C. Neyerlin, G. Wenbin, J. Jorne, and H. A. Gasteiger, *J. Electrochem. Soc.* **153**, A1955–A1963 (2006).  
22  
23 [14] R. Flückiger, Ph.D. thesis, ETH Zürich (2009).  
24  
25 [15] M. W. Chase, *NIST-JANAF Thermochemical Tables* (American Institute of Physics, 1998), 4th ed.  
26  
27 [16] M. J. Lampinen and M. Fomino, *J. Electrochem. Soc.* **140**, 3537–3546 (1993).  
28  
29 [17] P. Gode, F. Jaouen, G. Lindbergh, A. Lundblad, and G. Sundholm, *Electrochim. Acta* **48**, 4175–4187 (2003).  
30  
31 [18] R. Schweiss, C. Meiser, T. Damjanovic, I. Galbati, and N. Haak, *SIGRACET*  
32  
33 *Gas Diffusion Layers for PEM Fuel Cells, Electrolyzers and Batteries*,  
34  
35 [http://www.sglgroup.com/cms/\\_common/downloads/products/product-groups/su/fuel-cell-components/White-P](http://www.sglgroup.com/cms/_common/downloads/products/product-groups/su/fuel-cell-components/White-P)  
36  
37 (2016), accessed: 2018-11-24.  
38  
39 [19] W. Y. Hsu, J. R. Barkley, and P. Meakin, *Macromolecules* **13**, 198–200 (1980).  
40  
41 [20] T. E. Springer, T. A. Zawodzinski, and S. Gottesfeld, *J. Electrochem. Soc.* **8**, 2334–2341 (1991).  
42  
43 [21] D. R. Morris and X. Sun, *J. Appl. Polym. Sci.* **50**, 1445–1452 (1993).  
44  
45 [22] Y. Sone, P. Ekdunge, and D. Simonsson, *J. Electrochem. Soc.* **143**, 1254–1259 (1996).  
46  
47 [23] M. Eikerling, Y. I. Kharkats, A. A. Kornyshev, and Y. M. Volkovich, *J. Electrochem. Soc.* **145**, 2684–2699  
48  
49 (1998).  
50  
51 [24] T. Thampan, S. Malhotra, H. Tang, and R. Datta, *J. Electrochem. Soc.* **147**, 3242–3250 (2000).  
52  
53 [25] P. Costamagna, *Chem. Eng. Sci.* **56**, 323–332 (2001).  
54  
55 [26] C. A. Edmondson and J. J. Fontanella, *Solid State Ion.* **152–153**, 355–361 (2002).  
56  
57 [27] A. A. Kulikovskiy, *J. Electrochem. Soc.* **150**, A1432–A1439 (2003).  
58  
59  
60  
61  
62  
63  
64  
65



- 1  
2  
3 [28] A. Z. Weber and J. Newman, *J. Electrochem. Soc.* **151**, A311–A325 (2004).  
4  
5 [29] F. Meier and G. Eigenberger, *Electrochim. Acta* **49**, 1731–1742 (2004).  
6  
7 [30] C. Yang, S. Srinivasan, A. B. Bocarsly, S. Tulyani, and J. B. Benziger, *J. Membr. Sci.* **237**, 145–161 (2004).  
8  
9 [31] P. Choi, N. H. Jalani, and R. Datta, *J. Electrochem. Soc.* **152**, E123–E130 (2005).  
10  
11 [32] J. Fimrite, B. Carnes, H. Struchtrup, and N. Djilali, *J. Electrochem. Soc.* **152**, A1815–A1823 (2005).  
12  
13 [33] G. S. Hwang, M. Kaviani, J. H. Nam, M. H. Kim, and S. Y. Son, *J. Electrochem. Soc.* **156**, B1192–B1200  
14 (2009).  
15  
16 [34] L. Maldonado, J.-C. Perrin, J. Dillet, and O. Lottin, *J. Membr. Sci.* **389**, 43–56 (2012).  
17  
18 [35] Q. Zhao, N. Carro, H. Y. Ryu, and J. Benziger, *Polymer* **53**, 1267–1276 (2012).  
19  
20 [36] L. Holzer, O. Pecho, J. Schumacher, P. Marmet, O. Stenzel, F. Büchi, A. Lamibrac, and B. Münch, *Electrochim. Acta* **227**, 419–434 (2017).  
21  
22 [37] K. S. Babu, H. T. Chung, P. Zelenay, and S. Litster, *ACS Appl. Mater. Interfaces* **8**, 32764–32777 (2016).  
23  
24 [38] M. Khandelwal and M. M. Mench, *J. Power Sources* **161**, 1106–1115 (2006).  
25  
26 [39] M. L. Huber, R. A. Perkins, D. G. Friend, J. V. Sengers, M. J. Assael, I. N. Metaxa, K. Miyagawa, R. Hellmann,  
27 and E. Vogel, *J. Phys. Chem. Ref. Data* **41**, 033102 (2012).  
28  
29 [40] N. Zamel and X. Li, *Prog. Energy Combust. Sci.* **39**, 111–146 (2013).  
30  
31 [41] J. Yuan and B. Sundén, *Int. J. Heat Mass Transf.* **58**, 441–456 (2013).  
32  
33 [42] N. Alhazmi, M. Ismail, D. Ingham, K. Hughes, L. Ma, and M. Pourkashanian, *J. Power Sources* **241**, 136–145  
34 (2013).  
35  
36 [43] N. Alhazmi, D. B. Ingham, M. S. Ismail, K. Hughes, L. Ma, and M. Pourkashanian, *J. Power Sources* **270**,  
37 59–67 (2014).  
38  
39 [44] M. Ahadi, M. Tam, M. S. Saha, J. Stumper, and M. Bahrani, *J. Power Sources* **354**, 207–214 (2017).  
40  
41 [45] O. Burheim, P. J. S. Vie, J. G. Pharoah, and S. Kjelstrup, *J. Power Sources* **195**, 249–256 (2010).  
42  
43 [46] O. S. Burheim, J. G. Pharoah, H. Lampert, P. J. S. Vie, and S. Kjelstrup, *J. Fuel Cell Sci. Technol.* **8**, 021013  
44 (2011).  
45  
46 [47] R. L. Rowley, W. V. Wilding, J. L. Oscarson, Y. Yang, N. A. Zundel, T. E. Daubert, and R. P. Danner,  
47 *DIPPR<sup>®</sup> Data Compilation of Pure Chemical Properties* (Design Institute for Physical Properties, AIChE,  
48 New York, 2007).  
49  
50 [48] D. W. Green and R. H. Perry, *Perry’s Chemical Engineers’ Handbook* (McGraw-Hill, 2008), 8th ed.  
51  
52  
53  
54  
55  
56  
57  
58  
59  
60  
61  
62  
63  
64  
65

