

Calorimetric study of block-copolymers of poly(*n*-butyl acrylate) and gradient poly(*n*-butyl acrylate-*co*-methyl methacrylate)

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Abstract

The nanophase separation in diblock and triblock copolymers consisting of immiscible poly(*n* butyl acrylate) (block **A**) and gradient copolymers of methyl methacrylate (MMA) and *n* butyl acrylate (*n*BA) (block **M/A**) were investigated by means of their heat capacity, C_p , as a function of the composition of the blocks **M/A** and temperature. In all copolymers studied, both blocks are represented by their C_p and glass transition temperature, T_g , as well as the broadening of the transition temperature range. The low temperature transition of the blocks **A** is always close to that of the pure poly(*n* butyl acrylate) and is independent of the analyzed compositions of the block copolymer, but broadened asymmetrically relative to the homopolymer due to the small phase size. The higher transition is related to the glass transition of the copolymer block of composition **M/A**. Besides the asymmetric broadening of the transition due to the phase separation, it decreases in T_g and broadens, in addition, symmetrically with increasing acrylate content. The concentration gradient is not able to introduce a further phase separation with a third glass transition inside the **M/A** block.

Keywords: Differential scanning calorimetry; Gradient copolymer; Block copolymer

1. Introduction

The blending of different polymers often results in improvement of properties exhibited by the individual components. Most polymers are incompatible, therefore, in blends, they do not dissolve. For thermodynamic reasons, the phases should separate on a macroscopic scale. Combining long sequences of the different polymers into block copolymers complicates this phase separation because the strong chemical bonds linking the different parts of a block copolymer must be located at the interface between the different phases. This strongly-bound interface causes the most change in the properties of the block copolymers when compared to the homopolymers of the same components. As a major result, the otherwise expected separation into macrophases of the chemically different, immiscible components is restricted in the block copolymer, and depending on the size of the blocks, microphase or even nanophase-separation is observed [1]. The need to minimize

the surface free energy of the phase-separated samples develops a rich variety of periodic morphologies [2].

For linear copolymers, two extreme structures can be inferred: diblock copolymers, composed of two connected, incompatible polymer chains, and random copolymers, where different types of monomers are distributed along the chains. Besides these, more complicated distributions of the monomers were also studied, e.g. tapered-block copolymers [3–5], multiblock copolymers [6], gradient copolymers [7, 8], and alternating, copolyoligomers [9,10]. The gradient copolymers are actually intermediate between the aforementioned two extreme cases. They have a well-defined structure and composition which changes gradually from predominantly sequences of one comonomer to the other as a function of the copolymer chain length. Due to this composition distribution, the repulsive inter-chain interactions are smoothly changing along the chain. This structure is different from the case of simple block copolymers, where the repulsive interactions are confined to the junction of the blocks, and random copolymers, where the repulsive interactions are distributed along the chain. These atypical interactions are expected to result in unique thermal properties for the gradient copolymers. Up to date,

however, only few and qualitative publications are available which report on thermal analysis of gradient copolymers [11,12]. In the present work we use conventional, quantitative DSC methods to study miscibility and phase separation based on heat capacity and glass transition temperatures for a series of diblock and triblock copolymers of poly(butyl acrylate) and gradient copolymers of butyl acrylate and methyl methacrylate with different compositions.

2. Experimental

The diblock and triblock copolymers studies are represented by the generalized formulae $A\text{-}M/A$ and $A/M\text{-}A\text{-}M/A$, respectively. The block M/A is a copolymer of methyl methacrylate (MMA) and *n*-butyl acrylate (*n*BA), while the block A consists of pure poly(*n*-butyl acrylate) (*Pn*BA). The copolymers were synthesized by atom transfer radical polymerization (ATRP) [13–15] starting from macroinitiators of *Pn*BA of well-defined lengths which are either monofunctional (for the $A\text{-}M/A$ diblock copolymers) or bifunctional (for the $A/M\text{-}A\text{-}M/A$ triblock copolymers). The block M/A is not a statistical copolymer, but it is a gradient copolymer with higher MMA concentration than the average at the junction point with A and higher *n*BA concentration than the average at the free ends. In the copolymerization stage of the synthesis, MMA is initially consumed at a faster rate, so that the residual *n*BA concentration increases and leads to the higher *n*BA concentration towards the chain ends, thus, forming a natural gradient block.

