Electronic Supplementary Information for

Catalytic intramolecular hydroamination of aminoallenes using titanium complexes of chiral, tridentate, dianionic imine-diol ligands

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Complete description of solid-state structures of ligands

Four of the five ligands were characterized by X-ray crystallography. L-H2L4, L-H2L6, L-H₂L7, and L-H₂L8 each crystallized from 4:1 ether: isopropanol to give X-ray quality crystals. The molecular structure of L-H₂L4 is shown in Figure S1. Rather than the expected iminophenol tautomer, the crystal structure clearly showed the iminium-phenoxide tautomer, with the iminium N atom protonated and the phenol O atom deprotonated. There is an intramolecular hydrogen bond between N(1) and the phenoxide oxygen, O(1). Previously we have observed intramolecular hydrogen bonds between the aliphatic oxygen and the amine nitrogen in our bidentate amino alcohols.^{66, 67} The atoms in the phenoxide aromatic ring, O(1), C(7), N(1) and H(1), are all essentially coplanar, with deviations from planarity of 0.131 Å for C(7) and 0.169Å for N1. The C(8) atom is displaced 0.544 Å, and the C(7)-N(1)-C(8)-C(9) dihedral angle is -86.7° which places the aliphatic alcohol oxygen atom 2.710 Å from the plane. There is a significant difference in the bond lengths from C(1)-O(1), at 1.2944(19) Å and C(9)-O(2), at 1.4240(18) Å, suggesting that a more accurate description of the structure is the keto-iminium ion with delocalization of charge along C(1)-C(6)-C(7)-N(1). The bonds between C(1)-C(2), C(1)-C(6), C(3)-C(4) and C(5)-C(6) are all long, at 1.410 – 1.440 Å, while the C(2)-C(3) and C(4)-C(5) bond lengths are both shorter at around 1.37 Å. In contrast, the C-C bond lengths in the aromatic ring on the benzyl side chain are typical aromatic bond distances ranging from 1.383 to 1.394 Å. There is residual electron density in the electron density map around C(6), suggesting a buildup of negative charge at that atom.



Figure S1. Thermal ellipsoid drawing of L-H₂L4 (hydrogen atoms except H(1), H(2) and H(8) are omitted for clarity; ellipsoids shown at 50% probability). Selected bond distances (Å): C(1)-

C(2) 1.418(2), C(2)-C(3) 1.373(2), C(3)-C(4) 1.410(2), C(4)-C(5) 1.368(2), C(5)-C(6) 1.416(2), C(6)-C(1) 1.440(2), C(1)-O(1) 1.2944(19), C(9)-O(2) 1.4240(18).

The iminium phenoxide structure has been observed previously for imino phenols.⁶⁸ An interesting example by Salehi⁶⁹ has two similar structures exhibiting both possible tautomeric forms. One structure in that paper has the imino-phenol structure (HL¹), while another exhibits the iminium-phenoxide structure (HL²). HL² has the deprotonated phenol oxygen engaged in both inter- and intramolecular H bonds to iminium hydrogens and has almost exactly the same C-C and C-O bond lengths as seen in L-H₂L4. Examining the intermolecular contacts of L-H₂L4 shows that in addition to the intramolecular hydrogen bonding, there is a similar network of intermolecular hydrogen bonding of H(2) to O(1) of a neighboring molecule (Figure S2).



Figure S2. Hydrogen bonding network of $1-H_2L4$ showing intramolecular contacts between H(1) and O(1), and intermolecular contacts between H(2) and O(1).