- 1  
2  
3 [49] T. A. Zawodzinski, M. Neeman, L. O. Sillerud, and S. Gottesfeld, *J. Phys. Chem.* **95**, 6040–6044 (1991).  
4  
5 [50] K. D. Kreuer, *J. Membr. Sci.* **185**, 29–39 (2001).  
6  
7 [51] X. Gong, A. Bandis, A. Tao, G. Meresi, Y. Wang, P. T. Inglefield, A. A. Jones, and W.-Y. Wen, *Polymer* **42**,  
8 6485–6492 (2001).  
9  
10 [52] S. Tsushima, K. Teranishi, and S. Hirai, *Energy* **30**, 235–245 (2005).  
11  
12 [53] J.-C. Perrin, S. Lyonnard, and F. Volino, *J. Phys. Chem. C* **111**, 3393–3404 (2007).  
13  
14 [54] K. D. Kreuer, M. Schuster, B. Obliers, O. Diat, U. Traub, A. Fuchs, U. Klock, S. J. Paddison, and J. Maier,  
15 *J. Power Sources* **178**, 499–509 (2008).  
16  
17 [55] N. S. Schneider and D. Rivin, *Polymer* **51**, 671–678 (2010).  
18  
19 [56] K. Kidena, T. Ohkubo, N. Takimoto, and A. Ohira, *Eur. Polym. J.* **46**, 450–455 (2010).  
20  
21 [57] Z. Ma, R. Jiang, M. E. Myers, E. L. Thompson, and C. S. Gittleman, *J. Mater. Chem.* **21**, 9302–9311 (2011).  
22  
23 [58] Q. Zhao, P. Majsztrik, and J. Benziger, *J. Phys. Chem. B* **115**, 2717–2727 (2011).  
24  
25 [59] D. Rivin, C. E. Kendrick, P. W. Gibson, and N. S. Schneider, *Polymer* **42**, 623–635 (2001).  
26  
27 [60] K. Aotani, S. Miyazaki, N. Kubo, and M. Katsuta, *ECS Trans.* **16**, 341–352 (2008).  
28  
29 [61] Y. Tabuchi, R. Ito, S. Tsushima, and S. Hirai, *J. Power Sources* **196**, 652–658 (2011).  
30  
31 [62] C. K. Mittelsteadt and J. Staser, *ECS Trans.* **41**, 101–121 (2011).  
32  
33 [63] T. F. Fuller, Ph.D. thesis, University of California, Berkeley (1992).  
34  
35 [64] S. Motupally, A. J. Becker, and J. W. Weidner, *J. Electrochem. Soc.* **147**, 3171–3177 (2000).  
36  
37 [65] S. Ge, X. Li, B. Yi, and I.-M. Hsing, *J. Electrochem. Soc.* **152**, A1149–A1157 (2005).  
38  
39 [66] T. D. Myles, A. M. Kiss, K. N. Grew, A. A. Peracchio, G. J. Nelson, and W. K. S. Chiu, *J. Electrochem. Soc.*  
40 **158**, B790–B796 (2011).  
41  
42 [67] A. Lokkiluoto and M. M. Gasik, *Int. J. Hydrog. Energy* **38**, 10–19 (2013).  
43  
44 [68] X. Ye and M. D. LeVan, *J. Membr. Sci.* **221**, 147–161 (2003).  
45  
46 [69] D. A. Caulk, A. M. Brenner, and S. M. Clapham, *J. Electrochem. Soc.* **159**, F518–F529 (2012).  
47  
48 [70] R. Mills, *Ber. Bunsenges. Phys. Chem.* **75**, 195–199 (1971).  
49  
50 [71] M. Holz, S. R. Heil, and A. Sacco, *Phys. Chem. Chem. Phys.* **2**, 4740–4742 (2000).  
51  
52 [72] A. C. Olesen, T. Berning, and S. K. Kaer, *ECS Trans.* **50**, 979–991 (2012).  
53  
54 [73] T. Takamatsu, M. Hashiyama, and A. Eisenberg, *J. Appl. Polym. Sci.* **24**, 2199–2220 (1979).  
55  
56  
57  
58  
59  
60  
61  
62  
63  
64  
65

