

Reactive desorption of CO hydrogenation products under cold pre-stellar core conditions

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ABSTRACT

The astronomical gas-phase detection of simple species and small organic molecules in cold pre-stellar cores, with abundances as high as $\sim 10^{-8}$ - 10^{-9} n_H, contradicts the generally accepted idea that at 10 K, such species should be fully frozen out on grain surfaces. A physical or chemical mechanism that results in a net transfer from solid-state species into the gas-phase offers a possible explanation. Reactive desorption, i.e., desorption following the exothermic formation of a species, is one of the options that has been proposed. In astronomical models the fraction of molecules desorbed through this process is handled as a free parameter, as experimental studies quantifying the impact of exothermicity on desorption efficiencies are largely lacking. In this work, we present a detailed laboratory study with the goal to derive an upper limit for the reactive desorption efficiency of species involved in the CO – H₂CO – CH₃OH solid-state hydrogenation reaction chain. The limit for the overall reactive desorption fraction is derived by precisely investigating the solid-state elemental carbon budget, using reflection absorption infrared spectroscopy and the calibrated solid-state band strength values for CO, H₂CO and CH₃OH. We find for temperatures in the range of 10 to 14 K, an upper limit of 0.24 ± 0.02 for the overall elemental carbon loss upon CO conversion into CH₃OH. This corresponds with an effective reaction desorption fraction of ≤ 0.07 per hydrogenation step, or ≤ 0.02 per H-atom induced reaction, assuming that H-atom addition and abstraction reactions equally contribute to the overall reactive desorption fraction along the hydrogenation sequence. The astronomical relevance of this finding is discussed.

1. INTRODUCTION

Surface chemistry in the earliest phase of molecular cloud evolution is determined by a rather low particle density ($\sim 10^3$ - 10^4 cm⁻³) at temperatures around $\lesssim 15$ K. In such environments, chemical processes on dust grains are largely dominated by H-atom addition reactions, forming an H₂O-rich

polar ice layer on silicate dust grains by recombining oxygen and hydrogen atoms (Tielens & Hagen 1982; Tielens et al. 1991; Öberg et al. 2011; Boogert, Gerakines & Whittet 2015; Miyauchi et al. 2008; Ioppolo et al. 2008, 2010; Cuppen et al. 2010; van Dishoeck, Herbst & Neufeld 2013; Linnartz, Ioppolo & Fedoseev 2015). During this stage also other molecules, like NH_3 and CH_4 form in $\text{N}+\text{H}$ and $\text{C}+\text{H}$ addition reactions (Hasegawa, Herbst & Leung 1992; Hiraoka et al. 1995, 1998; Hidaka et al. 2011; Fedoseev et al. 2015). In the next phase, temperatures become as low as ~ 10 K and densities increase ($\sim 10^{4-5} \text{ cm}^{-3}$); gaseous CO accretes onto the pre-formed H_2O -rich ice and forms a CO-rich apolar ice coating (Pontoppidan 2006), an event often referred to as the ‘catastrophic CO freeze-out stage’. The time-scale for the CO freeze-out phase is about 10^{4-5} years ($1 \times 10^9/n_{\text{H}}$ years where $n_{\text{H}} = 2n(\text{H}_2) + n(\text{H})$, Willacy & Millar 1998), which is shorter than the typical lifetime of a low-mass prestellar core ($\sim 4.5 \times 10^5$ years, Enoch et al. 2008). The CO-coating is regarded as a starting point in the hydrogenation scheme that leads primarily to the formation of H_2CO and CH_3OH through successive H-atom additions. This reaction scheme has been studied in much detail before, in a gas-grain model by Tielens & Hagen (1982), in astrochemical simulations and theoretical studies (Charnley et al. 1997; Cuppen et al. 2009; Chang & Herbst 2012) and in a number of systematic laboratory experiments (Hiraoka et al. 1994; Watanabe & Kouchi 2002; Zhitnikov & Dmitriev 2002; Fuchs et al. 2009). The solid-state laboratory and astrochemical modelling conclusions are in agreement with the observed CH_3OH ice and gas abundances in star forming regions that cannot be explained through a pure gas-phase radiative association route (Geppert et al. 2005; Garrod et al. 2006). Analysis of Very Large Telescope (VLT) ice data shows that CH_3OH and CO must be mixed in interstellar ices and exist in a water poor environment (Cuppen et al. 2011; Penteadó et al. 2015), fully consistent with the picture that methanol is formed via sequential hydrogenation of accreted CO molecules. Moreover, recent theoretical and laboratory studies revealed that recombination of the reactive intermediates, HCO, CH_3O and CH_2OH formed through H-atom addition and abstraction reactions along the $\text{CO}-\text{H}_2\text{CO}-\text{CH}_3\text{OH}$ chain, also result in the formation of larger complex organic molecules (COMs), like ethylene glycol, glycolaldehyde, methyl formate, and even glycerol in the solid-state (Butscher et al. 2015; Chuang et al. 2016, 2017; Fedoseev et al. 2015, 2017; Garrod et al. 2006; Woods et al. 2013).

For the typical pressure and temperature conditions in dark and dense clouds, all these species except H_2 and He are expected to be fully depleted into the solid-state. However, astronomical observations show unexpectedly large abundances of gaseous CH_3OH and COMs (Bacmann et al. 2012; Cernicharo et al. 2012; Öberg et al. 2010, Jimenez-Serra et al. 2016). In dark and pre-stellar clouds it is too cold for thermal desorption to take place; thus, non-thermal mechanisms need to be considered to explain the observed gas phase abundances, but effective low temperature desorption processes remain poorly understood.

Different mechanisms have been proposed, investigated and discussed, to explain how frozen species can sublimate. Up to recently, UV induced photo-desorption was considered as an efficient process for both non-dissociative desorption, i.e., CO photo-desorption, and dissociative processes

eventually followed by fragment recombination in the ice or kick-out processes, i.e., H₂O photo-desorption (Andersson & van Dishoeck 2008; Arasa et al. 2015). However, experiments on methanol ice exhibit substantial molecular fragmentation, including CH₃O, CH₂OH and CH₃, and with photo-desorption rates far too low to transfer substantial amounts of intact CH₃OH into the gas phase (Bertin et al. 2016; Cruz-Diaz et al. 2016). It is expected that this will also apply to COMs larger than methanol; photo-desorption and photo-dissociation are here clearly connected. A possible way to circumvent fragmentation upon photo-excitation is UV induced co-desorption; the excitation of one species results in desorption of another species. For mixed CO and N₂ ices, the N₂ photo-co-desorption rate (defined as indirect DIET photo-desorption rate) was found to be a highly effective process, i.e., $\sim 10^{-2}$ molecule·cm⁻²·photon⁻¹ for photon energies between 7.9 and 9.5 eV (Bertin et al. 2013), but in ice mixtures of CO and CH₃OH no methanol could be detected upon CO excitation (Bertin et al. 2016). In a similar way, low temperature thermal co-desorption of a less volatile species like methanol with CO was only found to yield very low upper limits of 7×10^{-7} molecules per CO (Ligterink et al., submitted). Local heating of an ice mantle induced by the impacting cosmic rays offers another possible mechanism (Léger et al. 1985; Ivlev et al. 2015), but given the high energies involved, it is rather unlikely that this would be a less destructive process than photo-desorption and this process is generally less effective for strongly bound molecules like CH₃OH than for CO. An interesting alternative is reactive desorption.

