

FIGURE 1. (a) The effect of scCO₂ pressure on the conversion of reaction in Scheme 1 carried out with the silica-supported catalyst **2c**. Conversion data are averaged over two runs, after 5 h. (b) The density of scCO₂ at 50 °C, in the range of 75–150 bar.

Also, the use of supercritical CO₂ as a PTC-solvent is compared (on a kinetic basis) to the classical use of low polarity solvents, such as toluene and *n*-heptane.

Results and Discussion

The Pressure Effect. In the dense phase, most of the CO₂ properties are affected by its pressure. In particular the density which, at a reduced temperature (T_r) of 1.1, increases by around 80% when the reduced pressure (P_r) is augmented from 1 to 3.⁶ Accordingly, a sharp enhancement of the solubility of organic compounds in CO₂ is usually observed in the proximity and above the critical point.

Under *LSc*-PTC conditions, the effect of the CO₂ pressure was investigated, using the reaction of Scheme 1 as a model for nucleophilic displacements. Experiments were run at 50 °C in a stainless steel autoclave loaded with a mixture of compound **1**, an aqueous solution of KI, and the PT-catalyst (**1**:KI:**2** in a 1:2.5:0.05 molar ratio, respectively). Both catalysts **2a** and **2c** (Q = P; R = *n*-Bu) were used. The pressure of CO₂ was varied from 80 to 150 bar and the reaction conversion was measured by GC, after 5 h.

In the case of the supported catalyst **2c**, results are reported in Figure 1a (squares), and they are compared with the trend of the CO₂ density in the same pressure interval, at 50 °C (Figure 1b).⁷

As the pressure is enhanced from 80 to 150 bar, the conversion of **1** into *n*-octyl iodide (**4**) drops from 73% to 15%, with an overall decrease by a factor of 5. The CO₂ density instead increases about four times in the same pressure range.

A good parallel for this behavior can be found with the silica-promoted Diels–Alder cycloadditions carried out in scCO₂ by Weinstein et al.⁸ Also for these reactions, the

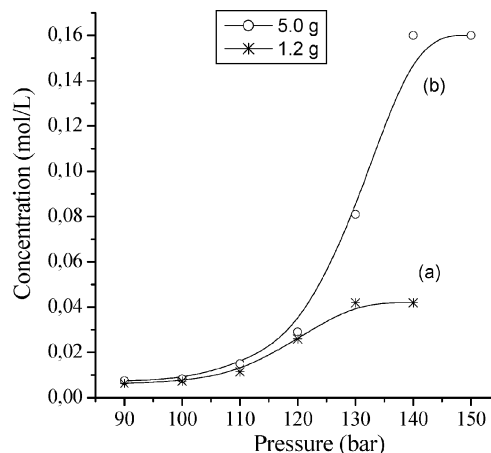


FIGURE 2. Molar concentration of *n*-octyl mesylate (○, 5.0 g; *, 1.2 g) in scCO₂ at different pressures and at 60 °C.

authors observed that the CO₂ pressure had a negative effect on the reaction rate, and they concluded that the result was due to a decreased adsorption of the reagents over the silica surface. As the pressure and hence the density of scCO₂ were augmented, the solubility of reagents in the supercritical phase was enhanced as well, and the rate dropped since reagents were removed from the promoter.

Likewise, in the low-pressure range of Figure 1a, it may be assumed that compound **1** is only partially solubilized in scCO₂ so that the high conversion is the result of a favorable partitioning of the reagent on the silica-supported catalyst. Meanwhile, at a higher pressure **1** desorbs from the support and the reaction is slowed.

An indirect proof for this reasoning comes from the reaction carried out in the presence of the conventional onium salt **2a** (Figure 1a, stars). In the interval pressure of 90–150 bar, the conversion shows only a slight decrease with an overall variation of ~20%. Although the reaction still takes place through the partitioning of **1** between the scCO₂ and the onium salt,⁹ the partitioning process does not involve the intermediacy of the silica matrix, so that the pressure effect is more limited. This may account for the different performance of catalyst **2a** with respect to **2c**, especially in the high-pressure region.

In light of these results, the measure of the concentration of **1** in scCO₂ at different pressures could give definite support for the interpretation of Figure 1a.

Concentration Measures. *n*-Octyl mesylate (both 1.2 and 5.0 g were used) was loaded in a stainless steel autoclave (150 mL) and pressurized with CO₂ in the range of 80–170 bar, at 60 °C (details are in the Experimental Section).