- 1  
2  
3 [74] A. Guillermo, G. Gebel, H. Mendil-Jakani, and E. Pinton, *J. Phys. Chem. B* **113**, 6710–6717 (2009).  
4  
5 [75] J. H. Wang, *J. Am. Chem. Soc.* **73**, 510–513 (1951).  
6  
7 [76] S. C. Yeo and A. Eisenberg, *J. Appl. Polym. Sci.* **21**, 875–898 (1977).  
8  
9 [77] The Chemours Company, *Nafion<sup>TM</sup> NR211 and NR212 - Solution Cast Membranes (Product Bulletin P-11)*,  
10 [https://nafionstore-us.americommerce.com/Shared/P11\\_C10610\\_Nafion\\_NR-211\\_\\_NR-212\\_P11.pdf](https://nafionstore-us.americommerce.com/Shared/P11_C10610_Nafion_NR-211__NR-212_P11.pdf)  
11 (2016), accessed: 2018-11-24.  
12  
13 [78] J. Peron, A. Mani, X. Zhao, D. Edwards, M. Adachi, T. Soboleva, Z. Shi, Z. Xie, T. Navessin, and S. Holdcroft,  
14 *J. Membr. Sci.* **356**, 44–51 (2010).  
15  
16 [79] T. F. Fuller and J. Newman, *J. Electrochem. Soc.* **139**, 1332–1337 (1992).  
17  
18 [80] S. Dutta, S. Shimpalee, and J. Van Zee, *Int. J. Heat Mass Transf.* **44**, 2029–2042 (2001).  
19  
20 [81] H. P. L. H. van Bussel, F. G. H. Koene, and R. K. A. M. Mallant, *J. Power Sources* **71**, 218–222 (1998).  
21  
22 [82] S. Ge, B. Yi, and P. Ming, *J. Electrochem. Soc.* **153**, A1443–A1450 (2006).  
23  
24 [83] T. A. Zawodzinski, J. Davey, J. Valerio, and S. Gottesfeld, *Electrochim. Acta* **40**, 297–302 (1995).  
25  
26 [84] D. Weng, J. S. Wainright, U. Landau, and R. F. Savinell, *J. Electrochem. Soc.* **143**, 1260–1263 (1996).  
27  
28 [85] X. Ye and C.-Y. Wang, *J. Electrochem. Soc.* **154**, B676–B682 (2007).  
29  
30 [86] W. Braff and C. K. Mittelsteadt, *ECS Trans.* **16**, 309–316 (2008).  
31  
32 [87] Y. Park and J. Caton, *Int. J. Hydrog. Energy* **33**, 7513–7520 (2008).  
33  
34 [88] Z. Peng, A. Morin, P. Huguet, P. Schott, and J. Pauchet, *J. Phys. Chem. B* **115**, 12835–12844 (2011).  
35  
36 [89] F. Xu, S. Leclerc, D. Stemmelen, J.-C. Perrin, A. Retournard, and D. Canet, *J. Membr. Sci.* **536**, 116–122  
37 (2017).  
38  
39 [90] Y.-K. Choe, E. Tsuchida, T. Ikeshoji, S. Yamakawa, and S.-a. Hyodo, *J. Phys. Chem. B* **112**, 11586–11594  
40 (2008).  
41  
42 [91] T. A. Zawodzinski, T. E. Springer, J. Davey, R. Jestel, C. Lopez, J. Valerio, and S. Gottesfeld, *J. Electrochem.*  
43 *Soc.* **140**, 1981–1985 (1993).  
44  
45 [92] G. Xie, *J. Chem. Soc.* **142**, 3057–3062 (1995).  
46  
47 [93] X. Ren, W. Henderson, and S. Gottesfeld, *J. Electrochem. Soc.* **144**, L267–L270 (1997).  
48  
49 [94] X. Ren and S. Gottesfeld, *J. Electrochem. Soc.* **148**, A87–A93 (2001).  
50  
51 [95] T. Okada, S. Møller-Holst, O. Gorseth, and S. Kjelstrup, *J. Electroanal. Chem.* **442**, 137–145 (1998).  
52  
53 [96] M. Ise, K. D. Kreuer, and J. Maier, *Solid State Ion.* **125**, 213–223 (1999).  
54  
55  
56  
57  
58  
59  
60  
