

The Influence of Protonation on the Electroreduction of Bi (III) Ions in Chlorates (VII) Solutions of Different Water Activity

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Abstract We examined the electroreduction of Bi (III) ions in chlorate (VII) solutions under varied protonation conditions of the depolariser using voltammetric and impedance methods. The results of the kinetic parameter correlation lead to the statement that the changes in the amount of chloric (VII) acid against the amount of its sodium salt in the supporting electrolytes of the low water activity have a significant influence on the rate of Bi (III) ion electroreduction. The increase of the concentration of chloric acid sodium salt, as well as the chloric (VII) acid alone within the particular concentration of the supporting electrolyte, inhibits the process of Bi (III) ion electroreduction. It should be associated with the reorganisation of the structure of the double layer connected with the slow dehydration inhibited by ClO_4^- ions. The standard rate constants k_s values with the increase of the chlorate (VII) concentrations for all the solutions examined of chlorates (VII) confirms the catalytic influence of the decrease of water activity on the process of Bi (III) ion electroreduction. The multistage process is confirmed by the non-rectilinear $\ln k_f = f(E)$ dependences.

Keywords Electrochemistry · Bi (III) electroreduction · Protonation · Double layer · Kinetic parameters · Catalytic activity

Introduction

The electrochemical properties of metal ions depend on the composition and concentration of the supporting electrolyte. The metal cations in aqueous solutions show strong interactions with water molecules. The dehydration steps play a big role in the deposition reactions [1].

The studies concerning the reduction process of Bi (III) in chlorates (VII) solution of different water activity point at strong interactions between a depolariser ion and water molecules [2, 3].

The hydrolysis of Bi (III) ions reduces the pH range in which you can study their electroreduction. It has to be emphasised that aqua ions $[\text{Bi}(\text{H}_2\text{O})_9]^{+3}$ only exist in strongly acidic noncomplexing solutions. In solutions with $\text{pH} > 0$, oxo and hydroxo complexes of BiOH^{+2} , $\text{Bi}(\text{OH})_2^+$ and BiO_{aq}^+ or condensed structures such as Bi_2O^{+4} , $\text{Bi}_6\text{O}_6^{+6}$, $\text{Bi}_6(\text{OH})_4^{+6}$, and $\text{Bi}_6\text{O}_6(\text{OH})_3^{+3}$ [1].

According to Lovrić et al., [3] the reduction of Bi (III) includes three partial dehydration steps. Additionally, in the first stage of Bi (III) electroreduction, the cations of the supporting electrolyte participate and they absorb as well the released water molecules in their hydration spheres, which can explain the inversely proportional dependence of the reaction rate on water activity. The studies conducted in 1–8 mol dm^{-3} chlorates (VII) [4] point at the significant role of water activity in the Bi (III) electroreduction process. The values of kinetic parameters determined in the examined solutions point at the

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slight influence of supporting electrolyte concentration on the kinetics of Bi (III) electroreduction in 1–3 mol dm⁻³ chlorates (VII), whereas in 4–8 mol dm⁻³ chlorates (VII) the values of the determined kinetic parameters indicate a significant increase of reversibility of Bi (III) ion electroreduction, with the increase of chlorates (VII) concentration. The character of the changing rate of constants, in the function of the potential, indicates at a multistage process of Bi (III) ions electroreduction, as well as a different mechanism of Bi (III) electroreduction in the solutions with low water activity in comparison with the solutions with high water activity [4].

The subject matter of the research will be the electroreduction of Bi (III) ions in chlorates (VII). The issue of Bi (III) ion electroreduction in weak complexing solutions entails the aspect of practical research. Mainly the possibility of directing and pointing towards increased accuracy of Bi (III) ion determination.

In the experiments, the chloric (VII) acid to sodium chlorate (VII) concentration ratios in the supporting electrolyte were varied, leading to various forms of the studied depolariser.

The methodology of the study is based on the electrochemical methods (voltammetry, Faradaic impedance).

Experimental

Technique of Measurement

The measurements were performed in a three-electrode cell containing the following: a dropping or hanging mercury-electrode with a controlled increase rate and a constant drop surface (0.014740 cm²), as a working electrode (MTM Poland); Ag/AgCl as a reference electrode and a platinum spiral, as an auxiliary electrode. The polarographic, voltammetric and impedance measurements were carried out in thermostated cells at 298 K by using an Autolab Fra 2/GPES (Version 4.9) frequency response analyser (Eco Chemie, Utrecht, Netherlands). The solutions were deaerated using nitrogen, which was passed over the solutions during the measurements. Analytical grade chemicals from Fluka were used.

