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Article

# Thalhammerite, Pd<sub>9</sub>Ag<sub>2</sub>Bi<sub>2</sub>S<sub>4</sub>, a New Mineral from the Talnakh and Oktyabrsk Deposits, Noril'sk Region, Russia

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**Abstract:** Thalhammerite, Pd<sub>9</sub>Ag<sub>2</sub>Bi<sub>2</sub>S<sub>4</sub>, is a new sulphide discovered in galena-pyrite-chalcopyrite and millerite-bornite-chalcopyrite vein-disseminated ores from the Komsomolsky mine of the Talnakh and Oktyabrsk deposits, Noril'sk region, Russia. It forms tiny inclusions (from a few μm up to about 40–50 μm) intergrown in galena, chalcopyrite, and also in bornite. Thalhammerite is brittle and has a metallic lustre. In plane-polarized light, thalhammerite is light yellow with weak bireflectance, weak pleochroism, in shades of slightly yellowish brown and weak anisotropy; it exhibits no internal reflections. Reflectance values of thalhammerite in air (R<sub>1</sub>, R<sub>2</sub> in %) are: 41.9/43.0 at 470 nm, 43.9/45.1 at 546 nm, 44.9/46.1 at 589 nm, and 46.3/47.5 at 650 nm. Three spot analyses of thalhammerite give an average composition: Pd 52.61, Bi 22.21, Pb 3.92, Ag 14.37, S 7.69, and Se 0.10, total 100.90 wt %, corresponding to the empirical formula Pd<sub>8.46</sub>Ag<sub>2.28</sub>(Bi<sub>1.82</sub>Pb<sub>0.32</sub>)Σ<sub>2.14</sub>(S<sub>4.10</sub>Se<sub>0.02</sub>)Σ<sub>4.12</sub> based on 17 atoms; the average of five analyses on synthetic thalhammerite is: Pd 55.10, Bi 24.99, Ag 12.75, and S 7.46, total 100.30 wt %, corresponding to Pd<sub>8.91</sub>Ag<sub>2.03</sub>Bi<sub>2.06</sub>S<sub>4.00</sub>. The density, calculated on the basis of the empirical formula, is 9.72 g/cm<sup>3</sup>. The mineral is tetragonal, space group *I4/mmm*, with *a* 8.0266(2), *c* 9.1531(2) Å, *V* 589.70(2) Å<sup>3</sup> and *Z* = 2. The crystal structure was solved and refined from the single-crystal X-ray-diffraction data of synthetic Pd<sub>9</sub>Ag<sub>2</sub>Bi<sub>2</sub>S<sub>4</sub>. Thalhammerite has no exact structural analogues known in the mineral system; chemically, it is close to coldwellite (Pd<sub>3</sub>Ag<sub>2</sub>S) and kravtsovite (PdAg<sub>2</sub>S). The strongest lines in the X-ray powder diffraction pattern of synthetic thalhammerite [*d* in Å (*I*) (*hkl*)] are: 3.3428(24)(211), 2.8393(46)(220), 2.5685(21)(301), 2.4122(100)(222), 2.3245(61)(123), 2.2873(48)(004), 2.2201(29)(132), 2.0072(40)(400), 1.7481(23)(332), and 1.5085(30)(404). The mineral honours Associate Professor Oskar Thalhammer of the University of Leoben, Austria.

**Keywords:** thalhammerite; platinum-group mineral; Pd<sub>9</sub>Ag<sub>2</sub>Bi<sub>2</sub>S<sub>4</sub> phase; reflectance data; X-ray-diffraction data; crystal structure; Komsomolsky mine; Talnakh deposit; Noril'sk region; Russia

## 1. Introduction

Thalhammerite, ideally Pd<sub>9</sub>Ag<sub>2</sub>Bi<sub>2</sub>S<sub>4</sub>, was observed in the same holotype specimen as kravtsovite, PdAg<sub>2</sub>S [1], and vymazalováite, Pd<sub>3</sub>Bi<sub>2</sub>S<sub>2</sub> [2]. The type sample (polished section) comes from

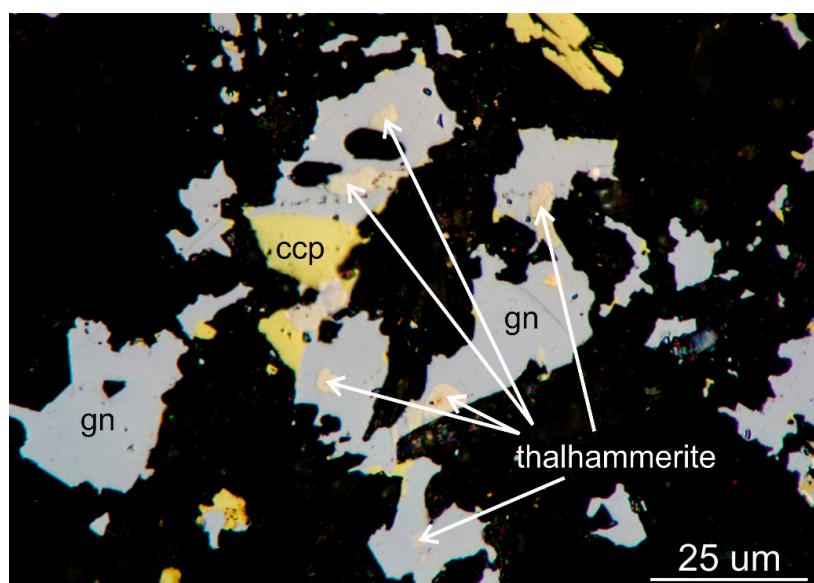
vein-disseminated pyrite-chalcopyrite-galena ore from the Komsomolsky mine in the Talnakh deposit of the Noril'sk district, Russia. The sample was found at coordinates: 69°30'20" N and 88°27'17" E. The mineralization is characterized by lack of Ni minerals and high galena content and Pt-Pd-Ag bearing minerals in an association of pyrite and chalcopyrite. The host rocks of pyrite-chalcopyrite-galena ore are diopside-hydrogrossular-serpentine metasomatites developed in diopside-monticellite skarns below the lower exocontact of the Talnakh intrusion (the eastern part of the Komsomolsky mine). Thalhammerite, in pyrite-chalcopyrite-galena ores, occurs in association with cooperite, braggite, vysotskite, stibiopalladinite, telargpalite, sobolevskite, kotulskite, sopcheite, insizwaite, kravtsovite, vymazalováite, Au-Ag alloys, and Ag-bearing sulphides, selenides, sulphoselenides, and tellurosulphoselenides. The mineral was also observed in vein-disseminated millerite-bornite-chalcopyrite ore from the Talnakh and Oktyabrsk deposits of the Noril'sk region [3]. The host rocks of millerite-bornite-chalcopyrite ore are pyroxene-hornfels at the lower exocontact of the Kharaelakh intrusion (the western part of the Komsomolsky mine). In millerite-bornite-chalcopyrite ore, thalhammerite occurs in association with kotulskite, telargpalite, laflammeite, and Au-Ag alloys.