61  
62  
63  
64  
65

- 1  
2  
3 [97] Z. Luo, Z. Chang, Y. Zhang, Z. Liu, and J. Li, *Int. J. Hydrog. Energy* **35**, 3120–3124 (2010).  
4  
5 [98] H. P. L. H. Van Bussel, F. G. H. Koene, and R. K. A. M. Mallant, *J. Power Sources* **71**, 218–222 (1998).  
6  
7 [99] H. Meng and C.-Y. Wang, *J. Electrochem. Soc.* **151**, A358–A367 (2004).  
8  
9 [100] M. Tasaka, T. Mizuta, and O. Sekiguchi, *J. Membr. Sci.* **54**, 191–204 (1990).  
10  
11 [101] V. M. Barragán and S. Kjelstrup, *J. Non-Equilib. Thermodyn.* **42** 217–236 (2017).  
12  
13 [102] J. P. G. Villaluenga, B. Seoane, V. M. Barragán, and C. Ruiz-Bauzá, *J. Membr. Sci.* **274**, 116–122 (2006).  
14  
15 [103] S. Kim and M. M. Mench, *J. Membr. Sci.* **328**, 113–120 (2009).  
16  
17 [104] M. Tasaka, T. Urata, R. Kiyono, and Y. Aki, *J. Membr. Sci.* **67**, 83–91 (1992).  
18  
19 [105] T. Suzuki, R. Kiyono, and M. Tasaka, *J. Membr. Sci.* **92**, 85–93 (1994).  
20  
21 [106] J. T. Hinatsu, M. Mizuhata, and H. Takenaka, *J. Electrochem. Soc.* **141**, 1493–1498 (1994).  
22  
23 [107] P. Futerko and I.-M. Hsing, *J. Electrochem. Soc.* **146**, 2049–2053 (1999).  
24  
25 [108] J. P. Meyers and J. Newman, *J. Electrochem. Soc.* **149**, A710–A717 (2002).  
26  
27 [109] P. Choi and R. Datta, *J. Electrochem. Soc.* **150**, E601–E607 (2003).  
28  
29 [110] T. A. Zawodzinski, S. Gottesfeld, S. Shoichet, and T. J. McCarthy, *J. Appl. Electrochem.* **23**, 86–88 (1993).  
30  
31 [111] P. Choi, N. H. Jalani, and R. Datta, *J. Electrochem. Soc.* **152**, E84–E89 (2005).  
32  
33 [112] H. Takata, N. Mizuno, M. Nishikawa, S. Fukada, and M. Yoshitake, *Int. J. Hydrog. Energy* **32**, 371–379  
34  
35 (2007).  
36  
37 [113] P. Costamagna, S. Grosso, and R. Di Felice, *J. Power Sources* **178**, 537–546 (2008).  
38  
39 [114] S. Ochi, O. Kamishima, J. Mizusaki, and J. Kawamura, *Solid State Ion.* **180**, 580–584 (2009).  
40  
41 [115] A. Kusoglu, M. H. Santare, and A. M. Karlsson, *Polymer* **50**, 2481–2491 (2009).  
42  
43 [116] C. K. Mittelsteadt and H. Liu, in *Handbook of Fuel Cells: Fundamentals, Technology, and Applications*, edited  
44  
45 by W. Vielstich, H. A. Gasteiger, A. Lamm, and H. Yokokawa (John Wiley & Sons, Hoboken, N.J., 2010).  
46  
47 [117] M. H. Eikerling and P. Berg, *Soft Matter* **7**, 5976–5990 (2011).  
48  
49 [118] Y. Li, Q. T. Nguyen, C. L. Buquet, D. Langevin, M. Legras, and S. Marais, *J. Membr. Sci.* **439**, 1–11 (2013).  
50  
51 [119] K.-D. Kreuer, *Solid State Ion.* **252**, 93–101 (2013).  
52  
53 [120] K.-D. Kreuer, *Solid State Ion.* **328**, 35–36 (2018).  
54  
55 [121] S. Didierjean, J. C. Perrin, F. Xu, G. Maranzana, M. Klein, J. Mainka, and O. Lottin, *J. Power Sources* **300**,  
56  
57 50–56 (2015).  
58  
59  
60  
61  
62  
63  
64  
65