Reactive desorption may contribute to the gas-phase enrichment of both simple and complex species; directly, upon formation, formaldehyde, methanol or other COMs are instantly released into the gas phase because of the available excess energy of several eV, or indirectly, by releasing intermediates that cannot be efficiently formed in gas-phase chemical networks, but may act as precursor in COM gas-phase networks. A reactive co-desorption process cannot be excluded either, but the excess energy upon bond formation may rapidly dissipate into the ice lattice before being absorbed by another species. These three scenarios may take place simultaneously, but with different efficiencies.

In astrochemical models, the reactive desorption fraction, i.e., the efficiency of product desorption from the surface after reaction, is usually a free parameter, chosen in a range of 0.01-0.1 to account for the steady state gas-phase abundances of simple species as observed in the later stages of dense clouds (Garrod et al. 2007; Vasyunin & Herbst 2013). The initial focus was on energy release upon H₂ formation (Duley & Williams 1993), and later extended to other reactions (Garrod et al. 2007; Vasyunin & Herbst 2013; Cazaux et al. 2016; Minissale et al. 2016; Fredon et al. 2017). In the latter work, reactive desorption of species was also proposed as a starting point for low temperature gas-phase formation of COMs (see also Balucani et al. 2015; Vasyunin et al. 2017; Taquet et al. 2016; Rivilla et al. 2017). Currently, the relative solid-state and gas-phase formation efficiencies for dark cloud conditions are under renewed debate.

So far, there have been few attempts to experimentally study the reactive desorption of the two main CO hydrogenation products, H₂CO and CH₃OH. This has a reason; the simultaneous

occurrence of H-atom induced addition and abstraction reactions, i.e., forward and backward steps along the reaction chain (Fig. 1), make it hard to de-convolute the individual reaction steps.

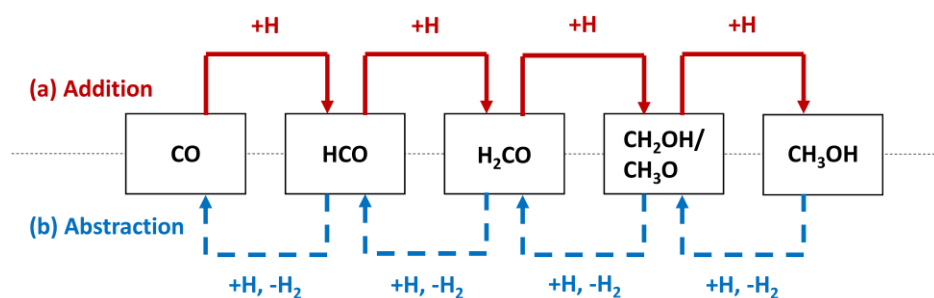


Figure 1. The CO-H₂CO-CH₃OH hydrogenation sequence scheme. The solid line and dashed line is for H-atom (a) addition and (b) abstraction reactions, respectively.

Previously reported experimental desorption fractions for reactive steps in the CO hydrogenation chain are high. Hidaka et al. (2004) stated that the H₂CO reactive desorption fraction during hydrogenation could be as high as 60 % to rationalize their finding that the column density of the newly formed CH₃OH product was less than the consumed column density of the predeposited H₂CO, i.e., $\Delta N(\text{CH}_3\text{OH})/\Delta N(\text{H}_2\text{CO}) < 1$. Minissale et al. (2016) found a reactive desorption fraction of 40 % for the single reaction $\text{HCO} + \text{H} \rightarrow \text{CO} + \text{H}_2$. This reaction was studied in the sub-monolayer region on a bar of oxidized highly-oriented pyrolytic graphite (HOPG) substrate using quadrupole mass spectrometry. For a more astronomically relevant thick ice (several tens of monolayers) these conditions are likely not fully representative and therefore reactive desorption fractions may be substantially different as well.

The work presented here is motivated by the lack of experimentally determined reactive desorption fractions for hydrogenation reactions in multi-layered CO-rich interstellar ice analogues, i.e., for an astrochemically relevant ice thickness of $\sim 0.01 \mu\text{m}$ (Boogert et al. 2015). Rather than trying to disentangle the individual reactions and to measure the individual reaction products in the gas-phase, we focus on the full CO-H₂CO-CH₃OH hydrogenation sequence. This is realized by estimating as precisely as possible, the difference in solid-state elemental carbon budget between the amount available from the original species and the resulting solid-state hydrogenation products. We assume that the observed difference is fully or at least partially due to reactive desorption, i.e., the resulting value reflects an upper limit. In order to perform such experiments, it is needed to derive setup specific band strength values in $\text{cm}\cdot\text{molecule}^{-1}$ for each of the involved species and their ice mixtures.

In the following section the experimental procedure is described. The results are presented in section 3. In the final section the astronomical relevance of the data and conclusions are presented.

2. EXPERIMENTAL

2.1 Approach

The total consumption of elemental carbon of the original species should be equal to the total amount of elemental carbon of the formed species, both in the solid state and, after reactive desorption, in the gas phase.

$$N_{consumed}(C) = \Delta N_{solid}(CO) = \Delta N_{solid}(H_2CO) + \Delta N_{solid}(CH_3OH) + N_{RD}(C) \quad (1)$$

Where $\Delta N_{solid}(CO)$ is the column density of CO consumption, $\Delta N_{solid}(H_2CO)$ ($\Delta N_{solid}(CH_3OH)$) is the net column density of the H₂CO (CH₃OH) formation, and $N_{RD}(C)$ is the column density of carbon-bearing species upon reactive desorption mechanism after H-atom bombardment. This means that an upper limit for $N_{RD}(C) / N_{consumed}(C)$ can be derived by comparing accurately the total difference in CO abundance and H₂CO and CH₃OH abundances. In practice, equation (1) is only an approximation, as it does not take into account other loss channels, such as the formation of other products than H₂CO and CH₃OH; reactive intermediates can recombine and form COMs larger than methanol, as recently shown by Chuang et al. (2016) and Fedoseev et al. (2017). The impact of this restriction is discussed in Section 3.