The mineral likely formed under the same conditions as kravtsovite and vymazalováite, with decreasing temperature [3], most likely below 400 °C. Thalhammerite was also observed, in intergrowths with sobolevskite, in PGE ores from the Fedorov-Pana Layered Intrusive Complex, Russia (V.V. Subbotin—per. communication). Furthermore, the occurrence of unknown phases corresponding to Pb- and Tl-analogues of thalhammerite from the Fedorov-Pana Layered Intrusive Complex has been reported [4].

Both the mineral and name were approved by the Commission on New Minerals, Nomenclature and Classification of the International Mineralogical Association (IMA No 2017-111). The mineral name is for Dr. Oskar Thalhammer (b. 1956) Associate Professor at the University of Leoben, Austria for his contributions to the ore mineralogy and mineral deposits of platinum group elements. The type specimen is deposited at the Department of Earth Sciences of the Natural History Museum, London, UK, catalogue no. BM 2016, 150.

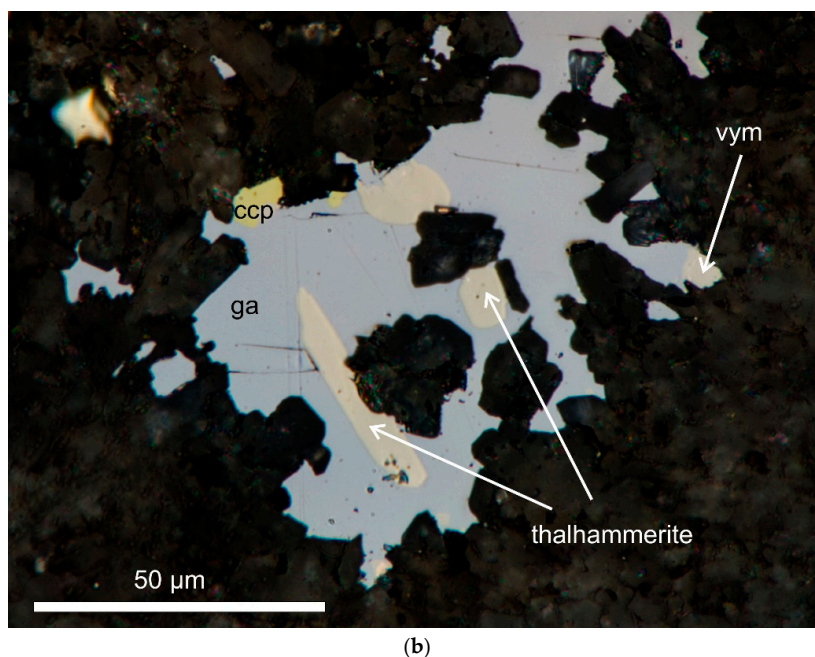
## 2. Appearance, and Physical and Optical Properties

Thalhammerite forms very small inclusions (from a few  $\mu\text{m}$  up to about 40–50  $\mu\text{m}$ ) in galena, chalcopyrite (Figure 1), and also in bornite.



(a)

Figure 1. Cont.



**Figure 1.** Digital image in reflected plane polarized light showing inclusions of thalhammerite in galena (gn) in association with (a) chalcopyrite (ccp) and (b) vymazalováite (vym).

The mineral occurs in aggregates (100–200  $\mu\text{m}$  in size) formed by intergrowths of telargpalite, braggite, vysotskite, sopcheite, stibiopalladinite, sobolevskite, moncheite, kotulskite, malyshevite, insizwaite, acanthite, aurian silver, kravtsovite, and vymazalováite in association with galena, chalcopyrite, bornite, millerite, and pyrite.

Thalhammerite is opaque with a metallic lustre. The mineral is brittle. The density calculated on the basis of the empirical formula is  $9.72 \text{ g/cm}^3$ . In plane-polarized light, thalhammerite is light yellow with weak bireflectance, weak pleochroism, in shades of slightly yellowish brown and weak anisotropy. It exhibits no internal reflections.

Reflectance measurements were made in air relative to a WTiC standard on both natural and synthetic thalhammerite using a J and M TIDAS diode array spectrometer attached to a Zeiss Axiotron microscope. The results are tabulated (Table 1) and illustrated in Figure 2.

**Table 1.** Reflectance data for natural and synthetic thalhammerite.

$\lambda$ (nm)	Natural		Synthetic	
	R <sub>1</sub> (%)	R <sub>2</sub> (%)	R <sub>1</sub> (%)	R <sub>2</sub> (%)
400	40.0	41.2	40.6	42.1
420	40.6	41.8	41.3	42.6
440	41.1	42.3	42.0	43.2
460	41.7	42.8	42.6	43.9
470	41.9	43.0	42.9	44.3
480	42.2	43.3	43.1	44.6
500	42.7	43.9	43.8	45.3
520	43.2	44.4	44.6	46.0
540	43.7	44.9	45.3	46.6
546	43.9	45.1	45.6	46.9
560	44.2	45.4	45.9	47.2
580	44.7	45.9	46.4	47.7
589	44.9	46.1	46.7	47.9
600	45.2	46.3	46.9	48.1
620	45.6	46.8	47.3	48.5
640	46.1	47.3	47.7	48.9



Table 1. Cont.

