

ORIGINAL ARTICLE

Exploring the Ni redox activity in polyanionic compounds as conceivable high potential cathodes for Na rechargeable batteries

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Although nickel-based polyanionic compounds are expected to exhibit a high operating voltage for batteries based on the Ni^{2+/3+} redox couple activity, some rare experimental studies on the electrochemical performance of these materials are reported, resulting from the poor kinetics of the bulk materials in both Li and Na nonaqueous systems. Herein, the electrochemical activity of the Ni^{2+/3+} redox couple in the mixed-polyanionic framework Na₄Ni₃(PO₄)₂(P₂O₇) is reported for the first time. This novel material exhibits a remarkably high operating voltage when cycled in sodium cells in both carbonate- and ionic liquid-based electrolytes. The application of a carbon coating and the use of an ionic liquid-based electrolyte enable the reversible sodium ion (de-)insertion in the host structure accompanied by the redox activity of Ni^{2+/3+} at operating voltages as high as 4.8 V vs Na/Na⁺. These results present the realization of Ni-based mixed polyanionic compounds with improved electrochemical activity and pave the way for the discovery of new Na-based high potential cathode materials.

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INTRODUCTION

Lithium-ion batteries have been largely recognized as the most efficient electrochemical energy storage devices for both portable electronics and electric vehicle applications.^{1,2} However, the growth and diversification of the energy storage market trigger interest in low-cost and environmentally friendly alternative systems. In addition, recent concerns over the cost and future availability of lithium highlight the urgent need to exploit alternative energy storage systems.³ In these terms, sodium (Na)-ion batteries are attractive candidates because of the electrochemistry that is similar to the well-established lithium-ion technology.⁴ To date, the LiFePO₄ olivine with (PO₄)³⁻ polyanionic framework, owing to its superior thermal stability and low cost, is considered to be one of the best electrode materials for lithium-ion batteries, mostly in view of its stable operating voltage and satisfactory specific capacity. However, the intrinsic tunability of the operating voltage of polyanionic compounds because of the presence of different transition metals such as Mn, Co and Ni⁵⁻⁷ has triggered interest in alternative frameworks exhibiting higher cell voltages. According to theoretical prediction, lithium nickel phosphate has a remarkably high working potential (~5.1 V vs Li^{+/Li}) because of the Ni^{2+/3+} redox activity.⁸ Unfortunately, several studies have demonstrated that the implementation of this material is restricted by several drawbacks such as the intrinsic sluggish kinetics attributable to the low electronic conductivity, the poor lithium transport in commonly used electrolyte systems and the structural

instability of the delithiated phases upon cycling that require further studies to obtain high-performance Ni-based phosphates as high voltage cathodes for battery application.⁸⁻¹²

With regard to Na-based systems, considerable research has also been devoted to NASICON (sodium (Na) Super Ionic CONductor)-type polyanionic materials, such as NaMPO₄,^{13,14} Na₂MP₂O₇ (where M is a transition metal)^{15,16} and Na₃V₂(PO₄)₃,^{17,18} as cathode materials for room temperature Na-ion batteries. However, the low energy densities resulting from the low theoretical capacity and, generally, cell voltage, as well as the poor rate capability compared with their lithium counterparts, have inhibited the widespread development of Na-ion batteries.¹⁹ Unlike LiMPO₄, the NaMPO₄ thermodynamically stable phase obtained under the conventional synthesis conditions is the maricite phase that is known to be electrochemically inactive with respect to the olivine phase.^{20,21} Among the polyanionic frameworks, pyrophosphate compounds offer a stable three-dimensional (P₂O₇)⁴⁻ framework with multiple sites for Na ions. Fe-based pyrophosphates have been reported to exhibit enhanced rate capability performance most likely because of the presence of open diffusion pathways for Na ions. Indeed, first principle calculations confirmed a relatively low migration barrier for diffusion of Na ions.¹⁵

In this context, a new polyanionic compound of the general formula Na₄M₃(PO₄)₂(P₂O₇) with the Fe^{2+/3+} redox couple has been proposed by Kang and colleagues^{22,23} as a promising cathode in terms

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of its superior Na mobility and thermal stability. The material has a theoretical capacity of 129 mAh g^{-1} and an operating potential of $\sim 3.2 \text{ V}$ vs Na^+/Na , higher than the potential exhibited by NaFePO_4 and $\text{Na}_2\text{FeP}_2\text{O}_7$.^{13,24} According to first principle calculations, the P_2O_7 dimers can buffer the volume change during cycling ($< 4\%$) and offer Na diffusion pathways in the three-dimensional open framework that may ensure long-term cycling performance. Cobalt-based polyanionic compounds show electrochemical properties similar to the Ni-based analogs in Li systems with low conductivity and poor thermal stability.^{25,26} However, the $\text{Na}_4\text{Co}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)$ synthesized by a sol-gel method shows an average working potential of $\sim 4.4 \text{ V}$ vs Na^+/Na and a high rate capability for Na ions.²⁷ A combination of different polyanionic groups could contribute to outstanding ion diffusion and voltage profile in Na-ion batteries as well as structure stability during the Na (de-)insertion process. Interestingly, a theoretical prediction reported by Islam and colleagues²⁸ has pointed out that doping the $\text{Na}_4\text{Fe}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)$ with Ni can lead to a significant increase in the cell voltage.

In light of such results, we report the investigation of a Ni-based polyanionic material as a potential high voltage cathode for rechargeable Na batteries. In this work, the pristine and carbon-coated materials, that is, $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)$ and $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)/\text{C}$, were stoichiometrically prepared by a conventional solid-state method. This study represents an intriguing investigation in terms of both structural and electrochemical characterization of the Ni redox activity in mixed polyanionic compounds. The electrochemical activity of the $\text{Ni}^{2+/3+}$ redox couple in Na cells, predicted by previous theoretical calculations, is experimentally reported in this paper for the first time. The attractive possibility of using high operating voltage Na-based cathode materials highlights the need of electrolyte media with a wide electrochemical stability window, opening a new unexplored research field for Na-ion batteries.

