

JUPITER: AEROSOL CHEMISTRY IN THE POLAR ATMOSPHERE

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ABSTRACT

Aromatic compounds have been considered a likely candidate for enhanced aerosol formation in the polar region of Jupiter. We develop a new chemical model for aromatic compounds in the Jovian auroral thermosphere/ionosphere. The model is based on a previous model for hydrocarbon chemistry in the Jovian atmosphere and is constrained by observations from *Voyager*, *Galileo*, and the *Infrared Space Observatory*. Precipitation of energetic electrons provides the major energy source for the production of benzene and other heavier aromatic hydrocarbons. The maximum mixing ratio of benzene in the polar model is 2×10^{-9} , a value that can be compared with the observed value of $(2_{-1}^{+2}) \times 10^{-9}$ in the north polar auroral region. Sufficient quantities of the higher ring species are produced so that their saturated vapor pressures are exceeded. Condensation of these molecules is expected to lead to aerosol formation.

Subject heading: planets and satellites: individual (Jupiter)

1. INTRODUCTION

The first ring compound, benzene ($c\text{-C}_6\text{H}_6$, also known as A_1 in this Letter), was detected on Jupiter by the *Voyager* Infrared Interferometer Spectrometer (IRIS) experiment around $10 \mu\text{m}$ (Kim et al. 1985). Recent observations by the *Infrared Space Observatory* (ISO) have also identified benzene on Jupiter (Encrenaz et al. 1997). In addition, the *Galileo* probe mass spectrometer has detected traces of benzene (Niemann et al. 1998). The formation of benzene on Jupiter has been attributed to successive additive reactions involving C_2H_2 (Allen & Yung 1985), but no detailed modeling has been reported. In particular, the energy source for the synthesis of ring compounds was not identified.

In this Letter, we present preliminary results of a chemical model of benzene and polycyclic aromatic hydrocarbons (PAHs) in the Jovian auroral region. We shall argue that charged energetic particles in addition to the enhanced auroral temperature are responsible for the production of benzene and PAHs. Our model, which incorporates the most important reaction pathways for synthesizing PAHs, is a combination of a photochemical model of hydrocarbons in the outer solar system and the H atom and ion chemistry of Perry et al. (1999).

2. PHOTOCHEMICAL MODEL

We base our model on the one-dimensional hydrocarbon photochemical model developed at Caltech/Jet Propulsion Laboratory (Gladstone, Allen, & Yung 1996; Lee, Yung, & Moses 2000). To this most comprehensive and recently updated model, we make four major modifications: (1) adding important reactions that produce PAHs; (2) including reactions that are the result of precipitating electrons; (3) using a temperature profile that is appropriate for the auroral region; and (4) taking into account the condensation of higher ring molecules. We use the same eddy diffusion coefficient as that used in the low-latitude model. The latitude of the polar model is 60° .

We add 98 PAH-related reactions to the model for the Jovian atmosphere. The most important reactions, along with kinetic rate constants, are summarized in Table 1. We adopt the no-

menclature used by Wang & Frenklach (1994). Our calculation includes aromatic molecules up to pyrene (A_4). The reaction scheme is adopted largely from laboratory and modeling studies in flame chemistry. In the absence of laboratory data for some reactions, we estimate their rate coefficients by analogy with known reactions. We use the same photodissociation coefficients for all A_n , since the cross section data for the $n > 1$ are unavailable. The major pathways for the formation of benzene and PAHs are shown in Figure 1.

The primary driving force for aerosol chemistry in the auroral atmosphere of Jupiter is dissociative recombination of methane ions that are produced from interaction with high-energy particles. Previous studies have shown that the power of the aurorae is in the range 10^{13} – 10^{14} W and is greater than the EUV power from the Sun (see, e.g., Atreya 1986, § 6.3). Perry et al. (1999) have carried out electron transport calculations for the effects of precipitating electrons with a range of energies from 20 to 100 keV and have presented dissociation and ionization rates of various ionic species. In our study, we adopt their calculated reaction rates (model B) for seven reactions involving the dissociative recombination reactions of H_3^+ and CH_5^+ , obtained from their Figure 4 (see reactions R315–321 in Table 1). The total input energy is about 11 ergs $\text{cm}^{-2} \text{s}^{-1}$. The destruction rate of methane by precipitating electrons is about 25 times higher than that by photolysis at 60° .

