

Scientific writing skill

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Victoria Manuel Delgado

University of Cádiz



Index

- Getting started at the lab: **Literature searches**
- **Experimental work:** The day a day protocol
- Obtaining and gathering results:
Writing a scientific paper
- **Make your scientific work visible:**
Poster elaboration

Literature searches

When starting a research work, the first thing to do is...
literature searching



Data bases for scientific bibliography (UCA source):

http://diana.uca.es/search*spl/y?SEARCH=base+de+datos&searchscope=3&SORT=D

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Science direct: <http://www.sciencedirect.com/>

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Literature searches

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Search for... Eg., "heart attack" AND stress Article Title, Abstract, Keywords

Limit to:

Date Range (inclusive)
 Published All years to Present
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Document Type
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Subject Areas
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Literature searches

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Search for... Eg. "heart attack" AND stress **Article Title, Abstract, Keywords**

+ Add search field

Limit to:

Date Range (inclusive)
 Published **All years** to **Present**
 Added to Scopus in the last **7** days

Document Type
ALL

Subject Areas
 Life Sciences (> 4,300 titles.)
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Literature searches

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- Review
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Refine Limit to Exclude

Year 2014 (507)

Author Name Schultz, T. (7) Hengerer, F.H. (7) Frings, A. (5) Kohnen, T. (5) Querques, G. (5)

Subject Area Medicine (411) Neuroscience (101) Physics and Astronomy (71) Materials Science (42) Engineering (40)

Document Type Article (397)

ABSTRACT

2- μ m fiber laser sources for sensing Ceng, J., Jiang, S. 2014 Optical Engineering 53 (6), 061609

2- μ m fiber lasers have become a research topic with an increased emphasis due to a variety of applications including eye-safe LIDAR, spectroscopy, remote sensing, and mid-infrared (mid-IR) frequency generation. We review our latest development on various 2- μ m fiber laser sources, including single-frequency fiber lasers, Q-switched fiber lasers, mode-locked fiber lasers, and mid-IR supercontinuum fiber sources. All these fiber laser sources are based on thulium and holmium ions using our proprietary glass fiber technology. Potential applications of these fiber laser sources for sensing are also briefly discussed. © 2014 Society of Photo-Optical Instrumentation Engineers (SPIE).

Document ID	Title	Author(s)	Year	Journal	Cited by
1	Periocular photodynamic therapy for squamous intra-epidermal carcinoma	Connelly, N., Hemmant, B., ...	2014	Journal of Dermatological Treatment	0
2	2- μ m fiber laser sources for sensing	Ceng, J., Jiang, S.	2014	Optical Engineering 53 (6), 061609	0
3	Hyperemic responses of the optic nerve head blood flow to chromatic equiluminant flicker are reduced by ocular hypertension and early glaucoma	Falsini, B., Riva, C.E., Salgarello, T., ...	2014	Optical Engineering	0
4	Patterns of Peripheral Retinal and Central Macula Ischemia in Diabetic Retinopathy as Evaluated by Ultra-widefield Fluorescein Angiography	Sim, D.A., Keane, P.A., Rajendram, R., ...	2014	American Journal of Ophthalmology	0
5	Management of Idiopathic Retinal Vasoproliferative Tumors by Slit-Lamp Laser or Endolaser Photocoagulation	Krivacic, V., Massin, P., Desjardins, L., ...	2014	American Journal of Ophthalmology	0

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- Kohnen, T. (5)
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Subject Area

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- Neuroscience (101)
- Physics and Astronomy (71)
- Materials Science (42)
- Engineering (40)

Document Type

- Periocular photodynamic therapy for squamous intra-epidermal carcinoma
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Optical Engineering

Volume 53, Issue 6, 1 June 2014, Article number 061609

2- μ M fiber laser sources for sensing

Wang, Q. , Geng, J., Jiang, S. 

AdValue Photonics Inc., 3708 E. Columbia Street, Tucson, Arizona 85714, United States

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Abstract

2- μ m fiber **lasers** have become a research topic with an increased emphasis due to a variety of applications including **eye-safe** LIDAR, spectroscopy, remote sensing and mid-infrared (mid-IR) frequency generation. We review our latest development on various 2- μ m fiber **laser** sources, including single-frequency fiber **lasers**, Q-switched fiber **lasers**, mode-locked fiber **lasers**, and mid-IR supercontinuum fiber sources. All these fiber **laser** sources are based on thulium and holmium ions using our proprietary glass fiber technology. Potential applications of these fiber **laser** sources for sensing are also briefly discussed. © 2014 Society of Photo-Optical Instrumentation Engineers (SPIE).

Author keywords

laser sensing; midinfrared; mode-locked **laser**; Q-switched **laser**; single frequency; Subject terms: 2- μ m fiber **laser**

Indexed keywords

Frequency generation; Glass fiber technology; **Laser** sensing; **Lasers**, Q-switched; Mid infrared (mid IR); Midinfrared; Mode-locked **laser**; Single frequency

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Geng, J. , Wang, C
(2011) Proceeding
Optical Engineering

Q-switched puls
silicate fibers
Wang, Q. , Geng, J
(2012) Progress in
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Mode-locked 2 μ m
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Wang, Q. , Geng, J
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Engineering controlled terms: Locks (fasteners); Q switching



