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PAPER

Nickel complexes with “click”-derived pyridyl-triazole ligands: weak intermolecular interactions and catalytic ethylene oligomerisation†

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The ligands 1-(cyclohexyl)-4-(2-pyridyl)-1,2,3-triazole (**1**), 1-(2,6-diisopropylphenyl)-4-(2-pyridyl)-1,2,3-triazole (**2**), 1-(4-butoxyphenyl)-4-(2-pyridyl)-1,2,3-triazole (**3**) and 1-(methyl)-4-(2-pyridyl)-1,2,3-triazole (**4**) were synthesized by the Cu(i) catalyzed “Click” reaction between 2-pyridylacetylene and the corresponding azides. The ligands were then reacted with NiBr₂·3H₂O to generate the complexes (1)₂NiBr₂ (**1a**), (2)₂NiBr₂ (**2a**), (3)₂NiBr₂ (**3a**) and (4)₂NiBr₂ (**4a**). Structural characterization of **1a** confirmed the mononuclear and distorted octahedral environment around the Ni(II) center, with the pyridyl-triazole ligands coordinating in a bis-chelating fashion. Bond length analysis inside the 1,2,3-triazole ring shows a short N=N double bond that is flanked by two longer C–N and N–N bonds pointing to the existence of “azo” character in the ring. The highly polar five-membered 1,2,3-triazole ring makes its C–H bond acidic, and these bonds participate in an extended weak intermolecular C–H⋯Br interactions with the Br-groups of neighboring molecules, resulting in a 3-D network. The nickel complexes with these “Click” ligands were tested as pre-catalysts for ethylene oligomerization, and the complexes showed moderate activity in that reaction with good selectivity towards C₄ oligomers.

Introduction

The Huisgen [3 + 2] cycloaddition reaction between azides and alkynes was reported in the thermal form decades back.¹ The Cu(i) catalyzed version of this reaction has been popularized as the “Click” reaction in recent years (Scheme 1).^{2,3}

This reaction works under mild conditions, delivers high product yields of exclusively one isomer, and is tolerant against a host of functional groups. Because of these attractive features, the “Click” reaction has become an important tool in every chemist’s synthetic toolbox.⁴ The existence of reasonably basic nitrogen atoms in the 1,2,3-triazole rings, and the possibility of introducing additional donor groups in the substituents, has led to these classes of compounds being used as ligands in coordination chemistry.⁵ Thus a host of metal centers have been coordinated to such “Click” derived ligands, and the metal complexes have found use in electrochemical^{6–12} and photochemical studies,^{11–16} in supramolecular chemistry,^{17–21} magnetism,^{22,23}

metal-ion sensing^{4a,24} and catalysis.^{25–30} We have recently investigated metal complexes of substituted 1,2,3-triazole ligands for their electrochemical properties,³¹ magnetic bistability³² and their application in homogenous catalysis.^{33,34} In this context, substituted pyridyl-triazoles were used in combination with d⁸ metal centers like Pd(II) and Pt(II) by us^{31,35} and others.^{36–41} In this report we present new complexes (1)₂NiBr₂ (**1a**), (2)₂NiBr₂ (**2a**), (3)₂NiBr₂ (**3a**) and (4)₂NiBr₂ (**4a**) that contain the d⁸ metal center Ni(II) with the pyridyl-triazoles, 1-(cyclohexyl)-4-(2-pyridyl)-1,2,3-triazole (**1**), 1-(2,6-diisopropylphenyl)-4-(2-pyridyl)-1,2,3-triazole (**2**), 1-(4-butoxyphenyl)-4-(2-pyridyl)-1,2,3-triazole (**3**) and 1-(methyl)-4-(2-pyridyl)-1,2,3-triazole (**4**). We also describe here the differences in the synthesis and properties of these complexes compared to those of their higher homologues. Structural characterization of one of the complexes is provided. We also present the use of these complexes as pre-catalysts for ethylene oligomerisation. To the best of our knowledge, these are the first metal complexes based on substituted “click” pyridyl-triazole ligands that have been reported as pre-catalysts for ethylene oligomerisation. There have been some recent reports on the use of nickel complexes of triazole substituted α-diimine ligands as precatalysts for the polymerization of norbornene and styrene.^{42,43}