2.1. Synthesis of macroinitiator

A flask was loaded with CuBr, 1,4-dimethoxybenzene and a stir bar, which was then vacuum/backfilled with N_2 (three times). Next, purged *n*-butyl acrylate (*n*BA) and N,N,N',N'' -pentamethyldiethyltriamine (PMDETA) were added, via N_2 purged syringes. The solution was stirred until homogeneous, and the appropriate initiator was added (methyl-2-bromopropionate monofunctional; dimethyl-2,6-dibromoheptadionate difunctional) via a purged syringe. An initial sample was taken, and the flask was placed in an 353 K oil bath for 240 min. Three samples were dissolved in $CDCl_3$, acetone, and THF for NMR, GC, and GPC analysis, respectively. The remaining sample was dissolved in acetone and passed through a column of alumina, excess solvent was removed followed by precipitation in 50:50 $H_2O/MeOH$, and finally dried under vacuum.

2.2. Synthesis of gradient copolymers

A flask was loaded with *n*BA macroinitiator and a stirrer bar, which was put under vacuum for at least 30 min. Purged anisole was added to dissolve the polymer. A second flask

was loaded with CuCl, bis(2-pyridylmethyl)-octadecylamine (BPMODA) and a stirrer bar, which was then vacuum/backfilled with N_2 (three times). Purged MMA and *n*BA were added to the second flask in accordance with the desired characteristics of the gradient segment. The solution was stirred until homogeneous, and then transferred by syringe to the first flask, which was then placed into an 353 K oil bath for 24 h. Three samples were dissolved in $CDCl_3$, acetone, and THF for NMR, GC, and GPC analysis, respectively. The remaining sample was dissolved in acetone and passed through a column of alumina, excess solvent was removed followed by precipitation in 50:50 $H_2O/MeOH$, and finally dried under vacuum. The structural characterization of the copolymers used are shown in Table 1. The composition is known from synthesis and quantitative NMR. The molar masses were measured by exclusion chromatography in THF (1 g l^{-1}) calibrated with polystyrene. The data were rounded to two digits. Homopolymers for comparative study were poly(*n*-butyl acrylate) (Aldrich, $M_w = 99,000$ Da) and poly(methyl methacrylate) (Scientific Polymer Products, Inc., $M_w = 35,000$ Da).

2.3. Calorimetric characterization

The DSC study was carried out using a power-compensated Perkin Elmer DSC7, known to yield heat capacities with a precision of about $\pm 1\%$ [16]. Dry N_2 gas with a flow rate of $20\text{ cm}^3\text{ min}^{-1}$ was purged through the DSC cell. Cooling was accomplished using a mechanical refrigerator (IntraCooler). Typical sample masses for the standard DSC experiments were 10–20 mg. The samples were weighed on a Cahn-28 automatic electro-balance with an accuracy of ± 0.001 mg.

The as-prepared samples were first heated to 423 K and kept at this temperature for 5 min to erase any prior thermal history. Then the samples were cooled to 205 K to set the structure to be analyzed by subsequent heating to 423 K. The heating and cooling rates were 10 K min^{-1} . The heat-flow rate was initially calibrated with the heat of fusion of indium (28.45 J g^{-1}). The heat capacity obtained was then refined by correction with a reference run of sapphire over the same temperature range as the sample [17]. The calorimeter asymmetry between the empty reference and sample calorimeters was eliminated with an empty-pan run used as a baseline for the heat-flow rate of the sample and calibration runs.