Although there is no direct measurement of the temperature profile, the temperature in the auroral region is known to be much higher than that in nonauroral regions because of the heat deposition from precipitating particles. We use a temperature profile deduced by Trafton et al. (1994), which is consistent with observed rotational temperatures of H_2 (Clarke et al. 1994; Liu & Dalgarno 1996; Kim, Fox, & Caldwell 1997). To reconcile the higher observed thermospheric temperature with the low temperature computed in the model of Perry et al. (1999), we suggest that the former is a sporadic phenomenon, whereas the model calculation is for the time-averaged atmosphere. Since the synthesis of PAHs occurs primarily during the transient high-temperature phase, the higher temperature profile should be more appropriate.

TABLE 1
PARTIAL LIST OF REACTIONS AND RATE COEFFICIENTS USED IN AEROSOL CHEMISTRY MODEL

Number	Reaction	Rate Coefficients	Reference
R211	$C_4H_5 + C_2H_2 \rightarrow l-C_6H_6 + H$	$9.63 \times 10^{-16} T^{1.02} e^{-5489/T}$	1
R235	$C_3H_3 + C_3H_3 + M \rightarrow A_1 + M$	$k_{\infty} = 1.66 \times 10^{-13} \text{ a}$	1
R241	$C_4H_3 + C_2H_2 \rightarrow l-C_6H_4 + H$	$6.14 \times 10^{-8} T^{-1.21} e^{-5589/T}$	1
R242	$C_4H_3 + C_2H_2 + M \rightarrow n-C_6H_5 + M$	$k_{\infty} = 7.08 \times 10^{-20} T^{1.97} e^{2820/T} \text{ a}$	2
R243	$C_4H_3 + C_2H_2 + M \rightarrow A_1 + M$	$k_{\infty} = 3.82 \times 10^{44} T^{-17.65} e^{-12287/T} \text{ a,b}$	3
R255	$l-C_6H_4 + H + M \rightarrow n-C_6H_5 + M$	$k_{\infty} = 5.48 \times 10^{20} T^{-10.04} e^{-9467/T} \text{ a,b}$	1
R256	$l-C_6H_4 + H + M \rightarrow A_1 + M$	$k_{\infty} = 5.98 \times 10^{53} T^{-20.09} e^{-14150/T} \text{ a,b}$	1
R258	$n-C_6H_5 + H \rightarrow l-C_6H_4 + H_2$	2.49×10^{-11}	4
R259	$n-C_6H_5 + H + M \rightarrow l-C_6H_6 + M$	$k_{\infty} = 1.83 \times 10^{18} T^{-9.65} e^{-3525/T} \text{ a}$	4
R261	$l-C_6H_6 + H \rightarrow A_1 + H$	$1.44 \times 10^{-7} T^{-1.34} e^{-1762/T}$	1
R262	$l-C_6H_6 + H + M \rightarrow n-C_6H_7 + M$	$k_{\infty} = 2.49 \times 10^{-8} T^{-1.69} e^{-805.7/T}$ $k_0 = 8.00 \times 10^{-31} T^{-0.52} e^{-503.6/T}$	1
R264	$l-C_6H_7 + M \rightarrow A_1 + H + M$	$1.39 \times 10^{-2} T^{-4.22} e^{-5690/T}$	1
R265	$A_1 + H + M \rightarrow A_1 + M$	$k_{\infty} = 1.66 \times 10^{-10}$ $k_0 = 1.82 \times 10^{28} T^{-16.3} e^{-3526/T}$	4
R266	$A_1 + C_2H_2 + M \rightarrow n-A_1C_2H_2 + M$	$k_{\infty} = 1.28 \times 10^{17} T^{-9.19} e^{-6748/T} \text{ a,b}$	1