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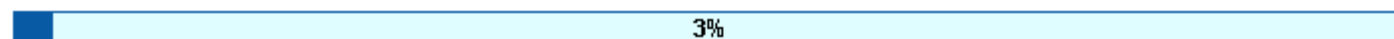
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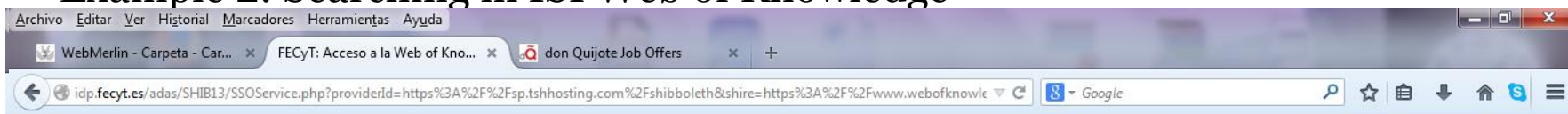
Indexed keywords

Frequency generation; Glass fiber technology; **Laser** sensing; **Lasers**, Q-switched; Mid infrared (mid IR); Midinfrared; Mode-locked **laser**; Single frequency

Engineering controlled terms: Locks (fasteners); Q-switching

Literature searches

Example 2: Searching in ISI Web of Knowledge



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All years

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Literature searches

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MORE SETTINGS

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- Authors
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- Source Titles
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- Languages
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By: Garcia, Veronica; Steeghs, Wil; Bouten, Marcel; et al.
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39. **Non-Dispersive Solvent Extraction of Uranium from Nitric Acid Medium by Several Amides and their Mixture with TODGA using a Hollow Fiber Contactor**

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Times Cited: 0
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By: Jagdale, Y. D.; Vernekar, P. V.; Patwardhan, A. W.; et al.
SEPARATION SCIENCE AND TECHNOLOGY Volume: 48 Issue: 16 Special Issue: SI Pages: 2454-2467
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JOURNAL OF CLEANER PRODUCTION Volume: 59 Pages: 274-283 Published: NOV 15 2013

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Non-Dispersive Solvent Extraction of Uranium from Nitric Acid Medium by Several Amides and their Mixture with TODGA using a Hollow Fiber Contactor

By: Raut, DR (Raut, D. R.)^[1]; Mohapatra, PK (Mohapatra, P. K.)^[1]

SEPARATION SCIENCE AND TECHNOLOGY

Volume: 48 Issue: 16 Pages: 2436-2443 Special Issue: SI

DOI: 10.1080/01496395.2013.807839

Published: NOV 2 2013

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Abstract

Non-dispersive solvent extraction (NDSX) of uranium from nitric acid feed solutions was studied using three dialkyl amides, viz. N,N-di-n-hexylhexanamide (DHHA), N,N-di-n-hexyloctanamide (DHOA), N,N-di-n-hexyldecanamide (DHDA), and their binary mixtures with TODGA (N,N,N',N'-tetra-n-octylidiglycolamide) in normal paraffinic hydrocarbon (NPH) as the carrier solvents. The feed solution usually contained 10.6g/L of uranium in 3M HNO₃ and the extraction studies were carried out using a commercial hollow fiber contactor operating in recirculation mode at a flow rate of 3L/h. In a separate set of experiments, the stripping studies were carried out in the NDSX mode using the U loaded organic extract and 0.01M HNO₃ or 1M Na₂CO₃ as the strippant. Use of TODGA along with the amides resulted in faster mass transfer rates though the extraction and stripping efficiencies of the dialkylamides were reversed in the mixture as compared to that observed in the absence of TODGA.

Keywords

Author Keywords: dialkylamide; hollow fiber contactor; non-dispersive solvent extraction; TODGA; uranium

KeyWords Plus: SUPPORTED LIQUID-MEMBRANE; FISSION-PRODUCTS; HEAVY-METALS; U(VI); SEPARATIONS

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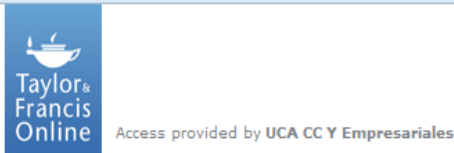
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Non-Dispersive Solvent Extraction of Uranium from Nitric Acid Medium by Several Amides and their Mixture with TODGA using a Hollow Fiber Contactor

DOI: 10.1080/01496395.2013.807839

D. R. Raut^a & P. K. Mohapatra^{a*}

pages 2436-2443

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Volume 48, Issue 16, 2013

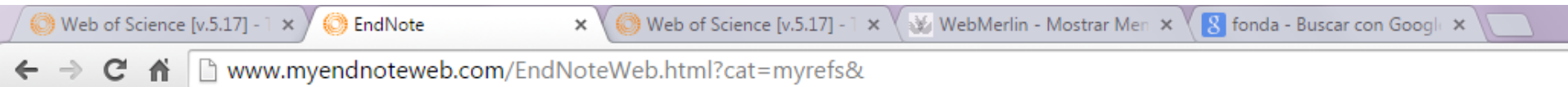
Uranium Extraction Studies

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Information in scientific papers

