



Iron(II) and Iron(III) Spin Crossover: Toward an Optimal Density Functional

Siig, Oliver S; Kepp, Kasper P.

Published in:

Journal of Physical Chemistry Part A: Molecules, Spectroscopy, Kinetics, Environment and General Theory

Link to article, DOI:

[10.1021/acs.jpca.8b02027](https://doi.org/10.1021/acs.jpca.8b02027)

Publication date:

2018

Document Version

Peer reviewed version

[Link back to DTU Orbit](#)

Citation (APA):

Siig, O. S., & Kepp, K. P. (2018). Iron(II) and Iron(III) Spin Crossover: Toward an Optimal Density Functional. *Journal of Physical Chemistry Part A: Molecules, Spectroscopy, Kinetics, Environment and General Theory*, 122(16), 4208-4217. <https://doi.org/10.1021/acs.jpca.8b02027>

General rights

Copyright and moral rights for the publications made accessible in the public portal are retained by the authors and/or other copyright owners and it is a condition of accessing publications that users recognise and abide by the legal requirements associated with these rights.

- Users may download and print one copy of any publication from the public portal for the purpose of private study or research.
- You may not further distribute the material or use it for any profit-making activity or commercial gain
- You may freely distribute the URL identifying the publication in the public portal

If you believe that this document breaches copyright please contact us providing details, and we will remove access to the work immediately and investigate your claim.

Fe(II) and Fe(III) Spin Crossover: Towards an Optimal Density Functional

Oliver Sørensen Siig, and Kasper P. Kepp

J. Phys. Chem. A, **Just Accepted Manuscript** • DOI: 10.1021/acs.jpca.8b02027 • Publication Date (Web): 09 Apr 2018Downloaded from <http://pubs.acs.org> on April 10, 2018**Just Accepted**

“Just Accepted” manuscripts have been peer-reviewed and accepted for publication. They are posted online prior to technical editing, formatting for publication and author proofing. The American Chemical Society provides “Just Accepted” as a service to the research community to expedite the dissemination of scientific material as soon as possible after acceptance. “Just Accepted” manuscripts appear in full in PDF format accompanied by an HTML abstract. “Just Accepted” manuscripts have been fully peer reviewed, but should not be considered the official version of record. They are citable by the Digital Object Identifier (DOI®). “Just Accepted” is an optional service offered to authors. Therefore, the “Just Accepted” Web site may not include all articles that will be published in the journal. After a manuscript is technically edited and formatted, it will be removed from the “Just Accepted” Web site and published as an ASAP article. Note that technical editing may introduce minor changes to the manuscript text and/or graphics which could affect content, and all legal disclaimers and ethical guidelines that apply to the journal pertain. ACS cannot be held responsible for errors or consequences arising from the use of information contained in these “Just Accepted” manuscripts.



Fe(II) and Fe(III) Spin Crossover: Towards an Optimal Density Functional

Oliver S. Siig and Kasper P. Kepp*,

Technical University of Denmark, DTU Chemistry, Building 206, 2800 Kgs. Lyngby, DK – Denmark.

* Corresponding Author. Phone: +045 45 25 24 09. E-mail: kpj@kemi.dtu.dk

Abstract

Spin crossover (SCO) plays a major role in biochemistry, catalysis, materials, and emerging technologies such as molecular electronics and sensors, and thus accurate prediction and design of SCO systems is of high priority. However, the main tool for this purpose, density functional theory (DFT), is very sensitive to applied methodology. The most abundant SCO systems are Fe(II) and Fe(III) systems. Even with average good agreement, a functional may be significantly more accurate for Fe(II) or Fe(III) systems, preventing balanced study of SCO candidates of both types. The present work investigates DFT's performance for well-known Fe(II) and Fe(III) SCO complexes, using various design types and customized versions of GGA, hybrid, meta-GGA, meta-hybrid, double-hybrid, and long-range-corrected hybrid functionals. We explore the limits of DFT performance and identify proficient Fe(II)-Fe(III)-balanced functionals. We identify and quantify remarkable differences in the DFT description of Fe(II) and Fe(III) systems. Most functionals become more accurate once Hartree-Fock exchange is adjusted to 10-17%, regardless of the type of functionals involved. However this typically introduces a clear Fe(II)-Fe(III) bias. The most accurate functionals measured by mean absolute errors < 10 kJ/mol are CAMB3LYP-17, B3LYP*, and B97-15 with 15-17% Hartree-Fock exchange, closely followed by

1
2 CAMB3LYP and CAMB3LYP-15, OPBE, rPBE-10, and B3P86-15. While GGA functionals
3
4 display a small Fe(II)-Fe(III) bias, they are generally inaccurate, except the O exchange
5
6 functional. Hybrid functionals (including B2PLYP double hybrids and meta hybrids) tend to
7
8 favor HS too much in Fe(II) vs. Fe(III), which is important in many studies where the oxidation
9
10 state of iron can vary, e.g. rational SCO design and studies of catalytic processes involving iron.
11
12 The only functional with a combined bias < 5 kJ/mol and a decent MAE (15 kJ/mol) is our
13
14 customized PBE0-12 functional. Alternatively one has to sacrifice Fe(II)-Fe(III) balance to use
15
16 the best functionals for each group separately. We also investigated the precision (measured as
17
18 the standard deviation of errors) and show that the target accuracy for iron SCO is 10 kJ/mol for
19
20 accuracy and 5 kJ/mol for precision, and DFT is probably not going to break this limit in the near
21
22 future. Importantly, all four types of functional behavior (accurate/precise, accurate/imprecise,
23
24 inaccurate/precise, inaccurate/imprecise) are observed. More generally, our work illustrates the
25
26 importance not only of overall accuracy but also *balanced* accuracy for systems likely to occur in
27
28 context.
29
30
31
32
33
34
35
36
37
38
39
40
41
42
43
44
45
46
47
48
49
50
51
52
53
54
55
56
57
58
59
60

Introduction.

Spin crossover (SCO) is a fundamental quantum mechanical process occurring in some molecular systems whereby two electronic states with different net spin, high-spin (HS) and low-spin (LS), interconvert upon perturbation, e.g. temperature or pressure¹⁻⁸. SCO requires that the free energy difference between the two electronic spin states is close to zero under the conditions of interest^{2,9,10}:

$$\Delta G_{\text{SCO}} = \Delta H_{\text{SCO}} - T \Delta S_{\text{SCO}} \approx 0 \quad (1)$$

The reaction enthalpy ΔH_{SCO} largely arises from the electronic changes of the first coordination sphere during SCO and typically favors LS, whereas the entropy of the process ΔS_{SCO} largely arises from changes in the geometries during SCO and typically favors HS because of the longer and weaker metal-ligand bonds due to occupation of the e_g orbitals pointing towards the ligands^{9,11}. Many SCO systems have been designed during the past many decades, and some have had their free energy decomposed into ΔS_{SCO} and ΔH_{SCO} contributions^{2,12-14}. Because HS states possess more entropic metal-ligand bonds, the enthalpy-entropy compensation is remarkably strong¹² and largely responsible for the thermally induced transition to HS that can be observed experimentally, as the $T\Delta S$ term begins to favor the HS state².

SCO is a fundamentally important feature of life, as it is required for biological control over the triplet O_2 in the Earth's atmosphere^{15,16}. SCO is also important to various technological applications such as molecular magnets, sensors, molecular electronics and transition-metal-based catalysis^{3,17-21}. Iron is the most common central metal ion in current SCO systems. The balance between the central metal ion and the ligand field together enables SCO as both central ion and ligand has systematic spin state preferences²². The spectrochemical series²³⁻²⁵ gives information about this via the (vertical) energy difference between the orbital levels involved in

1
2 SCO (in O_h symmetry the splitting Δ_o of the e_g and t_{2g} levels). The more relevant series for
3
4 rationalizing and predicting SCO is a thermochemical series of spin state propensities²⁶, which
5
6 corrects the spectrochemical series by accounting for the substantial contributions from
7
8 geometric relaxation, spin pairing, and entropy and vibrational effects^{12,26}. From such ligand-
9
10 metal considerations it can be seen that iron with nitrogen donor ligands is a hotspot for SCO²⁶.
11
12 6-coordinate iron(II) complexes with nitrogen donor ligands are archetypal⁵, with heme being
13
14 the natural reference benchmark of this type whose SCO plays a fundamental role in the oxygen
15
16 management of life²⁷.
17
18
19
20

21
22 Density functional theory (DFT) is well suited to study SCO systems, as the electron
23
24 correlation can be described accurately at a relatively modest computational cost^{28,29}. However,
25
26 the precise relative energy of HS and LS states is hard to obtain, because this requires a balanced
27
28 treatment of Fermi and Coulomb correlation⁹. Different density functionals produce very
29
30 different HS-LS energy gaps as has been discussed in detail^{9,12,14,26,30-41}. Because the exchange
31
32 integrals of the Hartree-Fock (HF) treatment only account for Fermi correlation, they explicitly
33
34 favor HS. Accordingly, the amount of HF exchange included in a hybrid functional dominates
35
36 the HS-LS gap, and accordingly this gap increases linearly with the HF exchange^{30,31,42,43}. Also,
37
38 the inclusion of gradient terms in meta functionals has been found to affect the HS-LS gap^{26,44}.
39
40
41
42

43
44 The major concern in DFT which also specifically relates to SCO is the predictive value
45
46 of a given functional once applied outside its parameterization range, i.e. “universality”⁴⁵⁻⁴⁷. We
47
48 have previously observed¹² that the ferric Fe(III) and ferrous Fe(II) systems are not described
49
50 equally well by a given functional, and this effect is very significant vs. the noise in the
51
52 methodology. From the SCOFE30 database¹², it can be seen that Fe(III) SCO systems tend to
53
54 have experimental ΔH_{SCO} about 10 kJ/mol smaller than Fe(II) systems, and correspondingly also,
55
56 due to the strong enthalpy-entropy compensation¹², the experimental ΔS_{SCO} is smaller by perhaps
57
58
59
60