2.2 Experimental setup

All experiments are performed using SURFRESIDE², an ultra-high vacuum setup fully optimized to study non-energetic atom addition reactions in interstellar ice analogues at cryogenic temperatures. Details of the design and operation of SURFRESIDE² are available from Ioppolo et al. (2013) and Fedoseev et al. (2016). The base pressure in the main chamber is $\sim 10^{-10}$ mbar, and the corresponding H₂O contamination through background gas accretion is negligible ($< 6.3 \times 10^{10}$ molecule·cm⁻²·s⁻¹). A gold-plated copper substrate is mounted in the center of the main chamber and cooled by a recently upgraded closed-cycle helium cryostat that allows to vary the substrate temperature between 8 and 450 K. A silicon diode is used to monitor the temperature with 0.5 K absolute accuracy. It should be noted that the temperatures used in the present manuscript are about ~ 2 K lower than the corresponding numbers used in Fuchs et al. (2009); new and more precise thermal sensors were installed on the substrate. Gaseous species, i.e., H₂CO obtained by thermal decomposition of paraformaldehyde (Sigma-Aldrich, 95%) at 60-80 °C under vacuum, CH₃OH purified through multiple freeze-pump-thaw cycles and CO (Linde 2.0), are admitted into the UHV chamber through high-precision leak valves mounted under an angle of 22 degrees from the substrate surface normal. The ices are grown with sub-monolayer precision with CO, H₂CO and CH₃OH deposition rates of $\sim 1.7 \times 10^{12}$, $\sim 1.4 \times 10^{12}$ and $\sim 2.2 \times 10^{12}$ molecules · cm⁻²·s⁻¹, respectively. **The corresponding ice column densities for CO, H₂CO and CH₃OH are $\sim 6.0 \times 10^{15}$, $\sim 5.0 \times 10^{15}$ and $\sim 8.0 \times 10^{15}$ molecules·cm⁻², respectively.** The details of the exact deposition rate

calculations are provided later.

After a preset ice thickness of the precursor species is realized, H-atoms are introduced using a Hydrogen Atom Beam Source (HABS; [Tschersich 2000](#)), mounted in a second UHV chamber. A nose-shape quartz pipe is placed along the path to efficiently quench and thermalize excited H-atoms and non-dissociated molecules through multiple collisions with the walls of the pipe. The used H-atom flux in this work amounts to $\sim 8 \times 10^{12}$ atoms \cdot cm $^{-2}$ s $^{-1}$ with an absolute uncertainty of < 50 % ([Ioppolo et al. 2013](#)) and H-atoms reach the surface under an incident angle of 45 degrees from the substrate surface normal.

The ice composition is monitored *in situ*, before and during H-atom bombardment by Fourier Transform Reflection-Absorption InfraRed Spectroscopy (FT-RAIRS) in the range from 700 to 4000 cm $^{-1}$, with 1 cm $^{-1}$ resolution. The ice column density N in cm $^{-2}$ is derived from a modified Beer's law:

$$N = \frac{\ln 10 \cdot \int Abs(\nu) d\nu}{A'} \quad (2)$$

where $Abs(\nu)$ is the band absorbance, $d\nu$ is the wavenumber differential in cm $^{-1}$, and A' is the calibrated band strength in cm \cdot molecule $^{-1}$ in reflection mode. This RAIR band strength cannot be taken from literature data available for the IR transmission spectroscopy, as the signals obtained in reflection are enhanced through substrate dipole couplings. **The determination of RAIR band strength cannot be compensated for effective pathways only.** Therefore, a series of extra experiments (see section 2.3) has been performed to determine the band strength for each of the involved species for our specific experimental settings by using the interference pattern of a HeNe laser that is reflected from the growing ice sample at 10 K ([Baratta & Palumbo 1998](#); [Brunetto et al 2008](#); [Fulvio et al. 2009](#); [Bouilloud et al. 2015](#)). For pure ice the absolute column density is calculated by the equation:

$$N = \frac{d \cdot \rho \cdot N_a}{M} \quad (3)$$

where d is the thickness of ice in nm, ρ is the density in g \cdot cm $^{-3}$, N_a is the Avogadro's constant (6.022×10^{23} mol $^{-1}$), and M is the molar mass of the species. The ice thickness ($D_{growing}$) is experimentally determined by laser refractive interference:

$$D_{growing} = k \times \frac{\lambda}{2n \cdot \cos(\theta_f)} \quad (4)$$

where $\lambda = 632.8$ nm is the HeNe laser wavelength, n is the refractive index of a specific ice, $\theta_f \cong 3^\circ$ is the angle of refraction in the ice in degrees and k is the number of involved fringes ([Hollenberg & Dows 1961](#); [Westley et al. 1998](#)). The exact band strength values may vary with changing ice composition, structure, and temperature. The reported absolute uncertainty of this method is 5 %, a value that is largely determined by the density and refractive index uncertainty ([Baratta & Palumbo 1998](#); [Fulvio et al. 2009](#)). Moreover, as the same technique is used to determine CO, H $_2$ CO and CH $_3$ OH band strengths, the relative error between values is expected to be lower. This makes it

possible, in principle, to establish the obtained $N_{RD}(C)/N_{consumed}(C)$ with high confidence, even when chemistry changes relative ice compositions.

2.3 Band strength measurement for pure ices

The left-hand panel of Figure 2 (a-c) shows the HeNe laser interference pattern that changes with time as the ice thickness increases upon deposition of CO, H₂CO and CH₃OH onto the 10 K substrate. One single fringe is typically observed for the low-deposition rates used in this study. Ice density (ρ) and refractive index (n) may vary for different conditions (Bossa et al. 2015). **It is important to note that the precise ice density in the amorphous phase is generally unknown for the species studied in this work. In this study, the commonly accepted values in literature are taken as good approximate values. For CO, the reported ice density by Roux et al. (1980) was obtained at 20 K. The density may not be derived under the same deposition conditions as in our work. For H₂CO, no density value has been reported in solid phase; as a rough estimate the value measured in liquid phase is employed (Weast & Astle 1985). For CH₃OH, the ice density is derived from the value available for its crystalline structure (Mate et al. 2009).** Based on the density and refractive index values summarized in Bouilloud et al. (2015) - $0.80 \pm 0.01 \text{ g}\cdot\text{cm}^{-3}$ and 1.25 ± 0.03 for CO, $0.81 \pm 0.03 \text{ g}\cdot\text{cm}^{-3}$ and 1.33 ± 0.04 for H₂CO, and $1.01 \pm 0.03 \text{ g}\cdot\text{cm}^{-3}$ and 1.33 ± 0.04 for CH₃OH - CO, H₂CO and CH₃OH deposition rates are derived with values of $\sim(1.3 \pm 0.1) \times 10^{14}$, $\sim(4.7 \pm 0.4) \times 10^{13}$ and $\sim(5.6 \pm 0.5) \times 10^{13} \text{ molecules}\cdot\text{cm}^{-2}\cdot\text{s}^{-1}$. Note that these values differ in these experiments from the regular hydrogenation settings. **These choices are not ideal, but given the lack of more precise values, the best we can provide. Obviously, in the case that the 'real' density turns out to be different, this also effects the final ratio of $N_{RD}(C)/N_{consumed}(C)$; a higher density for CO and H₂CO in the end means that the actual ratio of $N_{RD}(C)/N_{consumed}(C)$ is higher.**