The molecular structure of $L-H_2L6$ is shown in Figure S2. The structure has one disordered tert-butyl group on C(27) which was modeled with partial occupancies of 0.715 and 0.285. A diagram of the full structure is included in the ESI. The molecule crystallized with the expected imine-phenol tautomer structure, and contains an intramolecular hydrogen bond between N(1) and O(2). As with H_2L4 , the benzyl carbon C(22) and N(1) are both almost within the plane described by the aromatic phenol ring, with a displacement of 0.041 Å for C(22) and 0.179 Å for N(1). The C(2) is displaced by 0.150 Å, while the C(22)-N(1)-C(2)-C(1) dihedral angle is -118.5°, which places O(1) 1.750 Å from the plane. The diphenyl substitution on the aliphatic alcohol group prevents intermolecular contacts involving O(1). Some of our bidentate amino alcohol ligands have exhibited intermolecular hydrogen bonding when the aliphatic oxygen is primary,⁶⁶ but tertiary alcohols have not exhibited this behavior.^{66, 67} Aside from the intramolecular hydrogen bond between O(2)-H(2)-N(1), there are no additional contacts with the atoms H(2), O(1), O(2) or N(1) within 3Å. As seen in the structure of L-H₂L4, there is a significant difference in the bond lengths from C(24)-O(2), at 1.3570(18) Å, and C(1)-O(1), at 1.4323(19) Å. The bonds between C(24)-C(23), C(23)-C(28), C(27)-C(26) and C(24)-C(25) are all long, at 1.400 - 1.412 Å, while the C(28)-C(27) and C(26)-C(25) bonds are both shorter at around 1.39 Å. The differences in bond-lengths between the phenol aromatic ring and the side

chain aromatic ring are less pronounced than in the case with H_2L4 , presumably due to the fact that the phenol is protonated and there is less disruption of the electronic structure of the aromatic ring.



Figure S3. Complete thermal ellipsoid drawing of L-H₂L6 showing the disordered *tert*-butyl group on C(27) (hydrogen atoms except H(2) and H(2A) are omitted for clarity; thermal ellipsoids are shown at 50% probability). C(34), C(35) and C(36) are modeled at 0.715 occupancy while C(34A), C(35A) and C(36A) are modeled at 0.285 occupancy. Selected bond distances (Å): C(24)-C(25) 1.405(2), C(25)-C(26) 1.393(2), C(26)-C(27) 1.402(2), C(27)-C(28) 1.385(2), C(23)-C(28) 1.400(2), C(24)-O(2) 1.3570(18), C(1)-O(1) 1.4323(19).

The molecular structure of L-H₂L7 is shown in Figure S4. The structure is very similar to that of L-H₂L6, with the expected imine-phenol structure and an intramolecular hydrogen bond linking N(1) and O(2). The phenol aromatic ring is essentially coplanar with O(2), C(19) and N(1), though the deviations of those atoms are opposite in direction to the previous ligands, with C(19) 0.100 Å above the plane, N(1) 0.117 Å above the plane, and C(2) 0.156 Å above the plane. The C(19)-N(1)-C(2)-C(1) dihedral angle is -110.3 °, which places O(1) 1.389 Å below the plane. There are no observed intermolecular hydrogen bond contacts, as they are again blocked by the diphenyl substitution at C(1). The bond lengths from C(21)-O(2), at 1.349 Å, and C(1)-O(1), at 1.438 Å, are again different, intermediate between the previous two structures, but closer to that observed in L-H₂L6. Bond alternation in the phenol aromatic ring is again observed, with approximately the same magnitude as seen in L-H₂L6.



Figure S4. Thermal ellipsoid drawing of L-H₂L7 (hydrogen atoms except H(2) and H(2A) are omitted for clarity; ellipsoids shown at 50% probability). Selected bond distances (Å): C(20)-C(21) 1.412(2), C(21)-C(22) 1.397(2), C(22)-C(23) 1.379(2), C(23)-C(24) 1.376(2), C(24)-C(25) 1.374(2), C(21)-O(2) 1.3492(19), C(1)-O(1) 1.4377(18).