R267	$A_1 + C_2H_2 \rightarrow A_1C_2H + H$	$1.25 \times 10^3 T^{-3.96} e^{-8611/T} \text{ b}$	4
R270	$A_1 + H \rightarrow A_1 + H_2$	$4.15 \times 10^{-10} e^{-8057/T}$	5
R273	$A_1C_2H + H + M \rightarrow n-A_1C_2H_2 + M$	$k_{\infty} = 1.66 \times 10^{30} T^{-12.76} e^{-8661/T} \text{ a,b}$	1
R276	$n-A_1C_2H_2 + C_2H_2 \rightarrow A_2 + H$	$3.49 \times 10^{-9} T^{-1.07} e^{-2417/T} \text{ b}$	4
R277	$n-A_1C_2H_2 + H \rightarrow A_1C_2H + H_2$	2.49×10^{-11}	4
R285	$A_2-1 + H + M \rightarrow A_2 + M$	$k_{\infty} = 1.66 \times 10^{-10}, k_0 = 3.33 \times 10^{-12}$	4
R287	$A_2-1 + C_2H_2 + M \rightarrow A_2C_2H_2 + M$	$k_{\infty} = 7.47 \times 10^{15} T^{-8.71} e^{-7201/T} \text{ a,b}$	1
R288	$A_2-1 + C_2H_2 \rightarrow A_2C_2HA + H$	$2.32 \times 10^{-2} T^{-2.64} e^{-8762/T} \text{ b}$	1
R289	$A_2 + H \rightarrow A_2-1 + H_2$	$k_{289} = k_{270}$	
R293	$A_2C_2HA + H + M \rightarrow A_2C_2H_2 + M$	$k_{\infty} = 3.15 \times 10^{26} T^{-11.63} e^{-8158/T} \text{ a,b}$	1
R294	$A_2C_2H_2 + H \rightarrow A_2C_2HA + H_2$	2.49×10^{-11}	4
R297	$A_2C_2H_2 + C_2H_2 \rightarrow A_3 + H$	$k_{297} = k_{276}$	
R303	$A_3-4 + H + M \rightarrow A_3 + M$	$k_{\infty} = 1.66 \times 10^{-10}, k_0 = 2.21 \times 10^{-14}$	4
R304	$A_3 + C_2H \rightarrow A_3C_2H + H$	8.30×10^{-11}	4
R305	$A_3-4 + C_2H_2 + M \rightarrow A_3C_2H_2 + M$	$k_{\infty} = 1.11 \times 10^{22} T^{-10.55} e^{-10675/T} \text{ a,b}$	4
R306	$A_3-4 + C_2H_2 \rightarrow A_3C_2H + H$	$1.33 \times 10^{-6} T^{-1.21} e^{-11380/T} \text{ b}$	1
R308	$A_3C_2H + H + M \rightarrow A_3C_2H_2 + M$	$k_{\infty} = 8.63 \times 10^{23} T^{-11.05} e^{-7402/T} \text{ a,b}$	1
R310	$A_3C_2H_2 + M \rightarrow A_4 + H + M$	$1.12 \times 10^{25} T^{-11.86} e^{-14150/T} \text{ b}$	1
R312	$A_4 + H + M \rightarrow A_4 + M$	$k_{\infty} = 1.66 \times 10^{-10} \text{ a}$	4
R315	$H_3^+ + e \rightarrow H_2 + H$	$2.88 \times 10^{-8} (300/T)^{0.65}$	6
R316	$H_3^+ + e \rightarrow H + H + H$	$8.63 \times 10^{-8} (300/T)^{0.65}$	6
R317	$CH_5^+ + e \rightarrow {}^3CH_3 + H_2 + H$	$5.3 \times 10^{-7} (300/T)^{0.52}$	6
R318	$CH_5^+ + e \rightarrow CH_3 + H + H$	$1.9 \times 10^{-8} (300/T)^{0.52}$	6
R319	$CH_5^+ + e \rightarrow CH_3 + H_2$	$1.7 \times 10^{-8} (300/T)^{0.52}$	6
R320	$CH_5^+ + e \rightarrow CH_4 + H$	$9.0 \times 10^{-9} (300/T)^{0.52}$	6
R321	$CH_5^+ + e \rightarrow CH + H_2 + H_2$	$9.5 \times 10^{-9} (300/T)^{0.52}$	6