Figure 2 reports the results. The molar concentration of **1** in scCO₂ is plotted vs the pressure of the supercritical phase. In the pressure region up to 100 bar, the concentration of *n*-octyl mesylate in scCO₂ does not exceed 7 mmol/L.

(8) Weinstein, R. D.; Renslo, A. R.; Danheiser, R. L.; Tester, J. W. *J. Phys. Chem.* **1999**, *103*, 2878–2887.

(9) (a) Dillow, A. K.; Jimmy Yun, S. L.; Suleiman, D.; Boatright, D. L.; Liotta, C. L.; Eckert, C. A. *Ind. Eng. Chem. Res.* **1996**, *35*, 1801. (b) Chandler, K.; Culp, C. W.; Lamb, D. R.; Liotta, C. L.; Eckert, C. A. *Ind. Eng. Chem. Res.* **1998**, *37*, 3252–3259.

(6) Lucien, F. P.; Foster, N. R. In *Chemical Synthesis Using Supercritical Fluids*; Jessop, P. G., Leitner W., Eds.; Wiley-VCH: New York, 1999; Chapter 1.2, pp 39–53. $T_r = T/T_c$; T_c = critical temperature = 31.1 °C; $P_r = P/P_c$; P_c = critical pressure = 73.8 bar.

(7) <http://webbook.nist.gov>.

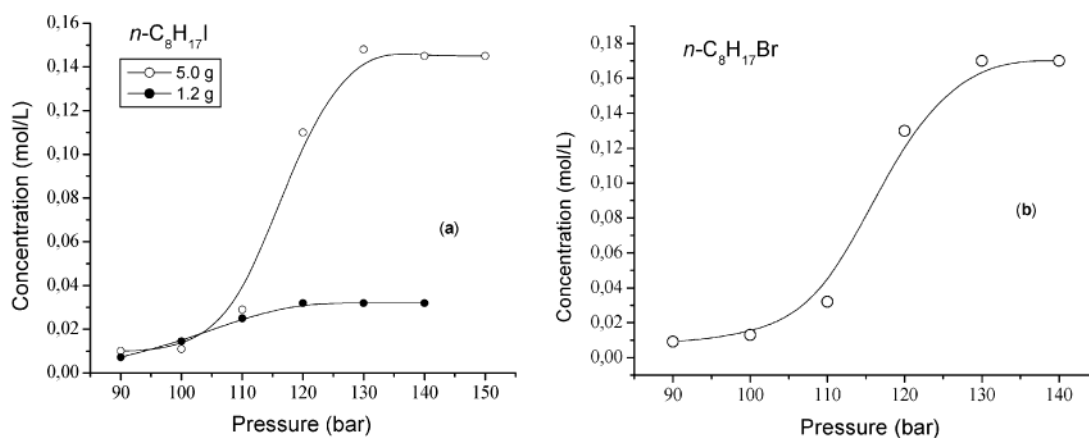


FIGURE 3. Pressure versus concentration plots at $T = 60\text{ }^{\circ}\text{C}$: (a) concentration of *n*-octyl iodide (○, 5.0 g; ●, 1.2 g) in scCO_2 at different pressures; (b) concentration of *n*-octyl bromide (○, 5.0 g) in scCO_2 at different pressures.

As the CO_2 pressure is increased over 110 bar, the concentration curve a (1.2 g of substrate, stars) has a steep enhancement and between 130 and 140 bar it reaches a *plateau* of 4.1×10^{-2} mol/L. A similar behavior is observed when 5.0 g of *n*-octyl mesylate are used (open circles, curve b). The concentration curves a and b show a very good overlap until pressures of 120 bar. Then, curve b goes up, and at 140 bar it reaches a maximum of about 0.16 mol/L. The *plateaus* of both curves match exactly the values expected for the complete solubilization of **1**.

At $60\text{ }^{\circ}\text{C}$, the CO_2 pressure has a threshold value of 130–140 bar below which the concentration of the substrate in scCO_2 rapidly falls (to only 4% of the initial 5.0 g, for pressure drops of 40 bar; curve b).

Importantly, these concentration trends validate the initial hypothesis: as Figure 2 is a good mirror image of Figure 1a (cat. **2c**), a clear dependence of the reaction conversion on the partitioning of the reagent between scCO_2 and the silica-supported catalyst can be inferred.

The effect of the CO_2 pressure was examined also for the product of the reaction ($n\text{-C}_8\text{H}_{17}\text{I}$), as well as for bromooctane. Experiments were run at $60\text{ }^{\circ}\text{C}$, as described for *n*-octyl mesylate. Results are reported in Figure 3.