The 2–7 mol dm⁻³ chlorates (VII) solutions of concentration ratio HClO₄:NaClO₄ such as (1:1) solution A, (1:4) solution B, (1:9) solution C, (4:1) solution D, (9:1) solution E were studied.

The enumerated solutions were designated according to the scheme, e.g.

- 3A Designates 3 mol dm⁻³ chlorates (VII), where HClO₄:NaClO₄=1:1
 3B Designates 3 mol dm⁻³ chlorates (VII), where HClO₄:NaClO₄=1:4
 3C Designates 3 mol dm⁻³ chlorates (VII), where HClO₄:NaClO₄=1:9
 3D Designates 3 mol dm⁻³ chlorates (VII), where HClO₄:NaClO₄=4:1
 3E Designates 3 mol dm⁻³ chlorates (VII), where HClO₄:NaClO₄=9:1

- 3C Designates 3 mol dm⁻³ chlorates (VII), where HClO₄:NaClO₄=1:9
 3D Designates 3 mol dm⁻³ chlorates (VII), where HClO₄:NaClO₄=4:1
 3E Designates 3 mol dm⁻³ chlorates (VII), where HClO₄:NaClO₄=9:1

The concentration of Bi (III) ions in the solutions studied was always 1 · 10⁻³ mol dm⁻³. Due to the weak solubility of Bi (NO₃)₃ in chlorates (VII), the solutions were sonicated.

In the DC polarography, SWV and CV voltammetry, the optimal experimental operating conditions were as follows: step potential 2 mV for DC, puls amplitude 20 mV, frequency 120 Hz and step potential 2 mV for SWV, and scan rate 5–1000 mVs⁻¹ and step potential 5 mV for CV. Impedance data were collected at 24 frequencies in the range from 200 to 50,000 Hz within the faradaic potential region at 10 mV intervals.

Elaboration of Experimental Data

The approximate diffusion coefficient (D_{ox}) Bi (III) ions in the studied solutions were calculated using the Ilkovič equation for diffusion-controlled limiting current. The DC waves of Bi (III) in 1–8 mol dm⁻³ chlorates (VII) solutions were used as a standard [4]. The formal potentials (E_f^0) of the electrode processes and values of the kinetic parameters (αn_α and k_s) were calculated based on voltammetric measurements. The details are described elsewhere [4]. According to the reversibility parameter of the electrode process (based on the dependence of the potential difference of anodic E_{pa} and cathodic E_{pc} peaks on the value $0.057/n$), the standard rate constants k_s were determined using two different equations. For the quasi-reversible processes, the k_s values were determined using the method elaborated by Nicholson [5] according to the equation:

$$\psi = \left(\frac{D_{\text{ox}}}{D_{\text{red}}} \right)^{\alpha/2} \frac{k_s (RT)^{1/2}}{(\pi n F \nu D_{\text{ox}})^{1/2}} \quad (1)$$

The function ψ was determined from the product of electron number exchanged in the electrode process (n) and the difference between the potentials of anodic and cathodic peaks ($E_{\text{pa}} - E_{\text{pc}}$), and its dependence on $n(E_{\text{pa}} - E_{\text{pc}})$ was tabled [6].

For the irreversible processes, the values k_s , which are dependent on the kinetic parameters, are described by the equation [4]:

$$E_{\text{pc}} = E_f^0 - \frac{RT}{\alpha n_\alpha F} \left[0.78 - \ln k_s + \ln \sqrt{D_{\text{ox}} b} \right] \quad (2)$$

where:

$$b = \frac{\alpha n_{\alpha} F v}{RT}$$

The values of the apparent rate constants k_f of Bi (III) ion electroreduction in the chlorates (VII) solutions as a function of the potential were calculated from impedance measurements. The details are described elsewhere [4].