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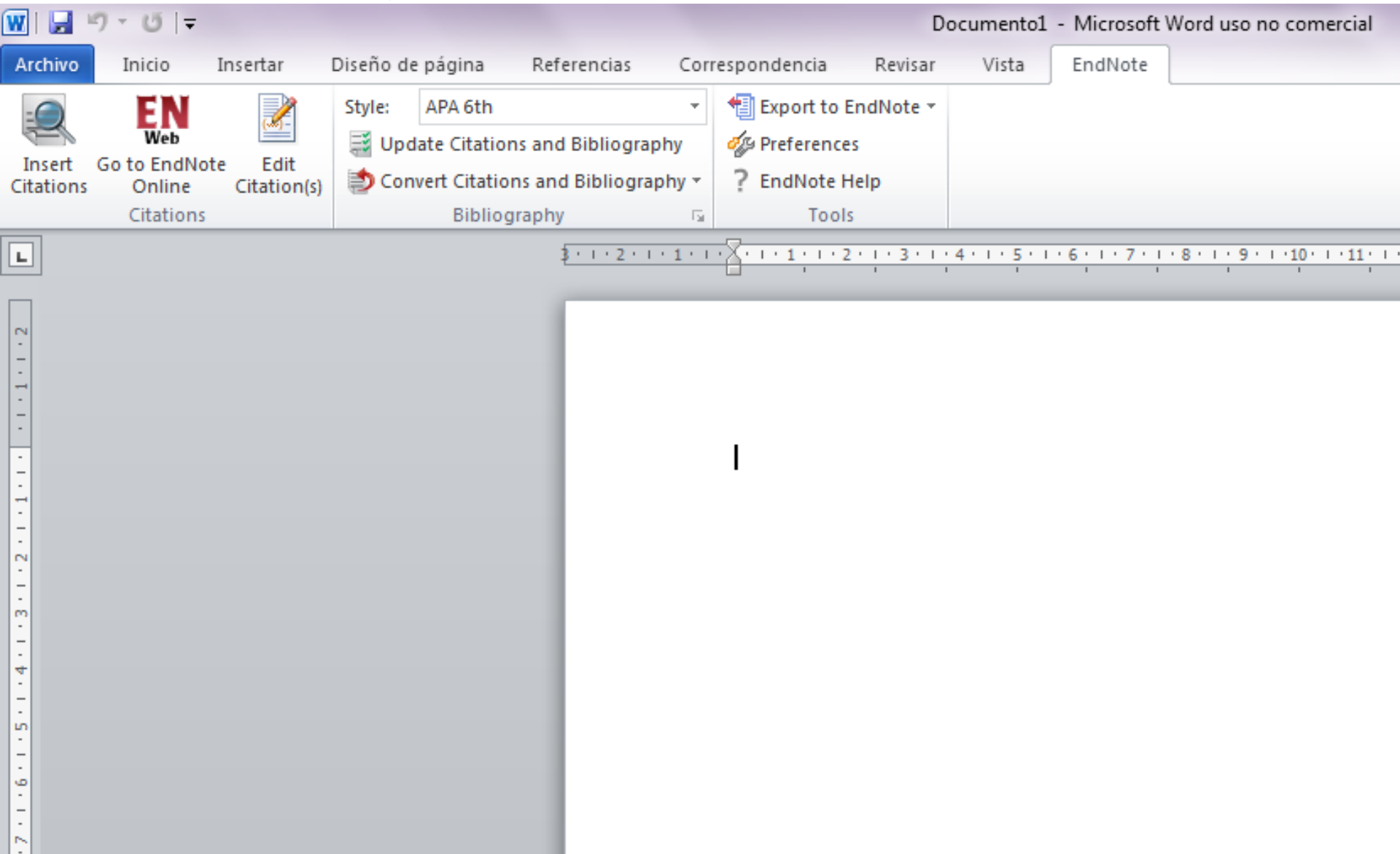
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- Getting started at the lab: **Literature searches**
- **Experimental work:** The day a day protocol
- Obtaining and gathering results:
Writing a scientific paper
- **Make your scientific work visible:**
Poster elaboration

A scientist in a white lab coat and purple gloves is working at a laboratory bench. The scientist is looking at a notebook with handwritten notes. In the background, there is a microscope, a calculator, and various laboratory glassware and bottles. The text "Day a day work in a lab" is overlaid on the image in a teal box.

Day a day work in a lab

Never forget,...

1) Keep a diary of your daily work:
EVERYTHING you did and did
wrong



2) Take pictures of the
important things



Never forget,...

3) Keep everything tidy
(cleaning is vital: Protocols)



4) Never leave a bottle/beaker,
etc. without the proper label



Index

- Getting started at the lab: **Literature searches**
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Writing a scientific paper



Writing a scientific paper

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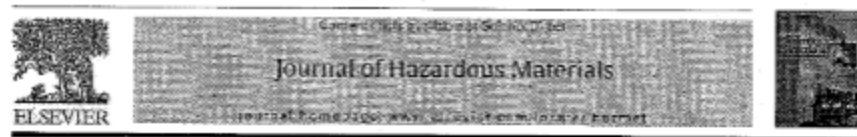
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Introduction



Isolation and preconcentration of Cd(II) from environmental samples using polypropylene porous membrane in a hollow fiber renewal liquid membrane extraction procedure and determination by FAAS

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FAAS

ABSTRACT

The use of polypropylene porous membrane in a hollow fiber renewal liquid membrane (HFRLM) procedure for determination of Cd(II) in water samples was assessed. Ammonium 0,0-diethyl dithiophosphate (DDTP) was used to complex cadmium (II) in an acid medium to obtain a neutral hydrophobic complex. The organic solvent used as donor in the sample extraction film complex from the aqueous solution and carries it over the polypropylene membrane pores. The organic solvent is immobilized inside the polypropylene membrane pores, breaking an homogeneous phase. The complex strips the lumen of the membrane when, at higher pH, the complex breaks down and Cd(II) is released into the stripping phase. EDTA was used to complex the cadmium (II), helping to trap the analyte in the stripping phase. The optimized variables were: sample pH, EDTA concentration, stripping pH, EDTA concentration, extraction temperature and time, extractor solvent and addition of salt to saturate the sample. The sample volume used was 10 mL and the stripping volume was 165 µL. The analyte enrichment factor was 167, limit of detection 1.3 µg L⁻¹, relative standard deviation 3.31 (15 µg L⁻¹, n=7) and the working linear range 1.3–20 µg L⁻¹.