- 1  
2  
3 [122] S. Shi, A. Z. Weber, and A. Kusoglu, *Electrochim. Acta* **220**, 517–528 (2016).  
4  
5 [123] A. Morin, G. Gebel, L. Porcar, Z. Peng, N. Martinez, A. Guillermo, and S. Lyonnard, *J. Electrochem. Soc.*  
6 **164**, F9–F21 (2017).  
7  
8 [124] W. Wagner and A. Pruss, *J. Phys. Chem. Ref. Data* **22**, 783–787 (1993).  
9  
10 [125] M. Escoubes and M. Pineri, in *Perfluorinated Ionomer Membranes*, edited by A. Eisenberg and H. L. Yeager  
11 (American Chemical Society, 1982), vol. 180, pp. 9–23.  
12  
13 [126] K. K. Pushpa, D. Nandan, and R. M. Iyer, *J. Chem. Soc. Faraday Trans. 1* **84**, 2047–2056 (1988).  
14  
15 [127] N. H. Jalani and R. Datta, *J. Membr. Sci.* **264**, 167–175 (2005).  
16  
17 [128] A. Kusoglu, A. Kwong, K. T. Clark, H. P. Gunterman, and A. Z. Weber, *J. Electrochem. Soc.* **159**, F530–F535  
18 (2012).  
19  
20 [129] A. Kusoglu, B. L. Kienitz, and A. Z. Weber, *J. Electrochem. Soc.* **158**, B1504–B1514 (2011).  
21  
22 [130] C. Yang, S. Srinivasan, A. B. Bocarsly, S. Tulyani, and J. B. Benziger, *J. Membr. Sci.* **237**, 145–161 (2004).  
23  
24 [131] V. Freger, E. Korin, J. Wisniak, and E. Korngold, *J. Membr. Sci.* **128**, 151–162 (1997).  
25  
26 [132] P. H. Emmett and S. Brunauer, *J. Am. Chem. Soc.* **59**, 1553–1564 (1937).  
27  
28 [133] J. Divisek, M. Eikerling, V. Mazin, H. Schmitz, U. Stimming, and Y. M. Volfkovich, *J. Electrochem. Soc.*  
29 **145**, 2677–2683 (1998).  
30  
31 [134] G. S. Hwang and A. Z. Weber, *J. Electrochem. Soc.* **159**, F683–F692 (2012).  
32  
33 [135] H. Fathi, A. Raoof, S. H. Mansouri, and M. T. van Genuchten, *J. Electrochem. Soc.* **164**, F298–F305 (2017).  
34  
35 [136] D. Wood, J. Davey, P. Atanassov, and R. Borup, *ECS Trans.* **3**, 753–763 (2006).  
36  
37 [137] Y. Ono, A. Ohma, K. Shinohara, and K. Fushinobu, *J. Electrochem. Soc.* **160**, F779–F787 (2013).  
38  
39 [138] C. Si, X.-D. Wang, W.-M. Yan, and T.-H. Wang, *J. Chem.* **2015**, 1–17 (2015).  
40  
41 [139] X. Wang and T. V. Nguyen, *J. Electrochem. Soc.* **155**, B1085–B1092 (2008).  
42  
43 [140] J. T. Gostick, M. A. Ioannidis, M. W. Fowler, and M. D. Pritzker, *J. Power Sources* **194**, 433–444 (2009).  
44  
45 [141] N. Zamel, X. Li, J. Becker, and A. Wiegmann, *Int. J. Hydrog. Energy* **36**, 5466–5478 (2011).  
46  
47 [142] L. Holzer, O. Pecho, J. Schumacher, P. Marmet, F. Büchi, A. Lamibrac, and B. Münch, *Electrochim. Acta*  
48 **241**, 414–432 (2017).  
49  
50 [143] J. T. Gostick, M. W. Fowler, M. D. Pritzker, M. A. Ioannidis, and L. M. Behra, *J. Power Sources* **162**,  
51 228–238 (2006).  
52  
53 [144] J. S. Yi and T. V. Nguyen, *J. Electrochem. Soc.* **146**, 38–45 (1999).  
54  
55  
56  
57  
58  
59  
60  
61  
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63  
64  
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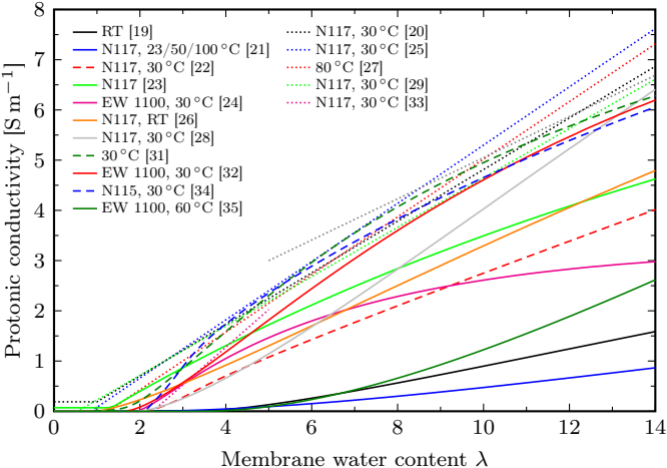
- 1  
2  
3 [145] M. L. Huber, R. A. Perkins, A. Laesecke, D. G. Friend, J. V. Sengers, M. J. Assael, I. N. Metaxa, E. Vogel,  