Results and Discussion

Polarographic and Voltammetric Measurements

Figure 1 presents SWV peaks of Bi (III) electroreduction in 2–7 mol dm⁻³ chlorates (VII), of the concentration ratio HClO₄:NaClO₄=1:1. With the increase of chlorates (VII) concentration, the SWV peaks of Bi (III) increase and are shifted towards the positive potentials. These results are inversely proportional to the water activity. These changes are practically identical with the results obtained in the earlier work, which points to the fact that the replacement of half of the chloric (VII) acid by the sodium chlorate (VII) has no influence on the height and position of the peaks.

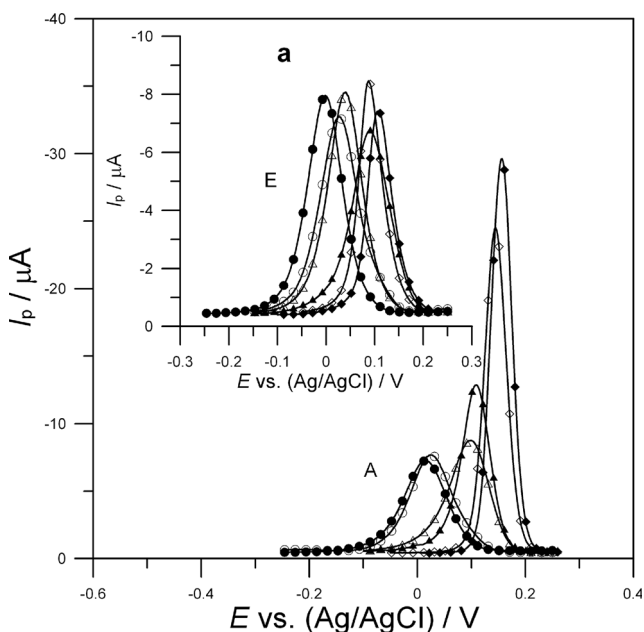


Fig. 1 The SWV peaks of the electroreduction of $1 \cdot 10^{-3}$ mol dm⁻³ Bi (III) in 2–7 mol dm⁻³ chlorates (VII), where HClO₄:NaClO₄=1:1 (A). The concentration of chlorates (VII) in mol dm⁻³: (○) 2; (●) 3; (Δ) 4; (▲) 5; (◇) 6; (◆) 7. The SWV peaks of the electroreduction of $1 \cdot 10^{-3}$ mol dm⁻³ Bi (III) in 2–7 mol dm⁻³ chlorates (VII), where HClO₄:NaClO₄=9:1 (E). The concentration of chlorates (VII) in mol dm⁻³: (○) 2; (●) 3; (Δ) 4; (▲) 5; (◇) 6; (◆) 7

Figure 1a presents SWV peaks of Bi (III) electroreduction in 2–7 mol dm⁻³ chlorates (VII), but at the concentration ratio HClO₄:NaClO₄=9:1. The changes of the peak potentials of Bi (III) electroreduction are similar to those in Fig. 1. It suggests that the composition of the active complex is independent on the concentration ratio HClO₄ and NaClO₄; whereas the peaks' height is practically unchanged with the change of chlorates (VII) concentration, which points to the fact that in the solutions with a large excess of chloric (VII) acid in comparison with sodium chlorate (VII), the rate of Bi (III) ion electroreduction is practically independent from the chlorate (VII) concentration.

In Fig. 2a, the dependences of SWV electroreduction peaks of $1 \cdot 10^{-3}$ mol dm⁻³ Bi (III) in 3 mol dm⁻³ chlorates (VII) (Fig. 2) and in 7 mol dm⁻³ chlorates (VII) (Fig. 2a) of the concentration ratios HClO₄:NaClO₄: 1 (A); 1:4 (B); 1:9 (C); 4:1 (D); 9:1 (E) were shown. In 3 mol dm⁻³ chlorates (VII), the changes in the concentration ratios HClO₄:NaClO₄ do not significantly influence the course of the curves $I_p = f(E)$. However, in 7 mol dm⁻³ chlorates (VII), the highest electroreduction peaks of Bi (III) are observed at HClO₄:NaClO₄=1:1. The increase of NaClO₄ concentration in the supporting electrolyte causes the decrease of the peak (Fig. 2, curves B, C) and the shift towards the positive potentials. The increase of HClO₄ concentration in chlorates (VII) solution (Fig. 2 curves D, E) causes the further decrease of

