

Marquette University
e-Publications@Marquette

Chemistry Faculty Research and Publications

Chemistry, Department of

7-1-2007

The Voltammetric Study of the Reduction of Tetraalkylammonium Perchlorate by $\text{Fe}(\text{TPP})^{2-}$

Chathra Shirani De Silva
Marquette University

Michael D. Ryan
Marquette University, michael.ryan@marquette.edu

Accepted version. *Journal of Porphyrins and Phthalocyanines*, Vol. 11, No. 7 (July 2007): 519-523.

DOI. © 2007 Wiley. Used with permission.

This is the peer reviewed version. This article may be used for non-commercial purposes in accordance with Wiley Terms and Conditions for self-archiving.

The Voltammetric Study of the Reduction of Tetraalkylammonium Perchlorate by $\text{Fe}(\text{TPP})^{2-}$

Michael D. Ryan
*Chemistry, Marquette University
Milwaukee, WI*

Chathra DeSilva
*Chemistry, Marquette University
Milwaukee, WI*

Abstract: Tetraalkylammonium ions react with FeTPP^{2-} to form $\text{Fe}(\text{TPP})(\text{R})^-$ and trialkylamine. The tetrabutylammonium cation was verified to be the source of the alkyl group in the product, $(\text{TPP})(\text{R})^-$, by using $(^1\text{H}_5\text{C}_2)_3(^2\text{H}_5\text{C}_2)\text{N}^+$ as the cation and ^2H NMR. The reaction of FeTPP^{2-} with Bu_4N^+ was monitored by cyclic voltammetry and thin layer spectroelectrochemistry. The activation parameters were measured, and were most consistent with an electron transfer (ET) mechanism. The rate of the reaction of tetramethyl and tetraethylammonium ions with $\text{Fe}(\text{TPP})^{2-}$ was also examined. The rate constant decreased significantly as the carbon chain length decreased, which was also consistent with an ET mechanism.

Introduction

Tetraalkylammonium ions are commonly used supporting electrolytes in voltammetric studies. They are generally considered to be resistant to reduction, but a few reports have indicated their reactivity [1,2,3]. The initial work by DeSilva et al. in our laboratory has shown that

iron(0) porphyrins are capable of reducing tetraalkylammonium ions while forming iron(II) porphyrin alkyls [1]. There have been a number of reports on the electrochemical or photochemical reduction of tetraalkylammonium salts. Aurbach and Gottlieb found that the direct reduction of the tetrabutylammonium ion (TBA⁺) at a gold electrode lead to the reductive cleavage of the C-N bond to form tributylamine and butyl radical [2]. Later Dahm and Peters studied the electrochemical reduction of tetraalkylammonium salts at a carbon electrode and found similar results [4]. Kunkely and Vogler studied the photochemical reaction of tetrabutylammonium tetrabutylborate and found that octane was the major hydrocarbon product [5].

Iron(0) porphyrins are quite reactive species which can react by either nucleophilic substitution (S_N2) or by single electron transfer (ET) processes. Iron(I) and iron(0) porphyrins have been shown to react with a variety of substrates including alkyl halides [6,7], carboxylic anhydrides [8], carbon dioxide [9,10,11] and weak acids to generate H₂ [12]. The reaction between iron(I) or iron(0) porphyrins with alkyl halides has been studied most extensively in order to investigate the competition between S_N2 and ET mechanisms [13,14]. Both processes were observed depending upon the temperature and structures of the substrates. These previous reports have emphasized the nucleophilic character of iron(0) porphyrins.

The dealkylation of tetraalkylammonium salts has been examined using halide ions are the nucleophile. The reactivity trend in molten salts and in solution was consistent with nucleophilic attack on tetraalkylammonium ions [15,16]. As reported previously, Fe(TPP)²⁻ has been similarly shown to dealkylated the tetraalkylammonium ion. Steric would be expected to favor the ET over the S_N2 mechanism. But, for small cations such as tetramethylammonium ion (TMA⁺) or tetraethylammonium (TEA⁺), one might expect a change to an S_N2 mechanism. Following up on our initial report on the formation of the iron porphyrin alkyls from iron(0) porphyrins, work continued in our laboratory to confirm that the reaction of Fe(TPP)²⁻ was indeed an ET reaction, and to determine whether smaller tetraalkylammonium cations led to a change in mechanism. This study would help to understand the nucleophilic properties of iron(0) porphyrins.