The behavior of both halides in scCO_2 is similar to that of **1**. However, it should be noted that the maxima of curves 3a ($n\text{-C}_8\text{H}_{17}\text{I}$) and 3b ($n\text{-C}_8\text{H}_{17}\text{Br}$) appear slightly shifted to a lower pressure (120–130 bar) with respect to the sulfonate ester (130–140 bar, Figure 2). The higher polarity of this latter compound may account for the difference.

The results of Figures 2–3 not only offer a good explanation for the effect of the CO_2 pressure on the reaction conversion, but even more importantly, they indicate that under *LSc*-PTC conditions, the supercritical CO_2 phase acts as a solvent only at a density above 0.5 g/mL ($P \geq 130$ bar at $60\text{ }^{\circ}\text{C}$).

The Kinetic Investigation. Under *LSc*-PTC conditions, a preliminary kinetic investigation of nucleophilic displacements on **1** was carried out with catalyst **2a** ($n\text{-Bu}_4\text{P}^+\text{Br}^-$), which was more efficient and less sensitive to pressure effects than the silica-supported onium salt. Compound **2a** was immiscible with scCO_2 .¹⁰

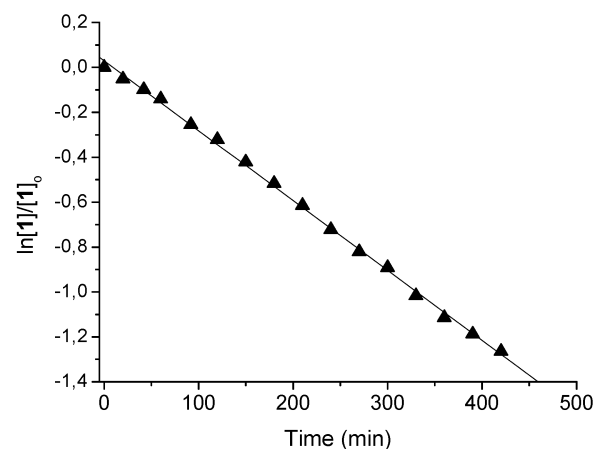


FIGURE 4. The reaction of *n*-octyl mesylate with KI: first-order rate constant plot for the reagent. $[\mathbf{1a}]_0$ and $[\mathbf{1a}]$: concentration of **1** at $t = 0$ and at a later time t .

Initial experiments were run at $60\text{ }^{\circ}\text{C}$, in a stainless steel autoclave (90 mL) loaded with a mixture of compound **1**, an aqueous solution of KI (or KBr), and the catalyst **2a** [**1**:KI (or KBr):**2a** in a 1:2.5:0.05 molar ratio, respectively]. A pressure of 160 bar was held throughout the reaction ($d_{\text{CO}_2} = 0.64$ g/mL), which was followed by GC and GC/MS.

In the case of the reaction with aq KI, results are reported in Figure 4.

During both reactions with KI and KBr, the concentration (determined by GC) of the sulfonate ester fits the expression $\ln[\mathbf{1a}]/[\mathbf{1a}]_0 = k_t([\mathbf{1a}]_0 \text{ and } [\mathbf{1a}])$. This behavior, in agreement with the general mechanism proposed by Starks for *ll*-PTC processes,^{1,11} allows the calculation of pseudo-first-order rate constants k_{obs} of $3.11 \times 10^{-3} \text{ min}^{-1}$ (KI, slope of Figure 4) and $2.05 \times 10^{-3} \text{ min}^{-1}$ (KBr) for the disappearance of the substrate.¹²

A second set of experiments was then carried out with the same procedure at $70\text{ }^{\circ}\text{C}$ and under 180 bar of CO_2

(10) Onium salts are virtually insoluble in scCO_2 ; ref 9a reported a solubility of 0.21 mmol/L for tetraheptylammonium bromide in scCO_2 (204 bar, $50\text{ }^{\circ}\text{C}$) in the presence of acetone (5%) as a cosolvent.

(11) (a) Starks, C. M. *J. Am. Chem. Soc.* **1971**, *93*, 195. (b) Starks, C. M.; Owen, R. M. *J. Am. Chem. Soc.* **1973**, *95*, 3613.

(12) Since an excess of the reactant salt is used, a pseudo-first-order rate constant is considered.