4 R. Mareš, and K. Miyagawa, *J. Phys. Chem. Ref. Data* **38**, 101–125 (2009).  
5  
6 [146] P. Berg, K. Promislov, J. S. Pierre, J. Stumper, and B. Wetton, *J. Electrochem. Soc.* **151**, A341–A353 (2004).  
7  
8 [147] M. B. Satterfield and J. B. Benziger, *J. Phys. Chem. B* **112**, 3693–3704 (2008).  
9  
10 [148] C. W. Monroe, T. Romero, W. Mérida, and M. Eikerling, *J. Membr. Sci.* **324**, 1–6 (2008).  
11  
12 [149] D. T. Hallinan and Y. A. Elabd, *J. Phys. Chem. B* **113**, 4257–4266 (2009).  
13  
14 [150] M. Adachi, T. Navessin, Z. Xie, F. H. Li, S. Tanaka, and S. Holdcroft, *J. Membr. Sci.* **364**, 183–193 (2010).  
15  
16 [151] A. Kongkanand, *J. Phys. Chem. C* **115**, 11318–11325 (2011).  
17  
18 [152] Q. He, A. Kusoglu, I. T. Lucas, K. Clark, A. Z. Weber, and R. Kostecky, *J. Phys. Chem. B* **115**, 11650–11657  
19 (2011).  
20  
21 [153] B. Kientiz, H. Yamada, N. Nonoyama, and A. Z. Weber, *J. Fuel Cell Sci. Technol.* **8**, 011013 (2011).  
22  
23 [154] P. W. Majsztrik, M. B. Satterfield, A. B. Bocarsly, and J. B. Benziger, *J. Membr. Sci.* **301**, 93–106 (2007).  
24  
25 [155] T. Okada, *J. Electroanal. Chem.* **465**, 1–17 (1999).  
26  
27 [156] T. V. Nguyen and R. E. White, *J. Electrochem. Soc.* **140**, 2178–2186 (1993).  
28  
29 [157] T. V. Nguyen, in *Tutorials in Electrochemical Engineering–Mathematical Modeling*, edited by R. F. Savinell,  
30 J. M. Fenton, A. C. West, S. L. Scanlon, and J. Weidner (The Electrochemical Society Proceedings Series,  
31 1999), vol. 99-14, pp. 222–241.  
32  
33 [158] D. Song, Q. Wang, Z.-S. Liu, and C. Huang, *J. Power Sources* **159**, 928–942 (2006).  
34  
35 [159] W. He, J. S. Yi, and T. V. Nguyen, *AIChE J.* **46**, 2053–2064 (2000).  
36  
37 [160] H. Meng, *J. Power Sources* **168**, 218–228 (2007).  
38  
39 [161] T. V. Nguyen and W. He, in *Handbook of Fuel Cells: Fundamentals, Technology, and Applications*, edited by  
40 W. Vielstich, H. A. Gasteiger, A. Lamm, and H. Yokokawa (John Wiley & Sons, Hoboken, N.J., 2010).  
41  
42 [162] D. Natarajan and T. V. Nguyen, *J. Electrochem. Soc.* **148**, A1324–A1335 (2001).  
43  
44 [163] J. H. Nam and M. Kaviany, *Int. J. Heat Mass Transf.* **46**, 4595–4611 (2003).  
45  
46 [164] A. Z. Weber, R. M. Darling, and J. Newman, *J. Electrochem. Soc.* **151**, A1715–A1727 (2004).  
47  
48 [165] E. Birgersson, M. Noponen, and M. Vynnycky, *J. Electrochem. Soc.* **152**, A1021–A1034 (2005).  
49  
50 [166] M. Eikerling, A. A. Kornyshev, and A. R. Kucernak, *Phys. Today* **59**, 38–44 (2006).  
51  
52 [167] W. H., X. Li, and P. Berg, *Electrochim. Acta* **54**, 6913–6927 (2009).  
53  
54  
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57  
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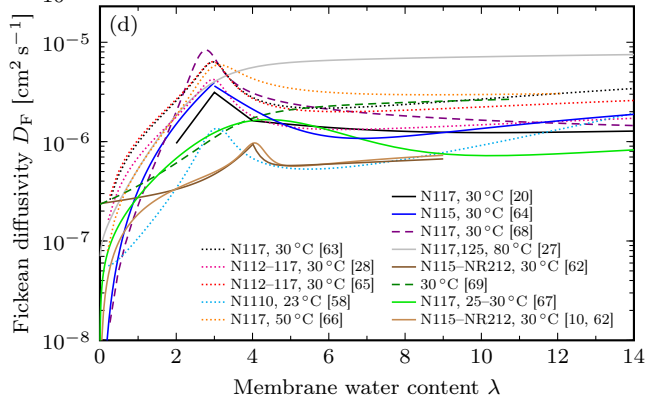
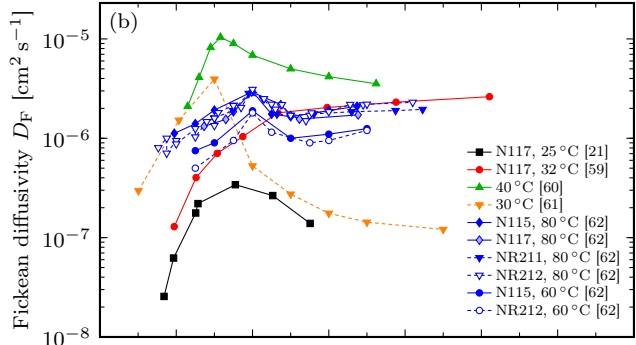
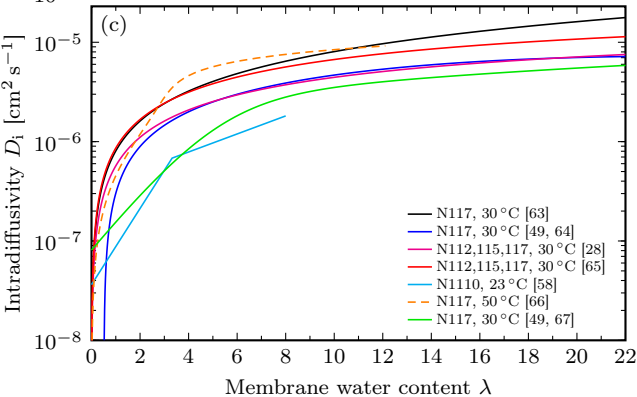
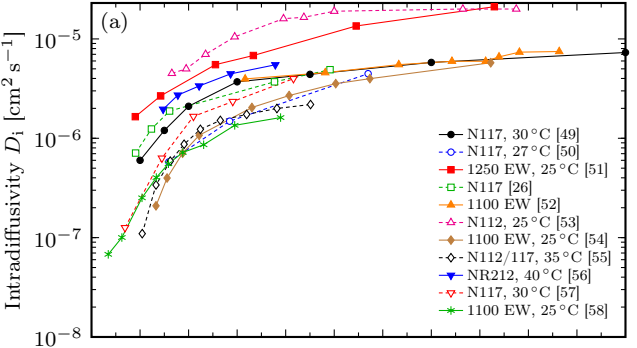
- 1  
2  
3 [168] I. V. Zenyuk, A. Lamibrac, J. Eller, D. Y. Parkinson, F. Marone, F. N. Büchi, and A. Z. Weber, *J. Phys.*  