1
2 30 J/molK. Because of the diverse ligands of this data set, these effects are averages with large
3
4 variations depending on the exact ligand field, but the trend is clear. The differential ligand field
5
6 stabilization energy and spin pairing effects are the same (10/5 and a change of $2P$). This
7
8 explains why the difference is relatively subtle.
9

10
11
12 Unfortunately, these differences pose a challenge to DFT: Thus, B3LYP*-D3 performs
13
14 best in a test once all physical effects (zero point energy, vibrational entropy, relativistic
15
16 corrections, dispersion) are included before comparison to experimental ΔH_{SCO} , but closer
17
18 inspection shows that this comes at a price of producing too much high-spin in Fe(II) systems
19
20 and too much LS in Fe(III) systems¹². The failure arises mainly for the hybrid functionals, as
21
22 they tend to not only favor HS, *but favor HS too much in Fe(II) compared to Fe(III)*. In the
23
24 following we refer to this as the “Fe(II)-Fe(III) bias”. This bias, upon reinspection of previous
25
26 results, easily passes 20 kJ/mol¹². This bias has so far been overlooked and has not previously, to
27
28 our knowledge, been described, although Friesner et al. have studied d-configuration-dependent
29
30 energies with DFT and addressed some related challenges⁴⁸.
31
32
33
34
35

36 Our computations in this paper show that, in the search for a proficient density functional
37
38 description of spin states of iron of major importance in many catalytic processes, the Fe(II)-
39
40 Fe(III) bias produces an important obstacle. Thus, we decided to explore this bias, and
41
42 investigate how far we can get with modern DFT towards optimal SCO performance, while
43
44 considering both accuracy, precision, and the Fe(II)-Fe(III) bias. In this search, we study a range
45
46 of different classes of density functionals, including GGAs, hybrids, meta GGAs, meta hybrids,
47
48 dopable hybrids, and range-separated hybrids, and also investigate customized versions to
49
50 identify the limits of accuracy and precision when applying DFT to the study of iron spin states.
51
52
53
54

55 **Methods.**

56
57
58
59
60

All computations were performed using the Turbomole software, version 7.0⁴⁹. The electron densities and energies were converged to 10^{-6} a.u., and the resolution of identify approximation was used to speed up the calculations^{50,51}.

In order to analyze Fe(II) and Fe(III) systems fairly, we used a balanced data set shown in **Figure 1**, consisting of five Fe(II) SCO systems and five Fe(III) SCO systems. The five Fe(II) systems are **1**: $[\text{Fe}(\text{paph})_2]^{2+}$ (paph = bis(2-(2-pyridylamino)-4-(2-pyridyl)thiazole)⁵²); **2**: $[\text{Fe}(\text{tacn})_2]^{2+}$ (tacn = 1,4,7-Triazacyclononane⁵³); **3**: $[\text{Fe}(\text{pyimH})_3]^{2+}$ (pyimH = 2-(2'-pyridyl)imidazole⁵³); **4**: $[\text{Fe}(\text{tpchxn})]^{2+}$ (tpchxn = N,N,N',N'-tetrakis(2-pyridylmethyl)-1*R*,2*R*-cyclohexanediamine⁵⁴); and the classic **5**: $[\text{Fe}(\text{phen})_2(\text{NCS})_2]$ (phen = 1,10-phenanthroline⁵⁵). The five Fe(III) systems studied are: **6**: $[\text{Fe}(\text{acac})_2\text{trien}]^+$ (acac = acetyl-acetonate-triethylenetetramine⁵⁶); **7**: $[\text{Fe}(\text{bzac})_2\text{trien}]^+$ (bzac = benzoyl-acetonate-triethylenetetramine⁵⁷); **8**: $[\text{Fe}(\text{bzacCl})_2\text{trien}]^+$,⁵⁷ **9**: $[\text{Fe}(\text{tfac})_2\text{trien}]^+$ (trifluoroacetyl-acetonate-triethylenetetramine⁵⁷); and **10**: $[\text{Fe}(\text{3-MeO-salenEt})_2]^+$ (3-MeO-salenEt = 3-methoxysalicylaldehyde-*N*-ethylethylenediamine⁵⁸). The geometries were optimized as previously described¹² using the BP86^{59,60} functional known to give accurate geometries for transition metal systems and the def2-SVP basis set⁶¹ including the Cosmo solvation model^{62,63}. The Cosmo radii for all atoms were the optimized default values of Turbomole, and 2.0 Å was used for iron. This protocol routinely provides excellent geometries for first row transition metal complexes with typical errors in metal-ligand bond lengths of 0.02–0.03 Å⁶⁴.

The entropy, in particular the vibrational entropy, plays an important role in determining the spin crossover tendency, which is given by the free energy in Equation (1). This entropy is larger for the HS state with the longer and weaker metal-ligand bonds. The entropy can be modeled using the molecular vibration state function of the low-spin and high-spin states after calculating the vibrational frequencies in both states. Thus, it is not very sensitive to the choice

1
2 of functional and is well modeled e.g. by a GGA functional such as BP86¹². The main challenge
3
4 in modeling SCO systems with DFT thus lies in the choice of functional used to computed
5
6 ΔH_{SCO} , and we focus our benchmark on the ability of DFT to reproduce this observable.
7
8

9
10 For assessing the energy difference between HS and LS states, energies for all functionals
11
12 were converged using the fully polarized def2-TZVPP basis set⁶¹; this basis set performed
13
14 accurately in previous benchmarks against experimental SCO enthalpies^{12,33}. In the present work,
15
16 a range of functionals were studied to investigate the Fe(II)-Fe(III) bias. Since some functionals
17
18 were already studied in a previous benchmark¹², these were excluded from the present study
19
20 except the three best functionals^{12,31,33}: the double-hybrid B2PLYP which includes both MP2
21
22 correlation energy and HF exchange corrections to the correlation and exchange functionals⁶⁵,
23
24 the meta hybrid TPSSh^{66,67} with 10% HF exchange, and B3LYP*³⁰ which is a 15% version of
25
26 B3LYP⁶⁸⁻⁷⁰, the latter two in their D3 corrected versions⁷¹. The TPSS⁶⁶ functional as the direct
27
28 non-hybrid counterpart of TPSSh was also included for strict comparison.
29
30
31
32

33
34 In addition, we studied the long range-corrected hybrid functional CAMB3LYP⁷², which
35
36 separates the exchange interaction into long- and short-range parts; M06 and M06-2X, which are
37
38 meta hybrids with 27% and 54% HF exchange, respectively⁷³; the local M06L functional⁷⁴; the
39
40 KT1 and KT2 functionals by Keal and Tozer⁷⁵, which are GGA type functionals developed
41
42 specifically for good performance for NMR parameters (KT1 obeys the uniform gas constraint;
43
44 KT2 is fitted); PW91-PW91⁷⁶; the two main revised versions of PBE⁷⁷, revPBE⁷⁸ and rPBE⁷⁹;
45
46 OPBE and OLYP using Handy and Cohen's optimized exchange functional⁸⁰ with the PBE⁷⁷ or
47
48 LYP⁶⁹ correlation functionals; B-VWN^{59,81}; B3P86^{60,68}; and PBEH-3C⁸².
49
50
51
52
53
54
55
56
57
58
59
60

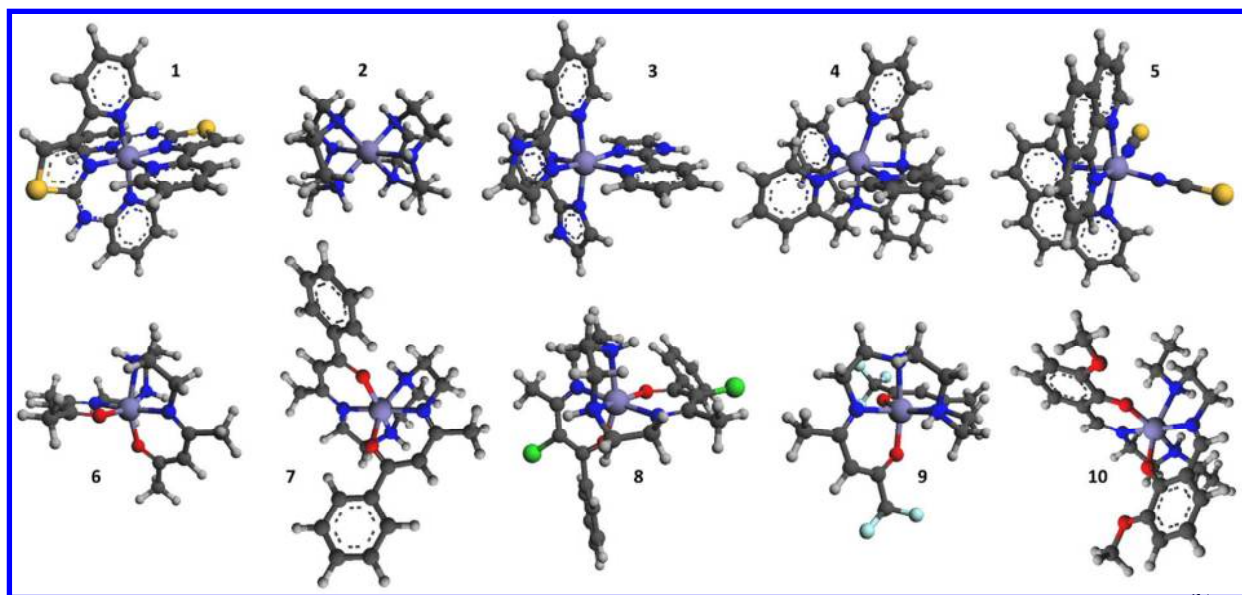


Figure 1. SCO complexes studied in this work: Fe(II) SCO systems: **1**: $[\text{Fe}(\text{paph})_2]^{2+}$, **2**: $[\text{Fe}(\text{tacn})_2]^{2+}$, **3**: $[\text{Fe}(\text{pyimH})_3]^{2+}$, **4**: $[\text{Fe}(\text{tpchxn})]^{2+}$, and **5**: $[\text{Fe}(\text{phen})_2(\text{NCS})_2]$. Iron(III) SCO systems: **6**: $[\text{Fe}(\text{acac})_2\text{trien}]^+$, **7**: $[\text{Fe}(\text{bzac})_2\text{trien}]^+$, **8**: $[\text{Fe}(\text{bzacCl})_2\text{trien}]^+$, **9**: $[\text{Fe}(\text{tfac})_2\text{trien}]^+$, and **10**: $[\text{Fe}(\text{3-MeO-salenEt})_2]$.