$\lambda$ (nm)	Natural		Synthetic	
	R <sub>1</sub> (%)	R <sub>2</sub> (%)	R <sub>1</sub> (%)	R <sub>2</sub> (%)
650	<b>46.3</b>	<b>47.5</b>	<b>47.9</b>	<b>49.1</b>
660	46.5	47.8	48.0	49.2
680	47.0	48.3	48.3	49.5
700	47.4	48.9	48.6	49.8

Note. The values required by the Commission on Ore Mineralogy are given in bold.

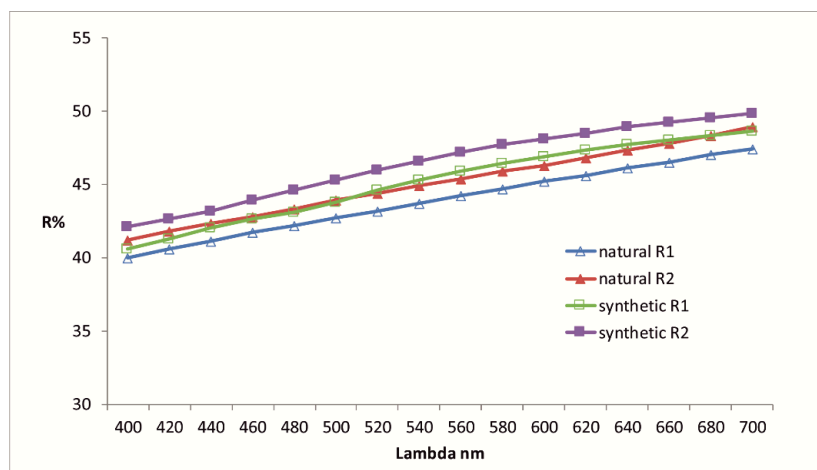


Figure 2. Reflectance data for thalhammerite compared to synthetic analogue, in air. The reflectance values (R%) are plotted versus the wavelength  $\lambda$  in nm.

### 3. Chemical Composition

Electron probe micro-analyses (EPMA) on grains of thalhammerite were obtained using a WDA Inca Wave 500 (Oxford Instruments NanoAnalysis, High Wycombe, UK) installed on an SEM Lyra 3GM (Tescan), with analytical conditions of 20 kV, 10 nA, and counting times of 30 s (on peak positions)/  $2 \times 15$  s (background on the left and right positions). The spectra were collected on PbM $_{\alpha}$ , BiM $_{\alpha}$ , PdL $_{\alpha}$ , AgL $_{\alpha}$ , SK $_{\alpha}$ , and SeL $_{\alpha}$  lines with standards of pure Se, Pd, Ag, Bi, synthetic PbTe, and natural FeS<sub>2</sub>. Other elements were below the detection limit.

EPMA on synthetic thalhammerite were obtained using a CAMECA SX-100 electron probe microanalyzer in wavelength-dispersive mode with an electron beam focussed to 1–2  $\mu$ m.

Pure elements and ZnS were used as standards and the radiations measured were BiM $_{\alpha}$ , PdL $_{\alpha}$ , AgL $_{\alpha}$ , and SK $_{\alpha}$ , with an accelerating voltage of 15 kV, and a beam current of 10 nA measured on the Faraday cup.

EPMA compared with literature data are given in Table 2. The empirical formulae calculated on the basis of 17 apfu are Pd<sub>8.46</sub>Ag<sub>2.28</sub>(Bi<sub>1.82</sub>Pb<sub>0.32</sub>) $\Sigma$ 2.14(S<sub>4.10</sub>Se<sub>0.02</sub>) $\Sigma$ 4.12 for thalhammerite and Pd<sub>8.91</sub>Ag<sub>2.03</sub>Bi<sub>2.06</sub>S<sub>4.00</sub> for its synthetic analogue, with the ideal formulae Pd<sub>9</sub>Ag<sub>2</sub>Bi<sub>2</sub>S<sub>4</sub>.

Table 2. Electron-microprobe analyses of natural and synthetic thalhammerite.

wt %	Pd	Ag	Pb	Bi	S	Se	Total
<b>Thalhammerite</b>							
	52.80	14.57	2.60	22.56	7.75	0.07	100.35
	53.40	14.29	3.05	22.09	7.62	0.03	100.47
	51.64	14.25	6.12	21.98	7.70	0.19	101.87
average	52.61	14.37	3.92	22.21	7.69	0.10	100.90
13/B-92 *	53.85	12.51		24.84	7.90		99.1
	52.77	12.27	1.77	24.29	7.45	0.57	99.12

Table 2. Cont.

wt %	Pd	Ag	Pb	Bi	S	Se	Total
<b>Thalhammerite</b>							
1/K-92 *	53.85	12.8		24.24	7.89		99.01
	52.77	11.83		25.73	7.99		99.53
	54.08	12.21		25.34	7.95		101.21
<b>Synthetic Sample</b>							
Exp37	54.18	13.69		25.02	7.59		100.48
	54.74	12.91		25.04	7.50		100.19
	54.66	12.46		25.90	7.39		100.42
	56.14	12.01		24.70	7.36		100.21
	55.78	12.67		24.27	7.44		100.16
	55.10	12.75		24.99	7.46		100.29
average	55.10	12.75		24.99	7.46		100.30

\* Sluzhenikin and Mohkov [2].

#### 4. Synthetic Analogue

The small size of thalhammerite embedded in galena (bornite) prevented its extraction and isolation in an amount sufficient for the relevant crystallographic and structural investigations. Therefore, these investigations were performed on the synthetic  $\text{Pd}_9\text{Ag}_2\text{Bi}_2\text{S}_4$ .

The synthetic phase of  $\text{Pd}_9\text{Ag}_2\text{Bi}_2\text{S}_4$  was prepared in an evacuated and sealed silica-glass tube in a horizontal furnace in the Laboratory of Experimental Mineralogy of the Czech Geological Survey in Prague. To prevent loss of material to the vapour phase during the experiment, the free space in the tube was reduced by placing a closely-fitting silica glass rod against the charge.

The temperature was measured with Pt-PtRh thermocouples and is accurate to within  $\pm 3$  °C. A charge of about 300 mg was carefully weighed out from the native elements. We used, as starting chemicals, palladium (99.95%), silver (99.999%), bismuth (99.999%), and sulphur (99.999%). The starting mixture was sealed and annealed, quenched, and then ground in an agate mortar under acetone and reheated to 350 °C for 134 days. The sample was quenched by dropping the capsule in cold water.

