Published in final edited form as:

Nat Chem. 2013 April ; 5(4): 293–299. doi:10.1038/nchem.1578.

Redox-Inactive Metals Modulate the Reduction Potential in Heterometallic Manganese-Oxido Clusters

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Abstract

Redox-inactive metals are found in biological and heterogeneous water oxidation catalysts, but their roles in catalysis are currently not well understood. A series of high oxidation state tetranuclear-dioxido clusters comprised of three manganese centers and a redox-inactive metal (M) of various charge is reported. Crystallographic studies show an unprecedented $Mn_3M(\mu_4\text{-}O)(\mu_2\text{-}O)$ core that remains intact upon changing M or the manganese oxidation state. Electrochemical studies reveal that the reduction potentials span a window of 700 mV, dependent upon the Lewis acidity of the second metal. With the pK_a of the redox-inactive metal-aqua complex as a measure of Lewis acidity, these compounds display a linear dependence between reduction potential and acidity with a slope of ca. 100 mV per pK_a unit. The Sr^{2+} and Ca^{2+} compounds show similar potentials, an observation that correlates with the behavior of the OEC, which is active only in the presence of one of these two metals.

Redox-inactive metal ions are critical components in many biological electron transfer reactions.[1,2] For example, Ca^{2+} is essential for activity in the oxygen-evolving complex (OEC) of Photosystem II (PSII), although its exact role in catalysis is still unclear.[3–5] There have been numerous studies of electron transfer to synthetic organic substrates,[1,6] but studies of electron transfer to metal oxo complexes relevant to the active sites of a number of metalloenzymes have been fewer, possibly due to the challenge of isolating complexes with bound metal ions. Recently, a non-heme $Fe^{IV}O$ complex with a bound Sc^{3+} ion was characterized crystallographically, and it was found that the presence of Sc^{3+} or Ca^{2+} in solution allowed the two-electron reduction of the complex using ferrocene.[7] The addition of other Lewis acidic metal ions to a different $Fe^{IV}O$ compound greatly enhanced electron

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Author Contributions

E.Y.T. and T.A. designed research. E.Y.T. and R.T. performed experiments. R.T. and J.Y. provided XANES characterization. E.Y.T., R.T., J.Y., and T.A. analyzed data. E.Y.T. and T.A. wrote the paper.

The authors declare no competing financial interests.

Supplementary information and chemical compound information are available in the online version of the paper.

transfer rates, although the adducts were not structurally characterized.[8] With a monomeric Mn^{II} complex, faster rates of oxygen reduction were observed in the presence of Ca^{2+}, and a Mn^{III}–(μ-OH)–Ca^{II} complex was isolated as the product.[9] Substitution of Sr^{2+} for Ca^{2+} in this complex showed a similar reduction potential, while substitution with Ba^{2+} resulted in a more negative reduction potential.[10] O-atom transfer from some manganese oxo complexes has also been promoted by the binding of redox-inactive metals such as lithium.[11]

Redox-inactive metal ions also play a role in non-biological electron transfer reactions such as water oxidation catalyzed by heterogeneous cobalt[12–14] and manganese oxides[15–17] containing alkali or alkali earth metals. In these examples, there has been speculation that the redox-inactive metal is associated with the transition metal catalyst in cubane-like structures reminiscent of the crystallographically determined structure of the OEC, an oxide-bridged $CaMn_4$ cluster in which the calcium center is associated with three of the manganese centers in a cubane motif (Fig. 1).[18,19] A recent study varying the redox-inactive metal (K^+, Ca^{2+}, Sr^{2+}, Mg^{2+}) in layered heterogeneous manganese oxides showed that, as in the OEC, the presence of Ca^{2+} allows for the highest catalytic activity.[15] Additionally, heterogeneous mixed oxides of cobalt show different water oxidation behavior dependent on the nature of the redox-inactive metal present in the mixture.[20] The role of the redox-inactive metal on electron transfer and catalysis within the material remains unclear in these systems.

Our recent report of a structural model of the $CaMn_3$ subsite of the OEC containing a high oxidation state heterometallic $Mn^{IV}{}_3CaO_4$ moiety and comparison to a tetramanganese analog suggested a significant influence of the calcium center on the redox properties of the cluster.[21] To study the scope and chemical basis of this phenomenon in multimetallic oxide clusters such as those found in the OEC and in heterogenous systems, access to well-defined and structurally related heterometallic oxido clusters of redox-active and inactive metals is desirable. Based on a multinucleating ligand-based synthetic strategy developed by our group,[22,23] we targeted heteronuclear clusters supported by the hexapyridyl trisalkoxido 1,3,5-triarylbenzene ligand (H_3L). In the present work, we describe the synthesis of a series of tetranuclear heterometallic trimanganese dioxo clusters [$Mn_3M(\mu_4\text{-}O)(\mu_2\text{-}O)$] containing a redox-inactive cation bridging via oxido moieties to manganese centers. Electrochemical characterization reveals that large changes in the Lewis acidity of the redox-inactive metal have a systematic effect on the redox properties of the cluster.

Results and discussion

Heterometallic clusters have been targeted by several groups as proposed structural models of the OEC as evidence it contained a mixed Mn-Ca-oxido cluster emerged.[24–30] Calcium-manganese clusters remain uncommon, though recently there have been a number of such complexes that have been isolated and structurally characterized.[9,21,25–30] Because heterometallic clusters are often synthesized by self-assembly, controlling the composition and relative arrangement of metals has been a challenge. Furthermore, these synthetic protocols are not necessarily easily extended to the incorporation of other redox-inactive metals instead of calcium. To develop general syntheses of heterometallic clusters, we employed a multinucleating ligand that affords versatile trimetallic ($M^{II}{}_3$) precursors.[23]

These Mn^{II}_3 species could be elaborated to site-differentiated tetramanganese cubanes, Mn_4O_4, as well as heteronuclear Mn_3CaO_4 clusters.[21] Strategies for general synthetic protocols to related heteronuclear complexes were then explored.