NOTE.—The complete set includes all the relevant reactions considered by Gladstone et al. 1996, updated by Lee et al. 2000, and some other less important PAH reactions studied in this work. Units for two- and three-body rate coefficients are $\text{cm}^3 \text{s}^{-1}$ and $\text{cm}^6 \text{s}^{-1}$, respectively; k_0 and k_{∞} refer to the low and high pressure limits of the rate coefficients.

^a $k_0 = 10^{-27}$, estimated.

^b Set at maximum value near 750 K.

REFERENCES.—(1) Wang & Frenklach 1994; (2) Colket 1986; (3) Westmoreland et al. 1989; (4) Wang & Frenklach 1997; (5) Keifer et al. 1985; (6) Perry et al. 1999.

Ring compounds A_1 through A_4 are condensed when their vapor pressures are exceeded. We define the loss rate of a condensing species by a procedure described in Summers & Strobel (1989), with a condensation time constant of 10^6 s. The vapor pressures are taken from Yaws (1994).

3. RESULTS AND DISCUSSION

The results of the model calculations are shown in Figure 2. Figure 2a presents the mixing ratios for CH_4 , C_2H_6 , C_2H_4 , and C_2H_2 . Compared with a solar-driven model at lower latitudes (Lee et al. 2000), the polar model has an order of magnitude more C_2 hydrocarbons. The reason is the greater rate of CH_4 dissociation, as discussed earlier. Figure 2b shows that the maximum mixing ratio of benzene is 2×10^{-9} . This result is in fairly good agreement with the *Voyager* observation of $(2_{-1}^{+2}) \times 10^{-9}$ in the north polar auroral region (Kim et al. 1985).

Above 1 mb, the column density is $9.3 \times 10^{13} \text{ cm}^{-2}$, which can be compared with the preliminary result of stratospheric midlatitude (not the auroral region) benzene column density at $2 \times 10^{14} \text{ cm}^{-2}$ from the *ISO* observation (B. Bézard 1999, private communication).

We study the response of the model to various effects including temperature, precipitating electrons, and condensation. Using a much lower temperature profile from *Voyager* measurements, we find that the total benzene concentration does not vary much, although the production and destruction rates are both decreased by a factor of 10. The concentrations of higher ring compounds A_2 , A_3 , and A_4 are decreased by factors of 2, 100, and 20,000, respectively. Without the ion-related reactions, benzene concentration is reduced by a factor of 550. This confirms that the energetic particles are essential to benzene formation. Test runs show that 62% of the higher ring

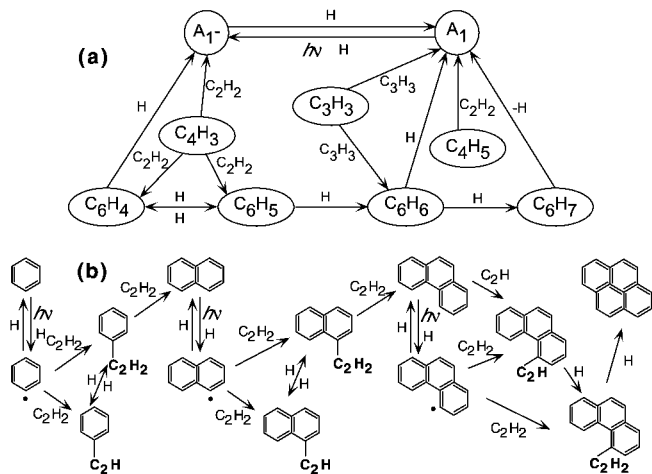


FIG. 1.—Schematic diagram showing the major pathways for (a) benzene (A_1) and (b) A_2 , A_3 , and A_4 formation in the auroral atmosphere of Jupiter.

molecules are condensed. The sensitivity of the benzene mixing ratio to the eddy diffusion coefficient will be tested in future works.