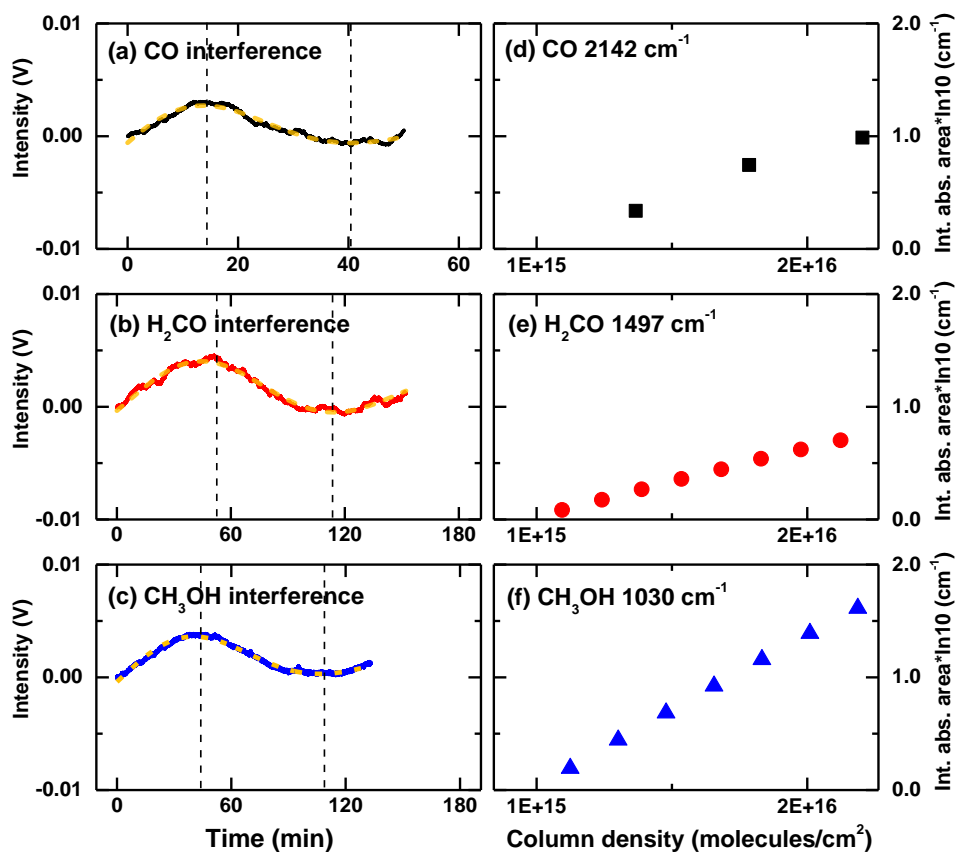


Figure 2. Left: Typical examples of the obtained interference fringes as function of time and their corresponding sinusoidal fittings (yellow dashed line) during 10 K substrate deposition of pure (a) CO, (b) H₂CO, and (c) CH₃OH. The vertical dashed lines (black) indicate the time for maximum and minimum amplitude. Right: The integrated absorption area for deposited (d) CO (2142 cm⁻¹ band), (e) H₂CO (1497 cm⁻¹ band) and (f) CH₃OH (1030 cm⁻¹ band) as a function of column density for the first 7 minutes of deposition.

The corresponding CO, H₂CO and CH₃OH plots, showing the column density upon ice deposition (horizontal axis) and the increasing IR absorption area (vertical axis) over time are presented in the right-hand panels of Figure 2 (d-e). As explained in the Experimental Section (Equation 2), from these plots the A' , band strength value, can be derived from molecule specific RAIR bands, following a linear fit. The fitted slopes for (d) CO (2142 cm⁻¹), (e) H₂CO (1497 cm⁻¹), and (f) CH₃OH (1030 cm⁻¹) yield absorption strength values A' of $(5.2 \pm 0.3) \times 10^{-18}$, $(3.2 \pm 0.3) \times 10^{-18}$, and $(7.1 \pm 0.6) \times 10^{-18}$ cm·molecule⁻¹, respectively. It should be noted that these values are only safe to use for identical experimental circumstances, i.e., identical IR incident angles, surface materials, etc.

Kerkhof et al. (1999) showed that the band strength of methanol in an ice mixture of CH₃OH:CO₂ (1:1) and in pure CH₃OH ice does not vary. We have checked this by performing a set of control experiments in which the CO, H₂CO and CH₃OH band strength values are compared with

those derived in CO:H₂CO (1:1), CO:CH₃OH (1:1) and CO:H₂CO:CH₃OH (1:1:1) mixed ices. This is realized in the following way. At first, a constant deposition rate of CO molecules is used through one of the dosing lines. Then, sequentially, H₂CO and CH₃OH are introduced through other dosing lines growing an ice mixture during co-deposition. Each time a new species is allowed into the main chamber, any changes in the CO absorbance integral area are carefully monitored. Then, in a similar way, changes in the H₂CO and CH₃OH absorbance integral area are measured. These changes are linearly correlated with their increasing column densities. From this it is found that the CO, H₂CO and CH₃OH band strengths are rather constant, with variations that do not exceed 2 % in CO, CO:H₂CO (1:1), CO:CH₃OH (1:1) and CO:H₂CO:CH₃OH (1:1:1) ices. RAIR spectra are used only as long as signals are far from saturation, to guarantee a linear correlation with the corresponding column densities (Teolis et al. 2007).

