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PAPER

Synthesis and reactivity of an unexpected highly sensitive 1-carboxymethyl-3-diazonio-5-nitrimino-1,2,4-triazole†

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During the synthesis of functionalized energetic triazole derivatives for biosensor detection devices 1 carboxymethyl 3 diazonio 5 nitrimino 1,2,4 triazole (**2**) was obtained by nitration of sodium 1 carboxymethyl 3,5 diamino 1,2,4 triazole (**1**). Zwitterionic (**2**) behaves like a diazonium cation which was proved from its reaction with sodium azide by the formation of 3 azido 5 nitrimino 1,2,4 triazole (**3**). Compounds **1–3** were fully characterized by low temperature single crystal X ray diffraction, vibrational (IR/Raman) spectroscopy, NMR spectroscopy, mass spectrometry, elemental analysis and DSC measurements. The sensitivities were determined by the BAM drop hammer, BAM friction tester and an electrostatic discharge device. The extremely high impact sensitivity of the diazonio compound in comparison to its azido analogue is explained by the use of electrostatic potentials.

Introduction

In the area of energetic materials research, insensitivity, sustain ability and detection of explosives have gained considerable interest over the past years and especially ecological aspects have become more and more important. For example, ongoing research tries to find suitable lead free primary explosives in order to replace lead azide and lead styphnate in primary compositions. Moreover RDX shows significant eco and human toxicity and research is underway to find suitable alternatives for this widely used high explosive.

The increasing number of terrorist attacks and efforts using explosives has demonstrated the need for methods to screen a wide range of explosive materials and demonstrate whether or not they pose a threat to security. There is also a need to detect and quantify trace levels of explosives contamination on items or surfaces. These trace levels could alert investigators to the possible presence of covered dangerous materials.¹ The need for fast, highly specific, reliable tests for explosives that can be deployed on a massive scale and produced at a relatively low cost is an important challenge. Systems must be easy to use and require minimal training to operate. Colorimetric tests on a simple basis could provide part of the solution.² Certainly the chemical tests often have a deficiency in the selectivity and specificity demonstrated by biosensors.³ Biosensors use biological molecules usually an enzyme, antibody, or nucleic acid to detect sample molecules of interest⁴ via hydrogen bonding,

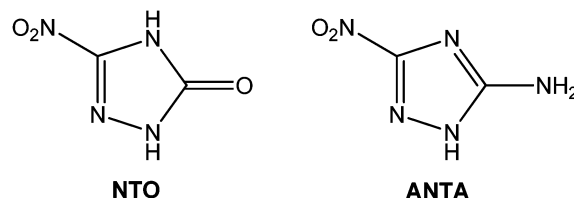


Fig. 1 Popular insensitive secondary explosives based on the 1,2,4 triazole moiety: NTO (5 nitro 1,2,4 triazol 3 one) and ANTA (3 amino 5 nitro 1,2,4 triazole).

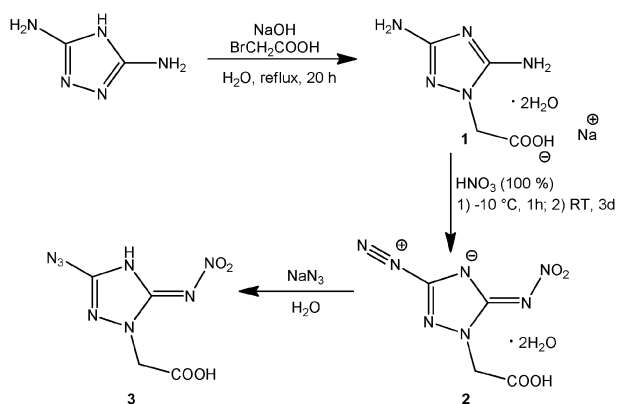
charge charge interactions, and other biochemical interactions. Thus the compounds require functional groups at first, a carboxy or carboxyalkyl group, so that the molecules could be linked to surfaces using amide bridges. Since many new energetic materials such as NTO (5 nitro 1,2,4 triazol 3 one)⁵ and ANTA (3 amino 5 nitro 1,2,4 triazole)⁶ (Fig. 1) are based on triazole heterocycles we tried to combine carboxyl groups with 1,2,4 triazoles connected to other energetic functionalities such as azide, nitro or nitramine groups. During the expected synthesis of 1 carboxymethyl 3,5 dinitrimino 1,2,4 triazole we obtained highly sensitive 1 carboxymethyl 3 diazonio 5 nitrimino 1,2,4 triazole (**2**), which was fully characterized and reacted to the corresponding azide derivative.

