

In-situ X-Ray Absorption Near Edge Structure Spectroscopy of a Solid Catalyst using a Laboratory-Based Set-up

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An *in-situ* laboratory-based X-ray Absorption Near Edge Structure (XANES) Spectroscopy set-up is presented, which allows performing long-term experiments on a solid catalyst at relevant reaction conditions of temperature and pressure. Complementary to research performed at synchrotron radiation facilities the approach is showcased for a Co/TiO₂ Fischer-

Tropsch Synthesis (FTS) catalyst. Supported cobalt metal nanoparticles next to a (very small) fraction of cobalt(II) titanate, which is an inactive phase for FTS, were detected, with no signs of re-oxidation of the supported cobalt metal nanoparticles during FTS at 523 K, 5 bar and 200 h, indicating that cobalt metal is maintained as the main active phase during FTS.

The developments in synchrotron radiation-based X-ray absorption (XAS) and X-ray emission (XES) techniques accomplished during the last decades, have positioned them as prominent approaches for in-depth characterization of materials and complex reactions systems.^[1] Specifically, for the characterization of solid catalysts the development of *in-situ* and *operando* spectroscopy as well as methods offering high spatial and temporal resolution has established X-rays techniques as crucial for studying the functionality of these materials.^[2,3] Among those techniques, X-ray Absorption Near Edge Structure (XANES) spectroscopy has become a powerful, highly versatile and useful characterization technique for the study of different catalytic systems of practical relevance.^[4-14]

However, one of the disadvantages of synchrotron-based techniques is their limited availability and accessibility, i.e. the

number of hours of beamtime access granted, which exclude certain characterization studies and developments. For instance, the restricted beamtime access typically confines *in-situ* studies of the deactivation of catalysts to a few days, which is often insufficient to conduct a conclusive study about long-term deactivation processes that can take place over much longer time frames, e.g. the long-term deactivation process of Fischer-Tropsch Synthesis (FTS) catalyst, which happens over days/weeks, and is a major challenge for FTS technology.^[15,16] Additional difficulties include the transport of the *in-situ* set-up and samples (e.g. air sensitive samples) to the synchrotron radiation facilities.^[17] To overcome this availability gap, during the last years research on the development of laboratory-based X-rays facilities has been intensified.^[17-26] By using laboratory-based set-ups, different successful *ex-situ* XANES studies have been performed on diverse systems.^[27-32] Interestingly, one can consider these developments as a revival of a somewhat older idea. For example, during the 1980s, Koningsberger *et al.* developed a laboratory-based EXAFS set-up, and successfully performed *ex-situ* experiments.^[33] However, limitations in the energy resolution of the monochromator crystal (~14 eV) limited the applicability of the set-up. We refer the reader for this older work to a book chapter by Koningsberger.^[34]

In the present study, we introduce the first *in-situ* XANES laboratory-based experiments of a solid catalyst, measured in transmission mode, under relevant conditions of temperature and pressure. We have performed such experiments for an industrial relevant Co/TiO₂ FTS catalyst. The experiments were compared and complemented by synchrotron-based measurements carried out on the same catalyst material. By using both set-ups, namely a laboratory-based and synchrotron-based system, it was possible to obtain fundamental physicochemical insights about the activation and deactivation process of a Co/TiO₂ FTS catalyst and compare the capabilities of both measurement approaches for collecting *in-situ* XANES data for this solid catalyst as well as relevant reference compounds, i.e. Co₃O₄, CoO, Co, and CoTiO₃.

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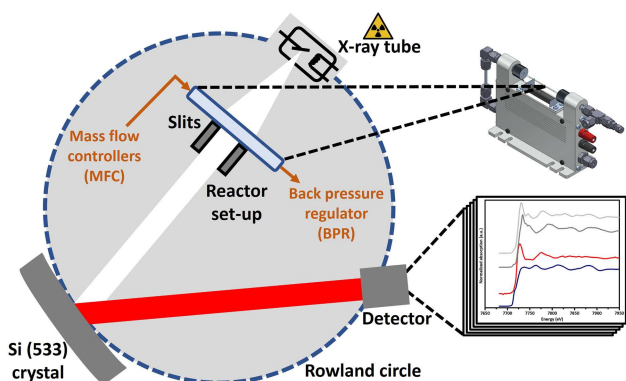


Figure 1. Schematic of the *in-situ* XANES set-up, principles and configuration.