3. RESULTS AND DISCUSSION

3.1 Elemental carbon budget in CO hydrogenation experiments

Figure 3 presents RAIR difference spectra obtained after hydrogenation of pre-deposited CO ice for 180 min by an H-atom flux of $\sim 8 \times 10^{12}$ atoms \cdot cm⁻²s⁻¹ at (a) 10 K, (b) 12 K, and (c) 14 K, respectively. The negative peak at 2142 cm⁻¹ shows the consumption of the originally pre-deposited CO ice, and the positive peaks indicate the formation of two newly formed main products, i.e. H₂CO (1727, 1497, and 1250 cm⁻¹) and CH₃OH (1030 cm⁻¹). Very similar results have been reported and discussed extensively in previous studies of Watanabe et al. (2002) and Fuchs et al. (2009). No entrapment of the formed intermediate radicals such as HCO, CH₃O, and CH₂OH is found, as these easily react with free H-atoms or other radical species.

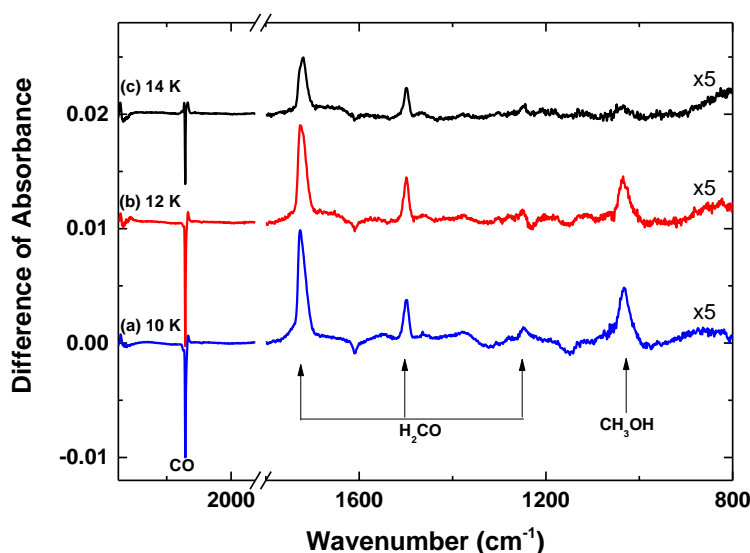


Figure 3. RAIR difference spectra obtained after the exposure of 6×10^{15} molecules cm⁻² pre-deposited CO to an H-atom flux of $\sim 8 \times 10^{12}$ atoms \cdot cm⁻²s⁻¹ for 180 min at (a) 10 K, (b) 12 K,

and (c) 14 K. Negative and positive signals reflect net consumption and formation, respectively.

Figure 4 shows the changing column density obtained from the integrated RAIR absorption areas of the CO (2142 cm^{-1}), H₂CO (1497 cm^{-1}), and CH₃OH (1030 cm^{-1}) signals during H-atom accretion. The negative CO-consumption peak at 10 K clearly saturates due to the limited penetration depth of H-atoms impacting on the surface. This observation is in agreement with previous findings (Fuchs et al. 2009) indicating that the CO consumption rate exhibits a strong temperature dependence; a lower temperature results in a higher CO conversion rate, fully in line with the proposed CO-H₂CO-CH₃OH reaction findings presented in the literature (Watanabe et al. 2004; Fuchs et al. 2009). This temperature dependence is explained by the different life (i.e. residence) time of H-atoms on the ice surface before desorbing from the ice surface or recombining to H₂ through interactions with other H-atoms. At 10 K and after 120 min of H-atom exposure, the CO depletion behavior starts slowing down, reaching a maximum value at approximately $2\times 10^{15}\text{ molecules}\cdot\text{cm}^{-2}$ (180 min), reflecting that only the few upper layers of CO molecules are accessible to accreting H-atoms.

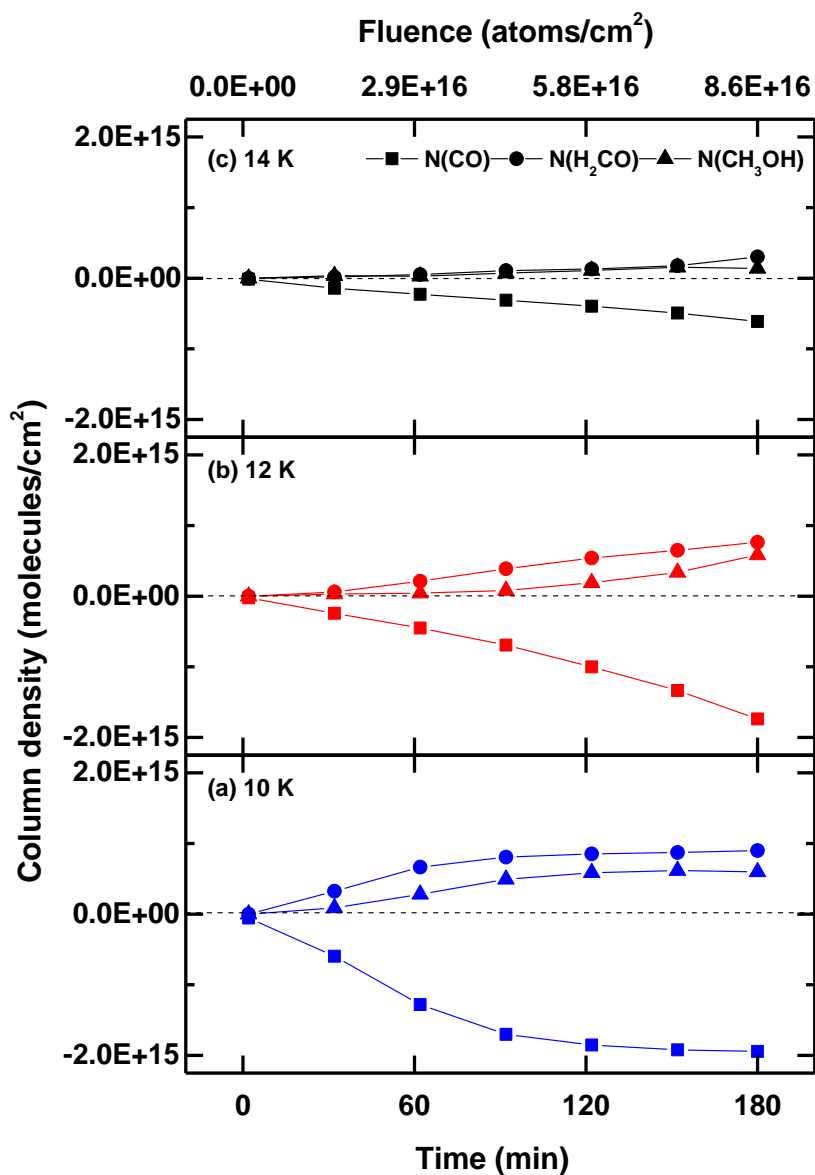


Figure 4. The evolution of the surface column density of CO, H₂CO, and CH₃OH during a CO hydrogenation experiment with an H-atom flux of $\sim 8 \times 10^{12}$ atoms \cdot cm⁻²s⁻¹ over 180 min at (a) 10 K, (b) 12 K, and (c) 14 K, respectively.