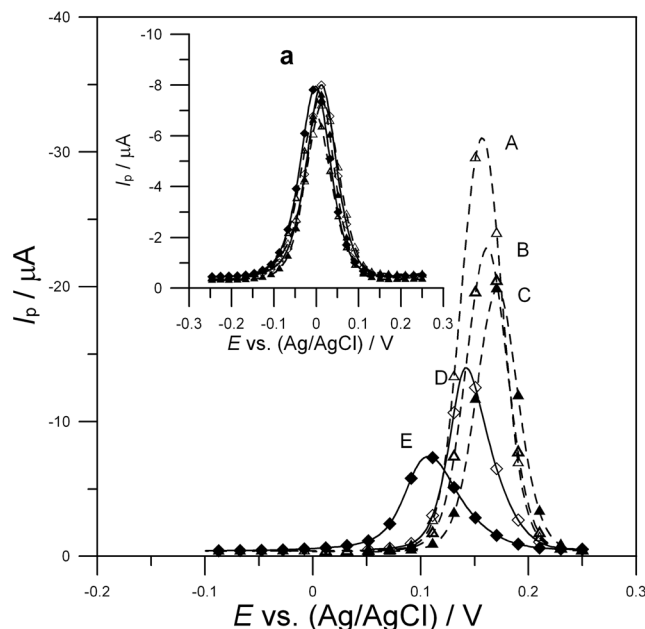


Fig. 2 The SWV peaks of the electroreduction of $1 \cdot 10^{-3}$ mol dm⁻³ Bi (III) in 7 mol dm⁻³ chlorates (VII), where HClO₄:NaClO₄=1:1 (—) 7A; HClO₄:NaClO₄=1:4 (—) 7B; HClO₄:NaClO₄=1:9 (—) 7C; HClO₄:NaClO₄=4:1 (—) 7D; HClO₄:NaClO₄=9:1 (—) 7E. The SWV peaks of the electroreduction of $1 \cdot 10^{-3}$ mol dm⁻³ Bi (III) in 3 mol dm⁻³ chlorates (VII), where HClO₄:NaClO₄=1:1 (—) 3A; HClO₄:NaClO₄=1:4 (—) 3B; HClO₄:NaClO₄=1:9 (—) 3C; HClO₄:NaClO₄=4:1 (—) 3D; HClO₄:NaClO₄=9:1 (—) 3E

SWV peaks of Bi (III) ions electroreduction, whereas the peak potential is shifted towards the more negative potentials. This suggests that the increase of HClO₄ concentration (curves D and E) causes significant changes in the composition of the active complex, which implicates the decrease of the rate of the Bi (III) ions electroreduction process.

The possibility of the formation of ionic pairs, e. g. Bi (III)—ClO₄⁻ has to be mentioned. If the electrode surface is charged negatively, the ionic pairs can favour decreasing the electroreduction rate [7].

The influence of water activity on the Bi (III) electroreduction process in (2–7 mol dm⁻³) chlorates (VII) for different content of NaClO₄ and HClO₄ also results from the course of the chronovoltammetric curves CV (Fig. 3a). With the increase of the chlorates (VII) concentration from 2 to 7 mol dm⁻³ in the solutions A, B and C of the supporting electrolytes, the decrease of ΔE_{ac} between the anodic and cathodic peaks is observed, which testifies to the increase of the reversibility of Bi (III) electroreduction process. Whereas for the solutions D and E, where the dominance of HClO₄ acid against NaClO₄ is increasing, the changes of ΔE_{ac} with the decrease of water activity are low.

Figure 4 presents CV curves of Bi (III) ions electroreduction in 6 mol dm⁻³ chlorates (VII), but for the different content of chloric acid and its sodium salt (6A, 6B, 6C, 6D, 6E). It should be noted that the anodic peaks of Bi (III) ions electroreduction are higher comparing to the cathodic peaks. The potentials of cathodic and anodic peaks are shifted towards the more positive potentials with the increase