Results and Discussion

The *in-situ* XANES laboratory-based set-up used in the experiments consists of an Ag-tube as the X-ray source, 40 mA and 20 kV, a spherically bent crystal monochromator utilizing the Johann geometry employing a Si (533) crystal with a bending radius $R=0.5$ m that was used for Co K-edge (7709 eV) measurements, and a NaI scintillator detector. The spectrometer uses the Bremsstrahlung spectrum and is designed to work in the 4–20 keV photon energy range. The detector, monochromator and X-ray source are located at a Rowland circle. The X-ray source is kept fixed and the monochromator and detector are moved during scanning across the Co K-edge. Details about the laboratory-based X-ray absorption spectrometer can be found in the work of Honkanen *et al.*^[31] The *in-situ* XANES set-up is illustrated in Figure 1.^[30]

i. Validation of the Laboratory-Based XANES Set-Up

The first set of experiments completed with the *in-situ* XANES laboratory-scale set-up were the measurements of the different cobalt reference materials. These data were then used to

identify the different cobalt species present during the different stages of reaction, and to elucidate the possible formation of CoO and CoTiO₃ during FTS. The second stage in our study was to test the Co/TiO₂ FTS catalyst under reaction conditions for 35 h at 523 K, 1 bar, and a H₂/CO ratio of 0.5. In order to validate the research accomplished by the *in-situ* XANES laboratory-scale set-up, analogous experiments were performed at the BM26 beamline at the European Synchrotron Radiation Facility (ESRF), wherein the different cobalt references were measured and the Co/TiO₂ catalyst was tested under FTS reaction conditions for 15 h at 523 K, 1 bar, and a H₂/CO ratio of 0.5. The *ex-situ* XANES measurements of the cobalt references performed using the laboratory-based set-up and at the BM26 beamline of ESRF are presented in Figure 2. The analysis of these *ex-situ* spectra shows that the laboratory-based set-up produces high quality data comparable to synchrotron measurements, although at reduced time-resolution (~0.5 h for the synchrotron-based measurements and ~3 h for the laboratory-based measurements). This is illustrated in Figure S1. The spectra collected during FTS at different times of reaction using the laboratory-based set-up, which shows a lower signal-to-noise ratio because of the relative small amount of cobalt in the sample and the high absorbance rate of the support (TiO₂), and at the BM26 beamline are presented in Figure 3. By analyzing the XANES data from the laboratory-based set-up, it was concluded that the catalyst consisted of a mixture of metallic Co (active phase for the FTS reaction), and CoTiO₃ (inactive phase for the FTS reaction), formed during the activation step of the catalyst at 673 K and 1 bar, in a pure H₂ flow, this inactive phase was formed because of the presence of hot-spot in the bed of the capillary reactor. By least squares linear combination fitting of the cobalt reference spectra to the unknown spectra collected during FTS reaction (Figure 4.I), it was established that ~33% of the cobalt was converted into CoTiO₃. After 35 h of FTS reaction, a small fraction of CoO was identified by the fit (Figure S2.1.4.a). However, this fraction was found within measurement uncertainty, i.e. at/below the detection limit and therefore does not provide conclusive evidence of a re-oxidation process during FTS reaction.