Experimental

Chemicals

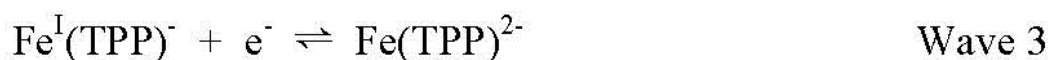
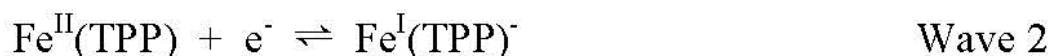
Chlorin-free iron(III) tetraphenylporphyrin chloride was purchased from MidCentury Chemicals, and was used as received. Dimethylformamide (DMF) (Aldrich Chemical Co.) was purified by heating at reflux temperatures over calcium hydride, followed by distillation under reduced pressure. Care has been taken to avoid exposure to the atmosphere by flushing the receiving containers with argon gas. All solvents were stored in an argon atmosphere. Tetramethylammonium (TMAP), tetraethylammonium (TEAP) and tetrabutylammonium (TBAP) perchlorate was obtained from GFS Chemical Co. Partially deuterated TBAP (d_9 -TBAP, Isotec Inc, 99.9% ^2H) was synthesized using the method of Missan et al.[17] with d_9 -butyl bromide replacing natural abundance butyl bromide. The crude product was recrystallized from anhydrous ethyl acetate[18]. Deuteration was then verified by ^2H NMR.

Instrumentation

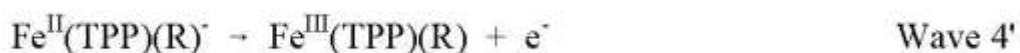
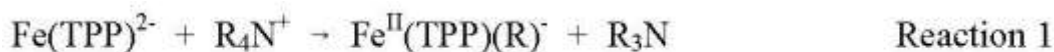
The cyclic voltammetric data were obtained with a Cypress System CYSY-1R potentiostat (Version 6.0 software). A three electrode cell was used for voltammetric measurements, consisting of a platinum wire or flag working electrode, a platinum flag auxiliary electrode, and a Ag/0.1 M AgNO₃ in acetonitrile reference electrode. The reference electrode was separated from the electrochemical test solution by a salt bridge filled with the appropriate solvent and supporting electrolyte. Digital simulations were carried out using Digisim 3 (Bioanalytical Systems). In the simulations, the diffusion coefficients of all species were assumed to be equal. The UV-visible spectra were recorded on a Hewlett-Packard 8452A diode array spectrophotometer. Controlled potential electrolysis were performed with an Electrosynthesis Co. Model 410 potentiostatic controller with an EG&G PARC model 379 digital coulometer. An OTTLE cell as described by Lin and Kadish was used for the spectroelectrochemical data [19].

Results and Discussion

The electrochemical reduction of Fe(TPP)Cl has been extensively studied and occurs in three reversible waves:



The product of Wave 3 is formally an iron("0") porphyrin but has considerable iron(I) porphyrin radical anion character. If the working electrode potential is poised for 3 min at a potential beyond the 3rd wave at room temperature, the appearance of new waves due to the alkylated iron porphyrin can be readily observed [1]. The following reaction was showed to have occurred:



where the prime waves indicate oxidation peaks, the unprimed waves indicate reduction peaks and *k* is the rate constant for Reaction 1. At elevated temperatures (above 47°C), Wave 4' can be readily observed in the cyclic scan (Figure 1B). Resonance Raman, visible spectroscopy and voltammetry have all shown that an alkyl iron porphyrin was formed [1]. No alkyl porphyrins were observed when the potential was held between Waves 2 and 3 (formation of Fe(TPP)⁻). In order to

exclude the solvent as the source of the alkyl group and to positively verify TBA⁺ as the source of the alkyl group, deuterated TBAP was synthesized (d₉-TBAP). For synthetic simplicity, only one of the alkyl groups was deuterated, as the d₉-TBAP was formed from the reaction of natural abundance tributyl amine and d₉-butyl bromide. Electrolysis of 4.0 mM FeTPP⁺Cl⁻ in THF and 0.050 M d₉-TBAP was carried out at -2.3 V vs Ag/AgClO₄ reference electrode. The reduced solution was characterized by ²H NMR. The results are summarized in Table 1, along with literature values as obtained by proton NMR. The results conclusively show that the supporting electrolyte is the source of the alkyl group.