#### 5. X-ray Crystallography

##### 5.1. Single-Crystal X-ray Diffraction

A small fragment of synthetic  $\text{Pd}_9\text{Ag}_2\text{Bi}_2\text{S}_4$  was mounted on a glass fibre and examined using a Rigaku Super Nova single-crystal diffractometer with an Atlas S2 CCD detector utilizing  $\text{MoK}\alpha$  radiation, provided by the microfocus X-ray tube and monochromatized by primary mirror optics. The  $\omega$  rotational scans were used for collection of three-dimensional intensity data. From a total of 3659 reflections, 221 were classified as unique observed with  $I > 3\sigma(I)$ . Corrections for background, Lorentz effects and polarization were applied during data reduction with the CrysAlis software. Empirical absorption correction was performed using the same software yielding  $R_{\text{int}} = 0.034$ . The crystal structure was solved with a charge-flipping method using the program Superflip [5] and subsequently refined by the full-matrix least-squares algorithm of JANA2006 program [6]. Because of the similarity of atomic number of Pd and Ag (46 and 47, respectively), it is nearly impossible to distinguish between these atoms from single-crystal ( $\text{MoK}\alpha$  radiation) diffraction data. The refinement indicated five metallic positions, which one of them was assigned as Bi. The remaining metallic sites show multiplicities 2:8:8:4. Considering the empirical chemical composition  $\text{Pd}_{8.91}\text{Ag}_{2.03}\text{Bi}_{2.06}\text{S}_{4.00}$  ( $Z = 2$ ) and coordination environment of the 4e site, which was very different from the others (see structure description), the 4e site was refined as Ag position. Next, refinement cycles included all anisotropic displacement parameters, which revealed too large a value for Pd(2) position ( $U_{\text{eq}}(\text{Pd}2) = 0.0146 \text{ \AA}^2$  cf. 0.0082 and 0.080  $\text{ \AA}^2$  for Pd(1) and Pd(3), respectively). Refinement of occupancy factors yielded 0.88 occupancy for the Pd(2) position; other positions were found to be fully occupied. Final refinement in the  $I4/mmm$  space group for 21 parameters converged smoothly to the

$R = 0.0310$  and  $wR = 0.0815$  for 221 observed reflections. Details of data collection, crystallographic data, and refinement are given in Table 3.

**Table 3.** Crystallographic data for the selected crystal of synthetic thalhammerite, Pd<sub>9</sub>Ag<sub>2</sub>Bi<sub>2</sub>S<sub>4</sub>.

Crystal Data	
Chemical formula (idealized)	Pd <sub>9</sub> Ag <sub>2</sub> Bi <sub>2</sub> S <sub>4</sub>
Space group	<i>I4/mmm</i> (No. 139)
<i>a</i> [Å]	8.0266(2)
<i>c</i> [Å]	9.1531(2)
<i>V</i> [Å <sup>3</sup> ]	589.70(2)
<i>Z</i>	2
Crystal size (mm)	0.034 × 0.027 × 0.013
Data Collection	
Diffractometer	SuperNova
Temperature (K)	293
Radiation	MoKα (0.7107 Å)
Theta range (°)	5.08–27.62
Reflections collected	3659
Independent reflections	226
Unique observed reflections [ <i>I</i> > 3(σ)]	221
Index ranges	−10 < <i>h</i> < 10 −10 < <i>k</i> < 10 −11 < <i>l</i> < 11
Absorption correction method	Empirical
Structure Refinement	
Refinement method	Full matrix least-squares on <i>F</i> <sup>2</sup>
Parameters/restraints/constraints	21/0/0
<i>R</i> , <i>wR</i> (obs)	0.0310/0.0815
<i>R</i> , <i>wR</i> (all)	0.0318/0.0817
Largest diff. peak and hole (e <sup>−</sup> /Å <sup>3</sup> )	1.20/−5.20

Atom coordinates and displacement parameters are listed in Table 4. Table 5 shows selected bond lengths.

**Table 4.** Fractional coordinates and anisotropic displacement parameters (Å<sup>2</sup>) for synthetic thalhammerite.

Atom	Pd(1)	Pd(2) *	Pd(3)	Ag	Bi	S
Wyckoff Position	2 <i>a</i>	8 <i>f</i>	8 <i>j</i>	4 <i>e</i>	4 <i>d</i>	8 <i>h</i>
<i>x</i>	1/2	1/4	1/2	1/2	1/2	0.2081(4)
<i>y</i>	1/2	1/4	0.2027(2)	1/2	0	0.2081(4)
<i>z</i>	1/2	1/4	0	0.1810(2)	1/4	0
<i>U</i> <sub>11</sub>	0.0069(8)	0.0104(8)	0.0077(7)	0.0098(6)	0.0090(4)	0.0071(12)
<i>U</i> <sub>22</sub>	0.0069(8)	0.0104(8)	0.0097(7)	0.0098(6)	0.0090(4)	0.0071(12)
<i>U</i> <sub>33</sub>	0.0109(13)	0.0051(10)	0.0077(7)	0.0082(9)	0.0079(6)	0.012(2)
<i>U</i> <sub>12</sub>	0	0.0017(6)	0	0	0	−0.0018(15)
<i>U</i> <sub>13</sub>	0	0.0011(4)	0	0	0	0
<i>U</i> <sub>23</sub>	0	0.0011(4)	0	0	0	0
<i>U</i> <sub>eq</sub>	0.0083(6)	0.0086(5)	0.0084(4)	0.0106(4)	0.0086(3)	0.0087(9)

\* Refined with 0.88 occupancy.

**Table 5.** Selected bond distances (Å) in the thalhammerite crystal structure.

Pd(1)	4 × S	2.362(3)	Ag	4 × Pd(3)	2.905(1)
	2 × Ag	2.919(2)		4 × Pd(2)	2.9073(4)
Pd(2)	2 × S	2.3372(7)			
	2 × Bi	2.8378(1)	Bi1	4 × Pd(3)	2.808(1)
	2 × Ag	2.9073(4)		4 × Pd(2)	2.8378(1)
	4 × Pd(3)	3.0670(2)			

Table 5. Cont.