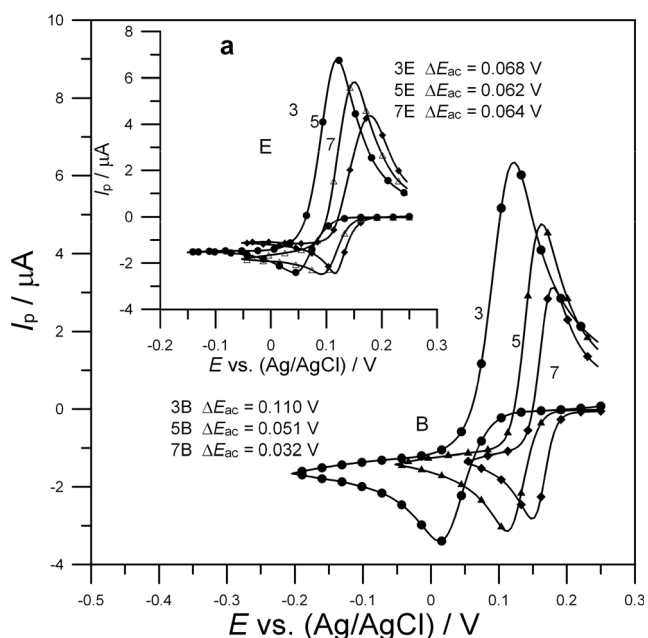


Fig. 3 Cyclic voltammogramme of $1 \cdot 10^{-3}$ mol dm⁻³ Bi (III) in chlorates (VII), where HClO₄:NaClO₄=1:4 (B). The concentration of chlorates (VII) in mol dm⁻³: (●) 3; (▲) 5; (◆) 7. Cyclic voltammogramme of $1 \cdot 10^{-3}$ mol dm⁻³ Bi (III) in chlorates (VII), where HClO₄:NaClO₄=9:1 (E). The concentration of chlorates (VII) in mol dm⁻³: (●) 3; (▲) 5; (◆) 7

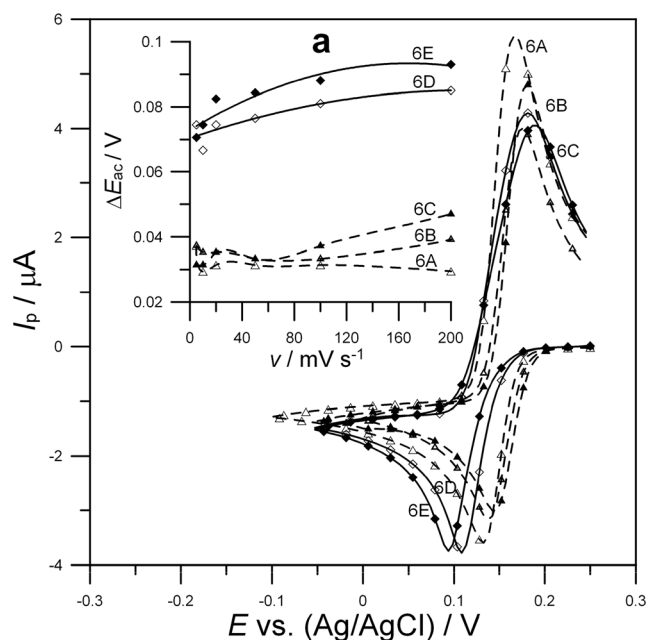


Fig. 4 The cyclic voltammogramme of $1 \cdot 10^{-3}$ mol dm⁻³ Bi (III) in 6 mol dm⁻³ chlorates (VII), where HClO₄:NaClO₄=1:1 (—) 6A; HClO₄:NaClO₄=1:4 (---) 6B; HClO₄:NaClO₄=1:9 (—) 6C; HClO₄:NaClO₄=4:1 (---) 6D; HClO₄:NaClO₄=9:1 (---) 6E. The influence of polarisation rate on the difference between the potentials of the anodic and cathodic peaks for the Bi (III)/Bi (Hg) couple in 6 mol dm⁻³ chlorates (VII), where HClO₄:NaClO₄=1:1 (—) 6A; HClO₄:NaClO₄=1:4 (---) 6B; HClO₄:NaClO₄=1:9 (—) 6C; HClO₄:NaClO₄=4:1 (---) 6D; HClO₄:NaClO₄=9:1 (---) 6E

of the amount of NaClO₄ to HClO₄ (6A, 6B, 6C). Instead, in the solutions with the preponderant concentration of chloric (VII) acid in the supporting electrolyte solution (6D, 6E), the shift of the cathodic peaks potentials towards more negative potentials is observed, whereas the anodic peaks are shifted towards the more positive potentials. Such oscillations of the cathodic and anodic peaks potentials in the function of the changes of NaClO₄ and HClO₄ concentration ratios in the supporting electrolyte suggest the differences in the mechanism of Bi (III) ions electroreduction, as well as the significant dependence of the active complex structure on the composition of the supporting electrolyte [3, 4].