4 *Chem. C* **120**, 28701–28711 (2016).  
5  
6 [169] M. A. Safi, N. I. Prasianakis, J. Mantzaras, A. Lamibrac, and F. N. Büchi, *Int. J. Heat Mass Transf.* **115**,  
7 238–249 (2017).  
8  
9 [170] S. Lal, A. Lamibrac, J. Eller, and F. N. Büchi, *J. Electrochem. Soc.* **165**, F652–F661 (2018).  
10  
11 [171] R. Hołyst, M. Litniewski, and D. Jakubczyk, *Soft Matter* **11**, 7201–7206 (2015).  
12  
13 [172] A. Z. Weber and J. Newman, *Chem. Rev.* **104**, 4679–4726 (2004).  
14  
15 [173] S. Basu, C.-Y. Wang, and K. S. Chen, *J. Electrochem. Soc.* **156**, B748–B756 (2009).  
16  
17 [174] W. M. Haynes, *CRC Handbook of Chemistry and Physics* (Taylor & Francis, 2016), 97th ed.  
18  
19 [175] V. E. Ostrovskii and B. V. Gostev, *J. Therm. Anal.* **46**, 397–416 (1996).  
20  
21 [176] P. J. Reucroft, D. Rivin, and N. S. Schneider, *Polymer* **43**, 5157–5161 (2002).  
22  
23 [177] W. L. and P. Jannasch, *J. Phys. Chem. B* **117**, 8561–8570 (2013).  
24  
25 [178] SGL Technologies GmbH, *SIGRACET GDL 24 & 25 Series Gas Diffusion Layer*,  
26 [http://www.fuelcellstore.com/spec-sheets/SGL-GDL\\_24-25.pdf](http://www.fuelcellstore.com/spec-sheets/SGL-GDL_24-25.pdf) (2009), accessed: 2018-11-24.  
27  
28 [179] E. C. Kumbur, K. V. Sharp, and M. M. Mench, *J. Electrochem. Soc.* **154**, B1305–B1314 (2007).  
29  
30 [180] O. S. Burheim, H. Su, H. H. Hauge, S. Pasupathi, and B. G. Pollet, *Int. J. Hydrog. Energy* **39**, 9397–9408  
31 (2014).  
32  
33 [181] D. M. Bernardi and M. W. Verbrugge, *J. Electrochem. Soc.* **139**, 2477–2491 (1992).  
34  
35 [182] R. Flückiger, S. A. Freunberger, D. Kramer, A. Wokaun, G. G. Scherer, and F. N. Büchi, *Electrochim. Acta*  
36 **54**, 551–559 (2008).  
37  
38 [183] S. Litster, W. K. Epting, E. A. Wargo, S. R. Kalidindi, and E. C. Kumbur, *Fuel Cells* **13**, 935–945 (2013).  
39  
40 [184] I. Nitta, O. Himanen, and M. Mikkola, *Electrochem. Commun.* **10**, 47–51 (2008).  
41  
42 [185] P. Zhou, C. W. Wu, and G. J. Ma, *J. Power Sources* **159**, 1115–1122 (2006).  
43  
44 [186] H. Sadeghifar, N. Djilali, and M. Bahrani, *J. Power Sources* **248**, 632–641 (2014).  
45  
46 [187] R. Vetter and J. O. Schumacher, in *14<sup>th</sup> Symposium on Fuel Cell and Battery Modelling and Experimental*  
47 *Validation (ModVal 14)* (Karlsruhe, Germany, 2017), <https://doi.org/10.21256/zhaw-3634>.  
48  
49 [188] A. Majumdar and C. L. Tien, *J. Heat Transfer* **113**, 516–525 (1991).  
50  
51 [189] Y. Zhou, G. Lin, A. J. Shih, and S. J. Hu, *J. Power Sources* **163**, 777–783 (2007).  
52  
53 [190] Z. Wu, Y. Zhou, G. Lin, S. Wang, and S. J. Hu, *J. Power Sources* **182**, 265–269 (2008).  
54  
55  
56  
57  
58  
59  
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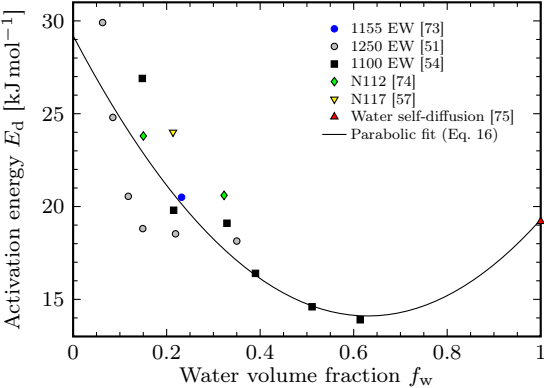
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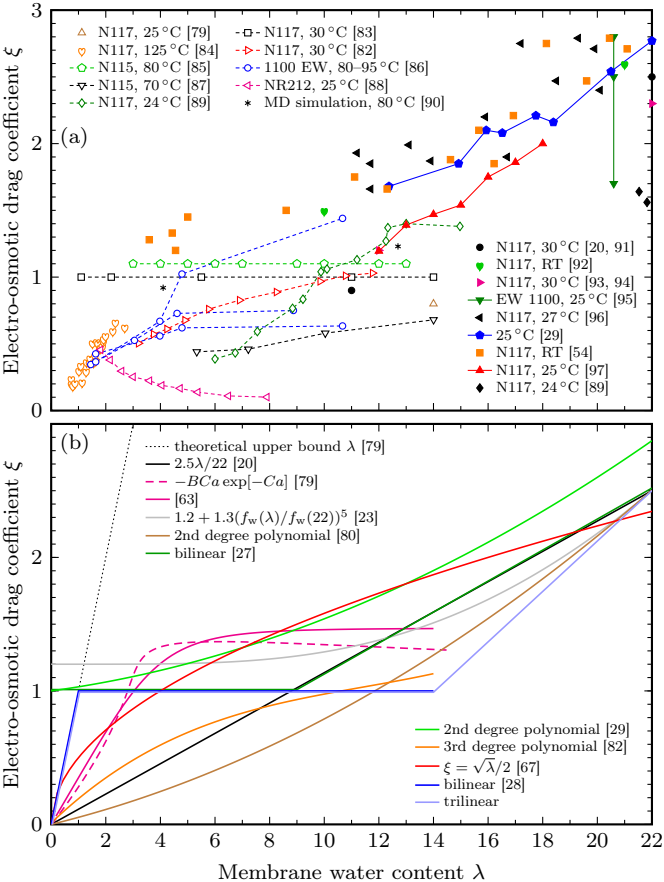
[191] T. A. Zawodzinski, C. Derouin, S. Radzinski, R. J. Sherman, V. T. Smith, T. E. Springer, and S. Gottesfeld, *J. Electrochem. Soc.* **140**, 1041–1047 (1993).