Furthermore, we used the xcfun library implemented in Turbomole to develop customized functionals⁸³ that we studied systematically as well: OPBE-15 (with 15% HF exchange); rPBE-10 (rPBE customized as a 10% hybrid); B97-15 (B97-D customized as a 15% hybrid); CAMB3LYP-15 and CAMB3LYP-17 which have reduced HF exchange relative to the native functional (which has 19% HF exchange); and customized versions of PBE0⁷⁷ with 15, 12, and 10% HF exchange (PBE0-15, PBE0-12, PBE0-10). The goal was to use a wide range of different types of functionals (GGA, meta, GGA hybrid, meta hybrid, double-hybrid), and then optimize HF exchange toward best performance. Most of these functionals have not been studied before in the context of SCO.

1
2 Dispersion interactions have been shown to affect the SCO equilibrium by selectively
3 favoring the more compact LS state⁸⁴. Accordingly all energy calculations included dispersion
4 corrections using Grimme's D3 method⁷¹. Some of the methods already include dispersion
5 effects (e.g. the MP2-corrected double hybrid B2PLYP and B97-D) and these were evaluated
6 based on their own dispersion corrections. Methods that do not include any correlation or
7 empirical dispersion correction were corrected by their parameterized version of D-3 (each
8 functional has a specific D3 set of parameters). Some functionals, including all the customized
9 ones, do not have parameterized dispersion corrections and for these, we used an average
10 correction based on previous work (averaged over BHLYP, PBE0, B3LYP, PW6B95, B3LYP*
11 (using the same correction as B3LYP), TPSSh, TPSS, BLYP, PBE, and BP86. These corrections
12 were in kJ/mol 8.3 (3.8) for **1**, 3.8 (2.2) for **2**, 12.8 (6.1) for **3**, 12.3 (5.7) for **4**, 5.2 (2.5) for **5**, 0.2
13 (0.5) for **6**, 1.4 (0.9) for **7**, 9.9 (4.6) for **8**, 0.4 (0.5) for **9**, and 6.2 (2.7) for **10**. The numbers in
14 parenthesis are standard deviations calculated from the corrections for the different functionals.
15 Because these are *differential* corrections for HS and LS states, they have relatively small
16 standard deviations. Thus, for SCO using any reasonable dispersion D3 correction is acceptable
17 within an expected uncertainty of 3 kJ/mol (the average standard deviation), which is
18 substantially below the target accuracy of 10 kJ/mol. Notice also the generally smaller values for
19 iron(III) systems.

20
21 The thermodynamic and zero-point energy corrections to the energies obtained from
22 numerical frequency calculations in Cosmo were included subsequently to the energy
23 calculation, using the corrections for each individual system as previously described¹². These
24 corrections are important to correct for before assessing the ability of a functional in predicting
25 the experimental enthalpy of SCO as the experimental numbers includes these effects.

1
2 All the electronic energies are listed in Supporting Information, **Table S1** for Fe(II)
3 systems and **Table S2** for Fe(III) systems. **Table S3** and **Table S4** show the corresponding HS-
4 LS energy gaps for Fe(II) and Fe(III) systems, respectively, corrected for ZPE, relativistic
5 effects, and dispersion. **Table S5** shows the experimental data, with references and error bars, as
6 well as the computed corrections to the direct electronic energy. The experimental errors are
7 within 1 kJ/mol for all ten systems studied here; thus the performance of the methods that we
8 identify is not dependent on uncertainties in the reported experimental numbers. Errors vs.
9 experimental data are summarized in **Tables S6-S9**, and the Fe(II)-Fe(III) bias is tabulated in
10 **Table S10**. **Table S11** shows the optimized xyz coordinates of all systems in both HS and LS
11 states.
12
13
14
15
16
17
18
19
20
21
22
23
24
25
26
27
28
29
30
31

32 **Results and Discussion.**

33
34

35 **General Accuracy of DFT for iron SCO.** **Figure 2** shows ΔH_{SCO} for the five Fe(II)
36 systems computed with the various functionals, and **Figure 3** shows ΔH_{SCO} for the Fe(III)
37 systems. As expected from previous work^{12,31-33}, the functionals perform quite distinctly, and
38 many favor either the HS or LS state by a large margin. As also expected, the amount of HF-
39 exchange included significantly impacts the results. Taking CAMB3LYP as an example, the
40 calculated spin gap decreases by approximate 9 kJ/mol for each 2% HF-exchange included in the
41 calculations.
42
43
44
45
46
47
48
49
50
51

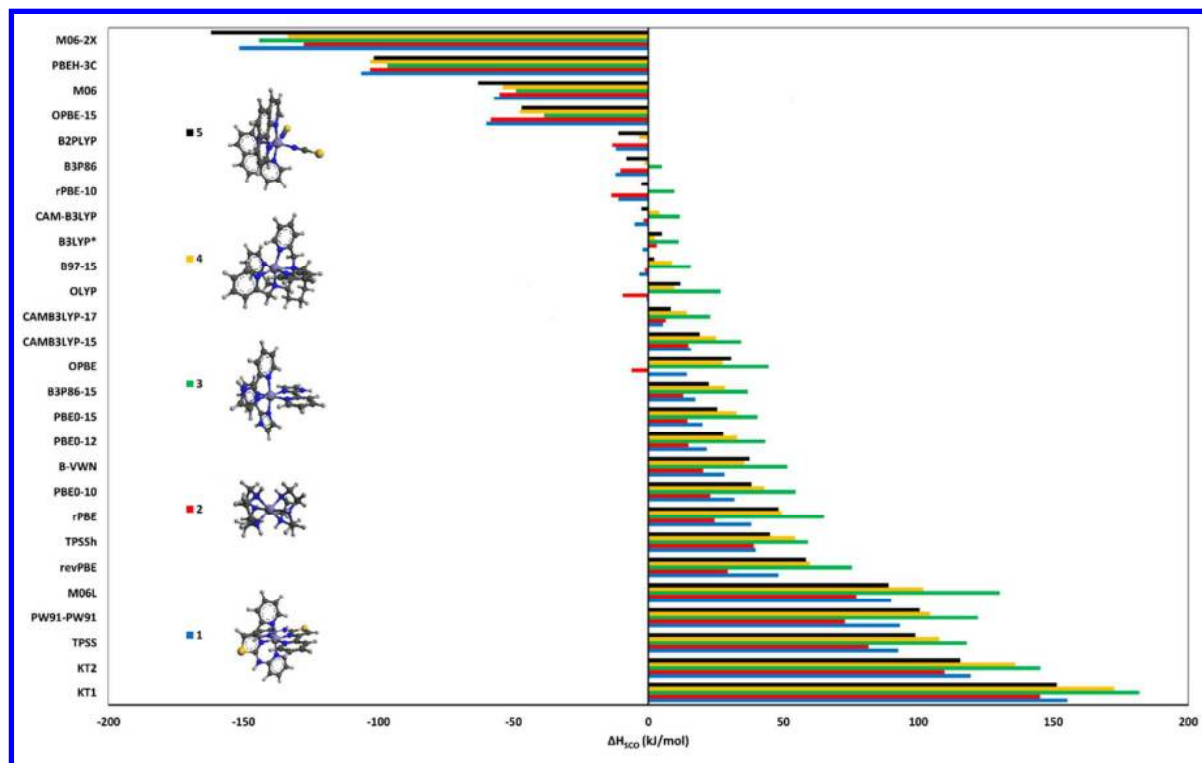
52 Because many of these functionals were customized to achieve good accuracy, many
53 perform reasonably well compared to what one would see in a test of “random” functionals, i.e.
54 many functionals in the center of **Figure 2** and **Figure 3** produce ΔH_{SCO} that are close to SCO
55
56
57
58
59
60

(i.e. slightly positive). However, there are also examples of extremely LS-biased non-hybrid functionals (KT1, KT2, TPSS, M06L, PW91-PW91), and functionals that are extremely HS-biased (M06-2X, M06, PBEH-3C). These functionals are importantly the same for both the Fe(II) and the Fe(III) systems, and errors can exceed 100 kJ/mol.

Figure 4 shows the means signed error (MSE, in black) in the computed ΔH_{SCO} vs. experimental data for all ten systems, once the electronic energies of the HS and LS states have been corrected for zero-point energy, dispersion effects, and relativistic contributions; the zero-point energy generally favors the HS state with the longer and weaker metal-ligand bonds, whereas the relativistic and dispersion effects favor the more compact LS state; these effects have been described in detail previously¹². The errors are shown in **Table S10**. The most accurate functionals in terms of systematic HS-LS balance, measured as the MSE for the full data set, are OPBE, the customized CAMB3LYP-15 and CAMB3LYP-17, B97-15 and B3LYP*. Other accurate functionals are CAMB3LYP in its normal form. The most accurate functionals measured by MAE < 10 kJ/mol are CAMB3LYP-17, B3LYP*, and B97-15 with 15-17% HF exchange, closely followed by CAMB3LYP and CAMB3LYP-15, OPBE, rPBE-10, and B3P86-15 with MAEs < 15 kJ/mol. These results show that customized HF exchange fractions remedy the spin balanced for a wide range of functionals. It is interesting to see that the GGA functional OPBE performs so well across the data series without any use of HF exchange; thus the O exchange functional has an effect that, from the spin state point of view, corresponds to the effect of ~15% HF exchange. This conclusion supports previous findings by Swart^{32,85} who used OPBE as a basis for his functionals. It is also encouraging to see the excellent performance of CAMB3LYP even in its normal 19% HF exchange form, but in particular in the customized versions with slightly less HF exchange. If one is not able to use customized versions of functionals, functionals such as OPBE, B3LYP*, and CAMB3LYP are among the most accurate

1
2 for iron SCO; B3LYP* was developed for this purpose⁴² and found in the most elaborate
3
4 benchmark so far to perform well¹².
5
6

7 The observation that rPBE requires only 10% HF exchange to reach its maximal accuracy
8 is interesting in the context that rPBE was developed to reduce the overbinding tendency of
9 PBE⁷⁹, and we have argued previously⁹ that the overbinding tendency (measured as too strong
10 metal-ligand bonds) and LS bias come together, partly because the HF exchange works to both
11 weaken bonds (by favoring the open-shell dissociated states with more exchange integrals) and
12 favor HS states (which also have more exchange integrals). The behavior of rPBE provides a
13 relevant example of this relationship.
14
15
16
17
18
19
20
21
22
23
24
25
26



27
28
29
30
31
32
33
34
35
36
37
38
39
40
41
42
43
44
45
46
47
48
49
50
51
52
53 **Figure 2.** Enthalpies of SCO in iron(II) systems (ΔH_{SCO} , in kJ/mol), computed as the energy
54 difference between HS and LS states, including differential zero-point energy, relativistic, and
55
56
57

thermal energy corrections, and D3 from B3LYP-D3 if dispersion is not already included in the method.