The molecular structure of L-H₂L8 is shown in **Figure S5**. The structure is similar to that of L-H₂L7, with the expected imine-phenol structure and an intramolecular hydrogen bond between N(1) and O(1). The phenol aromatic ring is again essentially co-planar with O(1), C(7) and N(1), and the deviations of those atoms are of the same sign as in L-H₂L7, with C(7) 0.080 Å above the plane, N(1) 0.142 Å above the plane, and C(8) 0.285 Å above the plane. The C(7)-N(1)-C(8)-C(12) dihedral angle is -122.6°, which places O(2) 0.789 Å below the plane. The isopropyl group allows a different geometric placement of the aliphatic oxygen atom, allowing it to be much closer to the plane as compared to the previous ligands. There are no observed intermolecular hydrogen bond contacts, though there is a long range intermolecular contact between O(2) and F(1) in an adjacent molecule at 3.36 Å. The bond lengths from C(1)-O(1), at 1.353(2) Å, and C(12)-O(2), at 1.437 Å, are again different, but closer to L-H₂L6 and L-H₂L7 than the other tautomer exhibited by L-H₂L4. Bond alternation in the phenol aromatic ring is again observed, though this structure exhibits the least deviation from an aromatic structure of the four ligands. This structure is almost identical to that of the previously reported molecule lacking a fluorine.^{60, 71}



Figure S5. Thermal ellipsoid drawing of L-H₂L8 (hydrogen atoms except H(2) and H(8) are omitted for clarity; ellipsoids shown at 50% probability). Selected bond distances (Å): C(1)-C(2) 1.398(3), C(2)-C(3) 1.380(3), C(3)-C(4) 1.380(3), C(4)-C(5) 1.374(3), C(5)-C(6) 1.400(3), C(6)-C(1) 1.406(3), C(1)-O(1) 1.353(2), C(12)-O(2) 1.437(2).



Figure S6. Calculated structures of imine phenol (left) and iminium phenoxide tautomers of L- H_2L4 (Gaussian B3LYP 6-31G(d,p)). Energies: iminium phenoxide, -903.79254086 au, imine phenol, -903.79907612 au, energy difference: 0.00653 au = 4.1 kcal/mol.

XYZ files output from Gaussian

T	Dhanani da	tautaman D'	2T V D / (2) 21 m (d m)	000 70054000	
Iminium	Phenoxide	tautomer, B.	3LYP/6-31g(a,p),	-903./9254086	a.u.
	2.603000	-0.452000	0.832000		
C	3.913000	-0.989000	1.095000		
C	4.843000	-1.124000	0.095000		
C	4.567000	-0.745000	-1.249000		
C	3.335000	-0.228000	-1.556000		
C	2.336000	-0.066000	-0.554000		
C	1.077000	0.466000	-0.886000		
C -	-1.186000	1.198000	-0.202000		
C -	-1.137000	2.756000	-0.020000		
C -	-2.504000	3.393000	-0.307000		
C -	-0.613000	3.166000	1.357000		
C -	-2.207000	0.460000	0.697000		
C -	-2.398000	-0.998000	0.326000		
C -	-3.439000	-1.373000	-0.534000		
C -	-3.622000	-2.706000	-0.901000		
C -	-2.761000	-3.689000	-0.410000		
C -	-1.721000	-3.328000	0.448000		
С -	-1.540000	-1.994000	0.815000		
H	0.477000	0.318000	0.967000		
н -	-0.583000	3.251000	-1.837000		
H	5.823000	-1.533000	0.333000		
H	5.326000	-0.865000	-2.015000		
H	3.100000	0.071000	-2.576000		
H	0.856000	0.758000	-1.911000		
H -	-1.460000	1.018000	-1.252000		
н -	-4.119000	-0.613000	-0.913000		
H -	-4.438000	-2.977000	-1.565000		
H -	-2.902000	-4.729000	-0.691000		
H -	-1.047000	-4.086000	0.836000		
H -	-0.719000	-1.726000	1.474000		
н -	-2.908000	3.052000	-1.270000		
н -	-3.243000	3.165000	0.465000		
н -	-2.387000	4.479000	-0.352000		
н -	-1.289000	2.841000	2.153000		
H	0.375000	2.738000	1.540000		
н -	-0.526000	4.255000	1.402000		
Н -	-3.169000	0.975000	0.629000		
Н -	-1.877000	0.542000	1.739000		
H	4.129000	-1.279000	2.118000		
N	0.128000	0.619000	0.017000		
0	1.729000	-0.324000	1.750000		
0 -	-0.177000	3.266000	-0.959000		