Results and discussion

Synthesis and characterization

The ligands 1-(cyclohexyl)-4-(2-pyridyl)-1,2,3-triazole (**1**),³⁵ 1-(4-butoxyphenyl)-4-(2-pyridyl)-1,2,3-triazole (**3**)⁴⁴ and

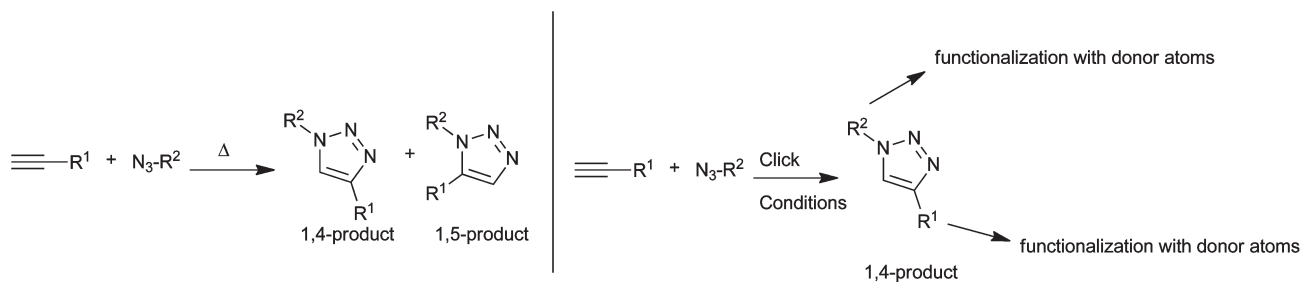
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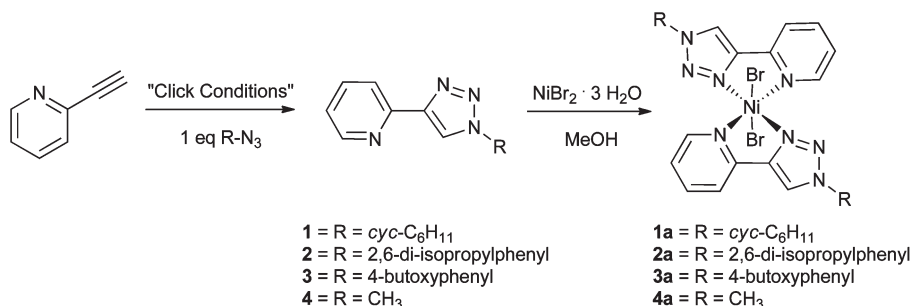
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Scheme 1 Thermal cycloaddition reaction between azides and alkynes (left) and the Cu(I) catalyzed "Click" reaction (right).



Scheme 2 Synthesis of complexes **1a–4a**.

1-(methyl)-4-(2-pyridyl)-1,2,3-triazole (**4**)⁴⁵ were synthesized according to the reported procedure, and the ligand 1-(2,6-diisopropylphenyl)-4-(2-pyridyl)-1,2,3-triazole (**2**) was prepared by a click protocol from 2-pyridylacetylene and 2,6-diisopropylphenylazide, using CuSO₄ together with ascorbic acid as the catalyst. Additionally, the tripodal ligand tris[(1-benzyl-1H-1,2,3-triazol-4-yl)methyl]amine, tbtta was used for improving the catalytic activity in that reaction (Experimental section). Reactions of the ligands **1**, **2**, **3** and **4** with NiBr₂·3H₂O afforded the complexes **1a**, **2a**, **3a** and **4a** in excellent yields (Scheme 2).