A dicationic calcium-bridged hexamanganese complex ($[LMn_2^{III}Mn^{II}O(OAc)_3]_2Ca(OTf)_2$) in which each trimanganese unit is coordinated by a μ_3-oxide was prepared and proposed to be an intermediate in the synthesis of $LMn^{IV}_3CaO_4(OAc)_3(THF)$.[21] In an effort to isolate other manganese clusters of high oxidation state relevant for the preparation of structural mimics of the OEC, $[LMn_2^{III}Mn^{II}O(OAc)_3]_2Ca(OTf)_2$ was treated with $Ca(OTf)_2$ and PhIO in 1,2-dimethoxyethane (DME) to form a red-purple compound (Fig. 2). The same compound was also independently synthesized in high yield (84%) in one step from the more reduced $LMn^{II}_3(OAc)_3$ precursor (Fig. 2). A single crystal X-ray diffraction (XRD) study of this species confirmed the material to be a calcium trimanganese dioxo complex ($[\mathbf{1}\text{-}Ca(DME)(OTf)]^{2+}$ ($\mathbf{1}$ = $LMn^{IV}Mn^{III}_2O_2(OAc)_2$), Fig. 3). The metal oxidation states were assigned based on crystallographic, XAS, and magnetism data (vide infra). In this complex, as in $LMn^{II}_3(OAc)_3$, the three manganese centers are bridged by three alkoxide donors from L, forming a six-membered ring, and the pyridine nitrogens of each dipyridyloxymethyl moiety coordinate to adjacent metal centers. Ca^{2+} is bridged to the trimanganese cluster by a μ_4-oxido, to the Mn^{IV} center by a μ_2-oxido, and to the remaining Mn^{III} centers by bridging acetate moieties. The Ca^{2+} is further coordinated by a bidentate DME ligand and a trifluoromethanesulfonate anion. Two trifluoromethanesulfonate ions remain outer-sphere. The isolated compounds reported here display diagnostic ^{1}H NMR spectra, although the paramagnetically broadened and shifted signals have not been assigned. The addition of excess water to a CD_2Cl_2 solution of $[\mathbf{1}\text{-}Ca(DME)(OTf)]^{2+}$ forms a new species by ^{1}H NMR spectroscopy (Supplementary Fig. S2) that was identified by XRD as the tris(aqua) complex $[\mathbf{1}\text{-}Ca(OH_2)_3]^{3+}$ (Supplementary Fig. S26). Addition of DME to the CD_2Cl_2 solution of $[\mathbf{1}\text{-}Ca(OH_2)_3]^{3+}$ converts the complex back to $[\mathbf{1}\text{-}Ca(DME)(OTf)]^{2+}$ (^{1}H NMR spectroscopy, Supplementary Fig. S9). In the solid state, $[\mathbf{1}\text{-}Ca(DME)(OTf)]^{2+}$ is stable for weeks under ambient conditions.

Treatment of a DME suspension of $[\mathbf{1}\text{-}Ca(DME)(OTf)]^{2+}$ with the one-electron reductant decamethylferrocene yields the singly reduced product $[\mathbf{2}\text{-}Ca(DME)(OTf)]^{+}$ ($\mathbf{2}$ = $LMn^{III}_3O_2(OAc)_2$). The chemical reversibility of this conversion is evidenced upon treatment of $[\mathbf{2}\text{-}Ca(DME)(OTf)]^{+}$ with AgOTf to convert back to $[\mathbf{1}\text{-}Ca(DME)(OTf)]^{2+}$. ^{1}H NMR analysis is consistent with clean interconversion of these two clusters. A XRD study of $[\mathbf{2}\text{-}Ca(DME)(OTf)]^{+}$ reveals a cluster isostructural to $[\mathbf{1}\text{-}Ca(DME)(OTf)]^{2+}$ with only slight bond distance changes (vide infra), indicating little rearrangement of the cluster upon reduction.

The variable temperature magnetic susceptibilites of $[\mathbf{1}\text{-}Ca(DME)(OTf)][OTf]_2$ and $[\mathbf{2}\text{-}Ca(DME)(OTf)][OTf]$ were studied (Supplementary Fig. S11). For $[\mathbf{1}\text{-}Ca(DME)(OTf)][OTf]_2$, dominant ferromagnetic coupling between Mn ions is observed (see Supplementary Table S1 for fitting parameters). At 14 K, the $\chi_M T$ value increases to a maximum of 18.7 $cm^3\ mol^{-1}$ K, which is close to the expected spin-only value of a S = 11/2 system (17.9 $cm^3\ mol^{-1}$ K, g = 2). Different values would be expected for a more oxidized S = 5 system (15 $cm^3\ mol^{-1}$K, g = 2) or a more reduced S = 6 system (21 $cm^3\ mol^{-1}$K, g = 2). These results

support the oxidation state assignment of [**1**-Ca(DME)(OTf)]$^{2+}$ as $Mn^{IV}Mn^{III}_2$. The $\chi_M T$ value of [**2**-Ca(DME)(OTf)][OTf] approaches 10.2 cm^3 mol^{-1} K at 300 K, which is near the expected spin-only value of three uncoupled Mn^{III} ions (S = 2 spins, 3 cm^3 mol^{-1} K, g = 2). The $\chi_M T$ value decreases at low temperatures, reaching 2.6 cm^3 mol^{-1} K at 4 K, which shows a dominant antiferromagnetic interaction (Supplementary Table S1). To further confirm the oxidation state assignment of the isolated species, Mn XANES spectra were collected for [**1**-Ca(DME)(OTf)](OTf)$_2$ and [**2**-Ca(DME)(OTf)](OTf) (Fig. 4a). The rising edge energy, taken as a zero-crossing point of the 2nd derivative spectrum (Fig. 4b), is shifted to a higher energy by ca. 1.0 eV step from [**2**-Ca(DME)(OTf)](OTf) (6548.66 eV) to [**1**-Ca(DME)(OTf)](OTf)$_2$ (6549.76 eV). Such step-wiseedge shift is consistent with an one-electron oxidation state change in redox-active Mn complexes,[31] when the geometry and type of ligands are highly conserved. This result, correlated with charge balance in the solid state structure, magnetism, and chemical reactivity, therefore, supports the formal oxidation state assignments for [**1**-Ca(DME)(OTf)](OTf)$_2$ and [**2**-Ca(DME)(OTf)](OTf) as $Mn^{IV}Mn^{III}_2$ and Mn^{III}_3 respectively.