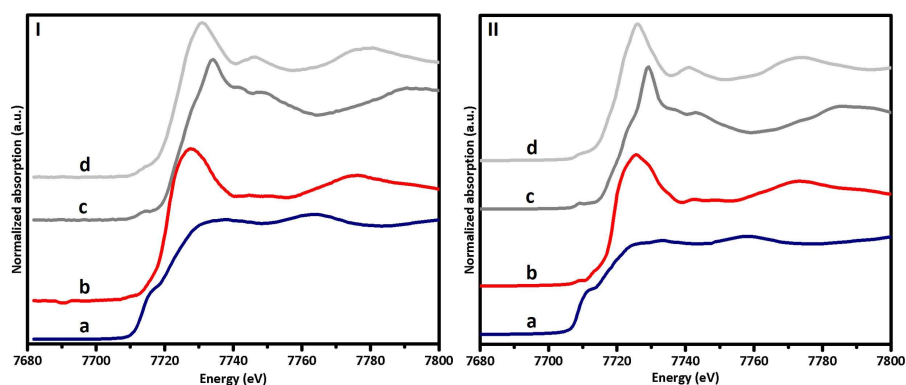


Figure 2. XANES of the cobalt reference materials measured: I) using the laboratory-based set-up; a) Co, b) CoTiO₃, c) Co₃O₄ and d) CoO, and II) at the BM26 beamline; a) Co, b) CoTiO₃, c) Co₃O₄ and d) CoO.

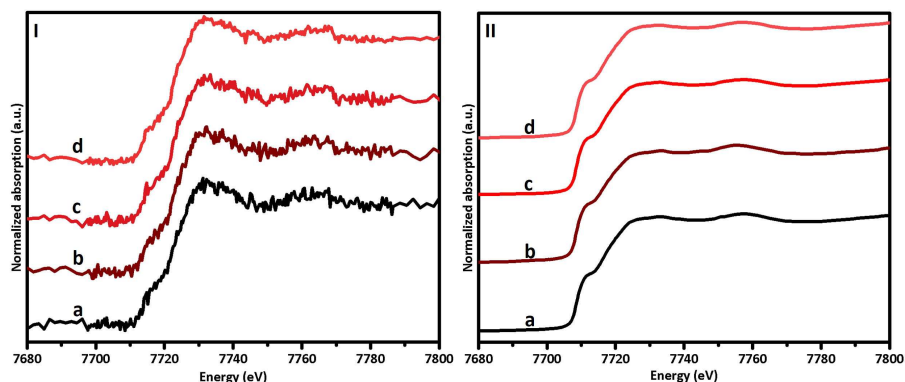


Figure 3. *In-situ* XANES spectra of the FTS reaction over a Co/TiO₂ catalyst at a reaction temperature of 523 K, a reaction pressure of 1 bar and a H₂/CO ratio of 0.5 at different reaction times: I) measured with a laboratory-based set-up at a) 8.75 h, b) 17.5 h, c) 26.25 h and d) 35 h, and II) measured at the BM26 beamline of ESRF at a) 6 h, b) 10 h, c) 12 h and d) 15 h. A glitch due to the monochromator is visible at 7764 eV in some of the spectra.