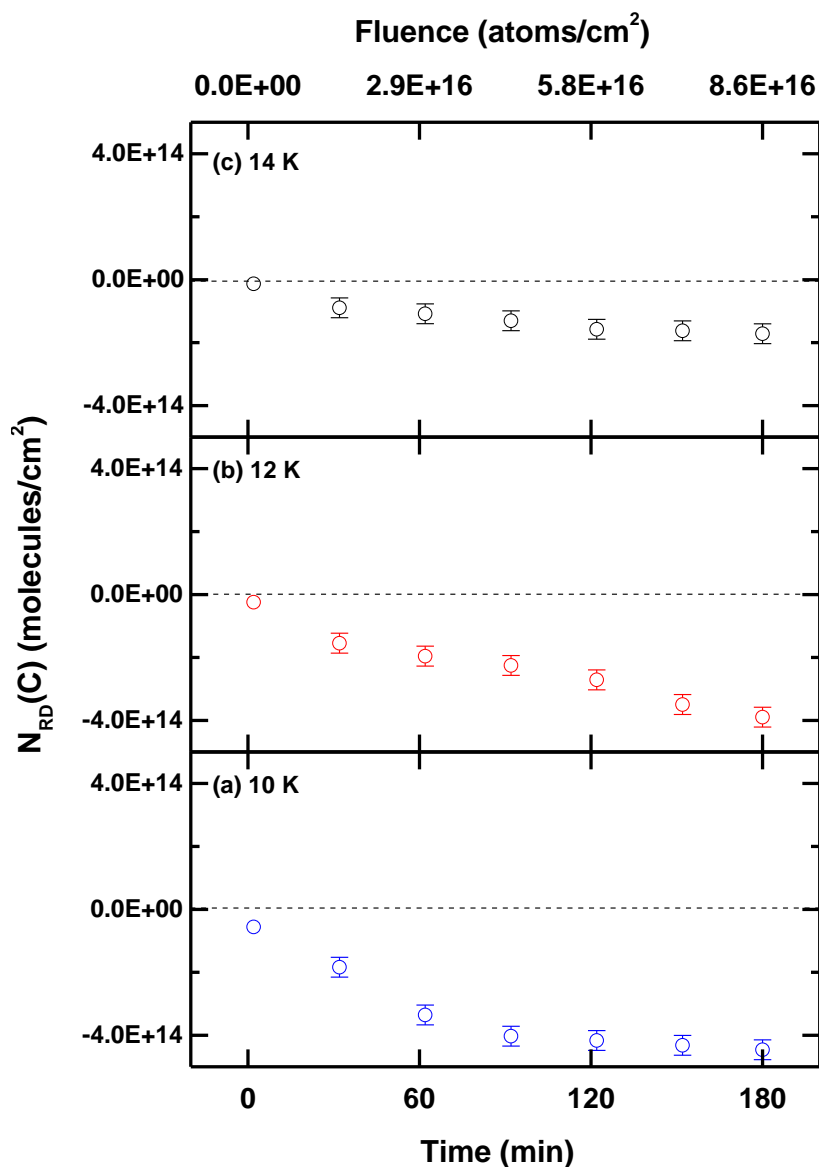


Figure 5. The evolution of the column density of the reactive desorption species, $N_{RD}(C)$, during a CO hydrogenation experiment with an H-atom flux of $\sim 8 \times 10^{12} \text{ atoms} \cdot \text{cm}^{-2} \cdot \text{s}^{-1}$ over 180 min at (a) 10 K, (b) 12 K, and (c) 14 K, respectively.

In Figure 5, the effective RD related column densities are shown for three different ice temperatures for 180 min H-atom bombardment. The increasing loss reflects that the sum of H_2CO and CH_3OH abundances is less than the amount of consumed CO molecules. At 180 min, the $N_{RD}(C)$ value reaches $\sim 4.5 \times 10^{14}$, $\sim 3.9 \times 10^{14}$ and $\sim 1.7 \times 10^{14} \text{ molecules} \cdot \text{cm}^{-2}$ for 10 K, 12 K and 14 K, respectively. As discussed in equation (1), $N_{RD}(C)$ reflects *all* products that are not recorded by the RAIRS technique in the solid-state, i.e., this variable includes both the effect of reactive desorption and other elemental carbon loss channels, such as the formation of COMs with 2, 3 or

more carbon-atoms. However, the expected final abundances of these COMs cannot account for more than a few percent of the consumed CO molecules, as discussed in Fedoseev et al. (2017). Moreover, their CO stretching vibration modes are very similar to the CO stretching vibration modes used here for the CH₃OH column densities calculations; the smaller fraction of carbon atoms conserved in the formed larger COMs is actually already contributing to the CH₃OH column densities. Non-overlapping (and generally weaker) IR absorption features of COMs, i.e., around 860 cm⁻¹ for glycolaldehyde and ethylene glycol are not visible, yielding upper limits of the order of $\sim 5 \times 10^{13}$ molecules · cm⁻²s⁻¹, i.e., <3 % of the consumed CO column density. It should be noted that previous studies on solid state COM formation by Chuang et al. (2015, 2017) and Fedoseev et al. (2017) were performed in co-deposition modus that are assumed to simulate interstellar ice conditions in a more representative way.

In Figure 6 the obtained upper limits of $N_{RD(C)}/N_{consumed(C)}$ ratios are summarized. For two different time periods; from the start to the end (0-180 min) and from 60-180 min. These ratios must be strictly treated as upper limits of the reactive desorption fraction for the performed hydrogenation experiments. The abundances obtained at the end of the CO ice hydrogenation experiments after 180 min of H-atom bombardment are used for the calculations (black column in Figure 6), as here the largest amounts with reaction products are reached. This results in the best peak-to-noise ratios and, therefore, the lowest uncertainties. The resulting $N_{RD(C)}/N_{consumed(C)}$ fractions are 0.23 ± 0.02 , 0.22 ± 0.02 and 0.28 ± 0.05 for 10 K, 12 K and 14 K, respectively. The averaged RD fraction ($N_{RD(C)}/N_{consumed(C)}_{10-14K}$) amounts to 0.24 ± 0.02 for overall 180 min.