MATERIALS AND METHODS

Synthesis of materials

All chemicals were used as received without further purification. Syntheses were carried out by simple and time-efficient solid-state reactions. For the precursor preparation, sodium pyrophosphate decahydrate ($\text{Na}_4\text{P}_2\text{O}_7 \cdot 10\text{H}_2\text{O}$, $\geq 99\%$, Alfa Aesar, Karlsruhe, Germany), nickel (II) acetate tetrahydrate (Ni (OCOCH_3)₂·4H₂O, $\geq 98\%$, Alfa Aesar) and ammonium dihydrogen phosphate ($\text{NH}_4\text{H}_2\text{PO}_4$, $\geq 99\%$, Merck Millipore, Darmstadt, Germany) were mixed by ball-milling in acetone according to stoichiometric proportions. Citric acid monohydrate ($\text{C}_6\text{H}_8\text{O}_7 \cdot \text{H}_2\text{O}$, $\geq 99\%$, Alfa Aesar) was used as a carbon source for the carbon-coated sample. After intensive milling, the mixture was dried at room temperature and ground. The precursor powder obtained was preheated at 300°C for 6 h and then annealed at 500°C (heating rate: $2^\circ \text{C min}^{-1}$) for 10 h in an aluminum oxide crucible using a tubular furnace under steady argon flow. Cooling to room temperature and grinding followed. The carbon-coated material was synthesized under the same conditions using citric acid monohydrate as the organic carbon source in the precursor preparation. A schematic representation of the synthesis procedure is reported in Supplementary Figure S1.

Material characterization

X-ray diffraction (XRD) analysis was performed in Bragg-Brentano geometry using a Bruker D8 Advance diffractometer (Bruker, Karlsruhe, Germany). All diffractograms were recorded with CuK_α radiation ($\lambda = 0.154 \text{ nm}$) in the 2θ range of 10° – 90° with a step size of 0.01° . The structures were refined using the Rietveld method with the DIFFRAC PLUS TOPAS 4.2 software (Bruker). Morphological and structural characterizations were conducted via ZEISS 1550VP Field Emission scanning electron microscopy (Carl Zeiss, Jena, Germany) and Cs-corrected FEI Titan 80–300 kV high-resolution transmission electron microscopy operated at an acceleration voltage of 80 and 300 kV (FEI, Eindhoven, The Netherlands). The elemental

distribution was investigated using an energy-dispersive X-ray spectrometer (Oxford Instruments, Munich, Germany). Thermal gravimetric analysis was performed using a thermogravimetric analyzer (TGA-209F, Netzsch, Selb, Germany) under O_2 flow of 20 ml min^{-1} within the temperature range of 30 – 600°C (final materials) and 30 – 800°C (precursors) at a $5^\circ \text{C min}^{-1}$ scan rate. Tap density measurements were performed on a tap density analyzer (Autotap, Quantachrome, Boynton Beach, FL, USA), with 1000 taps per sample.

Electrochemical characterization

The electrochemical tests were performed using three electrode Swagelok-type cells assembled in an argon-filled glove box. Sodium half-cells were assembled by using glass fiber filters (Whatman, Maidstone, UK) as separators soaked in various electrolyte solutions and sodium metal as the counter and reference electrodes. The 1 M solution of NaPF_6 (99%, Alfa Aesar) in ethylene carbonate/diethyl carbonate (1:1 w/w) (BASF, Ludwigshafen, Germany, Battery Grade) and the $10 \text{ mol}\%$ solution of NaTFSI (sodium bis(trifluoromethanesulfonyl)imide, 99.5%, SOLVIONIC, Varilhes, France) in Py_{14}FSI (*N*-butyl-*N*-methylpyrrolidinium bis(trifluoromethylsulfonyl)imide) were used as electrolytes. Sodium metal was cut from sodium pieces (99.8%, Acros Organics, Morris Plains, NJ, USA), roll-pressed and punched on the current collector. The electrodes were prepared by doctor-blade casting a slurry prepared by dispersing 80 wt% of active material, 10 wt% of carbon black as conductive agent (Super C65, IMERYS, Bodio, Switzerland) and 10 wt% of polyvinylidene fluoride (6020 Solef, Arkema Group, Paris, France) binder in *N*-methyl-2-pyrrolidone (anhydrous, 99.5%, Sigma-Aldrich, St Louis, MO, USA) on Al foil as the current collector. After drying, electrodes with a 12 mm diameter were punched, pressed and dried again. The mass loading was $\sim 2.8 \text{ mg cm}^{-2}$. The cells were assembled in an argon-filled glove box with H_2O and O_2 contents $< 1 \text{ p.p.m.}$ A Maccor 4000 Battery system (Maccor, Tulsa, OK, USA) was used for the galvanostatic charge/discharge test between 3.0 and 4.9 V vs Na^+/Na in the organic electrolyte and between 3.0 and 5.1 V vs Na^+/Na in the ionic liquid electrolyte at a current density of 10 mA g^{-1} . Cyclic voltammetry was performed on a multichannel potentiostat/galvanostat (VMP3, Biologic Science Instruments, Claix, France) using a scan rate of 0.01, 0.05, 0.1 and 0.5 mV s^{-1} in a voltage range of 3.0 – 5.0 V vs Na^+/Na (organic electrolyte) and 3.0 – 5.2 V vs Na^+/Na (ionic liquid electrolyte). All the potential values in this study are in reference to the Na^+/Na reference electrode. All measurements were performed in climatic chambers at a temperature of $20 \pm 1^\circ \text{C}$.