Alkylation of iron porphyrins can occur by ET or S_N2 processes. These processes can be distinguished by the determination of the activation parameters for Reaction 1. The products of the first two waves (Fe^{II}(TPP), Fe^I(TPP)⁻) are stable on the cyclic voltammetric time scale. The rate of Reaction 1 can be determined from the peak current ratio for the third wave (i_{pr}/i_{pf} , where i_{pr} = current due to the reverse wave corrected for the Cottrell decay of the forward wave, and i_{pf} = current due to the forward wave corrected for the Cottrell decay of the 2nd wave). A typical comparison of the experimental and simulated cyclic voltammograms are shown in Figure 1 for $k = 2.0 \text{ M}^{-1}\text{s}^{-1}$, using a pseudo-first order mechanism with respect to TBA⁺. The decrease in current for Wave 3' can be seen in both the simulated and experimental cyclic voltammograms, as well as the new peak for Wave 4' due to the oxidation of Fe(TPP)(Bu)⁻ (the poor fit to the reverse peak of Wave 1 is due to ligand kinetics which is not of relevance to this work). The peak current ratios for five temperatures are summarized in Table 2. Using the peak current ratios, the rate constants can be calculated [20] and the results are summarized in Figure 2.

The rate constant at room temperature was determined using spectroelectrochemistry as the reaction was too slow to be observed using cyclic voltammetry. The electrode potential was poised at -2.3 V, and the absorbance for Fe(TPP)²⁻ at 460 nm was monitored as a function of time in a thin-layer spectroelectrochemical cell. Linear pseudo-first order reaction kinetics were observed and a rate constant of $0.13 \text{ M}^{-1}\text{s}^{-1}$ was found. The enthalpy and entropy of activation were

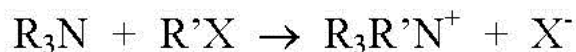
obtained from the Arrhenius plot, and the entropy of activation was found using the following equation [13]:

$$\ln k = \ln Z - (\Delta H^\ddagger/RT) + (\Delta S^\ddagger/R)$$

where Z is the collision frequency and estimated to be $3 \times 10^{11} \text{ M}^{-1} \text{ s}^{-1}$ [13]. Using this equation and the Arrhenius plot, the $\Delta H^\ddagger = 15 \pm 1 \text{ kcal}$ and $\Delta S^\ddagger = -7 \pm 2 \text{ cal/K}$.

The reaction of $\text{Fe}(\text{TPP})^{2-}$ with tetraethylammonium (TEA^+) and tetramethylammonium (TMA^+) ions was also examined. Slow scan rate cyclic voltammograms of $\text{Fe}(\text{TPP})\text{Cl}$ in DMF with 0.10 M tetraalkylammonium perchlorate are shown in Figure 3. From the cyclic voltammograms, it is clear that the reaction of TEA^+ with $\text{Fe}(\text{TPP})^{2-}$ is significantly slower than the reaction with TBA^+ , and an even slower reaction (if at all) could be observed with TMA^+ . In fact, in the presence of TMA^+ , the $\text{Fe}(\text{I})/\text{Fe}(\text{0})$ was nearly chemical reversible indicating either no reaction or a slow reaction between $\text{Fe}(\text{TPP})^{2-}$ and TMA^+ . Using thin-layer spectroelectrochemistry, the rate of the reaction between the tetraalkylammonium ion and $\text{Fe}(\text{TPP})^{2-}$ at room temperature was measured. The rate constant for TEA^+ with $\text{Fe}(\text{TPP})^{2-}$ was found to be $1.15 \pm 0.02 \times 10^{-2} \text{ M}^{-1} \text{ s}^{-1}$, compared to $0.13 \text{ M}^{-1} \text{ s}^{-1}$ for TBA^+ under the same conditions. With TMA^+ , the rate constant decreased to $7.2 \pm 0.6 \times 10^{-3} \text{ M}^{-1} \text{ s}^{-1}$.