Pd(3)	2 × S	2.343(3)
	2 × Bi	2.8080(8)
	2 × Ag	2.905(2)
	4 × Pd(2)	3.0670(2)

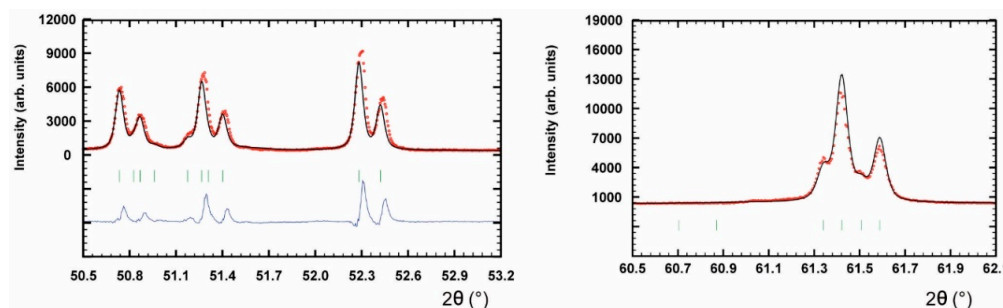
It should be noted that the refined tetragonal structure model of thalhammerite is only a substructure. As was revealed by subsequent Rietveld refinement (see below), the powder X-ray diffraction pattern of synthetic thalhammerite shows at medium and high diffraction angles a few very weak unindexed peaks and very subtle peak splitting, which cannot be fitted using the tetragonal model. Attempts to refine the structure from single-crystal data in rhombic subgroups of  $I4/mmm$  (i.e.,  $Fmmm$ ,  $Immm$ ) led to negligible lowering of  $R$ -factors (e.g., from 0.0313 to 0.0293) with a rapid increase of the refined parameters and correlations between them. Refinements in monoclinic subgroups failed. Additionally, neither of these low-symmetry models describe all peak splitting observed in powder diffraction patterns of synthetic thalhammerite. Therefore, we proposed only the tetragonal average substructure of thalhammerite, leaving some aspects of the structure unclear.

### 5.2. Powder X-ray Diffraction

The powder XRD pattern of synthetic thalhammerite was collected in the Bragg-Brentano geometry on a Bruker D8 Advance diffractometer equipped with the LynxEye XE detector and  $\text{CuK}\alpha$  radiation. The data were collected in the range from  $10^\circ$  to  $100^\circ$   $2\theta$  with a step size of  $0.005^\circ$   $2\theta$  and 2 s counting time per step. The structure model obtained from a single-crystal XRD study of synthetic thalhammerite was used as a starting structural model in the subsequent Rietveld refinement. The FullProf program [7] was used and the pseudo-Voigt function was used to generate the shape of the diffraction peaks. The refined parameters include those describing peak shape and width, peak asymmetry, unit-cell parameters, the occupancy parameter of the Pd(2) position, and six isotropic displacement parameters.

In total, 17 parameters were refined. No fractional coordinates were refined. The final cycles of Rietveld refinement converged to the agreement factors  $R_p = 0.077$  and  $R_{wp} = 0.115$ . The refinement indicated 7 wt %  $\text{Pd}_3\text{Bi}_2\text{S}_2$  ( $I2_13$ ) impurity in the investigated sample.

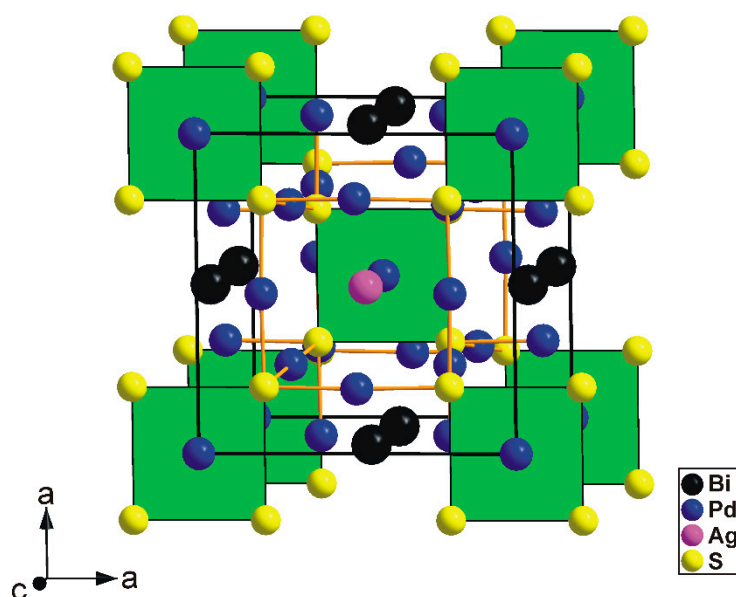
Figure 3 depicts two details of final Rietveld plot showing weak, however discernible, peak splitting at middle and high diffraction angles of  $2\theta$  (i.e., above  $50^\circ$ ). Attempts to index all observed diffractions in the powder pattern in the large and/or lower symmetry unit-cell remained unsuccessful and, therefore, the structure refinement was limited to the tetragonal substructure. Table 6 presents powder diffraction data for thalhammerite.



**Figure 3.** Details of the Rietveld profiles of synthetic thalhammerite showing the weak peak splitting, which cannot be fitted using the tetragonal cell. The observed (circles), calculated (solid), and difference profiles are shown. The vertical bars correspond to Bragg reflections.

## 6. Structure Description

The tetragonal substructure of thalhammerite contains three Pd, one Ag, Bi, and S sites, respectively. All sites, except the Pd(2) position, were found to be fully occupied. Its crystal structure is shown in Figure 4.



**Figure 4.** Crystal structure of thalhammerite showing the  $[PdS_4]$  squares and Pd–S bonds. Unit-cell edges are highlighted. Details show the Rietveld profiles of synthetic thalhammerite showing the weak peak splitting, which cannot be fitted using the tetragonal cell.