The glass transition temperature was chosen at the 50% change of the heat capacity which is close to the point of inflection. Also observed were the extrapolated temperatures of the beginning (T_1) and end of the glass transition (T_2), taken at the intersection of the tangent at the point of inflection with the extrapolated heat capacities of the solid and liquid phases, respectively. These temperature evaluations were made based on the experimental data, before comparison with data from the ATHAS data bank, as will be discussed below. The difference $T_2 - T_1$ is a measure for

Table 1
Characterization of the samples of block copolymers

Sample	Degree of polymerization			Composition (mol%)		Composition of block M/A MMA/n BA (mol%)		General composition (mol%)		M_n (kDa) (by GPC)	PDI
	Block A PnBA	BlockM/A MMA/nBA		Block A PnBA	Block M/A MMA/n BA	M/A	MMA	n BA	MMA		
		BlockM/A MMA/nBA	MMA/nBA								
<i>Diblock (A-M/A)</i>											
A-M/A1	166	206/0		45	55	100/0		45	55	42	1.27
A-M/A2	182	219/10		44	56	96/4		47	53	54	1.12
A-M/M/A3	182	209/19		44	56	92/8		49	51	58	1.17
A-M/A4	182	207/42		42	58	83/17		52	48	57	1.15
<i>Triblock (A/M-A-M/A)</i>											
A/M-A-M/A1	164	174/0		49	51	100/0		49	51	38	1.26
A/M-A-M/A2	164	198/12		44	56	94/6		47	53	42	1.21
A/M-A-M/A3	164	180/20		45	55	90/10		51	49	42	1.21
A/M-A-M/A4	164	180/56		41	59	76/24		55	45	46	1.20

the breadth of the glass transition. The asymmetry of the glass transitions was characterized, in addition, whenever possible, by finding the temperatures T_b and T_c , which were taken as the temperatures below and above the glass transition where the measured heat capacity begins to deviate from a linear the baseline of the glassy phase, and attains the linear baseline of the liquid phase, respectively. From the good fit to the data bank data on the homopolymers [18], one can estimate that the precision in heat capacity reached that common for the used instrumentation ($\pm 1\%$). Multiple measurements (3-5) were made on all samples.

3. Results

In Fig. 1(a-d), the heat capacities are shown for the four A-M/A-copolymers, measured after cooling from the melt. In the figures, the squares (\square) represent the averaged data-bank heat capacities of the homopolymers of the components calculated from the recommended heat capacity tables available in the ATHAS Data Bank [18]. In the region between glass transitions of PnBA and PMMA all PnBA repeating units were taken to be in the liquid state, whereas all PMMA repeating units were taken to be solid. The crosses (+) in Fig. 1(b-d) between the two glass transitions were calculated for the case that only the part of PnBA contained in the block A is in the liquid state, whereas in the gradient section, both the PnBA and the PMMA repeating units in the block M/A are solid, a case more likely, at least on the low-temperature side of the region between the glass transitions. The dashed and dotted lines of reference represent the sums of the heat capacities of the proper fractions of homopolymers in the solid and liquid states, respectively, excluding the transition effects by extrapolation into the temperature range of interest. These lines, together with the continuous recordings of the DSC experiment, were used as baselines for the calculation of the glass transition parameters which are listed in Table 2.

At temperatures below the glass transition of PnBA, when both components are in the solid state, and in the region above T_g of PMMA, when both components are in the liquid state, the heat capacities of the A-M/A-copolymers exhibit fully additive behavior. The experimental heat capacity data for the region of temperatures below T_g of PnBA are not completed in the figures, because the temperature range of the calorimeter is limited when using the IntraCooler as a cooling accessory. Heat capacities of all copolymers, however, were also investigated at lower temperatures (with lower precision) using liquid nitrogen cooling (see Fig. 1(a)). It has been found that ultimately the heat capacities are in good agreement with the baselines drawn from the ATHAS Data Bank using the assumption of additivity.