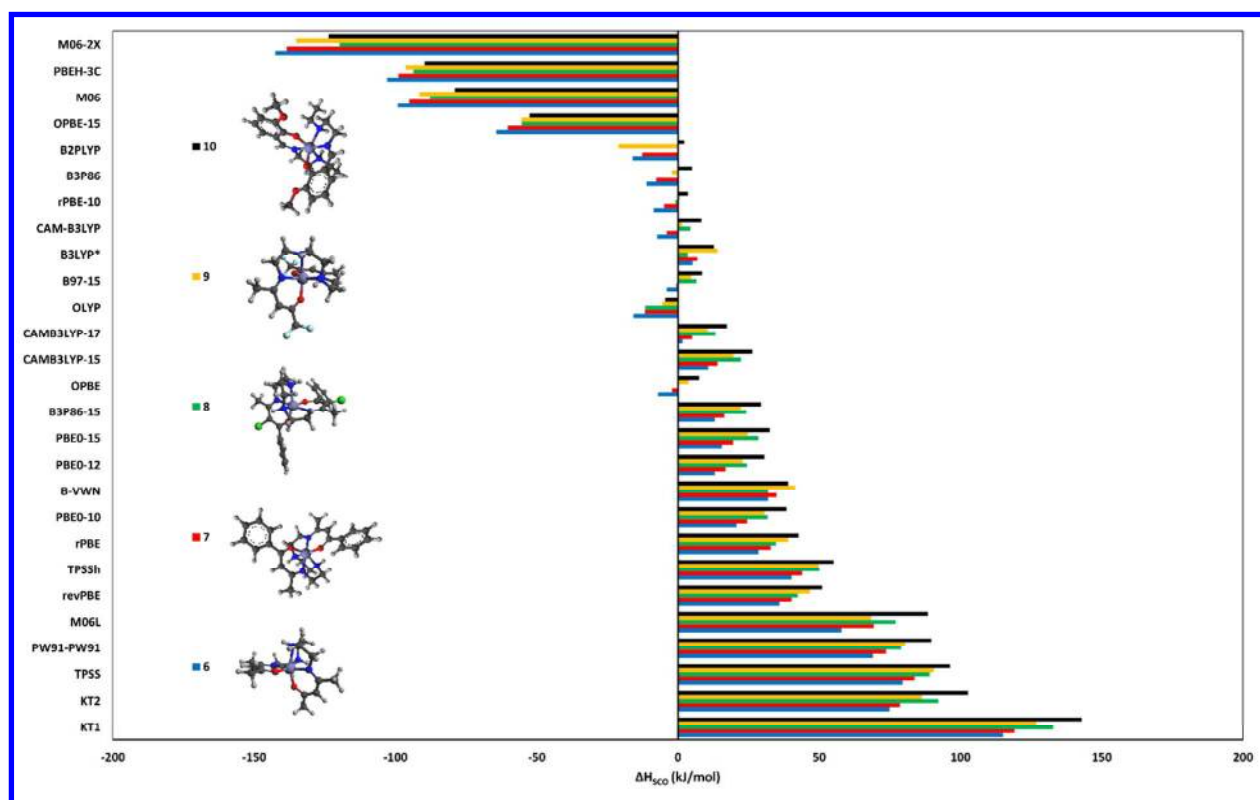


Figure 3. Enthalpies of SCO in iron(III) systems (ΔH_{SCO} , in kJ/mol), computed as the energy difference between HS and LS states, including differential zero-point energy, relativistic, and thermal energy corrections, and D3 from B3LYP-D3 if dispersion is not already included in the method.

Fe(II)-Fe(III) Bias. In addition to the MSE of each functional, **Figure 4** also shows the Fe(II)-Fe(III) bias, calculated as the MSE of the Fe(II) systems minus the MSE of the Fe(III) systems (red line). This property estimates the balance (transferability) of DFT, which should be a central focus in the search for universal functionals: It is not enough to show a small total MSE, or small total MAE, errors also need to be evenly distributed between the important

1 categories of systems likely to be of interest. It is notable from **Figure 4** that the Fe(II)-Fe(III)
2 bias can reach 20 kJ/mol for these systems and easily surpasses 10 kJ/mol in many of the good
3 functionals. Upon inspection of the data, such a tendency is also evident from the larger
4 functionals. Upon inspection of the data, such a tendency is also evident from the larger
5 functionals. Upon inspection of the data, such a tendency is also evident from the larger
6 functionals. Upon inspection of the data, such a tendency is also evident from the larger
7 functionals. Upon inspection of the data, such a tendency is also evident from the larger
8 functionals. Upon inspection of the data, such a tendency is also evident from the larger
9 functionals. Upon inspection of the data, such a tendency is also evident from the larger
10 functionals. Upon inspection of the data, such a tendency is also evident from the larger
11 functionals. Upon inspection of the data, such a tendency is also evident from the larger
12 functionals. Upon inspection of the data, such a tendency is also evident from the larger
13 functionals. Upon inspection of the data, such a tendency is also evident from the larger
14 functionals. Upon inspection of the data, such a tendency is also evident from the larger
15 functionals. Upon inspection of the data, such a tendency is also evident from the larger
16 functionals. Upon inspection of the data, such a tendency is also evident from the larger
17 functionals. Upon inspection of the data, such a tendency is also evident from the larger
18 functionals. Upon inspection of the data, such a tendency is also evident from the larger
19 functionals. Upon inspection of the data, such a tendency is also evident from the larger
20 functionals. Upon inspection of the data, such a tendency is also evident from the larger
21 functionals. Upon inspection of the data, such a tendency is also evident from the larger
22 functionals. Upon inspection of the data, such a tendency is also evident from the larger
23 functionals. Upon inspection of the data, such a tendency is also evident from the larger
24 functionals. Upon inspection of the data, such a tendency is also evident from the larger
25 functionals. Upon inspection of the data, such a tendency is also evident from the larger
26 functionals. Upon inspection of the data, such a tendency is also evident from the larger
27 functionals. Upon inspection of the data, such a tendency is also evident from the larger
28 functionals. Upon inspection of the data, such a tendency is also evident from the larger
29 functionals. Upon inspection of the data, such a tendency is also evident from the larger
30 functionals. Upon inspection of the data, such a tendency is also evident from the larger
31 functionals. Upon inspection of the data, such a tendency is also evident from the larger
32 functionals. Upon inspection of the data, such a tendency is also evident from the larger
33 functionals. Upon inspection of the data, such a tendency is also evident from the larger
34 functionals. Upon inspection of the data, such a tendency is also evident from the larger

This general tendency confirms that the effect is real. Unfortunately, most of the functionals that produce a low Fe(II)-Fe(III) bias perform relatively poorly for the overall ΔH_{SCO} . The six functionals that have a numerical Fe(II)-Fe(III) bias < 5 kJ/mol are OPBE-15, PBE0-12, PBE0-10, rPBE, revPBE, and TPSS. Of these, only PBE0-12 has an acceptable MAE of ~ 15 kJ/mol, the remaining five functionals having MAEs of 22-82 kJ/mol.

35 The most important example is the B3LYP* functional³⁰ which is much studied and
36 probably a first pick for many applications of iron chemistry. Its Fe(II)-Fe(III) bias is -15 kJ/mol
37 for this data set, and a similar large bias is evident upon reinspection of previous work¹², i.e. this
38 is a general feature. Since many applications involve the study of Fe(II) and Fe(III) in
39 combination, either in separate systems or in actual redox processes (and even when the formal
40 oxidation state does not change but back bonding is involved in Fe(II), e.g. heme systems) this
41 bias will cause an imbalanced treatment of Fe(II) and Fe(III) states by DFT. An imbalance of 15
42 kJ/mol in such a process is not a small error, but has so far never been investigated (until this
43 work) and is a typical example of the importance of balanced performance vs. average
44 performance.