Results and discussion

Synthesis

In order to synthesize a triazole derivative with two different energetic groups attached, 3,5 diamino 1,2,4 triazole was functionalized adapting the literature procedure by Einberg,⁷

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Scheme 1 Synthesis of compounds **1–3**.

in which 3,5 diamino 1,2,4 triazole was reacted with bromo acetic acid and sodium hydroxide yielding sodium 1 carboxy methyl 3,5 diamino 1,2,4 triazole dihydrate (**1**). The neutral compound could not be obtained even after addition of hydrochloric acid until the pH was below 2. Afterwards **1** was reacted with nitric acid (100%, traces of NO₂) at –10 °C and stirred for three days at room temperature in order to obtain the corresponding 1 carboxymethyl 3,5 dinitrimino 1,2,4 triazole. After the mixture was poured onto ice and stored at 0 °C we obtained orange crystals of 1 carboxy methyl 3 diazonio 5 nitrimino 1,2,4 triazole dihydrate (**2**) (Fig. 3). This could be due to the presence of NO₂ in the nitration mixture. Usually diazonium cations are prepared by the reaction of amines and nitrous acid (nitrous acid is formed by the disproportionation of NO₂ in water). Attempts to vary the conditions for this reaction (temperature 0 °C, 5 °C; nitration mixtures; sulfuric acid (98%)/nitric acid (100%) 50 : 50) lead to lower yield of **2**. The expected dinitrated compound could not be isolated or detected. In order to investigate the reactivity of **2** and to introduce a second energetic group, **2** was reacted with sodium azide in aqueous solution to form 1 carboxymethyl 3 azido 5 nitrimino 1,2,4 triazole (**3**) in 85% yield (Scheme 1).

Crystal structures

Suitable single crystals of the described compounds (**1–3**) were picked from the crystallization mixture and mounted in Kel F oil, transferred to the N₂ stream of an Oxford Xcalibur3 diffractometer with a Spellman generator (voltage 50 kV, current 40 mA) and a KappaCCD detector using a λ_{MoKα} radiation wavelength of 0.71073 Å. All structures were measured at –100 °C. The data collection and data reduction were carried out with the CRYSTALISPRO software.⁸ The structures were solved with SIR 92,⁹ refined with SHELXL 97¹⁰ and finally checked using the PLATON software¹¹ integrated in the WINGX software suite.¹² The non hydrogen atoms were refined anisotropically and the hydrogen atoms were located and freely refined. The absorptions were corrected by a SCALE3 ABSPACK multi scan method.¹³

Sodium 1 carboxymethyl 3,5 diamino 1,2,4 triazole dihydrate crystallizes in the monoclinic space group *P*2₁/*c* with four molecules per unit cell. The density of 1.631 g cm^{–3} is slightly lower than that of neutral monoamino 1,2,4 triazole monohydrate.¹⁴ The triazole ring forms an aromatic planar system involving

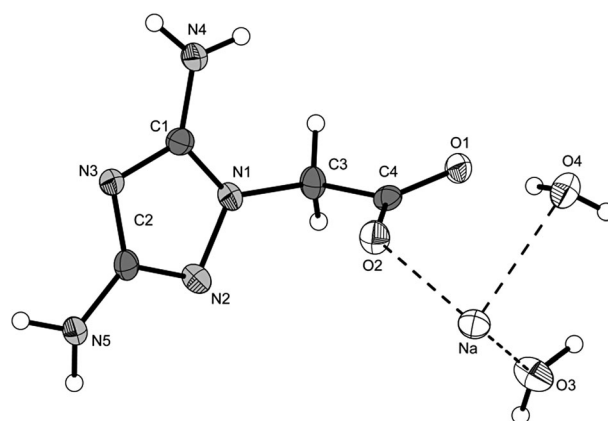


Fig. 2 Diamond plot of the molecular unit of **1**. Hydrogen atoms are shown as spheres of arbitrary radius and thermal displacements are set at 50% probability. Selected distances (Å): Na O3 2.319(1), Na O1 2.376(1), Na O2 2.392(1), Na O4 2.572(1), O1 C4 1.264(2), O2 C4 1.247(2), N1 C1 1.338(1), N1 N2 1.395(1), N1 C3 1.450(1), N2 C2 1.323(1), N3 C1 1.344(1), N3 C2 1.364(1), N4 C1 1.346(1), N5 C2 1.384(1), C3 C4 1.531(2); selected angles (°): O3 Na O1 92.1(1), O3 Na O2 91.7(1), O1 Na O2 100.1(1), O3 Na O4 84.7(1), O1 Na O4 176.6(1), O4 Na O4 86.2(1), O2 Na O4 79.3(1), O4 Na N5 78.5(1), C1 N1 N2 109.9(1), C1 N1 C3 129.3(1), N2 N1 C3 120.4(1), C2 N2 N1 101.5(1), C1 N3 C2 102.7(1), N1 C1 N3 110.1(1), N1 C1 N4 125.3(1), N3 C1 N4 124.6(1), N2 C2 N3 115.7(1), N2 C2 N5 122.1(1), N3 C2 N5 122.0(1), N1 C3 C4 113.9(1), O2 C4 O1 126.2(1), O2 C4 C3 118.8(1), O1 C4 C3 115.0(1).

both amino groups. The C4 atom of the carboxymethyl group is directed ~90° out of the plane. The molecular unit of **1** is depicted in Fig. 2. The carboxyl group is deprotonated and coordinated to the sodium atom. Two sodium atoms are linked by two water molecules and to one carboxyl group of the diamino molecule building a 3 dimensional network. Ribbons are formed by the parallel molecules, which are connected by the sodium ions. The triazole molecules are linked exclusively by the crystal water molecules.