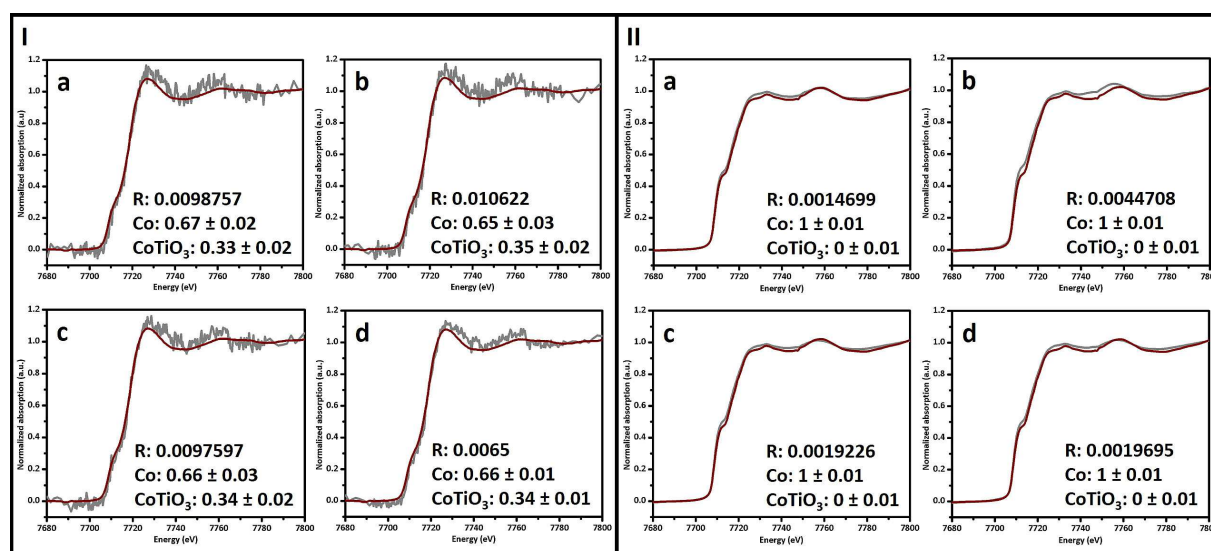


Figure 4. Least Squares Linear Combination fitting of the data recorded during the experiments performed: I) using the *in-situ* XANES laboratory-scale set-up at different reaction times; a) 8.75 h, b) 17.5 h, c) 26.25 h and d) 35 h, and II) at the BM26 beamline of the ESRF after a) 6 h, b) 10 h, c) 12 h and d) 15 h of FTS. Experimental data is plotted in grey, fits are plotted in red.

In the case of the synchrotron measurements, a similar procedure was applied (Figure 4.II) and no substantial amount of CoTiO₃ was identified by the fit in the different steps of both reactions, as illustrated in Figure 6.

ii. Long-Term FTS Reaction Experiments

A second set of *in-situ* experiments was completed with the XANES laboratory-scale set-up using FTS reaction conditions of 523 K and 5 bar, using a H₂/CO ratio of 0.5 and 2. These experiments were performed for 200 h each. The different spectra collected at different times for both long-term reactions are presented in Figure 5.

By least squares linear combination analysis of the long-term XANES data using the same procedure as presented in the

previous section, it was established that a small fraction of CoTiO₃ was identified by the fit in the different steps of both reactions, as illustrated in Figure 6.

For all the spectra measured the combination of the reference spectra of Co and CoTiO₃ gives the best fitting, except for one case (see section S2.2.12). However, as in the case of the data from the BM26 beamline, CoTiO₃ and CoO fractions were found within measurement uncertainty and does not provide conclusive evidence about their formation during the activation step and the FTS. The complete fitting data is presented in section S2.2.

In addition, no re-oxidation of the cobalt metal nanoparticles was observed in any of the FTS experiments performed at different pressures (1 and 5 bar), and H₂/CO ratios (0.5 and 2), which agrees with previous studies from our research group that indicate that re-oxidation is not observed during the activation or deactivation stages of a Co/TiO₂ FTS catalyst.^[35–37]

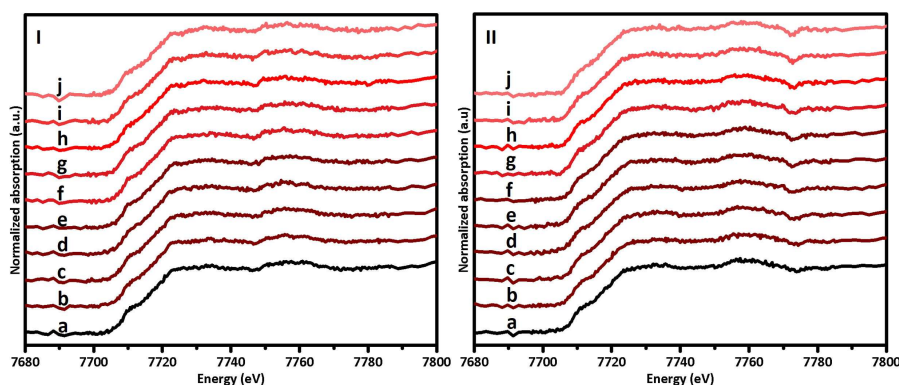


Figure 5. *In-situ* XANES spectra of the FTS reaction over a Co/TiO₂ catalyst at a reaction temperature of 523 K, and a reaction pressure of 5 bar: I) using a H₂/CO ratio of 0.5 at different reaction times; a) 20 h, b) 40 h, c) 60 h, d) 80 h, e) 100 h, f) 120 h, g) 140 h, h) 160 h, i) 180 h and j) 200 h; and II) using a H₂/CO ratio of 2 at different reaction times; a) 20 h, b) 40 h, c) 60 h, d) 80 h, e) 100 h, f) 120 h, g) 140 h, h) 160 h, i) 180 h and j) 200 h.