1
2 It is interesting to note that many functionals that have a low Fe(II)-Fe(III) bias include
3
4 the PBE correlation functional, which is largely nonempirical⁷⁷; this could support the notion that
5
6 more non-empirical functionals are more transferable (or, in some terminology, more
7
8 “universal”), i.e. it may have exact bounds that partly remedy the Fe(II)-Fe(III) bias.
9

10
11
12 Considering this, the number of good functionals is reduced substantially. As mentioned,
13
14 B3LYP* is no longer very suitable if the oxidation state can vary at the iron center. We also find
15
16 that the CAMB3LYP range-corrected hybrid functional produces a bias of -10 kJ/mol, whereas
17
18 the non-hybrid GGA OPBE, on the other hand, has a Fe(II)-Fe(III) bias of +11 kJ/mol. One
19
20 possible solution could be to predict SCO behavior using a combination of the proposed
21
22 functionals; both OPBE and CAMB3LYP-17 performed exceptionally well on the entire data set,
23
24 and the individual biases of the two functionals cancel to produce results that compare well with
25
26 both the overall data set, and the Fe(II) and Fe(III) species separately. However, for normal
27
28 studies in iron-based catalysis and chemistry, users will have to pick a functional either
29
30 considering minimal bias with decent accuracy (PBE0-12 being the recommended functional
31
32 with a bias of 5 kJ/mol and MAE of 15 kJ/mol), or, if only one oxidation state of iron is
33
34 consistently studied, the best functional for the oxidation state of interest, e.g. B3LYP*,
35
36 CAMB3LYP, or OPBE for Fe(III) and CAMB3LYP-15, CAMB3LYP-17, or B3P86-15 for
37
38 Fe(II).
39
40
41
42
43
44

45 **System Dependencies and Pathologies.** To investigate whether any of the conclusions
46
47 above are sensitive to the choice of data set, i.e. if there are pathological cases among the
48
49 systems that could give rise to erroneous conclusions, we divided errors into system as shown in
50
51 a radar plot in **Figure 5A**. The blue range represents the MAE for each of the systems, using data
52
53 for all functionals in the study, whereas the red range represents the system-specific MAEs for
54
55 the most accurate functionals (those with MAE < 15 kJ/mol for the full data set). We recommend
56
57
58
59
60

using radar plots of this type to identify unsuitable (pathological) systems in a benchmark data set: A healthy data set will have an almost circular form of both curves, as is indeed seen in **Figure 5A**; thus, our data set does not include pathological systems.

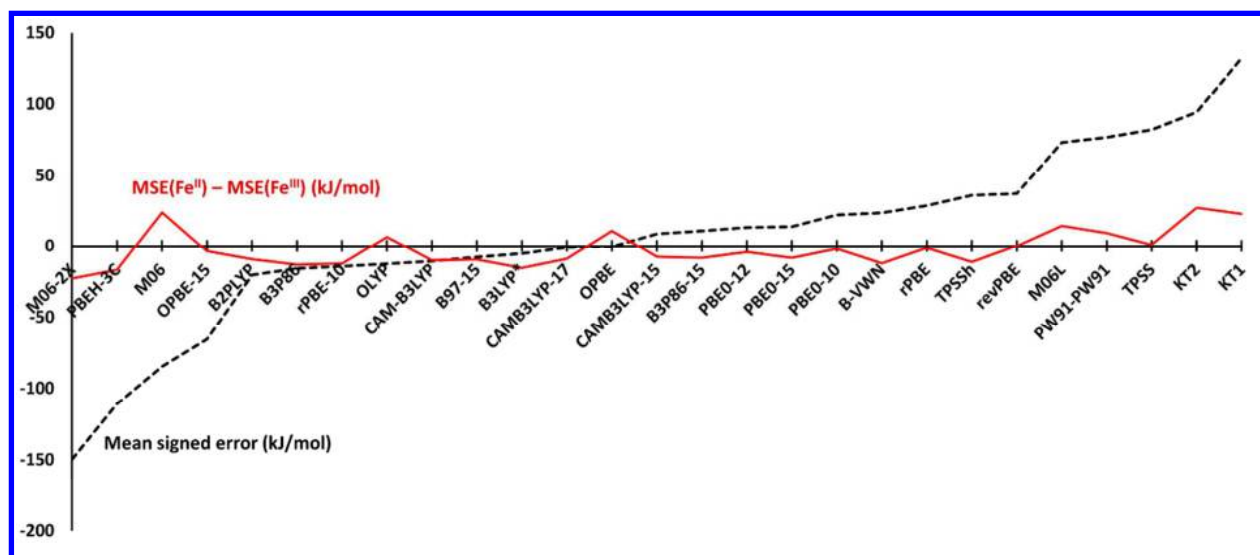
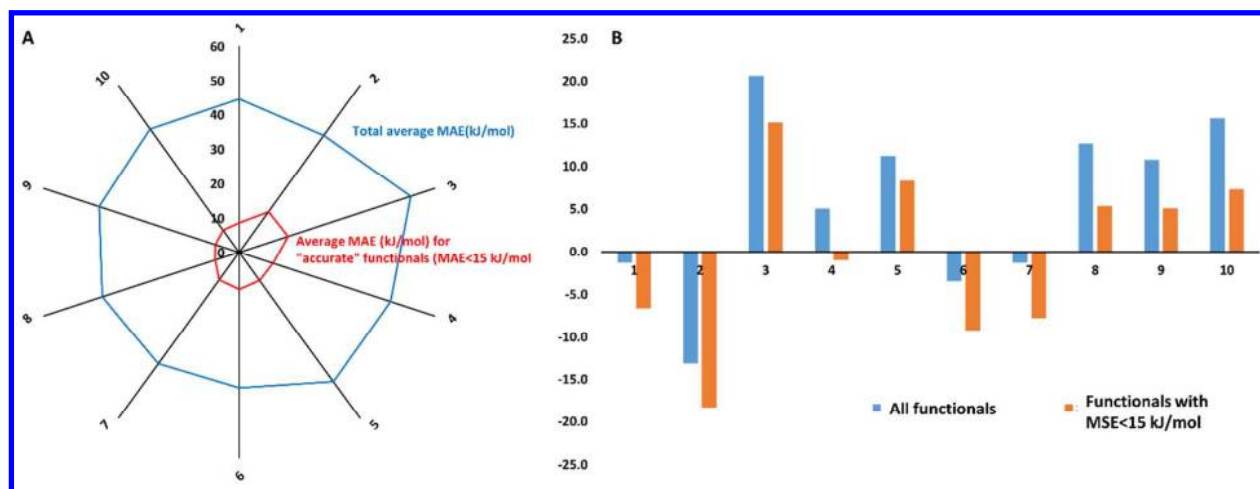


Figure 4. Mean signed error (kJ/mol, black) and the Fe(II)-Fe(III) bias, calculated as the difference in mean signed error for the Fe(II) and Fe(III) systems (kJ/mol, red).

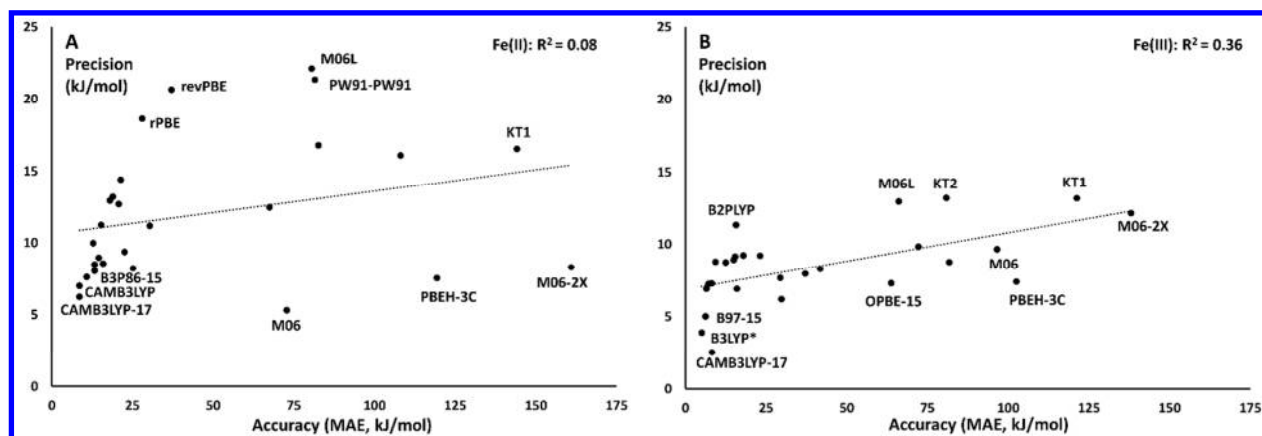


1
2 **Figure 5. A)** System-specific mean absolute error (kJ/mol) of all functionals (blue) vs. best
3 functionals (red) divided into system. **B)** Mean signed error (kJ/mol) for each system used in the
4 test set, averaged over all functionals (blue) and those functionals that have MSE < 15 kJ/mol.
5
6
7
8
9
10
11

12 Another important point is the magnitude of systematic errors divided into system type,
13 because they can reflect missing realism in the model of that system; such as the plot is shown in
14 **Figure 5B**. The blue bars represent the MSE for all studied functionals for a given system,
15 whereas the orange bars represent the MSE of each system using only accurate functionals (with
16 MSE < 15 kJ/mol). Systems **2** and **3** have errors up to 15-20 kJ/mol that could reflect either a
17 weakness in the realism of the chemical model used to compute ΔH_{SCO} (e.g. neglect of counter
18 ions) or a general weakness in the DFT treatment (dispersion, steric or electronic strain or
19 similar). These two are both Fe(II) systems, but the errors are otherwise well distributed for both
20 Fe(II) and Fe(III) systems.
21
22
23
24
25
26
27
28
29
30
31
32
33

34 **Precision vs. Accuracy of DFT Applied to Iron Spin States.** The final thing we wanted
35 to investigate was whether the errors that we observed above are distributed in a consistent way
36 or whether they spread, i.e. if there is a predictable error associated with each functional. We
37 decided to quantify the precision of the functional by the standard deviation of the errors: The
38 larger this value is, the less predictive are the errors obtainable with a functional. We also wanted
39 to see if the precision correlates with the accuracy of the density functionals. Precision is
40 arguably a very important (but overlooked) aspect of performance, because an apparently
41 accurate functional could obtain its good average performance with a large spread in
42 performance for the individual system, *which would make the performance of the functional*
43 *unpredictable*. To study this, we suggest to use scatter plots of the type shown in **Figure 6A** (for
44
45
46
47
48
49
50
51
52
53
54
55
56
57
58
59
60

1
2 Fe(II) systems) and **Figure 6B** (for Fe(III) systems), where precision (as defined above) is
3
4 plotted vs. accuracy.
5
6



21
22 **Figure 6.** Precision vs. accuracy of density functionals for spin crossover: A) Fe(II) systems; B)
23 Fe(III) systems.
24
25