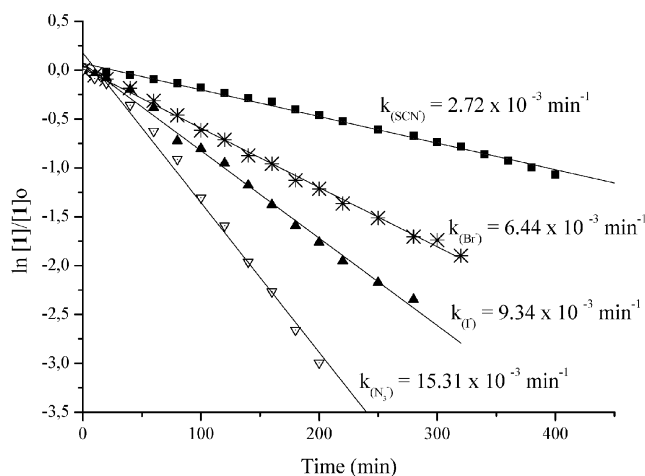


FIGURE 5. k_{obs} for the reaction of *n*-octyl mesylate (**1**) with aq KX [X = SCN (■), Br (*), I (▲), and N₃ (▼)] at 70 °C and 180 bar.

($d_{\text{CO}_2} = 0.61$ g/mL): compound **1** was made to react with aqueous solutions (4.8 M, 7.5 mL) of four different potassium salts KX (X = I, Br, N₃, and SCN) and the catalyst **2a** [**1**:KX:**2a** in a 1:2.5:0.05 molar ratio, respectively].¹³

As for the previous tests, linear logarithmic plots of the concentration of the substrate with time [$\ln [\mathbf{1a}]/[\mathbf{1a}]_0 = k_{\text{obs}}t$] were obtained for the reaction of **1** with all the examined anions. Results are reported in Figure 5, where k_{obs} (slope) is shown for each anion.

The overall scale of relative nucleophilicity of anions is in the order of Scheme 2,¹⁴ and it is in agreement with the one reported by Montanari et al.¹⁵ for the reaction of *n*-octyl mesylate with the same nucleophiles carried out in a biphasic water/chlorobenzene system and in the presence of C₁₆H₃₃PBu₃⁺Y⁻ (Y = N₃⁻, I⁻, Br⁻, and SCN⁻) as a PT-catalyst. It was demonstrated that the reactivity of anions was due to their hydration shells whose size was in turn related to the size of the nucleophile for the sequence of halide ions (I⁻ and Br⁻), and to the charge dispersion for the N₃⁻ and SCN⁻ ions.^{15b} Moreover, the authors claimed that the highest reaction rates were achieved in the presence of apolar solvents where the anion solvation was the poorest.

Although these considerations could fit also our *LSc*-system, a comparison between scCO₂ and conventional nonpolar media was necessary to understand the effect of the supercritical CO₂ phase as a PT-solvent. Solutions of *n*-octylmethane sulfonate (0.17 M, 82.5 mL) in either *n*-heptane or toluene were made to react at 70 °C, with aqueous solutions (4.8 M, 7.5 mL) of KX salts (X = I, Br, N₃, and SCN) in the presence of the catalyst **2a** [**1**:KX:

(13) Under *LSc*-PTC conditions, some anions (i.e. SCN⁻) appeared quite unreactive. A higher temperature (70 °C) was then necessary. Accordingly, the CO₂ pressure was increased (180 bar) to achieve a medium density comparable to that of previous experiments at 60 °C (see Figure 4).

(14) The solubility (if any) of the catalyst **2a** in scCO₂ was not known (ref 10), neither was it determined by us. However, since all reactions of Figure 5 were run at 70 °C and 180 bar, and most of all, using aqueous solutions of identical anion concentrations, corrections of k_{obs} to account for the catalyst concentration may not be influential.

(15) (a) Landini, D.; Maia, A. M.; Montanari, F.; Pirisi, M. *J. Chem. Soc., Chem. Commun.* **1975**, 950. (b) Landini, D.; Maia, A. M.; Montanari, F. *J. Am. Chem. Soc.* **1978**, *100*, 2796.

SCHEME 2. The Scale of Reactivity for Some Anions under *LSc*-PTC Conditions

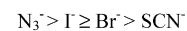


TABLE 1. Pseudo-First-Order k_{obs} for the Disappearance of *n*-Octylmethane Sulfonate in the Reaction with aq KX Salts at 70 °C, under Both *II*- and *LSc*-Conditions

entry	anion	$k_{\text{obs}} \times 10^{-3}$ (min ⁻¹)		
		<i>n</i> -heptane	toluene	scCO ₂
1	N ₃ ⁻	14.36	24.53	15.31
2	I ⁻	8.90	6.13	9.34
3	Br ⁻	5.81	6.28	6.44
4	SCN ⁻	1.90	1.29	2.72

2a in a 1:2.5:0.05 molar ratio, respectively]. In all cases, k_{obs} of pseudo-first-order were measured for the disappearance of the substrate.

