Inorganic Chemistry
including bioinorganic chemistry

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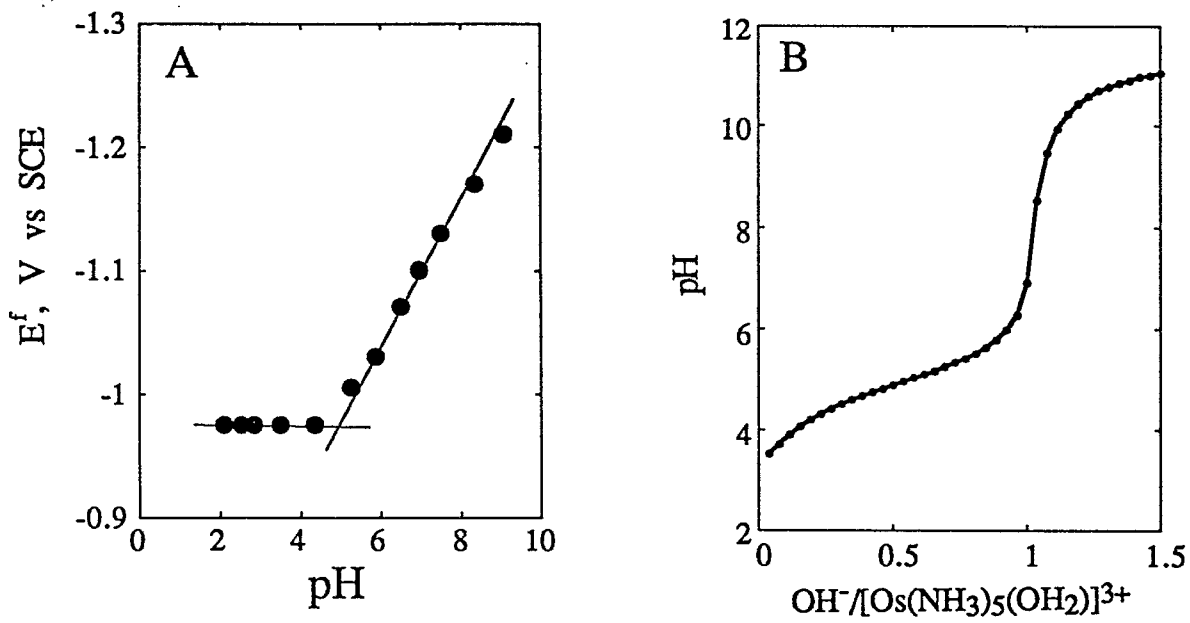

Figure S1. (A) Variation of the formal potential of the $[Os(NH_3)_5(OH_2)]^{3+/2+}$ couple with the pH of the supporting electrolytes which were buffered between pH 2.1 and 9.1. (B) pH titration curve for the titration of 5.4 mM $[Os(NH_3)_5(OH_2)]^{3+}$ with 0.01 M NaOH.

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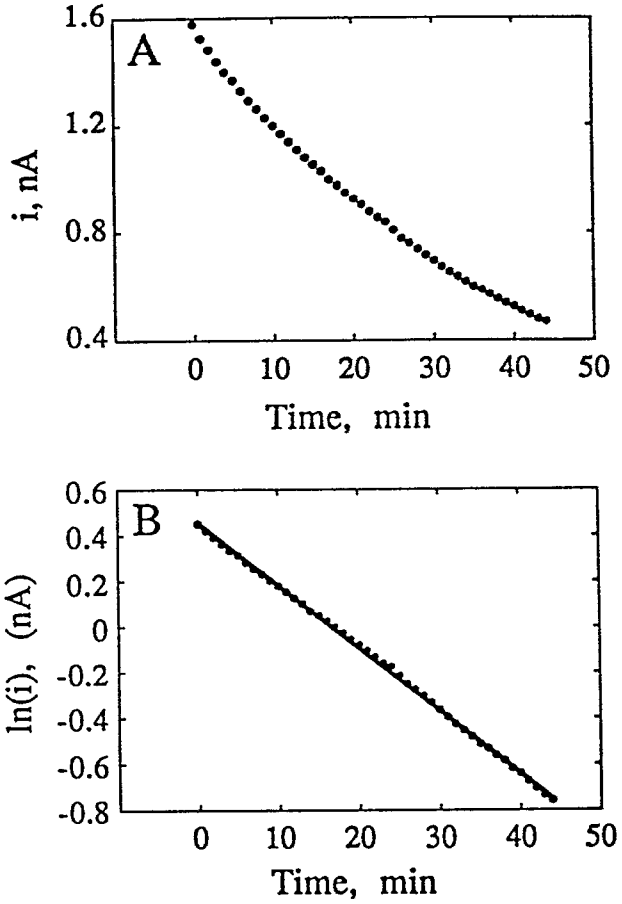