In Table 3 the activation parameters for the reaction of low valent iron porphyrins with substrate are summarized. The previously studied reactions of iron porphyrins with alkyl halides are more similar to the Menshutkin type reactions in which alkyl amines react with alkyl halides to form ionic products.

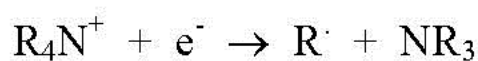


these cases, the origin of the large negative values of ΔS and ΔS^\ddagger is the electrostatic solvation of the charges formed [21]. Both these values are large and negative, and the solvation effects dominate over the translational entropy of bringing the reactants together

[21]. In Reaction 1 above, there is a decrease in solvation as a cation and a dianion react to form a mono-cation and neutral species. This reaction is more similar to a reverse Menschutkin reaction, which has been shown for pyridines to have positive ΔS^\ddagger , reflective of the ΔS for the reaction.

An examination of the activation enthalpies and entropies can provide insight into the ET and S_N2 reaction pathways. As was stated above, solvation changes will dominate the transition state for the S_N2 type reaction, leading to large negative or positive entropies of activation, depending upon whether ionic compounds are formed (large negative entropy) or charges are reduced (positive entropies). But relatively smaller values of ΔH^\ddagger are observed because of the concerted bond-breaking/bond-making process [13] (it should be noted that the reverse Menschutkin reaction for alkyl pyridiniums have a large ΔH^\ddagger , which is due to the endothermic and non-spontaneous nature of this reaction for alkyl halides and trialkylamines) [21]. In a similar manner, the ET mechanism, which is an outer sphere process, requires less organization in the transition state, but much higher values for ΔH^\ddagger because of the bond breaking process. An examination of Table 3 indicates that the reaction of $\text{Fe}(\text{TPP})^{2-}$ with TBA^+ is most consistent with an ET mechanism.

A stronger confirmation of the ET mechanism was obtained from the variation in the alkyl group in the tetraalkylammonium salts. The rate of the reaction of tetraalkylammonium ions and FeTPP^{2-} is consistent with the stability of the alkyl radical that would be initially formed.



Since the reaction appears to be primarily an ET mechanism rather than an S_N2 mechanism, the reaction rate would be controlled primarily by the energy required to break the C-N bond and form the radical. If there were significant bond formation in the transition state due to the nucleophilic character of FeTPP^{2-} , this should have been strongest for the TMA^+ ion because of the reduced steric hinderance in the transition state. The fact that tetramethylammonium ion reacts

substantially slower than TEA⁺ and TBA⁺ is consistent with an ET mechanism.

Conclusion

Tetraalkylammonium ions react with Fe(TPP)²⁻ to form Fe(TPP)(R)⁻ complexes and trialkylamines. The low value for ΔS^\ddagger and the high value for ΔH^\ddagger were most consistent with an ET mechanism instead of an S_N2 mechanism. Steric factors probably reduce the ability of the low valent iron complexes, which are known to have nucleophilic character, to react by an inner sphere process. The rate of the reaction increased as the carbon chain length from methyl to butyl, which was also consistent with an ET mechanism, reflecting the stability of the alkyl radical that was initially formed. The reactivity of the tetraalkylammonium ions indicates that they should be avoided whenever iron(“0”) porphyrins are formed.

Table 1

¹H and ²H NMR Spectroscopy of Iron Alkyl Porphyrins

Complex	Nuclei	$\delta,$				Reference
		ppm ^a				
		α	β	γ	δ	
Fe(TPP)(CH ₂ CH ₃) ⁻	¹ H	-6.19	-2.25	---	---	[22]
Fe(TPP)(CH ₂ CH ₂ CH ₃) ⁻	¹ H	-6.0	-1.5	-0.6	---	[22]
Fe(TPP)(CH ₂ CH ₂ CH ₂ CH ₃) ⁻	² H	-6.90	-2.15	-0.55	+0.01	This work

Table 2

Variation of the Peak Current Ratios for Wave 3
as a Function of Scan Rate and Temperature

Temperature	Peak current ratios		
Scan rate (mV/s)	50	100	200
47°	0.65	0.79	0.86
55°	0.46	0.65	0.75
62°	0.46	0.58	0.68
65°	0.42	0.55	0.65
72°	--	0.44	0.52

Solvent: DMF; working electrode: glassy carbon electrode; Electrolyte: 0.10 M TBAP