Results are reported in Table 1.

Although a direct kinetic comparison of the solvent effect is not possible,¹⁶ it should be noted that the reactivity scale underlined in Scheme 2 has a general validity for each of three solvents investigated.

Under biphasic water/supercritical conditions, CO₂ appeared a good PT-solvent: within limits, its behavior resembled that of apolar C7–C8 aliphatic or aromatic hydrocarbons.

However, a relevant point still needed to be clarified.

Under *II*-PTC conditions, a well-known concept is that the efficacy of onium salts as PT-catalysts largely depends on their organophilicity.^{1,17} Lipophilic salts have small but not negligible extraction constants (E_{QX} ¹⁸) in biphasic water/apolar solvent systems: for instance, in a water/chlorobenzene biphasic system, an E_{QX} of 0.1 was reported for Bu₄N⁺Br⁻.¹⁹

By contrast, although onium salts were substantially insoluble in scCO₂, *LSc*-reactions yielded results similar to those of *II*-processes carried out in toluene or *n*-heptane.

To understand how the onium salt promotes the reaction, the following experiment was carried out. An autoclave (150 mL) was loaded with an aqueous solution of the catalyst **2a** (9.6×10^{-3} M, 25 mL), heated at 70 °C under stirring, and finally pressurized with CO₂ at 180 bar. After 30 min, stirring was stopped and the mixture was allowed to settle. Known volumes of both the

(16) Second-order rate constants expressed as $k_{\text{obs}}/[\text{Bu}_4\text{P}^+\text{Br}^-]_{\text{org}}$ should be considered to account for the actual concentration of the PT-catalyst in each of three solvents. However, no data were available for $[\text{Bu}_4\text{P}^+\text{Br}^-]_{\text{scCO}_2}$ (see also ref 10 and footnote 14).

(17) (a) Dehmlöv, E. V.; Dehmlöv, S. S. In *Phase Transfer Catalysis*, 3rd ed.; VCH Publishers: New York, 1993. (b) *Phase Transfer Catalysis*; Halpern, M. E., Ed.; ACS Symp. Series no. 659; American Chemical Society: Washington, DC, 1997.

(18) Bar, R.; Karpuj-Bar, L.; Sasson, Y.; Blum, J. *Anal. Chim. Acta* **1983**, *154*, 203. E_{QX} is the ratio of concentration of the onium salt between the organic and the aqueous phases $\{E_{\text{QX}} = [\text{Q}^+\text{X}^-]_{\text{org}}/[\text{Q}^+\text{X}^-]_{\text{aq}}\}$.

(19) Montanari, F.; Landini, D.; Rolla, F. *Top. Curr. Chem.* **1982**, *101*, 147. Also, ref 13b reported that in a water/chlorobenzene system (1:0.75 v/v), the reaction of *n*-octylmethane sulfonate with aq KBr yielded $k_{\text{obs}} = 10.4 \times 10^5$ and 0.24×10^5 s⁻¹ in the presence of Bu₄P⁺Br⁻ (**2a**) and Pr₄N⁺Br⁻, respectively. Although a salting-out effect took place, the percent of the two catalysts in the organic phase was 97 and 2.5% (Akabori, S.; Ohtomi, M. *Bull. Chem. Soc. Jpn.* **1975**, *48*, 2991).

TABLE 2. Partitioning of Onium Salts in scCO₂/Water

entry	PT-catalyst ^a	reagent 1 ^b	aq phase (25 mL)		% of Br ⁻ ion in ^d	
			alone	added KX ^c (mol)	aq phase	sc phase
1	2a	no	yes	none	100	—
2		yes	yes	none	100	—
3a		no	yes	X = CO ₃ ²⁻	100	—
3b		yes	yes	(0.12)	82	—
4a		no	yes	X = NO ₃ ⁻	100	—
4b	yes	yes	(0.12)	100	—	
5	2d	yes	none	none	ND ^e	—
6a		no	yes	X = CO ₃ ²⁻	81	—
6b		yes	yes	(0.12)	74	—
7a		no	yes	X = NO ₃ ⁻	79	—
7b		yes	yes	(0.12)	79	—

