

# Proceedings of the Iowa Academy of Science

---

Volume 43 | Annual Issue

Article 35

---

1936

## Binary Systems with Acetamide

Robert Butikofer  
*Coe College*

Ben H. Peterson  
*Coe College*

Copyright ©1936 Iowa Academy of Science, Inc.

Follow this and additional works at: <https://scholarworks.uni.edu/pias>

---

### Recommended Citation

Butikofer, Robert and Peterson, Ben H. (1936) "Binary Systems with Acetamide," *Proceedings of the Iowa Academy of Science*, 43(1), 178-178.

Available at: <https://scholarworks.uni.edu/pias/vol43/iss1/35>

This Research is brought to you for free and open access by the Iowa Academy of Science at UNI ScholarWorks. It has been accepted for inclusion in Proceedings of the Iowa Academy of Science by an authorized editor of UNI ScholarWorks. For more information, please contact [scholarworks@uni.edu](mailto:scholarworks@uni.edu).

were similar to existing data on the streaming potential of such solutions through glass tubes and through cellulose diaphragms.

DEPARTMENT OF CHEMISTRY,  
STATE UNIVERSITY OF IOWA,  
IOWA CITY, IOWA.

---

## THE PREPARATION AND ANALYSIS OF MONOBROMO- OAMINE, DIBROMOAMINE, MONOCHLOROAMINE, AND NITROSYLCHLORIDE

GEORGE H. COLEMAN AND GILBERT E. GOHEEN

Monobromoamine and dibromoamine were prepared in ether solution by the reaction of bromine with ammonia in ether solution at the temperature attainable with a mixture of dry ice and acetone.

An ethereal solution of monochloroamine in concentration of one mole per liter was prepared by the action of sodium hypochlorite on ammonia.

Pure liquid nitrosyl chloride was prepared by the action of dry hydrogen chloride with nitrosyl sulfuric acid.

Apparatus and methods for the preparation, purification, and analysis of these substances were illustrated and described.

DEPARTMENT OF CHEMISTRY,  
STATE UNIVERSITY OF IOWA,  
IOWA CITY, IOWA.

---

## BINARY SYSTEMS WITH ACETAMIDE

ROBERT BUTIKOFER AND BEN H. PETERSON

The system acetamide — propionamide was investigated by the freezing point method. The solubility curve obtained showed a eutectic at approximately 0.5 mol. fraction. The experimentally determined solubility curve corresponds well with the ideal solubility calculated from the ideal equation.

DEPARTMENT OF CHEMISTRY,  
COE COLLEGE,  
CEDAR RAPIDS, IOWA.