42					
imine	phenol tauto	omer, B3LYP/6	-31g(d,p),	-903.79907612	a.u.
С	2.727000	-0.176000	0.804000		
С	4.039000	-0.610000	1.049000		
С	4.888000	-0.900000	-0.012000		
С	4.456000	-0.770000	-1.340000		
С	3.160000	-0.343000	-1.589000		
С	2.277000	-0.039000	-0.538000		
С	0.920000	0.401000	-0.824000		
С	-1.253000	1.146000	-0.226000		
С	-1.326000	2.708000	-0.097000		
С	-2.727000	3.232000	-0.448000		
С	-0.885000	3.210000	1.279000		
С	-2.274000	0.376000	0.650000		
С	-2.331000	-1.107000	0.345000		
С	-3.297000	-1.611000	-0.537000		
С	-3.352000	-2.972000	-0.842000		
С	-2.437000	-3.854000	-0.265000		
С	-1.471000	-3.365000	0.615000		
С	-1.419000	-2.004000	0.917000		
H	-0.710000	3.108000	-1.914000		
H	5.901000	-1.234000	0.195000		
H	5.128000	-1.000000	-2.161000		
H	2.806000	-0.237000	-2.612000		
H	0.636000	0.463000	-1.884000		
H	-1.477000	0.922000	-1.285000		
H	-4.018000	-0.931000	-0.985000		
H	-4.112000	-3.342000	-1.524000		
H	-2.479000	-4.915000	-0.497000		
H	-0.755000	-4.044000	1.070000		
H	-0.657000	-1.632000	1.596000		
H	-3.075000	2.828000	-1.408000		
H	-3.4/3000	2.979000	0.310000		
H	-2.686000	4.321000	-0.533000		
H	-1.544000	2.838000	2.069000		
H	0.136000	2.888000	1.494000		
H	-0.912000	4.303000	1.293000		
H	-3.268000	0.80/000	0.501000		
H	-2.008000	0.531000	1./01000		
П N	4.304000		2.080000		
	0.084000	0.091000	U.LU/UUU		
0	1.92/000	0.094000	1 010000		
0	-0.3/9000	J.200000	-1.019000		
п	T.023000	0.399000	1.405000		



Figure S7. Thermal ellipsoid drawing of complex **14** showing three solvent molecules (two disordered) of CH_2Cl_2 (hydrogen atoms are omitted for clarity; thermal ellipsoids are shown at 50% probability). C(3S), Cl(5S), Cl(6S), C(2S), Cl(3S) and Cl(4S) are modeled at 0.519 occupancy; C(3T), Cl(5T), Cl(6T), C(2T), Cl(3T) and Cl(4T) are modeled at 0.481 occupancy.



















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	- 22.9186						_
	7875.92					*	
	2469.05 —						
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lrbon	2857.821 6319.821 6319.821 77302.721 8002.721 8002.721 2541.721						- - -
T) SI) CE	8973.021 — — 150.5765						
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11 (NMe2) 2 (P	8282.491 - 8725.491 - 9460.9946						- 160





Hydroamination of hepta-4,5-dienylamine at 135 °C with *in situ* catalysts (5 mol% catalyst) to give tetrahydropyridine **2a**, *Z*- or *E*- α -vinylpyrrolidines *Z*- and *E*-**3a**























Hydroamination of 6-methyl-hepta-4,5-dienylamine 1c at 135 °C with *in situ* catalysts (5 mol% catalyst).

























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