

## A NEW OCCURRENCE OF MANGANILVAITE $\text{CaFe}^{2+}\text{Fe}^{3+}(\text{Mn}^{2+}, \text{Fe}^{2+})[\text{Si}_2\text{O}_7]\text{O}(\text{OH})$ AT DOGNECEA, SOUTHWESTERN BANAT, ROMANIA: CHEMICAL COMPOSITION, CRYSTAL STRUCTURE AND CATION ORDERING

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Samples of ilvaite systematically exceeding 0.5 *apfu* Mn (in formula normalized on the basis of 6 cations) and suggesting the presence of the recently discovered mineral manganilvaite (ZOTOV *et al.*, 2005; BONEV *et al.*, 2005) were identified in the skarn deposit at Dognecea, where they occur in association with Mn-hedenbergite and magnetite. A number of 14 EDS chemical analyses of such ilvaites resulted in the following ranges of variation: (in weight percents; numbers in parentheses read as average values and standard deviations, respectively): **Si**: 13.39–14.09 (13.92; 0.15), **Al**: 0.00–0.27 (0.09; 0.11), **Fe**: 32.88–34.53 (34.62; 0.64), **Mn**: 7.30–9.17 (7.85; 0.54), **Mg**: 0.00–0.26 (0.07; 0.10), **Ca**: 9.86–10.64 (10.26; 0.22). Atomic proportions of  $\text{Fe}^{2+}$  and  $\text{Fe}^{3+}$  for a chemical formula unit normalized for 6 cations, were calculated as  $\text{Fe}^{2+} = 2 - \text{Mn}^{2+}$  and  $\text{Fe}^{3+} = \text{Fe}_{\text{total}} - \text{Fe}^{2+}$ . The variation domains of  $\text{Fe}^{2+}$ ,  $\text{Fe}^{3+}$  and  $\text{Mn}^{2+}$  atomic proportions were the following: 1.34–1.48 (average 1.44; standard deviation 0.04), 0.98–1.04 (1.01; 0.021) and 0.52–0.66 (0.56; 0.04), respectively. Other varieties of Mn-poorer ilvaites were identified, too.

Manganilvaite from Dognecea is a  $P2_1/a$  polymorph ( $a = 13.014$  Å,  $b = 8.846$  Å,  $c = 5.848$  Å and  $\beta = 90.34^\circ$ ) with significant ordering of  $\text{Fe}^{2+}$  and  $\text{Fe}^{3+}$  among *Me11* and *Me12* structural sites. The single crystal structure determination of manganilvaite – based on 1509 unique reflections, was refined down to a final  $R = 5.41\%$  (with single generic metals – Fe or Mn – in *Me11*, *Me12* and *Me2* positions). The crystal structure data allowed the calculation of interatomic and per-polyhedra average distances for all relevant cation-oxygen pairs and the evaluation of  $\text{Fe}^{2+}$  and  $\text{Fe}^{3+}$  occupancies in *Me11* and *Me12* structural sites. Statistical  $\text{Fe}^{3+}$  occupancies were calculated on the basis of two classic atomic radii models, *i.e.* GHOSE (1966) – (G) and SHANNON (1976) – (S). Thus, *Me11* structural site is 53% (G) or 84% (S) occupied by  $\text{Fe}^{2+}$  (normalized values for *Me11* + *Me12* = 100%).  $\text{Fe}^{3+}$  occupancy in *Me11* and  $\text{Fe}^{2+}$  occupancy in *Me12* site is complementary to these values. The corresponding ordering parameter (TAKEUCHI *et al.*, 1983) was 0.34 (G) or 0.47 (S). Same calculations concerned the *Me2* structural site and allowed indirect evaluation of  $\text{Mn}^{2+}$  occupancy *vs.* complementary  $\text{Fe}^{2+}$ . Calculation of Mn occupancy in *Me2* based on a simple linear variation of *Me2*–O Shannon distances be-

tween 2.18 Å ( $\text{Fe}^{2+}$ –O) and 2.23 Å (Mn–O), resulted in 42%  $\text{Mn}^{2+}$  whereas the empirical model of Carrozzini (1994) yielded 52%  $\text{Mn}^{2+}$ . Similar discrepancies between the Shannon and Carrozzini models may be described in the case of the type manganilvaite from Ossikovo (BONEV *et al.*, 2005). The refinement of the structure with split *Me2* position (Mn and Fe) resulted in a slightly lower  $R = 5.33\%$ . Attempts to refine the structure with Mn in Ca, *Me11* or *Me12* position have failed, suggesting that Mn is confined to the *Me2* position.

The experimental  $\beta$  value is slightly discrepant with regard to the ones calculated on the basis of correlated ordering parameter and degree of monoclinicity (TAKEUCHI *et al.*, 1983 (T); FINGER & HAZEN, 1987 (FH)):  $90.16^\circ$  (T) and  $90.27^\circ$  (FH).

Local paragenetical relations and various published correlations between the ordering parameter and the temperature of formation, suggest that Mn-ilvaite from Dognecea represents a reaction product of 6 hed + 4 mgt + 3  $\text{H}_2\text{O} \rightarrow 6$  ilvaite +  $\frac{1}{2}$   $\text{O}_2$ , formed at temperatures not exceeding 300 °C, and in conditions of relatively low  $f\text{O}_2$ .

### References

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