To study the effects of the redox-inactive center on the properties of the [Mn_3O_2] core, the analogous Sr^{2+}-, Y^{3+}-, Na^+-, and Zn^{2+}-capped trimanganese dioxo complexes were targeted (Figs 2 and 3). Treatment of $LMn^{II}_3(OAc)_3$ with PhIO and M(OTf)$_n$ (M = Na, Sr, Y), led to new species with ^{1}H NMR spectroscopic characteristics similar to compounds [**1**-Ca(DME)(OTf)]$^{2+}$ and [**2**-Ca(DME)(OTf)]$^+$. Complexes [**1**-Sr(DME)(OTf)]$^{2+}$ and [**2**-Y(DME)(OTf)]$^{2+}$ are structurally analogous to [**1**-Ca(DME)(OTf)]$^{2+}$ with the redox-inactive metal bridged by two oxido moieties to the trimanganese cluster and further coordinated by a DME molecule and a triflate anion (Fig. 3d, h). The yttrium-capped dioxo compound was isolated in the more reduced Mn^{III}_3 state rather than the $Mn^{IV}Mn^{III}_2$ state observed under the same reaction conditions for the calcium and strontium dioxo compounds. Similar to the reduction of [**1**-Ca(DME)(OTf)]$^{2+}$ to [**2**-Ca(DME)(OTf)]$^+$, the reduced strontium compound [**2**-Sr(DME)(OTf)]$^+$ was prepared by treating [**1**-Sr(DME)(OTf)]$^{2+}$ with decamethylferrocene. The Na-capped dioxo complex [**1**-Na]$_2^{4+}$ was isolated in the solid-state as a dimer of Mn_3NaO_2 moieties via acetate bridges. Each Mn_3NaO_2 core is structurally similar to the Ca^{2+}, Sr^{2+} and Y^{3+} analogs, with the Na^+ cation bridged by two oxidos to the trimanganese moiety (Fig. 3f). Preparation of the Zn-capped compound [**1**-Zn(CH_3CN)]$^{3+}$was accomplished by addition of Zn(OTf)$_2$ to a CH_3CN solution of [**1**-Ca(DME)(OTf)]$^{2+}$, leading to substitution of Ca^{2+} with Zn^{2+}. A XRD study of [**1**-Zn(CH_3CN)]$^{3+}$ grown from an acetonitrile/diethyl ether mixture shows one acetonitrile coordinated to the Zn^{2+} center. The Mn_3ZnO_2 core is structurally similar to the other Mn_3MO_2 moieties reported here. The magnetic susceptibilities of the $Mn^{IV}Mn^{III}_2$ complexes [**1**-Sr(DME)(OTf)][OTf]$_2$, [**1**-Na]$_2$[OTf]$_4$, and [**1**-Zn(CH_3CN)][OTf]$_3$ were measured, and all demonstrate ferromagnetic coupling of spins close to that of [**1**-Ca(DME)(OTf)][OTf]$_2$, supporting the same formal oxidation state assignment (Supplementary Fig. S11).

Analysis of the solid-state structures show that the distances from the redox-inactive metals to the briding oxido moieties vary as expected based on the ionic radius of the ion M, with the M–O distances increasing from [**1**-Zn(CH_3CN)]$^{3+}$ [2.008(3), 2.090(3) Å] to [**1**-Sr(DME)

(OTf)]$^{2+}$ [2.510(5), 2.651(4) Å] (Table 1). Upon reduction of the $Mn^{IV}Mn^{III}_2$ clusters [**1**-Ca(DME)(OTf)]$^{2+}$ and [**1**-Sr(DME)(OTf)]$^{2+}$ to form the Mn^{III}_3 complexes [**2**-Ca(DME)(OTf)]$^{+}$ and [**2**-Sr(DME)(OTf)]$^{+}$, the M–(μ_4-O) distances contract, indicating a stronger interaction with the redox-inactive metal as the interaction of the μ_4-oxido with the trimanganese core weakens due to increased population of σ-antibonding orbitals in d^4, Mn^{III} centers. As expected, the Mn–O distances increase upon reduction (compounds **2** vs **1**), but only small variations are observed based upon the redox-inactive metal. [**2**-Y(DME)(OTf)]$^{2+}$ displays the longest Mn-oxido average distance, consistent with the yttrium center, the most Lewis acidic metal of the series, drawing more electron density from the oxido moieties and weakening the Mn-O interactions. Notably, the bond distances of the Mn_3O_2 core are essentially the same between [**1**-Ca(DME)(OTf)]$^{2+}$ and [**1**-Sr(DME)(OTf)]$^{2+}$. These observations are consistent with XAS studies of Sr-substituted PSII indicating no significant structural change in the OEC from Ca-substituted PSII.[3]

We and others have reported studies that suggest a significant effect of redox-inactive metals on the electron transfer properties of metal-oxido species, phenomenon of particular relevance in biological systems and heterogeneous mixed metal oxides.[7,8,15,21,24,32–34] With a series of well-defined and structurally analogous Mn_3MO_2 complexes in hand, the effect of metal M on the reduction potential of the clusters was investigated. The cyclic voltammogram (CV) of a 10:1 CH_2Cl_2/DME solution of [**1**-Ca(DME)(OTf)]$^{2+}$ (0.1 M NBu_4PF_6) shows two quasireversible redox couples at −70 and −530 mV versus the ferrocene/ferrocenium couple (Fc/Fc^+) (Supplementary Fig. S15). Since [**1**-Ca(DME)(OTf)]$^{2+}$ is chemically reduced to [**2**-Ca(DME)(OTf)]$^{+}$ by the addition of one equivalent of decamethylferrocene (E^O ~ −0.48 V vs. Fc/Fc^+ in CH_2Cl_2), the couple centered at −70 mV is assigned as the [$Mn^{IV}Mn^{III}_2/Mn^{III}_3$] couple. A 10:1 CH_2Cl_2/DME solution of [**1**-Ca(DME)(OTf)]$^{2+}$ (0.05 M LiOTf electrolyte) was electrolyzed at a potential of −0.25 V vs. Ag/Ag^+ to form a new species by ^{1}H NMR spectroscopy. The amount of current passed at that potential supports the assignment of the wave at −80 mV as the one electron reduction of [**1**-Ca(DME)(OTf)]$^{2+}$. Independently prepared [**2**-Ca(DME)(OTf)]$^{+}$ reacts with excess LiOTf to cleanly form a product whose ^{1}H NMR spectrum matches that of the controlled potential electrolysis product described above (Supplementary Fig. S13). Under the same conditions, [**1**-Ca(DME)(OTf)]$^{2+}$ does not react with LiOTf. The more negative redox couple at −530 mV is presumed to correspond to the [$Mn^{III}_3/Mn^{III}_2Mn^{II}$] couple, but the more reduced product has not yet been isolated.