^a PT-catalysts (Bu₄P⁺ Br⁻ and C₁₆H₃₃PEt₃⁺ Br⁻) were used in the same molar amount (0.24 mmol). ^b *n*-Octyl mesylate was used for all experiments. ^c KX (X = CO₃²⁻, NO₃⁻) added to the aqueous phase. ^d Determined by titration with aq AgNO₃ (5 × 10⁻³ M). ^e Not determined.

supercritical and the aqueous phase were withdrawn through calibrated loops and titrated for bromide with aq AgNO₃ (see the Experimental Section). CO₂ was then partially vented to 130 bar. *n*-Octyl mesylate (3.0 g, 1.4 × 10⁻² mol) was rapidly pumped into the autoclave with a stream of CO₂ until the original pressure of 180 bar was reached. A second sample of the aqueous phase was collected and titrated, after stirring.

The same experiment was then repeated by loading initially the autoclave with (i) an aqueous solution of **2a** (9.6 × 10⁻³ M, 25 mL) containing either K₂CO₃ or KNO₃ (4.8 M) and (ii) the solid onium salt (**2a**, 0.24 mmol).

Additionally, also the behavior of *n*-C₁₆H₃₃PEt₃⁺ Br⁻ (**2d**) was considered. Results are reported in Table 2.

It should be first noted that although Bu₄P⁺ Br⁻ is soluble in *n*-octyl mesylate, and this in turn is soluble in scCO₂, not even traces of the salt are detected in supercritical phase, either in the presence or in the absence of water (entries 1, 2, and 5, respectively).²⁰ At 180 bar, the sulfonate ester is totally dissolved in scCO₂ and clearly it does not act as a cosolvent for the PT-catalyst.

The addition of K₂CO₃ or KNO₃ to the aqueous phase does not modify the partitioning (entries 3a and 4a).

But if the reagent **1** and aqueous K₂CO₃ are added together, a salting-out effect is manifest (entry 3b).¹⁷ Though the PT-catalysts still does not dissolve in scCO₂, the phase transfer agent **2a**, which is mostly partitioned in water, can be accounted for by imagining that it forms a third liquid phase between the aqueous and the supercritical solvents.²¹

A similar situation reasonably holds during the reaction run in the presence of aq KBr (or KI or KSCN). Under such conditions, however, a large excess of KX (X

= Br, I, SCN; all anions detectable by potentiometric titration) is present so that the direct titration of the bromide of the onium salt is not possible.

At room temperature and atmospheric pressure, compound **2d** is miscible also with water and with the substrate **1**. Table 2 shows that in the supercritical-aqueous biphasic system, at 180 bar, a salt effect is possible with both aq K₂CO₃ and KNO₃ (entries 6a and 7), and it is slightly increased by the addition of the reagent **1** (entry 6b).²² Though the catalyst **2d** is not soluble in CO₂ and as for compound **2a**, it may partly distribute as an immiscible third liquid phase.

Conclusions

This study indicates that nucleophilic displacement (S_N) reactions under *LSc*-PTC conditions are controlled mainly by substrate concentration in the supercritical phase, which in turn is governed by the CO₂ density at the operative pressure. In particular, we have defined a protocol for S_N reactions under *LSc*-PTC conditions on the model substrate *n*-octylmesylate whereby CO₂ pressure must be greater than 140 bar at 60–70 °C, for a CO₂ density of 0.55–0.65 g/mL.

Kinetic data indicate that supercritical CO₂ is a good solvent for *LSc*-PTC and that its behavior, within limits, resembles that of *n*-heptane and toluene.

The scCO₂ does not dissolve any measurable amount of phase-transfer catalysts. However, in the presence of *n*-octylmesylate and of a saline aqueous solution, a salt effect is observed, whereby some of the phase-transfer catalyst is not accounted for, either in the supercritical or aqueous phases. At this stage, there is no clear rationale for this behavior. Likely, with respect to a scCO₂/water system, different interfacial phenomena occur in the biphasic system formed by water and the binary mixture of scCO₂-octylmesylate, which may either favor the formation of small droplets of PT-catalyst containing water or perhaps force part of the phosphonium salt to form a separate third liquid phase wherein the reaction takes place.

Future perspectives will be on the synthesis and use of onium salts with CO₂-philic perfluorinated fragments. Moreover, the pH-buffering effect of CO₂ will also be investigated to tune the anion nucleophilicity.