### 6.1. Coordination of Cations

The Pd(1) position is in the centre of regular square of S atoms with Pd(1)–S distances of 2.362(3) Å. The coordination is perfectly planar. Similar coordination was observed in vysotskite, PdS [8], which shows similar Pd–S separation of 2.34 Å. Such coordination geometry is typical for low-spin  $4d^8$  Pd<sup>2+</sup> cation in normal sulfides with M:S ratio equal to or smaller to one [9]. The Pd(1) coordination is further completed by two Ag atoms at 2.919(2) Å lying perpendicular to the  $[M(1)S_4]$  squares.

The Pd(2) (refined to 0.88 occupancy of Pd) and Pd(3) sites form complex polyhedron. Both Pd positions are coordinated by two S atoms at distances 2.3372(7) and 2.343(3) Å, a value very close to the Pd–S distance of 2.334(4) Å observed for the zig-zag chains in the structure of kravtsovite, PdAg<sub>2</sub>S [1]. Whereas the S–Pd(2)–S group is perfectly linear, the S–Pd(3)–S shows a bonding angle of 177.9(1)°. Pd(2) is further coordinated by two Bi (2.8378(1) Å), two Ag (2.9073(4) Å), and two Pd(3) (3.0670(2) Å) atoms. Pd(3) also shows two Bi (2.8080(8) Å), two Ag (2.905(2) Å), and four Pd(3) (3.0670(2) Å) short contacts.

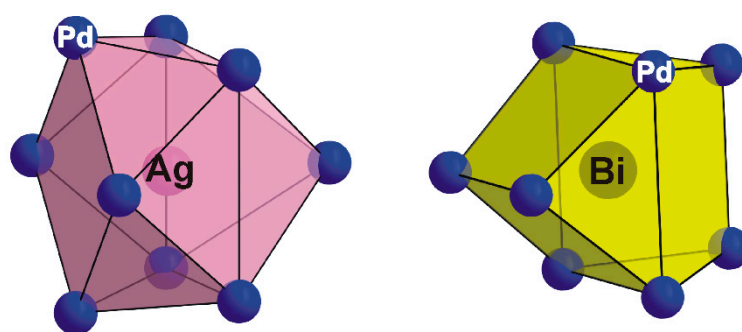
Ag site is surrounded by nine Pd atoms (Figure 5) forming a mono-capped tetragonal antiprismatic coordination. The Ag–Pd distances are in the range of 2.905(1) Å to 2.919(2) Å, comparable to those observed in lukkulaisvaaraita (Pd–Ag: 2.891(4)–3.037(4) Å; [10], where Ag atoms display tetragonal antiprismatic coordination.

As is shown in Figure 5, the Bi atom is coordinated by eight Pd atoms to form a bi-capped trigonal prism with Bi–Pd bond distances ranging from 2.808(1) to 2.8378(1) Å, values slightly shorter than those observed in structure of monoclinic PdBi (2.84–2.95) Å; [11]. There are no short (<3.5 Å) Bi–S contacts in the thalhammerite crystal structure. This contrasts with the environment of Bi in structure of chemically-related vymazalováite, Pd<sub>3</sub>Bi<sub>2</sub>S<sub>2</sub> [2,12], where Bi atoms show one additional S contact at 3.22(3) Å.



**Table 6.** X-ray powder diffraction data of thalhammerite (CuK $\alpha$  radiation, Bruker D8 Advance, Bragg-Brentano geometry). Only reflections with  $I_{(\text{obs})} \geq 1$  are listed.

$I_{(\text{obs})}$	$h$	$k$	$l$	$d_{(\text{meas})}$	$d_{(\text{calc})}$
11	1	0	1	6.0364	6.0338
11	1	1	0	5.6790	5.6767
13	0	0	2	4.5752	4.5736
8	2	0	0	4.0155	4.0140
18	1	1	2	3.5620	3.5615
24	2	1	1	3.3428	3.3420
2	2	0	2	3.0181	3.0169
9	1	0	3	2.8510	2.8504
46	2	2	0	2.8393	2.8383
21	3	0	1	2.5685	2.5684
100	2	2	2	2.4122	2.4117
61	1	2	3	2.3245	2.3241
48	0	0	4	2.2873	2.2868
29	1	3	2	2.2201	2.2197
2	2	3	1	2.1637	2.1634
17	1	1	4	2.1213	2.1212
40	4	0	0	2.0072	2.0070
3	3	3	0	1.8923	1.8922
8	4	0	2	1.8377	1.8378
15	2	3	3	1.7981	1.7982
18	2	2	4	1.7805	1.7807
23	3	3	2	1.7481	1.7485
4	1	3	4	1.6991	1.6991
2	4	2	2	1.6711	1.6710
5	1	4	3	1.6413	1.6410
1	2	1	5	1.6299	1.6300
4	4	3	1	1.5814	1.5814
1	5	1	0	1.5743	1.5744
8	0	3	5	1.5102	1.5103
30	4	0	4	1.5085	1.5085
9	1	1	6	1.4723	1.4724
13	4	4	0	1.4193	1.4192
7	4	4	2	1.3554	1.3554
12	2	2	6	1.3431	1.3431
9	2	5	3	1.3395	1.3393
1	3	5	2	1.3185	1.3184
19	3	1	6	1.3070	1.3070
7	1	5	4	1.2969	1.2968
9	6	2	0	1.2694	1.2693
3	1	2	7	1.2279	1.2279
18	6	2	2	1.2231	1.2231
1	1	6	3	1.2113	1.2112
10	4	4	4	1.2059	1.2058
11	3	3	6	1.1872	1.1871