The synthesis of the complexes, which at a first glance looks straightforward, is actually cumbersome. Hence, a careful control of reaction conditions is required to synthesize the "bis" complex. After monitoring a host of conditions, and screening a set of ligands, we were able to determine the ones required to preferentially stabilize the "bis" complex. We were not able to synthesize four coordinated complexes of the form (L)₂NiBr₂, L = pyridyl-triazole under the tested conditions. For the corresponding Pd(II) and Pt(II) cases, the four coordinated complexes were the exclusively isolated products.^{31,35} This is of course related to the strong tendency of Pd(II) and Pt(II) complexes to stabilize a square planar environment because of the huge gain through ligand field splitting.⁴⁶ The 3d metal ion Ni(II) is capable of adopting a variety of coordination modes depending on the electronic and steric nature of the ligands. In the present case, it was possible to isolate octahedrally coordinated Ni(II) complexes with two bis-chelating ligands and two additional bromide ligands.

The complexes were characterized by elemental analyses, UV-Vis spectroscopy and mass spectrometry (Experimental section). Mass spectra of the compounds showed peaks at *m/z* corresponding to the di-cations formed by the loss of two

bromide ions from the complexes (Experimental section). Ni(II) centers in an octahedral environment are usually characterized by three different d–d transitions in the visible or NIR region in their absorption spectroscopy.⁴⁷ These bands which are transitions from the ³A_{2g}(F) to ³T_{1g}(P), ³T_{1g}(F) and ³T_{2g}(F) states, respectively, are spin allowed, but parity forbidden, and are hence usually weak in intensity. The solubility of the complexes reported herein is poor in all tested organic solvents. Hence, it was possible to record UV-Vis-NIR spectra of these complexes only with a limited concentration. The recorded spectra for all the complexes are quite similar (Fig. S1†). Of the three transitions mentioned above, the one at the lowest wavelength usually is the most intense. For **3a** a band of weak intensity is observed at 370 nm and this is assigned to a ³T_{1g}(P) → ³A_{2g}(F) transition. For the other complexes a weak shoulder can be seen at around the same region (Fig. S1†). The other two expected d–d transitions are not observed in the reported complexes, possibly due to their intrinsic weak nature and the low solubility of the complexes. Low solubility of the complexes precluded a reliable determination of the extinction coefficients. Additionally, all complexes show absorptions in the UV region which are assigned to intra-ligand transitions of the pyridyl-triazole ligands. Such bands are reminiscent of bands observed for the corresponding Pd(II) and Pt(II) complexes with such ligands.^{31,35} Interestingly, the position of these bands in the UV-region seems to correlate with the electronic nature of the substituents on the triazole ring. Thus, for **1a** and **4a** with alkyl substituents on the triazole ring, this band appears at 275 nm, and for **2a** and **3a** with aryl substituents, the same band appears at 283 nm (Fig. S1†).

Despite their paramagnetism, which precluded their characterization by NMR spectroscopy, complexes **1a** and **2a** were EPR

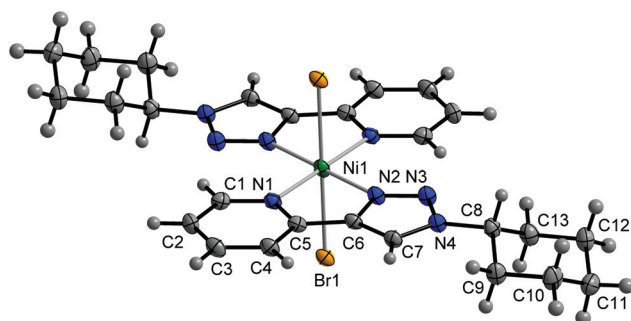


Fig. 1 ORTEP plot of **1a**. Ellipsoids are drawn at 50% probability. Ni1–N1, 2.084(2); Ni1–N2, 2.083(2); Ni1–Br1, 2.601(1); N2–N3, 1.312(3); N2–C6, 1.363(3); N3–N4, 1.348(3) Å. N1–Ni1–N2(chelate), 79.4(1); N1–Ni1–Br1, 90.3(1); N2–Ni1–Br1, 90.0(1); N1–Ni1–N2, 100.7(1).