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1. Introduction

The interest in miniaturization in the area of analytical chemistry has led to the introduction of alternative techniques to substitute the conventional liquid-liquid extraction and solid-phase extraction procedures. Among these alternative techniques, liquid-phase microextraction (LPME) introduced firstly by Liu and Dasgupta [1] and by Jeannot and Cantwell [2] is becoming a widely accepted and applied sample preparation technique, as it is a simple, relatively fast extraction and preconcentration procedure and is particularly attractive for the replacement of techniques that use solvents [3].

Among the several possible configurations in which LPME can be performed, the use of a hollow fiber membrane (HFPM) to stabilize the extracting phase was introduced by Pedersen-Rysgaard and Rasmussen [4]. HFPM can be conducted in two-phase or three-phase configurations [5] being that the three-phase system consists of immobilizing a water-immiscible organic solvent in the wall pores of the HF while an aqueous acceptor solution is held within its lumen. Thus, analytes are extracted into the interme-

diary organic phase and subsequently into the aqueous phase. In case of metal ion determination, the mass transfer process from a donor to an acceptor solution through a liquid membrane involves several stages, which include: a complexation reaction between the metal ion and the extractant at the membrane/donor solution interface, diffusion of the complex formed through the liquid membrane and break down of the complex at the membrane/acceptor solution interface with the release of the metal ion to the acceptor solution. This mechanism enables the transfer process to be carried out even at low analyte concentrations and even against an analyte concentration gradient, known as facilitated transport [6–11].

Recently, Kea et al. [12–14] introduced the concept of HFRLM, where an extractor solvent is introduced not only into the membrane porous but also into the sample. Due to the wetting affinity of the organic phase and hydrophobic membrane, a thin organic film of solvent develops at the interface between the donor phase and the membrane. The shear force due to the sample agitation causes the formation of organic microdroplets on the surface of the liquid membrane layer, which separate from the surface of the liquid membrane. At the same time, the organic microdroplets present in the sample greatly increase the contact area between the extractor solvent and the sample. In addition, the presence of organic solvent into the sample provides the liquid membrane renovation. The HFRLM was used for simultaneous extraction and concentration of copper (II) from wastewater

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Experimental procedure (instrumentation, reagents,...)



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[12], citric acid from dilute solutions [13] and penicillin G from aqueous solution [14]. The system of di(2-ethylhexyl) phosphoric acid in kerosene (extractor phase)+HCl (stripping phase): 30% N235 (trialkyl amine, R_3N , $R=C_8-C_{10}$): 20% n-octanol:50% kerosene (extractor phase)+NaOH (stripping phase) and 7% di-n-octylamine:30% iso-octanol:kerosene (extractor phase)+ Na_2CO_3 (stripping phase) were used to study the mass transfer characteristics of HFRLM process for each application aforementioned. In all these cases, a PVDF hollow fiber (effective length 29.8 cm, internal diameter 814 μm , external diameter 886 μm and membrane porosity 0.82), as well as a self-designed system with two 0–1 $dm^3 \text{ min}^{-1}$ peristaltic pumps and flow meters specifically designed for the experimental purposes were used. More recently, Carasek and co-workers [15] simplify the HFRLM presented by Ren et al. [12–14] adapting it to a U-shape configuration. In this new configuration a polydimethyl siloxane membrane was used to determine cadmium (II) from aqueous samples. DDTP was used to complex the Cd(II) to a neutral hydrophobic complex which was extracted by a solvent mixture n-butyl acetate and hexane (60:40, v/v) and re-extracted into EDTA pH 8.5.

In continuation to our previous communication [15], the purpose of this work is to assess the feasibility of the use of polypropylene porous membrane for HFRLM procedure in a U-shape configuration for determination of Cd(II) with FAAS. The method has been applied on environmental samples and the proposed system showed good results.

2. Experimental procedure

2.1. Instrumentation

A Varian Model SpectraAA 50 (Mulgrave, Australia) flame atomic absorption spectrometer equipped with a deuterium lamp (JISCHI IBA-4S, Tokyo, Japan) was used for the analysis. The instrument was operated under the conditions recommended by the manufacturer. The analytical signals were measured as peak area. A 120 Merlett Toledo pH meter was used to adjust the pH of the solutions. AMicroquímica MQAMA 301 stirrer (Santa Catarina, Brazil) was used to agitate the solutions and a Microquímica MQBTC 09-20 bath system was used to control the temperature. A Q 3/2 acanvel polypropylene hollow fiber membrane (600 μm id, 200 μm wall thickness and 0.2 μm pore size) was purchased from Membranas GmbH (Wuppertal, Germany). The hollow fiber was cut into 3.0 cm segments and was cleaned in acetone and dried before use. It was used as the barrier between the donor phase and the acceptor phase.