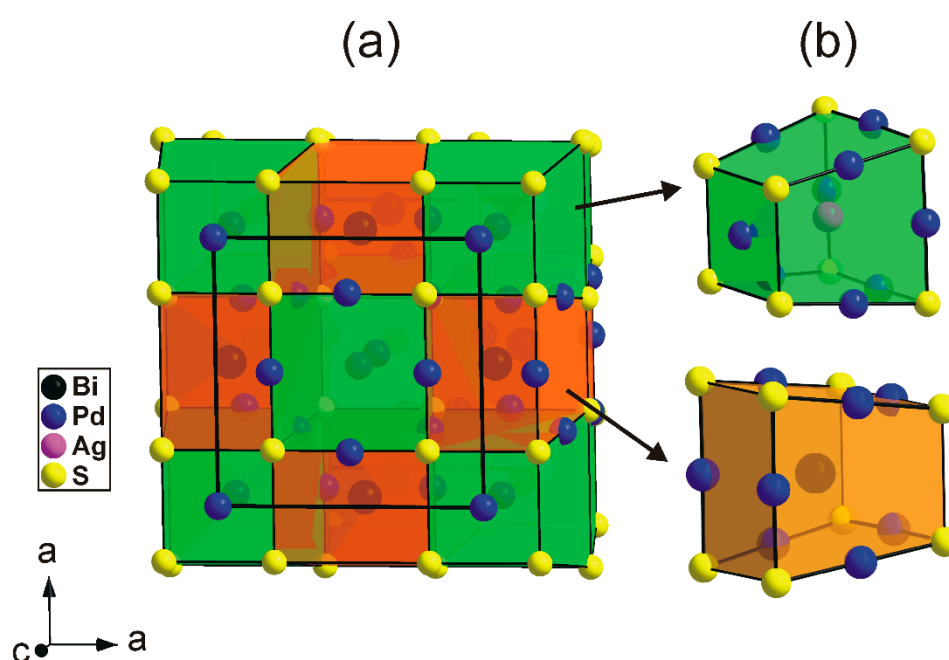
**Figure 5.** Coordination polyhedra of Ag (mono-capped tetragonal antiprism) and Bi (bi-capped trigonal prism) in the thalhammerite structure.

## 6.2. Modular Description

The thalhammerite crystal structure forms a three-dimensional framework. It contains features typical for intermetallic compounds (e.g., complex crystallochemical environment of metals) and, therefore, cannot be presented using a traditional cation-based coordination polyhedra approach.

Alternatively, the structure of thalhammerite can be conveniently described as an arrangement of two types of building blocks (cuboids) having common S atoms at the corners (Figure 6).

The first block (green in Figure 6) contains the  $[\text{PdS}_4]$  squares forming one face of the block and Ag atoms in its centre. Pd atoms are approximately located to the midpoints of the longer S–S edges. The second block (orange in Figure 6) contains Bi atoms in its centre. By analogy with the first block, the Pd atoms are located to the midpoints of the longer S–S edges. In the thalhammerite structure, two types of block alternate in a chess-board fashion within the (001) plane and form chains along the c axis (Figure 6). It should be mentioned that, neglecting the Ag and Bi atoms, the packing of the green blocks automatically generates their duals, and the orange block, vice versa.



**Figure 6.** (a) Arrangement of two types of building blocks in the thalhammerite structure. (b) Detailed view showing the block containing Ag (green) and Bi (orange) atoms.

## 6.3. Relation to Other Minerals

The thalhammerite structure represents a unique structure type, and no exact structural analogue is hitherto known. It is worth noting that its structure merges structure motives typical for polar chalcogenides and intermetallic compounds. The  $[\text{Pd}(1)\text{S}_4]$  square-planar coordination is a hallmark of Pd-bearing sulfides with an M:S ratio equal to, or slightly smaller than, one. Contrary to that, (almost) linear coordination of Pd by two S atoms and number of further metal-metal contacts resulting in complex coordination geometry, can be observed in sulphides with intermetallic behaviour (e.g., kravtsovite  $\text{PdAg}_2\text{S}$ , Vymazalová et al., 2017 [1]).

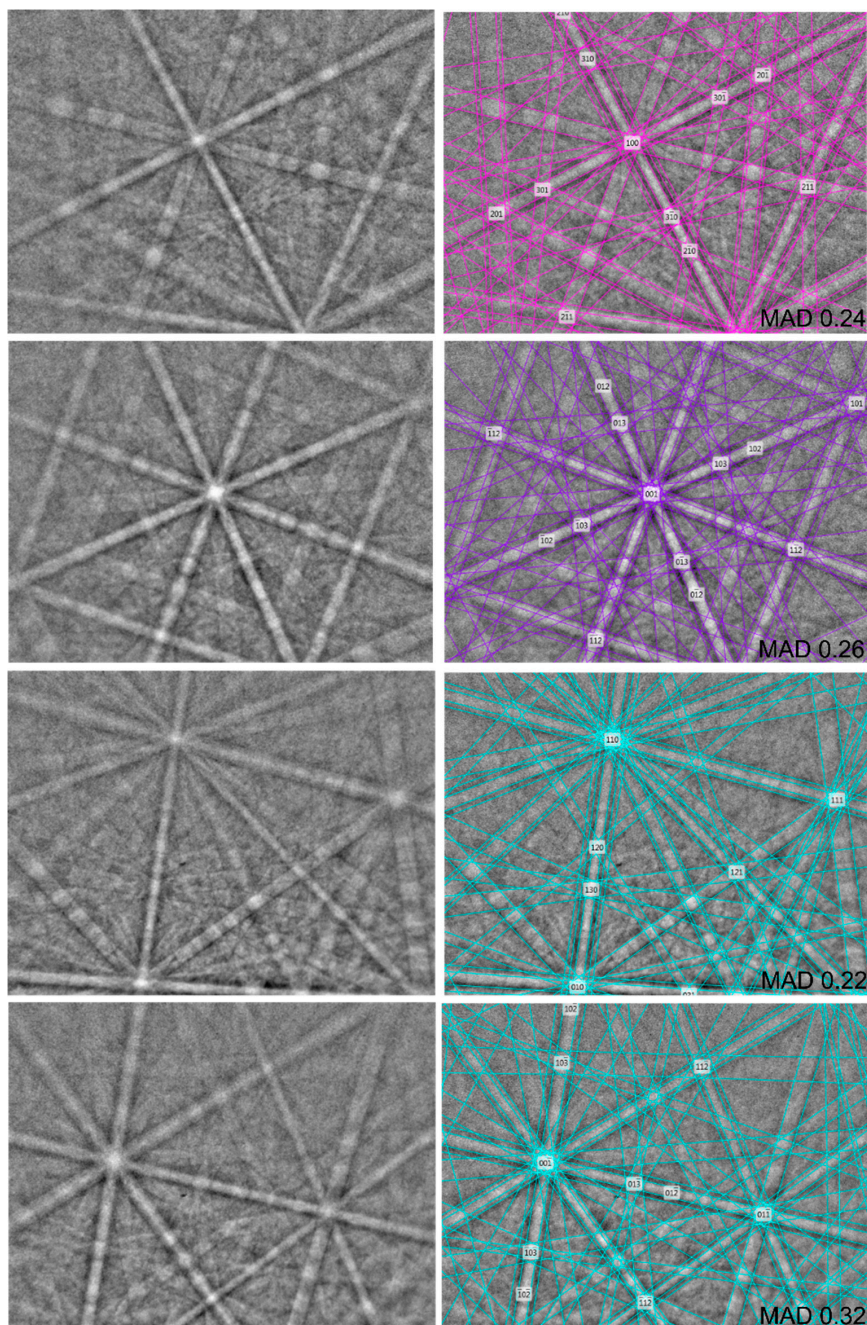
Another chemically-related mineral, coldwellite,  $\text{Pd}_3\text{Ag}_2\text{S}$  (McDonald et al., 2015 [13]), adopts a cubic  $\beta$ -Mn-like structure and, hence, differs substantially from that of thalhammerite.