RESULTS AND DISCUSSION

The structural properties of the $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)$ have been determined by XRD analysis. Figure 1 shows the Rietveld refinement and the schematic representation of the crystal structure deduced for $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)$. Rietveld refinement of the powder XRD pattern confirms that the $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)$ is isostructural to $\text{Na}_4\text{Fe}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)$ and crystallizes in the orthorhombic $Pn2_1a$ space group with $a = 18.0006(3) \text{ \AA}$, $b = 6.4933(1) \text{ \AA}$ and $c = 10.4115(2) \text{ \AA}$. The refinement has been performed with satisfactory agreement factors ($R_{\text{wp}} = 4.79\%$, $R_p = 3.45\%$, goodness of fit = 2.13) (ICSD (Inorganic Crystal Structure Database) 236316).²⁷ The diffractogram reported in Figure 1a presents no evidence of impurity phases, suggesting that the synthesis conditions adopted are suitable for obtaining pure $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)$. The refined cell parameters and atomic occupancy values are summarized and presented, respectively, in Supplementary Tables S1 and S2. Rietveld refinement of the XRD pattern of the carbon-coated analog ($\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)/\text{C}$) (Supplementary Figure S2) shows that the carbon coating process has no effect on the structural characteristics of the material that crystallizes in the same space group with comparable cell parameters. The crystal structures of $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)$ viewed along the [001] and [010] directions are displayed in Figure 1b. As the isostructural $\text{Na}_4\text{Co}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)$ and $\text{Na}_4\text{Fe}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)$, the crystal structure of $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)$ is composed of three-dimensional network layers of $[\text{Ni}_3\text{P}_2\text{O}_{13}]_\infty$ jointly composed by NiO_6 octahedra and

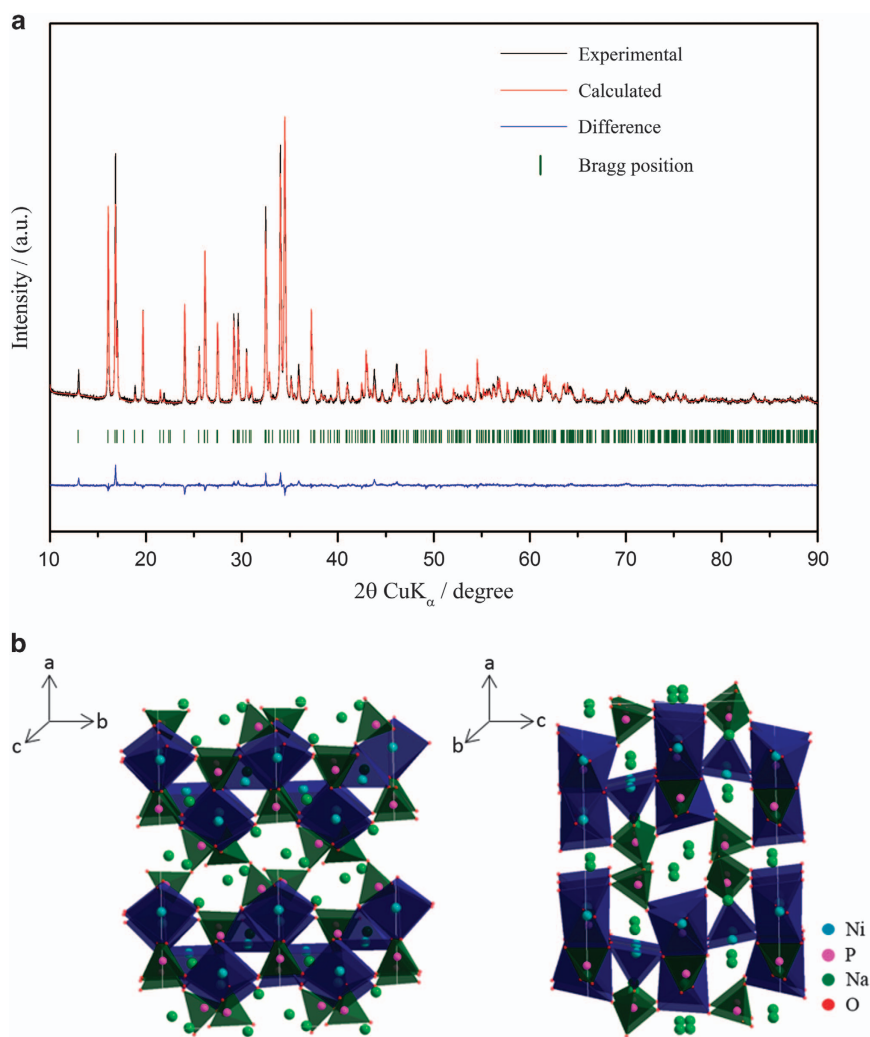


Figure 1 (a) Rietveld refinement and powder X-ray diffraction (XRD) for $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)$. (b) Schematic representation of the crystal structure of $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)$ viewed along the [001] and [010] directions.

$(\text{PO}_4)^{3-}$ tetrahedra in the bc plane that are bridged by the $(\text{P}_2\text{O}_7)^{4-}$ dimers generating large channels along the b -axis with a low activation barrier energy for Na^+ ion diffusion.^{23,29} The Na ions occupy four distinguishable sites, among which two (Na1 and Na4) are located in the large channels parallel to the b -axis. The Ni-based polyanionic materials have so far been considered as an electrochemically inactive cathode material in both Li and Na systems because of the low electronic conductivity resulting in the poor alkali ion diffusion activity.¹² However, in this study, the structure deriving from the mixed polyanionic $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)$ may stabilize the charged Ni analog and open pathways for the improved and reversible Na^+ diffusion with respect to the NaNiPO_4 and $\text{Na}_2\text{NiP}_2\text{O}_7$ polyanionic materials that can contribute to the improved electrochemical reactivity.