2.2. Reagents

Ultrapure water from a Milli-Q[®] (Bedford, MA, USA) water purification system (Millipore[®]) was used to prepare all solutions. All chemicals were of analytical grade and were used without previous purification, except for DDTP. The laboratory glassware was kept overnight in a 2% (v/v) Tritan[®] (Merck, Darmstadt, Germany) solution and then again overnight in a 10% (v/v) hydrochloric acid solution. Before use, the glassware was washed with deionized water and dried in a dust-free environment.

Cadmium (II) working standard solutions were prepared daily by dilution of a 2000 $mg L^{-1}$ cadmium (II) stock solution (atomic absorption grade, Carlo Erba, Italy).

Ammonium sulfate (Nuclear, São Paulo, Brazil) and sodium chloride (Nuclear) were used to evaluate the salt-out effect. Toluene (Tedia, Fairfield, OH, USA), hexane (Tedia, Fairfield, OH, USA) and n-butyl acetate (Neste, São Paulo, Brazil) were used as extractor solvents.

Fig. 1. Schematic illustration of three-phase HFRLM.

Arancaium D.D. diethyl diethoxyphosphate (DDTP) was supplied by Aldrich (Milwaukee, WI, USA) and used after purification with silica gel C18 columns (Merck).

The working solutions were prepared by the addition of proper amount of salt, organic solvent and DDTP solution 1% (v/v) to 20 ml of a working standard solution containing 100 $mg L^{-1}$ of Cd(II). An EDTA solution (Vetec) was used as stripping phase.

Water samples collected from Araraquá River (Araraquá, Santa Catarina, Brazil), Cassino River (Sertão, Rio Grande do Sul) and sea water sample (Rio Vermelho beach, Florianópolis, Santa Catarina, Brazil) were used to verify the accuracy of the proposed method.

2.3. Extraction procedure

An aliquot of 20 ml of sample and an adequate amount of salt, DDTP and extractor solvent were added to a 40 ml. vial. A polypropylene hollow fiber was cut carefully into 8 cm pieces and the two ends of each piece were connected to needles each with a 250 μL micro-syringe. One of these micro-syringes contained 105 μL of the acceptor solution, which was used to fill the lumen of the hollow fiber. The vial contain the sample was introduced into the temperature-controlled bath unit at the appropriate temperature. The extraction was carried out by placing the hollow fiber in the sample vial. After stirring for a predetermined time, the acceptor solution was collected by the second micro-syringe and the solution was directly injected into the FAAS. Fig. 1 represents schematically the three-phases of the HFRLM system used in this study.

3. Results and discussion

Firstly, the optimum experimental conditions for extraction of Cd(II) from aqueous solutions using the HFRLM system was determined. For this purpose a multivariable optimization was adopted. The best extracting organic solvent on the extraction efficiency was studied by triangular surface mixture design. Sample pH, DDTP concentrations, acceptor phase pH, EDTA concentrations, organic solvent volume and extraction temperature and time were optimized using a three-level full factorial design that generates surface response plots. In the optimized conditions, the analytical figures of merit were obtained and the proposed method was applied to the determination of Cd(II) in real system samples.

Writing a scientific paper

Results and discussion



3.1. Effect of salting-out

The addition of salts to the samples results in increase of the ionic strength of the solution influencing significantly the extraction process. In some cases, an increase in the ionic strength of the sample phase can decrease the solubility of the analytes in an aqueous medium because of the effect called salting-out [16]. In this process, the water molecules hydrate ions reducing the molecules of analytes dissolved in the aqueous medium through a competition mechanism. In other words, increasing the salt concentration in the donor phase the water molecules will have more interaction to salt ions, releasing the DDTP-Cd(II) complex to interact to the organic solvent.

In order to evaluate the salting-out effect, some preliminary experiments using sodium chloride and ammonium sulfate were carried out. The results have shown that both salts presented an improvement in the extraction efficiency. However, a saturated ammonium sulfate solution showed a very distinct improvement in the analytical signal as demonstrated previously in our recent work [15]. The possible explanation is the high solubility and the presence of double charge in the ammonium sulfate which increase considerably the ionic strength of the solution. Despite the difference in ionic strength of feed aqueous solution and stripping solution, there was no osmotic process observed in the proposed system. Probably, because of the fact that the polypropylene membrane was completely filled with organic solvent which prevents permeability of the water molecules. Thus, this condition was adopted in the subsequent tests.

3.2. Effect of extractor solvent

An organic solvent suitable for the HFRLM should have low solubility in water, good compatibility with polypropylene hollow membrane and high partition coefficient between the analytes and the solvent [17]. In this case, the polarity between the organic solvent and the DDTP-Cd(II) complex should be similar, so a high partition coefficient could be reached. The effect of some extractor solvents on the extraction efficiency was studied. The mixture of organic solvents was optimized through triangular surface response design, where toluene, hexane and butyl acetate were evaluated individually, through binary mixtures with 33% and 67% (v/v) of one solvent and a ternary mixture containing 33% of each solvent. The results obtained from this design can be seen in Fig. 2. Here, there is a region in which the response reaches high values, corresponding to the use of toluene. This condition was selected to continue the optimization of the proposed method.