The increase in both NaClO₄ (6A, 6B, 6C solutions) and HClO₄ (6D, 6E solutions) concentrations results in the increase in the distance between the anodic and cathodic peak potentials ΔE_{ac} :

6A $\Delta E_{ac} = 0.029$; 6B $\Delta E_{ac} = 0.036$; 6C $\Delta E_{ac} = 0.040$; 6D $\Delta E_{ac} = 0.074$; 6E $\Delta E_{ac} = 0.094$. This is evidence for the inhibiting effect on the kinetics of the studied electrode process. The increase of the concentration of sodium salt of the chloric (VII) acid against the HClO₄ concentration in the supporting electrolyte causes lower effects when comparing with the supporting electrolyte with the bigger amount of chloric (VII) acid compared with its sodium salt. In the solutions of the concentrated electrolytes (4–7 mol dm⁻³) of

chlorates (VII), the number of “free water molecules” decreases in the aftermath of the hydration process. In acid noncomplexing electrolyte solutions, the $\text{Bi}(\text{H}_2\text{O})_9^{3+}$ ion is described by the very low rate of hydration water release. Thereupon the cumulative electrode process will also be consisted of the chemical stages leading to the labilisation of $\text{Bi}(\text{H}_2\text{O})_9^{3+}$ hydration shell [8]. The dependence plot of the potential difference of the anodic and cathodic peaks ΔE_{ac} on the electrode polarisation rate (ν) (Fig. 4a) confirms these assumptions. In all the solutions of the supporting electrolyte (6A, 6B, 6C, 6D, 6E), the slight changes of ΔE_{ac} at low polarisation rates ($5\text{--}100\text{mv s}^{-1}$) are observed, which expresses the fact that the stage controlling the electroreduction rate of Bi (III) ions is the chemical reaction. This is certainly the reaction of Bi (III) ions dehydration mentioned before. The studies of Eyring [9] and Zelič [10] et al., concerning the rate of In (III) ions electroreduction in the solutions of noncomplexing electrolytes, confirm our assumptions.

It should be noticed as well that in the solutions 6A, 6B, 6C with the preponderant amount of NaClO_4 , the shape of $\Delta E_{\text{ac}}=f(\nu)$ (Fig. 4a) is completely different than in the solutions 6D, 6E, where the amount of chloric (VII) acid prevails. Such behaviour suggests differences in the electrode mechanism [4].

The research by Nazmutdinov et al. [11, 12], in which the quantum mechanical theory was used to describe the reduction of multivalent ions (e.g. In (III)), suggests the existence of the hydrolysed forms of Bi (III) ions in aqueous solutions. Moreover, the hydrolysed forms of Bi (III) ions can be more active as compared with Bi (III) aquacomplexes [12].

In the case of $[\text{Bi}(\text{H}_2\text{O})_9]^{3+}$, the acceptor molecular orbital is localised mostly on the Bi atom, which leads to relative slight dependence of the activation energy values on the first step of electron transfer. However, as the case of Bi (III) aquahydroxocomplex, a more strong decreasing the activation energy values starting from certain region of distances was observed as compared with $[\text{Bi}(\text{H}_2\text{O})_9]^{3+}$ [11, 12]. This adiabatic effect explains a high electrochemical activity of the Bi (III) aquahydroxocomplexes which may compete with Bi (III)

aquacomplexes in electroreduction and affect the changes in the mechanism of the process.

The formal potentials E_f^0 and the kinetic parameters αn_α and k_s (Table 1) were determined from the chronovoltammetric measurements.

The results of the parameters correlation lead to the statement that the changes in the amount of chloric (VII) acid against the amount of its sodium salt in the supporting electrolytes of the low water activity have a significant influence on the rate of Bi (III) ions electroreduction, particularly for the solutions A, B, C. It should be noticed that with the increase of the concentrations of both NaClO_4 (A, B, C) and HClO_4 (D, E) in the solution of supporting electrolyte, the standard rate constants k_s of Bi (III) ions electroreduction decrease. The direction of changes of the E_f^0 values shift suggests the differences in the mechanism of Bi (III) ions electroreduction in the solutions A, B, C of chlorates (VII), when comparing them with the solutions D, E.