Table 1 Crystallographic details for **1a**

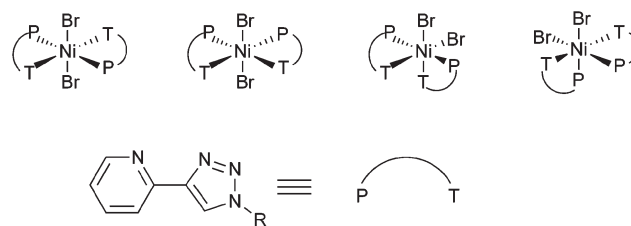
	1a
Chemical formula	C ₅₂ H ₆₄ Br ₄ N ₁₆ Ni ₂
<i>M_r</i>	1350.25
Cell setting, space group	Orthorhombic, <i>Pbca</i>
Temperature (K)	150(2)
<i>a</i> , <i>b</i> , <i>c</i> (Å)	13.0158(2), 9.8575(2), 20.8946(5)
<i>V</i> (Å ³)	2680.9(2)
<i>Z</i>	2
<i>D_x</i> (mg m ⁻³)	1.673
Radiation type	Cu Kα (λ = 1.54178 Å)
μ (mm ⁻¹)	4.790
Crystal size (mm)	0.25 × 0.20 × 0.18
Meas., indep. and obsvd refl.	6634, 2115, 1919
<i>R</i> [<i>F</i> ² > 2σ(<i>F</i> ²)], <i>wR</i> (<i>F</i> ²), <i>S</i>	0.024, 0.062, 1.018
No. of parameters	169
<i>R</i> _{int}	0.028
θ _{max} (°)	62.60
Δρ _{max} , Δρ _{min} (e Å ⁻³)	0.432, -0.331

silent down to 110 K in the X-band. This has to do with the large zero-field splitting usually observed for octahedral *S* = 1 nickel(II) centers, which is responsible for shifting their signals outside the range of the normal X-band EPR spectrometer.⁴⁶

Crystal structure of **1a**

1a (Fig. 1) was crystallized by slow evaporation of a methanol solution under ambient conditions. It crystallizes in the orthorhombic *Pbca* space group. Crystallographic details are given in Table 1. Selected bond lengths and bond angles are provided in the caption of Fig. 1.

The nickel center in **1a** is in a distorted octahedral coordination environment which includes the N3-triazole and pyridyl N atoms from two different ligands (**1**) in the equatorial plane. The additional axial positions are taken up by the bromide ions that are then trans to each other. In the equatorial plane the pyridine N as well as the triazole N donors from the two **1** ligands are mutually trans to each other. The Ni–N distances of 2.083(2) and 2.084(2) Å and the Ni–Br distances of 2.601(1) Å are in the expected range. The angles in the plane containing the two **1** ligands deviate from 90° because of the chelating nature of the



Scheme 3 Possible isomers of an octahedral nickel(II) complex with two substituted pyridyl-triazole ligands (adapted from ref. 49).

ligands **1**. The N2–N3 bond length inside the 1,2,3-triazole ring is 1.312(3) Å, and the N3–N4 and N2–C6 bond lengths are 1.348(3) and 1.363(3) Å, respectively. These distances show a short central N–N bond that is flanked by two relatively long N–N and N–C bonds. Such data point to the N=N “azo” character of the central N–N bond as opposed to an azine character. Similar observations have been made before by us for the corresponding Pd(II) and Pt(II) complexes of pyridyl-triazoles and by others for complexes with other metal centers.^{31,35,48} A similar coordination to that observed here for **1a**, was reported for a nickel complex with triazole-substituted α-diimine ligands.⁴² In that case the donor atoms from the tetradentate ligands were equatorially coordinated to the nickel centre. The bromide ligands take up the remaining axial sites around the octahedrally coordinated nickel center. In another report, predictions from DFT studies on nickel complexes containing ferrocenyl-substituted pyridyl-triazole ligands showed that among the various isomers that are possible (Scheme 3), the coordination observed here should be most stable.⁴⁹ In the reported compound, the co-ligands (perchlorates in the reported case and bromide in the present case) are in the axial positions and trans to each other, and the pyridine and triazole nitrogen donor atoms from the two different ligands occupy equatorial sites and are mutually trans to each other. However, no experimental structural evidence was reported for that nickel complex.⁴⁹ We have presented here structural evidence for the stability of this isomer in the solid state.