3.3. Effect of sample pH and DDTP concentration

Several factors were responsible for the choice of DDTP as complexing agent such as its solubility in water, high stability in acid medium and high formation constant to Cd(II) complexes [18]. The best experimental conditions for sample pH and DDTP concentration were determined through a three-level experimental design. From these results, a response surface was generated as shown in Fig. 3. Increasing the DDTP concentration from 0.025% to approximately 0.07% we can see an enhancement in the analytical signal. This indicates that with a DDTP concentration below 0.025% the extracting solvent does not contain enough DDTP to complex Cd(II), reducing the Cd(II) flux through the membrane. On the other hand, high DDTP concentrations lead to an increased membrane viscosity and, consequently, a lower extraction efficiency due to a reduced Cd(II) flux through the membrane. Also, high DDTP concentrations can lead to the formation of charged complexes like ML_2^{2-} (where M is the metal ion and L is the DDTP ligand). Thus, the amount of the

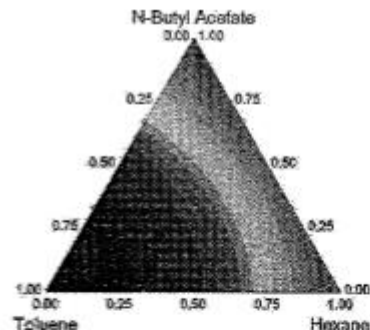


Fig. 2. Study of the effect of solvent choice on Cd(II) extraction by HFRLM with polypropylene membrane and detection by FAMS. Experimental conditions: $100 \mu\text{g L}^{-1}$ Cd(II), extraction time: 30 min, sample pH 3.5, DDTP concentration: 0.01% (v/v), stripping pH 8.5, EDTA concentration: $5 \times 10^{-3} \text{ mol L}^{-1}$, saturated solution $(\text{NH}_4)_2\text{SO}_4$ and extracting solvent volume: 500 μL .

essentially extracted into the organic solvent. Fig. 3 also shows an antagonistic effect between the sample pH and the DDTP concentrations. As can be observed, reducing the sample pH a higher DDTP concentration is required for the complexation of Cd(II), probably because the protonation of DDTP is increased. In accordance with this study the optimum values for the two variables are sample pH of 3.5 and DDTP concentration of 0.05% (v/v).

3.4. Effect of acceptor phase pH and EDTA concentration

As in the three-phase HFRLM system the DDTP-Cd(II) complex is formed in an acid medium, the pH acceptor phase should be basic promoting the broken down of the complex and release

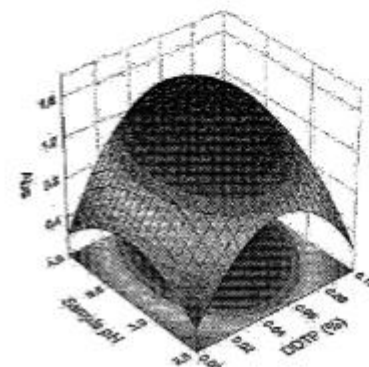


Fig. 3. Study of effect of sample pH and DDTP concentration on Cd(II) extraction by HFRLM with polypropylene membrane and detection by FAMS. Experimental conditions: $100 \mu\text{g L}^{-1}$ Cd(II), extraction time: 30 min, stripping pH 8.5, EDTA concentration: $5 \times 10^{-3} \text{ mol L}^{-1}$, saturated solution $(\text{NH}_4)_2\text{SO}_4$ and extraction solvent

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- Most important key words should be in the title.
- It must be powerful and with rhetoric strength.

Abstract

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- Put your thesis/question somewhere near the end.
- Put some important background information directly before.
- It should inform of your paper's general topic, the type of terminology, evidence and logic they can expect throughout the paper.
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- This section highlights the answers to the hypotheses you investigated.

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- It answers the questions posed in the introduction.
- It is considered *the heart of the paper* and usually requires several writing attempts.
- Try to organize your thoughts in a logical form.
- Should be as short as possible while clearly stating, supporting and explaining your answers and discussing other relevant issues.
- Explain how the results and conclusions of this study are important and how they influence our knowledge of the problem being examined.