The Impedance Measurements

The values of apparent rate constants k_f were obtained based on charge-transfer resistance [10] as a function of the potential. The increased values of the charge-transfer resistance R_{amin} determined at the formal potential (Table 2), clearly demonstrate the inhibitory effect of the supporting electrolyte (D, E).

It can be seen, however, that the distinct decrease of the charge-transfer resistance values, with the increase of the chlorates (VII) concentrations for all the solutions examined of chlorates (VII) (A, B, C, D, E), confirms the catalytic influence of the decrease of water activity on the process of Bi (III) ions electroreduction.

The dependences $\ln k_f=f(E)$ for all the chlorates concentrations studied are not linear (Fig. 5), and the curves' slopes change with the change of the potential and chlorates (VII) concentration. Some characteristics of k_f change in the function of the potential points at the existence of the multistage process of Bi (III) ions electroreduction [13–19]. It also

Table 1 The values of formal potentials (E_f^0), cathodic transition coefficients (αn_α), standard rate constants (k_s) of electroreduction of $10^{-3}\text{ mol dm}^{-3}$ Bi (III) in chlorates (VII) solutions of concentration ratio

Chlorate (VII)	2 mol dm^{-3}			4 mol dm^{-3}			6 mol dm^{-3}			7 mol dm^{-3}		
	E_f^0/V	αn_α	$10^4 k_s/\text{cm s}^{-1}$	E_f^0/V	αn_α	$10^4 k_s/\text{cm s}^{-1}$	E_f^0/V	αn_α	$10^4 k_s/\text{cm s}^{-1}$	E_f^0/V	αn_α	$10^4 k_s/\text{cm s}^{-1}$
A	0.102	0.28	1.43	0.125	0.40	24.4	0.150	0.70	61.7	0.160	0.74	100.3
B	0.097	0.33	1.50	0.131	0.41	13.5	0.157	0.60	46.6	0.158	0.69	90.6
C	0.095	0.35	1.52	0.125	0.40	12.4	0.158	0.54	27.1	0.168	0.60	85.1
D	0.098	0.31	1.52	0.111	0.39	10.4	0.143	0.40	10.9	0.153	0.43	9.91
E	0.102	0.27	1.28	0.110	0.39	7.13	0.140	0.39	6.91	0.146	0.40	6.61

$\text{HClO}_4:\text{NaClO}_4$ (1:1) solution A, (1:4) solution B, (1:9) solution C, (4:1) solution D, (9:1) solution E

Table 2 The values of formal potentials of electroreduction of $1 \times 10^{-3} \text{ mol dm}^{-3}$ Bi (III) in chlorates (VII) solutions of concentration ratio $\text{HClO}_4:\text{NaClO}_4$: (1:1) solution A, (1:4) solution B, (1:9) solution C, (4:1) solution D, (9:1) solution E, as well as the values of the charge transfer

resistance (R_{amin}) Bi (III) electroreduction in the studied systems determined at the formal potential

Chlorate (VII)	2 mol dm^{-3}		4 mol dm^{-3}		6 mol dm^{-3}	
	E_f^0/V	$R_{\text{amin}}/\Omega \text{ cm}^2$	E_f^0/V	$R_{\text{amin}}/\Omega \text{ cm}^2$	E_f^0/V	$R_{\text{amin}}/\Omega \text{ cm}^2$
A	0.102	224.3	0.125	28.96	0.150	2.72
B	0.097	242.9	0.131	56.9	0.157	6.94
C	0.095	252.1	0.125	74.2	0.158	11.98
D	0.098	253.1	0.111	89.4	0.143	74.0
E	0.102	249.1	0.110	139.3	0.140	117.6

confirms the earlier observed regularity of the differences in the mechanism of Bi (III) ions electroreduction apropos the change of $\text{NaClO}_4:\text{HClO}_4$ ratio in the supporting electrolyte. Assuming that the process of Bi (III) electroreduction is multi-staged, and the transfer of individual electrons proceeds consecutively, then at the less positive potentials, the process rate is controlled by the transfer of the first electron (Fig. 5) [13–15]. It occurs, presumably, in the outer Helmholtz plane or within the one water molecule from the electrode surface [20–22].