Scanning electron microscopy and high-resolution transmission electron microscopy analyses were conducted to assess the morphological and structural aspects of the synthesized powders. Figures 2a and b present the scanning electron microscopy images of $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)$ and $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)/\text{C}$. As shown in Figure 2a, the pristine $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)$ aggregates in shapeless micrometric particles (that is, polymorphic morphology). After the carbon coating treatment, the particle size of the powder is favorably reduced with well-defined edges. The tap density of carbon-coated $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2$

(P_2O_7) material is $\sim 1.01 \text{ g cm}^{-3}$, that is, smaller than the tap density of the pristine $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)$ material at 1.24 g cm^{-3} , that can be attributed to the reduced particle size and the carbon coating. The elemental mapping by energy-dispersive X-ray spectroscopy of Na, Ni, O, P and C has been performed to check the metal element distribution in $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)$ (Supplementary Figure S3). The analysis reveals a homogeneous distribution of the elements investigated, confirming the elemental composition. The transition electron microscopy analysis of $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)/\text{C}$, reported in Figure 2c, confirms the shapeless character of the particles, showing a sub-micrometric dimension with an average diameter of $\sim 500 \text{ nm}$. In addition, the high-resolution transmission electron microscopy image in Figure 2d clearly shows a lattice fringe of $d = 0.680 \text{ nm}$, corresponding to the (201) plane diffraction of $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)/\text{C}$. This material shows structural instability under an intense electron beam demonstrated by the disappearance of the lattice fringe after a prolonged electron beam irradiation (Supplementary Figure S4), opening the question about the crystal stability of the mixed polyanionic structure under nonconventional conditions.³⁰ Thermal gravimetric analysis (Supplementary Figure S5) was performed to investigate the thermal stability and carbon content of the prepared materials. The thermogram of pristine $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)$ shows a

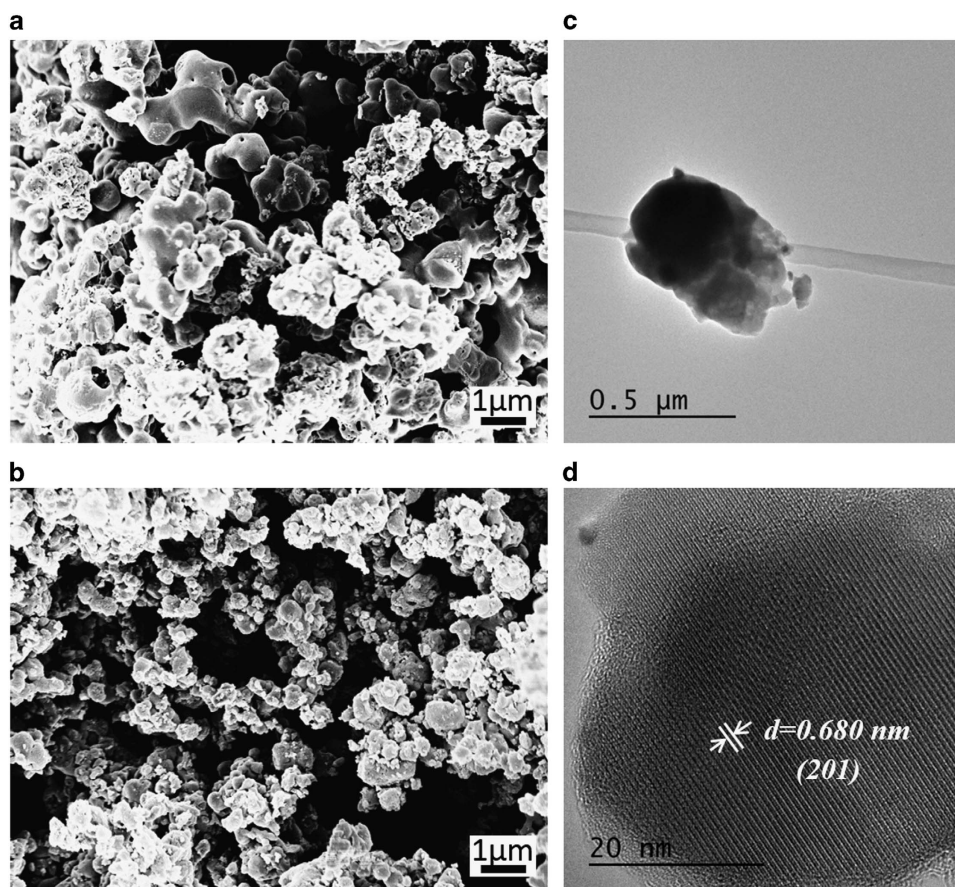


Figure 2 Scanning electron microscopy (SEM) images of (a) $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)$ and (b) $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)/\text{C}$ powders. (c, d) Transmission electron microscopy (TEM) images of $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)/\text{C}$ performed on an FEI Titan 80–300 kV electron microscope with image Cs-corrector at an accelerating voltage of 80 kV.

slight weight loss before 100 °C, attributable most likely to water desorption. The weight loss detected from 300 °C is related to the oxidation of residual carbon belonging to the organic salt precursor. Meanwhile, no weight loss is observed after 500 °C, even when exposed in O_2 , reflecting a considerably high thermal stability. The content of absorbed water in the carbon-coated material is higher than in the pristine material, suggesting a higher sensitivity to moisture, most likely because of a higher surface area. The carbon content is calculated as ~6.3% that corresponds to the weight detected within the 100–450 °C temperature range.