Both PnBA and PMMA are non-crystallizable, amorphous polymers, and all their block copolymers demonstrate

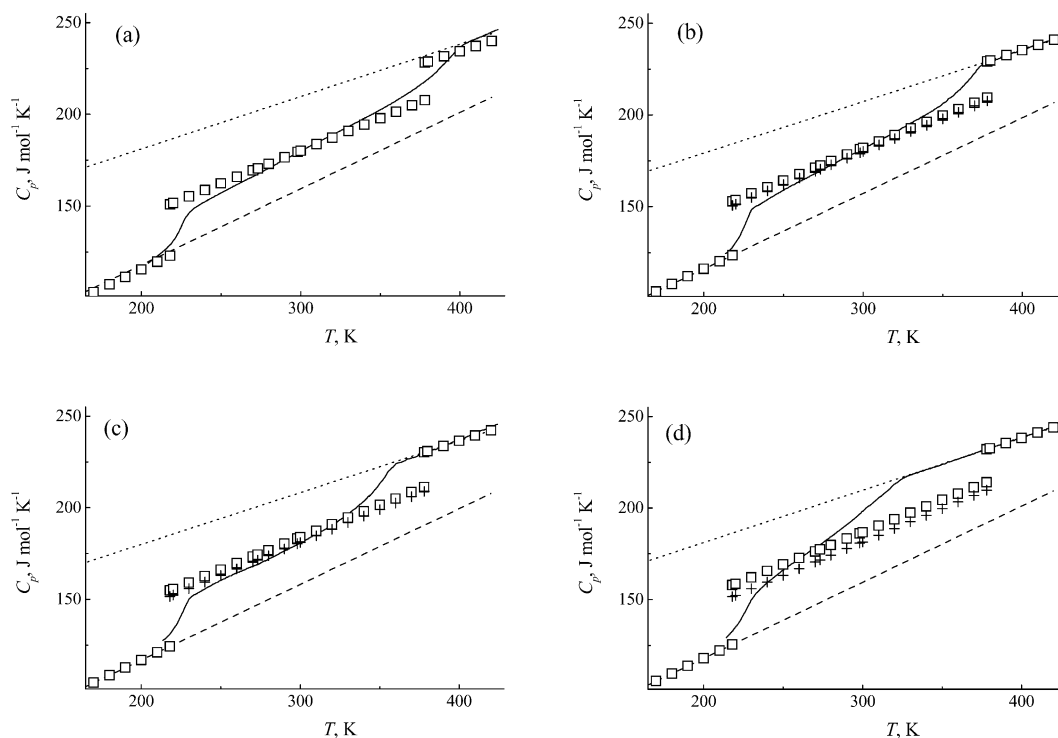


Fig. 1. Heat capacities of the **A M/A** copolymers as a function of temperature. Parts (a)–(d) display the data on **A M/A1**, **A M/A2**, **A M/A3**, and **A M/A4**, respectively, as shown in Table 1. For transition parameters see Table 2.

two glass transition temperatures. The parameters of the two glass-transitions are summarized in Table 2 and compared to the also measured transitions of the homopolymers. All low-temperature glass transition temperatures of the eight **A M/A** and **A/M A M/A** copolymers are very similar at 225 ± 2.5 K and agree with the newly measured *PnBA*. They somewhat exceed the T_g of *PnBA* in the ATHAS Data Bank, which is 218 K [18]. The jump in the heat capacity in the region of the first T_g of the copolymers is not directly comparable because of different content of *nBA* in the

blocks in the copolymer. The value of the first ΔC_p per mole x_1 of *nBA* from the block **M/A**, however, is only 49 ± 4 J mol⁻¹ K⁻¹ for the **A M/A**-copolymers and 51 ± 6 J mol⁻¹ K⁻¹ for all eight copolymers. This value is lower than the ΔC_p at the glass transition for pure *PnBA*, which is 61.4 J mol⁻¹ K⁻¹ according to Table 2, and 62.3 J mol⁻¹ K⁻¹ according to the literature [18]. The reason for the low value is the asymmetric broadening of the glass transition due to the size effect, to be described below.