Experimental Section

SAFETY WARNING: Operators of high-pressure equipment should take proper precautions to minimize the risks of personal injury.²³

The silica-supported PT-agent **2c** was synthesized and analyzed as previously described;²⁴ the loading of the onium salts was of 0.30 mequiv per g of silica. *n*-Octyl mesylate (**1**) was also prepared according to established procedures.²⁵ Before each reaction, compound **2c** was dried by heating under vacuum at 60 °C, overnight.

The *LSc*-reaction and the measures of concentration of **1** and octyl halides (Figures 2–3) in scCO₂ were performed in

(20) At atmospheric pressure and room temperature, the substrate **1** was able to dissolve an equimolar amount of the salt **2a**.

(21) The autoclave was equipped with sapphire windows for the visual observation of the mixture: also from this inspection, when the onium salt was used as a solid (entry 5, Table 2), it appeared totally immiscible in scCO₂. An analogous distribution was already reported by us when onium salts were in the presence of an apolar solvent and an aqueous phase. (a) Marques, C. A.; Selva, M.; Tundo, P. *J. Org. Chem.* **1994**, *59*, 3830. (b) Marques, C. A.; Selva, M.; Tundo, P. *J. Org. Chem.* **1995**, *60*, 2430. (c) Selva, M.; Perosa, A.; Tundo, P. *J. Org. Chem.* **1999**, *64*, 3934.

(22) Presently, there are no clear reasons for this unusual salting-out effect. Perhaps the presence of the sulfonate ester modifies the density/polarity of the sc-phase (and hence the nature of the interphase itself) so that onium salts may more easily be salted out.

(23) Jessop, P. G.; Ikariya, T.; Noyori, R. *J. Am. Chem. Soc.* **1996**, *118*, 345.

(24) (a) Tundo, P.; Venturello, P. *J. Am. Chem. Soc.* **1979**, *101*, 6606. (b) Tundo, P.; Venturello, P. *J. Am. Chem. Soc.* **1981**, *103*, 856.

(25) Sekera, V. C.; Marvel, C. S. *J. Am. Chem. Soc.* **1933**, *55*, 345.

stainless steel autoclaves of different volumes (25, 90, and 150 mL), all equipped with sapphire windows for the visual observation of the mixture, a manometer, and a thermocouple for the control of temperature. The 25-mL cell was electrically heated, while 90- and 150-mL autoclaves were thermostated by water-circulating jackets. The pressurization of all autoclaves with CO₂ was performed with an automatic syringe pump (ISCO model 260 D).

Pressure Effect: Figure 1. A mixture of **1** (200 mg, 0.96 mmol), the onium salt (Bu₄P⁺ Br⁻) or silica **2c** (both **2a** or **2c** were in 0.05 molar equiv with respect to **1**), and an aqueous solution of KI (0.5 mL, 4.8 M) was charged into a 25-mL hastelloy cell, which was pressurized with CO₂ (SFC/SFE grade) at approximately 30 bar. The autoclave was heated at the desired temperature (50 °C), while the reaction mixture was magnetically stirred. The final pressure ranging from 75 to 150 bar was reached by slowly adding the remaining CO₂ to the reactor. The reaction was allowed to proceed for 5 h; then, after cooling, CO₂ was slowly vented by bubbling it into acetone (5 mL). The content of the cell was washed with additional acetone (5 mL), and the combined acetone solutions were analyzed by GC/MS fitted with a 30 m HP5MS chromatographic column.

The Concentration of *n*-Octyl Mesylate (1**) and *n*-Octyl Halides in scCO₂: Figures 2 and 3.** In a typical experiment, *n*-octyl mesylate (**1**, 1.2 g, 5.77 mmol) was charged in a 150-mL autoclave, which was heated at 60 °C and brought to the desired pressure of CO₂ in the range of 80–150 bar. The mixture was kept under stirring at 750 rpm. After 30 min, the agitation was stopped, and an aliquot of the mixture was withdrawn through a stainless steel calibrated loop of 1.12 mL fitted to the autoclave head. The loop was vented into acetone (5 mL) and washed thoroughly with additional acetone (5 mL). The combined acetone solutions were rotary-evaporated and the amount of the liquid residue of **1** was determined gravimetrically. The measure was repeated twice for each of the chosen pressures.

Additionally, the same experiment was repeated with a higher quantity of **1** (5 g, 0.024 mmol).

The above-described procedure also was used to measure the concentration of *n*-octyl halides (C₈H₁₇X, X = Br, I), using 1.2 g of each compound.