The investigation of Ni redox activity in the polyanionic framework has been performed by galvanostatic cycling tests and cyclic voltammetry analysis. In an attempt to study the effect of the polyanionic framework in the Ni redox couple, three different Ni-based compounds were synthesized, that is, NaNiPO_4 (the phosphate), $\text{Na}_2\text{NiP}_2\text{O}_7$ (the pyrophosphate) and $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)$ (the mixed polyanionic compound). $\text{Na}_2\text{NiP}_2\text{O}_7$ and NaNiPO_4 polyanionic compounds have been prepared using the same solid-state method. As the Fe-based compound, the stoichiometric $\text{Na}_2\text{NiP}_2\text{O}_7$ crystallizes in the triclinic *P*-1 space group. The XRD pattern obtained is shown in Supplementary Figure S6a. NaNiPO_4 has also been prepared according to the synthesis conditions reported in the Supplementary Information. However, orthorhombic $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)$ was obtained instead of maricite NaNiPO_4 (the XRD pattern obtained is presented in Supplementary Figure S6b), reflecting the thermodynamic stability of the $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)$ phase under our experimental conditions.

To investigate the electrochemical processes occurring during sodium (de-)insertion, the carbon-coated materials have been selected as the material of choice, owing to its enhanced conductivity with respect to the pristine material.³¹ Moreover, in an attempt to enlarge the electrochemical stability window of the electrolyte and enable the investigation of the Ni redox activity at higher potential values, an ionic liquid-based electrolyte, that is, NaTFSI in Py_{13}FSI (1:9 in mole ratio), was employed.^{32,33} A comparative electrochemical investigation between the carbon-coated $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)$ and $\text{Na}_2\text{NiP}_2\text{O}_7$ is shown in Figure 3, showing the differential capacity vs voltage (dQ/dV) curves of $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)/\text{C}$ and $\text{Na}_2\text{NiP}_2\text{O}_7/\text{C}$. The dQ/dV curve of $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)/\text{C}$ shows three anodic peaks at 4.61, 4.67 and 5.00 V with two reversible cathodic peaks at 4.58 and 4.92 V, elucidating the redox activity of the Ni redox couple in a sodium cell accompanied by reversible Na ion (de-)insertion. For the dQ/dV curve of the $\text{Na}_2\text{NiP}_2\text{O}_7/\text{C}$ electrode, no electrochemical activity is observed, suggesting that Ni is not active in the pyrophosphate framework under our experimental conditions. From the comparison, it is incidentally determined that the mixed phosphate frameworks enable the $\text{Ni}^{2+/3+}$ redox activity in polyanionic structures compared with those of pyrophosphates.

To further understand the sodium (de-)insertion process associated with the Ni activity in such a framework, cyclic voltammetry and charging/discharging measurements were performed with NaTFSI in Py_{13}FSI (1:9 mole ratio) electrolyte, and the corresponding results are shown in Figures 4a and b. The cyclic voltammetry profile shown in Figure 4a exhibits three anodic peaks at 4.61, 4.67, and 5.00 V and

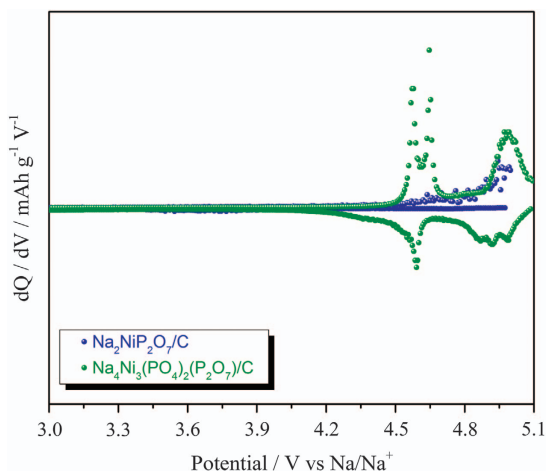


Figure 3 Initial differential capacities vs potential (dQ/dV) curves of the $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)/\text{C}$ and $\text{Na}_2\text{NiP}_2\text{O}_7/\text{C}$ in the NaTFSI: Py₁₃FSI (1:9 mole ratio) electrolyte at a current density of 10 mA g^{-1} .

only two cathodic peaks at 4.58 and 4.92 V, in accordance with the dQ/dV profile reported in Figure 3. The potential profile reported in Figure 4b reveals the de-insertion of 2.7 Na equivalents followed by the insertion of 1.3 Na equivalents, corresponding to a discharge capacity of 63 mAh g^{-1} , representing a promising degree of reversibility for this unexplored material. However, based on the $\text{Ni}^{2+/3+}$ redox reaction, a theoretical capacity of 127.2 mAh g^{-1} is expected, referring to 3 equivalents of Na^+ . The investigation of a wide potential window shows clearly that there is no further redox process at potentials higher than 5.0 V. In addition, electrolyte decomposition, aluminum current collector corrosion and possible oxidation of the carbon deriving from the carbon coating process, which may also affect the cycling performance, cannot be excluded at such high potential values vs Na^+/Na .^{34,35} Cyclic voltammetry tests have also been performed in standard carbonate-based electrolyte (1 M NaPF₆ in ethylene carbonate/diethyl carbonate, 1:1 w/w) (see Supplementary Figure S7). The cyclic voltammetry curves obtained at various scan rates (Supplementary Figure S7a) confirm the reversibility of the sodium (de-)insertion also at higher scan rates, despite the increased cell polarization. This observation is in good agreement with the charge/discharge voltage profile (Supplementary Figure S7b), suggesting a reversible redox process occurring at such a high working potential. The material reveals a satisfactory cycling stability. However, the redox peaks decrease upon cycling, most likely because of poor kinetics. More detailed description and discussion can be found in the Supplementary Information section. Regardless of its high working potential, the material exhibits relatively low coulombic efficiency and poor specific capacity compared with other reported Fe- or Co-based polyanionic compounds in both organic and ionic liquid electrolytes.^{36,37} The inferior electrochemical behavior of the Ni-based compound with respect to Fe- and Co-based analogs may be attributed to the intrinsic lower ionic conductivity of the former material. Sanz *et al.*¹² have reported that the Ni-based compound exhibits the lowest ionic conductivity of $\sim 2.1 \times 10^{-7} \text{ S cm}^{-1}$ at 300°C , with respect to Fe, Mn and Co.