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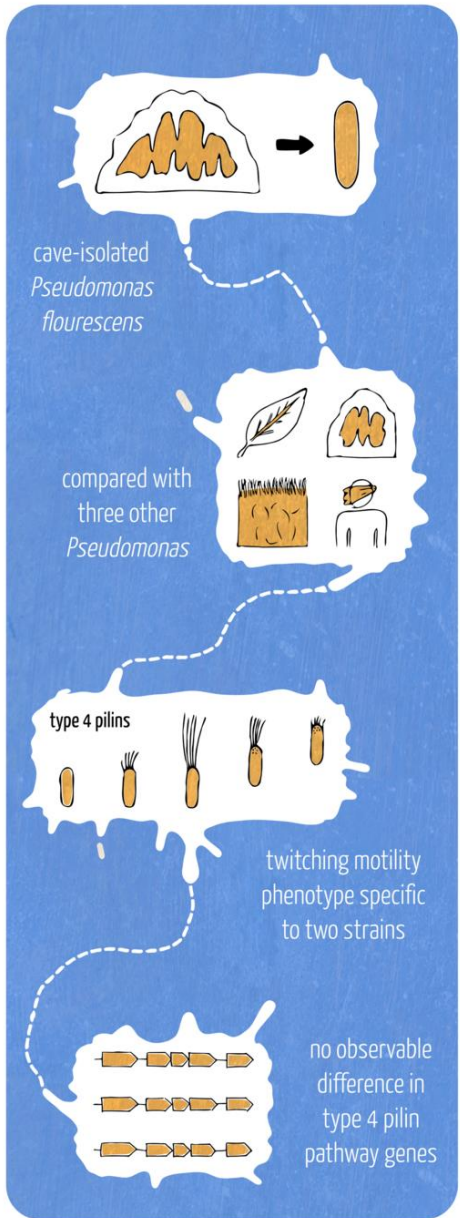
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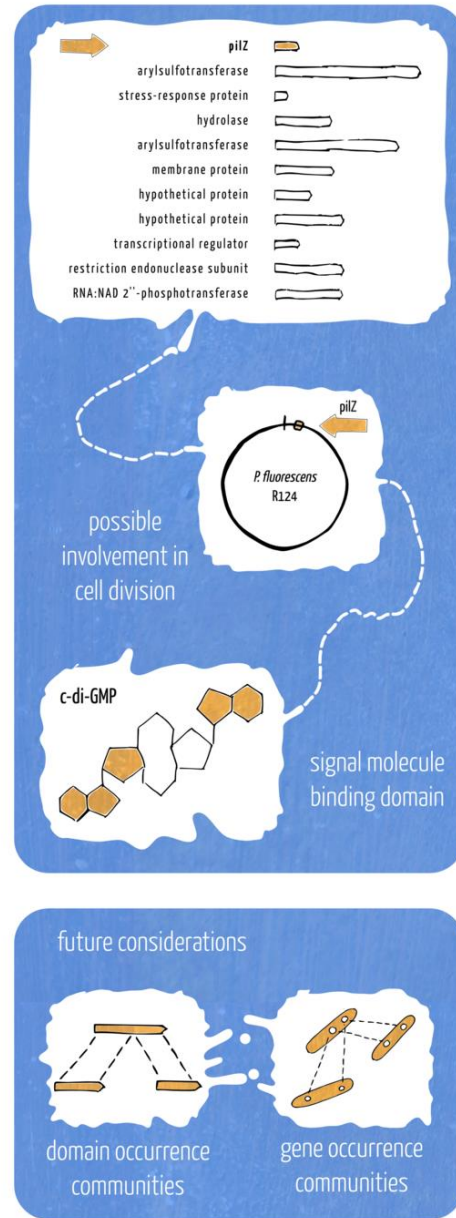
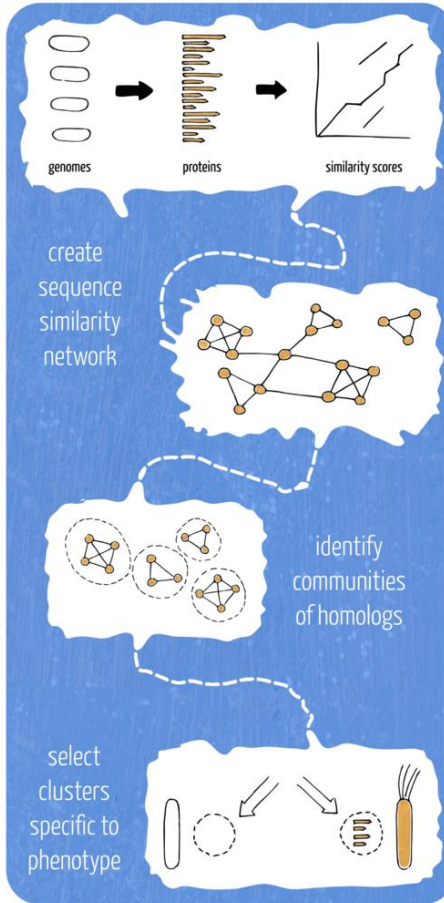
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Which problem do we need to answer to?

Theory of the system (basic)

Results

Conclusions

References (optional)

Acknowledgements

Modelling Water Flow and Transport in Real Soils and Catchments

Jessica Davies, Keith Beven (PI), Lancaster Environment Centre

Presented by MRC 09230017123413, 1st Jun 2010 - 10:00 AM - 1812

Email: j.davies@lancaster.ac.uk for more information

Why important?

- Understanding the processes by which water flows and transports solutes in catchments is vital to:
- a. predicting atmospheric soil loading
- b. assessing the movement of pesticides and herbicides
- c. predicting water quality and availability
- d. understanding hydrological processes

What's wrong with current models?

Simulation files produced from a wide range of current models of catchment flow hydrology are used to assess the realism of their representation of processes and fluxes. The results are compared to a comprehensive data set. Analysis reveals that current catchment hydrology models are a significant over-representation of the current catchment hydrology. This is due to the over-representation of the current catchment hydrology.



Our approach: the Multiple modelling framework

Multiple modelling framework (MUF) is a new approach to catchment hydrology modelling. It is based on the idea of using multiple models to simulate the same catchment. This allows us to compare the results of different models and to assess the realism of their representation of processes and fluxes. The MUF framework is based on the idea of using multiple models to simulate the same catchment. This allows us to compare the results of different models and to assess the realism of their representation of processes and fluxes.



5km scale experiment

The 5km scale experiment is a new approach to catchment hydrology modelling. It is based on the idea of using multiple models to simulate the same catchment. This allows us to compare the results of different models and to assess the realism of their representation of processes and fluxes. The 5km scale experiment is based on the idea of using multiple models to simulate the same catchment. This allows us to compare the results of different models and to assess the realism of their representation of processes and fluxes.

Landscape experiment

The landscape experiment is a new approach to catchment hydrology modelling. It is based on the idea of using multiple models to simulate the same catchment. This allows us to compare the results of different models and to assess the realism of their representation of processes and fluxes. The landscape experiment is based on the idea of using multiple models to simulate the same catchment. This allows us to compare the results of different models and to assess the realism of their representation of processes and fluxes.