The title compound 1 carboxymethyl 3 diazonio 5 nitrimino 1,2,4 triazole dihydrate (**2**) could be isolated in pure and crystalline forms from the reaction mixture. **2** crystallizes in the triclinic space group *P* $\bar{1}$ with two molecules in the unit cell (Fig. 3). The density of 1.671 g cm^{–3} is slightly higher in comparison with the starting material **1** (Table 1). Thus, the azo group seems not to have an observable effect on the density. Few zwitterionic diazonio triazole derivative can be found in the literature.^{15,16} However, to the best of our knowledge only alkylated imidazole diazonio compounds have been reported in ref. 17. The triazole ring and the nitrimino as well as the azo group form a planar system. The carboxy methyl group leans in the direction of the nitrimino group (~75° out of the ring plane). The nitro amino functionality should be described as nitrimine since the distance N4 C1 of 1.359(2) Å is significantly closer to a double bond than a C N single bond. The remaining proton is deprotonated balancing the zwitterionic molecule with the positive charge at nitrogen N6 and the negative one at the nitrogen atoms N3 or N4 (see electrostatic potential). The packing is again dominated by ribbons. Two molecules of **2** are packed together by their

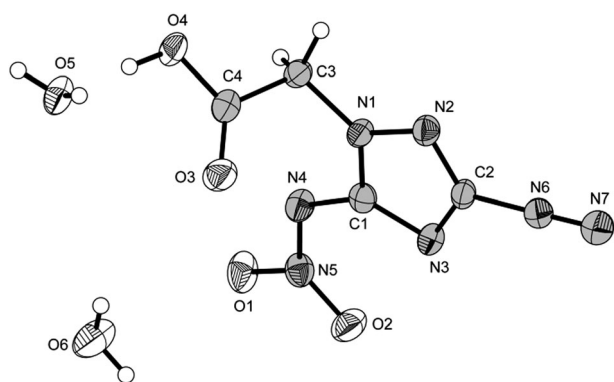


Fig. 3 Molecular unit of **2**. Hydrogen atoms are shown as spheres of arbitrary radius and thermal displacements are set at 50% probability. Selected bond distances (Å): N1 N2 1.350(2), N4 N5 1.329(2), N6 N7 1.088(2), N6 C2 1.398(2), N4 C1 1.359(2), N1 C1 1.367(2), N1 C3 1.441(2), C1 N3 1.342(2), N2 C2 1.309(2), N5 O2 1.234(2), N5 O1 1.260(2), N3 C2 1.337(2), O4 C4 1.307(2), O3 C4 1.202(2), C4 C3 1.512(2); selected angles (°): N7 N6 C2 179.5(2), N5 N4 C1 117.0(1), N2 N1 C1 110.9(1), N2 N1 C3 120.6(1), C1 N1 C3 128.5(1), N3 C1 N4 134.1(2), N3 C1 N1 109.8(1), N4 C1 N1 116.1(2), C2 N2 N1 99.0(1), O2 N5 O1 121.0(2), O2 N5 N4 124.5(1), O1 N5 N4 114.5(1), C2 N3 C1 99.4(1), O3 C4 O4 126.2(2), O3 C4 C3 123.1(2), O4 C4 C3 110.7(2), N2 C2 N3 121.0(1), N2 C2 N6 118.2(1), N3 C2 N6 120.8(1), N1 C3 C4 112.1(1); selected torsion angles (°): N5 N4 C1 N3 5.4(3), N5 N4 C1 N1 -177.3(1), C3 N1 C1 N4 3.5(2), C1 N4 N5 O2 4.9(2), N2 N1 C3 C4 -103.0(2), C1 N1 C3 C4 74.8(2), O3 C4 C3 N1 -5.7(2), O4 C4 C3 N1 175.3(1).

carboxy groups. The further connection is formed by hydrogen bonds involving both crystal water molecules.

