

# Inorganic Chemistry

including bioinorganic chemistry

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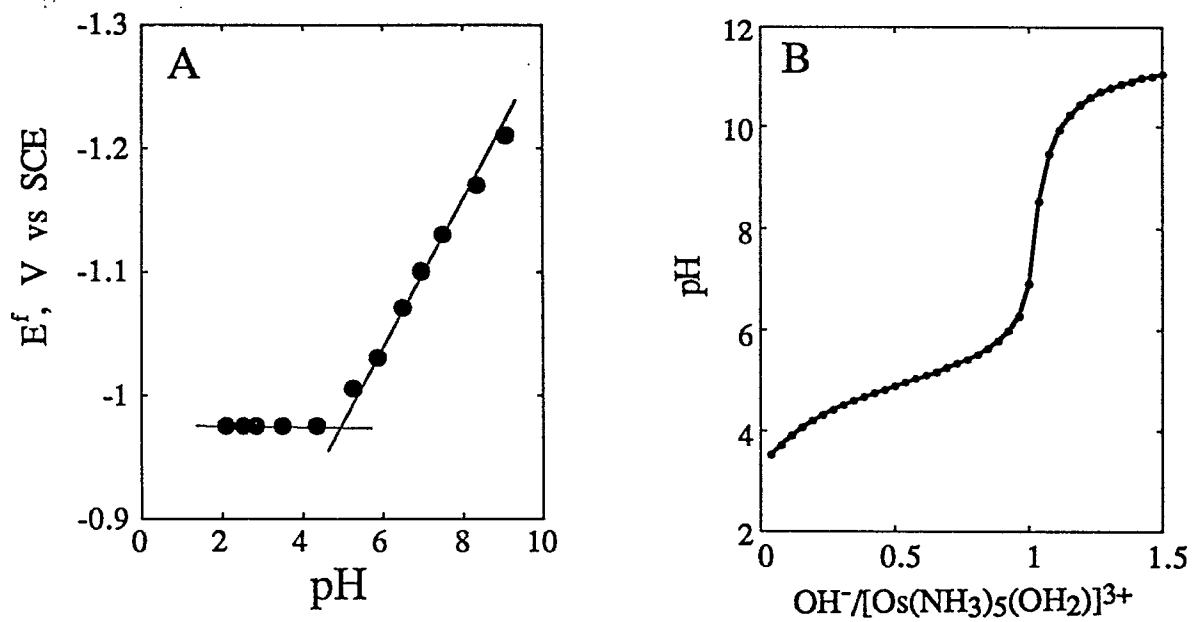


Figure S1. (A) Variation of the formal potential of the  $[\text{Os}(\text{NH}_3)_5(\text{OH}_2)]^{3+/2+}$  couple with the pH of the supporting electrolytes which were buffered between pH 2.1 and 9.1. (B) pH titration curve for the titration of 5.4 mM  $[\text{Os}(\text{NH}_3)_5(\text{OH}_2)]^{3+}$  with 0.01 M NaOH.

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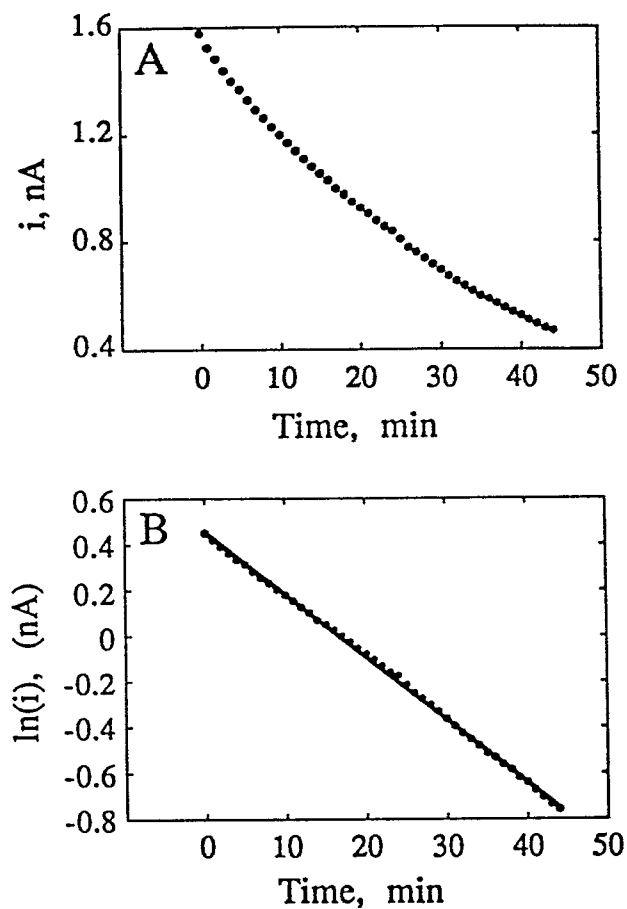


