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Letters to the Editor

Modified Bhatia's model for lattice dynamics of some b c.c. metals

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A phenomenological model for lattice dynamics of cubic metals has been proposed by Bhatia (1955) in which the ion-ion interaction was confined to the nearest neighbours only and the electron-ion interaction was calculated using the Thomas-Fermi theory of metals — This model has been fully exploited by Joshi & Hemkar (1961, 1962) and Homkar *et al* (1973). Recently Shukla & Camargo (1974) have pointed out correctly that Bhatia's approach to calculate the Thomas-Fermi screening parameter in terms of the elastic constants of the metal is erroneous, and gives an absurd value for this parameter. This has led Shukla & Salzberg (1973) to modify the electron ion interaction part of the dynamical matrix While Bhatia in his model has considered the electron screening through the Thomas-Fermi theory of metals, Shukla and Salzberg's modification allowed to vary this parameter in the range predicted by the theories of Bohm-Pines and Thomas-Fermi — They have applied this theory in the case of ficient and the copper and have obtamed result which agrees very well with experimental findings

The expression for the volume force obtained by Bhatia (1955) and Shukla & Camargo (1974) corresponds to the ionic centre. Physically it seems more justifiable to take an average value of the force (averaged over the whole Wigner-Seitz sphere) rather than at the ionic centre alone. With this modification we have applied thus theory in the case of block metals sodium, potassium and rubidium in the present study. The interionic interactions have been considered up to second neighbours.

The secular determinant used to determine the angular frequencies ω of the normal modes of vibration in cubic metal can be written as

$$|D - m\omega^2 I| = 0 \qquad \dots (1)$$

where D is the dynamical matrix, m is the ionic mass and I is the unit matrix

of order three. The typical diagonal and nondiagonal elements of the dynamical matrix are given by

$$D_{ii} = \frac{32}{9} (A_1 + 3B_1)(1 - C_i C_j C_k) + 4(A_2 + B_2)S_i^2 + 4B_2(S_j^2 + S_k^2) + \frac{K_e g_i^2 n^{-1}}{1 + (q^2/k_e^2)} G^2(qr_0)$$
(2)

$$D_{ij} = \frac{32}{9} A_1 S_i S_j C_k + \frac{K_e q_i q_j n^{-1}}{1 + (q^2/k_e^2)} G^2(qr_0) \qquad \qquad (2)$$

where $S_t = \sin(q_t a)$, $C_t = \cos(q_t a)$, a is the semi-lattice parameter, q_t the *i*th component of the phonon wave vector, K_e the bulk modulus of the electron gas $k_e = \beta(r_0/a_0)^{\frac{1}{2}}K_F$, where r_0 is the radius of the atomic sphere, a_0 the Bohr radius and K_F the Fermi wave vector, $0.353 \leq \beta \leq 0.814$.

$$G(x) = 3\left(\frac{\sin x - x \cos x}{x^3}\right), \quad x = (qr_0)$$

Each set of A_4 , B_4 represents force constants of the *i*th neighbours $n = \frac{1}{4a^3}$ is

the number of ions per unit volume

The five free parameters have been calculated by making use of the three experimental elastic constants and two zone boundary frequencies $(\nu T_o \text{ and } \nu T_2)$ both transverse in $[\zeta, 0, 0]$ and $[\zeta \zeta \zeta]$ direction respectively). The value of β which gives best result is :353 The elastic constants of sodium were the same as those used by Sharma & Joshi (1963), while for potassium and unbidnim these are due to measurements of Smith & Smith (1965) and Roberts & Meister (1966) The experimental frequencies for sodium, potassium and rubidium were taken from the measurements of Woods *et al* (1962). Cowley *et al* (1966) and Coplay *et al* (1973) respectively The calculated numerical values of the force constants are given below

Constants				Sodium	Potassium	Rubidium
A1	(10 ³ dyn cm ⁻¹)			3.1313	2.42705	1.91336
B ₁	(,,)	-0.13865	-0.22217	-0.18019
A_2	ć)	0.60404	1.30499	0.93694
B,	ć	.,)	0 0355	-0.28013	-0.15413
2aK	e (")	-0.2668	-1.62668	

The calculated and experimental dispersion curves of sodium, potassium and rubidium are presented in figure 1. Figure 1(a) shows that there is excellent agreement between the theoretical and experimental dispersion curves for sodium in all three directions Figures 1(b) and 1(c) show that in the case of potassium and rubidium there is excellent agreement in all branches except one transverse branch in $[\zeta \zeta 0]$ direction. This deviation is attributioable to the existance



Fig. 1b

Fig. 1 Phonon dispersion curves in the three symmetry directions in sodium Fig. 1(a); in potassum Fig. 1(b) and in rubidium Fig. 1(c). The experimental points are shown by the symbol \bigcirc for the longitudinal branches and the symbol \times and \blacktriangle for the transverse branches

of long range interatomic forces and inherent draw back of the Bhatia's model, in which the electron gas does not modify the transverse frequencies



However the results presented here show that modification of Bhatia's model by extending the interionic interaction up to second neighbour with the screening parameter on the lines of Bohm and Pines and averaging the volume force over the Wigner-Seitz sphere could explain the lattice dynamics of b c.c. metals sodium, potassium and rubidnum very successfully. To further justify this approach the work on other dubic metals is in progress.

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REFERENCES

- Bhatia A. B. 1955 Phys. Rev. 97, 379.
- Copley J. R D & Brockhouse B. N. 1973 Canad. J. Phys. 51, 657.
- Cowley R. A , Woods A. D. B. & Dolling G. 1966 Phys. Rev. 150, 487.
- Hemkar M. P. & Joshi S. K 1962 J. Phys. Soc. Japan 17, 754.
- Hemkar M. P., Prakash J. & Chaudra S. 1973 Acta Phys. Polonica A44, 3.
- Joshi S. K & Hemkar M. P 1961 J. Chem. Phys. 34, 1458; 1961 Phys. Rev. 122, 13; 1961 Physica 27, 793; 1962 Phys. Rev. 126, 1687.
- Roberts C. A. & Moister R. 1966 J. Phys. Chem. Solids 27, 1401.
- Sharma P. K. & Joshi S. K. 1963 J. Chem. Phys. 39, 2633.
- Shukla M. M. & Camargo F. P 1974 Oiencia e Cultura 26, 579.
- Shukla M. M. & Salzberg J. B. 1973 Phys. Lett. A 43A, 429.
- Smith P. A. & Smith C. S. 1965 J. Phys. Chem. Solids 26, 279.
- Woods A. D. B., Brockhouse B. N., March R. H., Stewart A. T. & Bower S. R. 1962 Phys.

Rev. 128, 1112.