Since the electrochemical studies are presumed to be of intact clusters, the solution stability of the [MMn_3O_2] core was studied. Isotopically labelled $LCaMn_3{}^{18}O_2(OAc)_2(DME)(OTf)_3$ ([**1***-Ca(DME)(OTf)][OTf]$_2$) was prepared and mixed with one equivalent of natural abundance [**1**-Ca(DME)(OTf)]$^{2+}$ in dichloromethane at room temperature, and the mixture was analyzed over time using electrospray ionization mass spectrometry (ESI-MS). Less than 20% isotopic scrambling occurs on the timescale of the electrochemical experiments (ca. 18% after 1 h at RT). Additionally, methylene chloride solutions of [**1**-Ca(DME)(OTf)]$^{2+}$ are stable at room temperature under anhydrous conditions for days (^{1}H NMR spectroscopy). The addition of 10 equivalents of $Ca(OTf)_2$ to the solution of [**1**-Ca(DME)(OTf)]$^{2+}$ does not change the reduction potential of the complex (Supplementary Fig. S17)

indicating that any equilibrium toward dissociation of the clusters lies toward Ca^{2+} association. To interrogate the effect of the Ca-coordinated ligands, under the same conditions, the reduction of [**1**-Ca$(OH_2)_3$]$^{3+}$ occurs within 30 mV of that of [**1**-Ca(DME)(OTf)]$^{2+}$ (Supplementary Fig. S16) suggesting that the capping ligands on the Ca^{2+} center do not significantly affect the reduction potentials of the clusters.

CVs of [**1**-Sr(DME)(OTf)]$^{2+}$, [**2**-Y(DME)(OTf)]$^{2+}$, [**1**-Zn(CH_3CN)]$^{3+}$, and [**1**-Na]$_2^{4+}$ all display the [$Mn^{IV}Mn^{III}_2/Mn^{III}_3$] couple observed for [**1**-Ca(DME)(OTf)]$^{2+}$**,** although at different potentials (Fig. 5a). Although the peak-to-peak separations of the observed couples are large when measured at a glassy carbon electrode (ΔE_p ~ 400 mV), indicating slow electron transfer to the complexes, the $E_{1/2}$ values calculated from the CVs of [**1**-Ca(DME)(OTf)]$^{2+}$, [**1**-Sr(DME)(OTf)]$^{2+}$, and [**1**-Na]$_2^{4+}$ are in good agreement with CVs collected using a hanging drop mercury electrode, with ΔE_p ~ 200 mV (Supplementary Fig. S14). The $E_{1/2}$ corresponding to the reduction of [**1**-Ca(DME)(OTf)]$^{2+}$ to [**2**-Ca(DME)(OTf)]$^{+}$ was also measured by monitoring the electronic absorption spectrum upon titration with dimethylferrocene (Supplementary Fig. S18), and the calculated values (−0.1 V vs. Fc/Fc^+) are close to the value from the CV (−0.08 V vs. Fc/Fc^+).

As structural characterization was obtained only for clusters displaying the $Mn^{IV}Mn^{III}_2$ and Mn^{III}_3 oxidation states, these are the reduction potentials that are compared below. Moreover, the more highly oxidized species are pertinent to the moieties present in the OEC and proposed for water oxidation catalysts. The $E_{1/2}$ values of the [$Mn^{IV}Mn^{III}_2/Mn^{III}_3$] couple become more positive as the charge of the redox-inactive metal increases. This trend suggests that the increased charge of the proximal redox-inactive cation facilitates reduction of the manganese centers. The $E_{1/2}$ value for the Zn^{2+} complex appears at potentials ca. 230 mV more positive compared to the Ca^{2+} and Sr^{2+} species. Although [**1**-Zn(CH_3CN)]$^{3+}$ is tricationic while [**1**-Ca(DME)(OTf)]$^{2+}$ and [**1**-Sr(DME)(OTf)]$^{+}$ are dicationic, the $E_{1/2}$ of the [$Mn^{IV}Mn^{III}_2/Mn^{III}_3$] couple of [**1**-Ca$(OH_2)$]$^{3+}$, also a tricationic complex, is more positive than that of [**1**-Ca(DME)(OTf)]$^{2+}$ by less than 30 mV. The more positive potential of the Zn^{2+} compound is inconsistent with a purely electrostatic explanation of the change in reduction potentials, as proposed for oxo-bridged manganese dimers with alkali and alkali earth metal ions associated via salen ligands modified with crown ether moieties.[33]

The effect of redox-inactive metals on the kinetics of electron transfer to a non-heme $Fe^{IV}O$ species has been previously linked to the Lewis acidity of the metal.[8,35] The $E_{1/2}$ values of the [$Mn^{IV}Mn^{III}_2/Mn^{III}_3$] couples measured above in organic solvents were plotted against the pK_a of the metal aqua ions, M(aqua)$^{n+}$ in water,[36] used here as a measure of the Lewis acidity of the metal M. A remarkable linear correlation is observed (Fig. 5b) that clearly links the effect of the redox-inactive metal to the cluster reduction potential in terms of the metal's Lewis acidity. The slope provides a quantitative measure of this correlation, with each pK_a unit shifting the potential by ca. 100 mV. This effect is likely a consequence of the interaction between the oxido moieties and redox-inactive metals vs. manganese centers. The stronger Lewis acid is expected to draw more electron density from the oxido ligands and to destabilize the higher oxidation state manganese centers.

The above findings have implications for biological and heterogeneous metal catalysts for water oxidation and other redox processes. Catalysis occurs at discrete multinuclear sites, consisting of five metal centers for the OEC in PSII and less than ten metal centers for heterogeneous manganese and cobalt oxides. The potential of the cluster is expected to vary based on structure and number of oxido ligands. For example, the reduction potential of the[$Mn^{IV}Mn^{III}_2CaO_2/Mn^{III}_3CaO_2$] couple reported here is more positive than that of the [$Mn^{IV}_3CaO_4/Mn^{IV}_2Mn^{III}CaO_4$] couple of the $CaMn_3$ cubane cluster,[21] despite the higher manganese oxidation states in the latter complex, likely due to the greater number of oxido ligands. Changing the nature of the redox-inactive metal component of the cluster, *without* a structural change of the cluster, allows for drastic variation of the reduction potential in both directions, potentially from values that render the chemistry of interest (e.g. water oxidation) thermodynamically unfavorable to values that make it favorable. Tuning of the reduction potentials of mixed metal-oxide cluster by Lewis acids is an appealing strategy for designing practical catalysts for water splitting.

It is notable in the present series that the Ca^{2+} and Sr^{2+} variants have essentially the same potentials, which is consistent with the observation that substitution of Sr^{2+} for Ca^{2+} in PSII retains reactivity and a similar electronic structure.[37] Although a previously proposed role of Ca^{2+} (or Sr^{2+}) of the OEC in facilitating the attack of a water or hydroxide moiety on an electrophilic manganese oxo is still possible,[38] the reported results support a significant role in redox tuning of the cluster.