Figure S2. (A) Time dependence of the steady-state current for the oxidation of  $[\text{Os}(\text{NH}_3)_5(\text{OH}_2)]^{2+}$  at a carbon fiber microelectrode maintained at  $-0.7$  V. The initial concentration of Os(II) was  $1.3$  mM. Supporting electrolyte:  $0.1$  M  $\text{CH}_3\text{SO}_3\text{Na}$ . Initial pH =  $5.6$ . (B) First-order plot of the data from (A).

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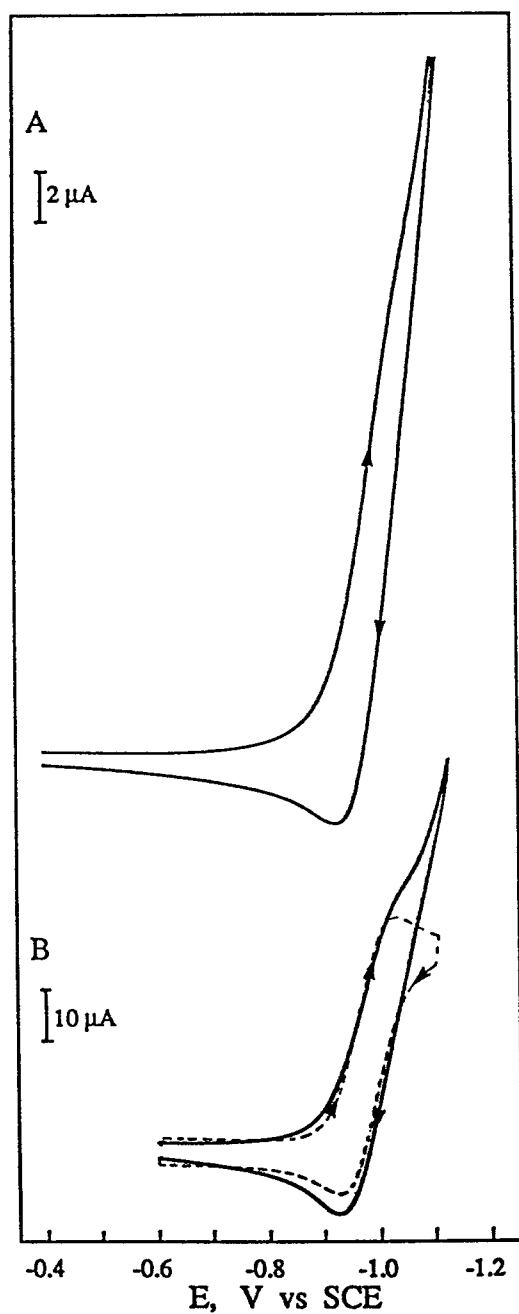


Figure S3. (A) Cyclic voltammogram for the reduction of  $0.74 \text{ mM}$   $[\text{Os}(\text{NH}_3)_5(\text{OH}_2)]^{3+}$  in  $0.1 \text{ M CH}_3\text{SO}_3\text{Na} + 0.48 \text{ M CH}_3\text{COOH}$  ( $\text{pH} = 2.53$ ). Scan rate:  $5 \text{ mV s}^{-1}$ . (B) Repeat of (A) at a scan rate of  $50 \text{ mV s}^{-1}$ . The dashed curve was calculated with the Digisim program (see text).

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