Potential profiles and cycling performance of $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)/\text{C}$ in both carbonate- and ionic liquid-based electrolytes are presented in Supplementary Figure S8. The profiles reveal first discharge capacities of ~ 40 and 63 mAh g^{-1} , respectively, in carbonate-based (3.0 to 4.9 V) and ionic liquid-based (3.0 to 5.1 V) electrolytes. However, the wider potential window inevitably leads to lower coulombic efficiencies

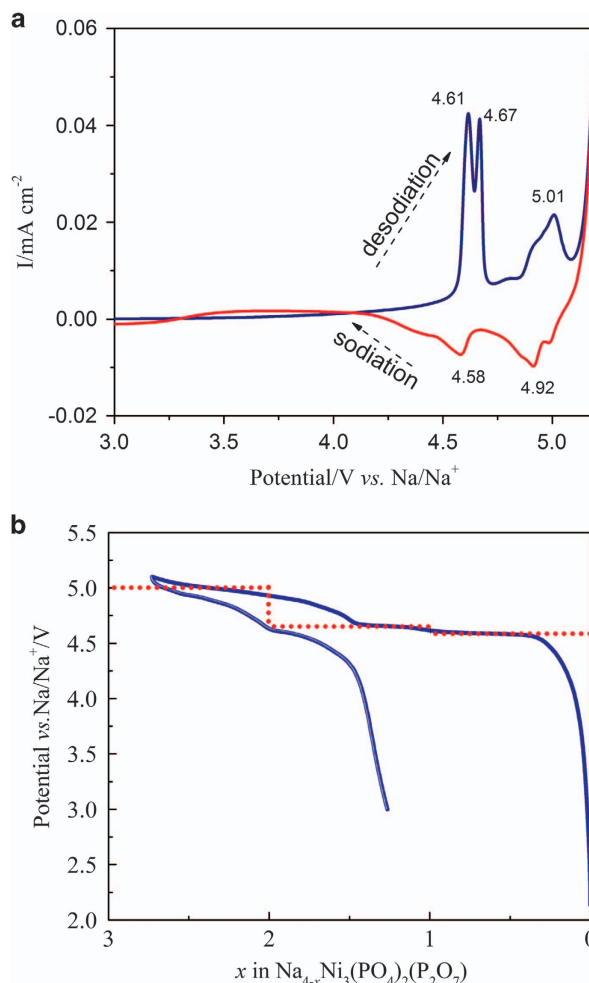


Figure 4 Cyclic voltammetry curve in the potential range of 3.0–5.2 V at a scan rate of 0.01 mV s^{-1} (a) and first cycle galvanostatic discharge/charge profile between 3.0 and 5.1 V at a current density of 10 mA g^{-1} (b) of $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)/\text{C}$ in NaTFSI:Py₁₃FSI (1:9 mole ratio) electrolyte.

because of oxidative decomposition reactions. As an example, at the highest potential (that is, 5.1 V vs Na^+/Na , corresponding to 5.4 V vs Li^+/Li), the decomposition of Py₁₃FSI ionic liquid is known to occur.³³ The conductive agent, Super C65, is playing a role at such potential values. The use of carbon nanotubes, for example, results in the 25% discharge capacity increase in carbonate-based electrolyte (see Supplementary Figure S9). However, extensive electrolyte decomposition occurs, resulting in an overall lower coulombic efficiency. A different electrode current collector (aluminum) and binder may also be used. For example, higher anodic stability ($\sim 5.5 \text{ V vs Na}^+/\text{Na}$) was reported using polytetrafluoroethylene as a binder, carbon paper as the current collector and 0.5 M NaTFSI in Py₁₃TFSI as the electrolyte. However, oxidative decomposition would always occur.³⁸

To better understand the factors affecting the electrochemical performance exhibited, further structural investigation has been performed. It has been proposed that structural reorganization upon cycling could affect the reversibility of the sodium (de-)insertion. Indeed, the structural changes during Na (de-)insertion may derive from the instability of the desodiated electrode, most likely because of the distortion of P_2O_7 dimers and changes in the unit cell volume.^{8,23}

The investigation of the Na (de-)insertion mechanism in the $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)/\text{C}$ electrode was carried out by *ex situ* XRD