A detailed analysis of the data presented in Figure 5 shows that $N_{RD(C)}$ does not decrease linearly with time; roughly during the first 60 min the decrease is faster (for all three investigated temperatures). This may hint at start-up effects, linked to structural weaknesses in the top surface layers because of any loosely bound CO molecules that get off much easier without the need of reactive desorption, for example, upon a collisional impact. For the period from 60 to 180 min, the $N_{RD(C)}$ decrease is much more linear. In order to exclude any perturbing contributions, therefore, the elemental carbon loss fraction is also calculated for the values derived between 60 and 180 min of H-atom bombardment (red column in Figure 6) resulting in somewhat lower upper values: 0.17 ± 0.05 , 0.15 ± 0.02 and 0.16 ± 0.08 for 10 K, 12 K and 14 K, respectively. The resulting averaged value (0.16 ± 0.03) amounts to ~ 70 % of the overall value. The latter value is likely more precise, as it does not include start-up effects. But in order to stay on the safe side, we report here the overall averaged RD fraction.

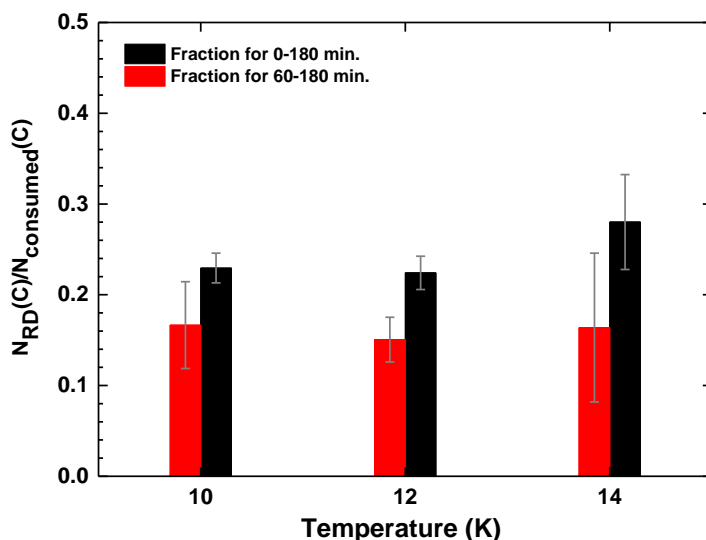


Figure 6. The $N_{RD}(C)/N_{consumed}(C)$ ratios obtained after the bombardment of pre-deposited CO ice with $\sim 8 \times 10^{12}$ H – atoms \cdot cm $^{-2}$ s $^{-1}$ for the full 180 minutes (black column) and the data recorded between 60 and 180 min (red column) at 10, 12 or 14 K, respectively.

3.2 Elemental carbon budget in H₂CO and CH₃OH hydrogenation experiments

A very similar set of experiments is used to investigate reactive desorption upon H₂CO and CH₃OH hydrogenation. Figure 7 (left panel) presents the RAIR difference spectra obtained after hydrogenation of pre-deposited (a) H₂CO and (b) CH₃OH ice for an H-atom flux of $\sim 8 \times 10^{12}$ atoms \cdot cm $^{-2}$ s $^{-1}$ for 180 min at 10 K. In spectrum (a), the negative peaks at 1727, 1497, and 1250 cm $^{-1}$ show the depletion of the initially deposited H₂CO, while the positive peaks indicate the formation of newly formed products, i.e., CO (2142 cm $^{-1}$) and CH₃OH (1030 cm $^{-1}$). The formation of CO through H-atom abstraction reactions from H₂CO and of CH₃OH through H-atom addition reactions to H₂CO have been reported previously by Hidaka et al. (2004) and Chuang et al. (2016).

In contrast, the CH₃OH hydrogenation experiments (Figure 7b) do not result in any efficient consumption of CH₃OH and the formation of hydrogenation products. Only a small negative CH₃OH peak (1050 cm $^{-1}$) is observed, while no secure identification of other carbon bearing species can be made. Moreover, the consumption peak accumulates at the very beginning of the H-atom bombardment and as discussed above, this may be due to other processes than reactive desorption. As in the CO hydrogenation experiments, there is no spectroscopic evidence for radical entrapments due to high reactivity.

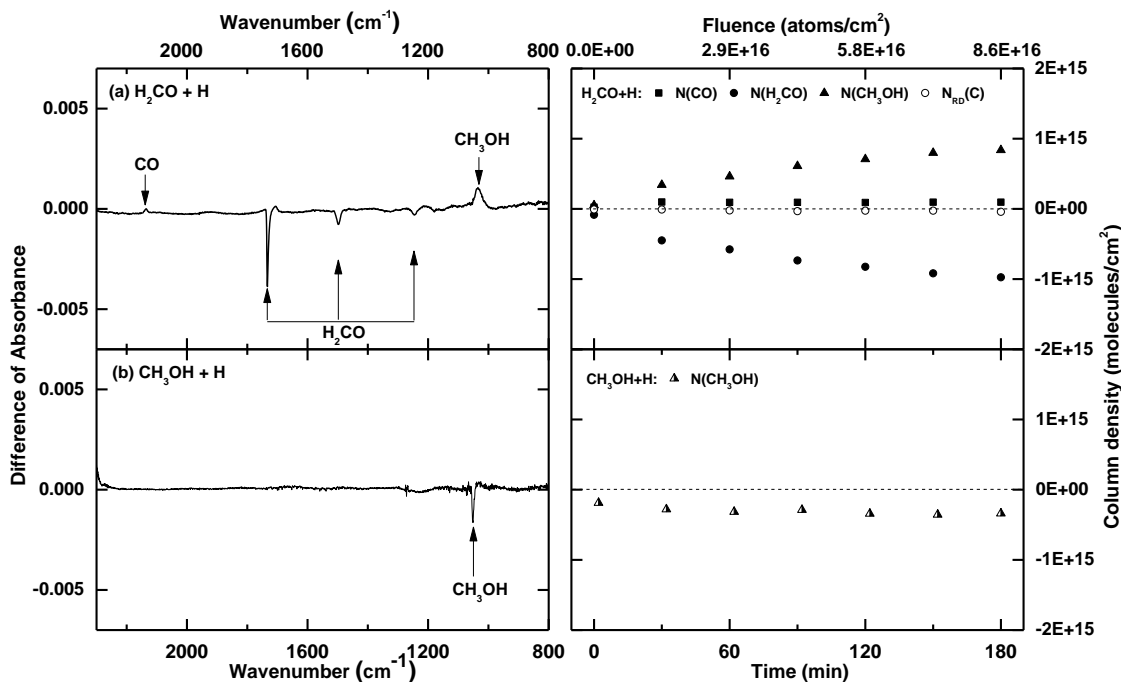


Figure 7. Left: RAIR difference spectra obtained after the exposure of 5×10^{15} molecules cm^{-2} pre-deposited (a) H_2CO and 8×10^{15} molecules cm^{-2} pre-deposited (b) CH_3OH to $\sim 8 \times 10^{12}$ $\text{H} - \text{atoms} \cdot \text{cm}^{-2} \cdot \text{s}^{-1}$ for 180 min at 10 K. Right: The kinetics of the surface column density of CO, H_2CO , and CH_3OH during H_2CO and CH_3OH hydrogenation over 180 min at 10 K.