The second glass transition temperature in the **A M/A1**

Table 2
Parameters of the glass transitions in the investigated homopolymers and their diblock and triblock copolymers

Sample	Block A							Block M/A						
	T_g (K)	ΔC_p (J mol ⁻¹ K ⁻¹)	$\Delta C_p/x_1$ (J mol ⁻¹ K ⁻¹)	T_1 (K)	T_2 (K)	T_c (K)	$T_2 - T_1$ (K)	T_g (K)	ΔC_p (J mol ⁻¹ K ⁻¹)	T_b (K)	T_1 (K)	T_2 (K)	T_c (K)	$T_2 - T_1$ (K)
<i>Homopolymers</i>														
<i>PnBA</i>	223	61.4	61.4	217	229	236	12							
<i>PMMA</i>								383	36.3	316	375	390	390	15
<i>Diblock (A M/A)</i>														
A M/A1	224	20.1	44.7	218	230	240	12	388	15.6	355	375	402	403	28
A M/A2	224	21.2	48.2	218	231	246	13	365	13.8	335	354	375	376	21
A M/A3	223	22.5	51.1	215	231	243	16	347	21.2	301	333	361	362	28
A M/A4	224	22.8	54.3	214	235	248	21	305	12.4	275	284	325	327	41
<i>Triblock (A/M A M/A)</i>														
A/M A M/A1	222	28.1	57.3	212	231	247	18	374	13.4	339	356	392	395	36
A/M A M/A2	225	23.1	52.5	215	236	249	21	358	10.6	332	344	371	375	27
A/M A M/A3	228	23.4	52.0	213	242	254	29	335	10.0	302	317	352	354	35
A/M A M/A4	229	18.0	43.9	209	248	254	39					319		

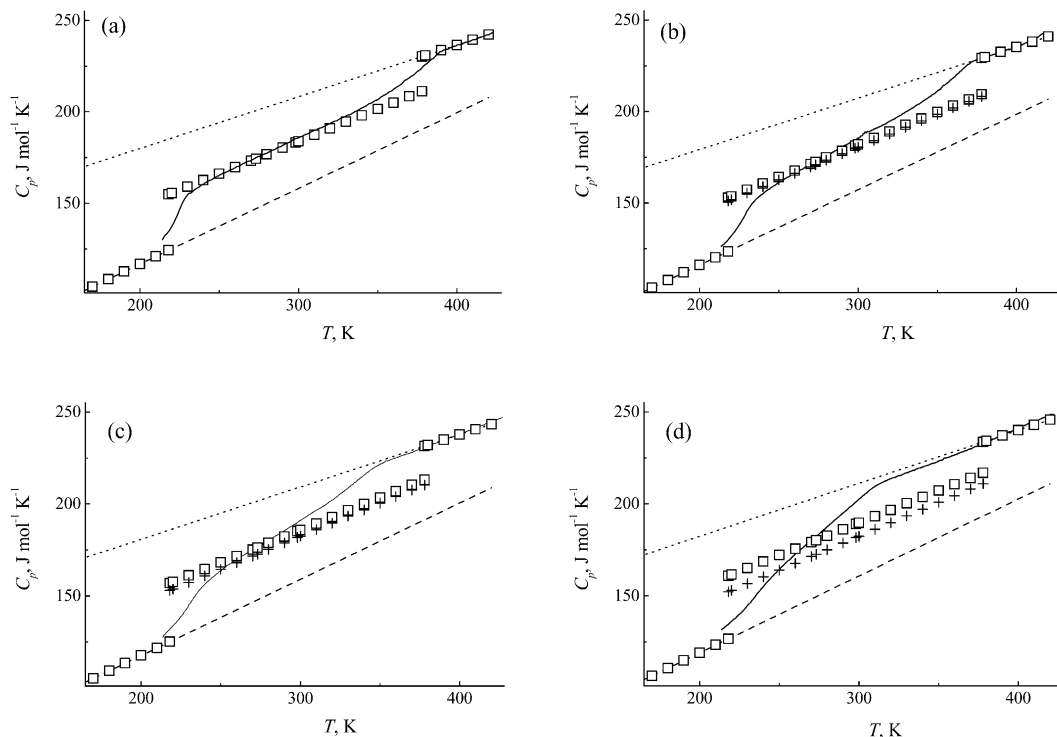


Fig. 2. Heat capacity of the A/M-A-M/A copolymers as a function of temperature. Parts (a)–(d) display the data on A/M-A-M/A1, A/M-A-M/A2, A/M-A-M/A3, and A/M-A-M/A4, respectively, as shown in Table 1. For transition parameters see Table 2.