Due to the presence of three basic nitrogen atoms within a five-membered ring, the C–H bond of the 1,2,3-triazole ring is known to be highly acidic. This in turn enormously enhances the ability of these C–H bonds to participate in hydrogen bonding. Such targeted hydrogen bonding has been used for anion binding and anion sensing.⁵⁰ In the case of **1a**, extensive hydrogen bonding is observed between the bromide ligands that are bound to the nickel centres and the C–H group of the 1,2,3-triazole rings of neighboring molecules (Fig. 2). Thus, the C–H group of a triazole ring is hydrogen-bonded to the bromide ligand of the next molecule in the same layer. The second bromide ligand forms hydrogen bonds to a C–H group on a triazole ring of the next layer. Relevant distances and angles are C–H⋯Br = 2.676(2) Å, C–H–Br = 169.1(1)° and Ni–Br–H = 101.4(1)°. In such a way extensive three-dimensional patterns are formed in the solid state. Similar head to tail arrangements stabilized by hydrogen bonding between the acidic C–H group of the 1,2,3-triazole ring of one molecule and the N2 atom of the 1,2,3-triazole ring of an adjacent molecule have been recently observed for Pd(II) complexes with pyridyl-triazole ligands.⁴¹

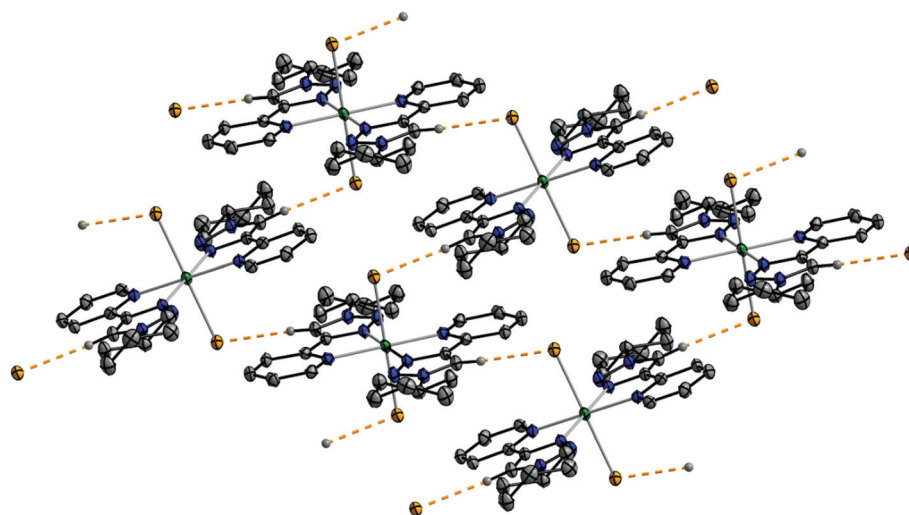


Fig. 2 Intermolecular C–H...Br interactions in **1a** resulting in an extended three dimensional network.