Table 3

Activation Parameters for the Reaction of Low Valent Iron Porphyrins with Substrates

Electron donor	substrate	"ET" behavior		"S _N 2" behavior		Reference
		ΔH^\ddagger , kcal	ΔS^\ddagger , cal/K	ΔH^\ddagger , kcal	ΔS^\ddagger , cal/K	
Fe(0)-ETIO	<i>n</i> -butyl bromide	--	--	2.3±0.4	-21±2	[13]
	<i>sec</i> -butyl bromide	11±2	4±5	--	--	[13]
	<i>t</i> -butyl bromide	10.0±0.5	-1.7±0.4	--	--	[13]
Fe(0)-TPP	<i>n</i> -butyl bromide	--	--	3.4±0.9	-22±3	[13]
Fe(I)-OEP	<i>n</i> -butyl bromide	--	--	6.5±0.8	-22±3	[13]
Fe(0)-TPP	TBA ⁺	15±1	-7±2	--	--	this work

Figure 1. A. Simulated cyclic voltammogram. B. Experimental cyclic voltammogram at a glassy carbon electrode. Concentration of Fe(TPP)Cl: 0.50 mM, concentration of TBAP: 0.10 M; solvent: DMF; scan rate: 100 mV/s; reference electrode: Ag/AgClO₄ in acetonitrile; temperature: 55°.

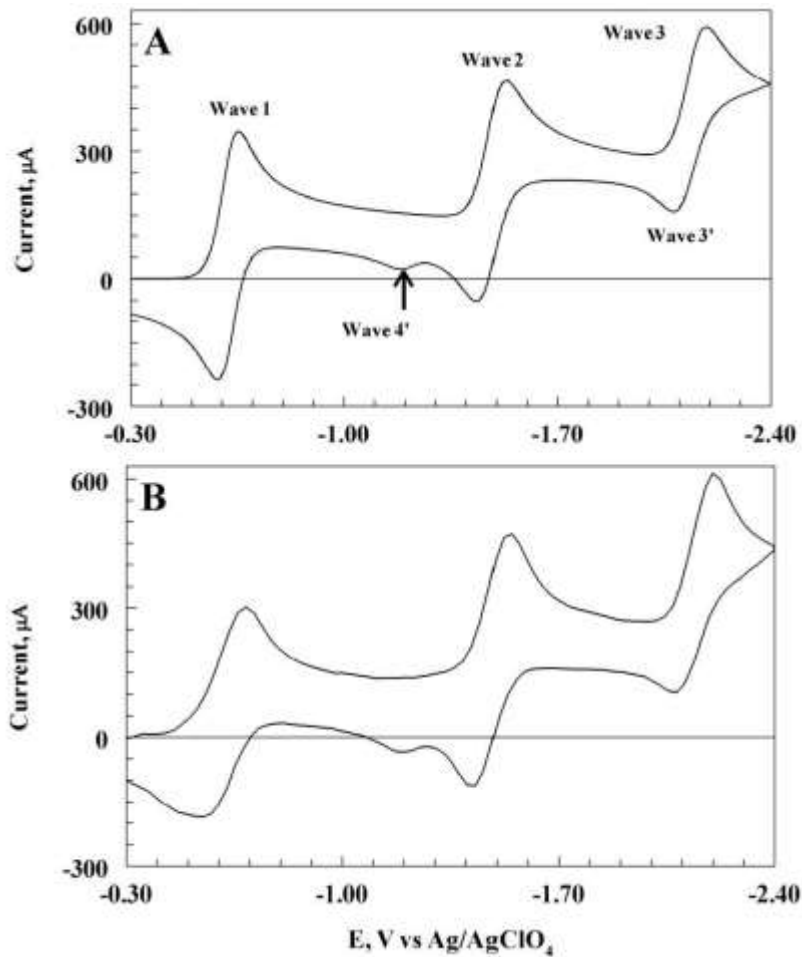


Figure 2. Arrhenius plot for the reaction of $\text{Fe}(\text{TPP})^{2-}$ with TBA^+ in DMF. Filled triangles: rate constants determined from cyclic voltammetry. Filled circle: rate constant determined from thin-layer spectroelectrochemistry.