26
27
28
29
30 It is evident from the analysis in **Figure 6** that the Fe(II) and Fe(III) systems behave
31 distinctly, which confirms the importance of studying Fe(II)-Fe(III) balance. Also important is
32 the observation that precision is weakly correlated to accuracy, with decent correlation for Fe(III)
33 systems ($R^2 \sim 0.36$) but not for Fe(II) systems ($R^2 \sim 0.08$). Thus, some functionals that are quite
34 accurate as measured by the standard approach of computed MAE, turn out to be very imprecise.
35 The analysis of the accuracy-precision relationship is quite interesting and we are not aware that
36 it has been discussed in detail. Thus, we see for example that the accurate hybrid functionals with
37 15-17% HF exchange are consistently also the most precise functionals (with precisions of the
38 order of 5 kJ/mol), which is a very important conclusion because it is required for predictive use
39 of DFT. Importantly, there are functionals that have good accuracy but poor precision, such as
40 rPBE and B2PLYP. In addition there are very inaccurate functionals with high precision, the
41 most prominent examples being PBE-H3C, M06, and M06-2X, as well as functionals that are
42
43
44
45
46
47
48
49
50
51
52
53
54
55
56
57

1 both inaccurate and imprecise, such as M06L and KT1. Thus, all four types of functional
2 behavior (accurate/precise, accurate/imprecise, inaccurate/precise, inaccurate/imprecise) are
3
4 observed, which should be of some interest in the future consideration of these and other
5
6 functionals.
7
8
9

10 11 12 13 14 **Conclusions.**

15
16
17 This work has studied the description of the HS and LS states of well-known SCO systems
18 containing Fe(II) and Fe(III), and compared DFT-derived enthalpies against available experimental
19 data. It is shown that Fe(II) and Fe(III) systems are typically not treated in a balanced way by
20 DFT, i.e. a functional is significantly more accurate for one of these oxidation states than the
21 other, and thus cannot distinguish fairly between SCO candidates of both types. To understand
22 this “Fe(II)-Fe(III) bias”, we deployed a range of functionals, including customized versions of
23 GGA, hybrid, meta-GGA, meta-hybrid, double-hybrid, and long-range-corrected hybrid
24 functionals, to search for the best-possible description of this problem by modern DFT.
25
26
27
28
29
30
31
32
33
34
35
36

37 We find that most functionals, regardless of the nature of the correlation and exchange
38 functionals, become more accurate in their hybrid forms once the included HF exchange is
39 adjusted to 10-17%. The most accurate functionals measured as mean absolute errors < 10
40 kJ/mol are CAMB3LYP-17, B3LYP*, and B97-15 with 15-17% HF exchange, closely followed
41 by CAMB3LYP and CAMB3LYP-15, OPBE, rPBE-10, and B3P86-15. The highest possible
42 accuracy however comes with a clear Fe(II)-Fe(III) bias of up to ~20 kJ/mol. Hybrid functionals,
43 regardless of design (including B2PLYP double hybrids and meta hybrids) tend to favor HS too
44 much in Fe(II) vs. Fe(III), which is important in many studies of iron spin states where the
45 oxidation state of iron can vary, e.g. rational SCO design and studies of catalytic processes
46
47
48
49
50
51
52
53
54
55
56
57
58
59
60

1
2 involving iron. The bias tends to grow with the amount of HF exchange such that for B3LYP the
3 imbalance amounts to 39 kJ/mol and for TPSSH it is 21 kJ/mol. In contrast, GGA functionals
4 display a small Fe(II)-Fe(III) bias, but are generally inaccurate, except those using the O
5 exchange functional.
6
7
8
9

10
11
12 Upon detailed analysis of both accuracy, precision, and balanced treatment (system bias),
13 we find that the only functional with a combined bias < 5 kJ/mol and a decent MAE (15 kJ/mol)
14 is the customized PBE0-12 functional. Alternatively, one has to either sacrifice Fe(II)-Fe(III)
15 balance, use the best functionals for each oxidation state separately, or use a combination of
16 functionals to directly estimate the SCO energetics.
17
18
19
20
21
22
23

24 The precision (measured as the standard deviation of errors) is not generally strongly
25 correlated to accuracy (measured by the MAE) of a functional, which is problematic because it
26 makes DFT less predictive *regardless* of a low MAE. Importantly, all four types of functional
27 behavior (accurate/precise, accurate/imprecise, inaccurate/precise, inaccurate/imprecise) are
28 observed. Thus, our work illustrates the importance of *balanced* accuracy for systems likely to
29 occur in context during a process.
30
31
32
33
34
35
36
37
38

39 In many catalytic processes both the oxidation and the spin state of iron changes, the
40 most prominent example perhaps being heme chemistry. In such cases, the Fe(II)-Fe(III) bias
41 will cause an error up to 20 kJ/mol in the potential energy surfaces that *will not* be removed by
42 error cancellation. In this context, our work also presents the first benchmark of the rPBE
43 functional widely used in iron-based catalysis such as the Haber–Bosch process, where iron spin
44 and oxidation states may change. The rPBE functional provides a much better LS-HS balance
45 and is less biased toward LS than PBE, but is also very imprecise, which may constitute a
46 problem for predictive use of DFT in catalysis when spin and oxidation states change. The
47
48
49
50
51
52
53
54
55
56
57
58
59
60

1
2 analysis provided in this work should be of value in establishing more accurate models of
3
4 catalytic processes of this type that take into account the Fe(II)-Fe(III) bias, accuracy, and
5
6 precision in a combined way.
7
8
9
10
11

12 **Supporting Information available.** The Supporting information file contains details of the
13
14 computations performed in this work, including all electronic energies, derived high-spin low-
15
16 spin gaps, experimental data used, ZPE and relativistic corrections, and coordinates of all studied
17
18 systems in both HS and LS states. This material is available free of charge via the Internet at
19
20 <http://pubs.acs.org>.
21
22
23
24

25 **Acknowledgements.** The authors acknowledge DTU for use of the Steno computer Cluster,
26
27 which was originally established by a grant to KPK from The Danish Council for Independent
28
29 Research | Natural sciences (FNU), grant 272-08-0041 (National Young Elite Researcher
30
31 Prize).
32
33
34
35
36
37

38 **References.**

- 39
40
41
42 (1) Halcrow, M. A. *Spin-Crossover Materials: Properties and Applications*; John Wiley &
43
44 Sons, 2013.
45
46
47 (2) Gütlich, P.; Goodwin, H. A. Spin Crossover—an Overall Perspective. In *Spin Crossover*
48
49 *in Transition Metal Compounds I*; Springer, 2004; pp 1–47.
50
51
52
53 (3) Létard, J.-F.; Guionneau, P.; Goux-Capes, L. Towards Spin Crossover Applications. *Spin*
54
55 *Crossover Transit. Met. Compd. III* **2004**, 1–19.
56
57
58
59
60

- 1
2 (4) Brooker, S. Spin Crossover with Thermal Hysteresis: Practicalities and Lessons Learnt.
3
4 *Chem. Soc. Rev.* **2015**, *44* (10), 2880–2892.
5
6
7 (5) Gütlich, P.; Garcia, Y.; Goodwin, H. A. Spin Crossover Phenomena in Fe(II) Complexes.
8
9 *Chem. Soc. Rev.* **2000**, *29* (6), 419–427.
10
11
12 (6) Harding, D. J.; Harding, P.; Phonsri, W. Spin Crossover in Iron (III) Complexes. *Coord.*
13
14 *Chem. Rev.* **2016**, *313*, 38–61.
15
16
17 (7) Guionneau, P. Crystallography and Spin-Crossover. A View of Breathing Materials. *Dalt.*
18
19 *Trans.* **2014**, *43* (2), 382–393.
20
21
22 (8) Ruiz, E. Charge Transport Properties of Spin Crossover Systems. *Phys. Chem. Chem.*
23
24 *Phys.* **2014**, *16* (1), 14–22.
25
26
27 (9) Kepp, K. P. Consistent Descriptions of Metal–ligand Bonds and Spin-Crossover in
28
29 Inorganic Chemistry. *Coord. Chem. Rev.* **2013**, *257* (1), 196–209.
30
31
32 (10) Toftlund, H. Spin Equilibrium in Solutions. *Monatshefte für Chemie/Chemical Mon.* **2001**,
33
34 *132* (11), 1269–1277.
35
36
37 (11) Kershaw Cook, L. J.; Kulmaczewski, R.; Mohammed, R.; Dudley, S.; Barrett, S. A.;
38
39 Little, M. A.; Deeth, R. J.; Halcrow, M. A. A Unified Treatment of the Relationship
40
41 between Ligand Substituents and Spin State in a Family of Iron (II) Complexes. *Angew.*
42
43 *Chemie Int. Ed.* **2016**, *55* (13), 4327–4331.
44
45
46 (12) Kepp, K. P. Theoretical Study of Spin Crossover in 30 Iron Complexes. *Inorg. Chem.*
47
48 **2016**, *55* (6), 2717–2727.
49
50
51 (13) Paulsen, H.; Duelund, L.; Winkler, H.; Toftlund, H.; Trautwein, A. X. Free Energy of
52
53
54
55
56
57
58
59
60