Table 2 Catalytic data for complexes **1a–4a** in ethylene oligomerisation with AlEtCl₂ as cocatalyst^a

Precatalyst	AlEtCl ₂ (equiv.)	Selectivity (mass%)			Activity [g C ₂ H ₄ /(g Ni·h)]	TOF [mol C ₂ H ₄ /(mol Ni·h)]	Selectivity for 1-butene (mol %) ^b
		C ₄	C ₆	C ₈			
1a	10	71	27	2	8000	28 900	7
	10	77	21	2	8500	30 600	7
2a	10	73	25	2	11 200	40 200	8
	10	71	26	3	12 200	43 800	10
3a	10	72	25	3	12 500	45 000	6
	10	74	23	3	11 500	41 300	7
4a	10	66	30	4	14 600	52 600	7
	10	63	33	4	15 900	57 300	6

^a Conditions: $T = 303$ K, 10 bar of C₂H₄, 35 min, 4×10^{-5} mol of complex; solvent: 10 mL of toluene (**1a** and **2a**) or chlorobenzene (**3a** and **4a**) for 10 equiv. of AlEtCl₂ (5 mL of AlEtCl₂ solution in toluene). ^b Within the C₄ fraction.

Catalytic ethylene oligomerisation

The demand for the production of linear α -olefins in the C₄–C₁₀ range is growing fast and this makes the catalytic oligomerisation of ethylene a topic of continuing academic and industrial interest.^{51–63} A range of nickel complexes with chelating ligands have been successfully used as catalysts for this purpose.⁶⁴ However, identifying and controlling the parameters that influence catalytic activity and selectivity still represents a major challenge in this field. Nickel complexes of 1,2,3-triazole-substituted α -diimine ligands have recently been used as precatalysts for the polymerization of norbornene and styrene.^{42,43} We present here results on the first use of nickel complexes of “click” derived pyridyl-triazole ligands in the catalytic oligomerisation of ethylene. Complexes **1a–4a** have been tested as precatalysts for this reaction with AlEtCl₂ as cocatalyst under 10 bar of ethylene pressure (Experimental section). Initial runs with **1a** showed AlEtCl₂ to be a better cocatalyst than MAO. Hence all subsequent runs with all the complexes were carried out with AlEtCl₂. Toluene turned out to be a better solvent for **1a** and **2a** and chlorobenzene for **3a** and **4a**. All the complexes exhibit reasonable activity in the range of about 8000–16 000 g C₂H₄/

(g Ni·h). The selectivity for the C₄ fraction was above 70 mass % for **1a–3a**. The complex **1a** with the cyclohexyl substituent on the triazole ring shows the lowest activity as well as TOF compared to **2a–4a** (see Table 2). The activity and TOF of complexes **2a–4a** are comparable, with **4a** containing the least bulky substituent on the 1,2,3-triazole ring, showing the least selectivity of all.

Conclusion

We have presented here the synthesis of a new pyridyl-triazole ligand, and have used three other literature reported pyridyl-triazole ligands to synthesise their corresponding nickel(II) complexes. Investigation through elemental analyses and mass spectrometry of the complexes confirmed the composition as NiL₂Br₂, L = pyridyl-triazole ligand. Structural characterization of one of the complexes showed that the bis-chelating ligands are in the equatorial plane of the octahedron, and the triazole and pyridine nitrogens from each of the ligands are *trans* to each other. Weak intermolecular C–H...Br interactions are observed in the solid state leading to an ordered three dimensional

structure. In the UV-Vis spectrum of the complexes only one of the three possible d-d transitions was observed. The nickel complexes turned out to be moderately good catalysts for the oligomerisation of ethylene, showing moderate activities and turnover numbers. This report thus presents new examples of nickel(II) complexes with the increasingly popular "Click" based 1,2,3-triazole ligands, and shows the utility of metal complexes of these ligands in yet another field of chemistry.

Experimental section

General

All solvents were dried and distilled using common techniques unless otherwise mentioned. ¹H-NMR spectra were recorded at 250.13 MHz on a Bruker AC250 instrument. Elemental Analyses was performed on a Perkin Elmer Analyser 240. Mass spectrometry measurements were carried out using an Agilent 6210 ESI-TOF instrument.