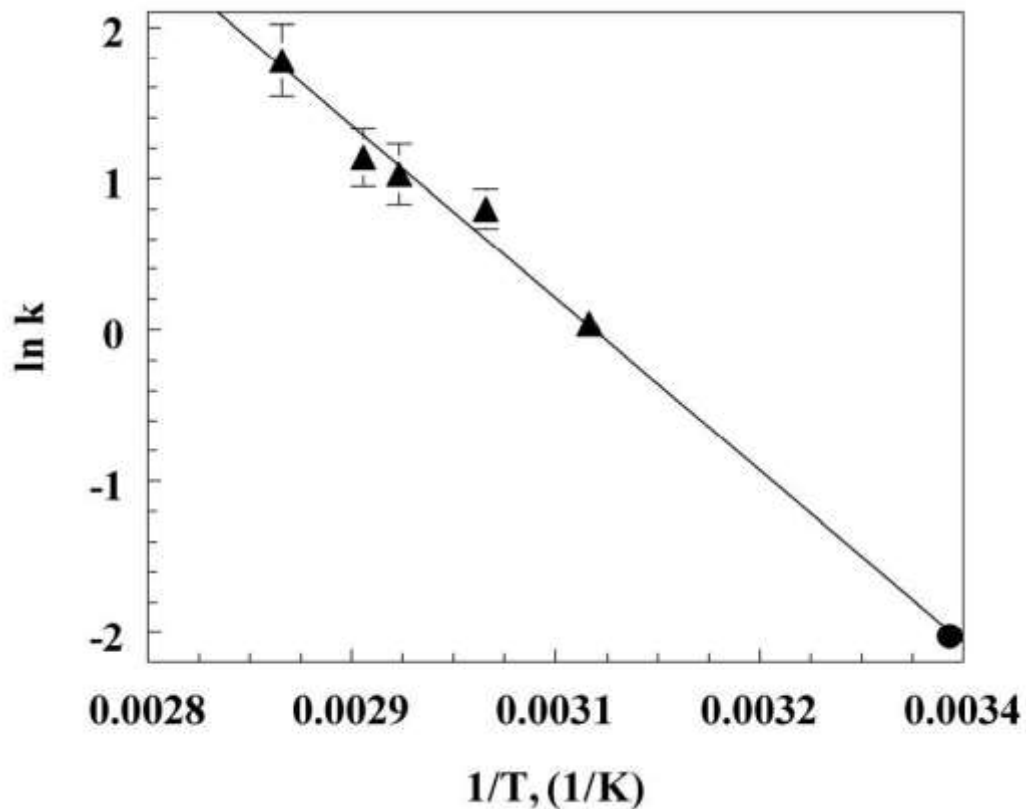
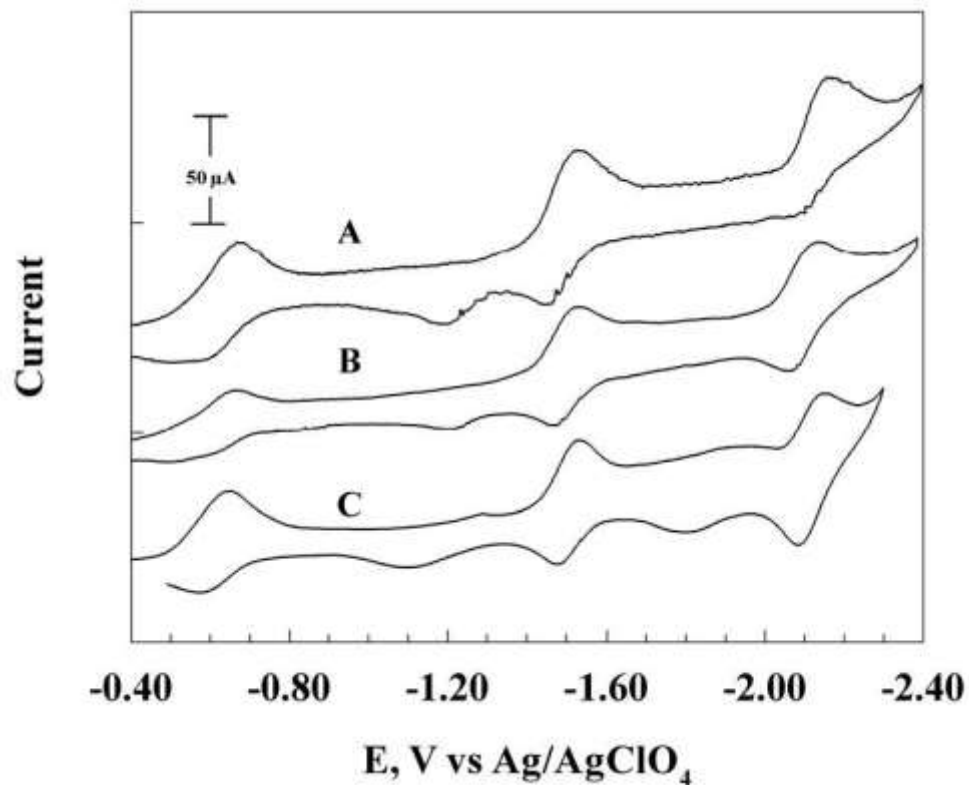