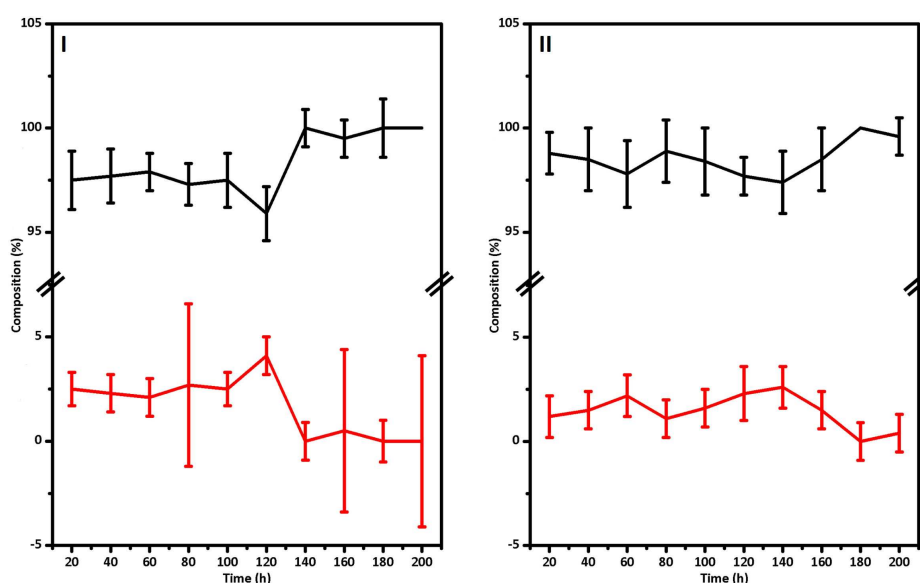


Figure 6. Composition of the different cobalt species obtained from Least Squares Linear Combination fitting of the data recorded during the experiments perform: I) using a H₂/CO ratio of 0.5 at different reaction times, and II) using a H₂/CO ratio of 2 at different reaction time, fitting values for Co in black and for CoTiO₃ in red.

Conclusions

A laboratory-scale set-up is presented that allows *in-situ* XANES measurements of a solid catalyst. This instrument provides the opportunity to carry out unique experiments, which are complementary to the research performed at synchrotron radiation facilities. Thanks to the stability of the set-up, as demonstrated in Figure S3, this technical development shows large potential for long-term experiments, e.g. when activation and/or deactivation studies of complex reaction systems have to be conducted over time frames of days or weeks. Such experiments are difficult to perform at synchrotrons due to the limited time available during granted beamtimes. By using this *in-situ* laboratory-scale XANES set-up it was not only possible to establish Co K-edge XANES of different Co reference compounds, namely Co₃O₄, CoO, Co and CoTiO₃, but also to follow

the activation and expected long-term deactivation^[15] of a Co/TiO₂ catalyst during realistic Fischer-Tropsch Synthesis (FTS) conditions; i.e., at elevated pressure and temperature and while flowing hydrogen and carbon monoxide. From the performed laboratory-scale *in-situ* FTS experiments, it is observed that at a temperature higher than 573 K, a small fraction of the cobalt can be converted into CoTiO₃ during the reduction process, which agrees with the concept of strong metal-support interaction.^[16] Furthermore, during FTS at different conditions (pressure and H₂/CO ratio) no cobalt re-oxidation was observed, which is in line with earlier reports^[35–37] from our research group, confirming that oxidation of the cobalt is not a main deactivation route for FTS, leaving sintering and coke formation as main reasons for the detrimental activity over time for Co-based FTS catalyst.^[15]