Figure S2. (A) Time dependence of the steady-state current for the oxidation of $[Os(NH_3)_5(OH_2)]^{2+}$ at a carbon fiber microelectrode maintained at -0.7 V. The initial concentration of Os(II) was 1.3 mM. Supporting electrolyte: 0.1 M CH_3SO_3Na. Initial pH = 5.6. (B) First-order plot of the data from (A).

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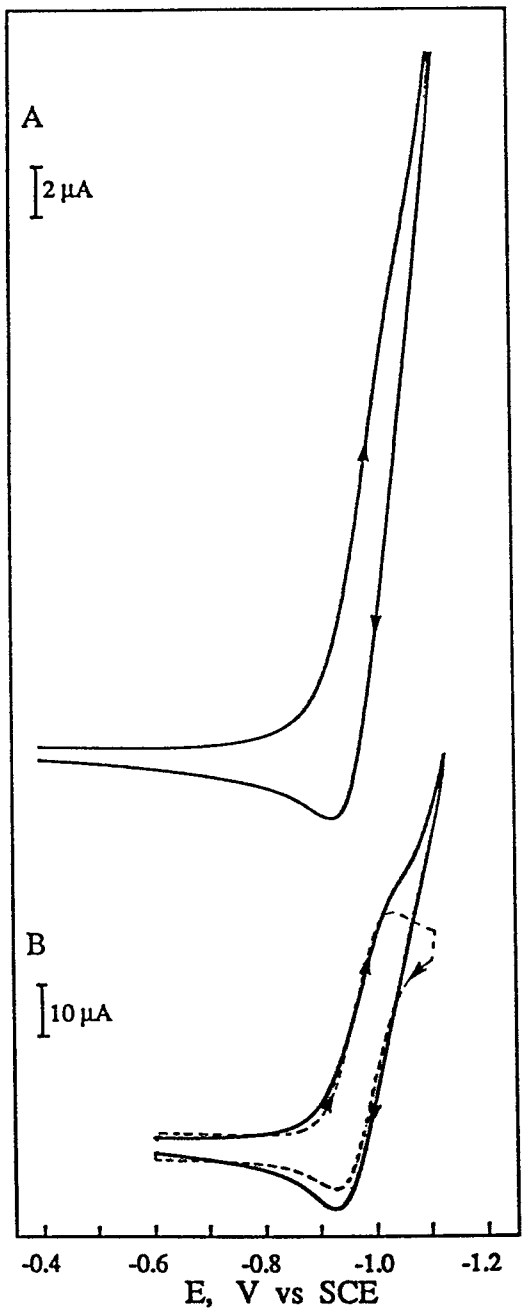

Figure S3. (A) Cyclic voltammogram for the reduction of 0.74 mM $[Os(NH_3)_5(OH_2)]^{3+}$ in 0.1 M CH_3SO_3Na + 0.48 M CH_3COOH (pH = 2.53). Scan rate: 5 mV s^{-1}. (B) Repeat of (A) at a scan rate of 50 mV s^{-1}. The dashed curve was calculated with the Digisim program (see text).

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