The Kinetic Investigation: Figures 4 and 5 and Table 1. An aqueous solution of KX (X = I, Br, N₃, and SCN; 7.5 mL, 4.8 M) and Bu₄P⁺ Br⁻ (0.25 g, 0.74 mmol) was charged into an 90-mL autoclave, heated at the desired temperature (60 or 70 °C, see Figures 4 and 5), and pressurized with CO₂ at 130 bar. The mixture was kept under magnetic stirring at 600 rpm. The autoclave was equipped with a calibrated loop (1.12 mL) by which *n*-octyl mesylate (**1**, 3 g, 1.44 × 10⁻² mmol)-was pumped into the reactor with a stream of CO₂ until the final pressure of 160 bar was reached. At intervals, through the same loop, aliquots of the reaction mixture were withdrawn, vented into acetone, and finally analyzed by GC and GC/MS. At the end of the reaction, the autoclave was cooled at room temperature, and its content was treated with diethyl ether (25 mL) and washed with water (25 mL). The organic extract was dried over Na₂SO₄ and filtered. After rotary evaporation, pure products [*n*-C₈H₁₇Br (**3a**); *n*-C₈H₁₇I (**3b**); *n*-C₈H₁₇N₃ (**3c**); and *n*-C₈H₁₇SCN (**3d**)] were isolated.

The structures of compounds **3a,b** were confirmed by comparison with authentic commercial samples. The characterization of compound **3c** was carried out by ¹H NMR and IR. Spectroscopic properties of compound **3d** were in agreement with those already reported.²⁶

***n*-Octyl Azide (**3c**).** ¹H NMR (CDCl₃, 300 MHz): δ 3.27 (2H, t, *J* = 7.0 Hz, N₃CH₂CH₂-), 1.62 (2H, tt, *J* = 7.0 Hz, -CH₂CH₂CH₂N₃), 1.5–1.2 (10H, m), 0.90 (3H, t, *J* = 7.0 Hz, CH₃CH₂-). IR (KBr), cm⁻¹: 2960 (C–H asym stretch CH₃), 2920 (C–H asym stretch CH₂), 2870 (C–H sym stretch CH₃), 2850 (C–H sym stretch CH₂), 2095 (N₃).

Experiments in Toluene and *n*-Heptane: Table 1. A 200-mL, three-neck round-bottomed flask equipped with a thermostating jacket, a reflux condenser, an adapter for the withdrawal of samples, and a magnetic bar was loaded with a solution of *n*-octyl mesylate (**1**, 3 g, 1.44 × 10⁻² mmol) in toluene or *n*-heptane (90 mL). An aqueous solution of KX (X = I, Br, N₃, and SCN; 7.5 mL, 4.8 M) and Bu₄P⁺ Br⁻ (0.25 g, 0.74 mmol) was added and the mixture was kept under stirring. The flask was thermostated at 70 °C. At intervals, samples of the reaction mixture were withdrawn and analyzed by GC and GC/MS.

Partitioning of Onium Salts in scCO₂/water Biphasic: Table 2. An aqueous solution of Bu₄P⁺ Br⁻ (9.6 × 10⁻³ M, 25 mL) was charged into an 150-mL autoclave and heated at 70 °C, under stirring. The autoclave head was equipped with two calibrated loops [(a) and (b): 4.1 mL], one of which (b) was also connected to a 1/8 in. tube dipping into the aqueous phase. The mixture was pressurized with CO₂ at 180 bar. After 30 min, the agitation was stopped and samples of both the supercritical and the aqueous phases were withdrawn. The two loops were vented and thoroughly washed with a methanol/water solution [1.1 v/v, 5 mL, loop a: sc-phase] or simply water [5 mL, loop b: aq phase]. These samples were then analyzed separately by a potentiometric titrator with aq AgNO₃ (5 × 10⁻³ M).

CO₂ was vented slowly until a pressure of 130 bar was reached. By using loop a, *n*-octyl mesylate (**1**, 3 g, 1.44 × 10⁻² mmol) was charged into the autoclave with a stream of CO₂ until the pressure of 180 bar was restored. Two more samples of the supercritical and the aqueous phases were withdrawn and analyzed through the above-described procedure.

The experiment was repeated with (i) an aqueous solution of **2a** (9.6 × 10⁻³ M, 25 mL) containing either K₂CO₃ or KNO₃ (4.8 M) and (ii) the solid onium salt (**2a**, 0.24 mmol). In this latter case, only loop a was used.

Both experiments also were carried with *n*-C₁₆H₃₃PEt₃⁺ Br⁻ (**2d**) either in aqueous solution (9.6 × 10⁻³ M, 25 mL) or as a solid salt (0.24 mmol).

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