Figure 3. Cyclic voltammetry of 0.50 mM Fe(TPP)(Cl) in DMF. A. 0.10 M TBAP; B. 0.10 M TEAP; C. 0.10 M TMAP. Scan rate: 1 mV/s; electrode: glassy carbon; reference electrode: Ag/AgClO₄ in acetonitrile; temperature: 22°.



Reference List

1. DeSilva, C., Czarnecki, K., and Ryan, M. D., *Inorg. Chim. Acta* 1994; **226**: 195-201.
2. Aurbach, D. and Gottlieb, H., *Electrochim. Acta* 1989; **34**: 141-156.
3. Saveant, J. M. and Su, B. K., *J. Electroanal. Chem.* 1978; **88**: 27-41.
4. Dahm, C. E. and Peters, D. G., *J. Electroanal. Chem.* 1996; **402**: 91-96.
5. Kunkely, H. and Vogler, A., *Z. Naturforsch. ,B: Chem. Sci.* 2001; **56**: 431-432.
6. Lexa, D., Mispelter, J., and Saveant, J.-M., *J. Am. Chem. Soc.* 1981; **103**: 6806-6812.
7. Lexa, D., Saveant, J.-M., and Wang, D. L., *Organometallics* 1986; **5**: 1428-1434.
8. Gueutin, C., Lexa, D., and Saveant, J. M., *J. Electroanal. Chem.* 1988; **256**: 223-227.
9. Hammouche, M., Lexa, D., Saveant, J. M., and Momenteau, M., *J. Electroanal. Chem.* 1988; **249**: 347-351.
10. Bhugun, I., Lexa, D., and Saveant, J. M., *J. Am. Chem. Soc.* 1994; **116**: 5015-5016.
11. Bhugun, I., Lexa, D., and Saveant, J. M., *J. Am. Chem. Soc.* 1996; **118**: 1769-1776.
12. Bhugun, I., Lexa, D., and Saveant, J. M., *J. Am. Chem. Soc.* 1996; **118**: 3982-3983.
13. Lexa, D., Saveant, J.-M., Su, K.-B., and Wang, D.-L., *J. Am. Chem. Soc.* 1988; **110**: 7617-7625.
14. Lexa, D., Saveant, J. M., Schaefer, H. J., Su, K. B., Vering, B., and Wang, D. L., *J. Am. Chem. Soc.* 1990; **112**: 6162-6177.
15. Gordon, J. E., *J. Org. Chem.* 1965; **30**: 2760-2763.
16. Gordon, J. E. and Varughese, P., *J. Chem. Soc. ,Chem. Commun.* 1971; 1160-1161.
17. Missan, S. R., Becker, E. I., and Meites, L., *J. Chem. Soc.* 1961; 58-61.
18. Luder, W., Kraus, P. B., Kraus, C. A., and Fuoss, R. M., *J. Am. Chem. Soc.* 1936; **58**: 255-258.
19. Lin, X. Q. and Kadish, K. M., *Anal. Chem.* 1985; **57**: 1498-1501.
20. Nicholson, R. S. and Shain, I., *Anal. Chem.* 1964; **36**: 706-723.
21. Arnett, E. M. and Reich, R., *J. Am. Chem. Soc.* 1980; **102**: 5892-5902.
22. Balch, A. L., Cornman, C. R., Safari, N., and Latos-Grazynski, L., *Organometallics* 1990; **9**: 2420-2421.