In summary, a series of tetrametallic dioxo complexes containing redox-inactive metal ions in the +1 to +3 oxidation states was synthesized by employing a trinucleating ligand framework. The structural characteristics of these complexes, with the oxido ligands bridging the redox-inactive metals and the manganese centers, make them particularly relevant to biological and heterogeneous metal-oxido clusters. Electrochemical studies of these compounds show that the reduction potentials are highly dependent upon the Lewis acidity of the redox-inactive metal, identifying the chemical basis for the observed differences in electrochemistry. This correlation provides evidence for the role of the Ca^{2+} ion in modulating the redox potential of the OEC and of other redox-inactive ions in tuning the redox potentials of other metal oxide electrocatalysts. The observed linear dependence between cluster potential and Lewis acidity provides a rational strategy for tuning the redox properties of heterometallic metal-oxido clusters of interest for catalysis. Current studies are focused on further understanding the relationship between the structures of mixed metal manganese-oxido clusters and their reactivity.

Methods

Unless indicated otherwise, reactions were carried out in oven-dried glassware in a glovebox under a nitrogen atmosphere. Anhydrous tetrahydrofuran (THF) was purchased from Aldrich in 18 L Pure-Pac™ containers. Anhydrous dichloromethane, diethyl ether, and THF were purified by sparging with nitrogen for 15 minutes and then passing under nitrogen pressure through a column of activated A2 alumina. Anhydrous 1,2-dimethoxyethane (DME) was dried over sodium/benzophenone ketyl and vacuum-transferred onto molecular sieves. CD_2Cl_2 was purchased from Cambridge Isotope Laboratories, dried over calcium

hydride, then degassed by three freeze-pump-thaw cycles and vacuum-transferred prior to use.

Iodosobenzene was prepared according to literature procedures.[39] $LMn_3(OAc)_3$ and$[LMn_3O(OAc)_3]_2Ca(OTf)_2$ were prepared according to previously published procedures.[21,23] **Caution!** Iodosobenzene is potentially explosive and should be used only in small quantities.

The syntheses of all reported compounds, characterization methods and equipment are described in the Supplementary Information.

Synthesis of [1-Ca(DME)(OTf)][OTf]$_2$

In the glovebox, a round-bottom flask equipped with a stir bar was charged with $LMn_3(OAc)_3$ (2.0 g, 1.67 mmol) and $Ca(OTf)_2$ (0.90 g, 2.67 mmol, 1.6 equiv). DME (200 mL) was added, and the yellow suspension was stirred at room temperature for 5 min. Iodosobenzene (0.81 g, 3.68 mmol, 2.2 equiv) was added as a solid, and the mixture was stirred at room temperature for 4 h, turning from yellow to purple. The purple solid was collected via filtration, washed with DME, then extracted with dichloromethane. The red-purple solution was concentrated *in vacuo* to yield the product as a red-purple solid (2.45 g, 84%). ^{1}H NMR (CD_2Cl_2, 300 MHz): δ 77.8, 76.5, 72.3, 69.6, 60.5, 53.7, 48.4, 38.6, 37.0, 24.2, 19.9, 17.5, 15.9, 8.1, 3.3, 2.9, −19.3, −23.5, −24.9, −26.4, −29.1 ppm. ^{19}F NMR (CD_2Cl_2): δ−74.4 ppm. UV-Vis (CH_2Cl_2, λ_{max} (ε)): 498 (1410 M^{-1} cm^{-1}), 846 (640 M^{-1} cm^{-1}) nm. Anal. Calcd. For $C_{68}H_{55}CaF_9Mn_3N_6O_{20}S_3$: C, 46.72; H, 3.17; N, 4.81. Found: C, 46.92; H, 3.27; N, 4.89.

Synthesis of [2-Ca(DME)(OTf)][OTf]

In the glovebox, a round-bottom flask equipped with a stir bar was charged with [**1**-Ca(DME)(OTf)]$^{2+}$ (0.750 g, 0.429 mmol) and decamethylferrocene (0.140 g, 0.429 mmol, 1 equiv). DME (30 mL) was added, and the purple mixture was stirred at room temperature over 1 h. The gray-purple precipitate was collected on a fritted glass funnel and washed with DME, then extracted with cold THF (40 mL). The purple filtrate was concentrated to ca. 20 mL *in vacuo*, then cooled to −35 °C to precipitate out more $Cp^*_2Fe^+$, which was filtered off over Celite. The purple filtrate was concentrated *in vacuo* to a purple solid, then recrystallized from DME/CH_2Cl_2/hexanes to yield the product as a purple solid (0.405 g, 59%). ^{1}H NMR (CD_2Cl_2, 300 MHz): δ 65.1, 57.8, 50.7, 41.6, 35.3, 29.6, 15.7, 15.0, 14.2, 8.2, 3.0, −9.6, −12.6, −17.0, −17.9 ppm. ^{19}F NMR (CD_2Cl_2): δ−74.6 ppm. UV-Vis (CH_2Cl_2, λ_{max} (ε)): 495 (710 M^{-1} cm^{-1}), 860 (310 M^{-1} cm^{-1}) nm. Anal. Calcd. for $C_{67}H_{55}CaF_6Mn_3N_6O_{17}S_2$: C, 50.32; H, 3.47; N, 5.26. Found: C, 50.04; H, 3.63; N, 5.06.

Electrochemical measurements were recorded using a Pine Instrument Company AFCBP1 bipotentiostat using the AfterMath software package. Cyclic voltammograms were recorded on ca. 1 mM solutions of the relevant complexes in the glovebox at 20 °C with an auxiliary Pt-coil electrode, a Ag/Ag+ reference electrode (0.01 M $AgNO_3$, 0.1 M ^{n}Bu_4NPF_6 in CH_3CN), and a 3.0 mm glassy carbon electrode disc (BASI). The electrolyte solutions were 0.1 M ^{n}Bu_4NPF_6 in CH_2Cl_2 ([**1**-Na]$_2^{4+}$) or 10:1 CH_2Cl_2/DME ([**1**-Ca(DME)(OTf)]$^{2+}$, [**1**-

Sr(DME)(OTf)]$^{2+}$, [**1**-Zn(CH_3CN)]$^{3+}$, and [**2**-Y(DME)(OTf)]$^{2+}$). For [**1**-Ca(DME)(OTf)]$^{2+}$, [**1**-Sr(DME)(OTf)]$^{2+}$, and [**1**-Na]$_2^{4+}$, CVs were also recorded in an electrolyte solution of 0.1 M ^{n}Bu$_4$NPF$_6$ in CH_2Cl_2 using a hanging mercury drop electrode (BASI CGME). Using the average mass of the mercury drop (0.037 g), the surface area of the drop (assumed to be a sphere) was calculated to be 0.095 cm^2. All reported values are referenced to an internal ferrocene/ferrocenium couple.

Supplementary Material

Refer to Web version on PubMed Central for supplementary material.