Syntheses

The ligands **1**, **3** and **4** were prepared according to reported procedures.^{35,44,45}

2: 2-Pyridylacetylene (206 mg, 2.0 mmol), 2,6-diisopropylphenylazide (406 mg, 2.0 mmol), CuSO₄·5 H₂O (25 mg, 0.1 mmol), sodium ascorbate (79 mg, 0.4 mmol) and TBTA (11 mg, 0.02 mmol) were dissolved in CH₂Cl₂/H₂O/*tert*-BuOH (2.5 mL/2.5 mL/5 mL) and stirred for 3 d at 50 °C. Then water (50 mL) was added and the reaction mixture was extracted with CH₂Cl₂ (3 × 20 mL). The organic phase was separated and washed with an EDTA/ammonia solution (1 M) (3 × 10 mL). Finally the solution was dried over Na₂SO₄ and the solvent was evaporated. After flash chromatography over silica (CH₂Cl₂: CH₃OH, 99:1) the product was isolated as a white solid (460 mg) in 75% yields. Anal. Calcd for C₁₉H₂₂N₄: C, 74.48; H, 7.24; N, 18.29; Found C, 74.06; H, 7.37; N, 18.10. HRMS (ESI) Calcd. for C₁₉H₂₃N₄ ([M + H]⁺): *m/z* 307.1917; found 307.1915. ¹H-NMR (250 MHz, CDCl₃): δ 1.15 (d, ³J_{H-H} = 6.8 Hz, 6H, CH₃); 1.17 (d, ³J_{H-H} = 6.8 Hz, 6H, CH₃); 2.35 (septet, ³J_{H-H} = 6.9 Hz, 2H, CH); 7.28 (t, ³J_{H-H} = 7.5 Hz, 1H, pyridyl); 7.32 (d, ³J_{H-H} = 7.9 Hz, 2H, phenyl); 7.51 (t, ³J_{H-H} = 7.8 Hz, 1H, phenyl); 7.84 (t, ³J_{H-H} = 7.7 Hz, 1H, pyridyl); 8.23 (s, 1H, 5-triazole-H); 8.32 (d, ³J_{H-H} = 7.9 Hz, 1H, pyridyl); 8.62 (multiplet, 1H, pyridyl). ¹³C-NMR (62.5 MHz, CDCl₃): δ 24.2; 24.4; 28.6; 120.7; 123.2; 124.0; 125.0; 131.0; 133.3; 137.1; 146.3; 148.3; 149.7; 150.5.

1a: NiBr₂·3H₂O (30 mg; 0.11 mmol) and **1** (50 mg; 0.22 mmol) were dissolved in MeOH (20 mL) and refluxed for 1 h. The solvent was evaporated to yield the product as a light blue solid (59 mg) in 80% yields. Slow evaporation of a methanol solution yielded crystals suitable for X-ray diffraction. Anal. Calcd for C₂₆H₃₂Br₂N₈Ni: C, 46.26; H, 4.78; N, 16.60; Found C, 45.74; H, 4.83; N, 16.33. UV-Vis (CH₂Cl₂): nm 276. HRMS (ESI) Calcd. for C₂₆H₃₂N₈Ni ([M - 2 Br]²⁺): *m/z* 257.1046; found 257.1047.

2a: The compound was prepared following the procedure for **1a** to yield a light blue solid (78 mg) in 85% yields. Anal. Calcd for C₃₈H₄₄Br₂N₈Ni: C, 54.90; H, 5.33; N, 13.48; Found C,

54.91; H, 5.36; N, 13.07. UV-Vis (CH₂Cl₂): nm 283. HRMS (ESI) Calcd. for C₃₈H₄₄N₈Ni ([M - 2 Br]²⁺): *m/z* 335.1516; found 335.1520.