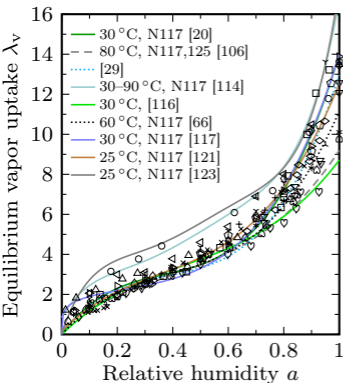




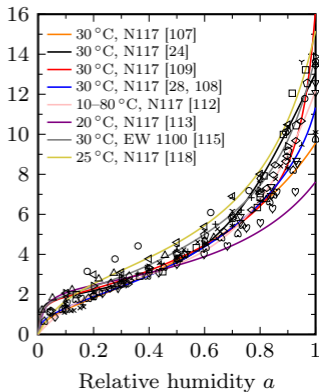




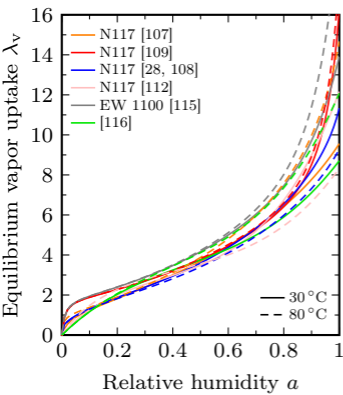
(a) Polynomial fits



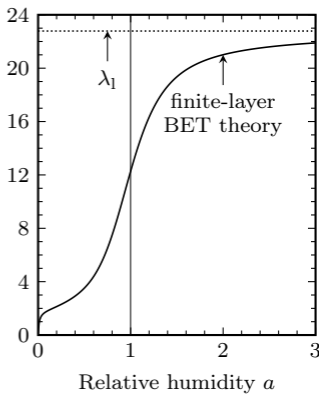
(b) Sorption models



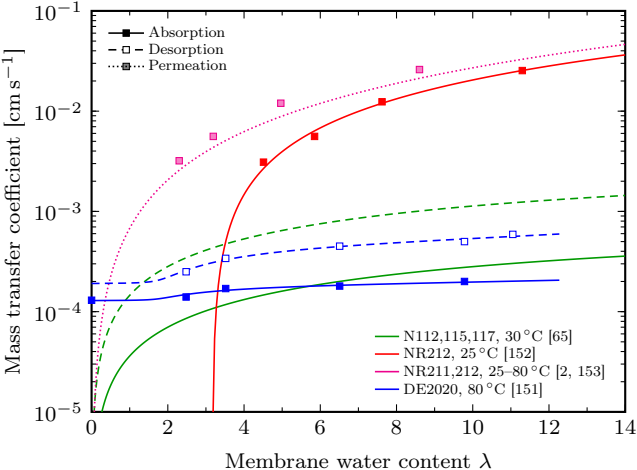
(c) Temperature dependence

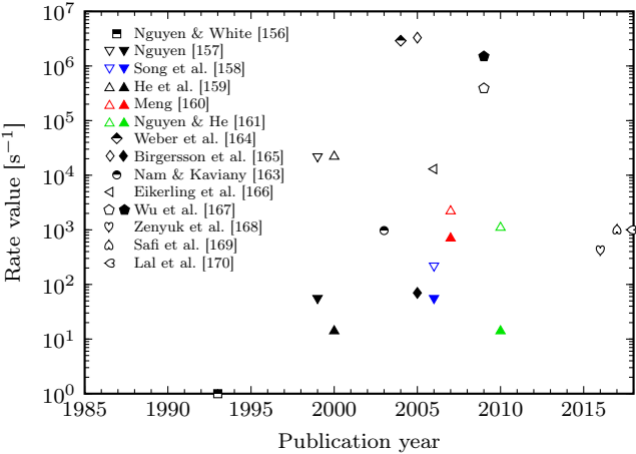


(d) Oversaturation

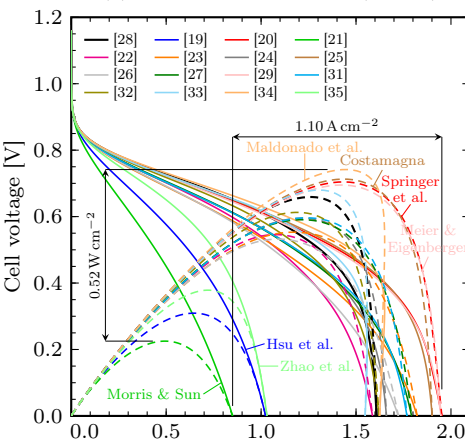


▽ 20 °C, N120 [125]	○ 25 °C, N117 [126]	□ 30 °C, N117 [20]	△ 25 °C, N117 [21]
◇ 20 °C, N117 [59]	▷ 25 °C, N117 [127]	◁ 90 °C, N117 [127]	◊ 30 °C, N117 [112]
⊖ 80 °C, N117 [106]	▽ 80 °C, N125 [106]	× 80 °C, N1110 [62]	○ 25 °C, NR211 [78]
▲ 40 °C, NR212 [56]	▽ 60 °C, NR212 [56]	+ 80 °C, NR212 [62]	* 25 °C, NR212 [128]

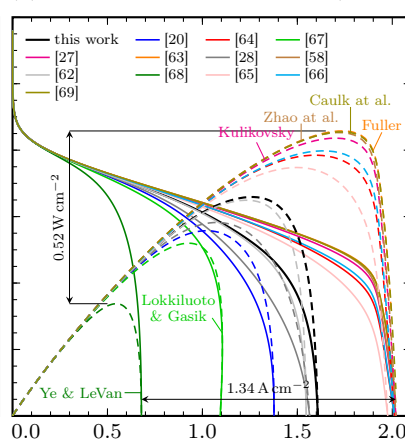




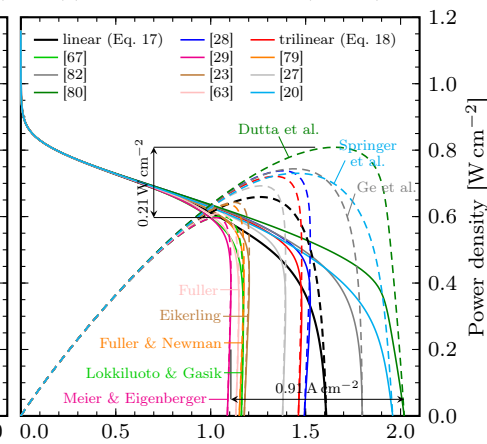
(a) Protonic conductivity (Tab. 2)



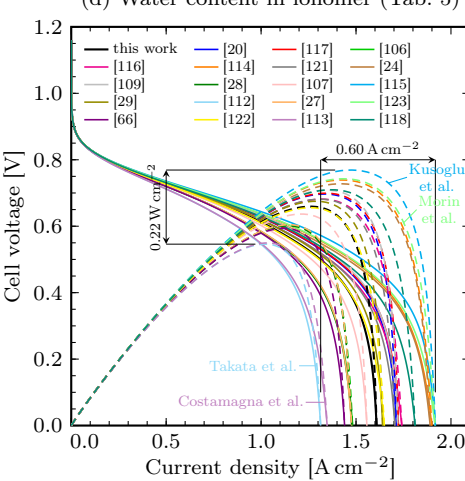
(b) Water diffusivity in ionomer (Tab. 3)



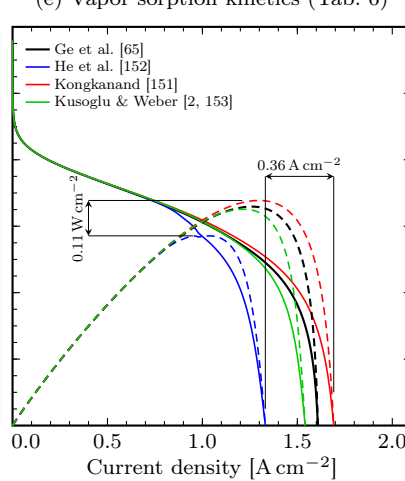
(c) Electro-osmotic drag (Tab. 4)



(d) Water content in ionomer (Tab. 5)



(e) Vapor sorption kinetics (Tab. 6)



(f) Evaporation/condensation (Tab. 7)