Acknowledgments

This work was supported by the California Institute of Technology, the Searle Scholars Program, the NSF CAREER CHE-1151918 (T.A.), and the NSF Graduate Research Fellowship Program (E.Y.T.). We thank Lawrence M. Henling and David E. Herbert for assistance with crystallography, and Po-Heng Lin for assistance with magnetic susceptibility studies. The Bruker KAPPA APEXII X-ray diffractometer was purchased via an NSF Chemistry Research Instrumentation award to Caltech (CHE-0639094). The X-ray spectroscopy work was supported by NIH Grant F32GM100595 (R.T.), and by the Director of the Office of Basic Energy Science, OBES, Division of Chemical Sciences, Geosciences, and Biosciences, DOE, under Contract DE-AC02-05CH11231 (J.Y.). Synchrotron facilities were provided by the Stanford Synchrotron Radiation Lightsource (SSRL) operated by the DOE, OBES.

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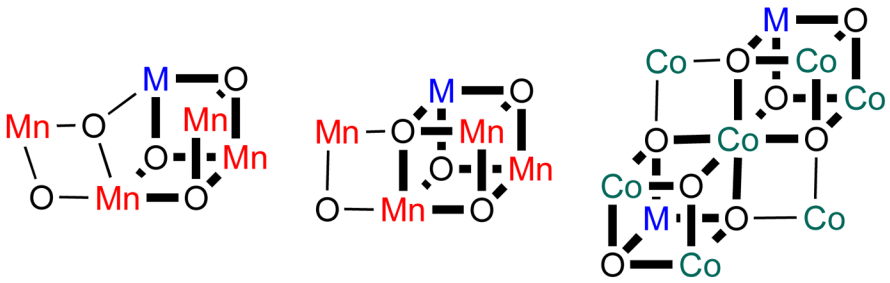

Figure 1.

Proposed structures of water oxidation catalysts containing redox-inactive metals (M) in the OEC (left, middle) and in heterogeneous cobalt oxide water oxidation catalysts (right).[12] The OEC is known to contain a Mn_3M core: one major model is based on EXAFS and electron paramagnetic resonance (EPR) studies (left)[41–43] and one on X-ray crystallography (middle).[18,19] Bold bonds emphasize the Mn_3M and Co_3M cluster cores.

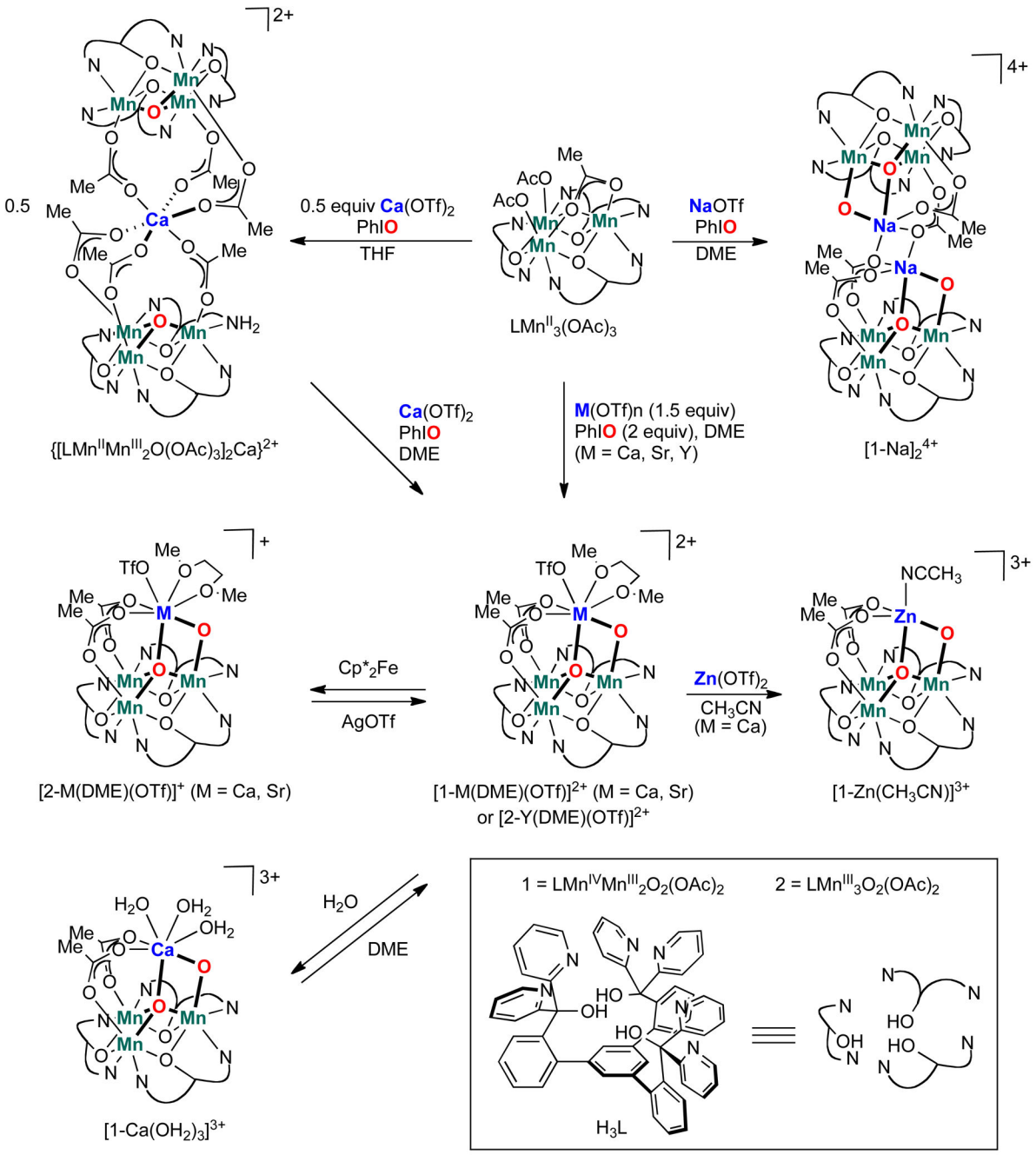

Figure 2.

Synthesis of tetrametallic trimanganese dioxido complexes. Reduced precursor $LMn_3(OAc)_3$ was oxidized in the presence of metal triflate salts to form $[LMn_3MO_2(OAc)_2]$ compounds (M = Na, Ca, Sr, Y). Complex [**1**-Ca(DME)(OTf)]$^{2+}$ (**1** = $LMn^{IV}Mn^{III}{}_2O_2(OAc)_2$) undergoes reversible chemical reduction, substitution of Zn^{2+} for Ca^{2+}, or exchange of the DME ligand for coordinated water molecules.