Similar conclusions regarding the $[\text{In}(\text{H}_2\text{O})_6]^{+3}$ electroreduction were made earlier in Ref [23] using a quantum mechanical theory.

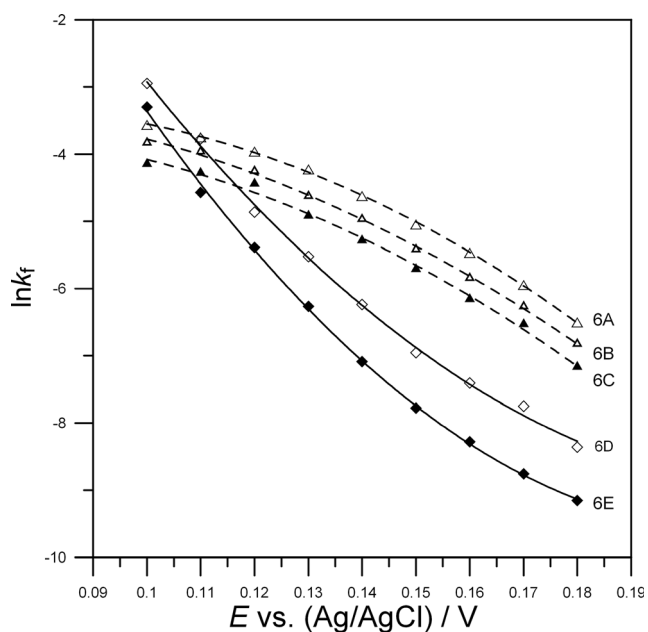


Fig. 5 The dependence of rate constants of $1 \times 10^{-3} \text{ mol dm}^{-3}$ Bi (III) electroreduction in 6 mol dm^{-3} chlorates (VII), where $\text{HClO}_4:\text{NaClO}_4 = 1:1$ (—) 6A; $\text{HClO}_4:\text{NaClO}_4 = 1:4$ (—) 6B; $\text{HClO}_4:\text{NaClO}_4 = 1:9$ (—) 6C; $\text{HClO}_4:\text{NaClO}_4 = 4:1$ (—) 6D; $\text{HClO}_4:\text{NaClO}_4 = 9:1$ (—) 6E on the electrode potential

The inner sphere contribution to the total reorganisation energy for $[\text{In}(\text{H}_2\text{O})_6]^{+3}$ is larger when compared with In (III) aquahydroxocomplex, since the solvent reorganisation is practically the same for both species. The electrode-reactant orbital overlap is stronger for $[\text{In}(\text{H}_2\text{O})_5\text{OH}]^{+2}$. Such findings agree with the structure of acceptor orbitals explored for both complex ions. The first electron transfer was shown to be rate controlling [23].

Conclusions

The described studies confirmed unequivocally the inversely proportional dependence of the rate of Bi (III) ions electroreduction on water activity [4]. The rate of Bi (III) ions electroreduction increases with the decrease of water activity [4].

Significant changes in the kinetics of the Bi (III) ions electroreduction process apropos the change of $\text{HClO}_4:\text{NaClO}_4$ ratio in the solutions ($4\text{--}7 \text{ mol dm}^{-3}$) of chlorates (VII) were found. The increase of the concentration of chloric acid sodium salt, as well as the chloric (VII) acid alone within the particular concentration of the supporting electrolyte, inhibits the process of Bi (III) ions electroreduction. It should be associated with the reorganisation of the structure of the double layer connected with the slow dehydration inhibited by ClO_4^- ions. Analysing water activity in the studied solutions of chloric (VII) acid and sodium salt of chloric (VII) acid [3], it should be mentioned that the increase of NaClO_4 concentration results in the changes of water activity from 0.965 in 1 mol dm^{-3} NaClO_4 to 0.628 in 7 mol dm^{-3} NaClO_4 ; whereas the change of HClO_4 concentration from 1 mol dm^{-3} to 7 mol dm^{-3} causes the decrease of water activity from 0.962 to 0.325. Such behaviour explains the change of the kinetics and presumably the mechanism of the process examined.

The question is raised, if the replacement of Na^+ ions with Me^{+n} ions, where $n > 1$, in the supporting electrolyte, will it lead to similar effects?

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