1 Carboxymethyl 3 azido 5 nitrimino 1,2,4 triazole (**3**) was isolated after recrystallization from water. **3** crystallizes in the monoclinic space group $P2_1/c$ with four molecules per unit cell (Fig. 4). The density of 1.767 g cm^{-3} is higher compared to compound **2**. Due to the higher double bond character of the C1 N4 (1.343(2) Å) bond and the proton connected to nitrogen atom N3, **3** should also be described as a nitrimino triazole instead of nitr(o)amino triazole. The former proton at nitrogen atom N4 is now found to be connected to the ring nitrogen N3. In addition, the N4 C1 bond length of 1.343(2) Å is close to a typical C N double bond. The shortest C N bond in the ring system is between atoms N2 C2 (1.301(2) Å). The azide nitrogen distances (N6 N7 1.254(2) Å, N7 N8 1.118(2) Å) are in the range of typical bond lengths found in the literature for organic azides, e.g. for phenyl azide.¹⁸ In contrast to the structure of **2** the crystal packing is formed without inclusion of water which leads to a lower number of hydrogen bonds. This may be a reason for the lower observed decomposition temperature of $125 \text{ }^\circ\text{C}$, in comparison to the corresponding diazonium triazole **2** ($T_{\text{dec}} = 142 \text{ }^\circ\text{C}$). Intermolecular hydrogen bonds are observed between the carboxyl groups of different molecules in the structure of **3**. In addition, an intramolecular hydrogen bond formed by N3 H3...O1 is found. The packing is characterized by layers which are connected through the hydrogen bonds of the carboxyl groups. They interact in the interspace of the parallel orientated triazole rings.

Table 1 XRD data and parameters

	1	2	3
Formula	$\text{C}_4\text{H}_{10}\text{N}_5\text{NaO}_4$	$\text{C}_3\text{H}_7\text{N}_7\text{O}_6$	$\text{C}_4\text{H}_4\text{N}_8\text{O}_4$
$M/\text{g mol}^{-1}$	215.16	249.17	228.15
Crystal system	Monoclinic	Triclinic	Monoclinic
Space group	$P2_1/c$ (14)	$P1$ (2)	$P2_1/c$
Color/habit	Colorless cons	Colorless rods	Yellow block
Size/mm	$0.30 \times 0.18 \times 0.09$	$0.40 \times 0.11 \times 0.10$	$0.29 \times 0.16 \times 0.13$
$a/\text{Å}$	9.5655(9)	5.9153(5)	10.3123(7)
$b/\text{Å}$	8.3881(7)	9.1960(8)	6.8821(5)
$c/\text{Å}$	10.9272(11)	9.4344(8)	12.2305(9)
$\alpha/^\circ$	90.00	93.949(7)	90.00
$\beta/^\circ$	91.718(8)	90.503(7)	98.873(7)
$\gamma/^\circ$	90.00	104.609(7)	90.00
$V/\text{Å}^3$	876.36(14)	495.25(7)	857.61(11)
Z	4	2	4
$\rho_{\text{calc}}/\text{g cm}^{-3}$	1.631	1.671	1.767
μ/mm^{-1}	0.181	0.154	0.156
$F(000)$	448	256	464
T/K	173	173	173
θ min max/ $^\circ$	4.26, 23.99	4.27, 25.99	4.18, 25.98
Dataset [h ; k ; l]	-7:10, -9:9, -11:12	-7:7, -7:11, -11:11	-8:12, -8:8, -15:15
Reflections collected	2707	3351	4250
Independent reflections	1325	1923	1677
R_{int}	0.0284	0.0218	0.0286
Observed reflections	881	1195	1125
No. parameters	100	182	161
Restraints	0	0	0
R_1 (all)/ R_1 ($I > 2\sigma$)	0.0547/0.0313	0.0625/0.0322	0.0557/0.0330
wR_2 (all)/ wR_2 ($I > 2\sigma$)	0.0519/0.0492	0.0556/0.0513	0.0705/0.0665
Goof	0.801	0.802	0.859
Resd. dens./ $e \text{ Å}^{-3}$	-0.175, 0.177	-0.168, 0.211	-0.216, 0.170
Solution	SIR 92	SIR 92	SIR 92
Refinement	SHELXL 97	SHELXL 97	SHELXL 97
CCDC	843029	843030	843031