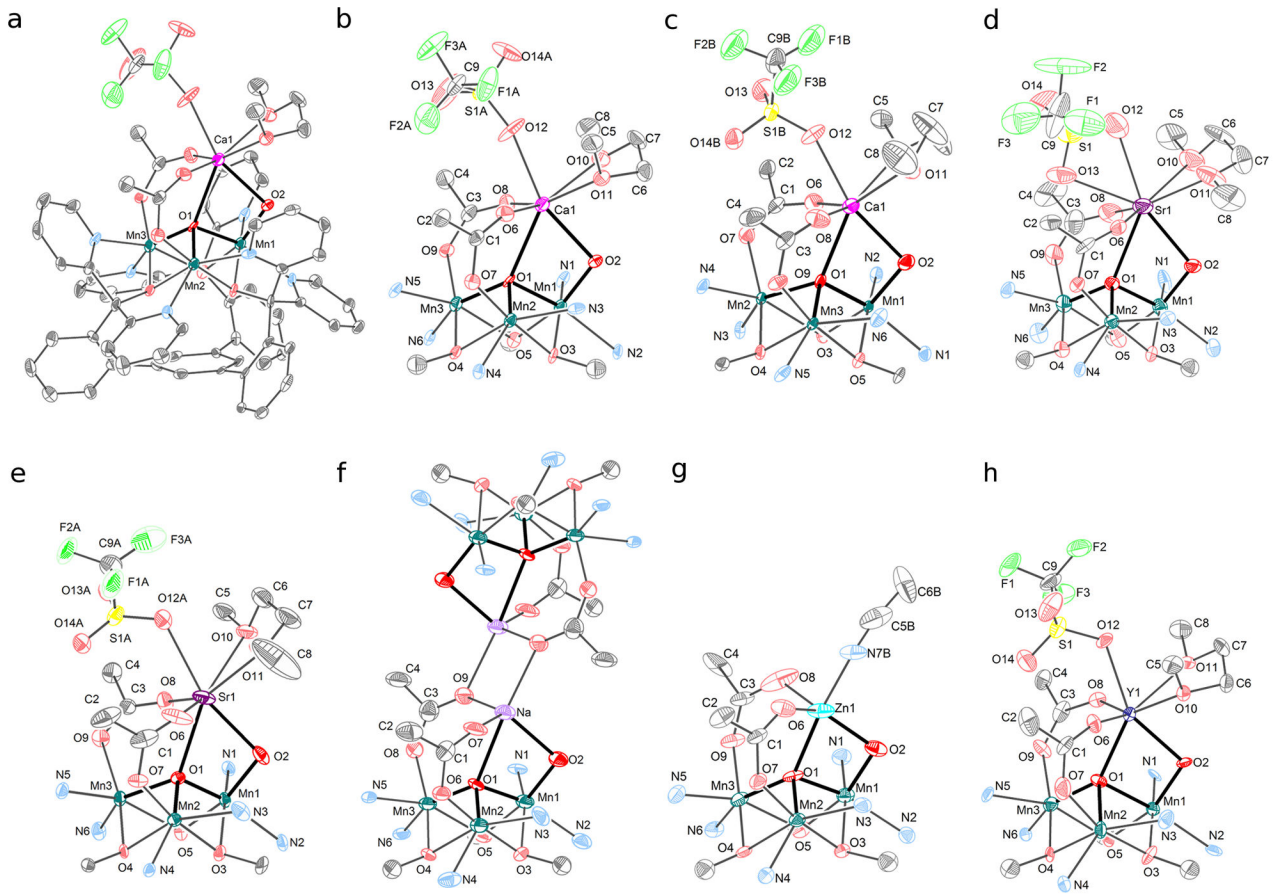

Figure 3.

Solid-state structures of reported complexes (thermal ellipsoids shown at 50% level). Hydrogen atoms and outer-sphere anions not shown for clarity. (a) Full structure of [**1**-Ca(DME)(OTf)][OTf]$_2$. (b) Truncated view of [**1**-Ca(DME)(OTf)][OTf]$_2$. (c) Truncated view of [**2**-Ca(DME)(OTf)][OTf]. (d) Truncated view of [**1**-Sr(DME)(OTf)][OTf]$_2$. (e) Truncated view of [**2**-Sr(DME)(OTf)][OTf]. (f) Truncated view of [**1**-Na]$_2$[OTf]$_4$. (g) Truncated view of [**1**-Zn(CH$_3$CN)][OTf]$_3$. (h) Truncated view of [**2**-Y(DME)(OTf)][OTf]$_2$.