Experimental Section

The Co/TiO₂ catalyst under study was synthesized by incipient wetness impregnation (IWI) using TiO₂ P90 as support (Evonik, 90% anatase and 10% rutile) and Co(NO₃)₂·xH₂O (Acros Organics, 99+%) as precursor. The cobalt loading obtained was 14.1 wt%, which was confirmed by Inductively Coupled Plasma-Atomic Emission Spectrometry (ICP-AES). The calcined catalyst consists of Co₃O₄ nano-crystallites of ~14 nm in diameter. The crystal size was estimated by X-Ray Diffraction (XRD) data analysis and confirmed by Scanning Transmission Electron Microscopy-Energy Dispersion X-ray Spectroscopy (STEM-EDX). These data can be found in Figures S4 and S5. The *in-situ* reactor set-up, used in the experiments, is made of a quartz capillary plug-flow reactor with an inner diameter of 0.98 mm and outer diameter of 1 mm.^[38,39] The catalyst was placed and fixed in the isothermal zone (Figures S6.1 and S6.2) of the capillary by quartz wool and the plug-flow reactor was heated by two IR heaters. The temperature was controlled by a thermocouple positioned in the bed of the capillary. During the experiments, two different measurements were performed: a) *ex-situ* measurements of cobalt references; Co₃O₄ (Sigma-Aldrich, 99.5%), CoO (Acros Organics, 99+%), CoTiO₃ (Alfa Aesar, 99.8%) and Co (Co foil, 99.9%); the powders were diluted with potato starch to increase the transmittance of the sample, and b) *in-situ* measurements by using the FTS reaction on Co/TiO₂ (14.1 wt%) catalyst at 523 K, 1 and 5 bar. The FTS reaction was divided into two different stages: i) activation (reduction of the Co₃O₄ nano-crystallites to Co⁰) at 673 K and 1 bar, in a H₂ flow of 1 ml/min (AGA, H₂ > 99.999%) and ii) the FTS reaction, feeding a mixture of H₂/CO (AGA, CO > 99.97%) of 0.5 (0.3 ml/min of H₂ and 0.6 ml/min of CO) and 2 (0.6 ml/min of H₂ and 0.3 ml/min of CO). Before starting the activation step, a Co K-edge spectrum was collected to identify the oxidation state of the cobalt in the sample. The different spectra obtained during the execution of the experiments were collected continuously. The spectra were collected every ~38 min and were summed up every 8.75 h (~14 scans per spectrum) for the experiment performed at 1 bar, and every 20 h (~32 scans per spectrum) for the experiments at 5 bar, which thereby determined the effective time resolution. During the activation stage, the temperature was raised from room temperature (RT) to 673 K with a heating rate of 5 K/min, and kept at 673 K for 2 h. Once the activation stage was concluded, the temperature was decreased to 523 K with a rate of 5 K/min and the gases were switched to a mixture of H₂/CO (0.5) in the case of the experiment at 1 bar. The FTS reaction was performed for 35 h. For the experiments performed at 5 bar, after the activation step was finished the temperature was reduced to RT, the system was pressurized in a mixture of H₂/CO (0.5 and 2 in each case), and the temperature was raised to 523 K with a heating rate of 5 K/min. The FTS reactions were performed for 200 h. A new set of experiments was performed at the Dutch-Belgian Beamline BM26 (DUBBLE) at the ESRF following an equivalent experimental procedure. These experiments were planned to complement and compare with the data obtained from the laboratory-based set-up. In this second set of experiments the different cobalt references (i.e., Co₃O₄, CoO, CoTiO₃ and Co) were measured and the FTS reaction was performed at 1 bar and for 15 h. The Athena XAS data processing software was used for the analysis of all sets of data. The Co/TiO₂ catalyst under study has been previously tested to check its catalytic activity by using a fixed-bed reactor. The FTS products were analyzed every hour by an on-line Gas Chromatograph (GC, Varian 430).^[40]

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Conflict of Interest

The authors declare no conflict of interest.

Keywords: X-ray spectroscopy · heterogeneous catalysis · hydrogenation reaction · Fischer-Tropsch Synthesis · Cobalt

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