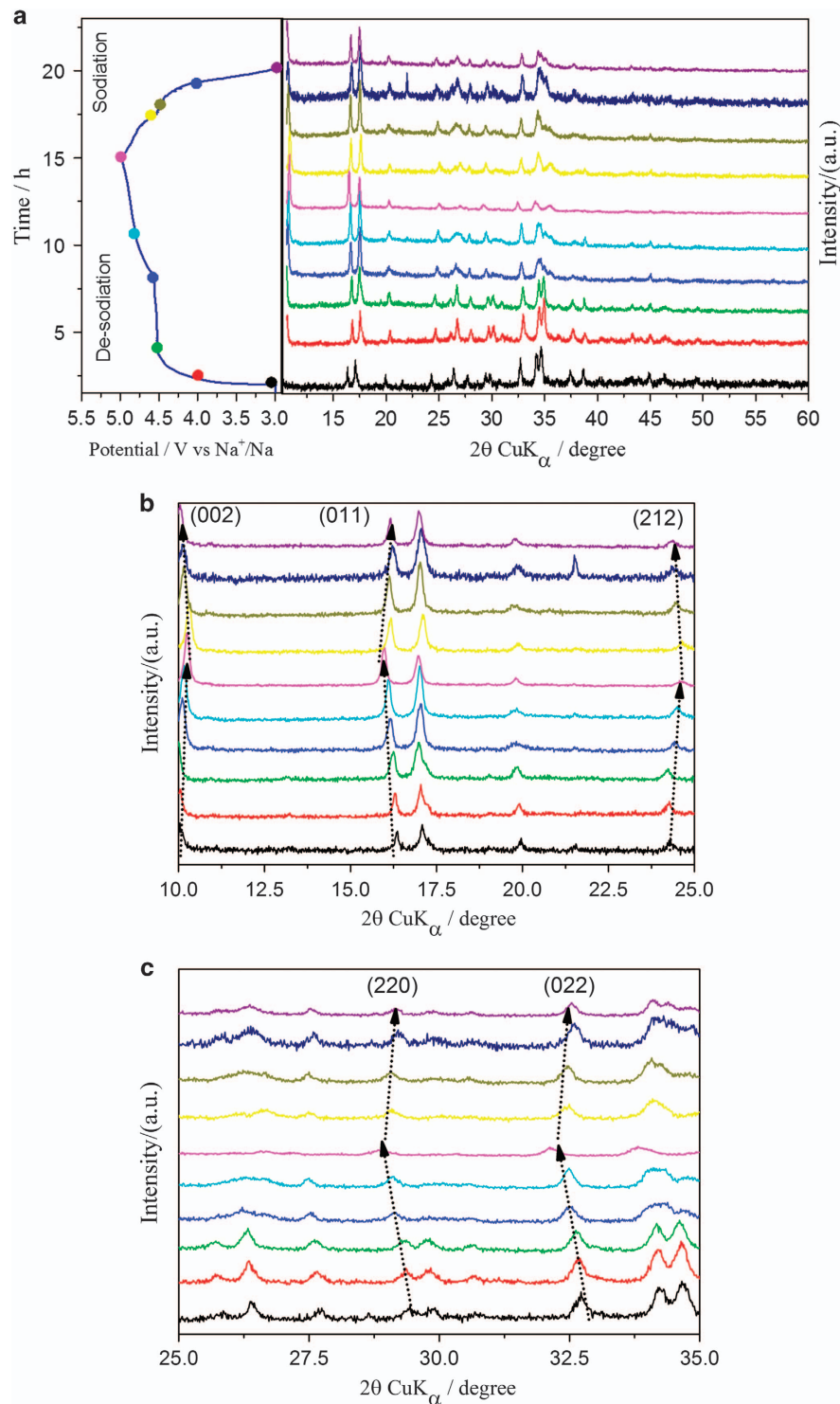


Figure 5 (a) Typical charge/discharge profile and *ex situ* X-ray diffraction (XRD) pattern of the $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)/\text{C}$ electrode at various charge and discharge states. (b, c) Amplified XRD patterns in the 2θ range of 10–25° and 25–35°, respectively.

measurements at different states of charge. Figure 5a displays the typical charge/discharge profile of a $\text{Na}/\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)/\text{C}$ cell with points and letters indicating the potential values at which XRD analysis has been performed. When charging the electrode to 5.0 V in the ethylene carbonate/diethyl carbonate-based electrolyte, the initial charge capacity is higher than the theoretical capacity that, at such a high working potential, may be attributed to the cathode–electrolyte

interface film formation of the Ni-based electrode materials and oxidative electrolyte decomposition.^{39,40} The patterns collected in the 10–60° 2θ range are displayed in Figure 5a. The main reflections of the mixed polyanionic phase are shown to be maintained upon charge and discharge, whereas moderate peak shifting and intensity changes occur, suggesting that neither phase transition nor irreversible structural reorganization occurs upon sodium de-insertion. The

amplified XRD patterns in the 2θ range of $10\text{--}25^\circ$ and $25\text{--}35^\circ$ in Figures 5b and c clearly show the peak shifts during the electrochemical reaction. Interestingly, although the (200) peak at $\sim 10^\circ$ is