Catchment scale experiment

The catchment scale experiment is a new approach to catchment hydrology modelling. It is based on the idea of using multiple models to simulate the same catchment. This allows us to compare the results of different models and to assess the realism of their representation of processes and fluxes. The catchment scale experiment is based on the idea of using multiple models to simulate the same catchment. This allows us to compare the results of different models and to assess the realism of their representation of processes and fluxes.

Conclusions

The MUF framework is a new approach to catchment hydrology modelling. It is based on the idea of using multiple models to simulate the same catchment. This allows us to compare the results of different models and to assess the realism of their representation of processes and fluxes. The MUF framework is based on the idea of using multiple models to simulate the same catchment. This allows us to compare the results of different models and to assess the realism of their representation of processes and fluxes.

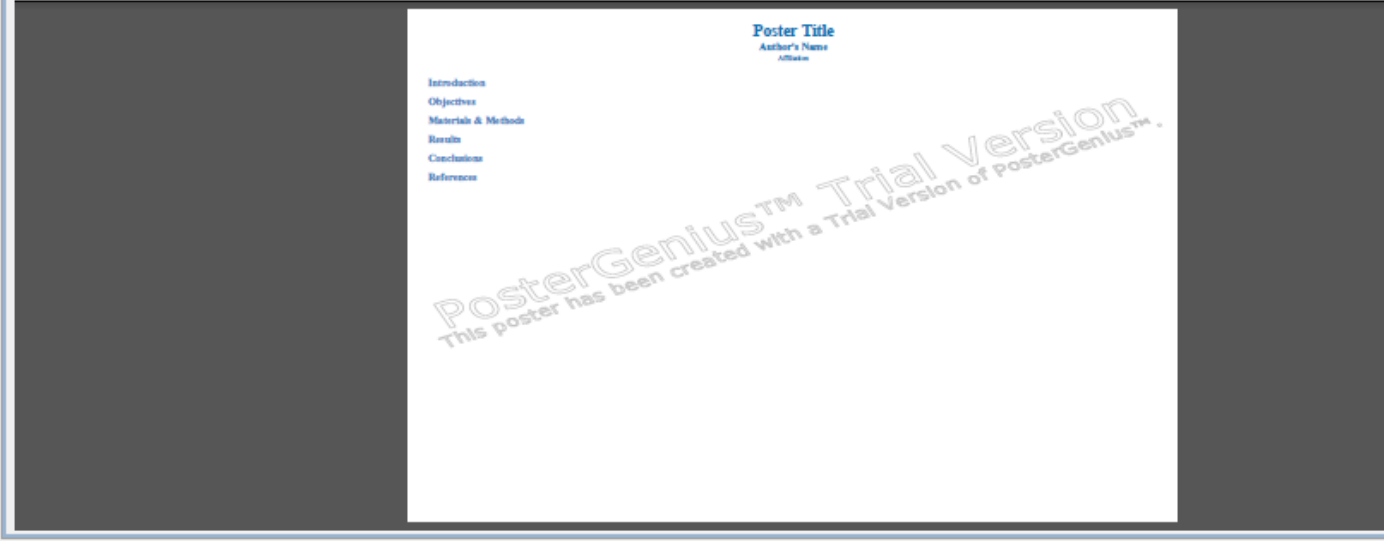
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The screenshot shows a web browser window displaying the ResearchGate profile of Victoria Manuel. The browser's address bar shows the URL: https://www.researchgate.net/profile/Victoria_Manuel?ev=hdr_xprf. The profile header includes a navigation menu with 'ResearchGate', 'Q&A', 'Publications', and 'Projects'. A search bar and user profile picture are also visible. The main profile section features a profile picture of Victoria Manuel, her name, and her affiliation: 'Chemical engineering', 'PhD Student, Researcher', and 'Universidad de Cádiz · Department of Analytical Chemistry'. A 'Add your publications' button is present. Below the header, a navigation bar lists 'OVERVIEW', 'CONTRIBUTIONS', 'INFO', 'STATS', and 'RG SCORE'. The 'OVERVIEW' section contains a tip: 'Show your career's best' and a statistics table. The statistics table shows: 2 PUBLICATIONS, 28 Views, 12 Downloads, 0 Citations, and 1.17 Impact Points. Below the statistics, there is a section for 'Add your publications' with a list of matches. One match is highlighted: 'Article: Simultaneous determination of benzodiazepines by ultraviolet-visible spectrophotometry in micellar media.' by M de la Guardia, M V Galdú, J Monzó, A Salvador. To the right, the 'ABOUT' section prompts the user to add a short introduction. The 'SKILLS AND EXPERTISE' section lists 'Metal Analysis', 'Spectrometry', and 'Analytical Environmental Chemistry'. The bottom of the image shows a Windows taskbar with various application icons and a system tray displaying the time as 19:18 on 21/05/2014.

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2 PUBLICATIONS 28 Views 12 Downloads 0 Citations 1.17 Impact Points

Article: **Simultaneous determination of benzodiazepines by ultraviolet-visible spectrophotometry in micellar media.**
M de la Guardia, M V Galdú, J Monzó, A Salvador

Universidad de Cádiz
Department of Analytical Chemistry
Andalusia, Spain

SKILLS AND EXPERTISE (3)
1 Metal Analysis Spectrometry
Analytical Environmental Chemistry



Victoria Manuel
Universidad de Cádiz

Article

Liquid phase micro-extraction: Towards the green methodology for ultratrace metals determination in aquatic ecosystems

JA López-López, C Vergel, C Mendiguch'ia, JJ Pinto, V Manuel, M Silva, M García-Vargas, C Moreno

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