The evolving column density of CO, H_2CO , and CH_3OH during H-atom exposure with time is shown in the right-hand panel of Figure 7. The consumed $N(\text{H}_2\text{CO})$ saturates at $\sim 1 \times 10^{15}$ molecules $\cdot \text{cm}^{-2}$ and is possibly limited by the H-atom penetration depth as discussed before. However, $N(\text{CH}_3\text{OH})$ saturates immediately, indicating the structural ice changes described before occurring on the surface comprising of CH_3OH molecules. As discussed for the CO hydrogenation experiments, the column density of the missing carbon during the first 60 min of H-atom bombardment likely includes other unwanted desorption processes. Following the same approach as described in Section 3.2, an extended calculation of the H_2CO hydrogenation experiment shows that the sum of the consumption of H_2CO and the production of CO and CH_3OH is very small; $\sim 2.0 \times 10^{13}$ molecules $\cdot \text{cm}^{-2}$, i.e., 4 % with respect to the CO hydrogenation experiments at 10 K. In the CH_3OH hydrogenation experiments, the $N_{RD}(\text{C})$ originates only from the consumption of $N(\text{CH}_3\text{OH})$ due to a poor detection of H_2CO and results in 4 % with respect to the CO hydrogenation experiments at 10 K. These small signals with large uncertainty complicate the determination of reactive desorption fractions. Future work is needed to provide more precise RD efficiencies for H_2CO and CH_3OH .

4. ASTROPHYSICAL IMPLICATION AND CONCLUSIONS:

The astronomical relevance of this work is that the RD fraction used in astrochemical models (for CO hydrogenation, but likely also more general) can be based on an experimentally determined value. The laboratory data obtained here indicate that the sum of consumption and production of carbon-bearing species along the CO hydrogenation in the solid state is not zero. The overall reactive desorption fraction $f_{overall} = N_{RD}(C)/N_{consumed}(C)_{10-14K} \leq 0.24 \pm 0.02$ for CO hydrogenation into CH₃OH is obtained under the assumption that the missing carbon fraction is fully or largely explained by the reactive desorption of hydrogenation products into the gas phase. This value is in the best case an upper limit and the actual effect of reactive desorption may be much smaller, in line with value derived from 60-180 min experiments.

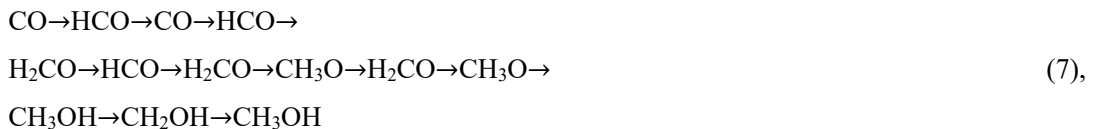
The obtained upper limit is the cumulative effect of all involved reactions. This includes H-atom addition and H-atom abstraction reactions with the stable species (CO, H₂CO and CH₃OH) as well as H-atom reactions with intermediate radicals (HCO, CH₃O, CH₂OH). Although the resulting number is an overall value, it is possible to derive the effective desorption fraction for a single H-atom reaction (f_{single}) by assuming an identical efficiency for each reaction in the hydrogenation process (Garrod et al. 2007). The total fraction of the species left on the surface, i.e., $(1 - f_{overall})$, is equal to the left fraction of a single reaction i.e., $(1 - f_{single})$, to the power n :

$$(1 - f_{single})^n = 1 - f_{overall} \quad (5),$$

where n is the number of reactions in the hydrogenation scheme. f_{single} can be obtained from the $f_{overall}$ measured in our experiments and depends on the average number of H-atom addition and abstraction steps occurring upon conversion of CO into CH₃OH (see Figure 1). For an astrochemical simulation in which only H-atom addition reactions are taken into account, the number of steps resulting in the conversion of CO molecules into CH₃OH is four ($n = 4$):



According to equation (5), this gives f_{single} is $\leq 0.066 \pm 0.005$. This value can be treated as the upper limit for the effective reactive desorption fraction at each of the H-atom addition steps. For a scenario in which also H-atom induced abstraction reactions are considered, this value further decreases. Assuming that both addition and abstraction reactions contributed equally efficient to the overall reactive desorption fraction along the hydrogenation sequence, e.g.:



the effective reactive desorption fraction per reaction of f_{single} is $\leq 0.023 \pm 0.002$, taking $n = 12$. Such effective reactive desorption fractions per hydrogenation reaction can be utilized in the models where both H-atom addition and abstraction steps are incorporated. It should be noted that at the current stage no clear understanding on the number of the involved reaction steps is achieved (Nagaoka et al. 2005; Hidaka et al. 2009; Minissale et al. 2015). In general, a higher number of reaction steps will yield a lower effective reactive desorption fraction per single reaction. Clear data on the relative contribution of the involved H-atom addition and abstraction steps do not exist in the literature and cannot be reliably concluded from the current study.

The effective reactive desorption fractions concluded in both the aforementioned scenarios are well within the 0.01 - 0.10 range of reactive desorption fractions currently used as a free parameter in astrochemical models to explain the transfer of species into the gas phase following their formation in the solid-state (Garrod et al. 2007; Vasyunin et al. 2013; Balucani et al. 2015; Chang & Herbst 2016). These values are clearly much lower than the 40-60 % values reported earlier for other surfaces, but as discussed before, these involved experiments in the sub-monolayer regime. Values of the order of 0.01-0.1, however, are large enough to support a solid state mechanism enhancing abundances of small carbon-bearing species into the gas-phase. This is relevant from two points of view. On the one hand, it provides an intrinsic non-thermal desorption mechanism that may explain the gas-phase observation of molecules that should be fully frozen out for the low temperatures in cold clouds. Garrod et al. (2007) show that models including chemical desorption only induced by association reactions require a reactive desorption fraction of 0.03 to optimally represent the observed gaseous CH₃OH abundances for L134N (Dickens et al. 2000). On the other hand, RD may offer the necessary precursors for proposed gas-phase reaction networks that result in the formation of complex molecules in addition to COMs already formed in the solid-state.

The astronomical take home message is that the experimental results presented here yield an effective reactive desorption fraction for CO hydrogenation that constrain the desorption efficiency previously used in astrochemical simulations to explain astronomical observations toward pre-stellar cores. The derived values give upper limits and the real effect of RD may be smaller.

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