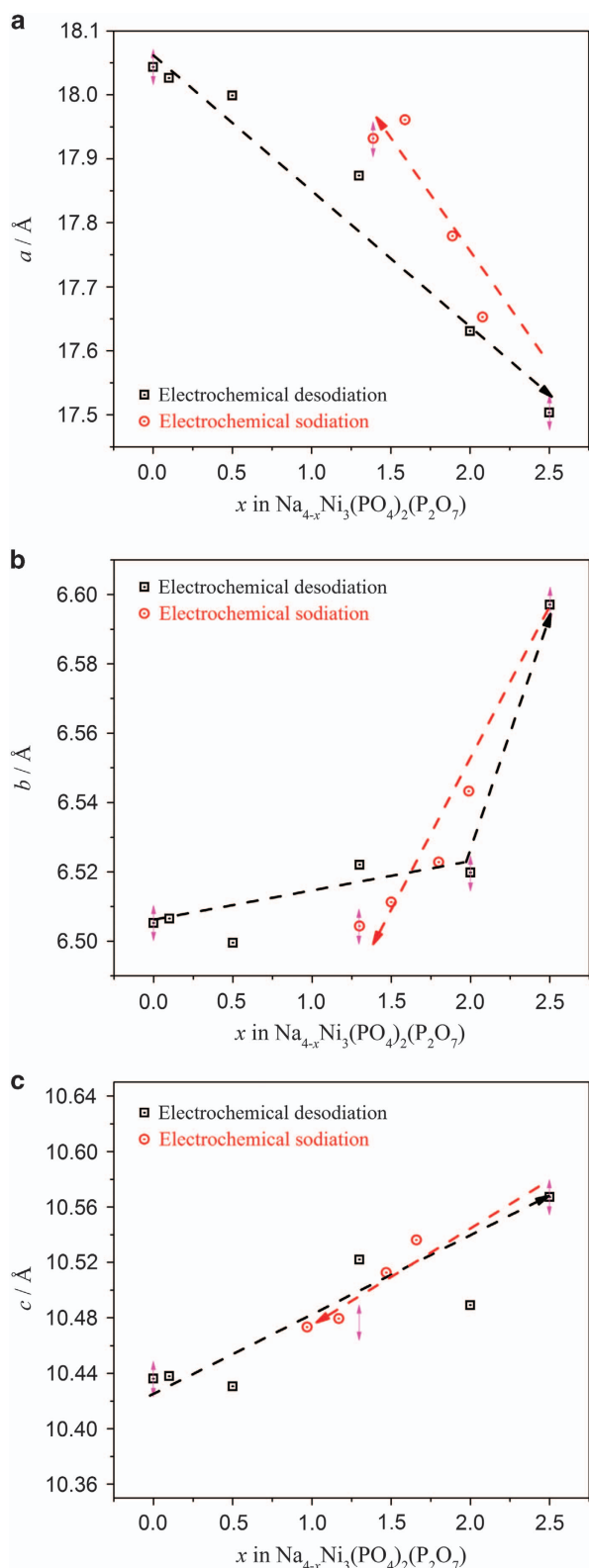


Figure 6 Lattice parameter (a , b , c) change during electrochemical desodiation/sodiation process.

moving to a higher 2θ value, other major peaks such as (011), (111), (220) and (022) show a slight shift to lower 2θ when the Na is extracted from the host material. During the discharge process, the peaks gradually move back to the original positions, suggesting the reversibility of the structural modifications associated with the volume change during the Na (de-)insertion process. To investigate the occupancy of the four different Na atoms upon sodium (de-)insertion, Rietveld refinement of the *ex situ* XRD results has been performed. The evolution of the refined lattice parameters (a , b , c) is shown in Figures 6a–c. The a lattice parameter value instantaneously decreases when the Na ions are electrochemically removed from the crystal structure, whereas the b and c axis parameters increase. The results herein presented differ slightly from those already reported for the Fe-based analog system, suggesting different structural changes in the potential range investigated.^{26,27} The Na occupancies, reported in Table S3, indicate that the sodium cations located in the large channels constructed by the P_2O_7 dimers, that is, the Na1 and Na4 sites, can be electrochemically removed first at relatively low potentials, resulting in the decrease of the a lattice parameter and leading to a volume change of 1.6%. When charging up to 5.0 V, the other two sodium cations, that is, from Na2 and Na3 sites, are extracted subsequently, leading to the rapid increase in the b parameter. When discharged back to 3.0 V, the observed lattice parameters are slightly different from those of the pristine sample, suggesting a not fully reversible electrochemical process, indicated as well by the electrochemical characterization above. The low reversibility may be ascribed to the collapse of $[\text{Ni}_3\text{P}_2\text{O}_{13}]_\infty$ layers along the a axis to a certain extent blocking the Na diffusion channels that confirms the instability of the desodiated electrode and reflects the low coulombic efficiency of the charge-discharge process. However, the superior thermal stability and small volume changes during cycling of the material demonstrate the advantages of using mixed polyanionic structures to construct a large diffusion channel for Na ions, proving this strategy to be an efficient way to develop Ni-based polyanionic materials as conceivable high voltage cathodes.

The integration of the two framework types in the same lattice, namely the $[\text{PO}_4]$ and $[\text{P}_2\text{O}_7]$ groups, enables the electrochemical activity of the $\text{Ni}^{2+/3+}$ redox reaction at high potential. This study represents a step forward in the design of high voltage cathode materials for Na-ion batteries. However, this research field faces several challenges. Electrochemically stable electrolytes and/or electrolyte additives, forming a cathode–electrolyte interphase with protective (electronic insulating, but ion-conducting) properties, thus mitigating or avoiding oxidative decomposition reactions, should be identified to enable such high voltage cathode materials in Na-ion batteries.

CONCLUSIONS

In summary, the $\text{Ni}^{2+/3+}$ redox activity in the mixed polyanionic framework has been experimentally demonstrated for the first time in sodium cells. Contrary to the pyrophosphate structure ($\text{Na}_2\text{NiP}_2\text{O}_7$), the combination of phosphate and pyrophosphate groups enables the reversible Na ion (de-)insertion accompanied by the $\text{Ni}^{2+/3+}$ redox activity. The $\text{Na}_4\text{Ni}_3(\text{PO}_4)_2(\text{P}_2\text{O}_7)$ has been synthesized via a conventional solid-state method. The carbon coating process and the implementation of an ionic liquid-based electrolyte enables the reversible (de-)insertion of 1.3 Na equivalents at an operating potential as high as 4.8 V vs Na^+/Na . Certainly, we are aware of the improvable electrochemical performance of the material; however, this is not the main objective of the present work. Meanwhile, searching for high voltage cathode materials opens a new and exciting perspective in the search for optimized and improved electrolyte solutions with a wider

electrochemical stability window. Nevertheless, we believe that this study holds promise for obtaining cathode materials of a high working potential. The mixed polyanionic structure, combining phosphate and pyrophosphate frameworks, allows, thanks to the P_2O_7 dimers, the construction of large Na diffusion channels and, in consequence, small volume changes during cycling, representing a step towards overcoming the drawbacks of Ni analogs with low ionic/electronic conductivity and poor structural stability. The work presented herein shows, for the first time, the realization of Ni-based mixed polyanionic compounds and paves the way for the discovery of new Na-based high potential cathode materials.

CONFLICT OF INTEREST

The authors declare no conflict of interest.

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