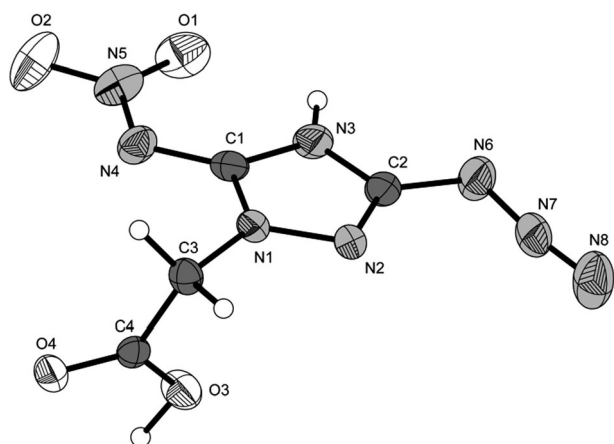


Fig. 4 Diamond plot of the molecular unit of **3**. Hydrogen atoms are shown as spheres of arbitrary radius and thermal displacements are set at 50% probability. Selected bond distances (Å): O4 C4 1.235(2), N1 C1 1.339(2), N1 N2 1.384(2), N1 C3 1.450(2), O3 C4 1.294(2), N4 C1 1.343(2), N4 N5 1.348(2), N2 C2 1.301(2), C4 C3 1.503(2), O1 N5 1.246(2), N3 C1 1.350(2), N3 C2 1.359(2), N7 N8 1.118(2), N7 N6 1.254(2), N6 C2 1.388(2), N5 O2 1.236(2); selected angles (°): C1 N1 N2 112.4(1), C1 N1 C3 125.2(1), N2 N1 C3 122.2(1), C1 N4 N5 115.9(1), C2 N2 N1 102.6(1), O4 C4 O3 125.3(2), O4 C4 C3 119.8(1), O3 C4 C3 114.9(1), C1 N3 C2 107.0(1), N1 C3 C4 110.9(1), N1 C1 N4 119.2(1), N1 C1 N3 105.0(1), N4 C1 N3 135.8(2), N8 N7 N6 170.6(2), N7 N6 C2 113.6(1), O2 N5 O1 121.5(2), O2 N5 N4 116.4(2), O1 N5 N4 122.1(2), N2 C2 N3 112.9(2), N2 C2 N6 128.2(2), N3 C2 N6 118.9(2); selected torsion angles (°): C1 N1 N2 C2 0.5(2), C3 N1 N2 C2 175.8(1), C1 N1 C3 C4 69.5(2), N2 N1 C3 C4 -105.1(2), O4 C4 C3 N1 -149.4(1), O3 C4 C3 N1 31.0(2), N2 N1 C1 N4 178.8(1), C3 N1 C1 N4 3.8(2), N2 N1 C1 N3 -0.7(2), C3 N1 C1 N3 -175.7(1), N5 N4 C1 N1 -179.0(1), N5 N4 C1 N3 0.3(3), C2 N3 C1 N1 0.5(2), C2 N3 C1 N4 -178.8(2), C1 N4 N5 O2 -176.3(1), C1 N4 N5 O1 3.7(2), N1 N2 C2 N3 -0.2(2), N1 N2 C2 N6 -178.4(2), C1 N3 C2 N2 -0.2(2), C1 N3 C2 N6 178.2(2), N7 N6 C2 N2 7.4(3), N7 N6 C2 N3 -170.7(2).

Spectroscopy

Sodium 1 carboxymethyl 3,5 diamino 1,2,4 triazole dihydrate **1** shows only the CH₂ signal in the ¹H NMR spectrum at 4.35 ppm as a singlet. The signals of the amino groups could not be observed in DMSO d₆ as solvent. The ¹³C NMR signals of **1** appeared at 169.9 ppm (deprotonated carboxy group), 160.0 ppm and 155.0 (triazole carbons) as well as at 47.5 ppm representing the CH₂ group.

The corresponding diazonium compound **2** shows also only one singlet in the ¹H NMR spectrum at a lower field (4.93 ppm) representing the CH₂ group. The carboxy group hydrogen could not be observed due to its very acidic character. In the ¹³C NMR spectrum most of the signals are shifted to higher fields, compared to the precursor compound **1**. Singlets at 167.4 ppm (COOH), 155.9 ppm (CNN), 151.2 ppm (C nitrimino) and 49.2 ppm (CH₂ group) were observed. The ¹H NMR spectrum of **3** also shows one singlet at 4.81 ppm, slightly shifted to a higher field. Most of the ¹³C NMR signals are shifted to higher fields except the carboxy group signal at 167.8 ppm. The singlets are observable at 150.0 ppm and 149.8 ppm, representing the triazole carbon atoms and at 48.9 ppm for the CH₂ group of the carboxymethyl moiety.

In the IR spectrum of the sodium salt **1** signals at 3362 cm⁻¹, 3278 cm⁻¹ and 3172 cm⁻¹ representing N O and N H vibrations are observable as well as signals for valence vibrations of C H at 2983 cm⁻¹ (Raman: 2984 cm⁻¹). The strong vibration at 1582 cm⁻¹ can be interpreted as the bending vibration of N H of the amino group. The signal at 1391 cm⁻¹ is assigned to the symmetric valence vibration of the negatively charged carboxyl group.¹⁹

The IR spectrum of the zwitterionic diazonium compound **2** shows O H vibrations at 3531 cm⁻¹ (water) and 3298 cm⁻¹. The vibration at 1730 cm⁻¹ can be assigned to the C O valence vibration of the carboxylic acid moiety. The diazonium (N₂⁺ substituent) valence vibration appears at 2274 cm⁻¹ as a strong signal (Raman: 2275 cm⁻¹). The asymmetric as well as the symmetric vibrations assigned to the nitrimino group are observable at 1522 cm⁻¹ and 1290 cm⁻¹. An almost similar IR and Raman spectrum is observed for compound **3**. Referring to this no vibration appears at 2274 cm⁻¹ but at 2153 cm⁻¹ reflecting the valence vibration of the azido group.¹⁹