and A/M-A-M/A1 samples at 388 and 374 K are not far from the T_g of pure PMMA of 383 K as one would expect for pure PMMA blocks. The ATHAS Data Bank lists 378 K [18] for the T_g of PMMA. Again, as will be shown below, the asymmetry of the glass transition may introduce the somewhat larger spread of the PMMA glass transitions. In the other three A-M/A samples, the second T_g is much lower than the T_g of PMMA. With increase in the content of *n*BA in the M/A block, it shifts to the region of lower temperatures, towards the T_g of P*n*BA. In addition, the second T_g broadens with an increase in content of *n*BA in the M/A block, as seen when inspecting Fig. 1(b)–(d).

The heat capacity plots for the four A/M-A-M/A-copolymers are shown in Fig. 2(a)–(d). The calculated lines are generated as described for Fig. 1. The behavior of this series of copolymers is similar to that of the A-M/A-copolymers. Their heat capacities are additive, both, in the region of temperatures below the glass transition of P*n*BA where both components are in the solid state, and in the region above T_g of PMMA, where both components are in the liquid state. All the A/M-A-M/A-copolymers also demonstrate two glass transition temperatures, as listed in Table 2. All the first glass transitions of the copolymers are similar to the A-M/A copolymers, as pointed out in the description of Fig. 1.

The second glass transition behaves also similar to the A-M/A-copolymers. Its temperature decreases and the region of glass transition broadens with increase in the *n*BA-content in the block M/A. The glass transition region

in the A/M-A-M/A4 copolymer is so broad, that it is hard to distinguish the two transitions. From the discussion of the heat capacity plot, however, it becomes clear that this copolymer still has two glass transitions, i.e. it remains incompatible over the full range of concentration.

4. Discussion

There are two major effects which can influence the glass transitions in block copolymers: (A) the sizes of the phases, and (B) the solubility of the components in each other. If both blocks of the copolymer are incompatible, the segment lengths determine the sizes of the phases as microphase or nanophase and the separating phases will have a morphology that depends on their size-ratio [2].

The glass transition temperatures, T_g , of both components are affected by the continuing molecules that cross the interface. If the second phase is less mobile, the end of the glass transition stretches to higher temperature as one approaches the interface. If, on the other hand, the second phase is more mobile, the beginning of the glass transition at the interface is broadened to lower temperature. When analyzing the two glass transitions of a block copolymer, one finds, thus, that the beginning of the overall glass transition region at the lower temperature and the end of glass transition at higher temperature will be constant, but the two midpoints of the transitions, the glass transition temperatures move towards each other because of the

asymmetric broadening of the transition. This effect should be seen by the changes in T_b and T_c . If the phase size is in the micrometer range or larger, this asymmetric broadening of the glass transition region is small because of the negligible specific interface area, as was shown by an earlier, quantitative analysis of the styrene/ α -methylstyrene block copolymer system [19].

In the present case, however, the Pn BA sections which define the interface for both types of the analyzed block copolymers have an average of only 171 repeating units, as can be seen from Table 1. This length corresponds to an extended-chain length of approximately 45 nm. The RMS end-to-end chain length of a random coil of this length in the melt or glass is about 10 nm, a value calculated by assuming the typical expansion coefficient of the freely-jointed, random-flight, mean-square end-to-end distance for an acrylate to its real dimensions to be about 7–10. The molar fraction of the Pn BA segments is not far from 0.5 for all samples, which results in a lamellar morphology [2] which allows to quantify the interface. Only a small amount of ordering is expected due to the location of the Pn BA segment-ends at the interface, i.e. the lamellar phases approach a nanophase thickness with a specific surface area of as much as $200 \text{ m}^2 \text{ g}^{-1}$ when assuming the RMS end-to-end length of 10 nm is a measure of the lamellar thickness. Such a large surface-to-volume ratio can lead to a substantial change in the breadth of the glass transition [19].