## 7. Proof of Identity of Natural and Synthetic Thalhammerite

The structural identity between the synthetic  $\text{Pd}_9\text{Ag}_2\text{Bi}_2\text{S}_4$  and the natural material was confirmed by electron back-scattering diffraction (EBSD) and Raman spectroscopy.

### 7.1. Electron Back-Scattering Diffraction

The structural identity between the natural material and the synthetic  $\text{Pd}_9\text{Ag}_2\text{Bi}_2\text{S}_4$  was confirmed by EBSD. A TESCAN Lyra 3GM field emission scanning electron microscope combined with EBSD system (Oxford Instruments AztecHKL system with NordlysNano EBSD camera) was used for the measurements. The surface of natural sample was prepared for investigation by broad beam argon ion milling using Gatan PECS II system operated at 1 kV. The solid angles calculated from the patterns were compared with our structural model for  $\text{Pd}_9\text{Ag}_2\text{Bi}_2\text{S}_4$  synthetic phase match containing 12 reflectors to index the patterns. The EBSD patterns (also known as Kikuchi patterns) obtained from the natural material (>50 measurements on different spots on natural thalhammerite grains) were found to match the patterns generated from our structural model for  $\text{Pd}_9\text{Ag}_2\text{Bi}_2\text{S}_4$  synthetic phase, Figure 7.

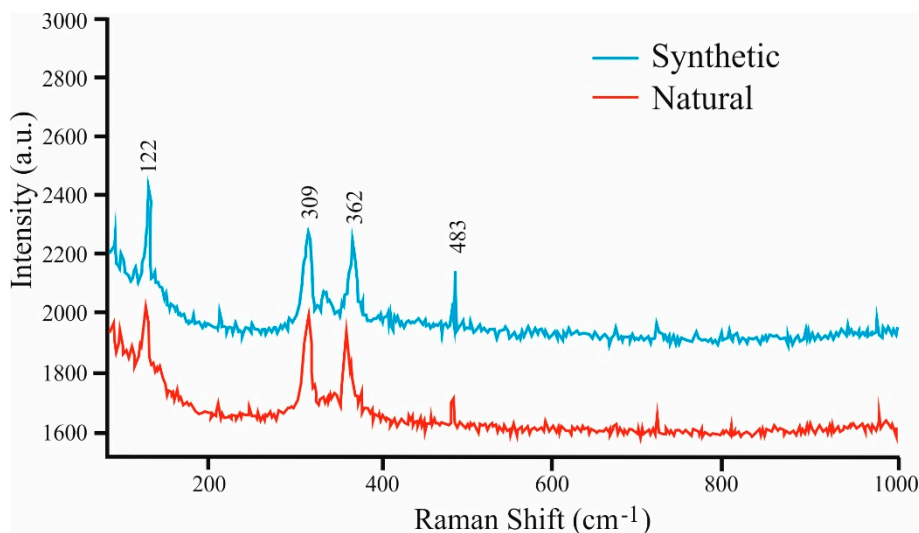


**Figure 7.** EBSD image of natural thalhammerite; in the right pane, the Kikuchi bands are indexed.

The values of the mean angular deviation (MAD, i.e., goodness of fit of the solution) between the calculated and measured Kikuchi bands range between  $0.22^\circ$  and  $0.48^\circ$ . These values reveal a very good match; as long as values of mean angular deviation are less than  $1^\circ$ , they are considered as indicators of an acceptable fit (HKL Technology, 2004).

## 7.2. Raman Spectroscopy

The Raman spectroscopy technique was applied to verify the structural identity between the synthetic  $\text{Pd}_9\text{Ag}_2\text{Bi}_2\text{S}_4$  and the natural material (Figure 8).



**Figure 8.** Comparison of Raman spectra in the synthetic  $\text{Pd}_9\text{Ag}_2\text{Bi}_2\text{S}_4$  and in the natural material.

Raman spectra were obtained using a LABRAM (ISA Jobin Yvon) instrument installed at the University of Leoben, Austria. A frequency-doubled 100 mW Nd:YAG laser with an excitation of a wavelength of  $\lambda = 532.6$  nm was used. The obtained Raman spectra of natural and synthetic  $\text{Pd}_9\text{Ag}_2\text{Bi}_2\text{S}_4$  show four discernible absorption bands at the following values: 122, 309, 362, and  $483\text{ cm}^{-1}$  (see Figure 8).

The EBSD study, Raman spectra, chemical identity and optical properties confirmed the identity of the natural and synthetic materials and thereby legitimise the use of the synthetic phase for the complete characterization of thalhammerite.

**Author Contributions:** All the authors (A.V., F.L., S.F.S., V.V.K., C.J.S., J.P., F.Z., G.G. and R.B.) discussed the obtained results, evaluated the data and wrote the article together. A.V. designed the article and conceived experiments; S.F.S. provided the samples and geological background; F.L., V.V.K., J.P. obtained the crystallographic data; C.J.S. provided optical properties; F.Z., G.G. and R.B. studied thalhammerite by Raman and evaluated the chemical data. All the authors revised and edited the manuscript.

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