Energetic properties

The sensitivities²⁰ of compounds **1–3** were investigated using the BAM drop hammer,²¹ BAM friction tester²¹ and a OZM electrostatic discharge device.²² Compound **1** is completely insensitive towards impact and friction stimuli.²³ In contrast, compounds **2** and **3** are very sensitive towards impact (**2**: < 1 J, **3**: 3 J) and should be handled with great care. With respect to the friction sensitivity **2** (20 N, very sensitive) is much more sensitive than **3** (144 N, sensitive). The electrostatic potentials of **2** and **3** were illustrated after computing an optimal structure at the B3LYP/6 31G* level of theory using the programs Gaussian09,²⁴ GaussView 5.0.8 (Fig. 5).²⁵ In many publications electrostatic potentials are used to predict sensitivities of energetic materials.^{26,27} Basically a higher charge separation leads to especially higher friction sensitivity values. Second, in contrast to non energetic organic molecules where the positive potential is larger but weaker in strength, in nitro and azo compounds usually more extensive regions with larger and stronger positive potentials are observed which can be related to the increased impact sensitivities. Especially a higher charge separation can be found for compound **2**, which could explain the much higher values in terms of impact and friction sensitivity.

The thermal behavior of **2** and **3** was investigated using a Linseis PT10 differential scanning calorimeter with a heating

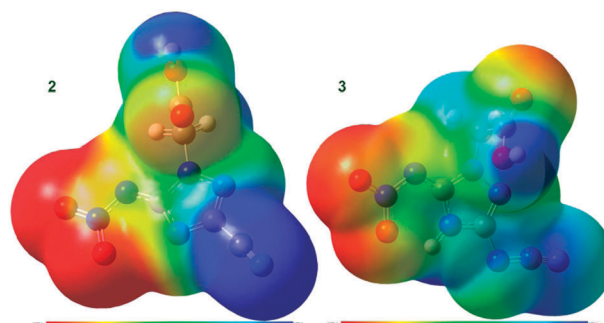


Fig. 5 Calculated (B3LYP/6 31G*) electrostatic potential of **2** (left) and **3** (right). The red regions represent electron rich regions, the blue regions extremely electron deficient regions.

Table 2 Overview of the sensitivity and performance data of compounds **1–3**

	1	2	3
Formula	C ₄ H ₁₀ N ₅ O ₄ Na	C ₄ H ₃ N ₇ O ₄ ·2H ₂ O	C ₄ H ₄ N ₈ O ₄
FW/g mol ⁻¹	215.14	249.14	228.13
IS ^a /J	> 100	< 1	3
FS ^b /N	> 360	20	> 144
Sieved grain size/μm	250 500	500 1000	< 100
ESD test ^c /J	1.50	0.012	0.20
N ^d [%]	32.55	39.35	49.12
Ω ^e [%]		-35.32	-42.08
T _{dec} ^f /°C	255	142	125
Density ^g /g cm ⁻³	1.631	1.671	1.767
Δ _f H _m ^h /kJ mol ⁻¹		-264	132
Δ _f U ⁱ /kJ kg ⁻¹		-960	664
EXPLO5.04 values			
-Δ _p U ^j /kJ kg ⁻¹		4315	4397
T _E ^k /K		3260	3556
p _C ^l /kbar		236	263
D ^m /m s ⁻¹		7674	7930
Gas vol. ⁿ /L kg ⁻¹		740	688

^a BAM drop hammer. ²⁰ ^b BAM methods. ²⁰ ^c Electric discharge tester. ^d Nitrogen content. ^e Oxygen balance. ^f Decomposition temperature from DSC (β = 5 °C). ^g From X ray diffraction. ^h Calculated (CBS 4M) heat of formation. ⁱ Energy of formation. ^j Energy of Explosion. ^k Explosion temperature. ^l Detonation pressure. ^m Detonation velocity. ⁿ Assuming only gaseous products.

rate of 5 degrees per minute.²⁸ Unfortunately, **2** and **3** have low thermal stabilities (T_{dec}: **2** 142 °C, **3** 125 °C), which probably excludes any practical applications, e.g. as novel lead free primary explosives.²⁹ Table 2 summarizes the physico-chemical properties of **1–3**. For the CHNO compounds the detonation parameters, which were calculated using the program EXPLO5 V5.04, are included. These values are based on the sum formula, heats of formation and the theoretical maximum density obtained by XRD. The heats of formation of **2** and **3** were calculated by the atomization method (see in ref. 30) based on CBS 4M electronic energies. The strongly negative value (Δ_fH_m^o -264 kJ mol⁻¹) of **2** is caused by the inclusion of two molecules of water in the crystal. The crucial calculated detonation parameters of **2** and **3** are higher in comparison to TNT (trinitrotoluene: D = 7253 ms⁻¹; p_{CJ} = 216 kbar), but lower than those of explosives like PETN (nitropenta: D = 8320 ms⁻¹; p_{CJ} = 320 kbar) and RDX (hexogen: D = 8748 ms⁻¹; p_{CJ} = 348 kbar).