The total dissociation rate of CH_4 in the model is 6.7×10^{10} molecules $\text{cm}^{-2} \text{s}^{-1}$. About 1% of this leads to the production of A_n species, resulting in a mass flux of about 10^{-14} g $\text{cm}^{-2} \text{s}^{-1}$. Polar haze particles occupy the region $\sim 0.3\text{--}50$ mb (West 1988; Pryor et al. 1991), probably concentrated near 10 mb (R. A. West 1999, private communication). Assuming the particles have a mean radius of $0.1 \mu\text{m}$, we estimate that a haze layer with optical depth of 0.1 in the near UV could be generated by the mass flux computed in our model. Our model yields for PAHs are considerably larger than what was generated in a laboratory study for Titan's atmosphere (Sagan et al. 1993).

4. CONCLUSION

Benzene was discovered over a decade ago, and the origin of polar aerosols on Jupiter has remained a puzzle for a long time. This is the first quantitative chemical model that attempts to simulate the production of aromatic compounds. Our simple model can account for most of the observations and suggests that the condensation of heavy aromatic molecules is a source for the aerosols. Our model can be tested by new observations. PAH concentrations should vary as the auroral temperature changes. The difference in benzene concentrations should be

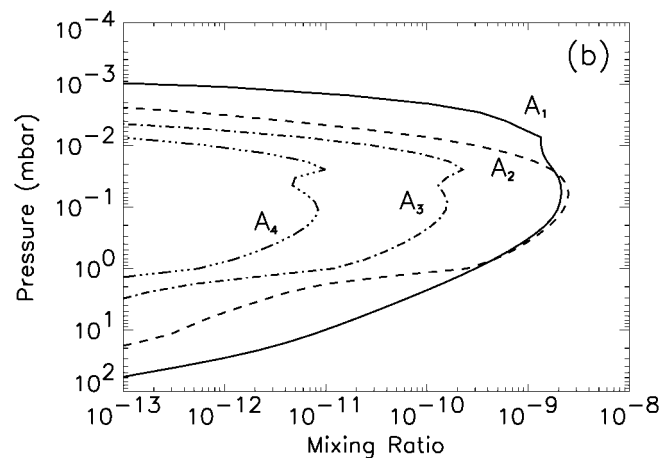
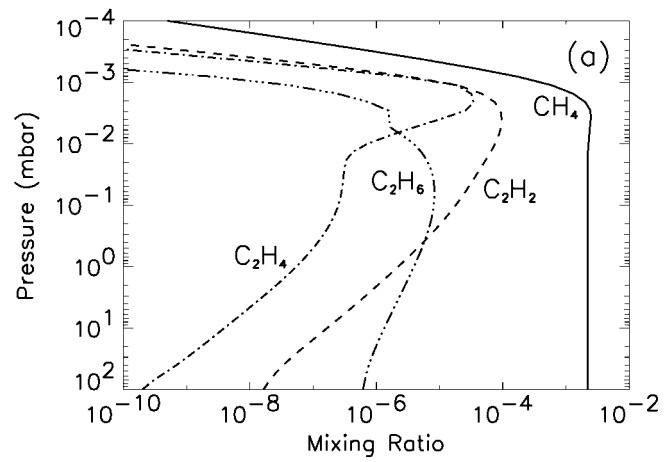


FIG. 2.—Hydrocarbon mixing ratio profiles as a function of pressure in the polar model for (a) CH_4 , C_2H_2 , C_2H_4 , and C_2H_6 and (b) A_1 , A_2 , A_3 , and A_4 .

observable at different latitudes, as the IRIS observations (Kim et al. 1985) suggested. More extensive observations of benzene and PAHs are needed to test our theory.

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