3a: The compound was prepared following the procedure for **1a** to yield a white solid (62 mg) in 70% yield. Anal. Calcd for C₃₄H₃₆Br₂N₈NiO₂: C, 50.59; H, 4.50; N, 13.88; Found C, 50.47; H, 4.63; N, 13.69. UV-Vis (CH₂Cl₂): nm 282, 370. HRMS (ESI) Calcd. for C₃₄H₃₆N₈NiO₂ ([M - 2 Br]²⁺): *m/z* 323.1152; found 323.1187.

4a: The compound was prepared following the procedure for **1a** to yield a white solid (27 mg) in 46% yield. Anal. Calcd for C₁₆H₁₆Br₂N₈Ni: C, 35.66; H, 2.99; N, 20.79; Found C, 35.25; H, 3.15; N, 20.38. UV-Vis (CH₂Cl₂): nm 277. HRMS (ESI) Calcd for C₁₆H₁₆N₈Ni ([M - 2 Br]²⁺): *m/z* 189.0420; found 189.0432.

Catalytic oligomerization of ethylene

The catalytic reactions were performed in a magnetically stirred (1200 rpm) 145 mL stainless steel autoclave. A 125 mL glass container was used to avoid corrosion of the autoclave walls. The precatalyst solutions were prepared by dissolving **1a** (33.2 mg, 4 × 10⁻⁵ mol) or **2a** (27 mg, 4 × 10⁻⁵ mol) in toluene (10 mL), and **3a** (32.3 mg, 4 × 10⁻⁵ mol) or **4a** (21.5 mg, 4 × 10⁻⁵ mol) in chlorobenzene (10 mL). In all cases, AlEtCl₂ was used as cocatalyst (8 × 10⁻⁵ mmol in toluene). The solution of the precatalyst was injected into the reactor under an ethylene flux, then 5 mL of the cocatalyst toluene solution (10 equiv.) was added (total volume: 15 mL). All catalytic reactions were started between 20 and 30 °C. No cooling of the reactor was done during the reaction. After injection of the catalyst and cocatalyst solutions under a constant low flow of ethylene, the reactor was pressurized to 10 bars. The temperature increased, due solely to the exothermicity of the reaction. The 10 bars working pressure was maintained through a continuous feed of ethylene from a bottle placed on a balance to allow continuous monitoring of the ethylene uptake. At the end of each test (35 min, unless otherwise specified) a dry ice bath was used to rapidly cool the reactor. When the inner temperature reached 0 °C, the ice bath was removed, allowing the temperature to slowly rise to 18 °C. The gaseous phase was then transferred into a 10 L polyethylene tank filled with water. An aliquot of this gaseous phase was transferred into a Schlenk flask, previously evacuated, for GC analysis. The amount of ethylene not consumed was thus determined. Although this method is of limited accuracy, it was used throughout and gave satisfactory reproducibility. The reaction mixture in the reactor was quenched *in situ* by the addition of ethanol (10 mL), transferred into a Schlenk flask, and separated from the metal complexes by trap-to-trap evaporation (20 °C, 0.8 mbar) into a second Schlenk flask previously immersed in liquid nitrogen in order to avoid loss of product.

X-ray crystallography

A suitable single crystal of **1a** was selected and mounted onto a thin glass fiber. X-ray intensity data were measured at 150 K on an Oxford Gemini S Ultra diffractometer with the Enhance X-ray Source of Cu-Kα radiation (λ = 1.54178 Å) using the ω-φ

scan technique.⁶⁵ Empirical absorption correction was applied using spherical harmonics implemented in SCALE3 ABSPACK scaling algorithm.⁶⁶ The structure was solved by direct methods and refined by full-matrix least-squares against F^2 of all data using the SHELXTL program package.⁶⁷ Anisotropic thermal factors were assigned to the non-hydrogen atoms, while the positions of the hydrogen atoms were generated geometrically, assigned isotropic thermal parameters, and allowed to ride on their respective parent atoms before the final cycle of least-squares refinement. CCDC-851805 contains the supplementary crystallographic data for this paper.

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