Experimental part

All reagents and solvents were used as received (Sigma Aldrich, Fluka, Acros Organics) if not stated otherwise. Melting and decomposition points were measured with a Linseis PT10 DSC using heating rates of 5 °C min⁻¹, which were checked with a Büchi Melting Point B 450 apparatus. ¹H, ¹³C and ¹⁵N NMR spectra were measured with a JEOL instrument. All chemical shifts are quoted in ppm relative to TMS (¹H, ¹³C) or nitromethane (¹⁵N). Infrared spectra were measured with a Perkin Elmer Spektrum One FT IR instrument. Raman spectra were measured with a Perkin Elmer Spektrum 2000R NIR FT Raman instrument equipped with a Nd:YAG laser (1064 nm). Elemental analyses were performed with a Netsch STA 429 simultaneous thermal analyzer.

Caution!

The compounds described in this paper are energetic materials and tend to explode under certain conditions. Appropriate safety precautions should be taken at all times, especially when manipulating **2**. Laboratories and personnel should be properly grounded and safety equipments such as Kevlar[®] gloves, leather coats, face shields, and ear plugs are necessary.

Sodium 1-carboxymethyl-3,5-diamino-1,2,4-triazole dihydrate

(1). To a solution of 9.9 g (0.10 mol) 3,5 diamino 1,2,4 triazole and 8.0 g (0.20 mol) sodium hydroxide in 500 mL methanol at 60 °C was added slowly 15.3 g (0.11 mol) bromoacetic acid. After refluxing the solution for 16 hours the solvent was removed under vacuum. The beige colored residue was dissolved in 200 mL H₂O and the pH of the solution was reduced to less than 2 with concentrated hydrochloric acid (25 mL, 37%). After recrystallization from H₂O the beige solid was obtained and dried under vacuum. Yield: 87%. Mp: 255 °C (DSC, 5 K min⁻¹); ¹H NMR (DMSO *d*₆): 4.35 (*s*, 2H, CH₂); ¹³C NMR (DMSO *d*₆): 169.9 (COO), 160.0 (CNH₂), 155.0 (CNH₂), 47.5 (CH₂); Raman (200 mW, 25 °C, cm⁻¹): 3176 (20), 2984 (50), 2945 (86), 1671 (60), 1642 (45), 1465 (56), 1430 (47), 1394 (83), 1311 (66), 1270 (42), 1152 (40), 1078 (29), 1025 (43), 918 (63), 871 (73), 825 (23), 814 (28), 766 (39), 661 (100), 598 (37), 536 (40), 449 (40), 354 (64), 307 (58), 272 (52); IR (Diamond ATR, cm⁻¹): 3362 (m), 3278 (m), 3172 (m), 3063 (m), 2983 (m), 1772 (vw), 1664 (vs), 1605 (m), 1582 (s), 1533 (m), 1472 (m), 1429 (m), 1391 (s), 1305 (m), 1265 (m), 1151 (w), 1082 (w), 1020 (m), 960 (w), 911 (w), 868 (w), 762 (w), 684 (m), 660 (m), 614 (vw); elemental analysis C₄H₈N₅NaO₃ (197.13 g mol⁻¹) calc.: C 24.37, H 4.09, N 35.53%; found: C 24.40, H 3.83, N 34.91%; BAM drop hammer: >100 J (250 500 μm); BAM friction tester: >360 N (250 500 μm); ESD: 1.5 J (250 500 μm).