- 1 Spin-Crossover Complexes Calculated with Density Functional Methods. *Inorg. Chem.*
2
3
4 **2001**, *40* (9), 2201–2203.
5
6
- 7 (14) Paulsen, H.; Schünemann, V.; Wolny, J. A. Progress in Electronic Structure Calculations
8 on Spin-Crossover Complexes. *Eur. J. Inorg. Chem.* **2013**, *2013* (5–6), 628–641.
9
10
- 11 (15) Scheidt, W. R.; Reed, C. A. Spin-State/stereochemical Relationships in Iron Porphyrins:
12 Implications for the Hemoproteins. *Chem. Rev.* **1981**, *81* (6), 543–555.
13
14
- 15 (16) Jensen, K. P.; Ryde, U. How O₂ Binds to Heme: Reasons for Rapid Binding and Spin
16 Inversion. *J. Biol. Chem.* **2004**, *279* (15), 14561–14569.
17
18
- 19 (17) Harvey, J. N.; Poli, R.; Smith, K. M. Understanding the Reactivity of Transition Metal
20 Complexes Involving Multiple Spin States. *Coord. Chem. Rev.* **2003**, *238*, 347–361.
21
22
- 23 (18) Sorai, M.; Nakano, M.; Miyazaki, Y. Calorimetric Investigation of Phase Transitions
24 Occurring in Molecule-Based Magnets. *Chem. Rev.* **2006**, *106* (3), 976–1031.
25
26
- 27 (19) Molnár, G.; Salmon, L.; Nicolazzi, W.; Terki, F.; Bousseksou, A. Emerging Properties
28 and Applications of Spin Crossover Nanomaterials. *J. Mater. Chem. C* **2014**, *2* (8), 1360–
29 1366.
30
31
- 32 (20) Liu, T.; Zheng, H.; Kang, S.; Shiota, Y.; Hayami, S.; Mito, M.; Sato, O.; Yoshizawa, K.;
33 Kanegawa, S.; Duan, C. A Light-Induced Spin Crossover Actuated Single-Chain Magnet.
34 *Nat. Commun.* **2013**, *4*, 2826.
35
36
- 37 (21) Cirera, J.; Ruiz, E. Theoretical Modeling of Two-Step Spin-Crossover Transitions in Fe II
38 Dinuclear Systems. *J. Mater. Chem. C* **2015**, *3* (30), 7954–7961.
39
40
- 41 (22) Jørgensen, C. K. Electron Transfer Spectra of Hexahalide Complexes. *Mol. Phys.* **1959**, *2*
42
43
44
45
46
47
48
49
50
51
52
53
54
55
56
57
58
59
60

- 1
2 (3), 309–332.
3
4
5 (23) Tsuchida, R. Absorption Spectra of Co-Ordination Compounds. I. *Bull. Chem. Soc. Jpn.*
6
7 **1938**, *13* (5), 388–400.
8
9
10 (24) Shimura, Y.; Tsuchida, R. Absorption Spectra of Co(III) Complexes. II. Redetermination
11
12 of the Spectrochemical Series. *Bull. Chem. Soc. Jpn.* **1956**, *29* (3), 311–316.
13
14
15 (25) Fajans, K. Struktur und Deformation Der Elektronenhüllen in Ihrer Bedeutung für Die
16
17 Chemischen und Optischen Eigenschaften Anorganischer Verbindungen.
18
19 *Naturwissenschaften* **1923**, *11* (10), 165–172.
20
21
22
23 (26) Mortensen, S. R.; Kepp, K. P. Spin Propensities of Octahedral Complexes From Density
24
25 Functional Theory. *J. Phys. Chem. A* **2015**, *119*, 4041–4050.
26
27
28
29 (27) Kepp, K. P. Heme: From Quantum Spin Crossover to Oxygen Manager of Life. *Coord.*
30
31 *Chem. Rev.* **2017**, *344*, 363–374.
32
33
34 (28) Kohn, W.; Becke, A. D.; Parr, R. G. Density Functional Theory of Electronic Structure. *J.*
35
36 *Phys. Chem.* **1996**, *100* (31), 12974–12980.
37
38
39
40 (29) Becke, A. D. Perspective: Fifty Years of Density-Functional Theory in Chemical Physics.
41
42 *J. Chem. Phys.* **2014**, *140* (18), 18A301.
43
44
45 (30) Salomon, O.; Reiher, M.; Hess, B. A. Assertion and Validation of the Performance of the
46
47 B3LYP- ω Functional for the First Transition Metal Row and the G2 Test Set. *J. Chem.*
48
49 *Phys.* **2002**, *117* (10), 4729–4737.
50
51
52
53 (31) Reiher, M. Theoretical Study of the Fe (Phen)₂(NCS)₂ Spin-Crossover Complex with
54
55 Reparametrized Density Functionals. *Inorg. Chem.* **2002**, *41* (25), 6928–6935.
56
57
58
59
60

- 1
2 (32) Swart, M. Accurate Spin-State Energies for Iron Complexes. *J. Chem. Theory Comput.*
3
4 **2008**, *4* (12), 2057–2066.
5
6
7 (33) Jensen, K. P.; Cirera, J. Accurate Computed Enthalpies of Spin Crossover in Iron and
8
9
10 Cobalt Complexes. *J. Phys. Chem. A* **2009**, *113*, 10033–10039.
11
12
13 (34) Isley III, W. C.; Zarra, S.; Carlson, R. K.; Bilbeisi, R. A.; Ronson, T. K.; Nitschke, J. R.;
14
15 Gagliardi, L.; Cramer, C. J. Predicting Paramagnetic ¹H NMR Chemical Shifts and State-
16
17 Energy Separations in Spin-Crossover Host–guest Systems. *Phys. Chem. Chem. Phys.*
18
19 **2014**, *16* (22), 10620–10628.
20
21
22 (35) Ioannidis, E. I.; Kulik, H. J. Towards Quantifying the Role of Exact Exchange in
23
24
25 Predictions of Transition Metal Complex Properties. *J. Chem. Phys.* **2015**, *143* (3), 34104.
26
27
28 (36) Coskun, D.; Jerome, S. V.; Friesner, R. A. Evaluation of the Performance of the B3LYP,
29
30 PBE0, and M06 DFT Functionals, and DBLOC-Corrected Versions, in the Calculation of
31
32 Redox Potentials and Spin Splittings for Transition Metal Containing Systems. *J. Chem.*
33
34 *Theory Comput.* **2016**, *12* (3), 1121–1128.
35
36
37 (37) Conradie, J.; Ghosh, A. DFT Calculations on the Spin-Crossover Complex Fe(salen)(NO):
38
39
40 A Quest for the Best Functional. *J. Phys. Chem. B* **2007**, *111* (44), 12621–12624.
41
42
43 (38) Verma, P.; Varga, Z.; Klein, J. E. M. N.; Cramer, C. J.; Que, L.; Truhlar, D. G.
44
45
46 Assessment of Electronic Structure Methods for the Determination of the Ground Spin
47
48 States of Fe(II), Fe(III) and Fe(IV) Complexes. *Phys. Chem. Chem. Phys.* **2017**, *19* (20),
49
50 13049–13069.
51
52
53 (39) Swart, M.; Groenhof, A. R.; Ehlers, A. W.; Lammertsma, K. Validation of
54
55
56 Exchange–Correlation Functionals for Spin States of Iron Complexes. *J. Phys. Chem. A*
57
58

- 1
2 **2004**, *108* (25), 5479–5483.
3
4
- 5 (40) Cirera, J.; Ruiz, E. Theoretical Modeling of the Ligand-Tuning Effect over the Transition
6 Temperature in Four-Coordinated FeII Molecules. *Inorg. Chem.* **2016**, *55* (4), 1657–1663.
7
8
9
- 10 (41) Houghton, B. J.; Deeth, R. J. Spin-State Energetics of Fe^{II} Complexes—The Continuing
11 Voyage Through the Density Functional Minefield. *Eur. J. Inorg. Chem.* **2014**, *2014* (27),
12 4573–4580.
13
14
15
16
17
- 18 (42) Reiher, M.; Salomon, O.; Hess, B. A. Reparameterization of Hybrid Functionals Based on
19 Energy Differences of States of Different Multiplicity. *Theor. Chem. Acc.* **2001**, *107* (1),
20 48–55.
21
22
23
24
25
- 26 (43) Gani, T. Z. H.; Kulik, H. J. Unifying Exchange Sensitivity in Transition-Metal Spin-State
27 Ordering and Catalysis through Bond Valence Metrics. *J. Chem. Theory Comput.* **2017**,
28 *13* (11), 5443–5457.
29
30
31
32
33
- 34 (44) Ioannidis, E. I.; Kulik, H. J. Ligand-Field-Dependent Behavior of Meta-GGA Exchange in
35 Transition-Metal Complex Spin-State Ordering. *J. Phys. Chem. A* **2017**, *121* (4), 874–884.
36
37
38
- 39 (45) Peverati, R.; Truhlar, D. G. Quest for a Universal Density Functional: The Accuracy of
40 Density Functionals across a Broad Spectrum of Databases in Chemistry and Physics.
41 *Philos. Trans. R. Soc. London A Math. Phys. Eng. Sci.* **2014**, *372* (2011), 20120476.
42
43
44
45
46
- 47 (46) Medvedev, M. G.; Bushmarinov, I. S.; Sun, J.; Perdew, J. P.; Lyssenko, K. A. Density
48 Functional Theory Is Straying from the Path toward the Exact Functional. *Science* **2017**,
49 *355* (6320), 49–52.
50
51
52
53
- 54 (47) Kepp, K. P. Comment on “Density Functional Theory Is Straying from the Path toward
55
56
57

- 1
2 the Exact Functional.” *Science* **2017**, *356* (6337), 496–497.
3
4
5 (48) Hughes, T. F.; Friesner, R. A. Correcting Systematic Errors in DFT Spin-Splitting
6
7 Energetics for Transition Metal Complexes. *J. Chem. Theory Comput.* **2011**, *7* (1), 19–32.
8
9
10 (49) Ahlrichs, R.; Bär, M.; Häser, M.; Horn, H.; Kölmel, C. Electronic Structure Calculations
11
12 on Workstation Computers: The Program System Turbomole. *Chem. Phys. Lett.* **1989**, *162*
13
14 (3), 165–169.
15
16
17 (50) Eichkorn, K.; Treutler, O.; Öhm, H.; Häser, M.; Ahlrichs, R. Auxiliary Basis Sets to
18
19 Approximate Coulomb Potentials. *Chem. Phys. Lett.* **1995**, *240* (4), 283–290.
20
21
22 (51) Weigend, F.; Häser, M. RI-MP2: First Derivatives and Global Consistency. *Theor. Chem.*
23
24 *Acc.* **1997**, *97* (1), 331–340.
25
26
27 (52) Beattie, J. K.; Binstead, R. A.; West, R. J. Intersystem Crossing Observed by Ultrasonic
28
29 Relaxation of the Singlet-Quintet Spin Equilibrium of Iron (II) Complexes in Solution. *J.*
30
31 *Am. Chem. Soc.* **1978**, *100* (10), 3044–3050.
32
33
34 (53) Turner, J. W.; Schultz, F. A. Intramolecular and Environmental Contributions to Electrode
35
36 Half-Reaction Entropies of $M(\text{tacn})_2^{3+/2+}$ (M= Fe, Co, Ni, Ru; Tacn= 1, 4, 7-
37
38 Triazacyclononane) Redox Couples. *Inorg. Chem.* **1999**, *38* (2), 358–364.
39
40
41 (54) Jesson, J. P.; Trofimenko, S.; Eaton, D. R. Spin Equilibria in Octahedral Iron(II)Poly((1-
42
43 Pyrazolyl)-Borates. *J. Am. Chem. Soc.* **1967**, *89* (13), 3158–3164.
44
45
46 (55) Sorai, M. Calorimetric Investigations of Phase Transitions Occurring in Molecule-Based
47
48 Materials in Which Electrons Are Directly Involved. *Bull. Chem. Soc. Jpn.* **2001**, *74* (12),
49
50 2223–2253.
51
52
53
54
55
56
57
58
59
60