If partial solubility exists for one component in the other, or if both components are partially soluble in each other, either one or both of the glass transition temperatures, T_g , will move towards the other, respectively. In addition the corresponding ΔC_p must change according to the solubility. Furthermore, in the dissolution of longer sequences of identical repeating units, the glass transition region broadens symmetrically about T_g . This is not the case in random copolymers where the glass transition also changes, but the breadth of the glass transition does not exceed that of the homopolymers, as was shown for example for increasingly brominated poly(2,6-dimethyl-1,4-phenylene oxide) [20]. For complete solubility of both components in block copolymers, as well as for blends of homopolymers, there is only one, broad glass transition and the beginning of the lower and the end of the upper glass transition move closer to each other, but never reaches the narrow glass transition range of a random copolymer, as was also documented on the styrene/ α -methylstyrene system of homopolymers of varying molar mass by DSC [21].

The two glass transitions in the DSC traces of the samples investigated in this paper which consist of blocks of pure Pn BA and PMMA of approximately equal contents (**A M/A1** and **A/M A M/A1**) show reasonable agreement with the glass transition temperatures of the homopolymers, suggesting that there is little miscibility between the components. The constancy of T_1 for block **A** of all nine homo- and copolymers, and T_2 and T_c for the PMMA and the **M/A1** and **A/M A M/A1** copolymer proves similarly

pure Pn BA and PMMA phases, at least within the interior of the phases. The Figs. 1(a) and 2(a), however, indicate a strong asymmetry between the glass transitions, seen best by the crossing of the measured and calculated heat capacities at about 300 K. For truly identical behavior, the measured C_p should follow the calculated C_p , marked by \square in the temperature region between the glass transitions. This change in slope of the measured heat capacity leads to the noticeable decrease in $\Delta C_p/x_1$ in Table 2 and masks the changes in T_c of **A** and T_b of **M/A**. Comparing these data to the similar analysis in the literature of a poly(styrene-block- α -methyl styrene) this should be caused by the smallness of the phases [19]. The mobile **A**-phase is bound to the glassy **M/A**-phase, i.e. the outer layers of n BA have their mobility reduced and a part of the glass transition moves to higher temperature. The opposite effect is observed for the **M/A**-phase where the surface attachment to the liquid **A**-phase lowers the glass transition, not because of solubility, but because of molecular mobility of the surface layer of the PMMA.

Additional changes in the glass-transition behavior are observed in the other three **A M/A** samples shown in Fig. 1. They also exhibit two glass transitions. The low-temperature T_g is, again, in reasonable agreement with T_g of Pn BA. This allows to suggest that the first T_g belongs to the blocks **A** with little change from **A M/A1** to **A M/A4**. In these copolymers the blocks of 160–180 n BA units are still phase-separated and only influenced by the rigidity of the second block due to the smallness of the phase. The value of $T_2 - T_1$ increases to double the value of Pn BA (see Table 2), which goes parallel with an increase in $\Delta C_p/x_1$. Both the value of $T_2 - T_1$ and $\Delta C_p/x_1$ are, thus, connected to the asymmetry caused by the small phase size. At the same time, the higher T_g , attributed to the block **M/A**, shifts to lower temperatures with higher concentration of n BA, as one would expect for such copolymers (see Tables 1 and 2). A comparison of T_g , T_b , and T_c of the gradient block copolymers with **A M/A1** reveals that this broadening is less asymmetric than the size effect, all three temperatures move to lower values, as seen best from Fig. 1(c) and (d). The gradient blocks in the copolymer with changing n BA concentration exhibit, thus, a broadening in the glass transition due to the copolymerization, but there is no additional phase separation within the **M/A**-block. The samples of **A/M A M/A2** to **A/M A M/A4** of Fig. 2 are again, as pointed out above, showing an almost constant behavior of the lower glass transition. With higher n BA content in the **M/A**-blocks, it gets increasingly difficult to establish the now overlapping glass transitions.