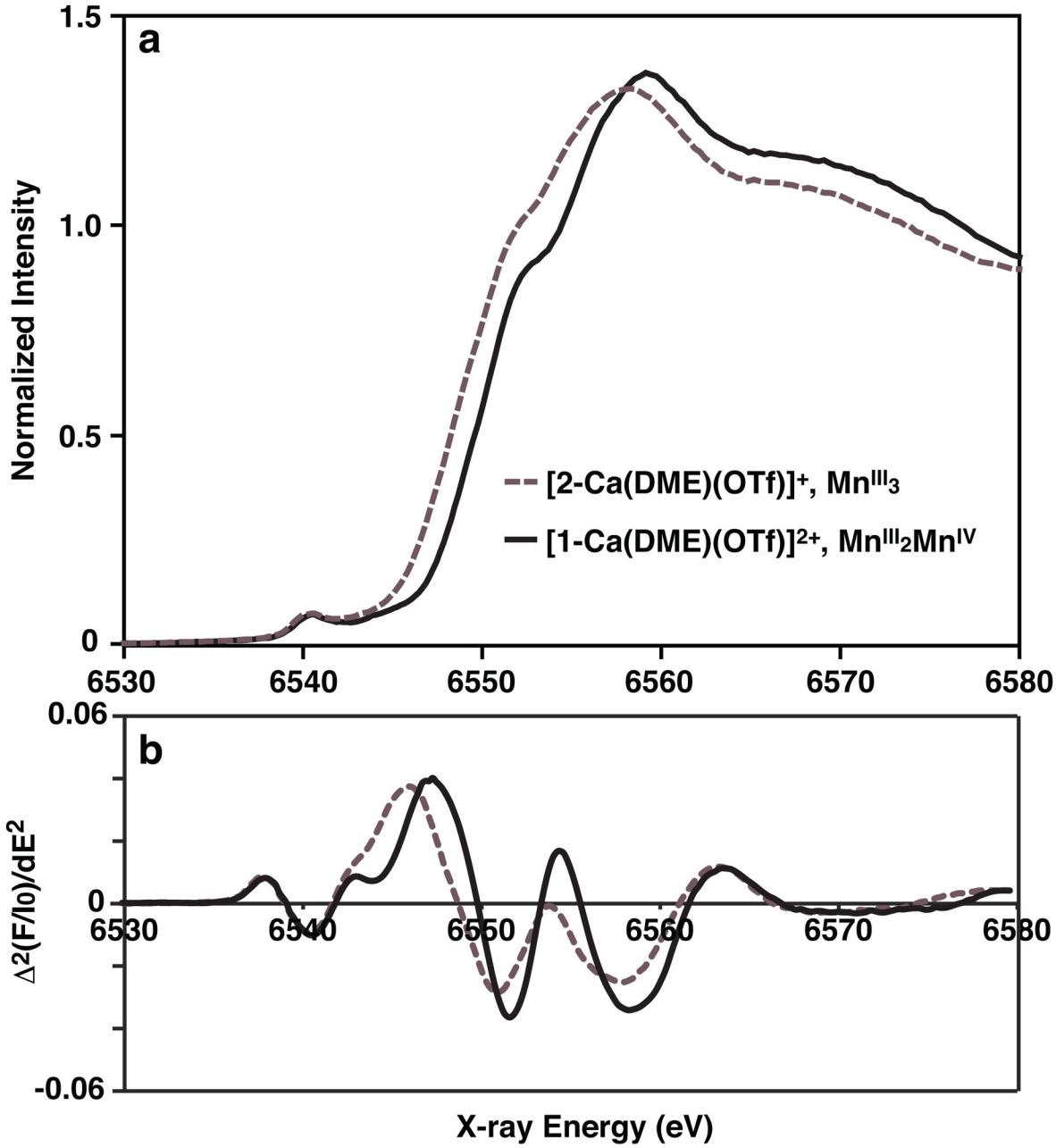

Figure 4.

(a) Mn XANES spectra and (b) second derivative spectra of [**2**-Ca(DME)(OTf)][OTf] (dotted) and [**1**-Ca(DME)(OTf)][OTf]$_2$ (solid). The shift in the rising edge energy of the two complexes (6548.66 and 6549.76 eV respectively) suggests a one-electron oxidation state change.

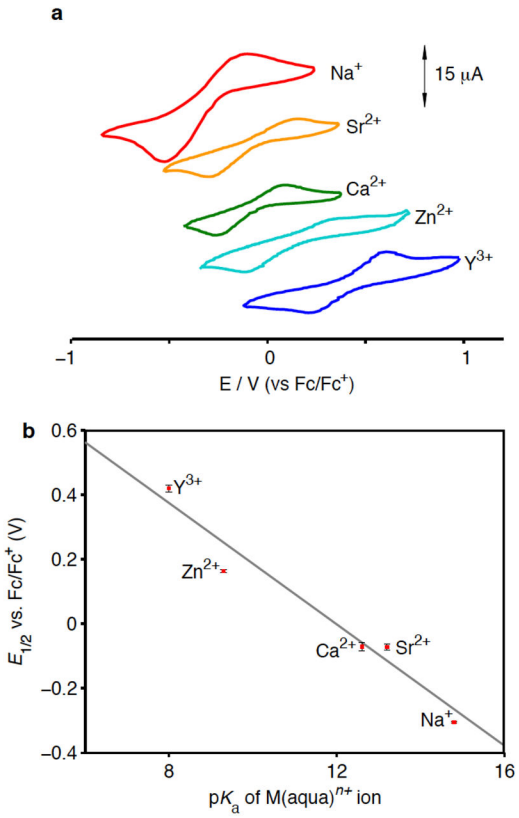

Figure 5.

The redox potentials of the $[MMn_3O_2]$ complexes are correlated with the Lewis acidity of the redox-inactive metal. (a) Cyclic voltammograms of reported complexes in 0.1 M NBu_4PF_6 10:1 CH_2Cl_2/DME (CV of $[\mathbf{1}\text{-Na}]_2^{4+}$ in CH_2Cl_2) using a glassy carbon disc electrode at a scan rate of 100 mV/s. (b) Dependence of $E_{1/2}$ of $Mn^{IV}Mn^{III}{}_2/Mn^{III}{}_3$ couple on pK_a of M(aqua)$^{n+}$ ion.[36] Error bars correspond to the standard deviation of the reduction potentials measured from three independent samples.

Table 1

Selected bond lengths (Å) for reported complexes. The Mn–O bond lengths vary depending on manganese oxidation state, but not significantly due to the redox-inactive metal M. M–O bond lengths vary with ionic radius. All e.s.d.s were calculated rigorously from the full covariance matrix.[40]

	[1-Ca(DME)(OTf)]$^{2+}$	**[2-Ca(DME)(OTf)]$^{+}$**	**[1-Sr(DME)(OTf)]$^{2+}$**	**[2-Sr(DME)(OTf)]$^{+}$**	**[1-Na]$_2^{4+}$**	**[1-Zn(CH$_3$CN)]$^{3+}$**	**[2-Y(DME)(OTf)]$^{2+}$**
M–Mn1	3.317(1)	3.283(1)	3.476(1)	3.424(1)	3.216(3)	3.0005(8)	3.3011(6)
M–Mn2	3.749(1)	3.802(1)	4.005(1)	3.923(1)	3.739(3)	3.3495(9)	3.7533(6)
M–Mn3	4.042(1)	4.034(1)	4.149(1)	4.166(1)	3.942(3)	3.7595(9)	3.8592(6)
Mn1–Mn2	3.0480(9)	3.0111(9)	3.062(1)	3.020(2)	3.076(2)	3.065(1)	3.2110(7)
Mn1–Mn3	3.0486(9)	3.1537(9)	3.051(1)	3.142(2)	3.000(2)	3.038(1)	3.0494(7)
Mn2–Mn3	3.0179(9)	3.0541(9)	3.025(1)	3.064(2)	3.004(2)	3.0739(9)	3.1223(7)
M–O2	2.349(3)	2.368(3)	2.510(5)	2.508(5)	2.300(6)	2.008(3)	2.269(2)
M–O1	2.452(3)	2.397(3)	2.651(4)	2.536(5)	2.422(6)	2.090(3)	2.232(2)
Mn1–O2	1.842(3)	1.887(3)	1.841(4)	1.869(5)	1.840(5)	1.862(3)	1.889(2)
Mn1–O1	2.017(3)	1.939(3)	2.022(4)	1.943(5)	1.995(5)	1.981(3)	2.150(2)
Mn2–O1	1.913(3)	2.159(3)	1.932(4)	1.856(5)	1.936(5)	1.946(3)	2.177(2)
Mn3–O1	1.958(3)	1.860(3)	1.936(4)	2.172(5)	1.889(5)	1.960(3)	1.895(2)
Avg. Mn– O1	1.96	1.99	1.96	1.99	1.94	1.96	2.07