- 1
2 (56) Dose, E. V.; Murphy, K. M. M.; Wilson, L. J. Synthesis and Spin-State Studies in Solution
3
4 Of γ -Substituted Tris (β -Diketonato) Iron (III) Complexes and Their Spin-Equilibrium. β -
5
6 Ketoimine Analogues Derived from Triethylenetetramine. *Inorg. Chem.* **1976**, *15* (11),
7
8 2622–2630.
9
10
11
12 (57) Koikawa, M.; B. Jensen, K.; Matsushima, H.; Tokii, T.; Toftlund, H. Syntheses and
13
14 Crystal Structures of Divalent Complexes with a New Hexadentate Ligand Derived from
15
16 1,4,7-Triazacyclononane. *J. Chem. Soc. Dalton Trans.* **1998**, No. 7, 1085–1086.
17
18
19
20 (58) Sorai, M.; Burriel, R.; Westrum, E. F.; Hendrickson, D. N. Mechanochemical Effect in the
21
22 Iron (III) Spin Crossover Complex [Fe(3-MeO-salenEt)₂]PF₆ as Studied by Heat Capacity
23
24 Calorimetry. *J. Phys. Chem. B* **2008**, *112* (14), 4344–4350.
25
26
27
28 (59) Becke, A. D. Density-Functional Exchange-Energy Approximation with Correct
29
30 Asymptotic Behavior. *Phys. Rev. A* **1988**, *38* (6), 3098–3100.
31
32
33 (60) Perdew, J. P. Density-Functional Approximation for the Correlation Energy of the
34
35 Inhomogeneous Electron Gas. *Phys. Rev. B* **1986**, *33* (12), 8822–8824.
36
37
38
39 (61) Weigend, F.; Ahlrichs, R. Balanced Basis Sets of Split Valence, Triple Zeta Valence and
40
41 Quadruple Zeta Valence Quality for H to Rn: Design and Assessment of Accuracy. *Phys.*
42
43 *Chem. Chem. Phys.* **2005**, *7* (18), 3297–3305.
44
45
46 (62) Schäfer, A.; Klamt, A.; Sattel, D.; Lohrenz, J. C. W.; Eckert, F. COSMO Implementation
47
48 in TURBOMOLE: Extension of an Efficient Quantum Chemical Code towards Liquid
49
50 Systems. *Phys. Chem. Chem. Phys.* **2000**, *2* (10), 2187–2193.
51
52
53
54 (63) Klamt, A.; Schüürmann, G. COSMO: A New Approach to Dielectric Screening in
55
56 Solvents with Explicit Expressions for the Screening Energy and Its Gradient. *J. Chem.*
57
58

- 1
2 *Soc. Perkin Trans. 2* **1993**, No. 5, 799–805.
3
4
5 (64) Jensen, K. P.; Roos, B. O.; Ryde, U. Performance of Density Functionals for First Row
6
7 Transition Metal Systems. *J. Chem. Phys.* **2007**, *126*, 14103.
8
9
10 (65) Grimme, S. Semiempirical Hybrid Density Functional with Perturbative Second-Order
11
12 Correlation. *J. Chem. Phys.* **2006**, *124* (3), 34108.
13
14
15 (66) Tao, J.; Perdew, J. P.; Staroverov, V. N.; Scuseria, G. E. Climbing the Density Functional
16
17 Ladder: Nonempirical Meta Generalized Gradient Approximation Designed for Molecules
18
19 and Solids. *Phys. Rev. Lett.* **2003**, *91* (14), 146401.
20
21
22
23 (67) Perdew, J. P.; Tao, J.; Staroverov, V. N.; Scuseria, G. E. Meta-Generalized Gradient
24
25 Approximation: Explanation of a Realistic Nonempirical Density Functional. *J. Chem.*
26
27 *Phys.* **2004**, *120* (15).
28
29
30
31 (68) Becke, A. D. Density functional Thermochemistry. III. The Role of Exact Exchange. *J.*
32
33 *Chem. Phys.* **1993**, *98* (7), 5648–5652.
34
35
36
37 (69) Lee, C.; Yang, W.; Parr, R. G. Development of the Colle-Salvetti Correlation-Energy
38
39 Formula into a Functional of the Electron Density. *Phys. Rev. B* **1988**, *37* (2), 785-789.
40
41
42 (70) Stephens, P. J.; Devlin, F. J.; Chabalowski, C. F.; Frisch, M. J. Ab Initio Calculation of
43
44 Vibrational Absorption and Circular Dichroism Spectra Using Density Functional Force
45
46 Fields. *J. Phys. Chem.* **1994**, *98* (45), 11623–11627.
47
48
49
50 (71) Grimme, S.; Antony, J.; Ehrlich, S.; Krieg, H. A Consistent and Accurate Ab Initio
51
52 Parametrization of Density Functional Dispersion Correction (DFT-D) for the 94
53
54 Elements H-Pu. *J. Chem. Phys.* **2010**, *132* (15), 154104.
55
56
57
58
59
60

- 1
2 (72) Yanai, T.; Tew, D. P.; Handy, N. C. A New Hybrid Exchange–correlation Functional
3 Using the Coulomb-Attenuating Method (CAM-B3LYP). *Chem. Phys. Lett.* **2004**, *393* (1),
4 51–57.
5
6
7
8
9 (73) Zhao, Y.; Truhlar, D. G. The M06 Suite of Density Functionals for Main Group
10 Thermochemistry, Thermochemical Kinetics, Noncovalent Interactions, Excited States,
11 and Transition Elements: Two New Functionals and Systematic Testing of Four M06-
12 Class Functionals and 12 Other Function. *Theor. Chem. Acc.* **2008**, *120* (1–3), 215–241.
13
14
15
16
17
18 (74) Zhao, Y.; Truhlar, D. G. A New Local Density Functional for Main-Group
19 Thermochemistry, Transition Metal Bonding, Thermochemical Kinetics, and Noncovalent
20 Interactions. *J. Chem. Phys.* **2006**, *125* (19), 194101.
21
22
23
24
25
26 (75) Keal, T. W.; Tozer, D. J. The Exchange-Correlation Potential in Kohn–Sham Nuclear
27 Magnetic Resonance Shielding Calculations. *J. Chem. Phys.* **2003**, *119* (6), 3015–3024.
28
29
30
31
32 (76) Perdew, J. P.; Wang, Y. Accurate and Simple Analytic Representation of the Electron-Gas
33 Correlation Energy. *Phys. Rev. B* **1992**, *45* (23), 13244.
34
35
36
37 (77) Perdew, J. P.; Burke, K.; Ernzerhof, M. Generalized Gradient Approximation Made
38 Simple. *Phys. Rev. Lett.* **1996**, *77* (18), 3865.
39
40
41
42 (78) Zhang, Y.; Yang, W. Comment on “Generalized Gradient Approximation Made Simple.”
43 *Phys. Rev. Lett.* **1998**, *80* (4), 890.
44
45
46
47 (79) Hammer, B.; Hansen, L. B.; Nørskov, J. K. Improved Adsorption Energetics within
48 Density-Functional Theory Using Revised Perdew-Burke-Ernzerhof Functionals. *Phys.*
49 *Rev. B* **1999**, *59* (11), 7413-7421.
50
51
52
53
54
55
56
57
58
59
60

- 1
2 (80) Handy, N. C.; Cohen, A. J. Left-Right Correlation Energy. *Mol. Phys.* **2001**, *99* (5), 403–
3 412.
4
5
6
7 (81) Vosko, S. H.; Wilk, L.; Nusair, M. Accurate Spin-Dependent Electron Liquid Correlation
8 Energies for Local Spin Density Calculations: A Critical Analysis. *Can. J. Phys.* **1980**, *58*
9 (8), 1200–1211.
10
11
12
13
14 (82) Grimme, S.; Brandenburg, J. G.; Bannwarth, C.; Hansen, A. Consistent Structures and
15 Interactions by Density Functional Theory with Small Atomic Orbital Basis Sets. *J. Chem.*
16 *Phys.* **2015**, *143* (5), 54107.
17
18
19
20
21
22 (83) Ekström, U.; Visscher, L.; Bast, R.; Thorvaldsen, A. J.; Ruud, K. Arbitrary-Order Density
23 Functional Response Theory from Automatic Differentiation. *J. Chem. Theory Comput.*
24 **2010**, *6* (7), 1971–1980.
25
26
27
28
29
30 (84) Kepp, K. P. The Ground States of iron(III) Porphines: Role of Entropy–enthalpy
31 Compensation, Fermi Correlation, Dispersion, and Zero-Point Energies. *J. Inorg.*
32 *Biochem.* **2011**, *105* (10), 1286–1292.
33
34
35
36
37
38 (85) Swart, M.; Solà, M.; Bickelhaupt, F. M. A New All-Round Density Functional Based on
39 Spin States and SN2 Barriers. *J. Chem. Phys.* **2009**, *131* (9), 94103.
40
41
42
43
44
45
46
47
48
49
50
51
52
53
54
55
56
57
58
59
60

TOC Graphic