5. Conclusions

This calorimetric analysis has shown for the first time that it is possible to separate size and solubility effects for block copolymers. The size-effect leads to an asymmetric

broadening of the glass transition. The solubility (copolymerization) leads to a shift in the glass transition, and, as long as there are homopolymeric sections in the copolymer, a symmetric broadening of the transition is superimposed. Furthermore, these first measurements allow the speculation that with higher precision in the calorimetry, as is possible by using multi-frequency temperature-modulated DSC [22], a more detailed description of the phase-contours within the samples may be possible. Of particular interest would be the analysis of *n*BA/MMA gradient copolymers which have sufficiently long *Pn*BA and PMMA ends to be able to produce a layer structure with a diffuse interphase.

Acknowledgments

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References

[1] Chen W, Wunderlich B. Nanophase separation of small and large molecules, invited feature article. *Macromol Chem Phys* 1999;200:283.

- [2] Matsen MW, Bates FS. *J Chem Phys* 1997;106:2436.
- [3] Zielinski RP, Childers CW. *Rubber Chem Technol* 1968;41:161.
- [4] Kraus G, Rollman KW. *Angew Makromol Chem* 1971;16/17:271.
- [5] Hashimoto T, Tsukahara Y, Tachi Y, Kawai H. *Macromolecules* 1983;16:648.
- [6] Legge NR, Holden G, Schroeder HH, editors. *Thermoplastic elastomers*. Munich: Hanser; 1987.
- [7] Greszta D, Matyjaszewski K, Pakula T. *Polym Prepr ACS* 1996;37(1):569.
- [8] Pakula T, Matyjaszewski K. *Macromol Theory Simul* 1996;5:987.
- [9] Di Lorenzo ML, Pyda M, Wunderlich B. *J Polym Sci, Part B: Polym Phys* 2001;39:1594.
- [10] Di Lorenzo ML, Pyda M, Wunderlich B. *J Polym Sci, Part B: Polym Phys* 2001;39:2969.
- [11] Matyjaszewski K, Greszta D, Pakula T. *Polym Prepr ACS* 1997;38(1):707.
- [12] Matyjaszewski K, Ziegler M, Arehart SV, Greszta D, Pakula T. *J Phys Org Chem* 2000;13:775.
- [13] Wang J S, Matyjaszewski K. *J Am Chem Soc* 1995;117:5614.
- [14] Matyjaszewski K, Xia J. *Chem Rev (Washington DC)* 2001;101:2921.
- [15] Matyjaszewski K, Shipp DA, McMurray GP, Gaynor SG, Pakula T. *J Polym Sci, Part A: Polym Chem* 2000;38:2023.
- [16] Wunderlich B. Heat capacity of polymers. In: Cheng SZD, editor. *Handbook of thermal analysis and calorimetry*, vol. 3. Amsterdam: Elsevier; 2002.
- [17] Archer DG. *J Phys Chem Ref Data* 1993;22:1441.
- [18] Pyda M, editor. *Advanced thermal analysis system (ATHAS). DATA Bank*, <http://web.utk.edu/~athas>.
- [19] Gaur U, Wunderlich B. *Macromolecules* 1980;13:1618.
- [20] Bopp RC, Gaur U, Kambour RP, Wunderlich B. *J Therm Anal* 1982;25:243.
- [21] Lau SF, Pathak J, Wunderlich B. *Macromolecules* 1982;15:1278.
- [22] Kwon YK, Androsch R, Pyda M, Wunderlich B. *Thermochim Acta* 2001;367/368:203.