1-Carboxymethyl-3-diazonio-5-nitrimino-1,2,4-triazole dihydrate

(2). At -10 °C, 329 mg (1.5 mmol) of sodium 1 carboxy methyl 3,5 diamino 1,2,4 triazole dihydrate was slowly added to 1 mL nitric acid (100%). After stirring for 72 hours at RT the solution was poured onto ice. The orange red crystals, which formed after one day, were filtered and washed with Et₂O. Yield: 28%. Mp: > 58 °C (H₂O), > 93 °C (H₂O), > 142 °C (dec., DSC 5 K min⁻¹); ¹H NMR (DMSO *d*₆): 4.93 (*s*, 2H, CH₂); ¹³C NMR (DMSO *d*₆): 167.4 (COOH), 155.9 (CNN), 151.2 (CNHNO₂), 49.2 (CH₂); Raman (100 mW, 25 °C, cm⁻¹): 3005 (3), 2963 (5), 2275 (100), 1516 (35), 1473 (51), 1397 (8), 1358 (5), 1264 (26), 1201 (25), 1139 (50), 1038 (22), 1019 (6), 972 (10), 938 (4), 790 (4), 761 (9), 729 (2), 483 (3), 455 (4), 376 (3), 293 (7); IR (Diamond ATR, cm⁻¹): 3531 (m), 3298 (m), 3004 (m), 2529 (w), 2274 (s), 1992 (w), 1730 (s), 1604 (w), 1522 (s), 1468 (m), 1421 (m), 1413 (7), 1408 (m), 1347 (m), 1290 (s), 1229 (vs), 1134 (m), 1018 (m), 970 (w), 937 (w), 898 (w), 789 (m), 778 (m), 760 (w), 725 (w), 670 (w), 648 (w); EA: C₄H₇N₇O₆ (249.14 g mol⁻¹) calc.: C 19.28, H 2.83, N 39.35%; found: C 19.56, H 2.45, N 39.18%; BAM drop hammer: <1.0 J (500 1000 μm); BAM friction tester: 20 N (500 1000 μm); ESD: 12 mJ (500 1000 μm).

1-Carboxymethyl-3-azido-5-nitrimino-1,2,4-triazole (3). 250 mg (1.2 mmol) of 1 carboxymethyl 3 azo 5 nitrimino 1,2,4 triazole dihydrate was dissolved in 5 mL water. At 0 °C, 77 mg

(1.2 mmol) of sodium azide, dissolved in 5 mL water, was added slowly with stirring. After stirring for 20 hours at RT the water was removed under vacuum. The solid was dissolved in acetone, the salt was filtered off and the acetone was removed under vacuum to yield pure **3** as a yellow solid. After recrystallization from water yellow crystals suitable for X ray analysis were formed. Yield: 85%. Mp: 125 °C (dec., DSC, 5 K min⁻¹); ¹H NMR (DMSO *d*₆): 4.81 (s, 2H, CH₂); ¹³C NMR (DMSO *d*₆): 167.8 (COOH), 150.0 (CNN), 149.8 (CNN), 48.9 (CH₂); ¹⁴N NMR (DMSO *d*₆): -18 (NNO₂); Raman (100 mW, 25 °C, cm⁻¹): 2994 (4), 2956 (9), 2859 (6), 2159 (5), 1579 (100), 1510 (8), 1448 (11), 1423 (9), 1404 (6), 1318 (8), 1248 (5), 1135 (4), 1035 (31), 962 (4), 900 (9), 810 (12), 748 (4), 678 (3), 652 (4), 575 (3), 503 (3), 451 (4), 387 (4), 340 (7), 271 (15); IR (Diamond ATR, cm⁻¹): 3044 (m), 2992 (m), 2953 (m), 2558 (m), 2248 (w), 2153 (vs), 1749 (m), 1723 (m), 1563 (vs), 1508 (m), 1420 (s), 1397 (s), 1299 (s), 1238 (vs), 1218 (vs), 1165 (m), 1130 (m), 1033 (m), 1018 (m), 916 (m), 892 (m), 774 (m), 747 (m), 718 (m), 657 (m); elemental analysis: C₄H₄N₈O₄ (228.04 g mol⁻¹) calc.: C 21.06, H 1.77, N 49.12%; found: C 22.20, H 1.78, N 47.00%; BAM drop hammer: 3 J (<100 μm); BAM friction tester: >144 N (<100 μm); ESD: 0.2 J (<100 μm).

Conclusions

Energetic materials containing carboxyl substituents are used to develop biosensor detection devices for explosive and hazardous materials. The nitration at -10 °C of sodium 1 carboxymethyl 3,5 diamino 1,2,4 triazole dihydrate (**1**) yielded highly sensitive 1 carboxymethyl 3 diazonio 5 nitrimino 1,2,4 triazole (**2**), which was also characterized by single crystal X ray diffraction. The molecular structure shows a zwitterionic constitution which can be also drawn by expedient Lewis representation. The positive charge of the diazonio group is balanced by the negatively charged electronegative triazole ring. DFT calculation and the depiction of the electrostatic potential confirm this charge separation which causes high sensitivity towards impact and friction also in its dehydrated form. The reactivity of the diazonio substituent was explored by the reaction with sodium azide. After release of dinitrogen 1 carboxymethyl 3 azido 5 nitrimino 1,2,4 triazole (**3**) was formed which shows lower sensitivity values, but also in the range of primary explosives. However, the diazonio compound shows higher thermal stability (*T*_{dec.}: **2** 142 °C) in comparison to **3** (*T*_{dec.}: 125 °C). Calculated detonation parameters are higher than those of TNT but lower than those of PETN and RDX.

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